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ELECTRICAL PROPERTIES OF INTERFACES-- COMPILATION OF DATA ON THE ELECTRICAL DOUBLE LAYER ON MERCURY ELECTRODES

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ELECTRICAL PROPERTIES OF INTERFACES

Compilation of Data on the Electrical Double Layer on Mercury Electrodes.

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October 1981

Interfacial Electrochemistry

COMPILATION OF DATA ON THE ELECTRICAL DOUBLE LAYER ON MERCURY ELECTRODES

The system consisting of a pure mercury electrode in contact with a solution containing at least one electrolyte is one of the best defined in interfacial electrochemistry. Mercury is readily obtained in a state of high purity and the possibility of renewing the interface allows measurements to be made under conditions where the interface is free from contamination. Consequently, this system is the most studied in interfacial electrochemistry and has yielded the largest quantity of data which can be accepted as reliable.

The present time seems to be appropriate for the collection of data for this system, because it marks a transition in the way this type of information is being acquired. In the past, measurements were largely made manually, point by point, using the capillary electrometer or the a.c. bridge. In the future it seems certain that the more automated methods developed in recent years will be preferred and the data will become available in the form of computer stores rather than printed tables.

When the compilation of these tables was first discussed, it was decided to present data in its primary form, that is as the measured interfacial tension γ or double layer capacity C, rather than in the form of data processed even by well-established thermodynamic methods. This has been adhered to in most tables, although sometimes the capacity has been given with its conjugate potential at integral values of the charge σ calculated by integrating the C-E curves from the potential of zero charge $E_{\sigma=0}$:

$$\sigma = \int_{E_{\sigma=0}}^{E} C\,d\,E \qquad (1)$$

This involves some interpolation of the measured C-E curves but the error introduced is generally negligible in comparison with the experiment error.

The tables have been made as complete as possible, subject to certain criteria of reliability discussed below. This means that they range in time from the earliest accurate measurements (Gouy's electrocapillary curves at the beginning of the century and Grahame's capacity curves in the 1940's) to about 1978. However, it has not been possible to obtain numerical data for all published work and data only available in graphical form have been rejected. In general it has not been possible to select a particular set of results from a large field of data because there is relatively little overlap of complete systems. What overlap there exists is usually in the form of a common system in a general study of two different types of behaviour, e.g. a common base solution in the study of two different series of added organic compounds. In such cases the system common to both is quoted in each set as there may be minor systematic differences, which will cancel in the analyses of each set.

Double layer capacity / potential curves

Introduction. Criteria for selection

This chapter contains double layer data on pure mercury electrodes in contact with electrolyte solutions which in some cases contain added non-electrolyte.

The principal parameter presented is the double layer capacity per unit area of the interface which is measured directly using an alternating current bridge under conditions such that no detectable electrode reaction occurs, i.e. the electrode is perfectly polarized.

The following general rules were used in the selection of data:

1. The mercury used should be highly purified, preferably by a wet method, followed by two or three distillations in a high vacuum still or a still with an air leak.
2. Solvent should be purified by distillation. The solute is preferably purified by recrystallisation, but in some cases analytical grade salts appear to be adequate. When a dropping electrode is used, analytical reagent grade salts may be adequate, but this should be demonstrated either by comparison with more highly purified salts or by the absence of Faradaic current.
3. The results should be obtained under equilibrium conditions. This is best demonstrated by double integration of the capacity/potential curves and comparison with the corresponding electrocapillary curves, but other indications such as invariance with time and frequency are also acceptable. If a genuine frequency effect is established (due to adsorption-desorption processes) the zero frequency extrapolated results are quoted here, unless otherwise stated.
4. The electrode should be spherical and surrounded symmetrically by a spherical or cylindrical counter electrode of negligible impedance. The electrode support should not shield the surface of the electrode and penetration of electrolyte between the mercury and the (glass) support must be eliminated.

5. In the case of dropping mercury electrodes the flow rate of the mercury should be constant. This is ensured by using a sufficiently high column of mercury, so that its pressure is large in comparison with the "back pressure" due to the interfacial tension of the drop-solution interface.

6. A well-defined reference electrode system must be used.

Accuracy

If the measurements are made with a sufficiently accurate bridge and timing device, the results obtainable with a dropping electrode can be internally consistent to about 0.1% but reproducibility between different workers may be as poor as 1% in some parts of the curve. The hanging drop electrode is more susceptible to contamination and to errors in area determinations, but under the best conditions comparable accuracy can be attained.

Presentation

The capacity (C) in microfarads per square centimetre ($\mu F\ cm^{-2}$) is given as a function of the potential (E) of the mercury electrode in volts (V) with respect to the reference electrode.

In some cases values of E and C are given at round values of the charge (σ) in microcoulombs per square centimetre ($\mu C\ cm^{-2}$). These have been obtained by fitting the experimental data three points at a time to a quadratic equation, integrating and interpolating. Errors involved in this procedure are less than 0.1% and hence are negligible. The integration constant is the potential of zero charge (p.z.c.) obtained usually by measuring the potential of an isolated streaming mercury electrode.

Electrocapillary curves

Introduction. Criteria for selection

This chapter contains double layer data for pure mercury electrodes in contact with electrolyte solutions which in some cases contain added non-electrolyte.

The principal parameter presented is the interfacial tension between the liquid metal and the electrolyte measured using a Lippmann electrometer under conditions that no detectable electrode reaction occurs, i.e. the electrode is perfectly polarized.

The following general rules were used in the selection of data:

1. The mercury used should be highly purified, preferably by a wet method followed by two or three distillations in a high vacuum still or

a still with an air leak.

2. Solvent should be purified by distillation. The solute is preferably purified by recrystallisation but in some cases analytical grade salts appear to be adequate.

   The capillary electrometer is usually operated with the mercury/electrolyte meniscus , a mm or so away from the end of a capillary which has a diameter of about 10 µm. Since no convection is possible in the electrolyte within this capillary, transport of material to the interface is by linear diffusion only. This is notably true for the early results of Gouy; for those systems where direct comparison is possible, there is good agreement between his data and more recent ones obtained in highly purified solutions.

3. The results should be obtained under equilibrium conditions. This is best demonstrated by comparison with the surface tension obtained by double integration of the corresponding capacity curves, but the absence of time variation is also an indication that conditions are satisfactory. The comments made under 2 above indicate that the attainment of equilibrium may be very slow, especially when the adsorption of a substance from more dilute solutions is being studied.

4. The mercury column must be at constant temperature and corrections must be applied for the height of the solution above the meniscus in the capillary.

5. The electrolyte must make a contact angle of zero with the glass in the capillary. If it does not, this may be indicated by "sticking" of the meniscus or other time effects. However, it may not be revealed in the measurements themselves without the comparison described in 3 above. The zero contact angle condition may not be necessary if the "maximum bubble pressure method" is used.

6. A well-defined reference electrode system must be used.

7. Drop time methods can give accurate electrocapillary curves, but errors can be large so that capillary electrometer results have been selected here whenever possible. Results from drop time measurements are indicated as such.

Accuracy

Measurements of the height of the column in the electrometer can be made to about 0.02% . Conversion of the corrected height to interfacial tension is done by multiplying by a factor derived from Gouy's absolute measurements of the interfacial tension at the p.z.c. for six electrolytes using the sessile drop method (Ann.Physique (9), 6 (1916), 5). It is unlikely that the resulting values of the interfacial tension are accurate to 0.02% but their self-consistency will be of this order and for most purposes, e.g. deriving surface excess data, the absolute accuracy is unimportant in comparison with the errors involved in processing the data. Hence, the data are quoted to 0.1 $\mathrm{mN\ m^{-1}}$ (= 0.1 $\mathrm{dyn\ cm^{-1}}$). Recent absolute measurements of the mercury-aqueous electrolyte interfacial tension by Vos and Los (J.Colloid Interface Sci. 74 (1980), 360) are certainly more accurate than the old values and are about 0.4 $\mathrm{mN\ m^{-1}}$ higher. This error in absolute value of 0.1% is carried proportionally through all calculations of derived quantities, and since these are rarely obtainable with a precision approaching this, it does not seem worthwhile to correct the interfacial tensions quoted here to the new standard. However, it is clearly desirable that this should be adopted in future work.

Presentation

The S.I. unit for interfacial tension is newtons divided by metres. However, the results collected here are given in $\mathrm{mN\ m^{-1}}$ because this is a unit of convenient size and also because it is numerically equal to the cgs unit $\mathrm{dyn\ cm^{-1}}$ which has been widely used in the past. γ is given as a function of the potential (E) of the mercury electrode in volts (V) with respect to the reference electrode. The latter is frequently a calomel electrode which may be one of three types: saturated (SCE), molar (normal) (NCE) or tenth molar (normal) (0.1 NCE), these terms referring to the concentration of KCl in the electrode.

Conversion of these scales may be achieved using the following values for the emf's of cells at 25 °C.

$$\mathrm{Hg}\ \Big|\ \underset{(\mathrm{sat})}{\mathrm{Hg_2Cl_2}}\ \Big|\ \underset{(\mathrm{sat})}{\mathrm{KCl}}\ \vdots\ \underset{(1.00\ \mathrm{M})}{\mathrm{KCl}}\ \Big|\ \underset{(\mathrm{sat})}{\mathrm{Hg_2Cl_2}}\ \Big|\ \mathrm{Hg} \qquad E = +0.0388\ \mathrm{V}$$

$$\mathrm{Hg}\ \Big|\ \underset{(\mathrm{sat})}{\mathrm{Hg_2Cl_2}}\ \Big|\ \underset{(\mathrm{sat})}{\mathrm{KCl}}\ \vdots\ \underset{(0.1\ \mathrm{M})}{\mathrm{KCl}}\ \Big|\ \underset{(\mathrm{sat})}{\mathrm{Hg_2Cl_2}}\ \Big|\ \mathrm{Hg} \qquad E = +0.0917\ \mathrm{V}$$

Acknowledgements

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Wageningen, The Netherlands

Paris, France.

October, 1981.

PRESENTATION OF TABULAR MATTER

The tables are not arranged in a particular order, because some publications cover wider grounds than others and because sometimes similar information is available from differing sources. As a trend, simple systems come first, the more complicated ones (i.e., the ones in the presence of organic additives) come later.

For easy reference the key on the following pages may be useful. In this key DL stands for double layer and ECC for electrocapillary curve.

Organic substances are listed with their trivial names. For a few lesser known substances the systematic name or a structural formula is given at the bottom of the pertaining table.

Survey

1. *Double layer on mercury* — *Page:*

1) Single electrolytes, monovalent cations — viii
2) Single electrolytes, multivalent cations — ix
3) Electrolyte mixtures, nonaqueous solutions, melts, etc. — x
4) Solutions containing organic substances — xi
5) Substances with ionic surfactant character. — xii

2. *Electrocapillary curves on mercury*

1) Single electrolytes, including some salts of organic acids — xiv
2) Electrolyte mixtures, including organic acids — xvii
3) Solutions containing organic substances — xviii
4) Substances with ionic surfactant character. — xxiii

INDEX

1. DOUBLE LAYER ON MERCURY (DL-Hg)

DL-Hg

1.1. Single electrolytes, monovalent cations. Aqueous solutions*)

$AgNO_3$ 125
CsCl 1-3, 155*-157*, 180*, 181*
$CsClO_4$ 298*, 313*
CsF 33
CsH_2PO_4 187*
CsI 282*, 289*
Formic acid 182*-194*, 365*
HAc 122-124
HCl 1-2, 32, 42, 71-76, 370*
$HClO_4$ 58-60, 106-111, 125, 126, 243*-250*, 266*-270*, 344*-348*, 368*, 369*
$HgClO_4$ 125
HI 69, 70
HNO_3 32
H_2SO_4 32, 51-53, 135, 179*, 252*, 264*, 265*, 359*, 371*, 374*, 375*, 399*- 410*
KAc 13-15
KBr 13-15, 28, 29, 31, 90-95, 179*, 294*
$KBrO_3$ 30
KCl 1-3, 13-15, 31, 34-36, 39, 43, 50, 78-88, 158*, 159*, 166*-168*, 259*-261*, 303*, 304*, 318*-322*, 376*-398*
$KClO_3$ 30
$KClO_4$ 16-18, 114-119, 297*, 313*
KCN 31
KCNS 13-15, 31
K_2CO_3 13-15, 32
KF 16-18, 31, 49, 65-68, 96-105, 147, 166*-170*, 174*-177*, 198*-203*
$KHCO_3$ 16-18, 32
KH_2PO_4 186*
KI 17-23, 31, 169*, 170*, 172*, 196, 197, 198*-200*, 262*, 263*, 305*, 306*
KIO_3 30

---

*) An asterisk means that the electrolyte occurs as a component of a mixture, or as a carrier electrolyte. See also 1.3 of this index for electrolytes in organic solvents either or not in the presence of water.

KNO_3	16-18, 31, 49, 63, 64, 129-134, 148, 153*, 173*, 207*-209*, 252*, 281*, 294*, 307*, 308*
KOH	13-15, 177*
KPF_6	54-57, 271*, 272*, 275*, 276*, 283*-293*, 314*, 317*
K_2SO_4	16-18, 32, 360*
$LiBF_4$	281*, 317*
LiCl	1-3, 77, 154, 155*-157*, 180*, 181*
$LiClO_4$	281*, 294*, 295*, 313*
LiH_2PO_4	183
$LiNO_3$	150-152, 153*
$NaBF_4$	136, 137
NaCl	1-3, 41, 89, 327*-330*
$NaClO_4$	27, 31, 50, 112, 113, 178*, 273*, 274*, 296*, 313*, 315*, 316*, 349*, 352*-354*
NaF	4-12, 40, 195*, 299*-302*, 309*, 310*, 331*-343*
Na-formate	189*-194*, 365
NaH_2PO_4	184*, 185*
$NaHSO_4$	30, 182*
NaN_3	44-48
$NaNO_2$	30
$NaNO_3$	41, 149, 153*
NaOH	37, 38, 122-124, 355*-358*, 366*, 367*
Na_2S	32
NaCNS	24-26
$Na_2S_2O_6$	30
Na_2SO_3	30
Na_2SO_4	61, 62, 188*, 253*-258*, 277-280*, 311*, 361*-364*, 372*, 373*, 411*-423*
NH_3 (liq.)	171*-173*
NH_4Cl	1-3, 160*, 161*, 171*, 282*, 294*
NH_4ClO_4	120, 121, 164*, 165*
NH_4F	162*-165*
NH_4NO_3	160*, 161*-163*, 171*
NH_4OH	120, 121
RbCl	1-3
TlOH	127, 128, 178*

1.2. Multivalent cations

$AlCl_3$	138, 139
$BaCl_2$	138-140, 145, 146

BaI_2 145, 146
$CaCl_2$ 138-140, 142
$Ca(ClO_4)_2$ 141
$Cd(NO_3)_2$ 125
$CoCl_2$ 138, 139
$CuSO_4$ 125
$LaCl_3$ 138-140
$MgCl_2$ 1-3, 143, 145, 146, 158*, 159*
$MgSO_4$ 144
$MnCl_2$ 138-140
$NiCl_2$ 138, 139
$PrCl_3$ 138-140
$SrCl_2$ 138-140

$(Bu)_4NClO_4$ 313*
$(Et)_4NBF_4$ 312*
$(Et)_4NBr$ 312*
$(Et)_4NClO_4$ 282*, 313*
$(Et)_4NI$ 312*
$(Me)_4NClO_4$ 313*
$(Prop)_4NBr$ 281*

1.3. Electrolyte mixtures, nonaqueous solutions, melts, etc.

KF 147, 210-213
KI 171
KNO_3 148, 173
LiCl 154
$LiNO_3$ 150-152
$NaClO_4$ 214-221
NaF 222-242
$NaNO_3$ 149
NH_4Cl 204
NH_4NO_3 171, 205, 206

CsH_2PO_4 + formic acid 187
(Cs + Li)Cl 155-157; (in MeOH) 180, 181
$(H + Na)ClO_4$ 350*, 351*
(H + Na) formate 189-194
H_2SO_4 + KBr 179
K(Cl + F) 166-168
KF + KHF_2 174-176

K(F + I) 169, 170; (in MeOH) 198-200
K(F + OH) 177
KH_2PO_4 + formic acid 186
(K + Li + Na)NO_3 153
(K + Mg)Cl 158-159
LiH_2PO_4 + formic acid 183
$NaClO_4$ + TlOH 178
NaH_2PO_4 + formic acid 184, 185
$NaHSO_4$ + formic acid 182
Na_2SO_4 + formic acid 188
NH_4(Cl + NO_3) 160, 161
NH_4(ClO_4 + F) 164, 165
NH_4(F + NO_3) 162, 163

1.4. Solutions containing organic substances

Note. For acetates (Ac), carbonate (CO_3), formates and methanol (MeOH), see also this index parts 1.1 and 1.3 .

acetamide (dimethyl) 291
acetamide (N-butyl) 283
acetamide (N-ethyl) 289
acetamide (N-methyl) 281, 282, 285, 286
acetonitrile 331, 332
acetylacetone 349
acridine + deriv. 318-320
alanine 339
benzene 259-268
benzoic acid 352-354
butanol 227, 228, 247-250
butanol (tert.-) 253-258
butylamine (tert.-) 359
butyrolactone 290, 317
camphene 311
camphor 277-279
cresol (o,m,p.) 340-343, 355-358
ethanol 222, 223
ether (dibutyl-) 365
ethylene carbonate 284, 312-314
formamide 288, 299-300, 303-308
formamide (dimethyl-) 288, 309, 310
formamide (N-ethyl-) 287

formamide (N-methyl-) 288, 295-298, 301, 302
formamide (N-methylpripionamide-) 288
formamide (N-tert.butyl-) 283
furan (tetrahydro-) 363
furfurylalcohol 361
furfurylamine 362
furfurylamine (tetrahydro-) 364
guanidinium chloride 323-330
hexanol 231, 232
hyamine 370
methanol 196-206, 210-221
methionine 366-369
naphtalene 280
nonylic acid 311
norvaline 337
norleucine 338
octanol 251, 252
α-pinene 311
pentanoic acid 344-348
pentanol 229-230
pentanol, chloro- 233-242
pentanol, iso- 243-246
polyacrylic acid 372-375
propanol 224-226
propionitrile 333, 334
propylene carbonate 284, 315, 316
pyridine 350, 351, 360
quinoline 321
quinoline, iso- 322
succinonitrile 335, 336
sulpholane 271-276
sulphoxide (dimethyl-) 292-294
thiourea 195, 201-203, 207-209
toluene 259-265, 269, 270

1.5. Substances with ionic surfactant character

For the sake of space the following abbreviations are used in this subsection: T = tri, M = methyl, E = ethyl, P = n-propyl; B = n-butyl, A = ammonium, Py = pyridinium, i = iso, Q = quinolinium. C_x stands for n-hydrocarbon chain with x C atoms.

$(C_6TMA)_2SO_4$	412
C_8TMABr	384, 385, 399
$(C_8TMA)_2SO_4$	408, 416, 417
$(C_{10}TMA)_2SO_4$	413
$C_{12}TMABr$	376, 400
$(C_{12}TMA)_2SO_4$	402, 414
$(C_{14}TMA)_2SO_4$	415
$C_{16}TMABr$	377, 395, 396, 401
$(C_{16}TMA)_2SO_4$	410, 411
$(C_8TEA)_2SO_4$	407, 422, 423
$C_{12}TEABr$	378
$C_{16}TEABr$	378, 393, 394
C_8TPABr	380, 381, 405
$(C_8TPA)_2SO_4$	409, 418, 419
$C_{12}TPABr$	386, 387
$C_{16}TPABr$	397, 398
C_8TBABr	382, 383, 403, 404
$(C_8TBA)_2SO_4$	406, 420, 421
$C_{12}TBABr$	388, 389
C_8PyBr	390
$C_{12}PyBr$	391
$C_{10}iQBr$	392

2. ELECTROCAPILLARY CURVES ON MERCURY (ECC-Hg)

2.1. Single electrolytes, including some salts of organic acids. Aqueous solutions*)

$Al_2(SO_4)_3$	92, 93
$As(Me)_4OH$	351*
$BaBr_2$	102
$BaCl_2$	100, 120*-123*
$Ba(H_2PO_2)_2$	89
$Ba(NO_3)_2$	85
$Ba(OH)_2$	84
$BaPt(CN)_6$	98
$Ba(SCNO)_2$	97
$BeSO_4$	93
$CaBr_2$	102
$CaCl_2$	100
$Ca(H_2PO_2)_2$	89
CaI_2	103
$Ca(NO_2)_2$	86
$Ca(NO_3)_2$	86
CsCl	11, 100, 147*
CsH_2PO_4	168*
CsI	12
CsOH	84
HBr	107, 347*, 349*, 352*, 355*
HCl	46-59, 106, 107, 120*-123*, 175*-183*, 234*-239*, 256*, 312*-314*, 318*
$HClO_4$	5-8, 22, 83, 148*, 162*, 163*, 188*, 198*-201*, 223*, 233*, 278*, 279*
HCNS	106
HI	107
HNO_3	104, 131*, 132*, 169*-172*
H_3PO_2	104
H_3PO_4	104, 347*, 349*-351*, 353*-357*
$H_2Pt(CN)_4$	106
H_2SO_4	104, 105, 208*, 346*-360*, 379*-383*, 392*-395*
$Hg_2(NO_3)$	131*, 132*, 169*-172*

---

*) An asterisk means that the electrolyte occurs as a component of a mixture, or as a carrier electrolyte. See also 2.3 of this index for electrolytes in organic solvents either or not in the presence of water.

K_3AsO_4 89, 90
KBr 2, 101, 149*-161*, 195*
KCl 1, 74, 99, 133*-146*, 189*-192*, 240*-250*, 257*-265*, 267*-269*, 363*-378*
$KClO_4$ 124, 128*-130*
KCNO 87
K_2CO_3 86, 87
$K_6Co(CN)_{12}$ 98
$KEtSO_4$ 96
KF 19, 81
$K_4Fe(CN)_6$ 97
KH_2AsO_4 90, 91
K_2HAsO_4 90
$KHCO_3$ 87
KH oxalate 95
K_2HPO_3 88
KH_2PO_4 167*
K_2HPO_4 87, 88
KI 35-45, 103, 194*, 346*
KNO_2 86
KNO_3 22, 85, 169*-172*, 220*-222*, 344*, 345*
K oxalate 95
KOH 84
$K(OH)SO_4$ 93
$K_4P_2O_7$ 88
$K_2Pt(CN)_4$ 98
KSCNO 96
K_2SeO_4 94
K_2SO_4 91, 266*
K succinate 95, 96
LiBr 76, 77, 101
LiCl 9, 32, 60-73, 99, 147*, 185*-187*
LiI 78-80, 103
LiH_2PO_4 165*
$LiNO_3$ 18
LiOH 84
Li_2SO_4 91, 92
$MgAc_2$ 95
$MgBr_2$ 102
$MgCl_2$ 100
$Mg(H_2PO_2)_2$ 89

$MgPt(CN)_4$	98, 99
$Mg(SCNO)_2$	97
$MgSO_4$	31, 92
NaAc	29, 94
$NaBF_4$	110
$Na_2B_4O_7$	86
NaBr	101
$NaBrO_3$	14
NaCl	75, 99
$NaClO_3$	15
$NaClO_4$	17, 20, 164*, 202*-207*, 280*
NaCN	16
Na_2CO_3	87
NaF	33, 34, 82, 193*, 196*, 197*, 224*-226*, 270*-277*
Na formate	28, 174*, 311*
Na fumarate	113, 114
$Na_2HA_sO_4$	91
NaH_2PO_2	89
NaH_2PO_4	3, 4, 166*, 230*-232*
Na_2HPO_3	88
Na_2HPO_4	88
NaI	103
Na maleate	111, 112
NaN_3	23-27
$NaNO_2$	85
NaOH	84, 227*-229*
Na_3PO_4	88
Na propionate	30
NaSCN	13
NaSCNO	96, 97
$Na_8SiW_{12}O_{42}$	94
Na_2SO_3	93
Na_2SO_4	91, 115-119, 173*, 209*-219*, 251*-255*, 315*-317*, 319*-344*, 347*, 353*-362*, 384*-391*
$Na_2S_2O_3$	94
$Na_2S_2O_6$	91
NH_4Br	109
NH_4CNS	109
$(NH_4)HC_2O_4$	109
$(NH_4)_4H_2(CO_3)_3$	109

$(NH_4)_2HPO_4$	108, 345*
NH_4NO_2	108
NH_4NO_3	21, 108, 184*
NH_4OH	108, 353*
$(NH_4)_2SO_4$	108
$N(Me)_4Br$	293, 294
$N(Me)_4Cl$	300, 301
$N(Me)_4F$	309, 310
$N(Me)_4I$	286
$N(Me)_4OH$	321*, 348*, 349*
$N(Et)_4Br$	291, 292
$N(Et)_4Cl$	298, 299
$N(Et)_4ClO_4$	125-127, 128*-130*
$N(Et)_4F$	306-308
$N(Et)_4I$	284, 285
$N(Et)_4OH$	348*-350*
$N(Prop)_4Br$	289, 290
$N(Prop)_4Cl$	296, 297
$N(Prop)_4F$	304, 305
$N(Prop)_4I$	282, 283
$N(But)_4Br$	287, 288
$N(But)_4Cl$	295
$N(But)_4F$	302, 303
$N(But)_4I$	281
RbBr	101
RbCl	10, 100
RbOH	84
$SrBr_2$	101
$SrCl_2$	100
$Sr(NO_3)_2$	85
$Sr(OH)_2$	84

2.2. Electrolyte mixtures, including organic electrolytes

(Ba + H)Cl	120-123
(Cs + Li)Cl	147
$(H + Hg)NO_3$	131, 132
$(H + Hg + K)NO_3$	169-172
(H + Na)formate	174
$[K + N(Et)_4]ClO_4$	128-130

formic acid + CsH_2PO_4 168
formic acid + KBr 161
formic acid + KCl 133
formic acid + KH_2PO_4 167
formic acid + LiH_2PO_4 165
formic acid + NaH_2PO_4 166
formic acid + Na_2SO_4 173
HAc + KBr 152, 160
HAc(mono-,di-,tri-Cl) + KBr 155-157
HAc + KCl 134, 135
HAc(mono-,di-,tri-Cl) + KCl 136-138
propionic acid + KBr 151, 159
propionic acid + KCl 139, 140
butyric acid + KBr 158
butyric acid (iso-) + KBr 149, 153
butyric acid + KCl 141, 142
butyric acid (iso-) + KCl 143, 144
valeric acid + KBr 150, 154
valeric acid + KCl 145, 146
valeric acid + $HClO_4$ 162, 163
benzoic acid(o-hydroxy) + $HClO_4$ 148
benzoic acid + $NaClO_4$ 164
Na(benzoate + SO_4) + phenol 322
Na(cinnamate + phtalate + SO_4) 320
Na(cinnamate + SO_4) 320
Na(cinnamate + SO_4) + K-benzene-sulphonate 320
Na(cinnamate + terephtalate + SO_4) 320
Na(cinnamate + triCl-acetate+SO_4) 320
Na(terephtalate + SO_4) 320
Na(terephtalate + triCl-acetate + SO_4) 321
Na(toluate + SO_4) 320
Na(toluate + salicylate + SO_4) 320
Na(triCl-acetate + SO_4) + K-benzene-sulphonate 321
$N(Me)_4OH$ + Na_2SO_4 + p-toluene-sulphonic acid 321

2.3. Solutions containing organic substances

(See also sec. 2.1 and 2.2 for some organic electrolytes.)

Note. Some of these substances occur only as a component of a mixture; inorganic carrier electrolyte is almost always present.

acetal 335
acetamide 341

acetanilide 230-232, 342
acetic acid 134, 135, 152, 160, 337
acetic acid (chloro-) 136-138, 155-157, 337
acetic aldehyde 334
acetone 335
acetone (diethyl) 336
acetone (methyl-ethyl) 336
acetonitrile 270, 271, 334
acetylacetone 280, 336
alanine 340
allyl alcohol 190, 326
amine (allyl) 356, 357
amine (benzyl) 360
amine (diisoamyl) 356
amine (heptyl) 356
amine (methyl) 353
amine (naphtyl) 360
amine (isoamyl) 355, 356
amine (isobutyl) 354
amine (triethyl) 354
amine (trimethyl) 353
amine (triisobutyl) 354, 355
amyl alcohol see pentanol
amygdaline 343
aniline 256, 322, 346, 358, 359
anisol 333
anisyl alcohol 326
asparagine 340
benzene 278
benzoic acid 340
benzoic acid (o,p-hydroxy) 148, 164, 332
benzoic aldehyde (p-oxy) 331
benzyl alcohol 326
biuret 343
butanol (n,i,tert) 187, 188, 200, 201, 209-218, 324
butyl chloral 336
butyl aldehyde (iso-) 334
butyramide 342
butyric acid 141, 142
butyric acid (iso-) 143, 144, 149, 153, 158, 337
butyric acid (oxy-) 338

caffeine 315-317, 321, 322, 345, 358
caproic acid 338
capronitrile 334
carboline + deriv. 319
chloroform 233, 337
choline 350
citric acid 339
cocaine 362
codeine 362
coniferyl alcohol 331
cresol (o,m,p-) 224-229
crotonic acid 338
cyclohexanol 223
dioxane 269
dioxynphtaline 331
erythritol 327
ethanol 324
ether (dibutyl-) 311
ether (diethyl-) 332
ethyl acetate 333
ethyl bromide 336
ethylene diamine 357
ethylene (trimethyl-) 324
ethyl formiate 333
formamide 341
formic acid 133, 161, 165-168, 311
fumaric acid 339
furan (2-carbonic acid) 341
furan (tetrahydro-) 255
furfural 335
furfuryl alcohol 250, 253
furfuryl alcohol (tetrahydro-) 251
furfurylamine 234-238, 240-246, 254
furfurylamine (N-methyl-) 248
furfurylamine (N-methyltetrahydro-) 249
furfurylamine (tetrahydro-) 239, 247, 252
galactitol 327
galactose 328
gallic acid 332
glucose 328

glyceric acid 338
glyceric mono-Cl-hydrine 333
glycerol 327
glycol (ethylene-) 192-195, 326, 340
glyceric acid 338
glyceric mono-Cl-hydrine 333
glycerol 327
glycol (isobutyl-) 326
glyoxal 335
guanidine 358
heptanol (n-) 325, 335
heptanone 335
hexanoic acid see caproic acid
hyamine 318
hydrazine (phenyl-) 360
hypoxanthine 343
inositol 328
lactic acid 338
lactose 321, 322, 329
leucine 340
malonic acid 339
maltose 329
mannitol 327
methanol 175-184, 202-207, 324
morpholine 268
naphtol 330
narcotine 312, 313
neurine 350
octanol 208
oxalic acid 105, 338
parabanic acid 343
paraldehyde 334
pentanoic acid see valeric acid
pentanol (n-,tert-,iso-) 198, 199, 325, 344, 345
pentanol (chloro-) 196, 197
perseitol 327
phenetol 333
phenol 219, 316, 317, 321, 322, 329, 344-346
phenol (isobutyl-) 329
phtalic acid (o-) 341

picoline	361
pilocarpine	362
pinacol hydrate	326
pinacolon	336
piperazine	357
piperidine	267, 362
piperonal	333
propanol (n-, iso-)	185, 186, 189, 324
propargyl alcohol	191
propioamide	342
propionic acid	139, 140, 151, 159, 337
propionic aldehyde	334
propionitrile	275-277, 334
protocatechic acid	332
protocatechic aldehyde	331
pyridine	257-266, 361
pyrocatechine	330
pyrogallol	331
pyrrol	358
quercitol	328
quinine	314
quinoline	220
quinoline (methyl-)	221, 222
quinone (hydro-)	330
raffinose	329
resorcitol	330, 346
saccharose	328
salicin , saligenin-β	331, 343
saligenine	see salicin
sarcosine	340
suberic acid	339
succinic acid	105
succinonitrile	272-274
sulfine (trimethyl-) hydr.	351-353
tartaric acid	339
tartaric acid (pyro-)	339
taurine	339
terpin hydrate	327
toluene	279
toluic acid (p-)	341
toluidine (o,m,p-)	322, 359

tyrosine 332
valeric acid (n,iso-) 145, 146, 150, 154, 162, 163, 337
valeric aldehyde 334
valeroacetone 341
vanillic acid 332
vanilline 332
xylidine (m-) 359, 360
xylose 328
urea 342
urea (methyl-) 342
urethane 342

2.4. Substances with ionic surfactant character

For the sake of space the following abbreviations are used in this subsection: A = ammonium; B = butyl; E = ethyl; M = methyl; P = propyl; Py = pyridinium; iQ = isoquinolinium; T = tri. C_x stands for n-hydrocarbon chain with x C atoms.

C_8TMABr 369, 392
$(C_8TMA)_2SO_4$ 381, 386
$(C_{10}TMA)_2SO_4$ 388
C_{12}TMABr 363, 393
$(C_{12}TMA)_2SO_4$ 389, 395
$(C_{14}TMA)_2SO_4$ 390
C_{16}TMABr 364, 394
$(C_{16}TMA)_2SO_4$ 383, 391
C_8TEABr 368, 377
$(C_8TEA)_2SO_4$ 380, 385
C_{12}TEABr 365
C_{16}TEABr 366
C_8TPABr 370, 378
$(C_8TPA)_2SO_4$ 382, 387
C_{12}TPABr 372
C_8TBABr 367, 376
$(C_8TBA)_2SO_4$ 379, 384
C_{12}TBABr 371
C_8PyBr 373
C_{12}PyBr 374
C_{10}iQBr 375.

Differential capacity of mercury in tenth-normal solutions of metallic chlorides. T = 25°C. Potentials measured relative to calomel electrode in same solution.

Reference : D.C. Grahame, J. Electrochem. Soc., 98, 343 (1951)

E Volts	Capacity in microfarads per square centimeter LiCl	NaCl	KCl	RbCl	CsCl	HCl	NH_4Cl	$MgCl_2$
-0.061			101.6	107.2	108.8	99.7		
-0.062			99.5	104.8	106.4	97.6	98.9	96.0
-0.063	94.1	94.9	97.2	102.4	103.8	95.4	96.6	94.1
-0.065	90.5	91.3	93.5	98.2	99.5	91.8	92.9	90.9
-0.070	83.1	83.9	85.5	89.3	90.4	84.0	85.0	84.4
-0.075	77.0	77.8	79.0	82.2	83.2	77.7	78.6	79.0
-0.080	72.0	72.7	73.8	76.5	77.4	72.6	73.4	74.7
-0.085	67.9	68.5	69.5	71.8	72.6	68.4	69.1	70.8
-0.090	64.4	64.9	65.8	67.8	68.5	64.8	65.5	67.4
-0.095	61.4	62.0	62.7	64.5	65.1	61.8	62.4	64.4
-0.100	59.1	59.6	60.3	61.8	62.4	59.5	60.0	62.0
-0.105	57.0	57.5	58.0	59.5	59.9	57.3	57.8	59.7
-0.110	55.3	55.6	56.2	57.5	57.8	55.5	55.9	57.7
-0.115	53.6	53.9	54.4	55.6	55.9	53.8	54.2	55.8
-0.120	52.1	52.4	52.8	54.0	54.2	52.2	52.6	54.2
-0.125	50.8	51.0	51.4	52.5	52.6	50.8	51.2	52.8
-0.130	49.6	49.8	50.1	51.1	51.2	49.6	50.0	51.7
-0.135	48.6	48.7	49.1	50.0	50.0	48.6	48.9	50.5
-0.140	47.6	47.7	48.0	48.9	48.9	47.5	47.9	49.2
-0.145	46.7	46.9	47.2	48.0	48.0	46.7	47.0	48.2
-0.150	45.9	46.0	46.3	47.1	47.1	45.8	46.1	47.2
-0.155	45.2	45.2	45.5	46.2	46.2	45.1	45.3	46.4
-0.16	44.52	44.61	44.79	45.51	45.51	44.39	44.66	45.64
-0.17	43.30	43.30	43.47	44.12	44.12	43.08	43.34	44.25
-0.18	42.26	42.26	42.39	43.03	42.98	42.05	42.26	43.11
-0.19	41.38	41.34	41.46	42.04	42.00	41.13	41.34	42.16
-0.20	40.63	40.55	40.67	41.20	41.16	40.34	40.55	41.32
-0.21	40.01	39.95	40.03	40.54	40.47	39.73	39.91	40.63
-0.22	39.46	39.39	39.46	39.93	39.85	39.17	39.35	40.01
-0.23	39.02	38.94	39.00	39.44	39.35	38.72	38.89	39.52
-0.24	38.67	38.58	38.64	39.04	38.95	38.37	38.54	39.14
-0.25	38.41	38.31	38.37	38.75	38.64	38.10	38.29	38.84
-0.26	38.22	38.12	38.17	38.51	38.40	37.91	38.12	38.61
-0.27	38.10	38.00	38.05	38.37	38.24	37.79	38.03	38.46
-0.28	38.05	37.96	38.01	38.29	38.17	37.76	38.02	38.39
-0.29	38.08	38.00	38.04	38.30	38.17	37.79	38.08	38.39
-0.30	38.17	38.09	38.13	38.36	38.24	37.89	38.19	38.44
-0.31	38.28	38.20	38.23	38.42	38.35	38.00	38.31	38.51
-0.32	38.43	38.35	38.37	38.53	38.51	38.15	38.45	38.62
-0.33	38.61	38.52	38.54	38.68	38.69	38.33	38.62	38.76
-0.34	38.79	38.71	38.72	38.83	38.89	38.52	38.80	38.91
-0.35	38.97	38.89	38.89	38.98	39.08	38.70	38.97	39.05
-0.36	39.19	39.12	39.12	39.19	39.31	38.94	39.19	39.26
-0.37	39.44	39.38	39.38	39.43	39.55	39.21	39.44	39.50
-0.38	39.60	39.55	39.55	39.58	39.71	39.40	39.61	39.66
-0.39	39.71	39.68	39.68	39.69	39.82	39.55	39.73	39.78
-0.40	39.82	39.82	39.82	39.82	39.94	39.70	39.86	39.90

cont.

Capacity in 0.1 N solutions of metallic chlorides (cont.)

E Volts	Capacity in microfarads per square centimeter							
	LiCl	NaCl	KCl	RbCl	CsCl	HCl	NH_4Cl	$MgCl_2$
-0.41	39.84	39.85	39.85	39.84	39.96	39.76	39.87	39.93
-0.42	39.76	39.76	39.78	39.75	39.88	39.71	39.79	39.86
-0.43	39.63	39.62	39.66	39.62	39.75	39.62	39.66	39.74
-0.44	39.40	39.38	39.44	39.40	39.52	39.43	39.44	39.52
-0.45	39.07	39.04	39.11	39.07	39.19	39.14	39.11	39.20
-0.46	38.62	38.58	38.67	38.63	38.75	38.73	38.66	38.77
-0.47	38.07	38.03	38.12	38.08	38.20	38.21	38.11	38.25
-0.48	37.47	37.44	37.53	37.49	37.61	37.65	37.52	37.70
-0.49	36.71	36.70	36.78	36.74	36.85	36.94	36.77	36.99
-0.50	35.82	35.81	35.89	35.85	35.96	36.08	35.89	36.15
-0.51	34.83	34.85	34.92	34.89	35.02	35.15	34.92	35.23
-0.52	33.84	33.88	33.95	33.93	34.09	34.23	33.96	34.31
-0.53	32.82	32.87	32.94	32.93	33.11	33.25	32.95	33.36
-0.54	31.75	31.82	31.88	31.88	32.08	32.21	31.90	32.37
-0.55	30.67	30.75	30.81	30.82	31.04	31.15	30.83	31.38
-0.56	29.65	29.75	29.80	29.82	30.07	30.14	29.83	30.45
-0.57	28.65	28.75	28.80	28.83	29.12	29.14	28.83	29.55
-0.58	27.68	27.79	27.84	27.88	28.20	28.17	27.88	28.71
-0.59	26.81	26.93	26.98	27.03	27.38	27.30	27.03	27.98
-0.60	25.97	26.08	26.15	26.20	26.59	26.45	26.20	27.23
-0.61	25.24	25.37	25.43	25.49	25.93	25.70	25.49	26.61
-0.62	24.58	24.72	24.78	24.85	25.32	25.03	24.85	26.03
-0.63	23.95	24.10	24.16	24.21	24.74	24.38	24.23	25.49
-0.64	23.39	23.55	23.62	23.70	24.26	23.82	23.70	25.00
-0.65	22.90	23.09	23.14	23.23	23.83	23.30	23.23	24.59
-0.66	22.43	22.60	22.69	22.79	23.42	22.78	22.78	24.16
-0.67	22.01	22.19	22.28	22.39	23.04	22.29	22.38	23.75
-0.68	21.64	21.83	21.93	22.04	22.70	21.86	22.03	23.38
-0.69	21.30	21.49	21.60	21.72	22.39	21.46	21.71	23.02
-0.70	21.00	21.19	21.31	21.44	22.12	21.10	21.43	22.68
-0.72	20.44	20.65	20.78	20.93	21.62	20.44	20.92	22.02
-0.74	19.94	20.16	20.30	20.46	21.16	19.83	20.45	21.41
-0.76	19.54	19.76	19.91	20.10	20.79	19.34	20.08	20.86
-0.78	19.19	19.42	19.57	19.77	20.47	18.95	19.75	20.33
-0.80	18.84	19.08	19.23	19.45	20.15	18.58	19.42	19.85
-0.82	18.50	18.74	18.89	19.13	19.83	18.22	19.08	19.38
-0.84	18.18	18.43	18.58	18.83	19.54	17.90	18.78	18.96
-0.86	17.90	18.14	18.30	18.57	19.27	17.62	18.50	18.59
-0.88	17.62	17.85	18.02	18.29	19.01	17.34	18.23	18.25
-0.90	17.35	17.58	17.76	18.04	18.77	17.08	17.98	17.91
-0.92	17.10	17.33	17.51	17.80	18.54	16.83	17.74	17.60
-0.94	16.85	17.08	17.27	17.58	18.31	16.60	17.52	17.31
-0.96	16.64	16.88	17.07	17.39	18.13	16.40	17.33	17.07
-0.98	16.46	16.71	16.90	17.23	17.98	16.24	17.18	16.87
-1.00	16.32	16.57	16.77	17.11	17.86	16.12	17.07	15.71
-1.02	16.12	16.37	16.58	16.93	16.69	15.93	16.89	16.48
-1.04	15.97	16.23	16.45	16.81	17.57	15.81	16.77	16.32
-1.06	15.86	16.13	16.36	16.73	17.50		16.69	16.21
-1.08	15.75	16.03	16.27	16.66	17.43		16.61	16.09
-1.10	15.64	15.93	16.18	16.58	17.35		16.53	15.97
-1.12	15.58	15.88	16.14	16.55	17.33		16.50	15.90

cont.

Capacity in 0.1 N solutions of metallic chlorides (cont.)

Capacity in microfarads per square centimeter

E Volts	LiCl	NaCl	KCL	RbCl	CsCl	NH_4Cl	$MgCl_2$
-1.14	15.51	15.82	16.09	16.52	17.30	16.46	15.83
-1.16	15.46	15.78	16.06	16.50	17.29	16.45	15.78
-1.18	15.42	15.75	16.05	16.51	17.30	16.46	15.75
-1.20	15.39	15.73	16.04	16.52	17.31	16.47	15.74
-1.22	15.38	15.73	16.05	16.55	17.34	16.50	15.74
-1.24	15.36	15.72	16.06	16.57	17.38	16.54	15.74
-1.26	15.37	15.73	16.09	16.62	17.42	16.60	15.77
-1.28	15.40	15.78	16.15	16.70	17.50	16.69	15.83
-1.30	15.43	15.82	16.21	16.77	17.58	16.78	15.89
-1.32	15.47	15.87	16.28	16.86	17.66	16.88	15.96
-1.34	15.51	15.92	16.34	16.94	17.73	16.98	16.01
-1.36	15.56	15.98	16.42	17.03	17.82	17.09	16.09
-1.38	15.66	16.09	16.55	17.18	17.96	17.26	16.22
-1.40	15.75	16.19	16.67	17.31	18.09	17.42	16.34
-1.42	15.83	16.29	16.79	17.45	18.22	17.58	16.46
-1.44	15.94	16.39	16.92	17.60	18.35	17.75	16.59
-1.46	16.04	16.50	17.06	17.75	18.50	17.95	16.73
-1.48	16.17	16.64	17.22	17.93	18.67	18.16	16.89
-1.50	16.29	16.76	17.38	18.10	18.84	18.38	17.06
-1.52	16.43	16.89	17.54	18.27	19.01	18.59	17.22
-1.54	16.56	17.01	17.70	18.45	19.18	18.82	17.40
-1.56	16.71	17.15	17.87	18.63	19.36		17.58
-1.58	16.89	17.32	18.08	18.85	19.58		17.81
-1.60	17.07	17.49	18.29	19.07	19.79		18.05
-1.62	17.24	17.64	18.49	19.29	20.00		18.27
-1.64	17.42	17.80	18.70	19.51	20.21		18.51
-1.66	17.60	17.97	18.92	19.74	20.44		18.75
-1.68	17.79	18.14	19.16	19.99	20.68		19.00
-1.70	17.98	18.32	19.41	20.25	20.94		19.27
-1.72	18.15	18.48	19.65	20.50	21.19		19.51
-1.74	18.33	18.67	19.90	20.76	21.44		19.74
-1.76	18.51	18.84	20.15	21.02	21.70		19.97
-1.78	18.74	19.02	20.44	21.33	21.99		20.23
-1.80	19.00	19.24	20.76	21.66	22.32		20.52
-1.82	19.29	19.49	21.11	22.02	22.67		20.83
-1.84	19.61	19.76	21.49	22.41	23.08		21.17
-1.86	19.97	20.07	21.93	22.87	23.54		
-1.88	20.34	20.42	22.39	23.34	24.02		
-1.90	20.70	20.81	22.90	23.86	24.57		
-1.92	21.06	21.26	23.44	24.41	25.14		
-1.94	21.37	21.76	24.03	25.02	25.76		

Observed values of differential capacity, C, and surface charge density, σ, for mercury in aqueous solutions of sodium fluoride. T = 25°C. Potentials relative to normal calomel electrode. Values of σ refer to potentials half way between listed potentials. Values of C and σ appear alternately. Values of C are in microfarads per sq. cm. Values of σ are in microcoulombs per sq. cm. and are listed without sign. Values of σ appearing near the top of the table are negative. They go to zero and become positive as one moves down the table.

Reference: D.C. Grahame, J.Amer. Chem. Soc., 76, 4819 (1954)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-1.90					19.19 C
			25.39 σ	24.08	21.74
-1.89			22.04 C	20.61 C	19.08 C
			25.17 σ	23.87	21.55
-1.88			21.82 C	20.45 C	18.97 C
			24.96 σ	23.67	21.36
-1.87			21.58 C	20.30 C	18.85 C
			24.74	23.46	21.17
-1.86			21.36 C	20.15 C	18.73 C
			24.53	23.26	20.99
-1.85			21.16 C	20.00 C	18.62 C
	25.55 σ		24.32	23.06	20.80
-1.84	21.81 C		20.96 C	19.86 C	18.52 C
	25.33 σ		24.10	22.86	20.62
-1.83	21.52 C		20.77 C	19.72 C	18.41 C
	25.12 σ		23.90	22.67	20.43
-1.82	21.30 C		20.62 C	19.60 C	18.31 C
	24.90	24.75	23.69	22.47	20.25
-1.81	21.10 C	21.50 C	20.43 C	19.48 C	18.21 C
	24.69	24.54	23.49	22.28	20.07
-1.80	20.90 C	21.33 C	20.26 C	19.33 C	18.10 C
	24.48	24.32	23.28	22.08	19.88
-1.79	20.72 C	21.15 C	20.11 C	19.23 C	18.01 C
	24.27	24.11	23.08	21.89	19.70
-1.78	20.60 C	20.97 C	19.96 C	19.10 C	17.92 C
	24.07	23.90	22.88	21.70	19.53
-1.77	20.45 C	20.81 C	19.80 C	18.99 C	17.83 C
	23.86	23.69	22.69	21.51	19.35
-1.76	20.31 C	20.64 C	19.67 C	18.89 C	17.73 C
	23.66	23.49	22.49	21.32	19.17
-1.75	20.20 C	20.50 C	19.53 C	18.78 C	17.64 C
	23.46	23.28	22.29	21.13	18.99
-1.74	20.09 C	20.33 C	19.40 C	18.66 C	17.55 C
	23.26	23.08	22.10	20.94	18.82
-1.73	19.96 C	20.19 C	19.29 C	18.55 C	17.47 C
	23.06	22.88	21.91	20.76	18.64
-1.72	19.83 C	20.04 C	19.19 C	18.44 C	17.38 C
	22.86	22.68	21.72	20.58	18.47
-1.71	19.72 C	19.91 C	19.07 C	18.33 C	17.29 C
	22.66	22.48	21.52	20.39	18.30

cont.

Capacity and charge in NaF-solutions (cont.)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-1.70	19.61 C	19.78 C	18.97 C	18.24 C	17.20 C
	22.47	22.28	21.33	20.21	18.12
-1.69	19.51 C	19.64 C	18.85 C	18.13 C	17.09 C
	22.27	22.08	21.15	20.03	17.95
-1.68	19.41 C	19.53 C	18.73 C	18.02 C	17.00 C
	22.08	21.89	20.96	19.85	17.78
-1.67	19.30 C	19.40 C	18.61 C	17.97 C	16.90 C
	21.88	21.69	20.77	19.67	17.62
-1.66	19.21 C	19.29 C	18.52 C	17.88 C	16.81 C
	21.69	21.50	20.59	19.49	17.45
-1.65	19.10 C	19.16 C	18.43 C	17.80 C	16.72 C
	21.50	21.31	20.40	19.31	17.28
-1.64	18.99 C	19.05 C	18.33 C	17.72 C	16.66 C
	21.31	21.12	20.22	19.13	17.11
-1.63	18.91 C	18.95 C	18.22 C	17.65 C	16.58 C
	21.12	20.93	20.04	18.96	16.95
-1.62	18.80 C	18.86 C	18.12 C	17.57 C	16.50 C
	20.93	20.74	19.86	18.78	16.78
-1.61	18.70 C	18.75 C	18.02 C	17.50 C	16.43 C
	20.75	20.55	19.68	18.61	16.62
-1.60	18.61 C	18.66 C	17.94 C	17.42 C	16.36 C
	20.56	20.36	19.50	18.43	16.45
-1.59	18.51 C	18.57 C	17.87 C	17.34 C	16.30 C
	20.38	20.18	19.32	18.26	16.29
-1.58	18.42 C	18.49 C	17.80 C	17.27 C	16.23 C
	20.19	19.99	19.14	18.09	16.13
-1.57	18.33 C	18.42 C	17.72 C	17.19 C	16.17 C
	20.01	19.81	18.96	17.92	15.97
-1.56	18.25 C	18.34 C	17.61 C	17.12 C	16.11 C
	19.83	19.63	18.79	17.74	15.81
-1.55	18.16 C	18.27 C	17.53 C	17.03 C	16.04 C
	19.64	19.44	18.61	17.57	15.65
-1.54	18.08 C	18.19 C	17.46 C	16.98 C	15.99 C
	19.46	19.26	18.44	17.40	15.49
-1.53	18.00 C	18.11 C	17.41 C	16.91 C	15.93 C
	19.28	19.08	18.26	17.23	15.33
-1.52	17.91 C	18.02 C	17.33 C	16.83 C	15.88 C
	19.10	18.90	18.09	17.07	15.17
-1.51	17.82 C	17.94 C	17.26 C	16.78 C	15.83 C
	18.93	18.72	17.92	16.90	15.01
-1.50	17.75 C	17.87 C	17.19 C	16.71 C	15.78 C
	18.75	18.54	17.74	16.73	14.85
-1.49	17.68 C	17.80 C	17.10 C	16.66 C	15.73 C
	18.57	18.36	17.57	16.56	14.69
-1.48	17.61 C	17.73 C	17.05 C	16.59 C	15.68 C
	18.40	18.19	17.40	16.40	14.54
-1.47	17.52 C	17.64 C	16.98 C	16.53 C	15.63 C
	18.22	18.01	17.23	16.23	14.38
-1.46	17.45 C	17.56 C	16.92 C	16.47 C	15.60 C
	18.05	17.84	17.06	16.07	14.23
-1.45	17.39 C	17.50 C	16.85 C	16.42 C	15.55 C
	17.87	17.66	16.90	15.90	14.07
-1.44	17.32 C	17.42 C	16.78 C	16.36 C	15.50 C
	17.70	17.49	16.73	15.74	13.92

cont.

Capacity and charge in NaF-solutions (cont.)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-1.43	17.26 C	17.35 C	16.72 C	16.30 C	15.47 C
	17.53	17.31	16.56	15.58	13.76
-1.42	17.19 C	17.25 C	16.65 C	16.26 C	15.42 C
	17.35	17.14	16.39	15.42	13.61
-1.41	17.11 C	17.18 C	16.60 C	16.20 C	15.38 C
	17.18	16.97	16.23	15.25	13.45
-1.40	17.06 C	17.13 C	16.54 C	16.15 C	15.33 C
	17.01	16.80	16.06	15.09	13.30
-1.39	16.99 C	17.06 C	16.49 C	16.10 C	15.29 C
	16.84	16.63	15.90	14.93	13.15
-1.38	16.93 C	17.02 C	16.42 C	16.06 C	15.26 C
	16.67	16.46	15.73	14.77	12.99
-1.37	16.88 C	16.96 C	16.35 C	16.01 C	15.22 C
	16.50	16.29	15.57	14.61	12.84
-1.36	16.81 C	16.91 C	16.32 C	15.97 C	15.18 C
	16.34	16.12	15.41	14.45	12.69
-1.35	16.78 C	16.85 C	16.27 C	15.92 C	15.13 C
	16.17	15.95	15.24	14.29	12.54
-1.34	16.71 C	16.80 C	16.22 C	15.89 C	15.09 C
	16.00	15.78	15.08	14.13	12.39
-1.33	16.66 C	16.73 C	16.18 C	15.85 C	15.05 C
	15.84	15.61	14.92	13.97	12.24
-1.32	16.61 C	16.69 C	16.14 C	15.81 C	15.01 C
	15.67	15.45	14.76	13.82	12.09
-1.31	16.58 C	16.63 C	16.09 C	15.79 C	14.99 C
	15.50	15.28	14.60	13.66	11.94
-1.30	16.52 C	16.58 C	16.04 C	15.77 C	14.95 C
	15.34	15.12	14.44	13.50	11.79
-1.29	16.49 C	16.54 C	16.00 C	15.73 C	14.92 C
	15.17	14.95	14.28	13.34	11.64
-1.28	16.44 C	16.49 C	15.98 C	15.71 C	14.90 C
	15.01	14.78	14.12	13.19	11.49
-1.27	16.40 C	16.44 C	15.93 C	15.69 C	14.89 C
	14.84	14.62	13.96	13.03	11.34
-1.26	16.35 C	16.39 C	15.91 C	15.67 C	14.87 C
	14.68	14.46	13.80	12.87	11.19
-1.25	16.33 C	16.34 C	15.88 C	15.64 C	14.86 C
	14.52	14.29	13.64	12.72	11.04
-1.24	16.30 C	16.29 C	15.85 C	15.62 C	14.84 C
	14.36	14.13	13.48	12.56	10.89
-1.23	16.26 C	16.25 C	15.83 C	15.60 C	14.83 C
	14.19	13.97	13.32	12.40	10.75
-1.22	16.23 C	16.22 C	15.80 C	15.58 C	14.82 C
	14.03	13.80	13.17	12.25	10.60
-1.21	16.20 C	16.19 C	15.78 C	15.57 C	14.81 C
	13.87	13.64	13.01	12.09	10.45
-1.20	16.18 C	16.15 C	15.76 C	15.56 C	14.80 C
	13.71	13.48	12.85	11.94	10.30
-1.19	16.14 C	16.12 C	15.73 C	15.55 C	14.80 C
	13.55	13.32	12.69	11.78	10.15
-1.18	16.12 C	16.08 C	15.73 C	15.54 C	14.81 C
	13.38	13.16	12.54	11.63	10.01

cont.

Capacity and charge in NaF-solutions (cont.)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-1.17	16.11 C	16.06 C	15.72 C	15.55 C	14.82 C
	13.22	13.00	12.38	11.47	9.86
-1.16	16.09 C	16.04 C	15.73 C	15.57 C	14.83 C
	13.06	12.84	12.22	11.31	9.71
-1.15	16.08 C	16.03 C	15.73 C	15.58 C	14.84 C
	12.90	12.68	12.06	11.16	9.56
-1.14	16.04 C	16.02 C	15.75 C	15.60 C	14.86 C
	12.74	12.52	11.91	11.00	9.41
-1.13	16.03 C	16.01 C	15.76 C	15.62 C	14.88 C
	12.58	12.36	11.75	10.85	9.26
-1.12	16.02 C	16.00 C	15.77 C	15.66 C	14.90 C
	12.42	12.20	11.59	10.69	9.11
-1.11	16.01 C	15.99 C	15.78 C	15.69 C	14.92 C
	12.26	12.04	11.43	10.53	8.97
-1.10	16.01 C	15.98 C	15.80 C	15.72 C	14.94 C
	12.10	11.88	11.28	10.38	8.82
-1.09	16.00 C	15.98 C	15.82 C	15.74 C	14.97 C
	11.94	11.72	11.12	10.22	8.67
-1.08	16.01 C	15.99 C	15.83 C	15.80 C	14.99 C
	11.78	11.56	10.96	10.06	8.52
-1.07	16.02 C	16.00 C	15.86 C	15.83 C	15.01 C
	11.62	11.40	10.80	9.90	8.37
-1.06	16.04 C	16.01 C	15.88 C	15.89 C	15.03 C
	11.46	11.24	10.64	9.74	8.22
-1.05	16.06 C	16.02 C	15.92 C	15.93 C	15.07 C
	11.30	11.08	10.48	9.58	8.07
-1.04	16.10 C	16.04 C	15.96 C	15.98 C	15.10 C
	11.14	10.92	10.32	9.42	7.91
-1.03	16.12 C	16.07 C	15.98 C	16.02 C	15.13 C
	10.98	10.76	10.16	9.26	7.76
-1.02	16.15 C	16.10 C	16.04 C	16.09 C	15.16 C
	10.82	10.60	10.00	9.10	7.61
-1.01	16.20 C	16.13 C	16.09 C	16.14	15.20 C
	10.65	10.44	9.84	8.94	7.46
-1.00	16.22 C	16.18 C	16.14 C	16.22 C	15.23 C
	10.49	10.27	9.68	8.78	7.31
-0.99	16.28 C	16.24 C	16.21 C	16.29 C	15.29 C
	10.33	10.11	9.52	8.62	7.15
-0.98	16.31 C	16.29 C	16.27 C	16.33 C	15.32 C
	10.17	9.95	9.36	8.45	7.00
-0.97	16.39 C	16.33 C	16.34 C	16.42 C	15.38 C
	10.00	9.79	9.19	8.29	6.85
-0.96	16.44 C	16.38 C	16.40 C	16.51 C	15.42 C
	9.84	9.62	9.03	8.12	6.69
-0.95	16.51 C	16.45 C	16.48 C	16.60 C	15.50 C
	9.67	9.46	8.86	7.96	6.54
-0.94	16.60 C	16.52 C	16.54 C	16.68 C	15.56 C
	9.51	9.29	8.70	7.79	6.38
-0.93	16.69 C	16.57 C	16.63 C	16.75 C	15.62 C
	9.34	9.13	8.53	7.62	6.23
-0.92	16.78 C	16.64 C	16.70 C	16.83 C	15.70 C
	9.17	8.96	8.37	7.46	6.07
-0.91	16.88 C	16.70 C	16.78 C	16.92 C	15.77 C
	9.00	8.79	8.20	7.29	5.91

cont.

Capacity and charge in NaF-solutions (cont.)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-0.90	16.98 C	16.77 C	16.86 C	17.01 C	15.83 C
	8.83	8.62	8.03	7.12	5.75
-0.89	17.09 C	16.86 C	16.95 C	17.10 C	15.91 C
	8.66	8.46	7.86	6.95	5.59
-0.88	17.18 C	16.95 C	17.03 C	17.20 C	15.97 C
	8.49	8.29	7.69	6.77	5.43
-0.87	17.29 C	17.04 C	17.12 C	17.30 C	16.02 C
	8.32	8.12	7.52	6.60	5.27
-0.86	17.41 C	17.14 C	17.24 C	17.40 C	16.09 C
	8.14	7.95	7.35	6.43	5.11
-0.85	17.52 C	17.24 C	17.35 C	17.50 C	16.14 C
	7.97	7.77	7.17	6.25	4.95
-0.84	17.66 C	17.35 C	17.45 C	17.60 C	16.20 C
	7.79	7.60	7.00	6.08	4.79
-0.83	17.80 C	17.51 C	17.54 C	17.70 C	16.26 C
	7.61	7.42	6.82	5.90	4.63
-0.82	17.94 C	17.65 C	17.66 C	17.81 C	16.31 C
	7.43	7.25	6.65	5.72	4.46
-0.81	18.09 C	17.80 C	17.80 C	17.91 C	16.36 C
	7.25	7.07	6.47	5.54	4.30
-0.80	18.22 C	17.94 C	17.92 C	18.02 C	16.41 C
	7.07	6.89	6.29	5.36	4.14
-0.79	18.41 C	18.09 C	18.04 C	18.17 C	16.47 C
	6.89	6.71	6.11	5.18	3.97
-0.78	18.58 C	18.28 C	18.15 C	18.28 C	16.51 C
	6.70	6.53	5.93	4.99	3.81
-0.77	18.79 C	18.45 C	18.28 C	18.39 C	16.53 C
	6.51	6.34	5.74	4.81	3.64
-0.76	18.95 C	18.63 C	18.42 C	18.46 C	16.54 C
	6.32	6.16	5.56	4.63	3.48
-0.75	19.17 C	18.82 C	18.53 C	18.52 C	16.53 C
	6.13	5.97	5.37	4.44	3.31
-0.74	19.35 C	19.02 C	18.65 C	18.54 C	16.51 C
	5.94	5.78	5.19	4.26	3.15
-0.73	19.55 C	19.23 C	18.78 C	18.55 C	16.46 C
	5.74	5.58	5.00	4.07	2.98
-0.72	19.75 C	19.44 C	18.89 C	18.53 C	16.39 C
	5.55	5.39	4.81	3.89	2.82
-0.71	19.96 C	19.64 C	19.02 C	18.50 C	16.25 C
	5.35	5.19	4.62	3.70	2.65
-0.70	20.19 C	19.84 C	19.13 C	18.43 C	16.11 C
	5.14	5.00	4.43	3.52	2.49
-0.69	20.40 C	20.04 C	19.24 C	18.37 C	15.96 C
	4.94	4.80	4.24	3.33	2.33
-0.68	20.61 C	20.23 C	19.34 C	18.28 C	15.79 C
	4.73	4.59	4.04	3.15	2.18
-0.67	20.83 C	20.43 C	19.43 C	18.14 C	15.59 C
	4.53	4.39	3.85	2.97	2.02
-0.66	21.10 C	20.62 C	19.53 C	18.00 C	15.37 C
	4.31	4.18	3.65	2.79	1.866
-0.65	21.35 C	20.82 C	19.60 C	17.83 C	15.11 C
	4.10	3.97	3.46	2.611	1.715

cont.

Capacity and charge in NaF-solutions (cont.)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-0.64	21.60 C	21.01 C	19.67 C	17.63 C	14.78 C
	3.88	3.76	3.26	2.435	1.568
-0.63	21.85 C	21.23 C	19.75 C	17.43 C	14.38 C
	3.67	3.55	3.06	2.260	1.424
-0.62	22.11 C	21.45 C	19.83 C	17.20 C	13.92 C
	3.46	3.34	2.87	2.088	1.285
-0.61	22.39 C	21.73 C	19.90 C	16.98 C	13.41 C
	3.22	3.12	2.666	1.919	1.150
-0.60	22.61 C	21.93 C	19.98 C	16.62 C	12.78 C
	3.00	2.90	2.467	1.752	1.023
-0.59	22.84 C	22.14 C	20.04 C	16.28 C	12.10 C
	2.767	2.679	2.266	1.590	0.902
-0.58	23.10 C	22.35 C	20.10 C	15.91 C	11.26 C
	2.536	2.456	2.065	1.430	0.789
-0.57	23.36 C	22.62 C	20.16 C	15.52 C	10.52 C
	2.302	2.230	1.864	1.275	0.684
-0.56	23.61 C	22.87 C	20.20 C	15.12 C	9.80 C
	2.066	2.001	1.662	1.124	0.586
-0.55	23.89 C	23.14 C	20.26 C	14.83 C	9.15 C
	1.827	1.770	1.459	0.976	0.494
-0.54	24.12 C	23.43 C	20.29 C	14.46 C	8.36 C
	1.586	1.535	1.256	0.831	0.411
-0.53	24.40 C	23.64 C	20.34 C	14.11 C	7.76 C
	1.342	1.299	1.053	0.690	0.333
-0.52	24.63 C	23.86 C	20.36 C	13.81 C	7.16 C
	1.096	1.060	0.849	0.552	0.262
-0.51	24.88 C	24.11 C	20.41 C	13.57 C	6.74 C
	0.847	0.819	0.645	0.416	0.194
-0.50	25.10 C	24.32 C	20.46 C	13.37 C	6.40 C
	0.596	0.576	0.440	0.283	0.130
-0.49	25.31 C	24.48 C	20.52 C	13.19 C	6.15 C
	-0.343	-0.331	-0.235	-0.151	-0.0687
-0.48	25.53 C	24.67 C	20.60 C	13.08 C	6.03 C
	-0.0875	-0.0841	-0.0291	-0.0198	-0.0084
-0.47	25.74 C	24.86 C	20.70 C	13.06 C	6.00 C
	+0.170	+0.164	+0.178	+0.111	+0.0516
-0.46	25.97 C	25.05 C	20.82 C	13.17 C	6.08 C
	0.430	0.415	0.386	0.242	0.112
-0.45	26.13 C	25.23 C	20.92 C	13.32 C	6.26 C
	0.691	0.667	0.595	0.376	0.175
-0.44	26.29 C	25.42 C	21.10 C	13.62 C	6.52 C
	0.954	0.921	0.806	0.512	0.240
-0.43	26.45 C	25.56 C	21.23 C	13.98 C	6.90 C
	1.218	1.177	1.019	0.652	0.309
-0.42	26.61 C	25.75 C	21.38 C	14.33 C	7.46 C
	1.484	1.434	1.232	0.795	0.384
-0.41	26.74 C	25.98 C	21.58 C	14.70 C	8.22 C
	1.752	1.694	1.448	0.942	0.466
-0.40	26.90 C	26.14 C	21.79 C	15.04 C	8.34 C
	2.021	1.955	1.666	1.092	0.554
-0.39	27.03 C	26.31 C	21.99 C	15.60 C	9.50 C
	2.291	2.218	1.886	1.248	0.649

cont.

Capacity and charge in NaF-solutions (cont.)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-0.38	27.18 C	26.44 C	22.20 C	16.14 C	10.21 C
	2.563	2.483	2.108	1.410	0.752
-0.37	27.31 C	26.62 C	22.45 C	16.56 C	11.08 C
	2.84	2.749	2.332	1.575	0.862
-0.36	27.42 C	26.78 C	22.70 C	17.20 C	12.00 C
	3.11	3.017	2.560	1.747	0.982
-0.35	27.52 C	26.92 C	22.96 C	17.79 C	12.75 C
	3.39	3.29	2.789	1.925	1.110
-0.34	27.58 C	26.98 C	23.21 C	18.35 C	13.60 C
	3.66	3.56	3.02	2.109	1.246
-0.33	27.65 C	27.04 C	23.42 C	18.90 C	14.40 C
	3.94	3.83	3.26	2.30	1.390
-0.32	27.71 C	27.11 C	23.67 C	19.41 C	15.15 C
	4.21	4.10	3.49	2.49	1.541
-0.31	27.77 C	27.17 C	23.93 C	19.98 C	15.89 C
	4.49	4.37	3.73	2.69	1.700
-0.30	27.82 C	27.22 C	24.15 C	20.40 C	16.63 C
	4.77	4.64	3.97	2.90	1.866
-0.29	27.86 C	27.24 C	24.38 C	20.77 C	17.40 C
	5.05	4.91	4.22	3.10	2.04
-0.28	27.90 C	27.26 C	24.61 C	21.22 C	18.00 C
	5.33	5.19	4.46	3.32	2.22
-0.27	27.94 C	27.29 C	24.79 C	21.63 C	18.65 C
	5.61	5.46	4.71	3.53	2.41
-0.26	27.97 C	27.33 C	24.94 C	22.03 C	19.20 C
	5.89	5.73	4.96	3.75	2.60
-0.25	27.99 C	27.36 C	25.10 C	22.39 C	19.70 C
	6.17	6.01	5.21	3.98	2.80
-0.24	28.00 C	27.42 C	25.25 C	22.71 C	20.20 C
	6.45	6.28	5.46	4.20	3.00
-0.23	28.01 C	27.45 C	25.40 C	23.03 C	20.65 C
	6.73	6.55	5.72	4.43	3.20
-0.22	28.02 C	27.47 C	25.55 C	23.32 C	21.09 C
	7.01	6.83	5.97	4.67	3.42
-0.21	28.05 C	27.51 C	25.68 C	23.61 C	21.41 C
	7.29	7.10	6.23	4.90	3.63
-0.20	28.11 C	27.53 C	25.81 C	23.85 C	21.73 C
	7.57	7.38	6.49	5.14	3.85
-0.19	28.14 C	27.56 C	25.92 C	24.09 C	22.10 C
	7.85	7.66	6.75	5.38	4.07
-0.18	28.20 C	27.60 C	26.06 C	24.30 C	22.45 C
	8.13	7.93	7.01	5.63	4.29
-0.17	28.27 C	27.64 C	26.17	24.51 C	22.71 C
	8.42	8.21	7.27	5.87	4.52
-0.16	28.33 C	27.71 C	26.27 C	24.70 C	22.99 C
	8.70	8.48	7.53	6.12	4.75
-0.15	28.43 C	27.76 C	26.40 C	24.88 C	23.18 C
	8.98	8.76	7.80	6.37	4.98
-0.14	28.51 C	27.82 C	26.52 C	25.00 C	23.31 C
	9.27	9.04	8.06	6.62	5.21
-0.13	28.61 C	27.88 C	26.64 C	25.12 C	23.48 C
	9.55	9.32	8.33	6.87	5.45
-0.12	28.71 C	27.98 C	26.76 C	25.28 C	23.61 C
	9.84	9.60	8.60	7.12	5.69

cont.

Capacity and charge in NaF-solutions (cont.)

E Volts	0.916 N	0.66 N	0.10 N	0.01 N	0.001 N
-0.11	28.32 C	28.05 C	26.87 C	25.41 C	23.72 C
	10.13	9.38	8.86	7.37	5.92
-0.10	28.95 C	28.15 C	26.98 C	25.57 C	23.83 C
	10.42	10.16	9.13	7.63	6.16
-0.09	29.08 C	28.25 C	27.09 C	25.70 C	23.92 C
	10.71	10.44	9.40	7.89	6.40
-0.08	29.22 C	28.36 C	27.18 C	25.81 C	24.01 C
	11.00	10.73	9.68	8.14	6.64
-0.07	29.39 C	28.49 C	27.29 C	25.94 C	24.10 C
	11.30	11.01	9.95	8.40	6.88
-0.06	29.53 C	28.64 C	27.39 C	26.08 C	24.19 C
	11.59	11.30	10.22	8.67	7.12
-0.05	29.70 C	28.76 C	27.49 C	26.20 C	24.26 C
	11.89	11.59	10.50	8.93	7.37
-0.04	29.90 C	28.92 C	27.62 C	26.32 C	24.33 C
	12.19	11.88	10.77	9.19	7.61
-0.03	30.11 C	29.06 C	27.74 C	26.43 C	24.40 C
	12.49	12.17	11.05	9.45	7.85
-0.02	30.31 C	29.25 C	27.87 C	26.55 C	24.47 C
	12.79	12.46	11.33	9.72	8.10
-0.01	30.58 C	29.44 C	28.00 C	26.68 C	24.53 C
	13.10	12.75	11.61	9.99	8.34
0.00	30.90 C	29.65 C	28.15 C	26.80 C	24.59 C
	13.41	13.05	11.89	10.25	8.59
+0.01	31.32 C	29.85 C	28.30 C	26.93 C	24.63 C
	13.72	13.35	12.18	10.52	8.83
+0.02	31.88 C	30.05 C	28.46 C	27.09 C	24.71 C
	14.04	13.65	12.46	10.80	9.08
+0.03	32.40 C	30.25 C	28.64 C	27.24 C	24.79 C
	14.36	13.95	12.75	11.07	9.33
+0.04	32.70 C	30.56 C	28.81 C	27.43 C	24.89 C
	14.69	14.26	13.03	11.34	9.58
+0.05	33.59 C	30.95 C	29.02 C	27.60 C	24.99 C
	15.02	14.57	13.32	11.62	9.83
+0.06	34.81 C	31.42 C	29.22 C	27.79 C	25.10 C
	15.37	14.88	13.62	11.90	10.08
+0.07	36.62 C	31.76 C	29.34 C	28.00 C	25.23 C
	15.74	15.20	13.91	12.18	10.33
+0.08	39.55 C	32.23 C	29.68 C	28.20 C	25.40 C
	16.13	15.52	14.21	12.46	10.59
+0.09	43.99 C	32.65 C	29.93 C	28.40 C	25.58 C
	16.57	15.85	14.51	12.74	10.84
+0.10	50.2 C	33.14 C	30.17 C	28.65 C	25.82 C
	17.08	16.18	14.81	13.03	11.10
+0.11	57.6 C	33.94 C	30.44 C	28.89 C	26.03 C
	17.65	16.52	15.11	13.32	11.36
+0.12	64.8 C	34.72 C	30.73 C	29.15 C	26.32 C
	18.30	16.86	15.42	13.61	11.62
+0.13	70.9 C	35.64 C	31.00 C	29.42 C	26.61 C
	19.01	17.22	15.73	13.90	11.89
+0.14	74.8 C		31.41 C	29.73 C	26.92 C
	19.76		16.04	14.20	12.16

cont.

Capacity and charge in NaF-solutions (cont.)

$\frac{E}{\text{Volts}}$	0.916 N	0.10 N	0.01 N	0.001 N
+0.15	78.4 C	31.77 C	30.04 C	27.21 C
	20.54	16.36	14.50	12.43
+0.16	82.6 C	32.22 C	30.40 C	27.59 C
	21.37	16.68	14.80	12.71
+0.17	89.2 C	32.73 C	30.79 C	27.90 C
	22.26	17.01	15.11	12.99
+0.18		33.19 C	31.19 C	28.21 C
		17.34	15.42	13.27
+0.19		33.80 C	31.58 C	28.60 C
		17.68	15.74	13.55
+0.20		34.51 C	32.02 C	28.98 C
		18.03	16.06	13.84
+0.21			32.55 C	29.41 C
			16.39	14.14
+0.22			33.09 C	29.91 C
			16.72	14.44
+0.23			33.77 C	30.44 C
			17.05	14.74
+0.24			34.36 C	31.02 C
			17.40	15.05
+0.25			35.15 C	31.70 C
			17.75	15.37
+0.26			36.00 C	32.32 C
			18.11	15.69
+0.27			37.00 C	33.10 C
			18.48	16.02
+0.28			38.00 C	34.04 C

Differential capacity of mercury in 0.1 N solutions of potassium salts. T = 25°C. Potentials measured relative to normal calomel electrode corrected for liquid junction potential between 0.1 N KCl and the solution named.

Reference: D.C. Grahame, M.A. Poth and J.I. Cummings, J. Amer, Chem. Soc., 74, 4422 (1952)

E Volts	KCl	KCNS	KOH	KAc	KBr	K_2CO_3
0.19				96.01		
0.18				83.30		
0.17				74.63		
0.16				68.51		
0.15				62.72		
0.14				58.00		
0.13				54.15		
0.12				50.88		
0.11				48.25		
0.10				45.91		
0.09				43.75		
0.08				41.88		
0.07				40.38		
0.06				39.05		
0.05				37.82		
0.04				36.60		
0.03				35.48		
0.02				34.50		
0.01				33.65		
0.00	123.1			32.98		
-0.01	97.4			32.25		
-0.02	81.38			31.60		
-0.03	71.19			31.05		
-0.04	63.91			30.57		
-0.05	58.50			30.15		
-0.06	55.20			29.80		
-0.07	51.95			29.45		
-0.08	49.40			29.13		
-0.09	47.45			28.83		57.31
-0.10	45.80			28.60		55.80
-0.11	44.39			28.40		54.95
-0.12	43.16			28.27		54.30
-0.13	42.11			28.19		53.83
-0.14	41.24		109.68	28.16		53.25
-0.15	40.48		87.75	28.11		52.68
-0.16	39.84	109.19	73.11	28.11		52.00
-0.17	39.31	100.65	64.78	28.13		51.76
-0.18	38.87	94.03	57.96	28.15		51.10
-0.19	38.54	85.47	52.72	28.21		50.35
-0.20	38.30	80.38	48.49	28.29	121.80	49.44
-0.21	38.11	77.26	44.91	28.37	108.70	47.75
-0.22	38.02	73.75	41.82	28.45	98.10	46.70
-0.23	38.00	69.08	39.15	28.53	89.45	45.35
-0.24	38.06	64.93	36.95	28.60	82.75	44.08
-0.25	38.16	61.62	35.07	28.70	77.28	42.82
-0.26	38.26	59.40	33.47	28.73	72.40	41.32

cont.

Capacity in potassium salts (cont.)

E Volts	KCl	KCNS	KOH	KAc	KBr	K_2CO_3
-0.27	38.42	57.32	32.03	28.79	68.25	39.50
-0.28	38.59	55.50	30.81	28.83	64.85	38.10
-0.29	38.77	54.11	29.80	28.89	62.00	36.61
-0.30	38.94	52.15	28.90	28.89	59.50	35.12
-0.31	39.19	50.50	28.15	28.83	57.38	33.80
-0.32	39.46	49.20	27.55	28.76	55.55	32.79
-0.33	39.58	47.98	27.00	28.68	53.95	31.70
-0.34	39.72	46.91	26.55	28.58	52.60	30.78
-0.35	39.86	45.92	26.00	28.49	51.50	29.94
-0.36	39.84	45.03	25.70	28.33	50.80	29.15
-0.37	39.76	44.25	25.33	28.23	50.14	28.40
-0.38	39.60	43.60	24.97	27.85	49.45	27.70
-0.39	39.38	43.05	24.60	27.58	49.10	26.83
-0.40	39.01	42.55	24.30	27.35	48.82	26.23
-0.41	38.53	42.20	23.98	27.00	48.65	25.60
-0.42	37.95	41.87	23.68	26.70	48.52	24.90
-0.43	37.36	41.55	23.40	26.38	48.55	24.30
-0.44	36.54	41.30	23.13	26.00	48.60	23.72
-0.45	35.62	41.08	22.91	25.70	48.60	23.10
-0.46	34.64	40.80	22.68	25.30	48.50	22.65
-0.47	33.67	40.68	22.47	25.00	48.33	22.20
-0.48	32.64	40.59	22.27	24.60	48.15	21.85
-0.49	31.57	40.53	22.07	24.25	48.10	21.54
-0.50	30.50	40.49	21.93	23.98	47.75	21.30
-0.51	29.52	40.45	21.75	23.65	47.31	21.05
-0.52	28.50	40.42	21.62	23.40	46.79	20.82
-0.53	27.57	40.40	21.47	23.05	46.15	20.61
-0.54	26.74	40.39	21.32	22.83	45.45	20.48
-0.55	25.91	40.35	21.21	22.57	44.52	20.35
-0.56	25.23	40.29	21.10	22.42	43.60	20.28
-0.57	24.59	40.19	20.98	22.20	42.39	20.20
-0.58	23.98	40.05	20.89	22.00	41.21	20.13
-0.59	23.47	39.82	20.80	21.77	39.95	20.10
-0.60	23.00	39.65	20.70	21.60	38.60	20.02
-0.61	22.56	39.48	20.62	21.43	37.20	19.98
-0.62	22.17	39.12	20.51	21.22	35.69	19.95
-0.63	21.83	38.87	20.41	21.03	34.23	19.90
-0.64	21.50	38.43	20.31	20.83	32.82	19.87
-0.65	21.23	38.02	20.20	20.73	31.38	19.80
-0.66	20.96	37.55	20.09	20.52	30.15	19.74
-0.67	20.70	37.00	19.97	20.38	28.91	19.68
-0.68	20.46	36.50	19.85	20.22	27.80	19.60
-0.69	20.24	35.73	19.73	20.05	26.66	19.52
-0.70	20.04	35.00	19.61	19.90	25.70	19.41
-0.71	19.86	34.20	19.47	19.72	24.80	19.32
-0.73	19.52	32.74	19.22	19.43	23.25	19.13
-0.75	19.18	31.13	18.97	19.18	22.00	18.92
-0.77	18.84	29.20	18.70	18.90	20.98	18.70
-0.79	18.54	27.47	18.45	18.63	20.11	18.43
-0.80	18.40	26.69	18.33	18.50	19.80	18.30
-0.85	17.72	23.02	17.70	17.80	18.45	17.68
-0.90	17.15	20.20	17.12	17.25	17.52	17.15
-0.95	16.74	18.45	16.64	16.80	16.90	16.73

cont.

Capacity in potassium salts (cont.)

E Volts	KCl	KCNS	KOH	KAc	KBr	K_2CO_3
-1.00	16.40	17.35	16.37	16.42	16.50	16.40
-1.05	16.17	16.65	16.17	16.18	16.28	16.18
-1.10	16.06	16.32	16.07	16.06	16.15	16.06
-1.15	16.04	16.20	16.04	16.01	16.10	16.05
-1.20	16.08	16.17	16.07	16.07	16.15	16.11
-1.25	16.22	16.21	16.17	16.18	16.25	16.23
-1.30	16.38	16.38	16.37	16.35	16.45	16.40
-1.35	16.69	16.59	16.61	16.60	16.68	16.66
-1.40	17.00	16.88	16.95	16.91	17.00	16.98
-1.45	17.40	17.22	17.33	17.30	17.35	17.32
-1.50	17.80	17.65	17.75	17.75	17.80	17.80
-1.55	18.32	18.12	18.20	18.22	18.29	18.23
-1.60	18.83	18.65	18.75	18.72	18.85	18.79
-1.65	19.44	19.25	19.37	19.34	19.45	19.38
-1.70	20.05	19.80	20.07	20.02	20.15	20.01
-1.75	20.81	20.52	20.84	20.80	20.90	20.79
-1.80	21.77	21.45	21.79	21.72	21.86	21.65
-1.85	22.98	22.54	22.95	22.82	23.01	22.62
-1.90	24.46	23.32	24.35	24.30	24.46	
-1.95			27.57			
Liq. Junct. Pot. mv.	0.00	1.6	-15.3	6.9	-0.4	5.8

Differential capacity of mercury in 0.1 N solutions of potassium salts. T = 25°C. Potentials measured relative to normal calomel electrode corrected for liquid junction potential between 0.1 N KCl and the solution named.

Reference : D.C. Grahame, M.A. Poth and J.I. Cummings, J. Amer. Chem. Soc., 74, 4422 (1952)

E Volts	$KClO_4$	KF	$KHCO_3$	KI	KNO_3	K_2SO_4
0.19	30.68	40.32			38.89	48.08
0.18	29.72	38.74			36.63	46.99
0.17	28.91	37.45			34.80	45.94
0.16	28.25	36.32			33.42	45.02
0.15	27.65	35.37			32.37	44.20
0.14	27.20	34.61			31.61	43.48
0.13	26.78	33.96			30.90	42.80
0.12	26.45	33.36			30.22	42.18
0.11	26.13	32.79			29.55	41.57
0.10	25.85	32.24			28.91	41.01
0.09	25.56	31.73			28.33	40.50
0.08	25.29	31.24			27.78	40.02
0.07	25.01	30.76			27.24	39.59
0.06	24.78	30.31			26.76	39.18
0.05	24.52	29.93			26.30	38.80
0.04	24.29	29.61			25.91	38.48
0.03	24.11	29.37			25.56	38.20
0.02	23.97	29.20			25.25	37.93
0.01	23.81	29.05	89.58		25.00	37.75
0.00	23.70	28.94	75.09		24.80	37.50
-0.01	23.67	28.83	65.89		24.55	37.40
-0.02	23.52	28.72	60.42		24.38	37.25
-0.03	23.40	28.59	56.50		24.21	37.19
-0.04	23.32	28.47	53.38		24.08	37.10
-0.05	23.25	28.33	50.80		23.97	37.05
-0.06	23.20	28.21	48.50		23.88	37.02
-0.07	23.16	28.14	46.45		23.78	37.02
-0.08	23.12	27.90	44.60		23.71	37.09
-0.09	23.11	27.75	43.09		23.67	37.18
-0.10	23.12	27.59	41.89		23.61	37.25
-0.11	23.14	27.43	40.40		23.60	37.35
-0.12	23.19	27.28	39.10		23.61	37.45
-0.13	23.22	27.14	38.14		23.65	37.58
-0.14	23.28	27.00	37.32		23.70	37.70
-0.15	23.35	26.88	36.65		23.76	37.80
-0.16	23.43	26.77	36.02		23.86	37.83
-0.17	23.53	26.65	35.52		24.02	37.90
-0.18	23.67	26.53	35.02		24.18	37.97
-0.19	23.81	26.42	34.53		24.35	37.93
-0.20	24.00	26.32	34.13		24.52	37.83
-0.21	24.19	26.20	34.00		24.72	37.72
-0.22	24.40	26.03	33.87		24.94	37.55
-0.23	24.60	25.87	33.79		25.19	37.45
-0.24	24.80	25.71	33.72		25.42	37.05
-0.25	25.01	25.52	33.65		25.70	36.53
-0.26	25.28	25.35	33.57		26.00	36.03

cont.

Capacity in potassium salts (cont.)

E / Volts	$KClO_4$	KF	$KHCO_3$	KI	KNO_3	K_2SO_4
-0.27	25.52	25.17	33.48		26.35	35.53
-0.28	25.77	24.97	33.32		26.69	35.00
-0.29	26.03	24.77	33.19		27.02	34.49
-0.30	26.25	24.57	33.03		27.40	33.68
-0.31	26.55	24.33	32.83		27.70	32.90
-0.32	26.80	24.12	32.62		28.09	32.10
-0.33	27.05	23.88	32.32		28.42	31.50
-0.34	27.30	23.63	31.93		28.75	30.65
-0.35	27.53	23.40	31.45		29.05	30.00
-0.36	27.72	23.20	30.93		29.39	29.15
-0.37	27.93	22.97	30.33		29.66	28.50
-0.38	28.10	22.75	29.65		29.93	27.70
-0.39	28.27	22.65	28.95		30.20	26.95
-0.40	28.42	22.45	28.30		30.45	26.30
-0.41	28.51	22.13	27.65		30.59	25.55
-0.42	28.57	21.88	27.00		30.66	24.85
-0.43	28.58	21.65	26.35		30.69	24.30
-0.44	28.57	21.40	25.75		30.68	23.70
-0.45	28.51	21.15	25.00		30.65	23.23
-0.46	28.35	20.95	24.32		30.52	22.70
-0.47	28.15	20.83	23.77		30.32	22.15
-0.48	27.95	20.71	23.28		30.11	21.70
-0.49	27.70	20.62	22.85		29.90	21.25
-0.50	27.45	20.52	22.50		29.65	20.85
-0.51	27.10	20.47	22.18		29.22	20.58
-0.52	26.77	20.42	21.90	117.0	28.75	20.40
-0.53	26.40	20.38	21.65	109.7	28.30	20.23
-0.54	26.00	20.36	21.43	104.1	27.75	20.12
-0.55	25.57	20.34	21.23	99.0	27.20	20.00
-0.56	25.18	20.31	21.05	94.2	26.60	19.95
-0.57	24.73	20.28	20.38	90.0	26.00	19.90
-0.58	24.33	20.23	20.75	86.25	25.52	19.87
-0.59	23.92	20.20	20.62	82.96	25.00	19.85
-0.60	23.52	20.15	20.50	80.13	24.50	19.82
-0.61	23.13	20.12	20.38	77.75	24.00	19.80
-0.62	22.78	20.08	20.28	75.45	23.52	19.75
-0.63	22.43	20.03	20.18	73.43	23.06	19.72
-0.64	22.08	19.98	20.07	71.70	22.69	19.68
-0.65	21.80	19.91	19.99	70.05	22.30	19.62
-0.66	21.50	19.84	19.88	68.58	21.95	19.58
-0.67	21.20	19.77	19.77	67.25	21.64	19.52
-0.68	20.95	19.69	19.66	66.05	21.32	19.45
-0.69	20.70	19.61	19.55	64.95	21.01	19.40
-0.70	20.48	19.52	19.47	63.87	20.75	19.32
-0.71	20.25	19.42	19.33	62.70	20.50	19.22
-0.73	19.83	19.20	19.11	60.30	20.00	19.02
-0.75	19.57	18.93	18.88	57.68	19.55	18.80
-0.77	19.07	18.68	18.63	54.75	19.24	18.60
-0.79	18.73	18.42	18.38	51.15	18.76	18.38
-0.80	18.58	18.28	18.25	49.35	18.60	18.28
-0.85	17.83	17.71	17.60	39.04	17.85	17.68
-0.90	17.25	17.18	17.10	29.62	17.25	17.15
-0.95	16.80	16.74	16.60	23.11	16.75	16.70

cont.

Capacity in potassium salts (cont.)

E Volts	$KClO_4$	KF	$KHCO_3$	KI	KNO_3	K_2SO_4
-1.00	16.45	16.39	16.28	19.45	16.40	16.38
-1.05	16.18	16.13	16.08	17.55	16.16	16.15
-1.10	16.01	16.00	15.98	16.72	16.04	16.01
-1.15	16.00	15.98	15.95	16.32	16.04	16.00
-1.20	16.05	16.03	16.00	16.22	16.10	16.01
-1.25	16.20	16.18	16.11	16.29	16.23	16.14
-1.30	16.39	16.35	16.28	16.46	16.43	16.35
-1.35	16.63	16.60	16.51	16.70	16.70	16.55
-1.40	16.95	16.90	16.80	17.02	17.02	16.85
-1.45	17.32	17.28	17.17	17.38	17.40	17.22
-1.50	17.76	17.72	17.59	17.80	17.83	17.68
-1.55	18.25	18.21	18.06	18.25	18.33	18.15
-1.60	18.80	18.75	18.60	18.78	18.87	18.68
-1.65	19.43	19.37	19.20	19.40	19.45	19.28
-1.70	20.11	20.05	19.86	20.10	20.08	19.95
-1.75	20.88	20.82	20.62	20.92	20.75	20.70
-1.80	21.70	21.70	21.48	21.83		21.60
-1.85	22.80	22.80	22.53	22.90		22.70
-1.90	24.43	24.20	23.90	24.45		24.25
-1.95	27.40	27.10	26.15	27.53		
Liq. Junct. Pot. mv.	1.5	4.1	6.1	0.0	0.9	6.1

Differential capacity on mercury in KI. T = 25°C. Concentration in mole/l. Potential with respect to normal KCl calomel electrode dipping into 0.1 M KI (aq.).

Reference: D.C. Grahame, J. Amer. Chem. Soc., 80, 4201 (1958)

conc.	0.0150		0.0250		0.0416		0.0500	
σ	E	C	E	C	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
18			-0.466	131.1				
17			-0.474	123.8				
16			-0.482	116.9				
15			-0.491	110.6	-0.507	108.6		
14			-0.500	104.8	-0.517	103.3	-0.523	103.4
13			-0.510	100.0	-0.527	98.15	-0.533	98.25
12			-0.520	95.18	-0.537	93.32	-0.543	93.44
11	-0.510	90.77	-0.531	90.06	-0.548	89.00	-0.554	88.97
10	-0.521	86.15	-0.542	85.32	-0.559	84.99	-0.566	84.87
9	-0.533	81.86	-0.554	81.07	-0.572	81.30	-0.578	81.13
8	-0.546	77.94	-0.567	77.57	-0.584	77.89	-0.590	77.79
7	-0.559	74.41	-0.580	74.49	-0.597	74.79	-0.604	74.73
6	-0.572	71.36	-0.594	71.61	-0.611	72.05	-0.617	71.97
5	-0.587	68.52	-0.608	68.99	-0.625	69.58	-0.631	69.57
4	-0.602	65.77	-0.623	66.59	-0.640	67.33	-0.646	67.42
3	-0.617	63.05	-0.638	64.31	-0.655	65.26	-0.661	65.37
2	-0.633	60.28	-0.654	61.94	-0.670	63.08	-0.677	63.26
1	-0.650	57.28	-0.670	59.28	-0.686	60.72	-0.693	61.00
0	-0.668	53.88	-0.688	56.15	-0.703	58.05	-0.709	58.47
- 1	-0.688	49.77	-0.706	52.43	-0.721	54.92	-0.727	55.52
- 2	-0.709	44.83	-0.726	48.03	-0.740	51.22	-0.746	52.06
- 3	-0.733	39.11	-0.748	42.95	-0.760	46.82	-0.766	47.85
- 4	-0.760	32.80	-0.773	37.18	-0.783	41.59	-0.788	42.88
- 5	-0.794	26.89	-0.802	31.07	-0.809	35.91	-0.812	37.22
- 6	-0.835	22.29	-0.838	25.61	-0.839	30.15	-0.842	31.53
- 7	-0.884	19.40	-0.881	21.32	-0.876	24.77	-0.877	26.07
- 8	-0.938	17.78	-0.931	18.62	-0.920	20.73	-0.919	21.62
- 9	-0.995	16.93	-0.988	17.07	-0.972	18.16	-0.969	18.61
-10	-1.055	16.44	-1.047	16.52	-1.029	16.91	-1.025	17.11
-11	-1.117	16.12	-1.108	16.18	-1.089	16.37	-1.085	16.45
-12	-1.179	16.01	-1.170	16.04	-1.151	16.13	-1.146	16.18
-13	-1.242	16.07	-1.233	16.07	-1.213	16.08	-1.209	16.10
-14	-1.303	16.30	-1.295	16.27	-1.275	16.23	-1.270	16.23
-15	-1.364	16.61	-1.356	16.59	-1.336	16.50	-1.332	16.49
-16	-1.424	16.99	-1.415	16.96	-1.396	16.88	-1.392	16.86
-17	-1.482	17.40	-1.473	17.39	-1.455	17.29	-1.450	17.28
-18	-1.539	17.86	-1.530	17.85	-1.512	17.76	-1.507	17.75

cont.

Capacity in potassium iodide (cont.)

conc.	0.0600		0.0833		0.1000		0.1200	
σ	E	C	E	C	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
18								
17							-0.524	115.8
16							-0.533	110.2
15							-0.542	104.9
14	-0.529	103.2					-0.552	99.97
13	-0.539	98.12					-0.562	95.50
12	-0.550	93.31	-0.561	92.21	-0.567	91.87	-0.573	91.51
11	-0.561	88.80	-0.572	88.06	-0.578	87.83	-0.584	87.42
10	-0.572	84.67	-0.583	84.15	-0.590	83.88	-0.596	83.38
9	-0.584	80.97	-0.596	80.51	-0.602	80.20	-0.608	79.69
8	-0.597	77.67	-0.608	77.18	-0.614	76.86	-0.621	76.37
7	-0.610	74.60	-0.622	74.15	-0.628	73.87	-0.634	73.49
6	-0.624	71.89	-0.635	71.47	-0.642	71.25	-0.648	70.93
5	-0.638	69.51	-0.650	69.22	-0.656	69.02	-0.663	68.68
4	-0.652	67.43	-0.664	67.16	-0.671	66.99	-0.677	66.72
3	-0.668	65.42	-0.679	65.22	-0.686	65.19	-0.693	65.03
2	-0.683	63.38	-0.695	63.42	-0.701	63.48	-0.708	63.36
1	-0.699	61.23	-0.711	61.53	-0.717	61.65	-0.724	61.56
0	-0.716	58.86	-0.727	59.37	-0.734	59.49	-0.741	59.56
- 1	-0.733	56.12	-0.745	56.72	-0.751	56.98	-0.758	57.27
- 2	-0.752	52.81	-0.763	53.64	-0.769	54.19	-0.776	54.57
- 3	-0.771	48.81	-0.782	50.18	-0.788	50.82	-0.795	51.36
- 4	-0.793	44.04	-0.803	45.86	-0.808	46.66	-0.815	47.54
- 5	-0.817	38.71	-0.826	40.80	-0.831	41.96	-0.837	43.10
- 6	-0.845	33.03	-0.852	35.40	-0.856	36.81	-0.862	38.09
- 7	-0.878	27.39	-0.883	29.82	-0.886	31.31	-0.890	32.70
- 8	-0.918	22.59	-0.920	24.59	-0.921	25.94	-0.923	27.33
- 9	-0.967	19.20	-0.964	20.59	-0.963	21.54	-0.963	22.69
-10	-1.022	17.38	-1.017	18.04	-1.014	18.58	-1.012	19.30
-11	-1.081	16.58	-1.074	16.58	-1.070	17.11	-1.066	17.47
-12	-1.142	16.25	-1.135	16.39	-1.130	16.50	-1.125	16.66
-13	-1.204	16.12	-1.196	16.21	-1.191	16.26	-1.186	16.34
-14	-1.266	16.23	-1.258	16.25	-1.252	16.27	-1.247	16.30
-15	-1.327	16.48	-1.319	16.47	-1.313	16.48	-1.308	16.49
-16	-1.387	16.84	-1.379	16.82	-1.373	16.81	-1.368	16.77
-17	-1.445	17.27	-1.438	17.26	-1.432	17.25	-1.427	17.25
-18	-1.502	17.74	-1.495	17.73	-1.489	17.71	-1.484	17.71

cont.

Capacity in potassium iodide (cont.)

conc.	0.2 6		0.250		0.300		0.416	
σ	E	C	E	C	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
18					-0.546	115.0		
17					-0.555	109.8		
16					-0.564	105.0		
15	-0.561	102.1	-0.568	101.4	-0.574	100.5	-0.583	98.36
14	-0.571	98.07	-0.578	97.06	-0.584	95.93	-0.594	93.84
13	-0.582	93.79	-0.589	92.34	-0.594	91.36	-0.605	89.53
12	-0.593	89.17	-0.600	87.84	-0.606	87.04	-0.617	85.45
11	-0.604	84.93	-0.611	83.80	-0.617	83.02	-0.629	81.62
10	-0.616	81.09	-0.624	80.10	-0.630	79.35	-0.641	78.12
9	-0.629	77.66	-0.636	76.75	-0.643	76.03	-0.654	75.05
8	-0.642	74.54	-0.650	73.72	-0.656	73.10	-0.668	72.22
7	-0.656	71.74	-0.663	71.07	-0.670	70.57	-0.682	69.70
6	-0.670	69.39	-0.678	68.83	-0.684	68.39	-0.696	67.53
5	-0.684	67.42	-0.693	66.93	-0.699	66.49	-0.712	65.67
4	-0.699	65.69	-0.708	65.20	-0.715	64.78	-0.727	63.95
3	-0.715	64.03	-0.723	63.61	-0.730	63.23	-0.743	62.37
2	-0.731	62.54	-0.739	62.10	-0.746	61.74	-0.752	60.94
1	-0.747	60.93	-0.755	60.55	-0.763	60.29	-0.776	59.65
0	-0.763	59.16	-0.772	58.97	-0.779	58.82	-0.793	58.23
- 1	-0.781	57.24	-0.789	57.22	-0.797	57.11	-0.810	56.68
- 2	-0.798	55.11	-0.807	55.05	-0.814	55.09	-0.828	54.89
- 3	-0.817	52.27	-0.826	52.53	-0.833	52.79	-0.846	52.75
- 4	-0.837	49.17	-0.845	49.52	-0.852	49.81	-0.866	50.19
- 5	-0.858	45.39	-0.866	46.03	-0.873	46.62	-0.886	47.24
- 6	-0.881	41.25	-0.889	42.06	-0.895	42.89	-0.908	43.92
- 7	-0.907	36.42	-0.914	37.55	-0.920	38.66	-0.932	40.14
- 8	-0.936	31.27	-0.943	32.65	-0.948	33.88	-0.958	35.83
- 9	-0.971	26.36	-0.976	27.66	-0.979	28.97	-0.988	31.07
-10	-1.013	22.25	-1.015	23.43	-1.017	24.50	-1.023	26.53
-11	-1.061	19.18	-1.062	20.10	-1.061	20.91	-1.064	22.68
-12	-1.116	17.46	-1.115	17.94	-1.112	18.48	-1.111	19.85
-13	-1.175	16.82	-1.172	17.07	-1.169	17.34	-1.164	18.14
-14	-1.235	16.57	-1.231	16.73	-1.227	16.92	-1.221	17.36
-15	-1.295	16.66	-1.291	16.76	-1.286	16.84	-1.279	17.12
-16	-1.355	16.90	-1.350	16.96	-1.346	17.03	-1.337	17.17
-17	-1.413	17.26	-1.409	17.28	-1.404	17.32	-1.395	17.41
-18	-1.470	17.70	-1.466	17.71	-1.461	17.72	-1.452	17.77

cont.

Capacity in potassium iodide (cont.)

conc.	0.500		0.600		0.8305	
σ	E	C	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
18			-0.566	109.5	-0.576	107.2
17			-0.576	104.8	-0.585	102.69
16			-0.585	100.3	-0.595	98.36
15	-0.590	97.11	-0.595	95.97	-0.606	94.12
14	-0.600	92.75	-0.606	91.74	-0.617	90.03
13	-0.611	88.56	-0.617	87.68	-0.628	86.24
12	-0.623	84.57	-0.629	83.84	-0.640	82.58
11	-0.635	80.82	-0.641	80.22	-0.652	78.99
10	-0.648	77.48	-0.654	76.88	-0.665	75.67
9	-0.661	74.43	-0.667	73.82	-0.679	72.65
8	-0.675	71.61	-0.681	71.04	-0.693	69.91
7	-0.689	69.14	-0.696	68.56	-0.707	67.41
6	-0.704	66.98	-0.710	66.38	-0.722	65.21
5	-0.719	65.08	-0.726	64.46	-0.738	63.34
4	-0.734	63.38	-0.741	62.77	-0.754	61.71
3	-0.750	61.85	-0.758	61.27	-0.770	60.22
2	-0.767	60.47	-0.774	59.96	-0.787	58.85
1	-0.783	59.14	-0.791	58.62	-0.804	57.55
0	-0.800	57.76	-0.808	57.26	-0.822	56.25
- 1	-0.818	56.28	-0.826	55.79	-0.840	54.84
- 2	-0.836	54.54	-0.844	54.11	-0.858	53.32
- 3	-0.855	52.46	-0.863	52.19	-0.877	51.62
- 4	-0.874	50.12	-0.882	50.02	-0.897	49.61
- 5	-0.895	47.42	-0.903	47.47	-0.918	47.44
- 6	-0.916	44.30	-0.925	44.58	-0.939	44.87
- 7	-0.940	40.80	-0.948	41.29	-0.962	41.99
- 8	-0.966	36.66	-0.973	37.42	-0.987	38.65
- 9	-0.995	32.17	-1.002	33.26	-1.014	34.94
-10	-1.028	27.77	-1.034	29.03	-1.045	31.19
-11	-1.067	23.85	-1.071	25.10	-1.079	27.22
-12	-1.112	20.86	-1.114	21.90	-1.118	23.75
-13	-1.163	18.79	-1.162	19.47	-1.163	21.11
-14	-1.218	17.68	-1.216	18.09	-1.213	19.08
-15	-1.275	17.29	-1.272	17.49	-1.267	18.10
-16	-1.333	17.27	-1.329	17.36	-1.323	17.74
-17	-1.391	17.46	-1.387	17.52	-1.379	17.76
-18	-1.448	17.82	-1.443	17.86	-1.435	18.02

cont.

Capacity in potassium iodide (cont.)

conc.	1.000		1.200	
σ C/cm^2	E Volts	C $\mu F/cm^2$	E Volts	C $\mu F/cm^2$
18	-0.581	105.9	-0.586	104.7
17	-0.591	101.5	-0.596	100.3
16	-0.601	97.25	-0.606	96.11
15	-0.611	93.07	-0.617	92.02
14	-0.622	89.10	-0.628	88.13
13	-0.634	85.38	-0.640	84.38
12	-0.646	81.69	-0.652	80.75
11	-0.658	78.15	-0.665	77.30
10	-0.671	74.90	-0.678	74.07
9	-0.685	71.89	-0.692	71.13
8	-0.699	69.14	-0.706	68.43
7	-0.714	66.68	-0.721	65.98
6	-0.729	64.55	-0.736	63.84
5	-0.745	62.68	-0.752	61.94
4	-0.761	61.01	-0.768	60.26
3	-0.778	59.50	-0.785	58.76
2	-0.795	58.13	-0.803	57.43
1	-0.812	56.89	-0.820	56.23
0	-0.830	55.63	-0.838	55.06
- 1	-0.848	54.32	-0.856	53.85
- 2	-0.867	52.91	-0.875	52.57
- 3	-0.886	51.32	-0.895	51.12
- 4	-0.906	49.51	-0.914	49.44
- 5	-0.926	47.42	-0.935	47.49
- 6	-0.948	45.05	-0.957	45.27
- 7	-0.971	42.37	-0.979	42.72
- 8	-0.995	39.31	-1.004	39.83
- 9	-1.022	35.78	-1.030	36.49
-10	-1.051	32.14	-1.059	32.93
-11	-1.085	28.27	-1.091	29.26
-12	-1.122	24.81	-1.127	25.84
-13	-1.165	22.08	-1.168	23.08
-14	-1.213	19.84	-1.214	20.70
-15	-1.266	18.58	-1.264	19.18
-16	-1.320	18.08	-1.318	18.49
-17	-1.376	18.00	-1.372	18.30
-18	-1.431	18.18	-1.427	18.38

Differential capacity on mercury in NaSCN. T = 25°C. Concentration in mole/l. Potentials with respect to a saturated KCl calomel electrode dipping into the working solution.

Reference : R. Parsons and P.C. Symons, Trans. Faraday Soc., 64,
1077 (1968)

conc.	0.001224		0.002136		0.006105		0.01223	
σ C/cm^2	E Volts	C μ F/cm^2	E Volts	C μ F/cm^2	E Volts	C μ F/cm^2	E Volts	C μ F/cm^2
9	-0.223	39.38	-0.250	40.64	-0.267	40.21	-0.290	40.63
8	-0.259	38.31	-0.275	39.80	-0.292	39.05	-0.315	39.43
7	-0.285	37.69	-0.300	39.08	-0.318	38.21	-0.341	38.77
6	-0.312	37.09	-0.326	38.56	-0.344	37.67	-0.367	38.26
5	-0.339	36.58	-0.352	37.97	-0.371	37.11	-0.393	38.02
4	-0.367	35.89	-0.379	36.81	-0.398	36.73	-0.419	37.74
3	-0.395	34.63	-0.406	35.79	-0.426	35.90	-0.446	37.53
2	-0.425	32.42	-0.435	33.90	-0.454	34.71	-0.473	36.74
1	-0.457	28.88	-0.466	30.84	-0.483	32.82	-0.500	35.74
0	-0.495	22.56	-0.501	25.98	-0.515	29.76	-0.529	32.92
- 1	-0.545	17.53	-0.544	20.85	-0.551	25.95	-0.560	31.40
- 2	-0.605	16.74	-0.596	18.45	-0.593	22.33	-0.593	28.30
- 3	-0.663	17.62	-0.650	18.52	-0.640	20.23	-0.631	24.94
- 4	-0.719	17.95	-0.704	18.72	-0.691	19.37	-0.673	22.16
- 5	-0.775	17.90	-0.758	18.42	-0.743	18.79	-0.721	20.22
- 6	-0.831	17.55	-0.812	18.00	-0.797	18.22	-0.772	18.91
- 7	-0.889	17.02	-0.869	17.50	-0.853	17.64	-0.826	17.96
- 8	-0.949	16.50	-0.927	16.99	-0.911	17.03	-0.883	17.19
- 9	-1.010	16.07	-0.986	16.48	-0.970	16.49	-0.942	16.62
-10	-1.073	15.74	-1.048	16.10	-1.032	16.14	-1.004	16.15
-11	-1.137	15.57	-1.111	15.86	-1.094	15.84	-1.066	15.87
-12	-1.201	15.53	-1.174	15.84	-1.158	15.70	-1.129	15.72
-13	-1.265	15.64	-1.237	15.87	-1.221	15.72	-1.193	15.73
-14	-1.329	15.85	-1.300	15.95	-1.285	15.85	-1.256	15.85
-15	-1.391	16.09	-1.362	16.14	-1.348	16.07	-1.319	16.05
-16			-1.423	16.48	-1.409	16.41	-1.381	16.35
-17			-1.483	16.87			-1.441	16.68

cont.

Capacity in NaSCN (cont.)

conc.	0.02446		0.06105		0.1169		0.1982	
σ	E	C	E	C	E	C	E	C
C/cm^2	Volts	$\mu F/cm^2$	Volts	$\mu F/cm^2$	Volts	$\mu F/cm^2$	Volts	$\mu F/cm^2$
9	-0.310	41.96	-0.335	41.46	-0.352	40.55	-0.357	41.56
8	-0.334	40.37	-0.359	39.99	-0.377	39.11	-0.382	38.31
7	-0.359	39.36	-0.385	39.03	-0.403	38.16	-0.409	36.27
6	-0.385	38.68	-0.410	38.43	-0.429	37.56	-0.437	35.11
5	-0.411	38.33	-0.437	37.96	-0.456	37.05	-0.466	34.41
4	-0.437	38.04	-0.463	37.64	-0.483	36.60	-0.495	33.86
3	-0.463	37.69	-0.490	37.34	-0.511	36.36	-0.525	33.62
2	-0.490	37.22	-0.517	37.09	-0.538	36.22	-0.555	33.50
1	-0.517	36.38	-0.554	36.63	-0.566	36.02	-0.584	33.54
0	-0.545	35.08	-0.571	35.98	-0.594	35.69	-0.614	33.57
- 1	-0.574	33.22	-0.599	34.92	-0.622	35.08	-0.644	33.47
- 2	-0.606	30.71	-0.629	33.54	-0.651	34.07	-0.674	33.11
- 3	-0.640	27.75	-0.660	31.34	-0.681	32.67	-0.705	32.36
- 4	-0.678	24.60	-0.693	28.75	-0.712	30.69	-0.736	31.10
- 5	-0.721	21.82	-0.729	25.76	-0.746	28.19	-0.769	29.22
- 6	-0.770	19.76	-0.771	22.74	-0.784	25.34	-0.805	26.84
- 7	-0.822	18.34	-0.818	20.16	-0.826	22.41	-0.844	24.14
- 8	-0.879	17.34	-0.870	18.31	-0.873	19.96	-0.888	21.51
- 9	-0.937	16.65	-0.926	17.10	-0.926	18.06	-0.937	19.32
-10	-0.998	16.17	-0.986	16.38	-0.983	16.88	-0.991	17.76
-11	-1.061	15.87	-1.048	15.99	-1.044	16.26	-1.049	16.81
-12	-1.124	15.73	-1.111	15.81	-1.106	15.96	-1.110	16.29
-13	-1.188	15.74	-1.174	15.81	-1.169	15.91	-1.172	16.11
-14	-1.251	15.86	-1.237	15.93	-1.231	16.01	-1.234	16.14
-15	-1.314	16.08	-1.300	16.15	-1.294	16.20	-1.295	16.33
-16	-1.375	16.37	-1.361	16.44	-1.355	16.50	-1.356	16.61
-17	-1.436	16.72			-1.415	16.84	-1.416	16.96
-18	-1.495	17.10			-1.473	17.25	-1.474	17.36
-19							-1.531	17.79
-20							-1.586	18.27

cont.

Capacity in NaSCN (cont.)

conc.	0.4943		0.9505	
σ	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
9	-0.408	39.97	-0.412	39.82
8	-0.434	38.94	-0.438	37.98
7	-0.460	38.43	-0.465	36.60
6	-0.486	37.98	-0.493	35.52
5	-0.512	37.08	-0.521	34.75
4	-0.540	38.45	-0.550	34.15
3	-0.567	36.32	-0.580	33.90
2	-0.595	36.34	-0.609	33.82
1	-0.622	36.41	-0.639	33.91
0	-0.650	36.46	-0.668	33.87
- 1	-0.677	36.35	-0.699	33.82
- 2	-0.705	36.04	-0.727	33.70
- 3	-0.733	35.39	-0.757	33.89
- 4	-0.761	34.39	-0.786	33.66
- 5	-0.791	32.99	-0.818	32.37
- 6	-0.822	31.20	-0.850	31.39
- 7	-0.855	29.03	-0.882	30.14
- 8	-0.891	26.59	-0.916	28.56
- 9	-0.931	24.05	-0.952	26.70
-10	-0.975	21.63	-0.991	24.70
-11	-1.023	19.62	-1.034	22.67
-12	-1.076	18.14	-1.080	20.84
-13	-1.133	17.23	-1.129	19.36
-14	-1.192	16.79	-1.183	18.36
-15	-1.252	16.69	-1.238	17.80
-16	-1.312	16.79	-1.295	17.56
-17	-1.371	17.05	-1.351	17.56
-18	-1.429	17.37		
-19	-1.486	17.76		
-20	-1.541	18.19		

Differential capacity on mercury in aqueous solutions of $NaClO_4$. Concentrations in mole/l. Temperature 25°C. 400 Hz. Data given as a function of surface charge σ.

Reference: A.N. Frumkin, V.F. Ivanov, B.B. Damaskin, A.A. Ivashchenko, N.I. Peshkova, Elektrokhimiya, 1, 279 (1965).

conc. 0.3		conc. 1.0		conc. 1.8		conc. 4.45		conc. 7.1		conc. 9.3	
σ	C	σ	C	σ	C	σ	C	σ	C	σ	C
$\mu C/cm^2$	$\mu F/cm^2$	$\mu C/cm^2$	$\mu F/cm^2$	$\mu C/cm^2$	$\mu F/cm^2$	$\mu C/cm^2$	$\mu F/cm^2$	$\mu C/cm^2$	$\mu F/cm^2$	$\mu C/cm^2$	$\mu F/cm^2$
18.1	27.7	18.3	27.7	18.5	27.2	19.9	26.9	20.2	26.5	20.6	26.4
15.7	25.2	15.8	25.2	16.1	24.8	17.5	24.7	17.8	24.4	18.2	24.2
13.2	23.4	13.4	23.3	13.7	23.2	15.1	23.1	15.4	23.0	15.8	22.9
10.9	22.7	11.1	22.6	11.4	22.6	12.8	22.6	13.2	22.5	13.5	22.4
8.8	23.3	8.9	22.9	9.2	22.8	10.6	22.8	11.0	22.7	11.5	22.6
6.4	25.1	6.6	24.2	6.9	23.8	8.3	23.7	8.7	23.5	9.1	22.9
3.5	27.6	3.9	26.3	4.4	25.6	3.8	25.2	6.4	24.6	6.7	23.7
0.6	28.6	1.2	28.1	1.6	27.2	3.0	26.8	3.7	26.1	4.2	24.8
- 2.1	26.4	- 1.5	28.1	- 1.0	28.4	0.3	28.2	0.9	27.5	1.7	26.1
- 4.5	21.9	- 4.1	25.0	- 3.8	27.3	- 2.6	28.3	- 1.9	28.2	- 1.0	27.3
- 6.6	18.7	- 6.4	21.0	- 6.5	23.9	- 5.3	26.6	- 4.7	28.2	- 3.8	27.8
- 8.4	17.0	- 8.6	18.4	- 8.7	20.4	- 7.9	23.4	- 7.5	27.0	- 6.5	28.0
-10.0	16.2	-10.3	16.7	-10.7	18.5	-10.0	20.6	-10.0	24.7	- 9.3	27.7
-11.7	15.9	-11.9	16.1	-12.4	17.4	-11.9	19.2	-12.3	22.0	-12.1	26.5
-13.3	16.1	-13.4	16.1	-14.1	17.1	-13.8	18.3	-14.4	20.6	-14.5	25.2
-14.8	16.3	-15.0	16.4	-15.8	17.1	-15.8	18.2	-16.4	19.8	-17.0	23.6
-16.5	16.7	-16.7	16.9	-17.6	17.3	-17.6	18.3	-18.3	19.7	-19.2	22.6
-18.2	17.4	-18.4	17.7	-19.4	18.1	-19.4	18.7	-20.3	19.9	-21.3	21.8
-20.0	18.2	-20.1	18.3	-21.1	18.7	-21.3	19.5	-22.3	20.5	-23.6	21.7
-21.7	20.4	-21.9	19.5	-22.9	19.7	-23.3	20.5	-24.4	21.4	-25.9	23.3

Differential capacity on mercury in aqueous solutions of KBr (concentration given in moles.l^{-1}).
Temperature: 25°C. Potentials measured with respect to a 0.1 M calomel electrode dipping into 0.1 M KBr.
Reference: J. Lawrence, R. Parsons and R. Payne, J. Electroanal. Chem. and Interfacial Electrochem., 16 (1968) 193.
Given are Potential and Capacitance as a function of surface charge σ.

c_{salt} =	0.005		0.010		0.020		0.050		0.100	
σ μC/cm^2	E Volts	C μF/cm^2	E Volts	C μF/cm^2	E Volts	C μF/cm^2	E Volts	C μF/cm^2	E Volts	C μF/cm^2
24	-0.123	227.6	-0.144	231.0	-0.175	236.0	-0.206	201.0	-0.230	189.5
23	-0.127	205.4	-0.154	181.7	-0.180	210.0	-0.211	183.2	-0.236	167.1
22	-0.132	182.5	-0.160	165.1	-0.185	183.6	-0.217	164.8	-0.242	144.4
21	-0.138	158.9	-0.166	148.2	-0.191	156.0	-0.223	146.1	-0.255	115.3
20	-0.150	119.2	-0.173	130.9	-0.206	112.2	-0.231	126.9	-0.265	105.7
19	-0.159	108.0	-0.181	113.4	-0.215	102.2	-0.234	107.7	-0.274	96.38
18	-0.169	96.95	-0.191	95.88	-0.225	92.28	-0.254	85.75	-0.285	87.43
17	-0.180	85.95	-0.206	78.79	-0.237	82.60	-0.267	79.58	-0.297	79.17
16	-0.192	75.38	-0.220	72.20	-0.250	72.93	-0.280	73.45	-0.312	73.86
15	-0.208	66.46	-0.234	65.90	-0.265	66.79	-0.244	67.72	-0.326	69.72
14	-0.224	60.49	-0.250	60.00	-0.280	61.07	-0.309	62.17	-0.341	65.56
13	-0.241	55.10	-0.268	55.46	-0.298	56.02	-0.326	57.16	-0.356	61.16
12	-0.260	50.73	-0.286	51.51	-0.316	52.56	-0.344	53.09	-0.373	56.50
11	-0.280	47.43	-0.306	48.39	-0.336	49.84	-0.363	50.51	-0.391	52.63
10	-0.302	45.37	-0.328	46.40	-0.357	47.69	-0.384	48.66	-0.411	50.29
9	-0.324	44.10	-0.349	45.58	-0.378	46.36	-0.405	47.46	-0.432	49.01
8	-0.347	43.35	-0.371	45.09	-0.400	45.83	-0.426	46.87	-0.452	48.20
7	-0.371	42.70	-0.394	44.39	-0.422	45.30	-0.447	46.60	-0.473	47.86
6	-0.394	41.68	-0.417	43.75	-0.444	44.73	-0.469	46.28	-0.494	47.87
5	-0.419	40.63	-0.440	43.12	-0.466	44.40	-0.490	46.06	-0.515	47.63
4	-0.444	39.61	-0.463	42.74	-0.489	43.73	-0.512	45.45	-0.536	47.56
3	-0.469	37.85	-0.487	40.90	-0.512	42.18	-0.534	44.60	-0.557	46.95
2	-0.497	35.32	-0.512	37.75	-0.536	39.99	-0.557	42.93	-0.579	45.58
1	-0.527	31.14	-0.540	33.84	-0.562	37.41	-0.581	40.20	-0.601	43.73
0	-0.564	22.14	-0.572	29.08	-0.590	34.33	-0.607	37.33	-0.624	41.56
-1	-0.614	18.01	-0.610	23.75	-0.624	25.61	-0.635	33.60	-0.650	37.72
-2	-0.672	17.30	-0.656	19.76	-0.663	23.12	-0.667	29.58	-0.678	33.90
-3	-0.729	17.89	-0.708	19.04	-0.709	20.75	-0.703	25.21	-0.709	30.09
-4	-0.784	18.21	-0.761	18.88	-0.758	19.67	-0.746	22.03	-0.745	25.87
-5	-0.839	18.17	-0.814	18.64	-0.810	18.95	-0.794	20.05	-0.787	22.48
-6	-0.895	17.86	-0.868	18.24	-0.864	18.32	-0.845	18.78	-0.833	20.17
-7	-0.952	17.23	-0.924	17.67	-0.919	17.68	-0.900	17.85	-0.885	18.60
-8	-1.011	16.62	-0.982	17.04	-0.977	17.09	-0.957	17.16	-0.940	17.57
-9	-1.072	16.26	-1.041	16.58	-1.036	16.61	-1.016	16.64	-0.999	16.82
-10	-1.134	15.94	-1.102	16.16	-1.097	16.26	-1.077	16.24	-1.059	16.40
-11	-1.197	15.74	-1.164	15.99	-1.159	16.06	-1.139	16.02	-1.120	16.13
-12	-1.261	15.75	-1.227	15.90	-1.222	15.95	-1.202	15.94	-1.182	16.03
-13	-1.324	15.81	-1.290	15.95	-1.284	16.03	-1.264	16.00	-1.245	16.08
-14	-1.387	16.03	-1.353	16.04	-1.346	16.20	-1.326	16.18	-1.307	16.24

Differential capacity on mercury in aqueous solutions of KBr (cont.)

	salt = 0.2		0.5		1.0		2.0		5.0	
σ	E	C	E	C	E	C	E	C	E	C
μC/cm^2	Volts	μF/cm^2	Volts	μF/cm^2	Volts	μF/cm^2	Volts	μF/cm^2	Volts	μF/cm^2
24	-0.242	162.7	-0.278	150.0	-0.307	132.2	-0.331	133.6	-0.370	118.3
23	-0.253	133.4	-0.285	135.5	-0.315	122.6	-0.339	120.5	-0.379	109.8
22	-0.261	123.5	-0.293	121.2	-0.323	113.1	-0.350	105.1	-0.388	101.6
21	-0.269	113.7	-0.304	106.3	-0.333	103.8	-0.360	98.22	-0.398	93.81
20	-0.278	103.9	-0.314	98.69	-0.343	94.77	-0.370	91.46	-0.410	87.63
19	-0.288	94.44	-0.324	91.22	-0.355	86.52	-0.382	84.89	-0.422	82.30
18	-0.300	84.79	-0.335	84.03	-0.367	80.57	-0.394	78.58	-0.434	77.11
17	-0.313	79.00	-0.348	77.27	-0.380	74.81	-0.408	72.39	-0.448	72.12
16	-0.326	73.36	-0.362	71.72	-0.394	69.34	-0.422	68.18	-0.462	67.52
15	-0.340	67.96	-0.376	66.72	-0.409	64.50	-0.437	63.65	-0.478	63.25
14	-0.355	62.94	-0.392	62.11	-0.425	60.42	-0.453	59.54	-0.494	59.38
13	-0.372	58.50	-0.408	58.03	-0.442	56.77	-0.471	56.04	-0.511	56.03
12	-0.390	54.71	-0.426	54.85	-0.460	53.67	-0.489	53.01	-0.530	53.10
11	-0.409	52.03	-0.445	52.55	-0.479	51.34	-0.508	50.64	-0.549	50.52
10	-0.428	50.28	-0.464	50.77	-0.499	49.63	-0.528	48.75	-0.569	48.43
9	-0.448	49.03	-0.484	49.56	-0.519	48.62	-0.549	47.63	-0.590	46.77
8	-0.469	48.30	-0.504	48.93	-0.540	48.48	-0.570	47.70	-0.612	46.48
7	-0.490	48.22	-0.525	48.88	-0.561	49.09	-0.591	49.22	-0.633	48.04
6	-0.510	48.20	-0.545	49.24	-0.581	49.56	-0.611	50.10	-0.654	49.35
5	-0.531	48.19	-0.566	49.67	-0.601	49.34	-0.631	49.24	-0.674	48.52
4	-0.552	48.17	-0.586	49.44	-0.621	48.80	-0.652	47.97	-0.695	45.59
3	-0.573	47.75	-0.606	49.07	-0.642	48.31	-0.673	47.06	-0.717	44.53
2	-0.594	47.02	-0.627	48.57	-0.663	48.05	-0.694	46.82	-0.740	44.05
1	-0.615	45.90	-0.647	47.64	-0.684	48.12	-0.716	46.36	-0.763	43.88
0	-0.638	44.00	-0.668	46.80	-0.715	45.87	-0.737	45.96	-0.785	43.50
- 1	-0.661	41.44	-0.690	44.57	-0.727	44.98	-0.759	44.84	-0.809	42.74
- 2	-0.686	38.47	-0.713	42.19	-0.750	43.16	-0.782	43.38	-0.833	41.77
- 3	-0.713	34.49	-0.738	39.32	-0.774	40.52	-0.806	41.55	-0.857	40.39
- 4	-0.744	30.53	-0.765	35.87	-0.779	37.50	-0.830	39.32	-0.882	38.62
- 5	-0.779	26.68	-0.794	32.12	-0.827	34.49	-0.857	36.69	-0.909	36.72
- 6	-0.820	22.90	-0.827	28.06	-0.858	30.94	-0.885	33.74	-0.937	34.52
- 7	-0.866	20.17	-0.865	24.59	-0.892	27.44	-0.916	30.82	-0.967	32.09
- 8	-0.918	18.39	-0.909	21.55	-0.931	24.10	-0.950	27.67	-0.999	29.73
- 9	-0.975	17.22	-0.958	19.26	-0.975	21.39	-0.989	24.76	-1.034	27.43
-10	-1.034	16.515	-1.012	17.81	-1.025	19.28	-1.031	22.28	-1.073	25.25
-11	-1.095	16.18	-1.070	16.87	-1.078	17.90	-1.078	20.28	-1.114	23.34
-12	-1.157	16.00	-1.130	16.45	-1.136	17.13	-1.129	18.89	-1.158	21.75
-13	-1.220	16.00	-1.191	16.30	-1.195	16.80	-1.183	18.14	-1.206	20.57
-14	-1.282	16.14	-1.252	16.37	-1.254	16.71	-1.239	17.71	-1.255	19.91

These data agree within 0.6 % with the data of Grahame et al (4th report pages 19-21) at potentials more negative than -0.4 V (vs 0.1 M Calomel electrode). At more positive potentials there are unexplained deviations which increase as the potential becomes more positive and the capacity higher.

Differential capacity on mercury in aqueous solutions of various salts. Temperature: 25°C. Potentials measured with respect to 0.1 M calomel electrode.
Reference: J.Lawrence, B.Sci. Thesis, Bristol (1964)

Given is $10^2 C$ in F/m^2 as a function of potential.

c salt	$KClO_3$	$KBrO_3$	KIO_3	$NaNO_2$	$NaHSO_4$	Na_2SO_3	$Na_2S_2O_6$
salt conc.	0.1 M	0.1 M	0.1 M	0.1 M	0.1 M	0.05 M	0.05 M
$\frac{E}{\text{Volts}}$							
0.25	55.63						
0.20	44.56	77.64			59.66		
0.15	38.63	62.56			50.80		
0.10	35.14	53.27			45.78		
+0.05	32.26	46.38		64.00	42.40		58.18
0.00	30.05	41.36	101.2	46.77	39.88		40.18
-0.05	28.31	37.85	75.42	37.03	38.19		35.16
-0.10	27.02	35.37	63.13	31.92	37.22		32.86
-0.15	26.14	33.78	54.51	29.30	36.78		31.62
-0.20	25.71	32.78	49.01	28.29	36.70		31.06
-0.25	25.60	32.16	45.23	28.34	36.68	123.2	30.94
-0.30	25.88	31.73	42.26	29.04	35.87	71.82	31.38
-0.35	26.48	31.18	39.41	30.03	34.07	62.68	31.80
-0.40	27.07	30.40	36.18	30.79	31.35	56.45	31.69
-0.45	27.46	29.37	32.27	30.47	28.36	47.04	30.38
-0.50	27.37	28.10	28.57	28.72	25.43	36.23	28.05
-0.55	26.54	26.62	25.70	26.18	23.02	27.67	25.48
-0.60	25.01	25.05	23.66	23.70	21.55	23.22	23.41
-0.65	23.37	23.45	22.25	21.88	20.73	21.31	21.90
-0.70	21.93	22.05	21.18	20.67	20.11	20.33	20.77
-0.75	20.67	20.86	20.28	19.73	19.47	19.54	19.80
-0.80	19.62	19.81	19.47	18.89	18.80	18.80	18.88
-0.85	18.68	18.88	18.69	18.13	18.10	18.08	18.10
-0.90	17.95	18.10	18.00	17.47	17.51	17.45	17.41
-0.95	17.31	17.46	17.37	16.91	16.94	16.91	16.84
-1.00	16.82	16.92	16.93	16.47	16.49	16.46	16.40
-1.05	16.46	16.57	16.51	16.15	16.12	16.13	16.08
-1.10	16.18	16.30		15.91	15.87	15.89	15.87
-1.15	16.06	16.18		15.74	15.73	15.75	15.76
-1.20	16.00	16.12		15.70	15.62	15.71	15.72
-1.25	16.06	16.16		15.74	15.67	15.76	15.78
-1.30	16.18	16.28		15.83	15.76	15.85	15.87
-1.35	16.39	16.47		16.00	15.87	16.01	16.05
-1.40	16.62	16.74		16.23	16.10	16.21	16.27
-1.45	16.95	17.03		16.46		16.49	16.53
-1.50	17.31	17.39		16.78		16.78	16.82

Differential capacity of mercury in 1 mol l^{-1} aqueous solutions. T = 25°C
Potential with respect to a saturated calomel electrode.

Reference: G.C. Barker and R.L. Faircloth. Advances in Polarography, ed. Longmuir, Pergamon, Oxford (1959) 319

E/volts	C/μF cm^{-2}							
	KF	KCl	KBr	KI	KSCN	KCN	KNO_3	$NaClO_4$
0.0	32.09	-	-	-	-	-	25.24	22.85
-0.1	29.54	58.13	-	-	-	-	23.48	22.57
-0.2	28.90	40.69	-	-	-	-	23.62	23.45
-0.3	28.42	38.70	74.34	-	65.47	-	25.42	25.22
-0.4	27.17	41.83	51.63	-	45.01	-	28.44	27.45
-0.5	24.90	39.92	48.80	-	39.18	-	30.43	28.15
-0.6	22.35	31.15	47.78	84.48	37.55	-	28.85	26.06
-0.7	19.85	23.28	39.19	62.50	36.32	-	24.38	22.38
-0.8	18.21	19.19	27.89	54.49	33.82	20.28	20.33	19.03
-0.9	17.20	17.38	20.87	45.96	29.12	16.65	17.89	16.89
-1.0	16.53	16.51	17.85	33.16	23.68	16.07	16.76	16.10
-1.1	16.42	16.39	17.04	23.19	19.76	16.12	16.45	16.05
-1.2	16.65	16.63	16.92	19.05	18.02	16.39	16.54	16.15
-1.3	17.15	17.15	17.31	18.00	17.50	16.87	16.85	16.47
-1.4	17.66	17.96	18.00	18.27	17.83	17.63	17.71	17.08
-1.5	18.76	18.89	18.95	19.05	18.68	18.36	-	18.05
-1.6	20.03	20.17	20.20	20.23	19.81	19.53	-	19.31
-1.7	-	21.74	-	-	-	-	-	-

Differential capacity of mercury in aqueous solutions. T= 25°C

Potential with respect to a saturated calomel electrode.

Reference: G.C. Barker and R.L. Faircloth. Advances in Polarography, ed. Longmuir, Pergamon, Oxford (1959) 319.

	C/μF cm^{-2}						
E/volts	0.5 M K_2CO_3	0.5 M $KHCO_3$	0.3 M HCl	0.3 M HNO_3	0.5 M K_2SO_4	0.25 M H_2SO_4	0.5 M Na_2S
0.0	-	40.92	-	23.96	38.73	31.52	-
-0.1	-	30.34	45.49	22.82	39.02	29.79	-
-0.2	48.79	31.08	36.98	23.23	37.98	31.09	-
-0.3	33.67	33.87	37.93	25.60	31.79	32.52	-
-0.4	27.44	31.31	39.88	28.86	26.66	29.66	-
-0.5	24.12	25.10	33.94	30.27	23.16	24.40	-
-0.6	21.21	21.06	24.82	26.97	20.57	20.97	-
-0.7	18.99	18.69	19.62	21.80	18.45	18.62	-
-0.8	17.39	16.98	17.10	18.12	17.25	17.10	-
-0.9	16.53	16.05	15.82	16.84	16.26	16.10	-
-1.0	16.07	15.64	15.20	-	16.05	15.54	17.34
-1.1	16.07	15.72	-	-	15.89	15.45	16.54
-1.2	16.43	16.05	-	-	16.05	-	16.21
-1.3	16.99	16.56	-	-	16.61	-	16.63
-1.4	17.74	17.25	-	-	17.38	-	17.20
-1.5	18.59	18.23	-	-	18.23	-	17.95
-1.6	19.80	19.48	-	-	19.53	-	18.89

Differential capacity of mercury in aqueous solutions of CsF. T = 25°C

Potential with respect to 1 mol l^{-1} CsCl calomel electrode in 0.00792 mol l^{-1} CsF.

Reference: F.R.M. Deane, A.E. Higinbotham and D.C. Grahame,
U.S. Dept. Office Tech. Serv., P.B. Rept., 143 (1958) 476.

c/mol l^{-1}	1.0		0.1		0.01		0.001	
σ/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
22	-1.5905	20.45						
20	-1.4901	19.41	-1.5381	19.42	-1.5852	19.37		
18	-1.3846	18.57	-1.4324	18.51	-1.4794	18.49	-1.5397	18.51
16	-1.2750	17.95	-1.3223	17.85	-1.3691	17.80	-1.4290	17.66
14	-1.1623	17.57	-1.2186	17.40	-1.2551	17.35	-1.3138	17.12
12	-1.0480	17.48	-1.0929	17.26	-1.1391	17.21	-1.1959	16.86
10	-0.9350	18.007	-0.9779	17.61	-1.0239	17.60	-1.0775	17.00
8	-0.8270	19.16	-0.8665	18.36	-0.9125	18.34	-0.9615	17.53
6	-0.7270	20.93	-0.7604	19.32	-0.8059	19.19	-0.8497	18.28
4	-0.6358	23.01	-0.6590	20.14	-0.7024	19.23	-0.7422	18.68
3	-0.5933	24.08	-0.6098	20.40	-0.6495	18.48	-0.6880	18.07
2	-0.5526	25.08	-0.5609	20.52	-0.5932	16.97	-0.6296	16.01
1.5	-0.5329	25.54	-0.5365	20.58	-0.5627	15.77	-0.5968	14.32
1	-0.5135	25.95	-0.5123	20.66	-0.5298	14.57	-0.5587	11.92
0.7	-0.5020	26.18	-0.4978	20.75	-0.5088	14.08	-0.5315	10.21
0.5	-0.4944	26.33	-0.4882	20.81	-0.4945	13.88	-0.5107	9.04
0.2	-0.4830	26.52	-0.4743	20.94	-0.4727	13.78	-0.4751	8.02
0.1	-0.4792	26.59	-0.4690	21.00	-0.4655	13.85	-0.4625	7.95
0	-0.4755	26.65	-0.4643	21.05	-0.4583	13.95	-0.4500	8.02
-0.1	-0.4717	26.72	-0.4595	21.11	-0.4512	14.10	-0.4377	8.23
-0.2	-0.4680	26.78	-0.4548	21.16	-0.4441	14.27	-0.4257	8.55
-0.5	-0.4569	26.94	-0.4407	21.35	-0.4235	14.90	-0.3932	10.04
-0.7	-0.4494	27.04	-0.4314	21.49	-0.4103	15.40	-0.3743	10.57
-1	-0.4384	27.19	-0.4175	21.72	-0.3913	16.20-	-0.3491	12.68
-1.5	-0.4201	27.40	-0.3947	22.16	-0.3618	17.69	-0.3129	15.02
-2	-0.4019	27.62	-0.3723	22.62	-0.3346	19.15	-0.2817	16.89
-3	-0.3659	27.94	-0.3291	23.68	-0.2854	21.38	-0.2269	19.49
-4	-0.3303	28.18	-0.2878	24.63	-0.2404	23.03	-0.1781	21.40
-6	-0.2599	28.61	-0.2092	26.21	-0.1572	24.81	-0.0906	24.10
-8	-0.1905	29.05	-0.1446	27.26	-0.0781	25.69	-0.0101	25.49
-10	-0.1225	29.81	-0.0631	28.47	-0.0113	26.31	0.0660	27.17
-12	-0.0566	30.97					0.1368	29.38
-14	0.0064	32.67					0.2022	31.98
-16	0.0651	35.80					0.2620	35.01
-17	0.0922	38.32					0.2898	36.92
-18	0.1170	43.32					0.3159	40.31
-19	0.1381	52.16					0.3390	47.33

Differential capacity of mercury in aqueous solutions of KCl. T = 25°C
Potential with respect to 1 mol l^{-1} KCl calomel.
Reference: D.C. Grahame and R. Parsons, J. Amer. Chem. Soc., 83 (1961) 291.

c/mol l^{-1}	0.02		0.04		0.0632		0.10	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm
-20	-1.642	18.79	-1.625	18.78	-1.612	18.78	-1.599	18.77
-18	-1.533	17.66	-1.515	17.66	-1.502	17.67	-1.489	17.69
-16	-1.416	16.77	-1.399	16.79	-1.386	16.82	-1.373	16.85
-14	-1.295	16.15	-1.277	16.19	-1.264	16.22	-1.252	16.25
-12	-1.170	15.93	-1.152	15.97	-1.140	15.99	-1.128	16.01
-10	-1.045	16.23	-1.028	16.28	-1.016	16.33	-1.004	16.39
-8	-0.924	17.09	-0.908	17.17	-0.897	17.24	-0.885	17.29
-6	-0.811	18.25	-0.795	18.36	-0.785	18.53	-0.774	18.80
-4	-0.704	19.02	-0.795	19.41	-0.681	19.91	-0.672	20.65
-2	-0.598	18.43	-0.690	20.21	-0.584	21.85	-0.581	23.90
0	-0.492	20.87	-0.497	24.78	-0.501	27.28	-0.506	29.73
2	-0.413	30.67	-0.427	32.69	-0.436	34.37	-0.445	36.01
4	-0.354	36.34	-0.371	37.30	-0.381	38.13	-0.392	39.22
6	-0.301	37.77	-0.318	38.38	-0.330	38.98	-0.342	39.63
8	-0.247	37.21	-0.266	37.84	-0.278	38.30	-0.291	38.75
10	-0.193	36.66	-0.212	37.24	-0.225	37.64	-0.239	38.01
12	-0.139	37.39	-0.159	37.88	-0.173	38.24	-0.186	38.55
14	-0.087	40.41	-0.108	40.91	-0.122	41.18	-0.136	41.47
16	-0.041	46.62	-0.062	47.02	-0.076	47.21	-0.091	47.79
18	-0.002	58.32	-0.023	57.97	-0.038	57.85	-0.052	58.08
20								

Capacity of mercury in aqueous KCl (cont.)

$c/\mathrm{mol\ l^{-1}}$	0.20		0.40		0.632	
$\sigma/\mu\mathrm{C\ cm^{-2}}$	E/volt	$C/\mu\mathrm{F\ cm^{-2}}$	E/volt	$C/\mu\mathrm{F\ cm^{-2}}$	E/volt	$C/\mu\mathrm{F\ cm^{-2}}$
-20	-1.584	18.80	-1.565	18.79	-1.554	18.80
-18	-1.475	17.72	-1.455	17.71	-1.445	17.74
-16	-1.359	16.88	-1.340	16.89	-1.329	16.92
-14	-1.239	16.30	-1.219	16.32	-1.209	16.39
-12	-1.114	16.05	-1.095	16.16	-1.086	16.30
-10	-0.991	16.44	-0.973	16.65	-0.965	16.92
-8	-0.873	17.47	-0.857	18.01	-0.851	18.49
-6	-0.763	19.23	-0.752	20.49	-0.750	21.53
-4	-0.666	22.03	-0.663	24.64-	-0.666	26.50
-2	-0.583	26.84	-0.590	30.48	-0.598	32.58
0	-0.516	33.09	-0.530	36.36	-0.542	38.17
2	-0.460	38.09	-0.478	40.51	-0.492	41.80
4	-0.409	40.48	-0.430	42.13	-0.445	42.97
6	-0.360	40.59	-0.382	41.55	-0.398	42.17
8	-0.310	39.42	-0.333	40.08	-0.350	40.42
10	-0.259	38.47	-0.283	38.97	-0.300	37.19
12	-0.207	38.97	-0.231	39.43	-0.249	39.65
14	-0.157	41.96	-0.182	42.37	-0.200	42.58
16	-0.112	48.01	-0.138	48.44	-0.156	48.65
18	-0.075	58.78	-0.100	58.64	-0.118	58.75
20	-0.044	77.39	-0.070	74.12	-0.088	74.55

Capacity of mercury in aqueous KCl (cont.)

c/mol l^{-1}	1.0		1.259		1.585	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20	-1.542	18.82	-1.535	18.86	-1.527	18.90
-18	-1.432	17.78	-1.425	17.82	-1.418	17.90
-16	-1.317	16.99	-1.311	17.06	-1.303	17.17
-14	-1.197	16.52	-1.192	16.63	-1.186	16.78
-12	-1.076	16.51	-1.071	16.65	-1.066	16.86
-10	-0.957	17.25	-0.953	17.52	-0.951	17.89
-8	-0.846	19.27	-0.845	19.82	-0.845	20.48
-6	-0.751	23.04	-0.753	23.87	-0.756	24.81
-4	-0.673	28.57	-0.678	29.53	-0.683	30.44
-2	-0.609	34.62	-0.616	35.51	-0.623	36.27
0	-0.556	39.77	-0.563	40.39	-0.572	40.88
2	-0.507	42.77	-0.516	42.99	-0.524	43.14
4	-0.461	43.48	-0.470	43.54	-0.478	43.47
6	-0.415	42.27	-0.423	42.18	-0.432	42.03
8	-0.366	40.48	-0.375	40.32	-0.383	40.11
10	-0.316	39.23	-0.324	39.15	-0.332	39.01
12	-0.265	39.84	-0.273	39.90	-0.281	39.97
14	-0.217	42.94	-0.225	43.16	-0.233	43.46
16	-0.173	49.16	-0.181	49.62	-0.190	50.12
18	-0.136	59.42	-0.145	59.72	-0.153	60.38
20	-0.105	75.12	-0.115	75.03	-0.124	75.99

Differential Capacity of a mercury electrode in aqueous NaOH T = 25°C

Potentials with respect to a hydrogen electrode in the same solution.

Reference: R.Payne, J. Electroanal. Chem.,1 (1964) 343

$c/\text{mol l}^{-1}$	0.01089		0.02178		0.04356		0.1089	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
6								
5			0.759	30.18	0.764	29.67	0.766	30.28
4	0.719	27.27	0.724	26.64	0.728	26.52	0.731	27.57
3	0.680	24.14	0.684	23.94	0.689	24.27	0.694	25.77
2	0.636	20.73	0.640	21.31	0.646	22.28	0.654	24.42
1	0.582	16.73	0.590	18.51	0.599	20.41	0.612	23.24
0	0.515	13.69	0.532	16.40	0.548	19.06	0.568	22.24
-1	0.443	14.82	0.471	16.57	0.495	18.70	0.522	21.49
-2	0.380	16.91	0.412	17.76	0.442	18.96	0.475	20.90
-3	0.324	18.08	0.358	18.54	0.389	19.17	0.427	20.41
-4	0.289	18.40	0.304	17.74	0.337	19.09	0.377	19.84
-5	0.214	18.24	0.250	18.49	0.284	18.69	0.326	19.18
-6	0.159	17.80	0.196	18.01	0.230	18.12	0.273	18.44
-7	0.102	17.24	0.139	17.42	0.174	17.48	0.217	17.71
-8	0.043	16.69	0.081	16.83	0.115	16.88	0.160	17.05
-9	-0.018	16.20	0.020	16.32	0.055	16.37	0.100	16.48
-10	-0.080	15.80	-0.042	15.94	-0.007	15.98	0.038	16.07
-11	-0.144	15.59	-0.105	15.71	-0.070	15.74	-0.024	15.81
-12	-0.209	15.54	-0.169	15.61	-0.134	15.64	-0.088	15.69
-13	-0.273	15.59	-0.233	15.65	-0.197	15.69	-0.152	15.70
-14	-0.337	15.74	-0.296	15.80	-0.261	15.81	-0.215	15.84
-15	-0.400	15.97	-0.359	16.03	-0.324	16.05	-0.278	16.06
-16	-0.462	16.31	-0.421	16.34	-0.386	16.34	-0.340	16.37
-17	-0.523	16.67	-0.482	16.71	-0.446	16.69	-0.400	16.72
-18	-0.582	17.09	-0.541	17.11	-0.505	17.15	-0.459	17.15
-19	-0.640	17.52	-0.598	17.76	-0.563	17.57	-0.517	17.57
-20	-0.696	18.02	-0.655	18.03	-0.619	18.05	-0.573	18.05
-21	-0.751	18.55	-0.709	18.53	-0.674	18.55	-0.628	18.55
-22			-0.762	19.06	-0.727	19.08	-0.681	19.10

Capacity in NaOH (cont.)

c/mol l^{-1}	0.2178		0.545		1.089	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
6	0.799	35.94	0.804	38.19	0.811	41.65
5	0.769	31.64	0.775	33.70	0.785	36.61
4	0.736	28.85	0.745	30.90	0.756	33.00
3	0.700	27.17	0.711	29.06	0.725	30.73
2	0.663	26.03	0.676	27.89	0.691	29.23
1	0.624	25.09	0.640	26.97	0.656	28.08
0	0.583	23.99	0.602	26.04	0.620	27.02
-1	0.541	23.34	0.563	25.07	0.582	25.91
-2	0.497	22.47	0.522	24.00	0.543	24.72
-3	0.452	21.57	0.479	22.84	0.501	23.47
-4	0.404	20.67	0.434	21.62	0.457	22.19
-5	0.355	19.75	0.387	20.47	0.411	20.92
-6	0.303	18.83	0.337	19.38	0.362	19.77
-7	0.249	17.98	0.284	18.38	0.310	18.72
-8	0.191	17.23	0.228	17.52	0.255	17.84
-9	0.133	16.62	0.170	16.85	0.198	17.15
-10	0.072	16.17	0.109	16.37	0.139	16.66
-11	0.009	15.88	0.048	16.08	0.078	16.33
-12	-0.054	15.75	-0.015	15.94	0.016	16.18
-13	-0.118	15.77	-0.078	15.95	-0.046	16.14
-14	-0.181	15.87	-0.140	16.06	-0.107	16.24
-15	-0.244	16.09	-0.202	16.28	-0.169	16.44
-16	-0.305	16.39	-0.283	16.57	-0.229	16.73
-17	-0.365	16.75	-0.323	16.91	-0.288	17.14
-18	-0.425	17.14	-0.381	17.30	-0.347	16.65
-19	-0.482	17.58	-0.438	17.72	-0.405	17.92
-20	-0.538	18.04			-0.460	18.33
-21	-0.593	18.54			-0.514	18.80
-22	-0.646	19.07			-0.567	19.36
-23	-0.698	19.65			-0.618	19.90

Differential capacity of mercury in aqueous solutions of KCl. T = 25°C

Potential with respect to a cation reversible electrode whose potential with respect to a 1 M. KCl calomel is given at the bottom of each column.

Reference: J.R. Sams, C.W. Lees and D.C. Grahame, J. Phys. Chem., 63 (1959) 2032.

c/mol l^{-1}	2.0	3.0	2.449	
E/volts	C/μF cm^{-2}	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-1.80	23.00	22.99	-1.27	17.30
-1.75	22.00	21.99	-1.22	17.15
-1.70	21.06	21.13	-1.17	17.11
-1.65	20.28	20.41	-1.12	17.16
-1.60	19.65	19.80	-1.07	17.40
-1.55	19.09	19.27	-1.02	17.84
-1.50	18.60	18.82	-0.97	18.55
-1.45	18.18	18.42	-0.92	19.63
-1.40	17.81	18.08	-0.87	21.18
-1.35	17.48	17.81	-0.82	23.33
-1.30	17.23	17.60	-0.76	26.93
-1.25	17.05	17.46	-0.70	31.39
-1.20	16.95	17.39	-0.66	34.47
-1.15	16.94	17.46	-0.60	38.06
-1.10	17.05	17.71	-0.56	39.17
-1.05	17.33	18.14	-0.52	39.07
-1.00	17.83	18.84	-0.50	38.69
-0.95	18.65	19.91	-0.46	37.61
-0.90	19.88	21.53	-0.42	36.53
-0.85	21.46	23.89	-0.40	36.16
-0.80	23.93	27.04	-0.38	36.05
-0.75	27.23	30.90	-0.36	36.08
-0.70	31.62	35.39	-0.34	36.49
-0.65	36.29	39.40	-0.30	38.05
-0.60	40.64	42.12	-0.26	41.29
-0.55	43.03	42.58	-0.22	47.06
-0.50	43.00	41.20	-0.20	51.39
-0.45	41.39	39.47		
-0.40	39.66	38.81		
-0.35	39.11	39.72		
-0.30	40.28	43.10		
-0.25	44.65	50.89		
(E-E(MCE))/V	-0.0176	-0.0292		-0.0234

Differential capacity of mercury in 0.1 mol l^{-1} aqueous NaF at various pressures. T = 25°C

Potential with respect to 0.1 mol l^{-1} NaCl calomel electrode at the pressure of the experimental.

Reference: G.J. Hills and R. Payne, Trans. Faraday Soc., 61 (1965) 326.

P/atm 1		121		255		384		520	
E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
0.079	32.87	0.026	30.38	0.082	32.81	0.067	31.84	0.070	31.94
0.024	30.47	-0.001	29.55	0.033	30.51	0.023	30.06	0.018	29.85
-0.039	28.78	-0.059	28.25	-0.013	29.10	-0.023	28.79	-0.032	28.55
-0.108	27.62	-0.113	27.43	-0.071	27.95	-0.076	27.78	-0.090	27.55
-0.171	26.88	-0.166	26.86	-0.126	27.18	-0.137	27.01	-0.141	26.94
-0.247	26.13	-0.208	26.49	-0.178	26.68	-0.189	26.57	-0.186	26.56
-0.317	25.15	-0.270	25.84	-0.233	26.20	-0.233	26.19	-0.247	26.06
-0.387	23.76	-0.333	24.90	-0.287	25.62	-0.278	25.72	-0.309	25.38
-0.453	22.16	-0.388	23.78	-0.327	25.03	-0.341	24.85	-0.372	24.34
-0.514	20.92	-0.449	22.34	-0.373	24.18	-0.396	23.76	-0.436	22.95
-0.587	20.25	-0.502	21.21	-0.436	22.73	-0.459	22.28	-0.487	21.80
-0.639	20.09	-0.549	20.56	-0.500	21.33	-0.502	21.38	-0.560	20.75
-0.713	19.70	-0.607	20.25	-0.548	20.65	-0.556	20.67	-0.612	20.51
-0.822	18.53	-0.663	20.08	-0.610	20.31	-0.606	20.41	-0.738	20.00
-0.915	17.40	-0.722	19.73	-0.662	20.18	-0.660	20.29	-0.683	20.34
-1.001	16.56	-0.788	19.06	-0.725	19.82	-0.728	19.93	-0.772	19.64
-1.110	15.91	-0.859	18.19	-0.798	19.08	-0.784	19.37	-0.849	18.71
-1.213	15.72	-0.934	17.29	-0.880	18.04	-0.853	18.53	-0.933	17.64
-1.298	15.81	-1.021	16.46	-0.958	17.14	-0.937	17.47	-1.023	16.72
-1.398	16.16	-1.093	16.02	-1.032	16.48	-1.032	16.54	-1.119	16.10
-1.499	16.72	-1.210	15.74	-1.120	15.97	-1.110	16.07	-1.213	15.83
-1.595	17.44	-1.331	15.87	-1.210	15.77	-1.210	15.80	-1.339	15.92
-1.703	18.34	-1.431	16.28	-1.334	15.89	-1.336	15.90	-1.439	16.26
-1.798	19.29	-1.531	16.88	-1.434	16.27	-1.436	16.27	-1.539	16.80
		-1.631	17.62	-1.534	16.84	-1.536	16.82	-1.639	17.50
		-1.731	18.49	-1.634	17.56	-1.636	17.53	-1.739	18.35
				-1.734	18.40	-1.736	18.37		

Differential capacity of mercury in aqueous solutions at high pressures

T = 25°C

Potential with respect to 0.1 mol l^{-1} NaCl calomel electrode at the pressure of the experiment.

Reference: G.J. Hills and R. Payne, Trans Faraday Soc., 61 (1965) 326

0.1 mol l^{-1} NaCl at 720 atm				0.1 mol l^{-1} $NaNO_3$ at 540 atm			
E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF^{-2}	E/volts	C/μF cm^{-2}
-1.886	20.18	-0.608	26.16	-1.724	18.47	-0.484	30.14
-1.830	19.44	-0.581	28.31	-1.672	18.00	-0.467	29.66
-1.766	18.78	-0.554	30.80	-1.622	17.60	-0.427	28.78
-1.698	18.16	-0.514	34.42	-1.571	17.22	-0.374	27.13
-1.635	17.66	-0.470	37.08	-1.498	16.76	-0.322	25.47
-1.562	17.12	-0.445	37.83	-1.454	16.49	-0.281	24.40
-1.474	16.57	-0.432	38.02	-1.399	16.24	-0.240	23.85
-1.372	16.11	-0.401	37.91	-1.333	15.99	-0.203	23.38
-1.287	15.91	-0.353	37.00	-1.234	15.83	-0.177	23.24
-1.223	15.88	-0.378	37.56	-1.156	15.88	-0.166	23.22
-1.147	16.02	-0.312	36.26	-1.082	16.15	-0.134	23.28
-1.082	16.27	-0.274	36.10	-1.018	16.62	-0.096	23.60
-1.002	16.85	-0.293	36.06	-0.929	17.57	-0.047	24.40
-0.947	17.40	-0.249	36.48	-0.878	18.28	-0.003	25.54
-0.888	18.17	-0.210	38.08	-0.811	19.54	0.044	27.16
-0.835	18.96	-0.180	40.46	-0.754	20.85	0.090	29.31
-0.776	19.97	-0.154	43.51	-0.696	22.83		
-0.731	20.87	-0.124	49.38	-0.645	24.14		
-0.683	22.22	-0.90	60.53	-0.594	27.75		
-0.644	23.39			-0.547	29.56		
				-0.510	30.20		

Differential capacity of a mercury electrode in aqueous solutions of HCl.

$T = 25^{\circ}C$

Potential with respect to calomel electrode in HCl of concentration equal to that of the working solution.

Reference: J.O'M. Bockris, K. Muller, H. Wroblova and Z. Kovac. J. Electroanal. Chem. 10 (1965) 416

	$C/\mu F\ cm^{-2}$					
$c/mol\ l^{-1}$ E/volts	2.94	1.0	0.3	0.1	0.03	0.01
- 0	1008	846	884	843	635	517
- 50	187	163	147	129	125	121
- 100	75.9	69.4	63.7	58.6	56.3	53.2
- 150	53.22	50.68	47.82	45.25	43.38	41.16
- 200	43.74	42.28	41.13	39.89	38.44	36.36
- 250	38.20	38.29	37.68	37.68	36.55	34.58
- 300	35.16	36.86	38.01	37.58	36.47	34.36
- 350	34.29	37.35	38.49	38.38	37.31	35.01
- 400	34.81	39.03	40.00	39.28	37.94	34.79
- 450	36.50	40.88	40.67	38.70	36.78	33.32
- 500	38.28	41.57	39.14	35.60	32.80	28.89
- 550	39.42	40.04	35.36	30.86	26.68	22.24
- 600	39.03	36.21	30.49	26.32	21.71	16.60
- 650	37.01	31.61	26.03	23.02	19.68	15.35
- 700	33.41	27.19	22.70	20.99	19.13	16.30
- 750	29.33	23.70	20.20	19.57	18.80	17.28
- 800	26.27	20.92	18.51	18.58	18.33	17.69
- 850	23.65	19.03	17.61	17.59	17.80	17.56
- 900	21.45	17.78	16.75	16.90	17.20	17.21
- 950	19.96	16.93	16.33	16.44	16.66	16.71
- 1000	19.29	16.53	15.86	16.00	16.20	16.26
- 1050		16.20	15.56	15.63	15.86	15.86
- 1100			15.36	15.44	15.54	15.47
- 1150			15.23	15.28	15.24	15.09
- 1200				15.28	15.10	14.91
- 1250				15.29	15.10	14.72
- 1300				15.33	15.13	14.67
- 1350						14.71

Differential capacity of mercury in aqueous solutions containing KCl.

T = 25°C

Potential with respect to Ag/AgCl electrode in the working solution.
Frequency 1.5 kHz

Reference: R. Parsons and S. Trasatti. Trans. Faraday Soc. 65 (1969) 3314

	$C/\mu F\ cm^{-2}$						
$c_{KCl}/mol\ l^{-1}$ E/volts	0.01	0.013	0.016	0.020	0.025	0.03	0.04
-1.70	18.45	18.63	18.80	18.93	19.15	19.27	19.41
-1.60	17.50	17.67	17.71	17.84	17.98	18.09	18.25
-1.50	16.77	16.85	16.89	17.00	17.12	17.20	17.36
-1.40	16.23	16.26	16.30	16.39	16.47	16.55	16.64
-1.35	16.06	16.06	16.09	16.15	16.22	16.29	16.36
-1.30	15.91	15.91	15.93	15.98	16.04	16.09	16.14
-1.25	15.86	15.85	15.86	15.88	15.90	15.94	16.00
-1.20	15.85	15.85	15.85	15.87	15.86	15.90	15.93
-1.15	15.99	15.92	15.96	15.94	15.91	15.92	15.94
-1.10	16.17	16.10	16.12	16.09	16.03	16.02	16.05
-1.05	16.50	16.34	16.39	16.32	16.24	16.24	16.24
-1.00	16.88	16.71	16.73	16.67	16.55	16.57	16.55
-0.95	17.29	17.16	17.17	17.11	16.98	16.98	16.95
-0.90	17.81	17.62	17.66	17.60	17.45	17.46	17.43
-0.85	18.22	18.04	18.15	18.09	17.96	17.99	17.94
-0.80	18.39	18.36	18.47	18.51	18.43	18.49	18.50
-0.75	18.26	18.40	18.58	18.74	18.75	18.89	19.00
-0.70	17.65	17.97	18.36	18.65	18.82	19.09	19.38
-0.65	16.40	17.06	17.67	18.22	18.64	19.08	19.67
-0.625	15.73	16.53	17.28	18.00	18.54	-	-
-0.60	15.32	16.22	17.05	17.91	18.64	19.32	20.29
-0.585	15.17	16.25	17.17	18.07	-	-	-
-0.575	15.36	16.40	17.33	18.29	19.18	-	-
-0.55	16.47	17.33	18.47	19.41	20.31	21.15	22.40
-0.50	22.05	22.56	23.64	24.33	25.21	25.85	26.96
-0.45	29.99	29.69	30.43	30.88	31.43	31.87	32.76
-0.40	35.23	34.56	35.12	35.53	35.68	36.02	36.67
-0.35	36.70	36.23	36.72	36.94	37.21	37.47	37.93
-0.30	36.49	36.22	36.51	36.79	37.05	37.31	37.51
-0.25	36.21	35.96	36.24	36.51	36.63	36.84	37.09
-0.20	36.23	36.37	36.46	36.75	37.06	37.26	37.47
-0.15	37.78	38.10	38.35	38.92	39.29	39.40	39.71
-0.10	42.88	43.20	43.52	43.96	44.37	44.66	44.07
-0.05	53.65	54.20	55.41	55.71	56.97	57.27	58.26

Differential capacity of mercury in aqueous solutions of NaN_3. T = 25°C

Potential with respect to 0.1 M KCl calomel electrode in contact with the working solution.

Reference: C.V. D'Alkaine, E.K. Gonzalez and R. Parsons.
J. Electroanal. Chem. 32 (1971) 57.

$c/mol\ l^{-1}$	0.008		0.0225		0.033	
$\sigma/\mu C\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$
18	-0.063	59.35	-0.096	60.13	-0.111	60.56
17	-0.081	54.48	-0.114	55.96	-0.128	56.49
16	-0.100	49.42	-0.132	51.60	-0.147	52.03
15	-0.121	45.44	-0.153	46.80	-0.167	47.28
14	-0.144	42.07	-0.175	43.10	-0.189	43.29
13	-0.169	39.32	-0.199	39.73	-0.213	40.25
12	-0.195	36.94	-0.225	37.27	-0.239	37.67
11	-0.223	35.32	-0.253	35.75	-0.266	36.13
10	-0.252	34.18	-0.281	34.69	-0.294	34.99
9	-0.281	33.75	-0.310	34.22	-0.323	34.47
8	-0.311	33.84	-0.339	34.28	-0.352	34.53
7	-0.340	34.15	-0.368	34.69	-0.381	35.05
6	-0.369	34.66	-0.397	35.31	-0.409	35.72
5	-0.398	35.15	-0.425	35.94	-0.437	36.48
4	-0.427	34.88	-0.453	36.11	-0.464	36.76
3	-0.456	33.67	-0.481	35.47	-0.491	36.62
2	-0.486	30.98	-0.509	33.87	-0.519	35.64
1	-0.521	27.17	-0.540	31.34	-0.548	33.52
0	-0.562	24.13	-0.574	27.78	-0.579	29.78
-1	-0.613	16.76	-0.613	23.60	-0.614	26.89
-2	-0.670	17.06	-0.658	20.74	-0.655	23.15
-3	-0.728	17.68	-0.708	19.52	-0.700	21.31
-4	-0.783	18.35	-0.760	18.91	-0.749	19.91
-5	-0.838	17.98	-0.814	18.25	-0.800	18.78
-6	-0.895	17.27	-0.870	17.61	-0.855	17.30
-7	-0.954	16.73	-0.927	17.06	-0.912	16.74
-8	-1.014	16.31	-0.986	16.59	-0.970	16.28
-9	-1.076	15.94	-1.048	16.16	-1.031	15.92
-10	-1.140	15.67	-1.110	15.82	-1.093	15.69
-11	-1.204	15.52	-1.174	15.62	-1.156	15.60
-12	-1.268	15.55	-1.238	15.59	-1.220	15.66
-13	-1.332	15.69	-1.302	15.67	-1.284	15.79
-14	-1.396	15.89	-1.366	15,83	-1.348	16.01
-15	-1.458	16.16	-1.428	16.07	-1.411	16.28
-16	-1.519	16.41	-1.490	16.34	-1.473	16.59
-17	-1.580	16.65	-1.551	16.65	-1.534	16.92
-18	-1.640	16.91	-1.610	16.93	-1.594	17.24

Capacity of mercury in aqueous NaN_3 (cont.)

$c/mol\ l^{-1}$	0.0443		0.0570	
$\sigma/\mu C\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$
18	-0.122	61.27	-0.131	62.05
17	-0.139	57.37	-0.147	57.77
16	-0.157	52.95	-0.165	53.19
15	-0.177	48.21	-0.185	48.52
14	-0.199	44.06	-0.207	44.44
13	-0.222	40.76	-0.230	41.13
12	-0.248	38.08	-0.255	38.41
11	-0.275	36.26	-0.282	36.46
10	-0.303	35.11	-0.310	35.33
9	-0.332	34.65	-0.339	34.84
8	-0.361	34.73	-0.367	34.98
7	-0.389	35.27	-0.396	35.47
6	-0.417	36.08	-0.423	36.41
5	-0.445	36.89	-0.451	37.15
4	-0.472	37.24	-0.477	37.53
3	-0.498	37.14	-0.504	37.53
2	-0.526	36.37	-0.531	36.85
1	-0.554	34.85	-0.558	35.56
0	-0.584	31.09	-0.588	31.96
-1	-0.617	29.58	-0.619	31.54
-2	-0.654	25.44	-0.654	27.12
-3	-0.695	22.76	-0.693	24.27
-4	-0.741	20.30	-0.737	21.77
-5	-0.791	19.37	-0.785	19.96
-6	-0.844	18.33	-0.836	18.75
-7	-0.900	17.53	-0.891	17.77
-8	-0.958	16.90	-0.949	17.07
-9	-1.018	16.38	-1.008	16.48
-10	-1.080	16.00	-1.070	16.09
-11	-1.143	15.76	-1.132	15.82
-12	-1.207	15.62	-1.196	15.66
-13	-1.271	15.66	-1.260	15.66
-14	-1.335	15.78	-1.324	15.76
-15	-1.398	15.98	-1.387	15.96
-16	-1.460	16.23	-1.449	16.21
-17	-1.521	16.56	-1.510	16.52
-18	-1.580	16.88	-1.570	16.86

Capacity of mercury in aqueous NaN$_3$ Ccont.)

c/mol l^{-1}	0.0677		0.086		0.140	
σ/μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
18	-0.136	62.68	-0.146	64.42	-0.164	65.48
17	-0.153	58.16	-0.163	59.27	-0.180	60.56
16	-0.171	53.61	-0.180	54.01	-0.197	55.22
15	-0.190	48.85	-0.200	49.49	-0.216	50.25
14	-0.212	44.76	-0.221	45.27	-0.237	46.03
13	-0.235	41.33	-0.244	41.95	-0.260	42.42
12	-0.260	38.52	-0.268	38.88	-0.284	39.54
11	-0.287	36.67	-0.295	36.96	-0.310	37.61
10	-0.314	35.43	-0.322	35.77	-0.337	36.34
9	-0.343	35.02	-0.351	35.38	-0.365	35.91
8	-0.371	35.16	-0.379	35.48	-0.393	35.97
7	-0.400	35.64	-0.407	35.98	-0.420	36.54
6	-0.427	36.57	-0.434	36.75	-0.447	37.23
5	-0.454	37.28	-0.461	37.52	-0.474	37.93
4	-0.481	37.71	-0.488	38.05	-0.500	38.58
3	-0.508	37.76	-0.514	38.20	-0.526	38.95
2	-0.534	37.12	-0.540	37.76	-0.552	38.87
1	-0.562	36.17	-0.567	37.22	-0.578	38.31
0	-0.590	32.40	-0.595	33.44	-0.604	35.00
-1	-0.620	33.08	-0.623	-	-0.633	-
-2	-0.655	28.08	-0.656	29.96	-0.662	32.82
-3	-0.693	25.11	-0.692	26.86	-0.694	29.64
-4	-0.735	22.43	-0.731	23.89	-0.731	26.66
-5	-0.782	20.41	-0.776	21.51	-0.771	23.53
-6	-0.832	19.04	-0.824	19.80	-0.816	21.13
-7	-0.887	17.92	-0.876	18.41	-0.866	19.26
-8	-0.944	17.16	-0.932	17.44	-0.919	18.07
-9	-1.003	16.57	-0.991	16.75	-0.976	17.14
-10	-1.064	16.12	-1.051	16.27	-1.036	16.52
-12	-1.190	15.68	-1.177	15.76	-1.160	15.86
-14	-1.318	15.78	-1.304	15.81	-1.286	15.86
-16	-1.443	16.22	-1.429	16.22	-1.411	16.27
-18	-1.564	16.87	-1.550	16.85	-1.532	16.87
-20	-1.680	17.57	-1.666	17.54	-1.648	17.59

Capacity of mercury in aqueous NaN_3 (cont.)

c/mol l^{-1}	0.237		0.467		0.716	
σ/μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
18	-0.183	66.30	-0.206	68.61	-0.222	70.07
17	-0.198	61.04	-0.222	63.55	-0.237	64.54
16	-0.215	55.96	-0.238	57.75	-0.253	59.43
15	-0.234	51.07	-0.256	53.52	-0.270	54.59
14	-0.255	47.29	-0.276	49.21	-0.289	50.27
13	-0.277	43.58	-0.297	45.43	-0.310	46.39
12	-0.300	40.62	-0.320	42.30	-0.332	43.42
11	-0.326	38.53	-0.344	39.87	-0.356	40.88
10	-0.352	37.04	-0.370	38.00	-0.381	38.70
9	-0.380	36.38	-0.396	36.98	-0.408	37.31
8	-0.407	36.31	-0.424	36.43	-0.435	36.66
7	-0.434	36.68	-0.451	36.59	-0.462	36.58
6	-0.462	37.34	-0.478	37.06	-0.489	36.94
5	-0.488	38.02	-0.505	37.77	-0.516	37.51
4	-0.514	38.69	-0.531	38.53	-0.543	38.24
3	-0.540	39.35	-0.557	39.26	-0.569	39.06
2	-0.565	39.60	-0.582	39.63	-0.594	39.60
1	-0.590	39.78	-0.608	40.13	-0.619	39.63
0	-0.616	36.48	-0.633	37.08	-0.645	37.13
-1	-0.643	-	-0.659	-	-0.669	43.13
-2	-0.671	35.58	-0.685	-	-0.692	40.27
-3	-0.701	31.95	-0.714	34.82	-0.724	35.98
-4	-0.734	28.90	-0.743	32.30	-0.752	33.81
-5	-0.771	25.74	-0.776	29.46	-0.783	31.25
-6	-0.812	22.84	-0.812	26.42	-0.817	28.35
-7	-0.858	20.65	-0.852	23.69	-0.854	25.57
-8	-0.909	18.90	-0.897	21.28	-0.895	23.05
-9	-0.964	17.67	-0.946	19.23	-0.941	20.69
-10	-1.022	16.83	-1.000	17.77	-0.992	18.82
-12	-1.144	16.03	-1.118	16.40	-1.105	16.96
-14	-1.270	15.94	-1.242	16.14	-1.225	16.38
-16	-1.394	16.35	-1.365	16.45	-1.347	16.63
-18	-1.514	16.96	-1.484	17.12	-1.465	17.26
-20	-1.629	17.72	-1.598	17.89	-1.578	18.02

Capacity of mercury in aqueous NaN_3 (cont.)

c/mol l^{-1}	0.905		1.192		1.400	
σ/μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
18	-0.230	70.52	-0.240	71.31	-0.247	71.44
17	-0.245	65.19	-0.255	65.99	-0.262	66.21
16	-0.261	60.18	-0.271	60.89	-0.278	61.12
15	-0.278	55.52	-0.288	56.09	-0.295	56.41
14	-0.297	51.05	-0.306	51.84	-0.313	52.45
13	-0.317	47.22	-0.326	48.08	-0.333	48.48
12	-0.339	44.09	-0.348	44.88	-0.354	45.28
11	-0.362	41.58	-0.371	42.07	-0.377	42.44
10	-0.387	39.21	-0.395	39.54	-0.402	39.78
9	-0.413	37.69	-0.421	38.07	-0.427	38.17
8	-0.440	36.91	-0.448	37.12	-0.454	37.30
7	-0.467	36.63	-0.475	36.77	-0.481	36.92
6	-0.495	36.86	-0.502	36.82	-0.508	36.90
5	-0.521	37.32	-0.529	37.18	-0.535	37.22
4	-0.548	38.06	-0.556	37.84	-0.562	37.82
3	-0.574	38.91	-0.582	38.79	-0.588	38.82
2	-0.600	39.60	-0.608	39.58	-0.613	39.58
1	-0.625	40.19	-0.633	39.98	-0.638	39.69
0	-0.650	37.20	-0.658	37.16	-0.664	37.20
-1	-0.677	38.12	-0.685	38.46	-0.690	38.62
-2	-0.703	38.01	-0.711	38.56	-0.716	38.78
-3	-0.730	36.85	-0.737	37.45	-0.742	37.70
-4	-0.758	34.70	-0.765	35.54	-0.769	35.98
-5	-0.788	32.29	-0.794	33.37	-0.798	33.87
-6	-0.820	29.49	-0.825	30.83	-0.829	31.48
-7	-0.856	26.62	-0.859	27.98	-0.862	28.79
-8	-0.896	24.09	-0.897	25.41	-0.898	26.20
-9	-0.939	21.67	-0.938	22.99	-0.939	23.74
-10	-0.988	19.64	-0.984	20.86	-0.983	21.48
-12	-1.097	17.40	-1.088	18.10	-1.084	18.54
-14	-1.215	16.68	-1.202	17.08	-1.196	17.38
-16	-1.335	16.79	-1.320	17.01	-1.312	17.22
-18	-1.452	17.37	-1.436	17.54	-1.427	17.66
-20	-1.565	18.12	-1.547	18.28	-1.538	18.38

Differential capacity of a mercury electrode in D_2O solutions of KF and KNO_3 T = 25°C

Potentials with respect to an aqueous 0.1 mol l^{-1} calomel electrode.

Reference: R. Parsons, R.M. Reeves and P.N. Taylor. J. Electroanal. Chem. 50 (1974) 14.

	KF			KNO_3			
c/mol l^{-1}	0.20	0.50	1.00	0.10	0.20	0.50	1.00
E/volts	C/μF cm^{-2}			C/μF cm^{-2}			
-0.050	28.11	28.99	30.44	23.20	23.83	24.58	26.55
-0.100	26.79	27.32	28.35	22.37	22.74	23.43	24.58
-0.150	25.97	26.36	27.10	21.87	22.14	22.52	23.26
-0.200	25.76	26.19	26.62	21.99	22.07	22.10	22.46
-0.250	25.48	26.32	26.79	22.59	22.48	22.16	22.15
-0.300	25.06	26.27	27.11	23.66	23.35	22.71	22.34
-0.350	24.66	26.07	27.26	25.16	24.68	23.79	23.00
-0.400	24.10	25.86	27.12	26.91	26.35	25.24	23.98
-0.450	23.43	25.41	26.74	28.53	28.12	26.94	25.41
-0.500	22.69	24.78	26.18	29.34	29.52	28.72	27.03
-0.550	22.05	23.94	25.21	28.82	29.93	30.01	28.59
-0.600	21.43	23.03	24.09	26.96	29.08	30.33	29.58
-0.650	20.84	21.93	22.70	24.37	27.03	29.32	29.77
-0.700	20.12	20.55	21.42	22.14	24.49	27.28	28.73
-0.750	19.38	19.78	20.10	20.34	22.11	24.81	27.04
-0.800	18.55	18.79	19.04	19.07	20.21	22.39	24.73
-0.850	17.82	17.93	18.04	18.03	18.75	20.25	22.39
-0.900	17.14	17.18	17.27	17.22	17.66	18.66	20.39
-0.950	16.58	16.60	16.66	16.57	16.87	17.47	18.77
-1.000	16.13	16.17	16.21	16.11	16.35	16.67	17.65
-1.050	15.80	15.85	15.90	15.74	15.89	16.14	16.85
-1.100	15.62	15.66	15.72	15.52	15.70	15.85	16.34
-1.150	15.51	15.57	15.65	15.48	15.58	15.68	16.08
-1.200	15.51	15.57	15.69	15.48	15.60	15.65	15.99
-1.250	15.58	15.69	15.78	15.59	15.88	15.85	16.13
-1.300	15.73	15.83	15.97	15.77	15.90	16.02	16.25
-1.350	15.97	16.08	16.22	16.00	16.14	16.22	16.45
-1.400	16.23	16.38	16.53	16.25	16.41	16.54	15.73
-1.450	16.58	16.73	16.90	16.62	16.78	16.90	17.08
-1.500	16.99	17.14	17.28	16.97	17.17	17.35	17.45
-1.550	17.43		17.75	17.43	17.61		17.88
-1.600	17.91		18.24	17.92			

Differential capacity of a mercury electrode in D_2O solutions of KCl and $NaClO_4$

T = 25°C

Potentials with respect to an aqueous 0.1 mol l^{-1} KCl calomel electrode.

Reference: R. Parsons, R.M. Reeves and P.N. Taylor. J. Electroanal. Chem. 50 (1974) 14.

	KCl					$NaClO_4$
c/mol l^{-1}	0.090	0.180	0.30	0.60	1.00	1.00
E/volts			C/μF cm^{-2}			C/μF cm^{-1}
-0.150	45.64	51.21	56.40	70.49	83.65	22.72
-0.175						22.49
-0.200	39.47	42.19	44.33	50.85	55.06	22.41
-0.225	37.72	39.82	41.46			22.37
-0.250	36.67	38.30	39.11	42.71	44.17	22.42
-0.275	37.18	37.26	37.72	40.52	41.44	22.59
-0.300	36.00	36.77	36.99	39.05	39.55	22.83
-0.325	36.05	36.68	36.73	38.23	38.36	23.13
-0.350	36.30	36.84	36.82	37.90	37.82	23.43
-0.375	36.43	37.16	37.23	37.98	37.77	23.80
-0.400	36.53	37.46	37.65	38.38	37.94	24.29
-0.425	36.31	37.61	38.14	38.79	38.47	24.75
-0.450	35.57	37.41	38.40	39.25	39.14	25.32
-0.475	34.32	36.74	38.28	39.49	39.67	25.89
-0.500	32.50	35.54	37.73	39.45	39.99	26.39
-0.525	30.24	33.91	36.51	38.90	40.04	26.82
-0.550	27.89	31.72	34.81	37.74	39.30	27.22
-0.575			32.60	36.22	38.18	
-0.600	23.59	27.29	30.42	34.16	36.23	27.66
-0.650	20.85	23.44	25.99	29.74	33.41	27.46
-0.700	19.28	20.79	22.60	25.46	27.71	26.56
-0.750	18.23	19.05	20.10	22.07	23.89	25.10
-0.800	17.54	17.91	18.52	19.81	21.06	23.19
-0.850	16.88	17.09	17.41	18.22	19.07	21.23
-0.900	16.37	16.43	16.63	17.13	17.75	19.50
-0.950	15.93	15.97	16.05	16.39	16.84	18.08
-1.000	15.62	15.61	15.65	15.93	16.26	17.08
-1.050	15.40	15.38	15.43	15.63	15.92	16.38
-1.100	15.26	15.27	15.29	15.48	15.75	15.98
-1.150	15.24	15.27	15.29	15.45	15.73	15.75
-1.200	15.28	15.34	15.36	15.54	15.83	15.68
-1.300	15.41	15.70	15.74	15.93	16.22	15.86
-1.325	15.61					15.94
-1.350	15.83	15.95	16.04	16.91	16.48	
-1.375	15.98	16.09				
-1.400		16.24			16.78	

Differential capacity of mercury in aqueous solutions of H_2SO_4 T = 25°C

Reference: J. Dec and E. Dutkiewicz (unpublished)

$c/\text{mol l}^{-1}$	2.5×10^{-3}		5×10^{-3}		0.1	
	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
	0.140	53.40	0.085	70.72	0.093	83.52
	0.190	45.01	0.080	73.00	0.120	64.86
	0.240	40.47	0.130	61.25	0.170	51.69
	0.290	38.09	0.180	48.32	0.220	45.75
	0.340	35.90	0.230	41.90	0.270	41.01
	0.390	35.22	0.280	38.74	0.320	38.15
	0.440	35.22	0.330	36.53	0.370	35.82
	0.490	35.63	0.380	34.98	0.420	34.01
	0.540	35.93	0.430	34.49	0.470	32.88
	0.590	35.40	0.480	34.65	0.520	32.13
	0.640	33.36	0.530	35.13	0.570	31.93
	0.690	29.58	0.580	35.44	0.620	31.93
	0.740	25.57	0.630	34.53	0.670	32.08
	0.790	20.77	0.680	32.17	0.720	31.75
	0.840	15.25	0.730	28.67	0.770	30.40
	0.873	11.57	0.780	24.64	0.820	27.92
	0.907	10.44	0.830	19.07	0.870	24.91
	0.940	11.25	0.880	14.19	0.920	22.54
	0.990	13.62	0.905	12.88	0.970	20.92
	1.040	15.94	0.917	12.65	1.020	19.79
	1.090	16.93	0.930	12.68	1.070	19.16
	1.140	17.17	0.980	14.31	1.120	18.26
	1.190	17.02	1.030	16.22	1.170	17.64
	1.240	16.64	1.080	17.29	1.220	17.01
	1.290	16.19	1.130	17.60	1.270	16.40
	1.340	15.71	1.180	17.41	1.320	15.84
	1.390	15.71	1.230	17.03	1.370	15.40
	1.440	14.99	1.280	16.45	1.420	15.10
	1.490	14.73	1.330	15.97	1.470	14.92
	1.540	14.55	1.380	15.59	1.520	14.79
	1.590	14.45	1.430	15.18	1.570	14.71
	1.640	14.38	1.480	14.91	1.620	14.71
	1.690	14.38	1.530	14.67		
	1.740	14.46	1.580	14.56		
	1.790	14.55	1.630	14.51		
	1.840	14.62	1.680	14.51		
	1.890	14.69	1.730	14.61		
			1.780	14.69		

Capacity in aqueous H_2SO_4 (cont.)

c/mol l^{-1}	1.0		1.75		3.5	
	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
	0.080	115.90	0.270	51.18	0.200	69.65
	0.130	62.14	0.320	36.40	0.246	62.71
	0.180	49.70	0.370	31.93	0.300	43.30
	0.230	43.14	0.420	29.16	0.350	37.58
	0.280	38.47	0.470	27.22	0.400	33.36
	0.330	34.75	0.520	25.86	0.450	30.28
	0.380	31.97	0.570	25.08	0.500	27.79
	0.430	29.71	0.620	24.85	0.550	25.75
	0.480	28.11	0.670	25.24	0.600	24.17
	0.530	27.02	0.720	26.29	0.650	22.97
	0.580	26.56	0.770	27.91	0.700	22.14
	0.630	26.72	0.820	29.97	0.750	21.69
	0.680	27.50	0.870	31.79	0.800	21.84
	0.730	29.01	0.920	32.75	0.900	23.57
	0.780	30.80	0.970	32.37	0.933	27.41
	0.830	32.44	1.020	30.61	0.950	25.83
	0.880	32.96	1.070	27.86	1.000	28.01
	0.925	31.97	1.120	25.09	1.050	30.12
	0.980	29.56	1.170	22.63	1.100	31.48
	1.030	26.68	1.220	21.08	1.200	30.65
	1.080	23.85	1.270	19.04	1.250	38.92
	1.130	21.59	1.320	17.88	1.300	26.73
	1.180	19.73	1.370	17.04	1.350	24.47
	1.230	18.43	1.420	16.49	1.400	22.44
	1.280	17.38	1.470	16.02	1.450	20.78
	1.330	16.55	1.520	15.76	1.450	19.43
	1.376	16.95	1.570	15.61	1.500	18.45
	1.430	15.63	1.620	15.57	1.550	17.77
	1.480	15.63			1.600	17.39
	1.526	15.21			1.650	17.02
	1.580	15.09			1.700	16.94
	1.616	15.13				
	1.645	15.15				

Capacity in aqueous H_2SO_4 (cont.)

c/mol l^{-1}	5.0		7.0		9.0	
	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
	0.107	75.66	0.107	108.13	0.100	82.16
	0.150	51.55	0.150	50.85	0.150	48.75
	0.200	42.02	0.200	43.58	0.200	41.97
	0.250	36.87	0.250	38.24	0.250	37.60
	0.300	32.76	0.300	34.29	0.300	34.47
	0.350	29.72	0.350	30.76	0.350	31.29
	0.400	27.16	0.400	28.13	0.400	28.38
	0.450	25.16	0.450	25.70	0.450	26.03
	0.500	23.47	0.500	23.79	0.500	23.98
	0.550	22.16	0.550	22.17	0.550	22.36
	0.600	21.16	0.600	20.90	0.600	20.89
	0.650	20.46	0.650	19.81	0.650	19.73
	0.700	20.07	0.700	19.05	0.700	18.67
	0.750	19.90	0.750	18.36	0.750	17.80
	0.800	20.05	0.800	18.02	0.800	17.08
	0.850	20.59	0.850	17.76	0.850	16.55
	0.900	21.63	0.900	17.74	0.900	16.14
	0.950	23.16	0.950	18.05	0.950	15.91
	1.000	25.18	1.000	18.79	0.982	15.84
	1.050	27.40	1.050	19.95	1.000	15.75
	1.100	29.19	1.100	21.52	1.050	15.94
	1.150	30.31	1.150	23.38	1.100	16.45
	1.200	30.36	1.198	25.46	1.150	17.20
	1.250	29.35	1.250	27.32	1.200	18.22
	1.300	27.83	1.292	28.56	1.250	19.74
	1.350	26.19	1.300	28.79	1.300	21.56
	1.402	24.03	1.350	27.40	1.350	23.36
	1.448	22.64	1.400	29.37	1.395	25.43
	1.500	21.24	1.450	29.14	1.425	26.49
	1.543	20.31	1.482	28.29		
	1.580	19.77				

Differential capacity for mercury in aqueous solutions of KPF_6

T = 25°C

Potential with respect to 0.1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} KPF_6 solution.

Reference: L.M. Baugh and R. Parsons. J. Electroanal. Chem. 40 (1972) 407.

			C/µF cm^{-2}		
c/mol l^{-1} E/volts	6.3×10^{-3}	0.010	0.016	0.025	0.040
+ 0.30		36.91	36.10	38.00	38.03
+ 0.29	30.57	33.56	32.45	33.30	33.93
+ 0.28		31.32	29.40	30.25	31.99
+ 0.27	28.34	29.46	28.36	29.20	29.38
+ 0.26	27.22	27.74	27.54	27.79	28.41
+ 0.25		26.25	26.42	27.04	27.44
+ 0.24	25.35	25.35	25.46	26.22	26.70
+ 0.23			24.93	25.40	25.95
+ 0.22	24.16	24.31	24.56	24.59	25.20
+ 0.21			24.19	23.99	24.61
+ 0.20	23.12	23.56	23.82	23.69	24.16
+ 0.19			23.45	23.39	23.71
+ 0.18	22.37	22.97	23.00	23.02	23.34
+ 0.17			22.63	22.80	23.04
+ 0.16		22.37	22.33	22.57	22.74
+ 0.15			22.11	22.35	22.45
+ 0.14	21.92	21.77	21.88	22.05	22.22
+ 0.13			21.81	21.90	21.99
+ 0.12	21.03	21.55	21.59	21.68	21.63
+ 0.09	20.73	21.18	21.29	21.23	21.33
+ 0.06	20.43	20.73	20.99	20.93	20.95
+ 0.03	20.36	20.43	20.69	20.79	20.81
+ 0.00	20.36	20.43	20.54	20.64	20.66

Capacity in aqueous KPF_6 (cont.)

	$C/\mu F\ cm^{-2}$				
$c/mol\ l^{-1}$ E/volts	0.063	0.10	0.16	0.25	0.40
+ 0.30	38.70	39.19	41.98		31.41
+ 0.29	34.24	35.31	37.52	40.57	29.09
+ 0.28	31.78	32.18	34.77	37.06	27.30
+ 0.27	29.84	30.47	32.54	34.60	26.32
+ 0.26	28.80	28.91	30.76	32.59	25.65
+ 0.25	27.69	28.09	29.64	31.10	25.20
+ 0.24	26.80	27.12	28.60	29.38	24.75
+ 0.23	25.98	26.30	27.56	28.41	24.30
+ 0.22	25.38	25.63	26.60	27.44	24.08
+ 0.21	24.79	25.03	25.78	26.47	23.71
+ 0.20	24.19	24.51	25.26	25.88	23.33
+ 0.19	23.89	24.06	24.59	25.06	23.03
+ 0.18	23.59	23.77	24.22	24.61	22.81
+ 0.17	23.30	23.39	23.77	24.16	22.58
+ 0.16	22.92	23.10	23.40	23.71	22.36
+ 0.15	22.63	22.80	23.11	23.34	22.21
+ 0.14	22.40	22.57	22.81	22.97	22.06
+ 0.13	22.18	22.35	22.44	22.74	21.91
+ 0.12	22.03	22.13	22.14	22.52	21.76
+ 0.09	21.57	21.61	21.77	21.92	21.46
+ 0.06	21.21	21.23	21.32	21.55	21.16
+ 0.03	20.84	20.93	20.95	21.35	20.94
+ 0.00	20.69	20.79	20.80	20.95	20.79

Capacity in aqueous KPF_6 (cont.)

			$C/\mu F\ cm^{-2}$		
$c/mol\ l^{-1}$ E/volts	6.3×10^{-3}	0.010	0.016	0.025	0.040
-0.72	17.52	17.82	18.31	18.70	19.09
-0.75	18.27	18.12	18.46	18.70	18.94
-0.78	18.27	18.12	18.38	18.63	18.79
-0.81	18.27	18.12	18.24	18.48	18.57
-0.84	18.12	18.12	18.09	18.18	18.27
-0.87	17.97	17.78	17.94	17.88	17.90
-0.90	17.75	17.60	17.64	17.58	17.60
-0.93	17.52	17.30	17.34	17.28	17.30
-0.96	17.00	17.00	17.04	16.99	17.00
-0.99	16.70	16.70	16.32	16.76	16.73
-1.02	16.41	16.41	16.60	16.54	16.55
-1.05	16.26	16.26	16.37	16.39	16.33
-1.08	16.18	16.18	16.23	16.24	16.11
-1.11	16.11	16.11	16.08	16.09	16.03
-1.14	16.03	16.03	16.00	15.94	15.96
-1.17	15.96	15.96	15.93	15.87	15.88
-1.20	15.88	15.88	15.85	15.87	15.88
-1.23	15.88	15.88	15.85	15.87	15.88
-1.26	15.88	15.88	15.85	15.87	15.88
-1.29	15.88	15.96	15.93	15.94	15.88
-1.32	15.96	16.03	16.00	16.02	16.03
-1.35	16.03	16.11	16.08	16.09	16.18
-1.38		16.18		16.24	16.33

Capacity in aqueous KPF_6 (cont.)

			C/μF cm^{-2}		
c/mol l^{-1} E/volts	0.063	0.10	0.16	0.25	0.40
-0.72	19.50	19.97	20.36	21.10	21.91
-0.75	19.28	19.59	18.84	20.43	21.09
-0.78	18.98	19.15	19.39	19.69	20.34
-0.81	18.61	18.77	18.95	19.16	19.59
-0.84	18.24	18.40	18.42	18.64	18.99
-0.87	17.94	18.03	17.98	18.20	18.40
-0.90	17.57	17.66	17.61	17.75	17.95
-0.93	17.27	17.28	17.24	17.37	17.50
-0.96	16.90	16.99	16.94	17.08	17.12
-0.99	16.67	16.69	16.64	16.78	16.83
-1.02	16.45	16.46	16.42	16.35	16.45
-1.05	16.30	16.32	16.27	16.33	16.38
-1.08	16.15	16.17	16.12	16.18	16.23
-1.11	16.00	16.02	15.97	16.11	16.15
-1.14	15.93	15.94	15.90	16.03	16.08
-1.17	15.85	15.94	15.90	16.03	16.00
-1.20	15.85	15.94	15.90	16.03	16.00
-1.23	15.85	15.94	15.90	16.03	16.08
-1.26	15.85	15.94	15.97	16.11	16.15
-1.29	15.93	15.99	16.05	16.18	16.23
-1.32	16.08	16.11	16.12	16.26	
-1.35	16.23	16.24	16.27	16.41	
-1.38	16.37	16.36		16.55	
-1.41		16.48		16.70	
-1.43		16.68			
-1.46		16.89			
-1.49		17.11			
-1.52		17.34			
-1.55		17.56			
-1.58		17.85			
-1.61		18.15			

Differential capacity of mercury in aqueous $HClO_4$ T = 25°C

Potential with respect to a hydrogen electrode in the working solution

Reference: R. Parsons and R. Payne, unpublished.

$c/\text{mol l}^{-1}$	0.01085		0.02713		0.05426	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
-14			-0.684	30.01		
-13	-0.873	15.09	-0.650	28.55		
-12	-0.807	15.02	-0.614	27.27		
-11	-0.740	15.08	-0.577	26.20	-0.738	15.25
-10	-0.674	15.28	-0.538	25.28	-0.673	15.48
-9	-0.609	15.61	-0.498	24.47	-0.609	15.84
-8	-0.546	16.06	-0.456	23.74	-0.547	16.35
-7	-0.485	16.62	-0.413	23.18	-0.487	16.96
-6	-0.426	17.27	-0.370	22.82	-0.429	17.73
-5	-0.369	17.84	-0.326	22.67	-0.374	18.56
-4	-0.314	18.24	-0.282	22.82	-0.322	19.49
-3	-0.259	18.18	-0.238	23.24	-0.272	20.53
-2	-0.203	17.43	-0.196	23.90	-0.224	21.90
-1	-0.144	16.51	-0.155	24.75	-0.180	23.62
0	-0.085	18.28	-0.115	25.68	-0.140	25.49
1	-0.036	22.40	-0.077	26.55	-0.102	27.03
2	0.006	25.27	-0.040	27.15	-0.065	27.90
3	0.045	26.34	-0.003	27.24	-0.030	28.12
4	0.082	26.55	0.034	26.56	0.006	27.76
5	0.120	26.16	0.073	24.97	0.043	27.03
6	0.159	25.38	0.115	22.65	0.080	26.08
7	0.199	24.53	0.162	20.52	0.119	25.09
8	0.241	23.73	0.212	19.48	0.160	24.19
9	0.283	23.07	0.264	19.08	0.202	23.51
10	0.327	22.74	0.316	18.68	0.245	23.06
11	0.371	22.63	0.371	18.12	0.289	22.88
12	0.415	22.83	0.427	17.45	0.332	22.98
13	0.459	23.24	0.485	16.80	0.376	23.31
14	0.501	23.86	0.546	16.21	0.418	23.82
15	0.542	24.59	0.609	15.71	0.459	24.50
16	0.582	25.47	0.673	15.35	0.500	24.50
17	0.621	26.52	0.739	15.13	0.539	26.11
18	0.658	27.88	0.805	15.08	0.576	27.05
19	0.692	29.70	0.871	15.11	0.612	28.12
20	0.725	31.30			0.647	29.47

Capacity of mercury in aqueous $HClO_4$ (cont.)

c/mol l^{-1}	0.1085		0.2170		0.5426	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-11	-0.740	15.26				
-10	-0.675	15.51	-0.677	15.64	-0.614	31.79
-9	-0.611	15.90	-0.614	16.13	-0.582	30.26
-8	-0.549	16.47	-0.553	16.35	-0.548	28.97
-7	-0.490	17.21	-0.495	17.88	-0.513	27.94
-6	-0.433	18.16	-0.441	19.20	-0.477	27.07
-5	-0.380	19.34	-0.391	20.86	-0.439	26.30
-4	-0.330	20.80	-0.346	22.79	-0.401	25.59
-3	-0.284	22.53	-0.303	24.79	-0.361	24.94
-2	-0.241	24.40	-0.265	26.62	-0.320	24.35
-1	-0.202	26.19	-0.228	28.00	-0.279	23.00
0	-0.164	27.61	-0.193	28.89	-0.237	23.50
1	-0.129	28.51	-0.159	29.29	-0.194	23.29
2	-0.094	28.82	-0.124	29.20	-0.151	23.24
3	-0.059	28.62	-0.090	28.69	-0.108	23.41
4	-0.024	28.02	-0.055	27.93	-0.066	23.78
5	0.012	27.13	-0.018	26.95	-0.024	24.37
6	0.050	26.11	0.020	25.99	0.016	25.14
7	0.089	25.12	0.059	25.04	0.056	26.03
8	0.130	24.24	0.100	24.24	0.093	26.97
9	0.171	23.55	0.141	23.65	0.130	27.90
10	0.214	23.14	0.184	23.28	0.165	28.66
11	0.258	22.98	0.227	23.18	0.200	29.19
12	0.301	23.08	0.270	23.29	0.234	29.39
13	0.344	23.39	0.313	23.59	0.268	29.19
14	0.387	23.90	0.355	24.08	0.302	28.56
15	0.428	24.53	0.396	24.69	0.338	27.44
16	0.468	25.29	0.436	25.39	0.376	25.90
17	0.507	26.11	0.475	26.16	0.416	24.03
18	0.545	27.01	0.512	27.00	0.459	22.04
19	0.581	28.01	0.549	27.96	0.507	20.11
20	0.616	29.23	0.584	29.10	0.559	18.47
21	0.649	30.75	0.617	30.48	0.615	17.20
22					0.674	16.36
23					0.737	15.81

Capacity of mercury in aqueous $HClO_4$ (cont.)

$c/mol\ l^{-1}$	0.9307		1.861		3.723	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-10			-0.708	19.93		
-9	-0.627	18.59	-0.660	21.51		
-8	-0.576	20.20	-0.615	23.17	-0.698	26.26
-7	-0.528	22.08	-0.573	24.77	-0.661	27.09
-6	-0.485	23.97	-0.534	26.18	-0.624	27.70
-5	-0.445	25.73	-0.497	27.32	-0.588	28.04
-4	-0.407	27.20	-0.461	28.12	-0.553	28.14
-3	-0.371	28.28	-0.426	28.57	-0.517	27.98
-2	-0.336	28.94	-0.391	28.71	-0.481	27.58
-1	-0.302	29.20	-0.356	28.51	-0.445	26.98
0	-0.267	29.09	-0.320	28.05	-0.407	26.25
1	-0.233	28.66	-0.284	27.38	-0.368	25.41
2	-0.197	28.00	-0.247	26.58	-0.328	24.53
3	-0.161	27.15	-0.209	25.69	-0.287	23.64
4	-0.124	26.25	-0.169	24.82	-0.244	22.86
5	-0.085	25.34	-0.128	24.00	-0.199	22.17
6	-0.045	24.54	-0.086	23.34	-0.154	21.65
7	-0.003	23.90	-0.043	22.87	-0.107	21.30
8	0.039	23.45	0.001	22.59	-0.060	21.13
9	0.082	23.22	0.046	22.51	-0.013	21.15
10	0.125	23.18	0.090	22.60	0.035	21.31
11	0.168	23.30	0.134	22.84	0.081	21.60
12	0.211	23.57	0.177	23.20	0.127	22.01
13	0.253	23.97	0.220	23.65	0.172	22.54
14	0.294	24.47	0.262	24.17	0.216	23.18
15	0.335	25.01	0.303	24.76	0.258	23.90
16	0.374	25.63	0.343	25.45	0.300	24.75
17	0.413	26.30	0.382	26.18	0.339	25.77
18	0.450	27.05	0.419	27.02	0.377	26.96
19	0.487	27.93	0.456	28.00	0.413	28.49
20	0.522	28.96	0.491	29.25	0.447	30.39
21	0.555	30.27	0.524	30.78	0.479	32.11
22	0.588	31.75	0.556	32.14	0.510	32.80
23			0.586	33.27	0.540	33.85
24					0.569	35.26
25					0.599	33.78

Differential capacity of a mercury electrode in aqueous Na_2SO_4 T = 25°C

Potential with respect to 0.1 M KCl calomel electrode dipping into 0.1 M Na_2SO_4

Reference: R. Payne, unpublished

c/mol l^{-1}	0.005		0.01		0.02		0.05	
σ/ μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
23							-0.197	46.47
22			-0.185	44.30	-0.197	40.63	-0.219	44.61
21			-0.208	42.41	-0.221	42.76	-0.242	42.94
20	-0.220	40.80	-0.232	40.76	-0.245	41.05	-0.265	41.46
19	-0.245	39.12	-0.257	39.32	-0.270	39.63	-0.290	40.19
18	-0.271	37.85	-0.283	38.06	-0.296	38.47	-0.315	39.12
17	-0.298	36.81	-0.310	37.06	-0.322	37.52	-0.341	38.20
16	-0.325	36.04	-0.337	36.39	-0.349	36.78	-0.367	37.53
15	-0.353	35.55	-0.365	35.83	-0.376	36.31	-0.394	37.07
14	-0.382	35.36	-0.393	35.58	-0.404	36.05	-0.421	36.83
13	-0.410	35.32	-0.421	35.51	-0.432	36.06	-0.449	36.84
12	-0.438	35.64	-0.449	35.74	-0.459	36.25	-0.476	37.04
11	-0.466	35.94	-0.477	36.14	-0.487	36.55	-0.503	37.25
10	-0.494	36.23	-0.504	36.53	-0.514	37.00	-0.529	37.59
9	-0.521	36.57	-0.532	36.79	-0.541	37.07	-0.556	37.71
8	-0.548	36.45	-0.559	36.75	-0.568	36.92	-0.582	37.45
7	-0.576	35.96	-0.586	36.17	-0.595	36.32	-0.609	36.80
6	-0.604	34.80	-0.614	35.06	-0.623	35.13	-0.637	35.51
5	-0.634	33.08	-0.643	33.32	-0.652	33.37	-0.666	33.72
4	-0.665	30.85	-0.674	31.04	-0.683	31.16	-0.696	31.57
3	-0.699	28.12	-0.708	28.33	-0.717	28.54	-0.729	29.10
2	-0.737	24.69	-0.745	25.11	-0.754	25.57	-0.765	26.54
1	-0.781	20.17	-0.789	20.88	-0.796	22.09	-0.805	23.91
0	-0.840	14.16	-0.843	16.47	-0.845	18.79	-0.849	21.61
-1	-0.914	14.01	-0.907	15.72	-0.901	17.59	-0.897	20.27
-2	-0.979	16.54	-0.967	17.29	-0.957	18.15	-0.947	19.82
-3	-1.037	17.94	-1.023	18.32	-1.011	18.73	-0.997	19.65
-4	-1.092	18.43	-1.077	18.62	-1.064	18.85	-1.049	19.36
-5	-1.146	18.30	-1.131	18.43	-1.117	18.57	-1.101	18.88
-6	-1.201	17.88	-1.186	17.96	-1.172	18.06	-1.155	18.24
-7	-1.258	17.34	-1.243	17.39	-1.228	17.44	-1.211	17.59
-8	-1.316	17.03	-1.301	16.80	-1.287	16.86	-1.269	16.97
-9	-1.377	16.29	-1.362	16.30	-1.347	16.36	-1.329	16.43
-10	-1.439	15.9	-1.424	15.91	-1.409	15.97	-1.390	16.03
-11	-1.502	15.68	-1.487	15.68	-1.472	15.73	-1.453	15.78
-12	-1.566	15.58	-1.551	15.61	-1.536	15.64	-1.517	15.68
-13	-1.630	15.61	-1.615	15.65	-1.600	15.66	-1.581	15.70
-14	-1.694	15.73	-1.679	15.79	-1.663	15.80	-1.644	15.82
-15	-1.757	15.97	-1.742	16.02	-1.726	16.02	-1.707	16.06
-16			-1.803	16.34	-1.788	16.32	-1.769	16.36
-17					-1.849	16.69	-1.829	16.52
-18					-1.908	17.07	-1.890	16.46
-19							-1.949	17.58

Capacity in Na_2SO_4 in H_2O (cont.)

	0.1		0.2		0.5		1.0	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
26							-0.203	46.38
25							-0.225	44.78
24							-0.248	43.51
23	-0.214	46.25	-0.227	45.61			-0.271	42.56
22	-0.236	44.57	-0.249	44.11	-0.273	43.42	-0.295	41.79
21	-0.258	43.05	-0.272	42.82	-0.296	42.47	-0.319	41.22
20	-0.282	41.71	-0.296	41.72	-0.320	41.68	-0.343	40.80
19	-0.306	40.56	-0.320	40.76	-0.344	41.03	-0.368	40.51
18	-0.331	39.57	-0.345	39.94	-0.369	40.51	-0.393	40.34
17	-0.357	38.77	-0.370	39.28	-0.364	40.14	-0.417	40.34
16	-0.383	38.18	-0.396	38.81	-0.419	39.91	-0.442	40.41
15	-0.409	37.82	-0.422	38.52	-0.444	39.82	-0.469	41.08
14	-0.436	37.67	-0.448	38.40	-0.469	39.84	-0.491	41.88
13	-0.462	37.70	-0.474	38.47	-0.494	39.98	-0.515	41.93
12	-0.489	37.88	-0.500	38.64	-0.519	40.20	-0.539	41.97
11	-0.515	38.15	-0.526	38.86	-0.544	40.39	-0.563	42.01
10	-0.541	38.38	-0.551	38.95	-0.568	40.49	-0.586	42.11
9	-0.567	38.41	-0.577	38.92	-0.563	40.36	-0.610	41.95
8	-0.593	38.10	-0.603	38.53	-0.618	39.86	-0.634	41.43
7	-0.620	37.31	-0.629	37.69	-0.643	38.95	-0.658	40.51
6	-0.647	35.99	-0.656	36.35	-0.669	37.61	-0.684	39.13
5	-0.675	34.18	-0.684	34.58	-0.697	35.86	-0.710	37.42
4	-0.706	32.06	-0.714	32.55	-0.725	33.86	-0.737	35.47
3	-0.738	29.76	-0.746	30.42	-0.756	31.90	-0.766	33.47
2	-0.773	27.52	-0.780	28.44	-0.788	30.03	-0.797	31.56
1	-0.811	25.37	-0.816	26.61	-0.822	28.38	-0.830	29.81
0	-0.852	23.50	-0.855	25.02	-0.859	24.88	-0.864	28.20
-1	-0.896	22.11	-0.896	23.67	-0.897	25.45	-0.900	26.66
-2	-0.942	21.21	-0.939	22.52	-0.938	24.11	-0.939	25.21
-3	-0.990	20.54	-0.985	21.49	-0.980	22.84	-0.980	23.78
-4	-1.040	19.89	-1.033	20.56	-1.025	21.59	-1.023	22.39
-5	-1.091	19.19	-1.082	19.63	-1.073	20.39	-1.069	21.06
-6	-1.144	18.44	-1.134	18.74	-1.123	19.29	-1.118	19.84
-7	-1.199	17.70	-1.189	17.91	-1.177	18.32	-1.170	18.76
-8	-1.257	-	-1.246	17.18	-1.233	17.49	-1.225	17.86
-9	-1.318	-	-1.305	16.58	-1.291	16.83	-1.282	17.13
-10	-1.380	16.04	-1.367	16.15	-1.351	16.34	-1.341	16.60
-11	-1.443	15.80	-1.429	15.87	-1.413	16.03	-1.402	16.25
-12	-1.506	15.70	-1.492	15.74	-1.476	15.88	-1.464	16.09
-13	-1.570	15.71	-1.556	15.75	-1.539	15.89	-1.527	16.06
-14	-1.633	15.84	-1.619	15.88	-1.602	16.00	-1.589	16.15
-15	-1.696	16.07	-1.682	16.10	-1.664	16.23	-1.650	16.35
-16	-1.758	16.38	-1.743	16.39	-1.725	16.52	-1.711	16.64
-17	-1.818	16.73	-1.804	16.74	-1.785	16.86	-1.770	16.98
-18	-1.877	17.12					-1.829	17.38
-19	-1.935	17.56						

Differential capacity of mercury in aqueous KNO_3

Potential with respect to 0.1 M calomel electrode dipping into 0.1 M KNO_3

Reference: R.Payne, J. Electrochem. Soc., 113 (1966) 999.

c / mol l^{-1}	0.01		0.025		0.05		0.1	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
21							0.193	37.72
20							0.166	34.66
19	0.193	32.73	0.191	32.64	0.153	32.41	0.136	32.25
18	0.161	30.49	0.160	30.42	0.121	30.30	0.104	30.23
17	0.127	28.59	0.126	28.53	0.087	28.50	0.070	28.45
16	0.091	27.03	0.090	26.99	0.051	26.96	0.033	26.91
15	0.053	25.69	0.052	25.65	0.013	25.66	-0.005	25.61
14	0.013	24.67	0.012	24.64	-0.027	24.64	-0.045	24.63
13	-0.028	23.96	-0.029	23.94	-0.068	23.93	-0.086	23.93
12	-0.070	23.56	-0.071	23.55	-0.110	23.58	-0.128	23.57
11	-0.113	23.44	-0.114	23.44	-0.153	23.55	-0.170	23.55
10	-0.155	23.74	-0.156	23.75	-0.195	23.86	-0.213	23.85
9	-0.197	24.25	-0.198	24.28	-0.237	24.49	-0.254	24.50
8	-0.237	25.10	-0.239	25.13	-0.277	25.42	-0.294	25.45
7	-0.276	26.07	-0.278	26.11	-0.315	26.54	-0.333	26.60
6	-0.314	27.05	-0.315	27.11	-0.352	27.71	-0.369	27.86
5	-0.350	27.84	-0.352	27.86	-0.387	28.77	-0.405	29.06
4	-0.386	28.25	-0.387	28.25	-0.422	29.56	-0.438	30.02
3	-0.422	27.80	-0.423	27.76	-0.455	29.87	-0.471	30.58
2	-0.459	26.15	-0.460	26.06	-0.489	29.45	-0.504	30.59
1	-0.499	22.85	-0.501	22.68	-0.523	28.29	-0.537	29.95
0	-0.548	18.30	-0.548	23.11	-0.560	26.34	-0.571	28.67
-1	-0.607	16.37	-0.600	20.65	-0.600	24.09	-0.607	28.87
-2	-0.667	17.36	-0.650	19.55	-0.643	22.15	-0.646	24.83
-3	-0.722	18.33	-0.701	19.36	-0.689	20.84	-0.688	22.91
-4	-0.776	18.64	-0.753	19.18	-0.739	19.98	-0.733	21.29
-5	-0.830	18.49	-0.806	18.81	-0.789	19.26	-0.782	19.99
-6	-0.885	18.06	-0.860	18.29	-0.823	18.53	-0.833	18.96
-7	-0.941	17.51	-0.915	17.67	-0.898	17.82	-0.887	18.06
-8	-0.999	16.96	-0.973	17.09	-0.955	17.19	-0.944	17.32
-9	-1.059	16.51	-1.033	16.59	-1.014	16.67	-1.002	16.74
-10	-1.120	16.17	-1.094	16.23	-1.075	16.28	-1.063	16.33
-11	-1.182	15.96	-1.156	16.02	-1.137	16.04	-1.125	16.09
-12	-1.245	15.89	-1.218	15.93	-1.199	15.96	-1.187	15.99
-13	-1.308	15.96	-1.281	15.99	-1.262	16.01	-1.250	16.03
-14	-1.370	16.13	-1.343	16.17	-1.324	16.17	-1.312	16.18
-15	-1.432	16.40	-1.405	16.42	-1.385	16.43	-1.373	16.45
-16	-1.492	16.76	-1.465	16.76	-1.446	16.78	-1.433	16.79
-17	-1.551	17.20	-1.524	17.19	-1.505	17.21	-1.492	17.22
-18	-1.609	17.68	-1.581	17.67	-1.562	17.68	-1.549	17.70
-19					-1.618	18.21	-1.605	18.24
-20							-1.659	18.82

Capacity of mercury in aqueous KNO_3 (cont.)

c/mol l^{-1}	0.2		0.5		1.0	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
21					0.132	39.13
20	0.147	34.61			0.106	36.68
19	0.117	32.34	0.099	33.41	0.078	34.40
18	0.085	30.35	0.068	31.35	0.048	32.30
17	0.051	28.60	0.035	29.49	0.016	30.36
16	0.015	27.04	0.000	27.82	-0.018	28.63
15	-0.023	25.76	-0.037	26.42	-0.054	27.15
14	-0.062	24.74	-0.076	25.28	-0.092	25.92
13	-0.103	24.04	-0.116	24.44	-0.131	24.95
12	-0.145	23.62	-0.158	23.89	-0.172	24.29
11	-0.188	23.55	-0.200	23.66	-0.213	23.92
10	-0.230	23.83	-0.242	23.76	-0.255	23.89
9	-0.271	24.44	-0.284	24.19	-0.297	24.17
8	-0.312	25.35	-0.324	24.95	-0.338	24.77
7	-0.350	26.51	-0.364	25.98	-0.377	25.66
6	-0.387	27.79	-0.401	27.19	-0.416	26.74
5	-0.422	29.08	-0.437	28.49	-0.452	27.96
4	-0.456	30.20	-0.472	29.72	-0.487	29.14
3	-0.489	31.00	-0.505	30.73	-0.521	30.22
2	-0.521	31.36	-0.537	31.44	-0.553	31.03
1	-0.553	31.16	-0.568	31.73	-0.585	31.54
0	-0.585	30.40	-0.600	31.53	-0.617	31.59
-1	-0.619	29.09	-0.632	30.83	-0.649	31.20
-2	-0.654	27.29	-0.665	29.61	-0.681	30.33
-3	-0.692	25.25	-0.700	28.00	-0.715	29.05
-4	-0.733	23.25	-0.737	26.07	-0.750	27.43
-5	-0.788	21.42	-0.777	24.01	-0.788	25.59
-6	-0.827	19.88	-0.820	22.03	-0.829	23.61
-7	-0.879	18.64	-0.868	20.26	-0.873	21.72
-8	-0.934	17.68	-0.919	18.84	-0.921	20.05
-9	-0.992	16.97	-0.974	17.73	-0.973	18.69
-10	-1.052	16.48	-1.031	16.99	-1.028	17.70
-11	-1.113	16.19	-1.091	16.53	-1.085	17.04
-12	-1.175	16.06	-1.152	16.31	-1.145	16.68
-13	-1.237	16.09	-1.214	16.24	-1.205	16.56
-14	-1.299	16.24	-1.275	16.36	-1.265	16.60
-15	-1.360	16.50	-1.336	16.58	-1.325	16.78
-16	-1.420	16.85	-1.395	16.90	-1.384	17.07
-17	-1.479	17.26	-1.454	17.30	-1.442	17.44
-18	-1.536	17.75	-1.511	17.74		
-19						
-20						

Differential capacity of mercury in aqueous KF solution T = 0° C

Potential with respect to 1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} KF solution.

Reference: D.J. Schiffrin, Trans. Faraday Soc., 67 (1971) 3318.

c/mol l^{-1}	0.01		0.02		0.04		0.06		0.10	
σ/μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-12	-1.1366	15.93	-1.1237	15.96	-1.1090	16.02	-1.1006	16.02	-1.0922	16.05
-10	-1.0124	16.43	-0.9999	16.50	-0.9857	16.54	-0.9775	16.60	-0.9693	16.63
-8	-0.8948	17.70	-0.8829	17.82	-0.8692	17.90	-0.8612	17.96	-0.8534	18.02
-6	-0.7869	19.41	-0.7757	19.57	-0.7627	16.69	-0.7551	19.80	-0.7478	19.93
-4	-0.6873	20.46	-0.6768	20.69	-0.6650	21.05	-0.6581	21.30	-0.6519	21.68
-2	-0.5870	18.74	-0.5738	19.58	-0.5700	20.74	-0.5652	21.51	-0.5617	22.54
0	-0.4620	14.14	-0.4676	16.91	-0.4704	19.74	-0.4713	21.32	-0.4738	23.08
2	-0.3418	20.25	-0.3596	21.03	-0.3742	22.42	-0.3811	23.37	-0.3896	24.60
4	-0.2530	24.24	-0.2727	24.62	-0.2905	25.13	-0.2997	25.52	-0.3110	26.09
6	-0.1731	25.50	-0.1937	25.74	-0.2129	21.11	-0.2227	26.28	0.2353	26.57
8	-0.0949	25.59	-0.1164	25.90	-0.1364	26.10	-0.1466	26.22	-0.1599	26.43
10	-0.0168	25.60	-0.0390	25.85	-0.0596	26.06	-0.0701	26.14	-0.0840	26.33
12	0.0606	26.15	0.0379	26.33	0.0166	26.55	0.0059	26.60	-0.0086	26.80
14	0.1354	27.59	0.1122	27.66	0.0903	27.87	0.0796	27.79	0.0647	28.00
16					0.1596	30.07	0.1491	29.91	0.1341	29.77

Capacity of mercury in aqueous KF (cont.)

c/mol l^{-1}	0.2		0.35		0.5		0.7		1.0	
σ/μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-12	-1.0828	16.06	-1.0765	16.07	-1.0692	16.17	-1.0595	16.29	-1.-512	16.42
-10	-0.9599	16.63	-0.9538	16.65	-0.9472	16.77	-0.9386	16.95	-0.9315	17.15
-8	-0.8441	18.07	-0.8381	18.12	-0.8324	18.27	-0.8254	18.55	-0.8196	18.81
-6	-0.7390	20.11	-0.7335	20.29	-0.7290	20.57	-0.7238	20.96	-0.7195	21.33
-4	-0.6445	22.23	-0.6405	22.73	-0.6375	23.19	-0.6342	23.76	-0.6316	24.26
-2	-0.5580	23.90	-0.5566	24.89	-0.5556	25.65	-0.5545	26.40	-0.5537	27.07
0	-0.4765	25.17	-0.4789	26.52	-0.4805	27.48	-0.4816	28.38	-0.4828	29.22
2	-0.3988	26.28	-0.4050	27.57	-0.4094	28.52	-0.4126	29.41	-0.4157	30.29
4	-0.3239	27.02	-0.3330	27.86	-0.3397	28.67	-0.3498	29.51	-0.3500	30.38
6	-0.2500	27.04	-0.2608	27.53	-0.2694	28.18	-0.2764	28.86	-0.2834	29.66
8	-0.1756	26.65	-0.1874	26.94	-0.1975	27.45	-0.2060	28.00	-0.2149	28.68
10	-0.1002	26.48	-0.1126	26.65	-0.1240	27.03	-0.1338	27.43	-0.1442	27.99
12	-0.0250	26.87	-0.0378	26.96	-0.0501	27.25	-0.0608	27.55	-0.0726	27.98
14	0.0479	28.07	0.0350	28.10	0.0221	28.30	0.0107	28.51	-0.0019	28.78
16	0.1170	30.07	0.1040	30.05	0.0909	30.04	0.0789	30.39	0.0657	30.50

Differential capacity of mercury in aqueous KF solution T = 25°C

Potential with respect to 1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} KF solution.

Reference: D.J. Schiffrin, Trans. Faraday Soc., 67 (1971) 3318.

c /mol l^{-1}	0.01		0.1		1.0	
σ/μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-12	-1.1854	15.86	-1.1314	16.00	-1.0888	16.33
-10	-1.0598	16.11	-1.0074	16.33	-0.9675	16.77
-8	-0.9382	16.88	-0.8879	17.23	-0.8518	17.95
-6	-0.8234	17.98	-0.7761	18.56	-0.7455	19.81
-4	-0.7147	18.60	-0.6719	19.79	-0.6499	22.12
-2	-0.6039	16.94	-0.5726	20.36	-0.5640	24.48
0	-0.4660	12.90	-0.4752	20.98	-0.4855	26.47
2	-0.3347	18.72	-0.3830	22.66	-0.4120	27.89
4	-0.2398	22.97	-0.2986	24.67	-0.3414	28.68
6	-0.1567	24.91	-0.2198	26.00	-0.2721	28.94
8	-0.0780	25.91	-0.1442	26.83	-0.2031	29.04
10	-0.0021	26.74	-0.0706	27.53	-0.1344	29.21
12			0.0009	28.49	-0.0665	29.79

Differential capacity of mercury in aqueous KF solution T = 15°C

Potential with respect to 1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} KF solution.

Reference: D.J. Schiffrin, Trans. Faraday Soc., 67 (1971) 3318.

c/mol l^{-1}	0.06		0.1		0.2		0.5		1.0	
σ/μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-12	-1.1289	15.99	-1.1177	16.01	-1.1052	16.04	-1.0861	16.18	-1.0761	16.33
-10	-1.0047	16.36	-0.9940	16.44	-0.9817	16.48	-0.9640	16.70	-0.9551	16.88
-8	-0.8860	17.44	-0.8760	17.55	-0.8640	17.65	-0.8482	18.00	-0.8407	18.21
-6	-0.7758	18.93	-0.7665	19.07	-0.7556	19.34	-0.7425	19.94	-0.7365	20.33
-4	-0.6736	20.07	-0.6656	20.46	-0.6567	21.09	-0.6474	22.20	-0.6437	22.84
-2	-0.5744	20.17	-0.5698	21.15	-0.5651	22.54	-0.5616	24.38	-0.5607	25.36
0	-0.4743	20.00	-0.4762	21.74	-0.4787	23.74	-0.4825	26.13	-0.4850	27.42
2	-0.3287	22.23	-0.3871	23.36	-0.3967	25.06	-0.4078	27.30	-0.4139	28.68
4	-0.2936	24.64	-0.3049	25.17	-0.3187	26.15	-0.3354	27.85	-0.3449	29.17
6	-0.2147	25.92	-0.2272	26.18	-0.2432	26.75	-0.2637	27.89	-0.2763	29.05
8	-0.1384	26.44	-0.1515	26.62	-0.1688	26.97	-0.1919	27.77	-0.2071	28.79
10	-0.0634	26.89	-0.0768	26.99	-0.0950	27.24	-0.1198	27.82	-0.1374	28.63
12	0.0101	27.69	-0.0035	27.73	-0.0223	27.89	-0.0484	28.29	-0.0671	28.48
14	0.0808	29.00	0.0678	29.07	0.0480	29.20	0.0209	33.14	0.0002	29.94

Differential capacity on mercury in 7.6 M HI at different temperatures. Given are the capacity in μFcm^{-2}, potential in mV vs Pt(H2)/H$^+$.

Reference: B.G. Dekker, M. Sluyters - Rehbach and J.H. Sluyters
J.Electroanal. Chem. 21 (1969) 137

	C/ u Fcm^{-2}					
T/°C -E/V	-35	-25	-15	-5	+5	+15
500	233		300			
520	146		171	220		
530						241
540	114		126			199
550			114			173
560	97.1			128		146
570					116	131
575			94.8			
580	85.6			103.5		119
590						111
600	76.3	79.2	82.3	89.0		104.2
610						98.9
620				81.0		92.4
625			73.2			
630	65.9				82.6	
640				74.4		
650	60.9	63.2	66.3	71.9	76.0	
660			64.0	70.2	73.0	
670			62.2			
675	56.5	58.7				
680			60.4			
690			58.8			
700	53.7	55.6	57.4			
710			56.4			
720	51.8	53.4	55.1			
730			54.1			
740	51.1	52.0	53.5			
750		51.1	52.8			
760	49.2	50.8	50.4			
780	48.5	49.3				
800	47.0					
820	46.1					

Differential capacity on mercury in 1 M HI at different temperatures. Given are the capacity in $\mu F\ cm^{-2}$, potential in mV vs $Pt(H_2)/H^+$

Reference: B.G. Dekker, M. Sluyters-Rehbach and J.H. Sluyters. J. Electroanal. Chem. 21 (1969) 137.

$C/\mu F\ cm^{-2}$

T/°C	0	15	25	T/°C	0	15	25
-E/mV				-E/mV			
200		270 -	264 -	725		38.2	
250	155 -	147 -	149 -	740			35.6
300	111 -	107.5	108.3	750	36.6	34.5	
350	87.6	85.0	85.7	760			33.1
400	71.8	70.7	71.4	775	32.9	31.3	
450	62.4	62.3	62.7	780			30.7
475	60.0	58.0	59.9	800	29.5	28.3	28.1
500	58.6	57.6	57.7	810			27.5
525	57.3	56.2	56.2	820			26.0
550	56.2	55.0	54.8	825	26.4	25.5	
575		53.3		830			25.3
600	54.4	52.0	51.4	840			24.2
620			49.6	850	24.1	23.4	23.7
625		49.8		860			22.6
640			47.8	870			22.0
650	49.9	47.6		875	21.8	21.5	
660			45.7	880			21.4
675		44.7		890			21.0
680			43.7	900	20.4	20.0	20.3
700	43.7	41.4	40.9	925	19.1	19.0	
720			38.3				

Differential capacity on mercury in 7.5 M HCl at different temperatures. Given are the capacity in $\mu F\ cm^{-2}$, potential in mV vs $Pt(H_2)/H_+$ in the same solution.

Reference: B.G. Dekker, M. Sluyters-Rehbach and J.H. Sluyters.
J. Electroanal. Chem. 21 (1969) 137.

	$C/\mu F\ cm^{-2}$									
T /°C	-35	-25	-15	-5	5	10	15	25	45	65
E/mV										
-50		59.4	61.7	65.8	69.2	71.4	71.1	79.5	101.3	113.4
-100	43.6	45.1	47.0	48.6	50.3	51.7	52.9	55.3	63.3	72.8
-150									40.4	
-200	30.9	31.7	32.6	33.6	34.5	35.1	35.8	36.9	31.9	43.9
-250										
-300	25.1	25.8	26.7	27.3	28.1	28.5	29.0	29.7	30.0	33.7
-350										31.5
-400	23.8	24.6	25.3	25.9	26.4	26.9	27.2	27.8	29.2	30.4
-425									29.1	
-450				26.7	27.1				29.2	29.9
-500	26.9	27.5	27.9	28.4	28.6	27.9	29.0	29.3	29.0	30.0
-550				30.8	30.8	30.9		30.9	30.8	30.4
-600	34.0	34.1	33.8	33.5	33.1	33.1	33.0	32.5	31.7	30.8
-650				35.8	34.9	34.7		33.6	32.3	30.9
-700	39.9	38.9	38.0	36.8	35.6	35.3	34.7	33.8	32.1	
-725										30.0
-750		38.4		36.2	34.9			33.1	31.3	30.0
-800	37.3	36.3	35.4	34.4	33.4	33.0	32.7	31.7	30.2	
-825									29.4	
-850		33.3				30.8	30.5	29.8		
-900	30.7	30.3	29.8	29.5	28.9	28.7	28.6	28.3		
-950				27.2	27.0	26.9	26.7			
-1000	25.8	25.7	25.7	25.5						
-1050		24.3	24.3							
-1100	23.2	23.2								
-1150	22.4									

Differential capacity on mercury in 1 M HCl at different temperatures. Given are the capacity in μF cm^{-2}, potential in mV vs Pt(H_2)H_+

Reference: B.G. Dekker, M. Sluyters-Rehbach and J.H. Sluyters. J. Electroanal. Chem. 21 (1969) 137.

	C/μF cm^{-2}			
T/°C	0	15	25	45
E/mV				
+200	75.0	84.7	91.2	117.2
+150	50.1	54.5	57.8	65.2
+100	40.0	43.2	45.4	50.7
+ 50	35.2	37.0	39.6	43.8
0	33.2	35.5	37.0	40.2
-850	33.2	35.3	36.9	39.2
-100	35.1	36.8	37.7	39.5
-150	38.5	39.5		
-200	42.1	41.8	41.1	40.4
-300	42.6	40.2	37.9	35.6
-400	32.3	30.6	28.7	27.4
-500	23.2	22.7	21.9	21.4
-600	18.7	18.6	18.3	18.1
-700	16.7	16.7	16.7	16.6
-800	15.9	16.0	15.9	15.9
-850				15.7
-900	15.8	15.8	15.8	15.7
-915				15.7
-925			15.8	15.7
-935				15.7
-950		15.9	15.8	15.7
-965			15.9	15.7
-975			15.9	15.7
-985			15.9	15.8
-1000	16.0	16.0	16.0	
-1015		16.0	16.0	
-1025	16.1	16.1	16.0	
-1035		16.1	16.0	
-1050	16.3	16.2	16.1	
-1065	16.3	16.3		
-1075	16.4	16.4		
-1085	16.5	16.4		
-1100	16.5	16.4		
-1115	16.7			
-1125	16.8			
-1135	16.8			

Differential capacity on mercury in HCl at 25^{o}C; given are the capacity in μF cm^{-2}, the potential in mV vs Pt(H_2)/H_+ in the same solution.

Reference: B.G. Dekker, M. Sluyters-Rehbach and J.H. Sluyters. J. Electroanal. Chem. 21 (1969) 137.

C/μF cm^{-2}

c/mol l^{-1}			HCl			
	1	2	4	6	7.5	10
E/mV						
+200	91.2					
+175	68.25					
+150	57.8	105.7				
+125	50.5	76.5				
+100	45.4	62.9	169.4			
+ 75	42.25	54.2	124.5			
+ 50	39.6	47.5	78.1			
+ 25	38.0	43.25	62.75			
0	37.0	40.0	53.5	97.2		
- 25	37.0	37.7	47.6	71.25		
- 50	36.9	36.4	42.5	58.8	79.5	
- 75	37.25	35.5	39.1	50.6	64.5	
-100	37.7	35.0	36.3	44.7	55.3	102.9
-125	38.7	35.0	34.5	40.5	49.0	77.25
-150	39.75	35.7	33.3	37.3	44.25	62.9
-175	40.6	36.3	32.2	34.8	40.15	53.9
-200	41.1	37.1	31.7	32.7	36.9	47.1
-225	41.0	38.1	32.0	31.4	34.5	42.7
-250	40.5	39.3	32.2	30.3	32.5	39.2
-275	39.5	39.75	32.7	29.6	31.0	36.25
-300	37.9	40.0	33.1	29.0	29.7	33.8
-325	35.75	40.0	34.1	28.85	28.65	31.9
-350	33.5	39.3.	34.9	28.7	28.1	30.2
-375	31.25	37.5	36.0	29.1	27.8	29.0
-400	28.7	35.9	36.7	29.8	27.8	27.7
-425	26.75	33.75	37.9	30.25	27.8	27.0
-450	25.0	32.1	37.4	31.0	28.2	26.4

Capacity in HCl (continued)

	$C/\mu F\ cm^{-2}$					
	HCl					
$c/mol\ l^{-1}$	1	2	4	6	7.5	10
E/mV						
-475	23.25	29.75	37.0	31.9	28.6	26.05
-500	21.9	27.9	36.1	32.7	29.3	25.7
-525	20.75	26.2	35.1	33.35	30.0	25.7
-550	19.75	24.6	34.2	33.8	30.9	25.7
-575	19.0	23.2	32.4	34.0	31.65	26.0
-600	18.3	21.8	31.0	34.1	32.5	26.3
-625	17.75	20.75	29.4	33.6	33.05	27.0
-650	17.35	20.0	27.7	33.1	33.6	27.7
-675	17.0	19.15	26.5	32.35	33.6	28.4
-700	16.7	18.6	25.2	31.3	33.8	29.1
-725	16.5	18.0	24.2	30.15	33.5	29.9
-750	16.25	17.5	23.0	28.9	33.1	30.5
-775	16.0	17.3	22.1	27.9	32.4	31.4
-800	15.9	17.1	21.2	26.8	31.7	
-825	15.9	16.75	20.75	25.8	30.8	
-850	15.9	16.5	20.0	24.9	29.8	
-875	15.8	16.5	19.4		29.0	
-900	15.8	16.4	18.7		28.3	
-925	15.8					
-950	15.8					
-975	16.0					
-1000	16.0					
-1025	16.1					
-1050	16.1					

Differential capacity in HCl in water at 0°C;[a] 25°C, 40°C and 60°C.[b] Given are potential in V, capacity in μF cm^{-2}. The reference electrode (aqueous saturated calomel electrode) was constantly at the same temperature as the system studied.

References: a) S. Minc and M. Jurkiewicz-Herbich. J. Electroanal. Chem. 40 (1972) 229.
b) S. Minc and M. Jurkiewicz-Herbich. J. Electroanal. Chem. 34 (1972) 351.

	0.01 M				0.1 M				0.2 M				0.4 M			
	C/μF cm^{-2}				C/μF cm^{-2}				0/μF cm^{-2}				C/μF cm^{-2}			
T/°C	0	25	40	60	0	25	40	60	0	25	40	60	0	25	40	60
-E/V																
0.200	33.16	33.84	42.00	41.81	36.10	38.60	39.70	42.10	37.20	39.47	40.18	41.08	34.45	38.12	40.81	42.44
0.250		33.84														
0.300	35.06	31.93	41.40	40.50	40.31	40.13	39.30	40.70	49.61	40.97	40.91	40.39	35.15	36.87	39.59	40.24
0.350	31.36	26.14	40.72	38.41	42.66	40.30	39.51	39.51	44.15	42.47	41.32	39.22	37.66	38.12	40.80	40.16
0.400	27.48	21.28	37.91	34.10	42.65	38.61	38.80	37.80	45.90	42.43	39.68	35.77	40.84	39.50	40.41	39.25
0.450	19.03	16.71	33.52	28.52	38.80	34.70	34.32	32.80	44.45	39.81	36.23	31.77	42.90	39.92	38.61	37.09
0.500	16.53	15.56	27.21	23.90	33.00	30.70	30.29	28.71	39.70	35.00	31.23	27.46	41.85	37.92	34.98	33.46
0.550	16.63	16.28	23.00	20.00	27.81	26.29	24.91	23.66	33.60	29.61	27.77	23.62	37.44	33.96	30.61	29.17
0.600	17.46	17.04	20.40	18.10	24.20	23.00	21.70	21.35	28.49	25.63	23.45	21.03	31.67	29.33	26.45	25.24
0.650	17.92						19.80						26.61	25.17	22.98	22.16
0.700	17.83	17.34	18.11	17.00	20.49	19.50	18.71	18.10	22.32	20.78	19.50	17.93	22.82	21.87	20.53	20.00
0.800	16.53	16.57	17.10		18.50	17.76	17.20	16.80	19.30	18.45	17.54	16.46	18.59	18.13	17.63	17.28
0.900	15.61		16.70	15.87	17.09	16.70	16.40	16.14	17.80	17.18	16.54	15.69	16.65	16.33	16.28	15.98
1.000				15.51	16.40		15.91	15.80	16.90	16.50			15.68	15.42	15.59	15.43
1.100																

Differential capacity in HCl in water at various temperatures (continued)

	0.6 M C/μF cm^{-2}				1.0 M C/μF cm^{-2}				1.4 M C/μF cm^{-2}			
T/°C	0	25	40	60	0	25	40	60	0	25	40	60
-E/V												
0.200	38.00	39.71	41.30	44.90	36.80	38.63	40.78	43.95	38.00	44.78	45.87	48.27
0.250												
0.300	38.05	39.60	40.60	41.31	33.61	36.15	37.71	40.63	32.73	36.91	37.96	40.53
0.350	41.10	41.70	41.41	41.40	35.00	37.83	38.70	40.71	32.81	36.52	37.49	40.04
0.400	44.72	43.21	41.90	41.20	37.21	39.60	40.13	40.88	34.60	37.13	37.79	40.03
0.450	47.20	42.90	40.60	40.38	41.15	40.93	41.08	40.50	37.52	38.87	39.06	40.08
0.500	46.91	39.80	37.52	36.81	43.30	40.44	40.39	38.11	40.70	40.35	40.21	39.75
0.550	42.30	35.11	33.02	32.50	43.03	37.39	37.36	34.58	42.80	40.56	40.20	37.90
0.600	36.15	30.00	28.30	28.50	40.22	33.10	33.03	30.38	41.98	38.87	39.23	34.81
0.650									38.37	35.52	37.57	31.07
0.700	25.70	23.00	21.80	22.02	29.42	24.47	24.98	23.36	33.30	32.26	30.77	27.41
0.800	20.60	19.15	18.60	18.61	22.00	19.47	20.04	19.45	24.50	26.65	24.00	21.93
0.900	18.37	17.33	16.95	16.95	18.35	17.03	17.53	17.44	19.56	20.22	20.13	18.76
1.000	17.34	16.55	16.22	16.55	16.74	15.93	16.54	16.60	17.18	17.87	17.79	17.20
1.100				16.15			16.10	16.20				

Mercury/ solution interface - LiCl in water at 25°C, ± 0.25°C.
Capacity-potential data. Given are the capacity in μF cm^{-2}, potential in V.
Reference electrode: 1 mol l^{-1} aq. KCl - calomel electrode.

Reference: J. Jastrzebska. Rocz. Chem. 44 (1970) 1779.

	C/μF cm^{-2}					
c/mol l^{-1}	0.05	0.1	0.5	0.8	1.0	2.0
-E/V						
0.1	36.93	40.37	48.61	52.25	55.47	62.40
0.2	34.20	36.13	38.07	40.09	40.81	41.59
0.3	34.86	37.01	38.45	38.23	38.65	37.77
0.4	29.37	34.45	41.23	41.81	42.16	40.43
0.5	22.45	26.50	37.00	40.16	41.32	42.28
0.6	18.43	20.84	27.63	30.49	31.87	35.41
0.7	17.61	18.92	21.82	22.83	23.44	27.05
0.8	16.84	16.91	17.65	18.90	19.26	21.33
0.9	16.06	16.16	16.48	17.22	17.63	18.15
1.0	15.30	15.54	15.66	15.69	16.17	16.91

Differential capacity on mercury in aqueous solutions of KCl at 25°C. Given are the potential in V and capacitance in μF cm^{-2} as a function of surface charge σ.

Reference: P. Holmqvist. J. Electroanal. Chem. 68 (1976) 31.

	0.1 M KCl		1.0 M KCl	
C/μC cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
-24	1.762	21.77	1.714	21.78
-22	1.666	20.12	1.619	20.13
-20	1.563	18.73	1.517	18.91
-18	1.453	17.65	1.410	18.16
-16	1.338	16.96	1.298	17.54
-14	1.217	16.36	1.182	16.62
-12	1.093	16.02	1.060	16.40
-10	0.969	16.12	0.940	17.11
-8	0.848	16.81	0.827	19.01
-6	0.733	18.22	0.727	22.06
-4	0.630	20.62	0.643	27.61
-2	0.540	24.25	0.576	33.60
0	0.465	29.73	0.520	39.40
2	0.404	35.42	0.472	42.96
4	0.351	39.38	0.426	44.00
6	0.300	39.95	0.380	43.33
8	0.250	39.15	0.333	41.28
10	0.198	38.24	0.283	39.35
12	0.146	38.60	0.232	39.18
14	0.096	40.83	0.184	42.65
16	0.050	45.95	0.139	49.94
18	0.012	55.53	0.101	58.39
20	-0.020	76.36	0.072	71.86

Capacity data for mercury in 0.1 M KCl solution
Reference electrode - SCE (NaCl solution)

Reference: K.G. Baikerikar and Robert S. Hansen. J. Colloid Interface Sci. 61 (1977) 239

$-E/V$	$C/\mu F\,cm^{-2}$	$-E/V$	$C/\mu F\,cm^{-2}$
1.600	19.19	0.550	23.92
1.550	18.61	0.525	25.57
1.500	18.11	0.500	27.65
1.450	17.65	0.475	30.15
1.400	17.23	0.450	32.78
1.350	16.87	0.425	35.25
1.300	16.58	0.400	37.23
1.250	16.30	0.375	38.68
1.200	16.13	0.350	39.46
1.150	16.01	0.325	39.60
1.100	15.99	0.300	39.25
1.050	16.80	0.275	38.69
1.000	16.20	0.250	38.08
1.950	16.49	0.225	37.60
0.900	16.86	0.200	37.38
0.850	17.37	0.175	37.47
0.800	17.98	0.150	38.13
0.750	18.68	0.125	39.38
0.700	19.50	0.100	41.36
0.650	20.44	0.075	44.46
0.625	21.04	0.050	49.13
0.600	21.76	0.025	56.52
0.575	22.68	0.0	70.62

Differential capacity on mercury electrode in KCl - given are the capacity in μF cm^{-2}, potential in V.

The 1 mol l^{-1} calomel electrode at a given temperature and at the same KCl concentration as the system studied was used as a reference.

Reference: S. Minc and J. Jastrzebska. Electrochim. Acta 10 (1965) 965.

0.0005 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.382	13.68	0.371	20.25	0.357	14.08	0.300	15.13
0.433	8.52	0.421	8.97	0.407	9.05	0.400	6.49
0.452	6.66	0.441	6.92	0.427	6.89	0.420	5.25
0.462	5.86	0.461	5.49	0.447	5.85	0.440	4.81
0.472	5.77	0.471	5.38	0.457	5.43	0.450	4.63
0.482	5.33	0.481	5.38	0.467	5.30	0.460	4.55
0.492	5.42	0.491	5.49	0.472	5.28	0.470	4.42
0.502	5.52	0.501	5.74	0.487	5.37	0.480	4.49
0.512	5.70	0.521	6.62	0.497	5.38	0.490	4.61
0.532	6.36	0.621	11.00	0.507	5.67	0.500	4.90
0.632	10.73	0.721	12.55	0.607	10.81	0.600	8.67
0.732	12.33	0.821	12.72	0.707	13.05	0.700	10.73
0.832	12.44			0.807	13.63	0.800	11.46

Differential capacity on mercury in KCl (continued)

0.003 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.373	21.38	0.360	20.00	0.454	11.59	0.350	20.1
0.423	19.31	0.410	14.81	0.474	10.24	0.450	10.7
0.473	13.58	0.460	10.42	0.484	9.77	0.470	9.95
0.493	10.18	0.480	9.80	0.494	9.30	0.480	9.60
0.503	9.42	0.490	9.64	0.504	9.70	0.490	9.26
0.513	9.32	0.500	9.25	0.514	10.00	0.500	9.58
0.523	9.96	0.510	9.80	0.524	10.30	0.510	9.70
0.533	10.00	0.520	10.00	0.554	11.70	0.520	9.80
0.573	11.94	0.530	10.50			0.550	9.95
0.673	14.00	0.560	11.86				

Differential capacity on mercury in KCl (continued)

0.005 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.385	25.30	0.377	29.40	0.367	23.58	0.359	24.44
0.435	18.80	0.427	20.50	0.417	17.88	0.409	17.37
0.485	13.80	0.447	16.31	0.437	15.43	0.429	15.70
0.495	13.25	0.467	14.63	0.457	13.67	0.449	13.35
0.510	13.20	0.477	13.72	0.467	13.16	0.459	12.93
0.515	13.34	0.487	13.01	0.477	12.60	0.469	12.50
0.525	13.52	0.497	12.75	0.487	12.45	0.479	12.20
0.535	13.71	0.507	12.60	0.497	12.20	0.489	11.95
0.585	15.90	0.517	12.74	0.507	12.20	0.504	11.80
0.685	18.18	0.527	12.96	0.517	12.34	0.509	11.82
0.785	18.04	0.577	14.84	0.527	12.54	0.519	11.97
0.885	17.35	0.677	17.70	0.567	13.84	0.559	13.04
0.985	16.35	0.777	18.02	0.667	17.19	0.659	16.50
		0.877	17.63	0.767	18.10	0.759	17.83
		0.977	16.90	0.867	17.74	0.859	17.60

Differential capacity on mercury in KCl (continued)

0.0075 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.394	29.00	0.386	26.40	0.378	27.00	0.370	28.22
0.444	21.50	0.436	20.70	0.428	20.40	0.420	20.56
0.474	18.00	0.456	18.60	0.478	16.00	0.450	17.56
0.494	16.60	0.476	18.60	0.488	15.58	0.470	15.95
0.504	16.30	0.486	16.38	0.498	15.30	0.480	15.37
0.514	16.06	0.496	15.96	0.508	15.00	0.490	15.11
0.524	16.03	0.506	15.57	0.518	14.98	0.500	14.90
0.534	16.03	0.516	15.48	0.523	14.80	0.510	14.75
0.544	16.30	0.526	15.48	0.528	14.88	0.520	14.64
0.554	16.60	0.536	15.50	0.538	15.16	0.530	14.66
0.594	17.85	0.546	15.64	0.548	15.32	0.540	14.75
0.694	19.30	0.556	15.82	0.578	16.00	0.550	14.95
0.794	18.90	0.586	16.68	0.678	18.65	0.570	15.30
0.894	17.60	0.686	19.90	0.778	19.32	0.670	18.00
		0.786	19.13	0.878	18.70	0.770	19.09
		0.886	18.31			0.870	18.89

Differential capacity on mercury in KCl (continued)

0.01 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.400	30.50	0.403	30.20	0.393	27.61	0.378	28.33
0.450	23.40	0.453	23.40	0.444	22.26	0.428	21.48
0.500	18.00	0.503	17.94	0.493	17.17	0.478	16.60
0.510	17.60	0.523	16.95	0.503	16.59	0.488	16.10
0.520	17.31	0.533	16.61	0.513	16.21	0.498	15.74
0.530	17.06	0.543	16.48	0.523	16.12	0.508	15.38
0.540	17.16	0.553	16.48	0.533	15.94	0.518	15.24
0.550	17.14	0.563	16.63	0.538	15.93	0.528	15.15
0.560	17.30	0.583	16.84	0.553	15.97	0.538	15.07
0.580	17.79	0.603	17.55	0.563	16.03	0.548	15.27
0.600	18.23	0.703	19.34	0.593	16.70	0.558	15.32
0.700	19.85	0.803	19.20	0.693	18.87	0.578	15.80
0.800	18.06	0.903	18.32	0.793	19.37	0.678	18.08
						0.778	18.87
						0.878	18.08

Differential capacity on mercury in KCl (continued)

0.0125 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.405	30.05	0.400	30.93	0.390	29.63	0.384	29.13
0.455	23.50	0.450	24.07	0.440	22.80	0.434	22.30
0.505	18.76	0.500	18.71	0.490	18.06	0.484	17.53
0.525	17.84	0.520	17.72	0.510	17.22	0.504	16.61
0.535	17.65	0.540	17.36	0.520	17.00	0.514	16.27
0.545	17.66	0.550	17.41	0.530	16.80	0.524	16.11
0.555	17.69	0.560	17.47	0.540	16.66	0.534	16.07
0.565	17.86	0.600	18.28	0.550	16.84	0.544	16.02
0.605	18.61	0.700	19.59	0.560	16.93	0.554	16.20
0.705	19.65	0.800	19.40	0.590	17.53	0.584	16.60
0.805	18.60			0.690	19.20	0.684	18.47
0.905	17.04			0.790	19.47	0.784	19.12
				0.890	18.86	0.884	18.90

Differential capacity on mercury in KCl (continued)

0.017 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
-E/V	C/µF cm^{-2}	-E/V	C/µF cm^{-2}	-E/V	C/µF cm^{-2}	-E/V	C/µF cm^{-2}
0.412	31.63	0.417	28.95	0.405	28.17	0.398	28.07
0.462	25.12	0.467	22.93	0.455	27.93	0.448	21.83
0.512	20.03	0.517	18.60	0.505	28.53	0.498	17.67
0.532	19.18	0.537	17.61	0.515	17.50	0.518	16.65
0.542	19.09	0.547	17.37	0.525	17.06	0.538	16.09
0.552	18.92	0.557	17.28	0.535	16.80	0.548	16.01
0.562	18.90	0.567	17.13	0.545	16.63	0.558	15.96
0.572	19.03	0.577	17.18	0.555	16.51	0.568	15.90
0.582	19.04	0.587	17.24	0.565	16.50	0.578	16.04
0.592	19.08	0.597	17.29	0.575	16.54	0.598	16.21
0.612	19.40	0.617	17.48	0.585	16.59	0.698	17.53
0.712	19.40	0.717	18.20	0.605	16.88		
0.812	18.31			0.705	17.94		

Differential capacity on mercury in KCl (continued)

$0.02\ \text{mol}\ l^{-1}$ KCl

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.412	30.45	0.410	29.30	0.404	28.53	0.398	28.60
0.462	24.88	0.460	24.00	0.454	23.09	0.448	22.73
0.512	20.49	0.510	19.87	0.504	19.17	0.498	18.52
0.532	19.50	0.530	19.01	0.524	18.27	0.518	17.66
0.552	19.05	0.550	18.48	0.544	17.84	0.538	17.21
0.562	19.04	0.560	18.49	0.554	17.70	0.548	17.03
0.572	18.97	0.570	18.45	0.564	17.69	0.558	16.99
0.582	19.10	0.580	18.45	0.574	17.73	0.568	16.97
0.592	19.15	0.590	18.50	0.584	17.77	0.578	16.90
0.612	19.16	0.610	18.53	0.604	17.86	0.588	17.12
0.662	19.16	0.710	19.10	0.704	18.59	0.598	17.74
0.712	19.02	0.810	18.47	0.804	18.55	0.648	18.11
0.812	17.96	0.910	17.41	0.904	17.74	0.798	18.48
0.912	16.46					0.898	17.88

Differential capacity on mercury in KCl (continued)

0.025 mol l^{-1} KCl

0°C		20°C		40°C		60°C	
$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}
0.470	25.69	0.415	33.39	0.408	30.99	0.403	29.29
0.520	22.09	0.465	26.93	0.458	24.82	0.453	24.10
0.540	20.99	0.515	21.89	0.508	20.82	0.503	20.09
0.550	20.86	0.535	20.68	0.548	19.45	0.523	19.08
0.560	20.69	0.555	20.12	0.568	19.33	0.563	18.45
0.570	20.59	0.575	19.93	0.588	19.24	0.583	18.40
0.580	20.51	0.585	19.88	0.608	19.37	0.593	18.40
0.590	20.47	0.605	19.95	0.648	19.53	0.603	18.58
0.610	20.46	0.615	19.95	0.688	19.75	0.643	19.06
0.620	20.36	0.635	20.00	0.708	19.78	0.663	18.93
0.640	20.32	0.715	19.96	0.758	19.51	0.683	19.06
0.660	20.25	0.775	19.57	0.808	19.19	0.703	19.06
0.680	20.21	0.815	19.25	0.908	18.24	0.753	18.88
0.700	20.12	0.915	17.96	1.008	17.87	0.853	18.48
0.720	19.95	1.015	17.17			0.953	17.92
0.770	19.35						
0.820	18.50						
0.920	17.18						
1.020	16.44						

Differential capacity on mercury electrode in 0.1 M NaCl at 25°. Given are the capacity in $\mu F\ cm^{-2}$, potential in V (Orion fluoride). reversible electrode).

Reference: K. Doblhofer and D.M. Mohilner. J. Phys. Chem. 75 (1971) 1698.

$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
1.6015	16.8	1.1401	15.8	0.6803	20.9	0.2885	36.2
1.5801	16.8	1.1204	15.8	0.6600	21.3	0.2734	36.3
1.5600	16.8	1.1001	15.8	0.6401	21.8	0.2589	36.4
1.5401	16.7	1.0805	15.9	0.6200	22.4	0.2439	36.8
1.5200	16.6	1.0600	15.9	0.6001	23.2	0.2292	37.4
1.5001	16.6	1.0400	16.0	0.5800	24.3	0.2143	37.8
1.4800	16.6	1.0201	16.1	0.5601	25.8	0.1997	38.3
1.4601	16.5	1.0000	16.2	0.5400	27.5	0.1847	38.9
1.4402	16.5	0.9803	16.3	0.5200	29.2	0.1704	39.9
1.4200	16.4	0.9603	16.5	0.5081	31.2	0.1562	41.0
1.4000	16.3	0.9401	16.6	0.5052	31.4	0.1424	42.2
1.3804	16.3	0.9202	16.8	0.4773	34.3	0.1288	44.5
1.3601	16.2	0.9001	17.1	0.4629	36.0	0.1147	47.1
1.3400	16.2	0.8800	17.4	0.4461	37.2	0.1014	50.7
1.3200	16.1	0.8601	17.8	0.4300	37.3	0.0878	55.1
1.3001	16.0	0.8402	18.1	0.4134	38.0	0.0742	60.3
1.2801	16.0	0.8202	18.4	0.3967	38.0	0.0611	65.8
1.2601	15.9	0.8001	18.6	0.3809	37.6	0.0482	71.6
1.2401	15.9	0.7800	19.0	0.3644	37.2	0.0358	77.4
1.2203	15.8	0.7615	19.2	0.3505	36.9	0.0232	83.0
1.2000	15.8	0.7400	19.6	0.3338	36.6	0.0108	88.0
1.1800	15.8	0.7208	20.0	0.3192	36.3		
1.1600	15.8	0.7001	20.4	0.3028	36.2		

Differential capacity on mercury electrode in KBr - given are the capacity in μF cm^{-2}, potential in V. Reference electrode: 1 mol l^{-1} KCl calomel electrode at the given temperature.

Reference: S. Minc and J. Jastrzebska. Electrochim. Acta 10 (1965) 965

0.0005 mol l^{-1} KBr

0°C		20°C		40°C		60°C	
$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}
0.382	22.11	0.369	23.44	0.357	20.32	0.345	19.54
0.432	13.29	0.389	19.72	0.407	12.91	0.395	12.67
0.452	11.16	0.399	17.92	0.417	12.13	0.415	10.47
0.462	10.35	0.419	12.55	0.427	10.33	0.435	8.75
0.477	9.56	0.429	12.43	0.437	9.95	0.445	8.29
0.482	9.61	0.439	10.26	0.447	9.22	0.455	7.94
0.492	9.96	0.449	9.86	0.457	9.06	0.465	7.76
0.532	10.89	0.459	9.33	0.467	8.88	0.475	7.87
0.632	14.53	0.469	9.18	0.477	8.90	0.485	7.96
0.732	14.79	0.479	9.41	0.487	9.02	0.495	8.39
0.832	14.57	0.489	9.51	0.497	9.31	0.595	12.21
		0.499	9.78	0.507	9.57	0.695	14.50
		0.519	10.58	0.607	13.85	0.795	14.96
		0.619	14.06	0.707	15.49		
		0.719	15.60	0.807	15.87		
		0.819	15.53	0.907	16.21		
		0.919	15.43				

Differential capacity on mercury in KBr (continued)

0.0006 mol l^{-1} KBr

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.386	20.55	0.384	22.47	0.372	24.21	0.350	24.78
0.436	16.82	0.434	16.23	0.422	18.64	0.400	17.63
0.456	13.42	0.454	13.80	0.442	15.32	0.420	13.75
0.476	12.18	0.474	11.70	0.462	11.98	0.440	11.46
0.486	11.34	0.484	10.66	0.472	10.39	0.460	9.41
0.496	10.37	0.494	10.05	0.482	9.40	0.470	9.36
0.506	10.29	0.504	9.71	0.492	9.16	0.480	8.71
0.516	10.00	0.514	9.35	0.507	9.02	0.490	8.34
0.526	10.03	0.524	9.38	0.512	9.05	0.500	8.24
0.536	10.13	0.534	9.49	0.522	9.06	0.510	8.25
0.546	10.51	0.544	9.94	0.532	9.35	0.520	8.45
0.636	13.27	0.634	13.72	0.632	13.20	0.600	11.71
0.736	15.20	0.734	15.33	0.732	15.17	0.700	14.76
0.836	14.45	0.834	15.65	0.832	15.83	0.800	15.67

Differential capacity on mercury in KBr (continued)

0.0007 mol l^{-1} KBr

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.440	14.55	0.378	22.57	0.366	22.83	0.354	22.94
0.480	12.20	0.428	17.39	0.416	16.93	0.404	16.92
0.490	11.28	0.448	14.40	0.436	14.88	0.424	14.06
0.500	11.20	0.468	11.40	0.456	12.70	0.444	11.92
0.510	11.10	0.478	10.52	0.476	9.80	0.464	10.44
0.520	11.08	0.488	9.17	0.486	9.03	0.484	8.57
0.530	11.38	0.498	8.98	0.496	8.54	0.494	8.49
0.540	11.72	0.508	8.72	0.511	8.16	0.504	8.00
0.640	13.69	0.518	8.52	0.516	8.32	0.514	8.47
0.740	13.61	0.528	8.73	0.526	8.41	0.524	9.63
0.840	13.65	0.538	8.78	0.616	10.94	0.604	12.58
		0.628	12.16	0.716	13.50	0.704	14.32
		0.728	14.46			0.804	18.09
		0.828	14.46				

Differential capacity on mercury in KBr (continued)

0.0008 mol l^{-1} KBr

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.393	21.16	0.300	21.25	0.369	22.79	0.358	22.63
0.443	16.38	0.430	16.38	0.419	17.35	0.408	16.22
0.463	14.29	0.450	14.43	0.439	14.97	0.448	11.26
0.483	12.23	0.470	11.40	0.459	12.18	0.468	9.58
0.503	11.08	0.490	10.07	0.479	10.65	0.478	9.12
0.513	10.87	0.500	9.83	0.489	10.25	0.488	8.78
0.523	10.75	0.510	9.82	0.499	9.98	0.498	8.77
0.533	11.06	0.520	9.75	0.514	9.70	0.508	8.76
0.543	11.34	0.530	10.23	0.519	9.78	0.518	8.79
0.643	13.55	0.630	13.36	0.619	13.31	0.528	9.80
0.743	14.34	0.730	14.51	0.719	15.50	0.608	12.05
0.843	13.96	0.830	14.28	0.819	15.61	0.708	14.17
						0.808	14.75

Differential capacity on mercury in KBr (continued)

0.0009 mol l^{-1} KBr

$0^{\circ}C$		$20^{\circ}C$		$40^{\circ}C$		$60^{\circ}C$	
-E/V	$C/\mu F\ cm^{-2}$	-E/V	$C/\mu F\ cm^{-2}$	-E/V	$C/\mu F\ cm^{-2}$	-E/V	$C/\mu F\ cm^{-2}$
0.396	21.02	0.384	21.08	0.362	21.40	0.361	22.14
0.446	16.20	0.434	16.07	0.412	16.51	0.441	15.57
0.466	14.18	0.454	14.76	0.452	11.73	0.451	11.21
0.486	12.18	0.474	12.03	0.472	10.52	0.471	10.09
0.506	11.18	0.484	11.56	0.482	10.11	0.481	9.62
0.516	10.79	0.494	10.81	0.492	9.96	0.491	9.39
0.526	10.76	0.504	10.49	0.502	9.85	0.501	9.42
0.536	11.01	0.514	10.43	0.512	9.84	0.511	9.30
0.546	11.02	0.524	10.42	0.522	10.11	0.521	9.54
0.556	11.18	0.534	10.43	0.612	13.12	0.531	9.66
0.646	13.80	0.544	10.84	0.712	14.19	0.611	12.44
0.746	14.41	0.634	13.65	0.812	15.07	0.711	14.09
		0.734	14.86			0.811	14.96

Differential capacity on mercury in KBr (continued)

0.001 mol l^{-1} KBr

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.398	20.80	0.387	21.71	0.375	21.37	0.364	21.87
0.448	17.15	0.437	17.52	0.425	17.63	0.414	17.40
0.498	12.17	0.477	13.28	0.475	12.31	0.464	11.24
0.508	11.41	0.497	11.42	0.485	11.30	0.474	10.67
0.518	10.98	0.507	11.03	0.495	10.53	0.484	10.07
0.528	10.84	0.517	10.62	0.505	10.34	0.494	9.68
0.533	10.83	0.527	10.56	0.520	10.15	0.504	9.57
0.548	10.88	0.537	10.58	0.525	10.18	0.514	9.53
0.648	12.90	0.547	10.99	0.535	10.30	0.524	9.54
0.748	13.40	0.637	13.34	0.545	10.61	0.534	10.07
0.848	14.02	0.737	13.80	0.625	12.98	0.564	10.49
		0.837	13.72	0.725	14.84	0.614	12.40
						0.714	14.74

Differential capacity on mercury electrode in KF - given are the capacity in $\mu F\ cm^{-2}$, potential in V. Reference electrode: 1 mol l^{-1} KCl calomel electrode at the given temperature.

Reference: S. Minc and J. Jastrzebska. Electrochim. Acta 10 (1965) 965.

0.0005 mol l^{-1} KF

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.232	10.70	0.324	9.64	0.202	10.31	0.210	11.22
0.332	7.74	0.374	6.19	0.302	8.14	0.310	8.53
0.432	4.63	0.414	5.03	0.402	5.48	0.410	5.24
0.452	4.17	0.444	4.40	0.422	4.65	0.430	4.82
0.472	4.10	0.469	4.06	0.442	4.52	0.450	4.09
0.492	4.15	0.484	4.18	0.452	4.35	0.460	3.97
0.512	4.29	0.504	4.25	0.462	4.03	0.470	4.66
0.532	5.43	0.524	4.93	0.472	4.38	0.480	4.91
0.632	7.18	0.574	6.53	0.482	4.40	0.510	5.69
0.732	7.37	0.624	7.44	0.502	4.81	0.610	8.61
0.832	7.40	0.724	7.71	0.602	8.17	0.710	9.31
		0.824	7.64	0.702	8.66	0.810	9.29
				0.802	8.67		

Differential capacity on mercury in KF (continued)

0.005 mol l^{-1} KF

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.285	23.94	0.276	21.05	0.268	20.05	0.291	18.78
0.385	15.87	0.376	13.53	0.368	12.92	0.391	12.05
0.435	11.83	0.426	11.01	0.418	10.80	0.441	10.35
0.445	11.45	0.436	10.82	0.428	10.57	0.451	10.16
0.455	11.22	0.446	10.56	0.438	10.40	0.461	10.12
0.465	11.20	0.456	10.48	0.448	10.32	0.471	10.16
0.475	10.86	0.471	10.43	0.458	10.31	0.481	10.20
0.485	10.90	0.476	10.54	0.468	10.30	0.491	10.38
0.495	11.20	0.496	10.88	0.488	10.78	0.591	13.61
0.535	11.82	0.526	11.87	0.568	13.64	0.691	16.24
0.585	15.15	0.476	14.00	0.668	16.44	0.791	16.74
0.685	16.78	0.676	16.83	0.768	17.18		
0.785	16.90	0.776	17.08				
0.885	16.29						

Differential capacity on mercury in KF (continued)

0.006 mol l^{-1} KF

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.284	10.60	0.272	10.62	0.262	11.07	0.259	11.30
0.384	8.19	0.372	8.12	0.362	8.72	0.359	8.57
0.474	6.27	0.422	6.76	0.412	7.28	0.409	7.40
0.504	6.53	0.432	6.54	0.432	6.71	0.429	7.02
0.514	6.64	0.442	6.44	0.442	6.59	0.449	6.58
0.534	7.14	0.452	6.36	0.452	6.65	0.459	6.19
0.584	8.38	0.462	6.30	0.462	6.55	0.469	6.16
0.684	9.78	0.472	6.28	0.472	6.20	0.479	6.31
0.784	9.65	0.482	6.34	0.482	6.54	0.489	6.76
0.884	10.37	0.492	6.44	0.512	6.86		
		0.572	8.11	0.562	8.08		
		0.672	9.65	0.662	9.96		
		0.772	9.92	0.762	10.67		
		0.872	9.67	0.862	10.59		

Differential capacity on mercury in KF (continued)

0.0075 mol l^{-1} KF

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.294	21.67	0.287	22.17	0.278	22.80	0.270	22.79
0.394	15.00	0.387	15.50	0.378	15.56	0.370	14.60
0.444	12.36	0.437	13.20	0.428	12.94	0.420	12.41
0.454	12.20	0.447	12.40	0.438	12.70	0.430	12.10
0.464	12.15	0.457	12.25	0.448	12.27	0.440	11.67
0.474	12.07	0.467	12.10	0.458	12.20	0.450	11.52
0.484	12.11	0.477	11.90	0.468	11.89	0.460	11.45
0.494	12.35	0.487	12.00	0.478	11.84	0.470	11.44
0.504	12.39	0.497	12.14	0.488	11.98	0.480	11.41
0.544	14.00	0.507	12.43	0.528	12.62	0.490	11.48
0.594	16.30	0.537	12.80			0.520	11.98
0.694	18.36	0.587	15.37			0.570	13.60
		0.687	17.76			0.670	16.69
						0,770	17.66

Differential capacity on mercury in KF (continued)

0.01 mol l^{-1} KF

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.300	20.73	0.293	21.15	0.285	19.69	0.278	20.10
0.400	15.00	0.393	14.95	0.385	14.82	0.378	14.90
0.450	13.33	0.443	13.03	0.435	13.18	0.428	13.11
0.460	13.10	0.453	12.87	0.445	13.00	0.438	12.77
0.475	13.00	0.463	12.75	0.455	12.80	0.448	12.50
0.480	13.05	0.478	12.70	0.465	12.63	0.458	12.36
0.490	13.15	0.483	12.72	0.480	12.59	0.468	12.28
0.500	13.30	0.493	12.74	0.485	12.69	0.483	12.23
0.510	13.60	0.543	13.83	0.495	12.76	0.488	12.24
0.550	14.80	0.593	15.77	0.585	15.20	0.498	12.32
0.600	16.90	0.693	17.79	0.685	17.67	0.578	14.37
0.700	18.55	0.793	17.82	0.785	18.22	0.678	16.80
0.800	19.33					0.778	17.13

Differential capacity on mercury in KF (continued)

0.017 mol l^{-1} KF

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.313	19.92	0.307	20.08	0.300	18.26	0.293	18.67
0.413	15.70	0.407	15.70	0.400	14.34	0.393	14.84
0.463	14.70	0.457	14.40	0.450	13.34	0.443	13.22
0.473	14.63	0.467	14.25	0.460	13.23	0.453	13.10
0.483	14.70	0.477	14.20	0.470	13.19	0.463	13.05
0.493	14.78	0.487	14.30	0.480	13.13	0.473	12.92
0.503	14.86	0.497	14.35	0.490	13.28	0.483	12.82
0.513	15.10	0.507	14.43	0.500	13.28	0.493	12.95
0.613	17.90	0.517	14.69	0.600	15.07	0.543	13.54
0.713	18.80	0.557	15.70	0.700	16.21	0.593	14.49
0.813	17.63	0.607	16.95	0.800	16.21	0.693	15.95
		0.707	18.10			0.793	16.23
		0.807	17.67				

Differential capacity on mercury in KF (continued)

$0.02\ \mathrm{mol\ l^{-1}}$ KF

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.313	21.41	0.307	31.30	0.303	19.00	0.297	19.13
0.413	16.47	0.357	19.32	0.403	15.88	0.397	16.12
0.463	15.55	0.407	16.82	0.453	15.30	0.447	15.30
0.473	15.48	0.457	15.55	0.463	15.29	0.457	15.13
0.483	15.51	0.467	15.39	0.473	15.22	0.467	15.13
0.493	15.59	0.477	15.38	0.483	15.20	0.477	15.00
0.503	15.70	0.487	15.49	0.493	15.24	0.487	14.96
0.513	16.00	0.497	15.50	0.503	15.35	0.497	14.98
0.613	18.31	0.507	15.57	0.603	16.53	0.507	14.98
0.713	18.93	0.557	16.58	0.703	17.25	0.597	15.76
0.813	17.74	0.607	17.71	0.803	16.88	0.697	16.75
0.913	16.53	0.707	18.85			0.797	16.85
		0.807	18.12				

Differential capacity on mercury in KF (continued)

0.025 mol l^{-1} KF

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.323	20.61	0.315	20.76	0.308	19.20	0.303	19.32
0.423	16.95	0.415	17.10	0.408	16.35	0.403	16.16
0.473	16.00	0.465	16.00	0.458	15.33	0.453	15.15
0.483	16.04	0.480	15.84	0.478	15.32	0.463	15.10
0.493	16.07	0.485	15.85	0.483	15.28	0.473	15.10
0.503	16.15	0.495	15.88	0.488	15.30	0.488	15.00
0.513	16.21	0.505	16.00	0.498	15.34	0.493	15.03
0.523	16.30	0.515	16.13	0.508	15.36	0.503	15.10
0.623	18.25	0.615	17.93	0.518	15.45	0.603	15.20
0.723	18.32	0.715	18.45	0.608	16.47	0.703	16.97
0.823	17.31	0.815	17.74	0.708	17.00	0.803	17.02
				0.808	17.00		
				0.908	16.46		

Differential capacity on mercury in KF (continued)

$0.03\ \mathrm{mol\ l^{-1}\ KF}$

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.325	20.07	0.420	17.67	0.313	20.28	0.408	16.78
0.425	17.24	0.450	17.00	0.413	17.24	0.438	16.13
0.445	16.87	0.470	16.65	0.463	16.41	0.458	15.79
0.465	16.72	0.480	16.40	0.473	16.26	0.468	15.75
0.475	16.42	0.490	16.45	0.483	16.21	0.478	15.70
0.485	16.50	0.500	16.46	0.490	16.24	0.488	15.69
0.495	16.55	0.510	16.57	0.503	16.35	0.498	15.71
0.505	16.59	0.520	16.77	0.513	16.36	0.508	15.83
0.525	16.79	0.570	17.48	0.523	16.41	0.518	15.75
0.625	18.19	0.620	18.30	0.563	16.90	0.528	15.88
0.725	18.17	0.720	18.63	0.613	17.65	0.558	16.23
0.825	16.98			0.713	18.27	0.608	16.89
				0.813	17.77	0.708	17.62
						0.808	17.71

Differential capacity on mercury in KF (continued)

0.035 mol l^{-1} KF

0 °C		20 °C		40 °C		60 °C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.318	20.48	0.313	20.37	0.297	19.00	0.290	19.15
0.418	17.07	0.413	17.14	0.397	15.90	0.390	15.58
0.448	16.61	0.443	16.54	0.427	15.25	0.410	15.11
0.458	16.57	0.453	16.43	0.447	15.06	0.420	14.90
0.468	16.50	0.463	16.32	0.467	15.02	0.430	14.67
0.478	16.46	0.473	16.24	0.477	15.01	0.440	14.61
0.488	16.47	0.483	16.20	0.487	15.00	0.450	14.41
0.498	16.53	0.493	16.24	0.497	15.06	0.460	14.37
0.508	16.57	0.503	16.32	0.547	15.45	0.470	14.37
0.538	17.04	0.513	16.40	0.597	15.90	0.480	14.36
0.618	17.82	0.613	17.79	0.697	16.02	0.490	14.35
0.718	17.48	0.713	17.70	0.797	15.56	0.500	14.40
0.818	16.41	0.813	16.72			0.540	14.72
						0.590	15.17
						0.690	15.73
						0.790	15.66

Differential capacity on mercury in $HClO_4$ at 25°C - given are the capacity in µF cm^{-2}, potential in mV vs $Pt(H_2)H^+$ in the same solution.

Reference: B.G. Dekker, M. Sluyters-Rehbach and J.H. Sluyters.
J. Electroanal. Chem. 21 (1969) 137.

	C/µF cm^{-2}						
c/mol l^{-1}	0.1	0.5	1.0	3.0	5.0	7.0	9.0
E/mV							
+350	23.0	24.7	25.2	25.6			
+325	22.7	24.35	24.7	25.1			
+300	22.5	24.1	24.35	24.6	25.0		
+275	22.4	23.8	24.0	24.2	24.3		
+250	22.4	23.6	23.7	23.8	23.7	39.6	
+225	22.45	23.4	23.45	23.4	23.2	33.0	
+200	22.6	23.3	23.2	23.1	22.8	26.2	31.4
+175	22.9	23.25	23.1	22.8	22.3	24.6	28.8
+150	23.3	23.3	23.1	22.5	21.9	23.5	26.9
+125	23.65	23.4	23.0	22.3	21.5	22.75	25.45
+100	24.1	23.6	23.0	22.1	21.15	22.0	24.3
+ 75	24.7	23.85	23.0	21.9	20.9	21.5	23.55
+ 50	25.4	24.1	23.1	21.8	20.7	20.9	22.9
+ 25	26.0	24.5	23.3	21.7	20.5	20.5	22.3
0	26.6	25.0	23.5	21.6	20.3	20.0	21.7
- 25	27.1	25.5	23.8	21.6	20.1	19.65	21.2
- 50	27.5	26.0	24.2	21.7	20.0	19.3	20.7
- 75	27.7	26.65	24.6	21.8	19.9	19.1	20.35
-100	27.7	27.3	25.2	21.9	19.9	18.9	20.0
-125	27.4	27.9	25.75	22.2	19.9	18.75	19.65
-150	26.8	28.4	26.4	22.4	20.0	18.6	19.3
-175	25.95	28.8	27.0	22.75	20.1	18.5	18.95
-200	25.0	29.1	27.6	23.1	20.3	18.4	18.6
-225	23.9	29.25	28.05	23.55	20.5	18.35	18.3
-250	22.9	29.2	28.4	24.05	20.7	18.3	18.1
-275	21.8	28.95	28.65	24.6	21.05	18.4	17.9
-300	20.9	28.3	28.8	25.1	21.4	18.5	17.7

Differential capacity on mercury in $HClO_4$ (continued)

	$C/\mu F\ cm^{-2}$						
$c/mol\ l^{-1}$	0.1	0.5	1.0	3.0	5.0	7.0	9.0
E/mV							
-325	20.1	27.4	28.75	25.6	21.8	18.6	17.55
-350	19.3	26.5	28.5	26.1	22.2	18.8	17.4
-375	18.7	25.35	28.1	26.6	22.65	19.0	17.35
-400	18.2	24.2	27.3	27.1	23.1	19.2	17.3
-425	17.6	23.0	26.5	27.5	23.65	19.6	17.3
-450	17.1	21.9	25.65	27.8	24.2	20.0	17.3
-475	16.8	20.95	24.5	27.95	24.8	20.45	17.35
-500	16.5	20.0	23.4	28.0	25.3	21.0	17.4
-525	16.2	19.25	22.3	27.9	25.8	21.5	17.6
-550	15.9	18.5	21.25	27.7	26.3	22.05	17.8
-575	15.7	17.85	20.3	27.3	26.7	22.65	18.1
-600	15.5	17.3	19.45	26.8	27.1	23.3	18.5
-625	15.3	16.9	18.7	26.3	27.3	24.0	18.9
-650	15.1	16.5	18.1	25.6	27.4	24.6	19.3
-675	15.0	16.2	17.6	24.9	27.5	2.52	19.9
-700	14.9	16.0	17.1	24.0	27.5	25.8	20.4
-725	14.8	15.8	16.7	23.2	27.4	26.3	21.1
-750	14.8	15.7	16.4	22.4	27.2	26.7	21.65
-775	14.8	15.6	16.2	21.6	26.9	27.3	22.4
-800	14.7	15.5	16.0	21.0	26.5	27.6	23.2
-825	14.7	15.4	15.8	20.3	26.0	27.8	23.8
-850	14.7	15.4	15.8	19.9	25.7	28.1	24.45

Differential capacity on mercury in 5.2 M $HClO_4$ at different temperatures. Given are the capacity in $\mu F\ cm^{-2}$, potential in mV vs $Pt(H_2)H^+$.

Reference: B.G. Dekker, M. Sluyters-Rehbach and J.H. Sluyters.
J. Electroanal. Chem. 21 (1969) 137.

	C/μF cm^{-2}		
t/°C	-35	-5	25
E/mV			
+500	29.8	30.8	32.3
+400	26.2	27.0	28.7
+300	23.1	24.0	25.2
+200	20.9	21.9	22.9
+100	19.35	20.4	21.4
0	18.5	19.6	20.5
-100	18.5	19.5	20.2
-200	19.4	20.0	20.3
-250	20.3	20.6	20.7
-300	21.6	21.6	21.4
-350	23.2	22.8	22.3
-400	25.1	24.1	23.3
-500	29.6	27.4	25.6
-600	33.5	30.1	27.5
-700	33.9	30.8	28.2
-800	30.6	29.0	27.2
-850		27.4	26.3
-875			25.7
-900	25.4	25.7	25.2
-925			
-940			
-950		24.1	24.0

	C/μF cm^{-2}		
t/°C	-35	-5	25
E/mv			
-960			
-975			
-980			
-1000	21.7	22.7	23.0
-1020			
-1025	21.1		
-1040			
-1050	20.6	21.6	
-1075	20.1		
-1090			
-1100	19.9	20.9	
-1110			
-1125	19.3		
-1130			
-1150			
-1160			
-1175			
-1190			
-1200			
-1210			
-1225			
-1250			

Differential capacity on mercury in 7.5 M $HClO_4$ at different temperatures. Given are the capacity in $\mu F\ cm^{-2}$, potential in mV vs $Pt(H_2)H^+$.

Reference: B.G. Dekker, M. Sluyters-Rehbach and J.H. Sluyters.
J. Electroanal. Chem. 21 (1969) 137.

	$C/\mu F\ cm^{-2}$	
$t/^{\circ}C$	25 - 26	-35 ± 0.5
E/mV		
+200	25.7	24.0
+100	22.3	20.8
+ 50	21.2	19.8
0	20.4	18.9
- 50	19.7	18.1
-100	19.1	17.6
-200	18.4	16.8
-300	18.3	16.7
-350	18.5	
-400	18.9	17.6
-450	19.6	
-500	20.4	20.1
-550	21.5	22.0
-600	22.7	24.5
-650	24.1	27.2
-700	25.3	29.7
-750	26.5	31.9

	$C/\mu F\ cm^{-2}$	
$t/^{\circ}C$	25 - 26	-35 ± 0.5
E/mV		
-800	27.6	33.3
-850	28.2	34.0
-900	28.7	33.9
-915	28.8	
-925	28.8	
-935	28.7	
-950	28.8	33.2
-965	28.8	
-975	28.0	
-985	28.8	
-1000		31.8
-1025		31.2
-1050		30.4
-1075		29.5
-1100		28.7
-1125		27.8
-1150		27.0

Capacity data for 0.1 M $HClO_4$ solution on mercury electrode - given are the capacity in μF cm^{-2}, potential in V.

T = 25°C

Reference electrode: SCE (NaCl solution)

Reference: K.G. Baikerikar, D.E. Broadhead and R.S. Hansen. J. Phys. Chem. 80 (1976) 370.

E/V	C/μF cm^{-2}
-1.100	15.27
-1.075	15.30
-1.050	15.34
-1.025	15.42
-1.000	15.52
-0.975	15.64
-0.950	15.77
-0.925	15.93
-0.900	16.14
-0.875	16.36
-0.850	16.62
-0.825	16.91
-0.800	17.26
-0.775	17.63
-0.750	18.06
-0.725	18.53
-0.700	19.07
-0.675	19.70
-0.650	20.39
-0.625	21.19
-0.600	22.13
-0.575	23.13
-0.550	24.26
-0.525	25.43
-0.500	26.57
-0.475	27.56
-0.450	28.35
-0.425	28.85
-0.400	29.04
-0.375	28.94
-0.350	28.59
-0.325	28.08
-0.300	27.44
-0.275	26.74
-0.250	26.05
-0.225	25.36
-0.200	24.72
-0.175	24.22
-0.150	23.76
-0.125	23.41
-0.100	23.14
-0.075	22.99
-0.050	22.91
-0.025	22.93
0.000	23.03
0.025	23.21
0.050	23.45
0.075	23.76
0.100	24.10
0.125	24.49
0.150	24.95
0.175	25.43
0.200	25.94

Differential capacity on mercury electrode in $HClO_4$ at 125 Hz; given are the capacity in $\mu F\ cm^{-2}$, potential in mV. T = 25°C
Reference electrode: 1 mol l^{-1} KCl calomel electrode.

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 875.

	$C/\mu F\ cm^{-2}$				
$c/mol\ l^{-1}$	1×10^{-5}	2×10^{-5}	5×10^{-5}	1×10^{-4}	1×10^{-3}
E/mV					
250	4.2				
220		6.3	11.2		
200	4.0			16.7	
170		5.3			
150	3.8			12.3	
120		3.4	6.6		
100	2.9			8.0	18.5
70		2.8			
60			3.5		
50	1.7				
30		1.7		3.8	11.7
0	1.4	1.6	2.3	2.6	7.4
-30		1.7	2.7		
-50	1.5			3.5	8.0
-60		2.4			
-70			4.6		
-100	2.4	3.6		5.7	10.9
-130			6.4		
-150				8.4	13.8
-170		5.0	7.6		
-200	3.4			11.5	15.7
-230		6.3			
-280			12.1		
-300	4.4			15.1	17.1
-310		8.2			
-380			13.1		
-400	4.4	8.1		17.1	16.9
-480			13.3		
-500	4.4	8.4		16.2	16.1
-580			12.6		
-600	4.4	8.4		15.7	15.6
-680			11.8		
-700	4.3	7.3		15.4	15.3
-800			12.5	15.3	15.3
-900	3.9			15.5	16.0

Differential capacity on mercury electrode in $NaClO_4$ at 125 Hz, 25°C.
Given are the capacity in $\mu F\ cm^{-2}$, potential in mV.
Reference electrode: 1 mol l^{-1} KCl calomel electrode.

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 875.

	$C/\mu F\ cm^{-2}$				
$c/mol\ l^{-1}$	5×10^{-5}	1×10^{-4}	2×10^{-4}	5×10^{-4}	1×10^{-3}
E/mV					
200	8.5	12.3			
180			15.6		
170				18.2	
150	7.5	10.5			20.0
130			12.0		
120				13.9	
100	4.0	6.4			15.3
80			6.9		
70				8.0	
50	2.3	3.2			9.0
40			3.9		
20				4.9	
0	1.6	2.5	3.3	4.8	6.5
-20			3.5		
-30				5.3	
-50	2.3	3.5			8.0
-70			5.3		
-100	3.9	5.5			10.9
-120			8.8		
-130				10.6	
-150	5.2				
-200	6.5	8.2			
-220			10.9	13.7	16.0
-230					
-250	7.5				
-300	8.5	10.0			18.1

Capacity in $NaClO_4$ (continued)

	$C/\mu F\ cm^{-2}$				
$c/mol\ l^{-1}$	5×10^{-5}	1×10^{-4}	2×10^{-4}	5×10^{-4}	1×10^{-3}
E/mV					
-320			13.2		
-330				15.9	
-400	8.0	11.7			17.9
-420			14.5		
-430				15.9	
-500	7.5	11.4			16.8
-520			13.7		
-530				15.9	
-600	7.5	10.5			16.2
-620			13.7		
-630				15.5	
-700	7.5	10.5			15.9
-720			13.7		
-730				14.5	
-800	7.5	10.5			15.9
-820			13.7		
-830				15.5	
-900					16.2
-930				15.4	
-1000					16.5

Differential capacity on mercury electrode in $KClO_4$; given are the capacity in $\mu F\ cm^{-2}$, potential in V.
The calomel electrode at a given temperature and at the same KCl concentration as the system studied was used as a reference.
All potential values /E/ presented in tables were recalculated vs the normal calomel electrode at 20°C.

Reference: S. Minc, M. Jurkiewicz and J. Jastrzebska. J. Electroanal. Chem. 10 (1965) 437.

0.0005 mol l^{-1} $KClO_4$

0°C		20°C		40°C		60°C	
-E/V	C/$\mu F\ cm^{-2}$	-E/V	C/$\mu F\ cm^{-2}$	-E/V	C/$\mu F\ cm^{-2}$	-E/V	C/$\mu F\ cm^{-2}$
0.232	18.71	0.219	20.23	0.207	19.78	0.196	21.41
0.282	16.44	0.269	19.32	0.257	18.89	0.264	20.39
0.332	11.99	0.319	17.24	0.307	16.80	0.296	17.23
0.362	9.00	0.369	11.99	0.357	11.57	0.346	12.64
0.382	7.40	0.389	9.76	0.407	6.95	0.396	7.74
0.402	6.02	0.399	9.09	0.427	5.82	0.416	6.49
0.412	5.54	0.419	6.36	0.437	5.44	0.426	6.08
0.432	5.29	0.429	5.97	0.447	5.18	0.436	5.65
0.442	5.23	0.439	5.58	0.457	4.96	0.446	5.35
0.452	5.41	0.449	5.37	0.467	4.91	0.456	5.51
0.462	5.54	0.459	5.11	0.477	4.92	0.466	5.11
0.472	5.72	0.469	5.21	0.487	5.06	0.476	5.11
0.482	6.27	0.479	5.39	0.497	5.35	0.486	5.11
0.532	9.04	0.499	5.97	0.507	6.14	0.496	5.26
0.632	11.53	0.519	8.20	0.607	11.34	0.596	10.56
0.732	12.16	0.619	12.46				
0.832	17.77	0.719	12.63				
		0.819	12.65				
		0.919	12.65				
		1.019	12.77				

Differential capacity on mercury in $KClO_4$ (continued)

0.001 mol l^{-1} $KClO_4$

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.248	20.54	0.237	22.96	0.225	21.91	0.215	23.11
0.348	20.12	0.337	19.54	0.325	19.70	0.315	19.15
0.448	15.52	0.387	15.07	0.375	14.73	0.415	9.72
0.468	12.49	0.437	8.97	0.395	13.30	0.435	8.33
0.478	11.67	0.457	7.68	0.425	9.68	0.455	7.24
0.488	11.22	0.477	6.89	0.445	8.44	0.465	6.86
0.498	7.07	0.487	6.49	0.465	6.89	0.475	6.61
0.508	7.05	0.497	6.59	0.475	6.65	0.485	6.44
0.518	7.07	0.507	6.68	0.485	6.62	0.495	6.31
0.528	7.07	0.517	7.21	0.495	6.52	0.505	6.41
0.538	7.24	0.537	8.06	0.505	6.49	0.515	6.40
0.548	7.49	0.637	12.83	0.515	6.59	0.525	6.70
0.648	12.86	0.737	14.33	0.525	6.73	0.615	10.17
				0.535	7.34	0.715	12.41
				0.625	11.72		

Differential capacity on mercury in $KClO_4$ (continued)

0.005 mol l^{-1} $KClO_4$

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.285	25.40	0.276	24.45	0.267	24.39	0.259	25.94
0.385	25.13	0.376	22.41	0.367	21.33	0.359	21.87
0.485	16.00	0.476	13.80	0.467	13.17	0.459	13.38
0.505	14.55	0.496	12.94	0.487	12.32	0.479	12.38
0.515	14.30	0.516	12.42	0.497	12.00	0.499	11.81
0.525	14.09	0.526	12.42	0.507	11.80	0.509	11.75
0.535	14.16	0.536	12.57	0.517	11.76	0.519	11.94
0.545	14.49	0.546	12.72	0.527	11.89	0.529	12.16
0.585	15.90	0.576	13.80	0.527	12.02	0.559	12.82
0.685	18.87	0.676	16.92	0.567	13.03	0.659	15.91
0.785	19.18	0.776	17.68	0.667	16.40		
0.885	18.24	0.876	16.71	0.767	17.25		
				0.867	17.44		

Differential capacity on mercury in $KClO_4$ (continued)

0.01 mol l^{-1} $KClO_4$

0°C		20°C		40°C		60°C	
$-E$/V	$C/\mu F\ cm^{-2}$	$-E$/V	$C/\mu F\ cm^{-2}$	$-E$/V	$C/\mu F\ cm^{-2}$	$-E$/V	$C/\mu F\ cm^{-2}$
0.300	25.83	0.293	24.78	0.285	25.22	0.278	28.31
0.400	25.13	0.393	23.04	0.385	22.87	0.378	22.26
0.500	18.13	0.493	16.35	0.485	16.04	0.478	15.48
0.520	17.39	0.513	15.65	0.505	15.04	0.508	14.13
0.530	17.17	0.523	15.43	0.515	14.87	0.518	14.05
0.540	17.09	0.533	15.43	0.525	14.74	0.528	13.94
0.550	17.09	0.543	15.48	0.535	14.70	0.538	13.80
0.560	17.22	0.553	15.56	0.545	14.74	0.548	13.77
0.600	18.04	0.593	16.43	0.555	14.87	0.558	13.86
0.700	19.39	0.693	17.91	0.585	15.52	0.578	15.10
0.800	18.65	0.793	17.70	0.685	17.22	0.678	17.24
0.900	17.56	0.893	16.74	0.785	17.56	0.778	17.87
				0.885	16.91	0.878	17.63

Differential capacity on mercury in $KClO_4$ (continued)

0.015 mol l^{-1} $KClO_4$

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.309	24.57	0.303	24.48	0.305	22.09	0.289	24.74
0.409	24.70	0.403	22.74	0.405	21.83	0.389	20.52
0.509	19.52	0.503	17.74	0.505	16.83	0.489	15.91
0.529	18.52	0.523	17.39	0.525	16.56	0.509	15.56
0.539	18.17	0.543	17.35	0.535	16.48	0.529	15.52
0.549	18.04	0.553	17.39	0.545	16.52	0.539	15.53
0.559	17.91	0.563	17.43	0.555	16.52	0.549	15.61
0.569	17.91	0.573	17.61	0.565	16.56	0.559	15.70
0.579	17.95	0.603	17.96	0.575	16.69	0.589	16.04
0.609	18.17	0.703	18.48	0.595	16.77	0.689	16.39
0.709	18.70	0.803	17.74	0.695	17.78		
0.809	17.65						

Differential capacity on mercury in $KClO_4$ (continued)

$0.02\ mol\ l^{-1}\ KClO_4$

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.316	25.21	0.309	23.88	0.303	24.30	0.297	24.19
0.416	25.63	0.409	23.53	0.403	22.87	0.397	20.93
0.516	20.97	0.509	18.69	0.503	17.93	0.497	16.10
0.536	20.29	0.529	17.93	0.523	17.34	0.517	15.72
0.546	19.91	0.549	17.55	0.543	16.96	0.537	15.47
0.556	19.75	0.559	17.47	0.553	16.88	0.547	15.42
0.566	19.58	0.569	17.38	0.563	16.79	0.557	15.38
0.576	19.41	0.579	17.34	0.573	16.71	0.567	15.42
0.586	19.41	0.589	17.38	0.583	16.75	0.577	15.47
0.596	19.37	0.599	17.38	0.593	16.79	0.587	15.51
0.606	19.33	0.609	17.40	0.603	16.71	0.597	15.55
0.616	19.37	0.709	17.64			0.697	16.14
0.636	19.66						
0.716	19.25						

Differential capacity on mercury electrode in various aqueous solutions at 125 Hz, 25°C. given are the capacity in µF cm^{-2}, potential in mV.
Reference electrode : 1 mol l^{-1} calomel electrode.
Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 875.

C/µF cm^{-2}

	NH_4OH				NH_4ClO_4	
c/mol l^{-1}	1 x 10^{-4}	2 x 10^{-4}	1 x 10^{-3}	1 x 10^{-2}	1 x 10^{-4}	1 x 10^{-3}
E/mV						
360		16.9				
330			19.5			
320	14.8					
280					18.2	
260		14.9				
240				19.5		
230			15.4			
220	10.7					
190				16.5		
180			13.3		16.8	
160		11.2				19.4
140				12.5		
130			9.6		12.9	
120	6.6					
110		7.1				15.5
90				7.9		
80			5.4		6.3	
70	3.1					
60		3.4				9.5
40			3.3	5.5		
30					3.0	
20	1.5					
0	1.4	1.8	2.5	3.8	2.4	6.7
-30	1.7					
-40		2.3				8.7
-60			3.8	5.8	4.2	
-80	1.9					
-90		4.0				10.5
-110			5.7	8.5	6.5	
-130	4.6					
-140		6.3				13.1
-160			8.5	11.4	9.9	

Capacity in NH_4OH and NH_4ClO_4 (continued)

	$C/\mu F\ cm^{-2}$					
	NH_4OH				NH_4ClO_4	
$c/mol\ l^{-1}$	1×10^{-4}	2×10^{-4}	1×10^{-3}	1×10^{-2}	1×10^{-4}	1×10^{-3}
E/mV						
-180	6.8					
-190		8.9				15.2
-210				14.0	13.0	
-240		11.6				
-260			14.2			
-280	10.1					
-290						17.1
-310				17.9	14.3	
-340		14.0				
-360			17.9			
-380	12.9					
-390						17.5
-410				18.7	15.6	
-440		15.0				
-460			18.4			
-480	13.0					
-490						17.2
-510				18.5	15.5	
-540		15.4				
-560			17.5			
-580	12.1					
-590						16.9
-610				17.4	15.6	
-640		15.0				
-660			17.7			
-680	12.1					
-690						16.8
-710				17.0	14.3	
-740		13.6				
-760			16.9			
-780	12.1					
-790						17.0
-810				17.0	14.3	
-890						17.6

Differential capacity on mercury electrode in various aqueous solutions at 125 Hz, 25°C. Given are the capacity in μF cm^{-2}, potential in mV vs NCE.

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 875.

$C/\mu F\ cm^{-2}$

	NaOH				CH_3COOH		
$c/mol\ l^{-1}$	$1x10^{-4}$	$2x10^{-4}$	$5x10^{-4}$	$1x10^{-3}$	$4.5x10^{-5}$	$6.5x10^{-4}$	$1x10^{-2}$
E/mV							
320					13.6		
270	16.0						
250		18.4					
220					10.9		
200			13.1				
190							21.4
170	12.5					15.8	
150		12.2		17.0			
140							15.4
130						11.2	
120	8.4				4.9		
100		7.7	10.1				
90							10.0
80						5.1	
70	4.5				3.5		
50		4.1	6.2	12.5			
40							5.8
30						3.0	
20					1.5		
0	2.6	3.4	4.8	7.9	1.3	2.3	4.4
-10							4.5
-20						2.4	
-30	3.0				1.8		
-50		4.6	6.0	6.2			
-60							5.9
-70						2.7	
-80	4.7				2.9		
-100		7.3	8.6	7.4			

Capacity in NaOH and CH_3COOH (continued)

$C/\mu F\ cm^{-2}$

	NaOH				CH_3COOH		
$c/mol\ l^{-1}$	$1x10^{-4}$	$2x10^{-4}$	$5x10^{-4}$	$1x10^{-3}$	$4.5x10^{-5}$	$6.5x10^{-4}$	$1x10^{-2}$
E/mV							
-110							8.8
-120						6.3	
-130	7.2						
-150		9.8	11.7	10.4			
-160							12.3
-180					5.9		
-200			14.0	13.4			
-210							14.8
-220						11.7	
-230	10.5						
-250		14.0	16.0		9.5		
-280							
-300			17.1	16.0			
-310							17.6
-320						15.8	
-330	13.3						
-350		16.1					
-380					12.3		
-400			17.2	18.0			
-410							17.7
-420						16.9	
-430	13.7						
-450		15.8					
-480					13.6		
-500			16.5	17.6			
-510							17.3
-520						17.1	
-530	13.8						
-550		15.1					
-580					13.7		

Capacity in NaOH and CH_3COOH (continued)

$C/\mu F\,cm^{-2}$

	NaOH				CH_3COOH		
$c/mol\,l^{-1}$	$1x10^{-4}$	$2x10^{-4}$	$5x10^{-4}$	$1x10^{-3}$	$4.5x10^{-5}$	$6.5x10^{-4}$	$1x10^{-2}$
E/mV							
-600			15.8	16.7			
-610							16.1
-630	13.2						
-640						16.5	
-650		14.6					
-680							
-700			15.4	16.0	13.6		
-710							15.7
-730	12.6						
-740						16.3	
-750		14.4					
-780					13.7		
-800			15.3	15.6			
-810							15.7
-840						16.2	
-850	12.3	14.4					
-880					14.7		
-900			15.4	15.6			
-950		14.5					
-1000				16.0			

Differential capacity $C/\mu F\ cm^{-2}$ on mercury electrode for various solutions, measured at 125 Hz, 25°C.

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 1157.

$C/\mu F\ cm^{-2}$

Depolarizer	$AgNO_3$			$HgClO_4$	$CuSO_4$	$Cd(NO_3)_2$	$HClO_4$**
Concn./ mmol l-1	3.0	1.0	0.5	1.0	0.5	0.5	1.0
pH of soln.	4.9	4.9	4.8	4.3	5.0	5.0	3.0
E*/V							
+0.1	-	-	-	-	-	-	18.5
0.05	-	-	-	-	-	-	11.9
0.00	-	3.5	2.0	5.5	2.7	-	7.4
-0.05	-	4.0	3.0	7.0	4.9	19.0	8.0
-0.1	10.7	6.2	4.8	9.8	7.0	13.0	10.9
-0.15	13.0	7.7	7.0	12.9	9.1	12.9	13.8
-0.2	14.8	10.8	9.3	14.9	11.3	14.0	15.7
-0.25	15.7	12.8	11.7	16.4	12.9	14.8	16.5
-0.3	16.2	14.1	14.1	17.3	14.2	15.0	17.1
-0.35	16.4	14.8	15.5	17.4	14.8	15.2	17.1
-0.4	16.4	15.2	15.7	17.1	14.7	15.2	16.9
-0.5	16.0	15.0	15.0	16.3	14.4	15.2	16.1
-0.6	15.5	14.6	14.2	15.6	13.9	15.0	15.6
-0.7	15.0	14.0	13.8	15.3	13.5	14.4	15.3
-0.8	15.1	14.0	13.5	15.5	12.9	13.7	15.3
-0.9	16.7	14.2	13.2	15.5	12.2	13.1	16.0
-1.0	-	14.1	12.8	16.1	11.5	-	-
-1.1	-	9.7	-	-	-	-	-
-1.2	-	3.6	-	-	-	-	-

* The electrode potential relative to the pzc of 1 $mmol\ l^{-1}$ $HClO_4$

** Reference data for polarised interface

- No observation

Differential capacity $C/\mu F\ cm^{-2}$ on mercury electrode at 125 Hz, at 25°C for 1 mmol l^{-1} solutions of various pH values adjusted by the successive addition of $HClO_4$

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 1157.

	$C/\mu F\ cm^{-2}$							
pH of soln.	11	10	9	8	6	5	4	$TlNO_3$ 1 mmol l^{-1} pH 5.2
$-\xi/V$								
0.2	9.0	10.0	10.5	10.7	11.0	10.8	12.1	11.4
0.25	7.0	9.8	10.7	11.0	11.5	10.5	12.0	12.2
0.3	7.4	11.8	12.1	13.0	13.2	12.2	13.1	14.0
0.35	8.7	13.5	13.9	14.7	15.5	14.0	14.1	15.8
0.4	9.1	14.7	15.4	15.4	16.4	14.8	14.8	16.8
0.5	7.8	14.3	15.4	15.9	16.0	15.2	15.4	16.7
0.6	5.7	12.8	14.4	15.1	15.2	14.8	15.5	15.9
0.7	3.5	10.5	13.3	14.4	14.6	14.6	15.4	15.4
0.8	2.5	7.1	11.0	13.8	13.9	14.0	15.2	14.9
0.9	2.1	4.3	8.0	12.3	13.4	13.5	15.7	15.2
1.0	2.0	3.0	5.1	10.0	12.0	13.3	16.4	14.1
1.1	2.0	2.8	3.5	6.5	9.4	12.2	16.8	13.2
1.2	1.6	2.3	2.4	3.4	5.8	10.0	16.6	8.1
1.3	1.0	1.6	1.6	1.9	2.6	6.2	16.0	4.8
1.4	0.6	1.1	1.1	1.3	1.8	2.8	14.5	1.8

ξ : the electrode potential relative to the half-wave potential of Tl(I) reduction

Differential capacity $C/\mu F\ cm^{-2}$ on mercury electrode at 125 Hz, at 25°C for various concentrations of TlOH.

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 1157.

	$C/\mu F\ cm^{-2}$					
Concn. of $TlOH/mmol\ l^{-1}$	0.5	1.0	2.0	3.0	4.0	8.0
$-\xi/V$						
0.2	8.0	9.0	9.8	12.0	13.9	16.2
0.25	6.1	7.0	9.0	10.4	11.4	13.9
0.3	6.2	7.4	10.4	11.4	11.7	15.4
0.35	7.5	8.7	11.0	12.1	13.0	17.0
0.4	7.8	9.1	11.0	12.0	13.0	16.6
0.5	6.8	7.8	9.0	10.5	11.4	13.7
0.6	5.0	5.7	7.0	8.2	9.3	11.1
0.7	2.8	3.5	4.6	5.6	7.1	8.8
0.8	2.3	2.5	3.0	4.0	5.0	6.2
0.9	1.7	2.1	2.4	3.3	3.0	4.5
1.0	1.8	2.0	2.2	2.6	2.6	3.6
1.1	1.8	2.0	2.0	2.3	2.3	3.0
1.2	1.5	1.6	1.6	1.9	1.9	2.6
1.3	0.9	1.0	1.1	1.4	1.4	2.5

ξ : the electrode potential relative to the half-wave potential of Tl(I) reduction

Frequency dependence of the differential capacity (μF cm^{-2})
on mercury electrode at 25°C. Electrolyte 10^{-3} mol l^{-1} TlOH.

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 1157.

Frequency/Hz	31	62.5	125	250	500
$-\xi$/V					
0.15	-	23.0	16.0	8.8	7.4
0.2	12.5	9.2	9.0	6.4	5.5
0.25	11.9	9.7	7.0	6.9	5.7
0.3	14.5	10.8	7.4	7.9	6.5
0.35	15.7	11.3	8.7	8.0	6.5
0.4	15.5	11.2	9.1	7.7	6.2
0.5	14.0	10.2	7.8	6.5	5.0
0.6	11.3	8.2	5.7	4.9	3.6
0.7	7.5	5.2	3.5	3.1	2.4
0.8	5.2	3.7	2.5	2.0	1.1
0.9	6.9	3.4	2.1	1.0	0,6
1.0	6.8	3.0	2.0	0.9	0.5
1.1	6.7	3.0	2.0	0.9	0.4
1.2	5.1	2.4	1.6	0.6	0.3
1.3	3.0	1.2	1.0	0.4	0.2

ξ : the electrode potential relative to the half-wave potential of Tl(I) reduction

Differential capacity on mercury electrode in KNO_3 - given are the capacity in μF cm^{-2}, potential in V. Potentials given with respect to 1 mol l^{-1} calomel electrode at 20°C.

Reference: S. Minc, M. Jurkiewicz and J. Jastrzebska. J. Electroanal. Chem. 10 (1965) 437.

0.0005 mol l^{-1} KNO_3

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.232	18.69	0.219	25.17	0.207	26.48	0.195	26.00
0.332	17.74	0.319	22.17	0.307	23.04	0.295	20.83
0.432	8.22	0.399	13.30	0.407	11.39	0.395	11.30
0.442	7.26	0.419	10.13	0.417	9.96	0.415	8.96
0.452	6.43	0.429	9.04	0.427	9.00	0.425	8.09
0.462	5.83	0.439	7.74	0.437	7.96	0.435	7.17
0.472	5.83	0.449	6.91	0.447	7.04	0.445	6.55
0.482	5.87	0.459	6.43	0.457	6.61	0.455	6.04
0.492	6.04	0.469	6.09	0.467	6.17	0.465	5.65
0.502	6.35	0.479	6.09	0.477	6.09	0.475	5.43
0.532	7.91	0.489	6.22	0.487	6.09	0.485	5.35
0.632	12.39	0.499	6.48	0.497	6.26	0.495	5.39
0.732	13.17	0.519	7.39	0.507	6.61	0.595	9.23
0.832	15.22	0.619	11.96	0.607	11.96	0.695	11.74
0.932	16.87	0.719	15.48	0.707	14.26	0.795	13.13
1.032	16.30	0.819	15.52	0.807	15.65	0.895	13.26
				0.907	15.52	0.955	13.39

Differential capacity on mercury in KNO_3 (continued)

0.001 mol l^{-1} KNO_3

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.248	18.51	0.286	21.19	0.275	21.19	0.264	22.84
0.348	18.54	0.336	21.30	0.325	21.27	0.314	20.66
0.448	14.84	0.436	10.22	0.425	10.30	0.414	10.29
0.468	10.91	0.456	8.18	0.445	8.40	0.434	7.57
0.478	8.00	0.466	7.58	0.455	7.46	0.444	7.65
0.488	7.49	0.476	7.00	0.465	7.05	0.454	6.99
0.498	7.27	0.486	6.59	0.475	6.57	0.464	6.58
0.508	7.02	0.496	6.45	0.485	6.23	0.474	6.26
0.518	6.91	0.506	6.41	0.495	6.08	0.484	6.05
0.528	6.98	0.516	6.56	0.505	6.16	0.494	5.97
0.538	7.34	0.526	6.70	0.515	6.23	0.504	6.05
0.598	10.36			0.525	6.42	0.514	6.13
						0.614	10.25
						0.714	13.66

Differential capacity on mercury in KNO_3 (continued)

0.002 mol l^{-1} KNO_3

0°C		20°C		40°C		60°C	
$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}	$-E$/V	C/μF cm^{-2}
0.264	23.10	0.253	22.00	0.243	22.91	0.234	24.31
0.364	22.84	0.353	21.58	0.343	20.43	0.334	20.41
0.464	13.02	0.453	11.75	0.443	10.72	0.434	10.81
0.484	11.03	0.473	9.86	0.463	9.28	0.454	9.15
0.494	10.34	0.483	9.21	0.473	8.74	0.464	8.66
0.504	10.04	0.493	8.75	0.483	8.39	0.474	8.22
0.514	9.47	0.503	8.67	0.493	8.15	0.484	7.94
0.524	9.51	0.513	8.56	0.503	7.96	0.494	7.78
0.534	9.71	0.523	8.67	0.513	8.10	0.504	7.74
0.564	11.18	0.533	8.94	0.523	8.21	0.514	7.82
0.664	14.82	0.553	9.60	0.543	8.85	0.534	8.23
		0.653	11.37				

Differential capacity on mercury in KNO_3 (continued)

0.003 mol l^{-1} KNO_3

0°C		20°C		40°C		60°C	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.270	21.84	0.263	22.13	0.254	23.81	0.245	24.39
0.370	21.80	0.363	21.42	0.354	21.59	0.395	20.36
0.470	13.10	0.463	12.57	0.454	12.26	0.445	11.66
0.490	11.76	0.483	11.07	0.474	10.79	0.465	10.31
0.510	10.70	0.503	10.00	0.494	9.75	0.485	9.42
0.520	10.54	0.513	9.84	0.504	9.46	0.495	9.19
0.530	10.70	0.523	9.80	0.514	9.33	0.505	9.01
0.540	10.77	0.533	9.84	0.524	9.33	0.515	8.97
0.550	11.05	0.543	10.20	0.534	9.46	0.525	9.06
		0.563	10.76			0.545	10.58
		0.663	14.18				
		0.763	15.34				
		0.863	15.22				

Differential capacity on mercury in KNO_3 (continued)

$0.005\ \mathrm{mol\ l^{-1}}\ KNO_3$

0°C		20°C		40°C		60°C	
$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.285	17.43	0.276	18.46	0.267	18.63	0.259	19.74
0.385	17.23	0.376	17.54	0.367	16.70	0.359	16.28
0.485	11.34	0.476	11.08	0.467	10.41	0.459	9.70
0.505	10.57	0.496	10.12	0.487	9.47	0.479	8.88
0.515	10.08	0.506	9.66	0.497	8.97	0.489	8.62
0.525	10.00	0.516	9.63	0.507	8.85	0.499	8.45
0.535	9.83	0.526	9.54	0.517	8.70	0.509	8.36
0.545	9.94	0.536	9.57	0.527	8.70	0.519	8.36
0.555	10.03	0.546	9.63	0.537	8.79	0.529	8.42
0.585	10.63	0.556	9.85	0.547	8.85	0.539	8.52
0.685	12.23	0.576	12.62	0.567	9.26	0.559	8.88
0.785	12.28	0.676	13.00	0.667	11.36	0.659	10.95
0.885	11.71	0.776	12.62	0.767	11.98	0.759	11.91
0.985	11.06	0.876	12.06	0.867	11.92	0.859	11.97
1.085	10.60	0.976	11.56	0.967	11.50	0.959	11.77
				1.067	11.24	1.059	11.61
						1.159	11.71

Differential capacity on mercury in KNO_3 (continued)

$0.006\ mol\ l^{-1}\ KNO_3$

$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$	$-E/V$	$C/\mu F\ cm^{-2}$
0.289	17.76	0.290	17.94	0.288	18.87	0.270	19.86
0.389	17.33	0.390	16.68	0.388	15.82	0.370	15.76
0.489	11.39	0.490	10.59	0.488	9.88	0.470	9.66
0.509	10.73	0.510	10.03	0.508	9.49	0.490	9.16
0.519	10.54	0.530	9.85	0.518	9.44	0.500	9.02
0.529	10.52	0.540	9.91	0.528	9.47	0.510	8.96
0.539	10.52	0.550	10.00	0.538	9.55	0.520	8.99
0.549	10.70	0.590	10.79	0.548	9.67	0.530	9.09
0.569	11.06	0.690	12.38	0.588	10.56	0.570	9.83
0.589	11.42	0.790	12.47				

Capacity data for mercury in 0.05M H_2SO_4 solution. T = 25°C
Reference electrode: SCE (NaCl solution).

Reference: K.G. Baikerikar and R.S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

E/V	$C/\mu F\ cm^{-2}$	E/V	$C/\mu F\ cm^{-2}$
-1.100	15.18	-0.250	33.09
-1.050	15.28	-0.225	33.72
-1.000	15.43	-0.200	34.06
-0.950	15.69	-0.175	34.18
-0.900	16.02	-0.150	34.13
-0.850	16.45	-0.125	34.00
-0.800	16.99	-0.100	33.84
-0.750	17.64	-0.050	33.73
-0.700	18.29	0.000	34.11
-0.650	18.94	0.050	35.13
-0.600	19.52	0.100	36.77
-0.550	19.93	0.150	39.15
-0.500	20.37	0.175	40.56
-0.450	21.72	0.200	42.28
-0.425	22.60	0.225	44.23
-0.400	24.03	0.250	46.45
-0.375	25.79	0.275	49.13
-0.350	27.62	0.300	52.66
-0.325	29.25	0.325	57.50
-0.300	30.88	0.350	66.25
-0.275	32.11		

Differential capacity on a mercury electrode in aqueous $NaBF_4$ at 25°C.
Reference electrode: saturated KCl calomel for solutions up to 0.15 mol l^{-1},
saturated NaCl calomel for those of 0.20 M and above.
Given are the capacity in μF cm^{-2}, potential in V.
Reference: A. de Battisti and S. Trasatti. J. Electroanal. Chem. 59 (1975) 137.

c/mol l^{-1}	0.01	0.015	0.02	0.028	0.04	0.054	0.07	0.1
E/V								
0.20	22.57	22.71	23.04	23.05	23.10	23.27	23.33	23.50
0.15	21.86	21.95	22.09	22.15	22.22	22.36	22.44	22.54
0.10	21.43	21.49	21.64	21.66	21.72	21.80	21.85	21.91
0.05	21.36	21.39	21.52	21.48	21.51	21.57	21.59	21.63
0.00	21.66	21.65	21.71	21.68	21.64	21.64	21.66	21.66
-0.05	22.34	22.11	22.24	22.16	22.13	22.06	22.03	22.05
-0.10	22.95	22.87	23.00	22.90	22.90	22.80	22.75	22.75
-0.15	23.63	23.68	23.84	23.74	23.82	23.73	23.66	23.68
-0.20	23.93	24.20	24.44	24.50	24.66	24.66	24.67	24.76
-0.25	23.24	23.90	24.36	24.66	25.04	25.23	25.38	25.63
-0.30	21.36	22.34	22.97	23.75	24.50	24.97	25.35	25.94
-0.35	17.95	19.46	20.46	21.63	22.80	23.63	24.36	25.32
-0.40	14.34	16.34	17.46	18.96	20.49	21.64	22.63	23.95
-0.45	13.02	14.72	15.86	17.31	18.84	19.97	20.97	22.41
-0.50	13.93	15.20	16.17	17.20	18.40	19.27	20.07	21.27
-0.55	15.52	16.42	17.16	17.81	18.66	19.21	19.72	20.56
-0.60	16.81	17.46	17.84	18.37	18.91	19.19	19.49	20.04
-0.65	17.65	17.98	18.18	18.54	18.86	18.96	19.10	19.48
-0.70	17.72	18.00	18.30	18.35	18.55	18.54	18.59	18.84
-0.75	17.57	17.75	17.91	17.96	18.07	18.03	18.01	18.17
-0.80	17.27	17.33	17.38	17.47	17.54	17.46	17.42	17.55
-0.85	16.74	16.93	16.92	16.97	17.02	16.93	16.90	16.99
-0.90	16.21	16.50	16.47	16.53	16.55	16.48	16.44	16.52
-0.95	15.90	16.12	16.13	16.16	16.18	16.12	16.07	16.16
-1.00	15.68	15.80	15.79	15.87	15.89	15.85	15.84	15.89
-1.05	15.45	15.60	15.64	15.68	15.71	15.71	15.70	15.76
-1.10	15.37	15.53	15.56	15.60	15.64	15.54	15.64	15.71
-1.15	15.30	15.53	15.56	15.61	15.62	15.62	15.67	15.73
-1.20	15.45	15.56	15.72	15.66	15.70	15.73	15.74	15.80
-1.25	15.60	15.69	15.87	15.75	15.81	15.84	15.90	15.95
-1.30	15.83	15.78	15.90	15.94	16.00	16.06	16.07	16.14
-1.35	15.98	15.99	16.06	16.14	16.23	16.29	16.33	16.38
-1.40	16.13	16.23	16.40	16.38	16.51	16.52	16.61	16.69
-1.45	16.36	16.53	16.59	16.66	16.81	16.87	16.93	16.94
-1.50	16.58	16.82	16.77	17.07	17.13	17.24	17.29	17.42
pzc/V	0.4344	0.4352	0.4354	0.4365	0.4371	0.4378	0.4392	0.4400

Differential capacity on mercury in aqueous $NaBF_4$ (continued)

$c/mol\,l^{-1}$	0.15	0.20	0.27	0.38	0.50	0.65	0.8	1.0
E/V								
0.20	23.65	23.57	23.71	23.85	23.91	24.02	24.12	24.20
0.15	22.70	22.64	22.74	22.86	22.93	23.07	23.12	23.20
0.10	22.04	22.04	22.07	22.15	22.21	22.30	22.40	22.45
0.05	21.67	21.67	21.50	21.73	21.76	21.78	21.80	21.85
0.00	21.64	21.58	21.54	21.60	21.62	21.70	21.79	21.88
-0.05	21.92	21.85	21.84	21.86	21.86	21.88	21.90	21.93
-0.10	22.56	22.45	22.50	22.43	22.45	22.40	22.37	22.35
-0.15	23.50	23.37	23.38	23.27	23.28	23.19	23.16	23.14
-0.20	24.60	24.50	24.50	24.45	24.48	24.35	24.27	24.20
-0.25	25.64	25.66	25.60	25.64	25.59	25.58	25.56	25.42
-0.30	26.25	26.48	26.61	26.74	26.85	26.76	26.82	26.71
-0.35	26.10	26.63	27.00	27.35	27.65	27.64	27.82	27.73
-0.40	25.09	26.00	26.61	27.23	27.71	27.89	28.20	28.18
-0.45	23.67	24.75	25.53	26.32	26.95	27.32	27.78	27.86
-0.50	22.34	23.32	24.12	24.89	25.65	26.09	26.62	26.77
-0.55	21.28	22.04	22.68	23.36	24.02	24.43	25.00	25.22
-0.60	20.43	20.95	21.29	21.85	22.27	22.72	23.23	23.43
-0.65	19.61	19.95	20.24	20.48	20.88	21.12	21.55	21.73
-0.70	18.83	19.05	19.12	19.35	19.55	19.77	20.10	20.24
-0.75	18.10	18.24	18.33	18.38	18.58	18.64	18.91	19.01
-0.80	17.42	17.51	17.58	17.52	17.72	17.75	17.96	17.99
-0.85	16.86	16.94	16.97	16.92	17.04	17.03	17.22	17.24
-0.90	16.40	16.45	16.49	16.44	16.50	16.52	16.69	16.70
-0.95	16.05	16.11	16.15	16.10	16.16	16.18	16.32	16.33
-1.00	15.82	15.89	15.90	15.86	15.93	15.98	16.09	16.10
-1.05	15.71	15.76	15.78	15.78	15.82	15.89	15.98	16.01
-1.10	15.70	15.73	15.73	15.78	15.81	15.90	15.96	16.01
-1.15	15.73	15.76	15.77	15.83	15.88	15.97	16.03	16.07
-1.20	15.82	15.87	15.91	15.95	16.02	16.10	16.16	16.23
-1.25	15.97	16.04	16.06	16.12	16.19	16.27	16.36	16.41
-1.30	16.19	16.25	16.28	16.35	16.44	16.52	16.60	16.65
-1.35	16.44	16.51	16.53	16.63	16.72	16.80	16.89	16.94
-1.40	16.75	16.78	16.83	16.95	17.04	17.09	17.19	17.26
-1.45	17.05	17.11	17.18	17.28	17.39	17.47	17.54	17.65
-1.50	17.44	17.48	17.56	17.66	17.76	17.86	17.95	17.99
pzc/V	0.4409	0.4475	0.4480	0.4486	0.4480	0.4475	0.4488	0.4499

Differential capacity of mercury in tenth normal solutions of metallic chlorides. $T = 25^{\circ}C$. Potentials measured with respect to calomel electrode in same solution.

Reference: D.C. Grahame, J. Electrochem. Soc., 98, 343 (1951) #

E	Capacity in microfarads per square centimeter								
Volts	$CaCl_2$	$SrCl_2$	$BaCl_2$	$MnCl_2$	$CoCl_2$	$NiCl_2$	$AlCl_3$	$LaCl_3$	$PrCl_3$
-0.061				103.4		104.1		110.9	112.4
-0.062		98.0	106.0	101.2	100.8	101.8	100.4	108.5	109.9
-0.063	99.2	96.4	103.6	99.1	98.7	99.7	98.5	106.0	107.4
-0.065	95.4	93.8	99.6	95.3	94.9	95.9	95.3	101.9	103.1
-0.070	87.6	88.4	91.1	87.4	87.1	87.9	88.6	93.0	94.0
-0.075	81.2	82.8	84.2	80.8	80.7	81.3	82.4	85.8	86.7
-0.080	75.9	77.5	78.7	75.6	75.4	76.0	77.0	80.0	80.7
-0.085	71.4	72.9	73.9	71.1	71.0	71.5	72.5	75.2	75.8
-0.090	67.6	69.0	69.9	67.4	67.3	67.8	68.7	71.0	71.6
-0.095	64.5	65.6	66.4	64.2	64.1	64.6	65.5	67.5	68.0
-0.100	62.0	63.0	63.7	61.7	61.6	62.1	62.9	64.7	65.2
-0.105	59.7	60.5	61.2	59.4	59.4	59.7	60.5	62.2	62.6
-0.110	57.7	58.5	59.1	57.4	57.4	57.7	58.5	60.1	60.4
-0.115	55.9	56.6	57.1	55.6	55.6	55.9	56.7	58.1	58.4
-0.120	54.3	54.9	55.3	54.0	54.0	54.2	55.0	56.3	56.6
-0.125	52.8	53.3	53.7	52.5	52.6	52.7	53.4	54.6	54.9
-0.130	51.5	51.9	52.3	51.2	51.2	51.4	52.0	53.2	53.5
-0.135	50.3	50.8	51.1	50.1	50.1	50.3	50.8	52.0	52.2
-0.140	49.2	49.6	49.9	49.0	49.1	49.2	49.7	50.7	51.0
-0.145	48.3	48.7	49.0	48.1	48.2	48.3	48.8	49.8	50.0
-0.150	47.4	47.8	48.0	47.2	47.2	47.3	47.8	48.7	49.0
-0.155	46.6	46.9	47.1	46.3	46.4	46.5	46.9	47.8	48.1
-0.16	45.82	46.09	46.36	45.60	45.64	45.78	46.18	46.98	47.25
-0.17	44.43	44.64	44.86	44.21	44.25	44.38	44.73	45.47	45.64
-0.18	43.28	43.45	43.62	43.07	43.11	43.24	43.58	44.21	44.38
-0.19	42.25	42.46	42.58	42.04	42.12	42.25	42.54	43.12	43.24
-0.20	41.40	41.56	41.69	41.20	41.28	41.40	41.69	42.17	42.30
-0.21	40.72	40.84	40.96	40.50	40.61	40.71	40.97	41.40	41.49
-0.22	40.08	40.20	40.30	39.87	39.99	40.09	40.34	40.71	40.77
-0.23	39.59	39.67	39.76	39.37	39.49	39.58	39.82	40.16	40.18
-0.24	39.16	39.27	39.34	38.98	39.10	39.18	39.41	39.71	39.73
-0.25	38.84	38.95	39.00	38.68	38.79	38.86	39.11	39.38	39.34
-0.26	38.62	38.70	38.78	38.48	38.56	38.65	38.87	39.11	39.02
-0.27	38.47	38.54	38.58	38.35	38.41	38.51	38.72	38.94	38.78
-0.28	38.41	38.47	38.48	38.31	38.34	38.45	38.66	38.86	38.63
-0.29	38.42	38.46	38.47	38.34	38.34	38.47	38.66	38.87	38.55
-0.30	38.48	38.52	38.52	38.44	38.41	38.55	38.74	38.93	38.52
-0.31	38.55	38.59	38.60	38.53	38.50	38.64	38.84	39.02	38.58
-0.32	38.65	38.71	38.73	38.67	38.63	38.78	38.98	39.15	38.65
-0.33	38.79	38.86	38.89	38.83	38.79	38.94	39.16	39.31	38.71
-0.34	38.93	39.02	39.07	39.00	38.96	39.11	39.34	39.48	38.81
-0.35	39.06	39.17	39.24	39.16	39.12	39.28	39.51	39.63	38.93
-0.36	39.25	39.39	39.46	39.38	39.34	39.49	39.74	39.86	39.08
-0.37	39.48	39.65	39.71	39.63	39.60	39.74	39.98	40.10	39.27
-0.38	39.62	39.80	39.88	39.79	39.76	39.90	40.14	40.26	39.38
-0.39	39.72	39.92	40.00	39.91	39.88	40.02	40.24	40.37	39.47
-0.40	39.83	40.06	40.14	40.02	40.02	40.14	40.35	40.51	39.56

cont.

# See note on p. Dl-Hg-9

Capacity in 0.1 N solutions of metallic chlorides (cont.)

E	Capacity in microfarads per square centimeter								
Volts	$CaCl_2$	$SrCl_2$	$BaCl_2$	$MnCl_2$	$CoCl_2$	$NiCl_2$	$AlCl_3$	$LaCl_3$	$PrCl_3$
-0.41	39.82	40.09	40.16	40.03	40.03	40.17	40.34	40.49	39.56
-0.42	39.72	40.01	40.08	39.94	39.93	40.10	40.20	40.39	39.47
-0.43	39.56	39.89	39.95	39.80	39.78	39.99	40.02	40.25	39.32
-0.44	39.31	39.67	39.72	39.57	39.54	39.77	39.76	40.00	39.09
-0.45	38.96	39.34	39.40	39.23	39.19	39.46	39.39	39.66	38.73
-0.46	38.51	38.92	38.98	38.79	38.74	39.02	38.94	39.22	38.31
-0.47	37.96	38.40	38.47	38.25	38.20	38.46	38.39	38.67	37.78
-0.48	37.38	37.84	37.92	37.66	37.61	37.86	37.80	38.10	37.21
-0.49	36.65	37.13	37.22	36.92	36.88	37.10	37.05	37.36	36.51
-0.50	35.81	36.27	36.36	36.07	36.03	36.21	36.17	36.50	35.67
-0.51	34.88	35.35	35.44	35.16	35.08	35.26	35.23	35.59	34.77
-0.52	33.98	34.41	34.51	34.25	34.22	34.32	34.30	34.70	33.95
-0.53	33.06	33.49	33.58	33.30	33.29	33.36	33.34	33.79	33.09
-0.54	32.10	32.50	32.59	32.29	32.31	32.39	32.37	32.82	32.18
-0.55	31.14	31.54	31.59	31.29	31.35	31.48	31.45	31.86	31.28
-0.56	30.28	30.70	30.63	30.38	30.44	30.55	30.67	31.06	30.53
-0.57	29.42	29.87	29.81	29.48	29.55	29.72	29.90	30.25	29.75
-0.58	28.59	29.07	29.03	28.66	28.70	28.92	29.18	29.49	28.99
-0.59	27.92	28.36	28.34	27.94	28.01	28.09	28.50	28.87	28.40
-0.60	27.23	27.65	27.66	27.23	27.34	27.46	27.85	28.31	27.75
-0.61	26.65	27.05	27.05	26.61	26.73	26.84	27.28	27.82	27.30
-0.62	26.12	26.49	26.53	26.06	26.14	26.27	26.83	27.30	26.79
-0.63	25.58	25.93	26.00	25.50	25.56	25.72	26.25	26.78	26.27
-0.64	25.10	25.43	25.51	25.03	25.07	25.23	25.94	26.30	25.85
-0.65	24.67	24.98	25.07	24.61	24.61	24.80	25.55	25.86	25.42
-0.66	24.25	24.53	24.62	24.18	24.19	24.35	25.13	25.44	25.02
-0.67	23.86	24.10	24.18	23.77	23.79	23.92	24.70	25.02	24.61
-0.68	23.52	23.74	23.80	23.39	23.41	23.54	24.32	24.63	24.22
-0.69	23.17	23.38	23.44	23.03	23.05	23.16	23.96	24.24	23.81
-0.70	22.85	23.06	23.11	22.70	22.72	22.81	23.63	23.88	23.42
-0.72	22.20	22.37	22.43	22.06	22.07	22.15	22.94	23.14	22.73
-0.74	21.56	21.73	21.78	21.44	21.46	21.54	22.22	22.43	22.07
-0.76	21.03	21.14	21.22	20.92	20.94	21.02	21.60	21.79	21.49
-0.78	20.55	20.62	20.68	20.45	20.47	20.55	21.04	21.16	20.94
-0.80	20.07	20.15	20.17	19.97	20.00	20.06	20.48	20.61	20.37
-0.82	19.59	19.66	19.67	19.49	19.52	19.57	19.94	20.10	19.83
-0.84	19.17	19.25	19.25	19.06	19.09	19.12	19.45	19.65	19.38
-0.86	18.81	18.89	18.89	18.67	18.71	18.71	19.02	19.26	18.98
-0.88	18.46	18.53	18.53	18.31	18.36	18.36	18.61	18.88	18.60
-0.90	18.12	18.19	18.19	17.99	18.02	18.05	18.25	18.54	18.27
-0.92	17.80	17.87	17.87	17.68	17.70	17.76	17.91	18.25	17.96
-0.94	17.51	17.58	17.58	17.39	17.40	17.48	17.58	17.91	17.67
-0.96	17.27	17.34	17.33	17.16	17.14		17.31	17.67	17.43
-0.98	17.07	17.13	17.13	16.95	16.92		17.08	17.45	17.24
-1.00	16.92	16.98	16.97	16.79	16.74		16.90	17.29	17.09
-1.02	16.70	16.77	16.76	16.58	16.51		16.67	17.07	16.89
-1.04	16.56	16.62	16.62	16.44	16.34		16.49	16.92	16.75
-1.06	16.45	16.52	16.52	16.33	16.20		16.37	16.81	16.65
-1.08	16.35	16.43	16.43	16.23	16.08			16.72	16.56
-1.10	16.25	16.35	16.34	16.13	15.96			16.62	16.47
-1.12	16.20	16.31	16.31	16.09	15.89			16.58	16.44
-1.14	16.15	16.27	16.27	16.04	15.82			16.54	16.40

cont.

Capacity in 0.1 N solutions of metallic chlorides (cont.)

E Volts	Capacity in microfarads per square centimeter					
	$CaCl_2$	$SrCl_2$	$BaCl_2$	$MnCl_2$	$LaCl_3$	$PrCl_3$
-1.16	16.12	16.25	16.25	16.02	16.51	16.39
-1.18	16.12	16.24	16.25	16.01	16.51	16.40
-1.20	16.11	16.25	16.26	16.01	16.52	16.41
-1.22	16.13	16.26	16.29	16.03	16.56	16.45
-1.24	16.14	16.29	16.31	16.06	16.59	16.49
-1.26	16.18	16.33	16.36	16.09	16.65	16.56
-1.28	16.25	16.41	16.44	16.17	16.74	16.65
-1.30	16.32	16.49	16.53	16.24	16.84	16.75
-1.32	16.40	16.59	16.62	16.33	16.94	16.86
-1.34	16.48	16.67	16.70	16.40	17.04	16.97
-1.36	16.58	16.77	16.79	16.50	17.15	17.09
-1.38	16.72	16.91	16.94	16.66	17.32	17.27
-1.40	16.86	17.05	17.07	16.80	17.50	
-1.42	17.00	17.18	17.20	16.95	17.67	
-1.44	17.15	17.32	17.33	17.11	17.86	
-1.46	17.31	17.47	17.47			
-1.48	17.49	17.63	17.64			
-1.50	17.67	17.80	17.80			
-1.52	17.86	17.96	17.96			
-1.54	18.04	18.12	18.12			
-1.56	18.24	18.29	18.29			
-1.58	18.47	18.50	18.50			
-1.60	18.70	18.71	18.71			
-1.62	18.91	18.91	18.91			
-1.64	19.13	19.11	19.11			
-1.66	19.36	19.32	19.32			
-1.68	19.58	19.51	19.54			
-1.70	19.80	19.71	19.77			
-1.72	19.99	19.89	19.97			
-1.74	20.19	20.07	20.18			
-1.76	20.39	20.23	20.39			
-1.78	20.63	20.43	20.66			
-1.80	20.91	20.63	20.96			
-1.82	21.22	20.88	21.30			
-1.84	21.57	21.14	21.68			

Note. The pH of these solutions was not mentioned bij the author but is expected to be about neutral. The presence of hydrolyzed species, if any (important in the case of non-metallic double layers) can probably be neglected because the work of Damaskin and Nikolaeva-Fedorovich (Nauchnye Doklady Vysshei Shkoly, Khim i Khim Tekhn, 1959, No 1) has shown that there is very little effect until the pH becomes as high as 11.

Differential capacity on mercury in aqueous solutions of $Ca(ClO_4)_2$.
Concentration in eq/l; Data given as a function of surface charge σ.
Temperature 25°C. 400 Hz.
Reference: A.N. Frumkin, V.F. Ivanov, B.B. Damaskin, A.A. Ivashchenko, N.I. Peshkova, Elektrokhimiya, 1, 279 (1965).

conc. 1.0		conc. 5.8		conc. 6.6		conc. 8.1		conc. 9.3	
σ μ C/cm^2	C μ F/cm^2	σ μ C/cm^2	C μ F/cm^2	σ μ C/cm^2	C μ F/cm^2	σ μ C/cm^2	C μ F/cm^2	σ μ C/cm^2	C μ F/cm^2
21.8	36.8	20.3	29.5	19.5	27.5	21.4	28.0	19.7	24.9
18.4	31.3	17.6	26.3	16.9	25.1	18.7	25.3	17.3	22.7
15.5	27.5	15.0	24.5	14.4	23.6	16.3	23.7	15.1	21.4
12.8	25.4	12.6	23.9	12.1	22.9	14.0	22.6	13.0	20.7
10.4	24.5	10.2	23.7	9.8	22.9	11.7	22.1	11.0	20.3
7.9	24.5	7.8	24.2	7.52	23.3	9.5	22.0	8.9	20.4
5.4	25.5	5.3	25.5	5.1	24.3	7.3	22.6	6.9	21.1
2.8	27.1	2.7	26.7	2.6	25.6	5.0	23.6	4.7	22.1
0	28.2	0	28.0	0	27.0	2.6	24.8	2.4	23.5
- 2.76	27.0	- 2.8	28.8	- 2.8	28.5	0	26.2	0	25.0
- 5.3	23.5	- 5.7	28.9	- 5.7	29.3	- 2.7	27.5	- 2.6	26.5
- 7.5	20.1	- 8.6	28.0	- 8.6	29.1	- 5.5	28.8	- 5.3	27.8
- 9.3	18.2	-11.3	26.5	-11.4	28.2	- 8.4	29.7	- 8.1	28.8
-11.1	17.2	-13.8	24.5	-14.2	26.7	-11.4	29.8	-11.0	29.7
-12.8	16.7	-16.2	23.0	-16.8	24.9	-14.3	29.3	-14.0	30.1
-14.4	16.9	-18.5	22.4	-19.2	23.9	-17.2	28.3	-17.0	30.1
-16.2	17.6	-20.7	22.2	-21.6	23.7	-20.0	27.5	-20.0	29.8
-18.0	18.5	-23.0	22.6	-24.0	24.2	-22.7	26.8	-23.0	29.1
-20.0	19.5	-25.6	23.6	-26.5	24.9	-25.4	26.3	-25.9	28.7
-21.9	19.6	-27.7	24.9	-29.0	26.0	-28.1	27.0	-28.8	28.7
-24.0	21.9	-30.3	26.3	-31.7	27.4	-30.9	28.0	-31.7	29.6

Differential capacity on mercury in aqueous solutions of $CaCl_2$.
Concentration in eq/l; Data given as a function of surface charge σ.
Temperature 25°C. 400 Hz.
Reference: A.N. Frumkin, V.F. Ivanov, B.B. Damaskin, A.A. Ivashchenko, N.I. Peshkova, Elektrokhimiya, 1, 279 (1965).

conc. 1.0		conc. 9.3	
σ / $\mu C/cm^2$	C / $\mu F/cm^2$	σ / $\mu C/cm^2$	C / $\mu F/cm^2$
16.0	46.8	13.7	40.5
11.8	39.0	9.8	35.6
7.9	38.2	6.5	33.2
4.0	41.6	3.3	33.4
0	37.5	0	33.9
- 3.3	29.0	- 3.3	33.9
- 6.0	23.7	- 6.5	31.2
- 8.0	19.8	- 9.8	28.2
- 9.9	18.0	-12.5	25.1
-11.7	17.1	-14.8	23.3
-13.4	16.8	-17.0	22.4
-15.2	17.1	-19.3	22.4
-16.9	17.7	-21.5	23.1
-18.7	18.6	-23.9	23.8
-20.6	19.5	-26.3	25.0
-22.6	20.7	-28.9	26.5
-24.7	22.0		
-27.0	23.3		

Differential capacity of mercury in aqueous solutions containing $MgCl_2$

$T = 25^{\circ}C$

Potential with respect to Ag/AgCl electrode in the working solution.
Frequency 1.5 kHz

Reference: R. Parsons and S. Trasatti. Trans. Faraday Soc. 65 (1969) 3314

	$C/\mu F\ cm^{-2}$						
c_{MgCl_2}/mol l^{-1} E/volts	0.0025	0.0032	0.004	0.005	0.006	0.0075	0.01
-1.60	17.30	17.43	17.57	17.63	17.73	17.85	18.01
-1.50	16.59	16.65	16.71	16.76	15.85	16.95	17.03
-1.40	15.96	16.01	16.07	16.09	16.18	16.24	16.31
-1.35	15.70	15.77	15.82	15.89	15.95	15.99	16.06
-1.30	15.54	15.61	15.66	15.69	15.77	15.80	15.87
-1.25	15.56	15.60	15.64	15.65	15.70	15.71	15.73
-1.20	15.61	15.63	15.66	15.66	15.70	15.69	15.70
-1.15	15.84	15.83	15.84	15.80	15.79	15.75	15.73
-1.10	16.19	16.12	16.07	15.99	15.97	15.93	15.90
-1.05	16.50	16.43	16.38	16.30	16.30	16.24	16.19
-1.00	16.94	16.88	16.80	16.71	16.71	16.64	16.57
-0.95	17.51	17.46	17.37	17.29	17.27	17.20	17.11
-0.90	18.21	18.12	18.05	17.95	17.94	17.86	17.78
-0.85	19.03	18.97	18.87	18.79	18.74	18.69	18.57
-0.80	19.87	19.79	19.72	19.65	19.53	19.56	19.45
-0.75	20.33	20.38	20.42	20.41	20.43	20.37	20.29
-0.70	20.59	20.76	20.90	21.02	21.08	21.05	21.00
-0.67	-	-	-	20.88	21.10	21.30	-
-0.65	19.66	20.05	20.42	20.73	21.09	21.27	21.70
-0.62	18.19	18.78	19.38	20.00	20.48	21.03	21.68
-0.60	16.76	17.73	18.63	19.28	20.07	20.66	21.54
-0.57	-	-	17.46	18.38	19.35	20.14	-
-0.55	15.17	16.29	17.32	18.46	19.29	20.41	21.97
-0.53	16.46	-	-	-	-	-	-
-0.52	-	-	19.18	-	-	-	-
-0.50	20.11	20.72	21.37	22.16	22.78	23.77	25.23
-0.45	27.36	27.93	28.49	28.89	29.52	30.17	31.03
-0.40	33.06	33.48	33.73	34.17	34.36	34.62	34.88
-0.35	34.80	35.37	35.70	36.11	36.28	36.45	36.63
-0.30	35.07	35.45	35.79	36.13	36.33	36.41	36.57
-0.25	35.54	35.78	35.97	36.16	36.32	36.36	36.52
-0.20	36.14	36.39	36.60	36.80	36.93	36.96	37.10
-0.15	38.09	38.31	38.55	38.69	38.86	38.17	39.44
-0.10	42.51	42.94	43.34	43.70	44.11	44.38	44.88
-0.05	54.27	54.50	54.83	55.31	55.80	56.46	57.36

Differential capacity of mercury in aqueous solutions of $MgSO_4$ T = 25°C

Potentials with respect to a mercury - mercurous sulphate electrode in 1.52 mol l^{-1} $MgSO_4$

Reference: J.A. Harrison, J.E.B. Randles and D.J. Schiffrin, J. Electroanal. Chem., 25 (1970) 197.

c/mol l^{-1}	0.104	0.172	0.207	0.233	0.440	0.581	1.00	1.55	2.07	2.59
E/volts					C/μF cm^{-2}					
0.25	42.72	41.80	41.26	41.40	39.38	38.60	36.99	35.73	34.71	33.86
0.30	41.29	40.65	40.27	40.36	38.44	37.60	35.72	34.14	32.91	31.82
0.34	40.52	40.27	40.06	40.20	38.53	37.70	35.70	33.75	32.28	30.98
0.38	39.97	40.19	40.11	40.36	39.21	38.44	36.42	34.18	32.40	30.76
0.42	39.64	40.19	40.33	40.70	40.16	39.69	37.82	35.37	33.25	31.21
0.46	39.49	40.30	40.62	41.07	41.17	41.05	39.77	37.42	34.93	32.44
0.50	39.59	40.43	40.88	41.38	41.27	42.30	41.93	40.15	37.59	34.72
0.54	39.65	40.49	40.95	41.41	42.58	42.96	43.76	43.19	41.11	38.05
0.58	39.30	40.10	40.45	40.87	42.15	42.69	44.37	45.45	44.68	42.26
0.62	38.07	38.34	38.97	39.23	40.52	41.02	43.26	45.83	46.98	46.26
0.66	35.68	36.01	36.45	36.65	37.18	38.19	40.59	43.81	46.41	48.44
0.70	32.67	32.90	33.33	33.48	34.38	34.84	37.06	40.28	43.44	47.15
0.74	29.52	29.79	30.23	30.39	31.35	31.77	33.77	36.46	39.24	43.17
0.78	26.58	27.15	27.54	27.64	28.85	29.30	31.07	33.10	35.24	38.47
0.82	24.10	25.00	25.48	25.80	26.93	27.43	28.89	30.49	32.10	34.40
0.86	22.54	23.53	24.05	24.37	25.48	25.88	27.11	28.35	29.46	31.03
0.90	22.09	22.88	23.28	23.54	24.38	24.65	25.63	26.56	27.30	28.39
0.95	22.11	22.42	22.65	22.79	23.21	23.32	23.94	24.59	25.06	25.77
1.00	21.81	21.77	21.91	21.99	22.06	22.09	22.44	22.88	23.17	23.63
1.05	21.10	20.86	21.00	21.06	20.94	20.86	21.06	21.33	21.53	21.86
1.10	20.20	19.98	20.01	20.08	19.85	19.75	19.83	20.04	20.15	20.42
1.15	19.26	19.07	19.06	19.11	18.87	18.76	18.79	18.93	19.01	19.22
1.20	18.39	18.21	18.20	18.24	18.04	17.93	17.93	18.04	18.10	18.25
1.25	17.64	17.47	17.48	17.51	17.32	17.24	17.22	17.30	17.36	17.51
1.30	17.02	16.92	16.88	16.96	16.76	16.69	16.67	16.75	16.84	16.95
1.35	16.54	16.46	16.43	16.53	16.34	16.28	16.27	16.37	16.43	16.54
1.40	16.19	16.13	16.11	16.21	16.05	16.00	16.01	16.10	16.16	16.27
1.45	15.97	15.94	15.91	16.02	15.89	15.84	15.85	15.94	16.02	16.12
1.50	15.85	15.84	15.83	15.94	15.82	15.78	15.80	15.89	15.98	16.10
1.55	15.84	15.85	15.82	15.96	15.84	15.80	15.85	15.94	16.05	16.17
1.60	15.90	15.95	15.90	16.05	15.93	15.93	15.97	16.09	16.19	16.32
1.70	16.28	16.35	16.32	16.47	16.38	16.39	16.43	16.56	16.70	16.87
1.80	16.95	17.03	16.99		17.09	17.12	17.16	17.30	17.49	17.70

Differential capacity for mercury in $BaCl_2$, BaI_2 and $MgCl_2$ at 25°C. Given are the capacity in $\mu F\ cm^{-2}$, potential in mV vs SCE.

Reference: M. Sluyters-Rehbach, J.S.M.C. Brenkel and J.H. Sluyters. J. Electroanal. Chem. 48 (1973) 411.

$-E/mV$	$C\ /\mu F\ cm^{-2}$		
	0.5 M $MgCl_2$	0.5 M $BaCl_2$	0.5 M BaI_2
100		59.0	
150	43.7	46.5	
200	39.5	40.4	
250	38.2	38.2	
300	38.8	37.7	
350	40.9	39.2	
400	42.8	41.4	ca. 400
450	43.1	41.7	190
500	40.8	40.4	124
550	36.3	36.8	94.3
600	31.4	31.5	76.2
650	26.9	26.8	65.2
700	23.4	23.4	58.4
750	20.9	21.0	54.5
800	19.1	19.1	51.8
850	17.9	18.0	49.2
900	17.0	17.2	45.2
950	16.4	16.6	40.5
1000	16.1	16.4	34.7
1050	15.9	16.2	29.4
1100	15.8	16.2	25.1
1150	15.8	16.4	22.0
1200	15.9	16.5	20.2
1250	16.2	16.7	19.1
1300	16.4	17.2	18.5
1350	16.8	17.6	18.5
1400	17.2	17.9	18.7
1450	17.7	18.3	18.9
1500	18.2	18.8	19.3

Differential capacity for mercury in $BaCl_2$, BaI_2 and $MgCl_2$ (continued)

$-E/mV$	$C_d/\mu F\ cm^{-2}$		
	0.5 M $MgCl_2$	0.5 M $BaCl_2$	0.5 M BaI_2
1550	18.8	19.4	19.8
1600	19.4	20.1	20.4
1650	20.2	20.6	21.0
1700	20.8	21.5	21.8
1750	21.7	22.5	22.8
1800	22.4	23.8	23.9
1810		23.9	24.2
1820			24.7
1830		24.6	24.8
1839			25.4
1840		25.3	
1849		25.6	25.8
1850	23.4		
1858			26.1
1859		26.4	
1867			26.9
1868		26.9	
1875			27.3
1876		27.75	
1882			28.3
1883		29.3	
1900	24.3		
1950	25.5		
2000	26.9		
2050	28.8		

Differential double layer capacity on mercury in concentrated aqueous KF and in aqueous KF melts of various temperatures and compositions. Potentials relative to calomel electrode, saturated at 20°C.

Reference: V.S. Palankev, A.M. Skundin and V.S. Bagotskii, Elektrokhimiya, 3, 370 (1967).

electrolyte	KF 1 mole.l^{-1}	KF 4 mole.l^{-1}	KF. 6 H_2O melt^2		KF. 3 H_2O melt ^{2}	
Temp.	20°C	20°C	20°C	60°C	60°C	95°C
E / volts	C / μ F/cm^2	C / μ F/cm^2	C / μ F/cm^2	C / μ F/cm^2	C / μ F/cm^2	C / μ F/cm^2
0.200	40.5	-	-	-	-	-
0.150	33.8	-	-	-	-	-
0.100	31.5	-	-	-	-	-
0.050	30.2	-	-	-	-	-
0.000	28.8	32.8	42.7	-	-	-
-0.050	28.2	32.3	35.6	49.4	-	-
-0.100	28.2	32.1	34.1	40.5	55.3	-
-0.150	28.5	32.4	34.8	39.0	44.1	70.2
-0.200	28.5	33.2	36.2	39.0	42.8	45.7
-0.250	28.6	33.6	38.0	38.8	43.1	44.1
-0.300	28.6	33.6	38.6	38.1	43.1	41.5
-0.350	28.0	33.0	38.6	36.8	42.1	38.3
-0.400	27.3	31.6	37.2	34.5	39.4	34.6
-0.450	26.1	29.8	34.7	31.8	35.8	31.3
-0.500	24.9	28.0	31.6	29.4	32.2	28.4
-0.550	23.5	25.8	28.9	27.2	28.8	25.6
-0.600	22.2	24.0	26.3	25.1	26.1	23.6
-0.650	20.8	22.2	24.1	23.2	24.0	21.9
-0.700	20.1	21.0	22.2	21.7	22.3	20.4
-0.750	18.8	19.8	21.0	20.7	20.9	19.5
-0.800	17.9	18.8	19.8	19.7	20.0	18.8
-0.850	17.2	18.2	19.0	19.0	19.1	18.3
-0.900	16.8	17.6	18.4	18.6	18.6	18.0
-0.950	16.4	17.4	17.9	18.1	18.3	17.6
-1.000	16.3	17.2	17.6	17.7	18.1	17.6
-1.050	16.2	17.1	17.4	17.7	18.0	17.4
-1.100	16.2	17.2	17.4	17.7	17.9	17.5
-1.200	16.4	17.3	17.5	17.8	18.3	18.1
-1.300	17.0	17.8	18.0	18.2	18.8	18.6
-1.400	17.7	18.4	18.7	19.0	19.6	19.4
-1.500	18.9	19.4	19.6	19.7	20.7	20.2
-1.600	20.0	20.6	20.8	20.9	22.0	22.0
-1.700	21.6	22.6	22.7	22.6	23.8	-
-1.800	23.7	25.0	25.5	-	-	-

Differential double layer capacity in aqueous KNO_3 melts of various compositions and temperatures. Potentials relative to Ag/0.15 M $AgNO_3$ electrode in $LiNO_3 . 3H_2O$.
Reference: V.S. Palankev, A.M. Skundin, V.S. Bagotskii,
Elektrokhimiya, 2, (1966), 640.

electrolyte	$KNO_3 . 6\ H_2O$		$KNO_3 . 3\ H_2O$
temperature	60°C	95°C	95°C
E / volts	C / $\mu F/cm^2$	C / $\mu F/cm^2$	C / $\mu F/cm^2$
-0.200	51.5	47.0	48.4
-0.250	46.4	43.5	45.0
-0.300	42.5	40.6	42.5
-0.350	39.4	38.2	39.3
-0.400	35.8	35.1	36.4
-0.450	33.0	32.4	33.8
-0.500	30.8	31.0	32.2
-0.550	29.4	29.9	30.7
-0.600	28.2	28.9	29.5
-0.650	27.4	28.0	28.5
-0.700	27.0	27.9	27.8
-0.750	27.0	27.9	27.3
-0.800	27.2	27.7	27.2
-0.850	27.4	27.4	26.8
-0.900	27.7	27.2	26.8
-0.950	28.2	26.9	26.2
-1.000	28.2	26.4	26.0
-1.050	28.2	25.6	25.8
-1.100	27.4	25.0	25.0
-1.150	26.7	24.0	24.3
-1.200	25.6	23.0	23.5
-1.250	24.4	22.0	23.0
-1.300	23.2	21.2	22.1
-1.350	22.8	20.4	-
-1.400	21.2	19.9	20.8
-1.450	-	19.4	-
-1.500	19.7	19.1	20.1
-1.600	18.8	18.7	19.5
-1.700	18.6	18.8	19.4
-1.800	18.7	-	-

Differential double layer capacity in aqueous $NaNO_3$ melts of various compositions and temperatures. Potentials relative to Ag/0.15 M $AgNO_3$ electrode in $LiNO_3 . 3\ H_2O$.
Reference: V.S. Palankev, A.M. Skundin, V.S. Bagotskii,
Elektrokhimiya, 2, (1966), 640.

electrolyte	$NaNO_3 . 6\ H_2O$			$NaNO_3 . 3\ H_2O$
temperature	20°C	60°C	95°C	95°C
E / volts	C / μF/cm^2	C / μF/cm^2	C / μF/cm^2	C / μF/cm^2
-0.200	48.2	45.5	45.1	45.3
-0.250	43.4	42.0	41.9	42.2
-0.300	39.8	38.0	39.0	39.7
-0.350	36.2	35.5	36.2	36.9
-0.400	33.5	33.1	33.6	34.7
-0.450	31.3	31.2	32.0	32.9
-0.500	29.5	30.6	30.4	31.0
-0.550	27.6	28.1	29.0	29.8
-0.600	26.4	27.2	28.2	28.8
-0.650	25.7	26.6	27.6	28.0
-0.700	25.4	26.4	27.3	27.5
-0.750	25.1	26.6	27.0	26.8
-0.800	25.4	26.6	27.0	26.6
-0.850	25.8	27.1	27.1	26.6
-0.900	26.8	27.6	26.8	26.5
-0.950	27.7	27.6	26.6	26.2
-1.000	28.6	27.9	26.6	25.9
-1.050	29.6	27.6	26.0	25.5
-1.100	30.0	27.4	25.4	25.4
-1.150	29.8	26.6	24.6	25.0
-1.200	29.2	25.6	23.8	24.3
-1.250	28.4	24.7	22.9	23.8
-1.300	27.0	23.7	22.1	23.2
-1.350	-	22.6	-	-
-1.400	24.1	21.9	20.7	22.0
-1.450	-	-	-	-
-1.500	22.0	20.4	20.0	21.2
-1.600	20.4	19.7	19.2	20.7
-1.700	19.8	19.2	19.2	-
-1.800	19.4	-	-	-

Differential double layer capacity in aqueous $LiNO_3$ melts of varying compositions. Potentials relative to Ag/0.15 M $AgNO_3$ in test electrolyte.

Reference: V.S. Palankev, A.M. Skundin and V.S. Bagotskii,
Elektrokhimiya, 2, (1966), 640.

electrolyte	$LiNO_3 \cdot 1,5\ H_2O$	$LiNO_3 \cdot 2\ H_2O$
temperature	130°C	130°C
E / volts	C / $\mu F/cm^2$	C / $\mu F/cm^2$
-0.050	-	-
-0.100	-	-
-0.150	-	-
-0.200	44.8	46.7
-0.250	40.5	40.6
-0.300	35.5	35.5
-0.350	33.4	32.7
-0.400	31.2	30.6
-0.450	29.8	28.8
-0.500	28.2	27.8
-0.550	27.2	26.6
-0.600	26.1	25.5
-0.650	25.4	24.8
-0.700	24.5	24.1
-0.750	23.8	23.9
-0.800	23.5	23.6
-0.850	23.3	23.6
-0.900	23.0	23.6
-0.950	23.0	23.6
-1.000	23.0	23.6
-1.050	23.0	23.4
-1.100	23.0	23.3
-1.150	23.0	23.1
-1.200	22.9	23.0
-1.250	22.8	23.0
-1.300	22.8	23.1
-1.350	22.8	-
-1.400	22.5	-
-1.450	-	-
-1.500	-	-

Double layer in aqueous $LiNO_3$ melts (cont.)

electrolyte	$LiNO_3 \cdot 2\,H_2O$			$LiNO_3 \cdot 3\,H_2O$			
temperature	80°C	95°C	115°C	40°C	60°C	80°C	95°C
E volts	C $\mu F/cm^2$	C $\mu F/cm^2$	C $\mu F/cm^2$	C $\mu F/cm^2$	C $\mu F/cm^2$	C $\mu F/cm^2$	C $\mu F/cm^2$
-0.200	42.3	-	42.6	-	-	-	-
-0.250	-	41.3	38.5	43.2	-	-	-
-0.300	36.6	37.7	35.0	38.9	38.4	38.8	36.8
-0.350	-	34.2	32.6	34.5	-	-	-
-0.400	31.6	31.9	30.9	31.5	31.6	33.1	31.0
-0.450	-	29.8	29.1	28.8	-	-	-
-0.500	27.5	28.4	27.4	26.8	27.3	27.3	26.6
-0.550	25.9	26.9	26.2	25.2	-	-	25.6
-0.600	24.6	25.8	25.5	23.8	24.2	25.2	24.5
-0.650	23.6	24.8	24.6	22.9	23.6	24.9	24.1
-0.700	22.9	24.0	24.0	22.2	23.1	24.0	23.7
-0.750	22.5	23.7	23.8	21.9	22.7	23.5	23.6
-0.800	22.4	23.4	23.4	22.2	22.7	23.6	23.6
-0.850	22.3	23.4	23.4	22.5	23.2	23.8	23.8
-0.900	22.5	23.4	23.4	22.8	23.3	24.1	23.9
-0.950	22.9	23.6	23.4	23.8	24.0	24.5	24.2
-1.000	23.2	24.0	23.7	24.6	24.5	25.0	24.4
-1.050	23.5	24.2	23.7	25.3	25.0	25.3	24.4
-1.100	23.9	24.4	23.7	26.4	25.5	25.8	24.6
-1.150	24.2	24.7	23.7	26.8	25.8	25.7	24.6
-1.200	24.3	24.5	23.7	27.0	26.0	25.4	24.3
-1.250	24.3	24.4	23.6	27.2	25.8	25.3	-
-1.300	24.3	24.4	23.5	27.2	25.6	25.0	24.2
-1.350	24.2	24.2	23.2	26.8	-	25.0	-
-1.400	24.0	24.2	23.1	26.6	25.1	24.6	23.4
-1.450	23.8	24.1	22.9	-	-	-	-
-1.500	23.7	23.8	-	25.9	24.5	24.2	23.0
-1.550	23.8	23.8	-	-	-	-	-
-1.600	-	-	-	25.2	24.0	23.3	-
-1.650	-	-	-	24.6	-	-	-

Differential double layer capacity in aqueous $LiNO_3$ melts of various compositions and temperatures. Potentials relative to Ag/0.15 M $AgNO_3$ electrode in $LiNO_3 \cdot 3\ H_2O$.
Reference: V.S. Palankev, A.M. Skundin, V.S. Bagotskii,
Elektrokhimiya, 2, (1966), 640.

electrolyte $LiNO_3 \cdot 6\ H_2O$

temperature	20°C	40°C	60°C	80°C	95°C
E / Volts	C / μ F/cm^2	C / μ F/cm^2	C / μ F/cm^2	C / μ F/cm^2	C / μ F/cm^2
-0.300	36.6	38.0	39.8	39.5	39.4
-0.350	32.6	-	34.0	33.6	-
-0.400	29.4	30.0	30.8	30.2	30.2
-0.450	26.9	27.5	28.4	28.5	-
-0.500	25.0	25.6	26.5	26.6	27.0
-0.550	23.6	24.4	25.5	25.4	-
-0.600	22.4	23.4	24.8	24.8	25.3
-0.650	21.9	23.2	24.4	24.3	25.0
-0.700	21.6	22.9	23.8	24.4	24.7
-0.750	21.6	22.9	23.9	24.4	24.8
-0.800	22.2	23.4	24.3	24.7	24.9
-0.850	23.2	24.0	25.1	25.4	25.2
-0.900	24.2	25.0	25.7	-	25.5
-0.950	25.6	26.0	26.2	25.8	25.8
-1.000	26.9	27.0	26.9	26.0	25.6
-1.050	28.0	27.8	27.8	26.2	25.5
-1.100	28.6	28.3	27.4	26.0	25.2
-1.150	29.4	28.4	27.0	25.6	24.8
-1.200	29.2	28.6	26.7	25.0	24.1
-1.250	29.1	27.5	26.2	24.4	-
-1.300	28.7	26.8	25.5	23.9	23.0
-1.350	27.9	-	-	-	-
-1.400	26.2	25.1	23.8	22.7	21.8
-1.450	-	-	-	-	-
-1.500	24.4	23.6	22.4	21.6	21.0
-1.550	-	-	-	-	-
-1.600	22.8	22.2	21.4	-	-
-1.650	22.1	-	-	-	-
-1.700	21.6	21.2	20.7	-	-
-1.750	21.1	-	-	-	-
-1.800	-	-	-	-	-

cont.

Differential double layer capacity on mercury in aqueous solutions and aqueous melts of some nitrate mixtures. Potentials relative to Ag/0.15 M $AgNO_3$ electrode in test electrolyte.
Reference: V.S. Palankev, A.M. Skundin and V.S. Bagotskii.
Elektrokhimiya, 2,(1966), 640.

electrolyte	$LiNO_3$-KNO_3 39-61 mole %		$LiNO_3$ 34.5 mole % $NaNO_3$ 18.5 mole % KNO_3 47.0 mole %	$LiNO_3$ - $NaNO_3$ - KNO_3 eutectic	
				+ 0.1 mole H_2O/ mole of nitrates	+1.5 mole H_2O/ mole of nitrates
temperature	260°C	160°C	130°C	130°C	130°C
E volts	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2
-0.050	48.7	-	-	-	-
-0.100	43.8	46.2	49.3	-	-
-0.150	40.5	43.2	-	-	-
-0.200	37.7	40.7	42.7	44.5	47.7
-0.250	34.2	38.5	40.5	42.4	44.3
-0.300	32.5	36.1	38.0	39.9	41.3
-0.350	30.9	33.8	36.0	36.5	37.3
-0.400	29.2	31.8	34.0	35.4	34.5
-0.450	27.7	30.2	31.8	33.2	32.4
-0.500	26.2	28.4	29.8	31.3	30.6
-0.550	25.1	26.7	28.2	29.6	29.2
-0.600	24.0	25.4	26.6	28.0	28.0
-0.650	23.0	24.1	25.1	26.6	27.2
-0.700	22.2	23.2	23.8	25.2	26.4
-0.750	21.3	22.0	22.5	24.1	25.8
-0.800	21.0	21.5	21.7	23.0	25.6
-0.850	20.2	20.8	20.7	22.2	25.2
-0.900	19.7	20.5	20.0	21.6	24.7
-0.950	19.6	20.0	19.5	21.1	24.6
-1.000	19.2	19.9	19.2	20.8	24.4
-1.050	19.0	19.7	18.8	20.4	24.0
-1.100	18.8	19.8	18.8	20.4	23.8
-1.150	18.8	19.8	18.8	20.3	23.5
-1.200	19.8	20.0	18.8	20.4	23.2
-1.250	-	20.3	19.0	20.6	22.9
-1.300	-	21.0	19.4	21.0	22.8
-1.350	-	-	-	21.2	22.5
-1.400	-	-	-	21.5	22.2
-1.450	-	-	-	-	22.2
-1.500	-	-	-	-	22.0

Differential double layer capacity in aqueous LiCl melts of various compositions and temperatures. Potentials relative to Ag/0.15 M $AgNO_3$ electrode in $LiNO_3$. 3 H_2O.
Reference: V.S. Palankev, A.M. Skundin and V.S. Bagotskii, Elektrokhimiya, 3, 370 (1967).

electrolyte	1 mole.l^{-1}	4 mole.l^{-1}	LiCl.6 H_2O melt		LiCl. 3 H_2O melt			LiCl.2H_2O melt
temperature	20°C	20°C	20°C	20°C	60°C	95°C	130°C	130°C
E volts	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2	C μ F/cm^2
-0.100	62.5	-	-	-	-	-	-	-
-0.150	47.6	-	-	-	-	-	-	-
-0.200	40.4	41.5	44.5	42.0	49.5	-	-	-
-0.250	37.5	37.5	39.1	38.9	45.2	50.8	-	-
-0.300	37.3	35.6	35.7	36.8	42.6	46.7	51.8	49.0
-0.350	38.7	35.6	34.2	35.7	40.8	44.5	49.0	46.7
-0.400	40.8	37.2	34.0	35.4	40.0	43.0	46.8	44.8
-0.450	43.0	40.0	34.9	36.2	39.5	41.6	44.2	43.3
-0.500	43.3	42.5	36.4	37.0	39.1	40.4	41.8	41.8
-0.550	40.5	43.5	38.4	37.6	38.8	38.6	39.7	39.9
-0.600	35.6	42.2	39.6	38.2	38.1	37.0	37.7	38.0
-0.650	30.7	38.5	39.5	38.3	37.0	35.4	35.5	36.1
-0.700	26.0	34.5	37.7	37.6	35.8	33.6	33.4	34.1
-0.750	22.7	30.0	35.3	36.3	34.1	31.7	31.4	32.2
-0.800	20.4	26.3	32.4	34.6	32.4	30.0	29.7	30.5
-0.850	18.6	23.3	29.6	32.9	30.8	28.5	28.0	29.3
-0.900	17.6	21.3	27.2	31.2	29.2	27.1	26.7	28.0
-0.950	17.0	19.8	25.1	29.5	27.7	26.0	25.6	27.0
-1.000	16.3	18.7	23.4	27.9	26.7	25.0	24.8	26.0
-1.050	16.0	18.0	22.2	26.6	25.3	24.0	23.9	25.2
-1.100	15.8	17.6	21.3	25.4	24.6	23.3	23.3	24.7
-1.200	15.7	17.4	20.3	23.9	23.5	22.5	22.7	23.9
-1.300	16.0	17.5	20.0	23.2	22.9	22.3	22.4	23.6
-1.400	16.5	17.9	20.2	23.0	22.9	-	22.7	23.7
-1.500	17.3	18.6	-	23.3	-	-	23.4	24.9
-1.600	18.0	-	-	24.4	-	-	-	-

Differential capacity of mercury in aqueous solutions of
x mol l^{-1} CsCl + (1-x) mol l^{-1} LiCl. T = 25°C

Potential with respect to a saturated calomel electrode in contact with the working solution.

Reference: R. Parsons and A. Stockton, J. Electroanal. Chem. 25 (1970) App. 10.

	$C/\mu F\ cm^{-2}$					
x E/volts	0.50	0.20	0.10	0.04	0.02	0.00
-0.12	52.30	49.90	48.25	47.65	47.50	47.55
-0.15	-	-	-	-	-	43.00
-0.18	42.30	41.20	40.65	40.45	40.35	40.30
-0.21	40.05	39.40	39.00	38.90	38.85	38.90
-0.24	39.05	38.75	38.40	38.30	38.30	38.25
-0.27	38.90	38.65	38.50	38.50	38.45	38.50
-0.30	39.40	39.35	39.25	39.20	39.20	39.25
-0.33	40.35	40.30	40.30	40.30	40.35	40.40
-0.36	41.50	41.60	41.60	41.65	41.60	41.65
-0.39	42.55	42.70	42.65	42.70	42.70	42.80
-0.42	43.15	43.25	43.20	43.30	43.30	43.30
-0.45	43.05	43.10	42.95	43.10	43.10	43.10
-0.48	42.10	41.90	42.00	42.05	42.05	41.90
-0.51	40.15	39.85	39.70	39.90	39.90	39.85
-0.54	37.55	37.20	37.15	37.20	37.20	37.15
-0.57	34.65	34.20	34.10	34.20	34.20	34.10
-0.60	31.75	31.55	31.20	31.20	31.20	31.20
-0.63	29.10	28.60	28.45	28.55	28.40	28.35
-0.66	26.80	26.15	25.90	25.90	25.90	25.85
-0.69	24.65	24.05	23.90	23.80	23.85	23.80
-0.72	22.95	22.40	22.20	22.10	22.10	22.05
-0.75	21.70	21.05	20.90	20.75	20.75	20.70
-0.78	20.65	19.95	19.80	19.70	19.65	19.60
-0.81	19.75	19.10	18.90	18.75	18.75	18.70
-0.84	19.05	18.45	18.25	18.05	18.00	18.00
-0.87	18.55	17.90	17.70	17.45	17.50	17.40
-0.90	18.15	17.55	17.25	17.05	16.95	16.90
-0.93	17.90	17.25	17.00	16.75	16.60	16.55

Capacity of mercury in x mol l^{-1} CsCl + (1-x) mol l^{-1}LiCl (cont.)

			C/μF cm^{-2}			
x	0.50	0.20	0.10	0.04	0.02	0.00
E/volts						
-0.96	17.65	17.00	16.70	16.45	16.35	16.30
-0.99	17.50	16.85	16.55	16.30	16.15	16.10
-1.02	17.45	16.80	16.45	16.15	16.00	15.95
-1.05	17.35	16.80	16.40	16.05	15.90	15.80
-1.08	17.35	16.80	16.40	16.00	15.85	15.70
-1.11	17.45	16.85	16.45	16.05	15.85	15.65
-1.14	17.45	16.95	16.55	16.05	15.90	15.70
-1.17	17.60	17.00	16.60	16.15	16.00	15.70
-1.20	17.65	17.15	16.75	16.30	16.05	15.80
-1.23	17.80	17.30	16.90	16.45	16.15	15.85
-1.26	17.95	17.55	17.05	16.60	16.30	15.95
-1.29	18.20	17.75	17.25	16.75	16.45	16.10
-1.32	18.40	18.00	17.50	16.95	16.70	16.25
-1.35	18.60	18.20	17.80	17.20	16.80	16.40
-1.38	18.85	18.50	18.00	17.45	17.05	16.55
-1.41	19.05	18.75	18.30	17.75	17.35	16.70
-1.44	19.30	19.05	18.70	18.05	17.65	16.90
-1.47	19.60	19.40	19.00	18.35	17.95	17.15
-1.50	19.95	19.80	19.35	18.65	18.25	17.35
-1.53	20.25	20.15	19.75	19.05	18.50	17.55
-1.56	20.65	20.45	20.10	-	-	17.80
-1.59	-	20.90	-	-	-	18.10

Capacity of mercury in x mol l^{-1} CsCl + (1-x) mol l^{-1} LiCl (cont.)

	x = 1.00		
E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-0.14	50.70	-0.75	22.10
-0.18	44.00	-0.78	21.10
-0.22	40.75	-0.81	20.20
-0.24	39.90	-0.84	19.50
-0.26	39.45	-0.87	19.00
-0.28	39.40	-0.90	18.55
-0.30	39.60	-0.93	18.25
-0.32	40.00	-0.96	18.00
-0.34	40.75	-0.99	17.80
-0.36	41.40	-1.02	17.70
-0.38	42.15	-1.05	17.65
-0.40	42.65	-1.08	17.60
-0.42	43.05	-1.11	17.65
-0.44	43.05	-1.14	17.70
-0.46	42.75	-1.17	17.80
-0.48	41.95	-1.20	17.85
-0.50	40.90	-1.23	18.00
-0.52	39.30	-1.26	18.15
-0.54	37.55	-1.29	18.30
-0.56	35.65	-1.32	18.45
-0.58	33.75	-1.35	18.70
-0.60	31.95	-1.41	19.15
-0.62	30.30	-1.47	19.65
-0.66	27.40	-1.50	19.95
-0.69	24.95	-1.53	20.25
-0.72	23.25	-1.56	20.65

Differential capacity of mercury in aqueous solutions containing KCl and $MgCl_2$ T = 25°C

Potential with respect to Ag/AgCl electrode in the working solution.
Frequency 1.5 kHz

Reference: R. Parsons and S. Trasatti. Trans. Faraday Soc. 65 (1969) 3314

	C/μF cm^{-2}						
c_{KCl}/mol l^{-1}	0.0204	0.0205	0.0206	0.0208	0.0210	0.0211	0.0214
c_{MgCl_2}/mol l^{-1}	0.0025	0.0032	0.004	0.00555	0.0065	0.0075	0.0100
E/volts							
-1.60	17.96	18.03	18.11	18.14	18.19	18.27	18.31
-1.55	17.46	17.51	17.61	17.64	17.69	17.75	17.80
-1.50	17.00	17.06	17.13	17.16	17.19	17.24	17.30
-1.45	16.62	16.66	16.74	16.77	16.79	16.83	16.88
-1.40	16.30	16.33	16.40	16.41	16.45	16.49	16.54
-1.35	16.04	16.08	16.13	16.14	16.17	16.20	16.25
-1.30	15.86	15.90	15.92	15.94	15.95	15.99	16.03
-1.25	15.75	15.77	15.80	15.81	15.82	15.85	15.88
-1.20	15.72	15.74	15.76	15.75	15.77	15.49	15.81
-1.15	15.79	15.80	15.82	15.81	15.82	15.83	15.83
-1.10	15.95	15.94	15.95	15.94	15.94	15.94	15.94
-1.05	16.23	16.22	16.21	16.19	16.18	16.19	16.16
-1.00	16.62	16.59	16.58	16.57	16.56	16.55	16.51
-0.95	17.13	17.10	17.08	17.06	17.05	17.03	16.99
-0.90	17.75	17.73	17.71	17.68	17.66	17.63	17.59
-0.85	18.48	18.45	18.44	18.41	18.40	18.39	18.34
-0.80	19.26	19.26	19.25	19.24	19.23	19.23	19.18
-0.75	19.95	20.00	20.04	20.06	20.08	20.10	20.08
-0.70	20.44	20.54	20.64	20.74	20.81	20.87	20.94
-0.65	20.44	20.66	20.86	21.10	21.28	21.43	21.64
-0.60	20.12	20.49	20.81	21.25	21.55	21.83	22.28
-0.57	20.34	20.76	21.15	21.65	22.03	22.32	22.91
-0.55	21.10	21.53	21.87	22.38	22.79	22.15	23.78
-0.50	25.41	25.78	25.93	26.49	26.69	27.03	27.57
-0.45	31.46	31.60	31.74	31.98	32.20	32.52	32.85
-0.40	35.73	35.74	35.76	35.98	36.11	36.42	36.57
-0.35	37.12	37.15	37.24	37.40	37.50	37.68	37.81
-0.30	37.01	37.04	37.13	37.25	37.32	37.42	37.55
-0.25	36.77	36.75	36.82	36.90	36.96	37.08	37.19
-0.20	37.12	37.22	37.42	37.51	37.57	37.66	37.84
-0.15	39.35	39.43	39.76	39.84	39.94	40.05	40.27
-0.10	44.61	44.78	45.20	45.29	45.46	45.65	45.81
-0.05	57.01	57.43	58.37	58.60	59.02	59.33	59.90

Capacity in aqueous KCl and $MgCl_2$ (cont.)

				$C/\mu F\ cm^{-2}$			
c_{KCl}/mol l^{-1}	0.01	0.013	0.016	0.0208	0.025	0.03	0.04
c_{MgCl_2}/mol l^{-1}	0.0053	0.00538	0.00545	0.00555	0.00564	0.00573	0.0059
E/volts							
-1.60	17.93	17.97	18.00	18.14	18.22	18.25	18.41
-1.55	17.42	17.49	17.52	17.64	17.72	17.75	17.87
-1.50	16.97	17.03	17.05	17.16	17.22	17.26	17.38
-1.45	16.56	16.65	16.69	16.78	16.82	16.85	16.98
-1.40	16.27	16.32	16.35	16.41	16.48	16.51	16.60
-1.35	16.00	16.06	16.09	16.14	16.20	16.23	16.30
-1.30	15.82	15.88	15.88	15.94	15.98	16.00	16.07
-1.25	15.71	15.76	15.76	15.81	15.85	15.88	15.91
-1.20	15.69	15.72	15.71	15.75	15.79	15.80	15.84
-1.15	15.76	15.78	15.77	15.81	15.83	15.83	15.85
-1.10	15.94	15.94	15.93	15.94	15.96	15.95	15.96
-1.05	16.24	16.23	16.19	16.19	16.19	16.18	16.16
-1.00	16.63	16.61	16.57	16.57	16.57	16.54	16.48
-0.95	17.16	17.14	17.09	17.06	17.04	17.00	16.93
-0.90	17.81	17.78	17.72	17.68	17.66	17.60	17.51
-0.85	18.56	18.53	18.45	18.41	18.39	18.33	18.21
-0.80	19.39	19.34	19.28	19.24	19.19	19.12	19.00
-0.75	20.23	20.18	20.10	20.06	20.03	19.95	19.83
-0.70	20.88	20.84	20.75	20.74	20.73	20.71	20.62
-0.65	21.04	21.08	21.03	21.10	21.19	21.24	21.29
-0.60	20.62	20.85	20.97	21.25	21.53	21.75	22.13
-0.57	20.53	20.94	21.22	21.65	22.05	22.46	23.10
-0.55	20.99	21.48	21.86	22.38	22.89	23.38	24.12
-0.50	24.79	25.24	25.90	26.49	26.99	27.47	28.29
-0.45	30.92	31.17	31.62	31.98	32.46	32.85	33.53
-0.40	35.10	35.44	35.77	35.98	36.35	36.62	37.06
-0.35	36.85	36.96	37.29	37.40	37.64	37.83	38.12
-0.30	36.74	36.85	37.06	37.25	37.37	37.58	37.80
-0.25	36.53	36.64	36.73	36.90	36.96	37.16	37.35
-0.20	37.06	37.25	37.19	37.51	37.57	37.69	37.93
-0.15	39.27	39.41	39.49	39.84	39.95	40.12	40.40
-0.10	44.63	44.74	44.82	45.29	45.38	45.69	46.21
-0.05	57.40	57.58	57.88	58.60	59.04	59.31	60.24

Differential Capacity of mercury in aqueous x mol l^{-1} NH_4Cl + (1-x) mol l^{-1} NH_4NO_3

T = 25°C

Potential with respect to 1 mol l^{-1} KCl calomel electrode

Reference: R. Payne, unpublished

	0		0.0020		0.0050		0.010	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-17	-1.365	18.32	-1.365	18.32	-1.365	18.32	-1.365	18.32
-16	-1.309	17.85	-1.309	17.87	-1.309	17.84	-1.309	17.85
-15	-1.253	17.49	-1.253	17.52	-1.253	17.49	-1.253	17.51
-14	-1.195	17.27	-1.195	17.28	-1.195	17.28	-1.195	17.30
-13	-1.137	17.19	-1.137	17.20	-1.137	17.21	-1.137	17.22
-12	-1.079	17.31	-1.079	17.34	-1.079	17.35	-1.079	17.36
-11	-1.022	17.70	-1.022	17.71	-1.022	17.74	-1.022	17.77
-10	-0.966	18.40	-0.967	18.42	-0.967	18.44	-0.967	18.49
-9	-0.913	19.44	-0.914	19.47	-0.914	19.48	-0.914	19.54
-8	-0.864	20.85	-0.864	20.86	-0.864	20.87	-0.865	20.95
-7	-0.818	22.52	-0.818	22.53	-0.818	22.56	-0.819	22.63
-6	-0.775	24.36	-0.775	24.36	-0.775	24.38	-0.777	24.47
-5	-0.735	26.20	-0.736	26.20	-0.736	26.21	-0.737	26.32
-4	-0.698	27.90	-0.699	27.90	-0.699	27.93	-0.700	28.02
-3	-0.663	29.28	-0.664	29.30	-0.664	29.35	-0.666	29.47
-2	-0.630	30.33	-0.630	30.38	-0.631	30.43	-0.632	30.54
-1	-0.597	30.99	-0.598	31.05	-0.598	31.10	-0.600	31.20
0	-0.565	31.18	-0.566	31.27	-0.566	31.31	-0.568	31.42
1	-0.533	30.92	-0.534	31.05	-0.534	31.08	-0.536	31.20
2	-0.500	30.31	-0.501	30.46	-0.502	30.50	-0.504	30.59
3	-0.466	29.40	-0.468	29.58	-0.468	29.61	-0.471	29.72
4	-0.432	28.30	-0.433	28.52	-0.434	28.54	-0.436	28.64
5	-0.396	27.09	-0.398	27.33	-0.398	27.35	-0.401	27.45
6	-0.358	25.95	-0.360	26.17	-0.361	26.21	-0.363	26.33
7	-0.319	24.92	-0.321	25.13	-0.322	25.18	-0.325	25.33
8	-0.278	24.13	-0.281	24.36	-0.282	24.42	-0.285	24.59
9	-0.236	23.62	-0.239	23.85	-0.240	23.94	-0.244	24.18
10	-0.194	23.43	-0.197	23.66	-0.198	23.83	-0.202	24.16
11	-0.151	23.57	-0.155	23.82	-0.156	24.10	-0.161	24.57
12	-0.108	24.00	-0.113	24.31	-0.116	24.78	-0.121	25.49
13	-0.068	24.74	-0.073	25.20	-0.076	25.95	-0.083	26.98
14	-0.028	25.75	-0.034	26.44	-0.039	27.64	-0.047	29.19
15	+0.010	27.01	+0.003	28.18	-0.004	29.97	-0.014	32.32
16	0.046	28.54	0.037	30.48			+0.015	36.84
17	0.080	30.20						

Capacity for Hg in aqueous x mol l^{-1} NH_4Cl + (1-x) mol l^{-1} NH_4NO_3 (cont.)

	0.020		0.050		0.10		0.90	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu Fcm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-17	-1.365	18.32	-1.365	18.32	-1.362	18.32	-1.365	18.32
-16	-1.309	17.86	-1.309	17.84	-1.309	17.85	-1.309	17.84
-15	-1.253	17.49	-1.253	17.49	-1.253	17.49	-1.253	17.45
-14	-1.195	17.29	-1.195	17.28	-1.195	17.28	-1.195	17.18
-13	-1.137	17.22	-1.137	17.21	-1.137	17.20	-1.136	17.02
-12	-1.079	17.36	-1.079	17.32	-1.079	17.32	-1.078	17.01
-11	-1.022	17.73	-1.022	17.70	-1.022	17.71	-1.019	17.22
-10	-0.967	18.43	-0.966	18.41	-0.966	18.40	-0.962	17.67
-9	-0.914	19.49	-0.914	19.45	-0.914	19.44	-0.906	18.44
-8	-0.864	20.91	-0.864	20.83	-0.864	20.82	-0.853	19.57
-7	-0.818	22.58	-0.818	22.53	-0.818	22.51	-0.804	21.14
-6	-0.776	24.40	-0.775	24.38	-0.775	24.34	-0.759	23.12
-5	-0.736	26.26	-0.735	26.25	-0.735	26.23	-0.718	25.50
-4	-0.699	27.97	-0.698	27.98	-0.698	28.00	-0.680	28.11
-3	-0.665	29.39	-0.664	29.41	-0.664	29.48	-0.646	30.85
-2	-0.631	30.46	-0.630	30.51	-0.630	30.66	-0.615	33.54
-1	-0.599	31.11	-0.598	31.19	-0.598	31.46	-0.587	35.94
0	-0.567	31.35	-0.566	31.45	-0.567	31.81	-0.559	37.99
1	-0.535	31.11	-0.534	31.28	-0.535	31.74	-0.534	39.64
2	-0.502	30.54	-0.502	30.73	-0.503	31.30	-0.509	40.73
3	-0.469	29.66	-0.469	29.92	-0.471	30.63	-0.485	41.32
4	-0.435	28.62	-0.435	28.96	-0.438	29.77	-0.460	41.47
5	-0.399	27.48	-0.400	27.91	-0.404	28.85	-0.436	41.21
6	-0.362	26.41	-0.363	26.96	-0.369	28.04	-0.412	40.66
7	-0.323	25.49	-0.326	26.22	-0.333	27.46	-0.387	40.02
8	-0.284	24.85	-0.287	25.80	-0.296	27.26	-0.362	39.41
9	-0.243	24.60	-0.248	25.84	-0.259	27.55	-0.336	38.98
10	-0.203	24.80	-0.210	26.42	-0.224	28.42	-0.311	38.80
11	-0.165	25.51	-0.173	27.67	-0.189	29.98	-0.285	39.02
12	-0.124	26.81	-0.138	29.67	-0.157	32.33	-0.259	39.72
13	-0.088	28.83	-0.106	32.54	-0.128	35.46	-0.235	41.01
14	-0.055	31.73	-0.077	36.41	-0.101	39.51	-0.211	42.90
15	-0.026	35.81	-0.051	41.45	-0.077	44.48	-0.188	45.64
16			-0.028	47.96	-0.056	51.02		
17					-0.037	58.83		

Differential capacity of a mercury electrode in aqueous solutions of $xM\ NH_4NO_3 + (1-x)M\ NH_4F$

T = 25°C

Potential with respect to 1 mol l^{-1} KCl calomel electrode

Reference: R. Payne, J. Phys. Chem., 69 (1965) 4113

x	0		0.001		0.005		0.01	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-20	-1.522	20.02	-1.522	20.02	-1.522	20.02	-1.522	20.02
-19	-1.471	19.33	-1.471	19.33	-1.471	19.33	-1.471	19.33
-18	-1.418	18.72	-1.418	18.72	-1.418	18.72	-1.418	18.72
-17	-1.364	18.16	-1.364	18.16	-1.364	18.16	-1.364	18.16
-16	-1.308	17.68	-1.308	17.68	-1.308	17.68	-1.308	17.68
-15	-1.251	17.29	-1.251	17.29	-1.251	17.29	-1.251	17.29
-14	-1.192	16.98	-1.192	16.99	-1.192	16.98	-1.192	16.99
-13	-1.133	16.79	-1.133	16.80	-1.133	16.79	-1.133	16.81
-12	-1.073	16.74	-1.074	16.74	-1.073	16.74	-1.074	16.75
-11	-1.014	16.82	-1.014	16.82	-1.014	16.80	-1.014	16.84
-10	-0.955	17.08	-0.955	17.09	-0.955	17.07	-0.955	17.11
-9	-0.897	17.53	-0.897	17.54	-0.897	17.52	-0.897	17.56
-8	-0.841	18.16	-0.841	18.16	-0.841	18.16	-0.841	18.22
-7	-0.787	18.95	-0.787	18.97	-0.787	18.96	-0.788	19.05
-6	-0.735	19.89	-0.736	19.90	-0.735	19.93	-0.736	20.03
-5	-0.686	20.94	-0.687	20.94	-0.686	21.02	-0.688	21.15
-4	-0.640	22.05	-0.640	22.07	-0.640	22.18	-0.642	22.36
-3	-0.596	23.16	-0.596	23.22	-0.596	23.39	-0.598	23.64
-2	-0.553	24.26	-0.554	24.32	-0.554	24.57	-0.557	24.91
-1	-0.513	25.25	-0.513	25.34	-0.515	25.72	-0.518	26.14
0	-0.474	26.13	-0.475	26.27	-0.477	26.74	-0.480	27.27
1	-0.436	26.90	-0.437	27.04	-0.440	27.59	-0.444	28.20
2	-0.400	27.48	-0.401	27.68	-0.404	28.25	-0.409	28.88
3	-0.363	27.92	-0.365	28.11	-0.369	28.66	-0.375	29.25
4	-0.328	28.20	-0.330	28.39	-0.334	28.83	-0.341	29.30
5	-0.292	28.20	-0.294	28.51	-0.299	28.75	-0.307	29.07
6	-0.257	28.46	-0.259	28.50	-0.264	28.51	-0.272	28.62
7	-0.222	28.50	-0.224	28.45	-0.229	28.18	-0.237	28.06
8	-0.187	28.53	-0.189	28.40	-0.194	27.86	-0.201	27.52
9	-0.152	28.64	-0.154	28.42	-0.157	27.65	-0.164	27.12
10	-0.117	28.83	-0.119	28.53	-0.121	27.63	-0.127	26.95
11	-0.083	29.11	-0.084	28.79	-0.085	27.78	-0.090	27.04
12	-0.049	29.54	-0.049	29.22	-0.049	28.22	-0.053	24.42
13	-0.015	30.07						

Capacity in aqueous xM NH_4NO_3 + (1-x) M NH_4F (cont.)

	0.02		0.05		0.1		0.2	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20	-1.522	20.02	-1.522	20.02	-1.522	20.02	-1.522	20.02
-19	-1.471	19.33	-1.471	19.33	-1.471	19.33	-1.471	19.33
-18	-1.418	18.72	-1.418	18.72	-1.418	18.72	-1.418	18.72
-17	-1.364	18.17	-1.364	18.18	-1.364	18.17	-1.364	18.19
-16	-1.308	17.72	-1.308	17.70	-1.308	17.69	-1.308	17.71
-15	-1.251	17.31	-1.251	17.30	-1.251	17.31	-1.251	17.33
-14	-1.192	17.00	-1.193	17.00	-1.193	17.01	-1.193	17.05
-13	-1.134	16.82	-1.133	16.82	-1.133	16.84	-1.134	16.89
-12	-1.074	16.75	-1.074	16.76	-1.074	16.82	-1.075	16.87
-11	-1.014	16.84	-1.014	16.86	-1.014	16.95	-1.015	17.02
-10	-0.955	17.11	-0.955	17.15	-0.956	17.28	-0.957	17.41
-9	-0.898	17.56	-0.898	17.65	-0.899	17.82	-0.901	18.05
-8	-0.842	18.22	-0.842	18.38	-0.844	18.64	-0.847	19.00
-7	-0.788	19.08	-0.789	19.33	-0.792	19.71	-0.796	20.24
-6	-0.737	20.10	-0.739	20.48	-0.743	21.03	-0.748	21.79
-5	-0.688	21.28	-0.692	21.84	-0.697	22.62	-0.704	23.56
-4	-0.643	22.59	-0.647	23.38	-0.654	24.34	-0.663	25.45
-3	-0.600	23.98	-0.606	25.01	-0.615	26.10	-0.625	27.31
-2	-0.559	25.39	-0.567	26.63	-0.578	27.81	-0.590	28.96
-1	-0.521	26.76	-0.531	28.10	-0.543	29.29	-0.556	30.22
0	-0.484	27.96	-0.496	29.33	-0.509	30.39	-0.523	31.06
1	-0.449	28.93	-0.462	30.17	-0.477	31.02	-0.491	31.40
2	-0.415	29.55	-0.429	30.57	-0.444	31.17	-0.459	31.24
3	-0.381	29.78	-0.397	30.51	-0.412	30.80	-0.427	30.67
4	-0.348	29.66	-0.364	30.02	-0.379	30.08	-0.394	29.75
5	-0.314	29.21	-0.330	29.22	-0.346	29.12	-0.360	28.61
6	-0.279	28.54	-0.295	28.25	-0.311	27.97	-0.324	27.41
7	-0.244	27.79	-0.259	27.25	-0.274	26.85	-0.287	26.25
8	-0.207	27.07	-0.222	26.37	-0.236	25.87	-0.248	25.27
9	-0.170	26.52	-0.183	25.68	-0.197	25.13	-0.208	24.58
10	-0.132	26.23	-0.144	25.27	-0.157	24.71	-0.167	24.20
11	-0.094	26.23	-0.104	25.21	-0.116	24.62	-0.125	24.14
12	-0.056	26.57	-0.065	25.51	-0.076	24.90	-0.084	24.43
13			-0.026	26.26	-0.036	25.52	-0.044	25.07

Differential capacity for mercury in aqueous x mol l^{-1} NH_4ClO_4 (1-x) mol l^{-1} NH_4F. T = 25 °C.

Potential with respect to 1 mol l^{-1} KCl calomel electrode

Reference: R. Payne, J. Phys. Chem., 70 (1966) 204.

x	0		0.0025		0.0050		0.010	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20	-1.522	20.02	-1.522	20.02	-1.522	20.02	-1.522	20.02
-19	-1.471	19.30	-1.471	19.30	-1.471	19.30	-1.471	19.30
-18	-1.418	18.67	-1.418	18.68	-1.418	18.68	-1.418	18.68
-17	-1.364	18.12	-1.364	18.13	-1.364	18.15	-1.364	18.19
-16	-1.308	17.64	-1.308	17.66	-1.308	17.67	-1.308	17.70
-15	-1.250	17.26	-1.250	17.26	-1.250	17.27	-1.251	17.30
-14	-1.192	16.95	-1.192	16.95	-1.192	16.97	-1.192	16.99
-13	-1.132	16.77	-1.133	16.77	-1.133	16.77	-1.133	16.80
-12	-1.073	16.71	-1.073	16.70	-1.073	16.71	-1.073	16.73
-11	-1.013	16.79	-1.013	16.78	-1.013	16.79	-1.014	16.82
-10	-0.954	17.06	-0.954	17.04	-0.954	17.05	-0.955	17.07
-9	-0.896	17.50	-0.896	17.48	-0.896	17.49	-0.897	17.51
-8	-0.840	18.14	-0.840	18.11	-0.840	18.12	-0.841	18.15
-7	-0.786	18.94	-0.785	18.91	-0.786	18.92	-0.787	18.96
-6	-0.734	19.87	-0.734	19.85	-0.734	19.88	-0.735	19.94
-5	-0.685	20.92	-0.685	20.91	-0.685	20.94	-0.686	21.02
-4	-0.638	22.02	-0.638	22.03	-0.639	22.08	-0.640	22.20
-3	-0.594	23.15	-0.594	23.19	-0.594	23.26	-0.596	23.43
-2	-0.552	24.23	-0.552	24.31	-0.553	24.42	-0.555	24.62
-1	-0.511	25.21	-0.511	25.33	-0.512	25.47	-0.515	25.70
0	-0.473	26.10	-0.473	26.23	-0.474	26.38	-0.477	26.66
1	-0.435	26.85	-0.453	26.99	-0.437	27.15	-0.440	27.43
2	-0.398	27.45	-0.398	27.57	-0.400	27.72	-0.404	27.98
3	-0.362	27.86	-0.362	27.96	-0.364	28.07	-0.368	28.28
4	-0.326	28.14	-0.327	28.17	-0.329	28.23	-0.333	28.33
5	-0.291	28.29	-0.291	28.25	-0.293	28.22	-0.298	28.19
6	-0.255	28.37	-0.256	28.23	-0.258	28.11	-0.262	27.97
7	-0.220	28.44	-0.221	28.19	-0.222	27.98	-0.226	27.69
8	-0.185	28.54	-0.185	28.19	-0.186	27.89	-0.190	27.49
9	-0.150	28.73	-0.150	28.28	-0.151	27.93	-0.153	27.47
10	-0.115	28.97	-0.114	28.47	-0.115	28.12	-0.117	27.62
11	-0.081	29.26	-0.080	28.82	-0.079	28.49	-0.081	27.98
12	-0.047	29.80	-0.045	29.30	-0.045	29.04	-0.046	28.57

Capacity of Hg in aqueous x mol l^{-1} NH_4ClO_4 + (1-x) mol l^{-1} NH_4F (cont.)

x	0.020		0.050		0.10		0.20	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-20	-1.522	20.02	-1.522	20.02	-1.522	20.02	-1.522	20.02
-19	-1.471	19.30	-1.471	19.30	-1.471	19.40	-1.471	19.31
-18	-1.418	18.69	-1.418	18.69	-1.418	18.77	-1.418	18.78
-17	-1.364	18.20	-1.364	18.20	-1.364	18.22	-1.364	18.23
-16	-1.308	17.68	-1.308	17.71	-1.309	17.74	-1.308	17.74
-15	-1.251	17.27	-1.251	17.29	-1.252	17.34	-1.251	17.36
-14	-1.192	16.99	-1.192	16.99	-1.193	17.02	-1.193	17.06
-13	-1.133	16.80	-1.133	16.81	-1.134	16.83	-1.134	16.87
-12	-1.073	16.73	-1.074	17.75	-1.075	16.79	-1.075	16.85
-11	-1.014	16.82	-1.014	16.85	-1.015	16.92	-1.016	16.99
-10	-0.955	17.09	-0.955	17.13	-0.957	17.23	-0.957	17.35
-9	-0.897	17.54	-0.897	17.62	-0.899	17.74	-0.901	17.95
-8	-0.841	18.19	-0.842	18.30	-0.844	18.52	-0.846	18.83
-7	-0.787	19.04	-0.788	19.21	-0.792	19.51	-0.795	19.98
-6	-0.736	20.03	-0.738	20.30	-0.742	20.74	-0.746	21.41
-5	-0.687	21.17	-0.690	21.57	-0.695	22.15	-0.701	23.00
-4	-0.641	22.42	-0.645	22.95	-0.651	23.70	-0.659	24.64
-3	-0.598	23.72	-0.603	24.40	-0.611	25.22	-0.620	26.25
-2	-0.557	24.96	-0.563	25.77	-0.572	26.67	-0.583	27.63
-1	-0.518	26.12	-0.525	27.01	-0.535	27.88	-0.547	28.68
0	-0.480	27.12	-0.489	27.98	-0.500	28.75	-0.513	29.34
1	-0.444	27.88	-0.453	28.63	-0.466	29.21	-0.479	29.55
2	-0.408	28.34	-0.419	28.91	-0.432	29.28	-0.445	29.39
3	-0.373	28.52	-0.384	28.83	-0.397	28.99	-0.411	28.88
4	-0.338	28.39	-0.349	28.44	-0.362	28.39	-0.376	28.13
5	-0.303	28.10	-0.313	27.85	-0.327	27.64	-0.340	27.26
6	-0.267	27.69	-0.277	27.17	-0.290	26.80	-0.302	26.35
7	-0.230	27.24	-0.240	26.51	-0.252	26.03	-0.264	25.51
8	-0.193	26.91	-0.202	25.98	-0.213	25.40	-0.224	24.85
9	-0.156	26.77	-0.163	25.68	-0.173	25,00	-0.183	24.41
10	-0.119	26.86	-0.124	25.65	-0.133	24.86	-0.142	24.23
11	-0.082	27.21	-0.085	25.90	-0.093	25.01	-0.101	24.33
12	-0.045	27.85	-0.047	26.47	-0.054	25.47	-0.060	24.69

Differential Capacity of mercury in aqueous solutions of x mol l^{-1} KCl + (1-x) mol l^{-1} KF

T = 25°C

Potentials with respect to a 1 mol l^{-1} KCl calomel electrode

Reference: R. Payne, Trans. Faraday Soc., 64 (1968) 1638.

x	0		0.0010		0.0025		0.0050	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20	-1.545	18.89	-1.545	18.90	-1.545	18.90	-1.545	18.90
-19	-1.491	18.34	-1.491	18.35	-1.491	18.35	-1.491	18.34
-18	-1.436	17.84	-1.436	17.84	-1.436	17.84	-1.436	17.83
-17	-1.379	17.41	-1.379	17.39	-1.379	17.39	-1.379	17.39
-16	-1.321	17.02	-1.321	16.90	-1.321	16.99	-1.321	17.00
-15	-1.261	16.70	-1.262	16.69	-1.262	16.67	-1.262	16.69
-14	-1.201	16.47	-1.201	16.44	-1.201	16.44	-1.201	16.45
-13	-1.140	16.32	-1.140	16.32	-1.140	16.31	-1.140	16.32
-12	-1.079	16.28	-1.079	16.29	-1.079	16.28	-1.079	16.28
-11	-1.017	16.39	-1.018	16.41	-1.017	16.38	-1.018	16.39
-10	-0.956	16.68	-0.957	16.72	-0.957	16.68	-0.957	16.70
-9	-0.898	17.16	-0.898	17.16	-0.898	17.17	-0.898	17.20
-8	-0.841	17.84	-0.841	17.84	-0.841	17.85	-0.841	17.87
-7	-0.786	18.68	-0.786	18.69	-0.786	18.71	-0.786	18.73
-6	-0.734	19.69	-0.734	19.69	-0.734	19.72	-0.734	19.75
-5	-0.684	20.81	-0.685	20.81	-0.684	20.86	-0.685	20.87
-4	-0.637	21.99	-0.638	22.01	-0.638	22.07	-0.638	22.10
-3	-0.593	23.18	-0.594	23.22	-0.594	23.31	-0.594	23.36
-2	-0.551	24.33	-0.522	24.41	-0.552	24.52	-0.553	24.63
-1	-0.511	25.40	-0.512	25.49	-0.512	25.65	-0.513	25.83
0	-0.472	26.32	-0.473	26.46	-0.474	26.69	-0.475	26.96
1	-0.435	27.09	-0.436	27.31	-0.437	27.62	-0.439	28.03
2	-0.398	27.70	-0.400	28.00	-0.401	28.44	-0.404	29.03
3	-0.362	28.14	-0.364	28.56	-0.367	29.12	-0.370	29.90
4	-0.327	28.46	-0.330	29.01	-0.333	29.70	-0.337	30.67
5	-0.292	28.66	-0.395	29.35	-0.299	30.20	-0.305	31.33
6	-0.257	28.76	-0.262	29.64	-0.266	30.58	-0.273	31.83
7	-0.222	28.75	-0.228	29.79	-0.234	30.83	-0.242	32.16
8	-0.188	28.65	-0.194	29.77	-0.202	30.94	-0.211	32.34
9	-0.123	28.52	-0.161	29.77	-0.169	31.04	-0.180	32.48
10	-0.117	28.59	-0.127	29.89	-0.137	31.21	-0.149	32.68
11	-0.083	28.97	-0.094	30.24	-0.105	31.63	-0.119	33.21
12	-0.049	29.82	-0.061	31.01	-0.074	32.45	-0.089	34.07
13	-0.016	30.70	-0.029	32.24	-0.044	33.78	-0.060	35.49
14	0.017	31.88			-0.015	35.71	-0.033	37.54
15	0.047	33.52					-0.007	40.38

Capacity of mercury in aqueous KCl + KF (cont.)

x	0.010		0.025		0.050		0.10	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20	-1.545	18.90	-1.545	18.90	-1.545	18.90	-1.545	18.90
-19	-1.491	18.34	-1.491	18.36	-1.491	18.33	-1.491	18.33
-18	-1.436	17.83	-1.436	17.85	-1.436	17.84	-1.436	17.84
-17	-1.379	17.39	-1.379	17.40	-1.379	17.40	-1.379	17.40
-16	-1.321	17.00	-1.321	17.02	-1.321	17.02	-1.321	17.02
-15	-1.262	16.68	-1.262	16.71	-1.262	16.71	-1.262	16.71
-14	-1.201	16.44	-1.202	16.47	-1.202	16.47	-1.202	16.47
-13	-1.140	16.31	-1.141	16.31	-1.140	16.31	-1.140	16.31
-12	-1.079	16.29	-1.079	16.28	-1.079	16.29	-1.079	16.29
-11	-1.018	16.42	-1.018	16.39	-1.018	16.40	-1.018	16.40
-10	-0.957	16.72	-0.957	16.67	-0.957	16.69	-0.957	16.69
-9	-0.898	17.21	-0.898	17.17	-0.898	17.19	-0.898	17.21
-8	-0.841	17.89	-0.841	17.84	-0.841	17.87	-0.841	17.94
-7	-0.787	18.75	-0.786	18.73	-0.786	18.77	-0.787	18.89
-6	-0.735	19.78	-0.734	19.79	-0.735	19.86	-0.736	20.11
-5	-0.685	20.95	-0.685	20.99	-0.686	21.14	-0.687	21.57
-4	-0.639	22.21	-0.639	22.33	-0.640	22.61	-0.643	23.30
-3	-0.595	23.51	-0.596	23.75	-0.597	24.27	-0.602	25.27
-2	-0.554	24.83	-0.555	25.29	-0.558	26.11	-0.564	27.50
-1	-0.515	26.16	-0.516	26.92	-0.521	28.10	-0.529	29.87
0	-0.477	27.47	-0.480	28.60	-0.486	30.18	-0.497	32.29
1	-0.442	28.74	-0.446	30.30	-0.454	32.23	-0.467	34.50
2	-0.408	29.96	-0.414	31.90	-0.424	43.06	-0.439	36.43
3	-0.375	31.09	-0.384	33.33	-0.396	35.65	-0.412	37.88
4	-0.343	32.08	-0.354	34.51	-0.368	36.74	-0.386	38.90
5	-0.313	32.89	-0.326	35.36	-0.341	37.46	-0.360	39.37
6	-0.283	33.50	-0.298	35.92	-0.315	37.77	-0.335	39.38
7	-0.253	33.87	-0.270	36.19	-0.288	37.81	-0.309	39.16
8	-0.223	34.08	-0.242	36.25	-0.262	37.69	-0.284	38.78
9	-0.194	34.18	-0.215	36.33	-0.235	37.57	-0.258	38.47
10	-0.165	34.43	-0.187	36.51	-0.208	37.63	-0.232	38.38
11	-0.136	34.91	-0.160	34.03	-0.182	38.04	-0.206	38.69
12	-0.108	35.82	-0.133	38.03	-0.156	38.97	-0.180	39.51
13	-0.080	37.34	-0.107	39.61	-0.130	40.49	-0.155	41.02
14	-0.054	39.46	-0.083	41.90	-0.106	42.76	-0.131	43.26
15	-0.029	42.66	-0.059	45.02	-0.084		-0.109	46.33
16	-0.007	46.57	-0.038	49.14	-0.063		-0.088	50.33
17			-0.019	55.17	-0.043		-0.069	55.60
18			-0.002	62.14	-0.026		-0.052	62.52
19							-0.037	71.38

Capacity of mercury in aqueous KCl + KF (contd.)

x	0.250		0.50		1.00	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-20	-1.545	18.90	-1.545	18.90	-1.545	18.90
-19	-1.491	18.37	-1.491	18.37	-1.491	18.37
-18	-1.436	17.87	-1.436	17.87	-1.436	17.87
-17	-1.380	17.43	-1.380	17.43	-1.380	17.43
-16	-1.322	17.05	-1.322	17.05	-1.322	17.05
-15	-1.262	16.72	-1.262	16.74	-1.262	16.75
-14	-1.202	16.48	-1.202	16.51	-1.202	16.55
-13	-1.141	16.33	-1.141	16.37	-1.142	16.44
-12	-1.080	16.31	-1.080	16.35	-1.081	16.44
-11	-1.019	16.43	-1.019	16.50	-1.020	16.54
-10	-0.958	16.75	-0.959	16.88	-0.960	16.81
-9	-0.899	17.31	-0.901	17.53	-0.903	17.77
-8	-0.843	18.13	-0.845	18.48	-0.848	18.94
-7	-0.789	19.24	-0.793	19.80	-0.797	20.55
-6	-0.739	20.67	-0.744	21.55	-0.751	22.65
-5	-0.693	22.49	-0.700	23.74	-0.709	25.27
-4	-0.650	24.67	-0.660	26.36	-0.671	28.17
-3	-0.611	27.20	-0.624	29.25	-0.638	31.21
-2	-0.576	29.92	-0.592	32.25	-0.607	34.26
-1	-0.545	32.71	-0.562	35.11	-0.579	36.96
0	-0.515	35.28	-0.535	37.64	-0.553	39.34
1	-0.488	37.53	-0.509	39.72	-0.528	41.07
2	-0.462	39.25	-0.484	41.22	-0.504	42.32
3	-0.437	40.45	-0.460	42.08	-0.481	42.90
4	-0.412	41.03	-0.437	42.40	-0.458	42.95
5	-0.388	41.08	-0.413	42.20	-0.434	42.51
6	-0.364	40.78	-0.389	41.64	-0.411	41.76
7	-0.339	40.25	-0.365	40.87	-0.386	40.89
8	-0.314	39.64	-0.340	40.05	-0.362	40.01
9	-0.288	39.14	-0.315	39.43	-0.336	39.30
10	-0.263	38.95	-0.289	39.10	-0.311	38.95
11	-0.237	39.20	-0.264	39.20	-0.285	39.05
12	-0.212	39.98	-0.239	39.88	-0.260	39.77
13	-0.187	41.46	-0.214	41.22	-0.235	41.18
14	-0.164	43.68	-0.190	43.32	-0.211	43.34
15	-0.141	46.68	-0.168	46.31	-0.189	46.34
16	-0.121	50.56	-0.147	50.19	-0.168	50.24
17	-0.102	55.73	-0.128	55.02	-0.149	55.17
18	-0.085	62.22	-0.111	61.30	-0.132	61.32
19	-0.070	70.69	-0.095	69.22	-0.116	68.84
20	-0.056	82.50	-0.081	79.53	-0.102	78.43
21	-0.044	98.88	-0.068	93.05	-0.090	89.71
22	-0.035	116.08	-0.059	113.7	-0.080	106.78
23			-0.051	135.3	-0.071	125.14
24			-0.044	156.8		

Differential capacity of mercury in aqueous solutions of x mol l^{-1} KI + $(1-x)$ mol l^{-1} KF

Potentials with respect to a 1 mol l^{-1} KCl calomel electrode in contact with the working solution

Reference: E. Dutkiewicz and R. Parsons, J. Electroanal. Chem., 11 100 (1966)

x	0.30		0.10		0.030		0.010	
$\sigma/\mu C\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$
-21	-1.582	19.29	-1.582	19.29	-1.582	19.29	-1.582	19.29
-20	-1.529	18.71	-1.529	18.68	-1.529	18.68	-1.529	18.64
-19	-1.475	18.31	-1.475	18.16	-1.475	18.15	-1.475	18.15
-18	-1.420	18.11	-1.419	17.72	-1.419	17.79	-1.419	17.77
-17	-1.364	17.80	-1.362	17.35	-1.362	17.40	-1.362	17.35
-16	-1.308	17.64	-1.304	17.05	-1.304	17.03	-1.304	16.98
-15	-1.251	17.73	-1.245	16.91	-1.245	16.77	-1.244	16.71
-14	-1.195	18.26	-1.186	17.03	-1.185	16.65	-1.184	16.54
-13	-1.142	19.52	-1.127	17.48	-1.125	16.75	-1.123	16.50
-12	-1.093	21.79	-1.072	18.76	-1.066	17.21	-1.063	16.69
-11	-1.050	25.08	-1.022	21.28	-1.009	18.38	-1.004	17.31
-10	-1.013	28.80	-0.978	25.15	-0.958	20.52	-0.948	18.63
-9	-0.981	33.48	-0.942	29.80	-0.913	24.64	-0.897	21.14
-8	-0.952	37.81	-0.911	34.83	-0.876	29.47	-0.854	25.06
-7	-0.928	41.69	-0.884	39.57	-0.845	34.93	-0.817	30.04
-6			-0.860	43.92	-0.818	40.17	-0.787	35.55
-5					-0.795	44.83	-0.761	40.76
-4								
-3								
-2								
-1								
0								

Capacity of Hg in x mol l^{-1} KI + (1 - x) mol l^{-1} KF (cont.)

x	0.0030		0.0010		0.0005		0.00	
$\sigma/\mu C\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$
-21	-1.582	19.29	-1.582	19.29				
-20	-1.529	18.68	-1.529	18.68				
-19	-1.475	18.15	-1.475	18.16				
-18	-1.419	17.77	-1.419	17.73	-1.419	17.73		
-17	-1.362	17.40	-1.362	17.32	-1.361	17.24	-1.367	17.35
-16	-1.304	17.03	-1.303	16.96	-1.303	16.88	-1.309	16.97
-15	-1.245	16.74	-1.244	16.67	-1.243	16.58	-1.250	16.69
-14	-1.184	16.51	-1.183	16.47	-1.182	16.44	-1.189	16.46
-13	-1.124	16.43	-1.123	16.38	-1.121	16.32	-1.128	16.36
-12	-1.063	16.51	-1.062	16.43	-1.060	16.36	-1.067	16.38
-11	-1.003	16.83	-1.001	16.64	-0.999	16.55	-1.006	16.56
-10	-0.944	17.53	-0.942	17.14	-0.940	17.00	-0.947	16.92
-9	-0.889	18.91	-0.885	18.03	-0.882	17.72	-0.888	17.46
-8	-0.839	21.33	-0.831	19.52	-0.827	18.87	-0.832	18.21
-7	-0.796	24.99	-0.783	21.89	-0.776	20.67	-0.779	19.12
-6	-0.759	29.78	-0.740	25.33	-0.730	23.26	-0.728	20.19
-5	-0.728	34.99	-0.703	29.38	-0.690	26.97	-0.679	21.36
-4	-0.701	39.83	-0.671	33.44	-0.656	32.45	-0.634	22.57
-3			-0.644	39.82	-0.627	35.95	-0.591	23.83
-2			-0.620	46.89	-0.601	39.28	-0.550	25.04
-1					-0.576	42.36	-0.511	26.05
0					-0.553	45.40	-0.473	26.82
1					-0.532	47.77	-0.436	27.61
2					-0.511	50.20	-0.401	28.15
3					-0.492	52.75	-0.365	28.49
4					-0.473	55.01	-0.330	28.73
5					-0.456	57.11	-0.296	28.89
6					-0.438	59.04	-0.261	29.00
7					-0.422	60.97	-0.227	29.17
8					-0.406	62.86	-0.192	29.42
9					-0.390	64.84	-0.159	29.87

Differential capacity of mercury in solutions of ammonium salts in liquid ammonia. T = -36.5°C

Potentials with respect to Pb / 0.1 mol l^{-1} $PbNO_3$ electrode in liquid ammonia except the /mol l^{-1} NH_4NO_3 where they are with respect to a hydrogen electrode in the working solution.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 41.

0.01 mol l^{-1} NH_4Cl		0.1 mol l^{-1} NH_4Cl		0.1 mol l^{-1} NH_4NO_3		1 mol^{-1} NH_4NO_3	
E/volts	C/µF cm^{-2}	E/volts	C/µF cm^{-2}	E/volts	C/µF cm^{-2}	E/volts	C/µF cm^{-2}
-1.071	17.35	-1.082	19.31	-1.136	30.13	-0.805	21.64
-0.939	14.42	-1.037	17.80	-1.108	24.33	-0.803	21.55
-0.818	12.85	-0.934	15.38	-1.056	21.69	-0.772	20.32
-0.703	11.13	-0.800	13.28	-0.959	18.62	-0.741	19.30
-0.583	10.03	-0.674	11.62	-0.869	16.66	-0.707	18.31
-0.455	8.93	-0.545	10.28	-0.774	15.08	-0.641	16.86
-0.303	8.28	-0.444	9.48	-0.657	13.28	-0.592	15.89
-0.196	7.99	-0.332	8.90	-0.550	11.96	-0.521	14.64
-0.092	7.84	-0.227	8.52	-0.441	10.91	-0.455	13.71
-0.074	8.15	-0.120	8.35	-0.331	10.19	-0.412	13.09
0.125	9.72	-0.087	8.38	-0.192	9.60	-0.351	12.27
0.199	10.97	-0.061	8.64	-0.106	9.40	-0.300	11.73
0.263	15.68	0.088	9.00	-0.049	9.43	-0.251	11.28
0.337	27.27	0.124	9.14	-0.044	9.43	-0.192	10.75
0.373	46.08	0.204	10.10	0.234	13.45	-0.151	10.47
0.383	68.0	0.270	12.99	0.283	15.52	-0.096	10.15
		0.337	21.03	0.322	18.77	-0.052	9.91
		0.301	15.62	0.354	22.86	0.0	9.71
				0.375	27.00	0.056	9.53
				0.393	36.3	0.110	9,42
						0.167	9.36
						0.195	9.32
						0.252	9.35
						0.299	9.40
						0.355	9.64
						0.390	9.82
						0.424	10.17
						0.446	10.42
						0.472	10.93

Differential capacity of mercury electrode in solutions of KI in liquid ammonia. T = -36.5°C

Potentials with respect to 0.0165% potassium amalgam electrode in contact with the working solution.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 41.

c/mol l^{-1}	0.01		0.1		1.0			
	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-}
	1.403	18.7	1.293	13.99	1.474	179.4	1.061	9.28
	1.354	13.70	1.246	11.87	1.469	151.5	1.040	9.15
	1.242	9.13	1.151	10.10	1.466	140.6	1.017	9.08
	1.280	10.15	1.110	8.94	1.463	128.3	0.920	9.11
	1.171	8.79	1.065	8.67	1.460	124.2	0.833	9.32
	1.087	8.46	1.055	8.63	1.452	107.8	0.773	9.46
	1.022	8.46	0.969	8.67	1.438	109.9	0.693	10.07
	0.913	8.42	0.886	8.87	1.412	90.7	0.572	11.12
	0.843	8.71	0.785	9.21	1.380	84.9	0.447	13.53
	0.765	8.93	0.718	9.55	1.349	72.7	0.319	18.52
	0.667	9.82	0.591	10.58	1.318	60.9	0.253	23.2
	0.591	10.50	0.473	12.28	1.286	42.6	0.166	30.7
	0.494	11.69	0.391	14.30	1.248	26.5	0.137	36.8
	0.403	13.60	0.300	17.88	1.203	15.8		
	0.311	16.79	0.177	26.07	1.170	11.93		
	0.228	20.61	0.166	27.03	1.126	10.11		
	0.118	35.28	0.100	60.8	1.090	9.49		
	0.146	27.20						
	0.092	54.40						

Differential capacity of mercury electrode in solutions of KNO_3 in liquid ammonia. T = -36.5°C

Potentials with respect to 0.0165 % potassium amalgam electrode in contact with the working solution.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 41.

$c/mol\ l^{-1}$	0.3 (saturated)		0.01	
	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$
	1485	26.32	1528	55.9
	1506	30.17	1471	31.5
	1452	23.16	1451	24.38
	1378	17.30	1409	18.07
	1280	11.70	1377	14.34
	1271	11.44	1313	9.46
	1205	10.16	1256	8.63
	1134	9.25	1151	8.32
	1065	8.70	1179	8.37
	968	8.80	1063	8.32
	869	8.99	963	8.49
	770	9.35	872	8.80
	666	9.94	788	9.06
	564	10.91	804	8.94
	446	12.87	686	9.48
	450	12.66	586	10.29
	343	15.94	481	11.81
	242	21.21	376	14.03
	156	29.47	320	16.16
	130	33.75	333	16.02
	111	51.56	240	19.92
			149	27.12
			104	50.9

Differential capacity on mercury in $KF + KHF_2$ at 25°C.

Aqueous solution of small quantities of HF in p M KF (p = ion strength), resulting in a composition of $(F^-) = x$; $(HF_2^-) = p - x$ and

$$(HF) = \frac{1}{K_2} \cdot \frac{p-x}{x}; \quad K_2 = \frac{(HF_2^-)}{(HF)(F^-)}; \quad 4 < K_2 < 5.5 .$$

x and p in mol l^{-2}.

Given are the capacity in $\mu F\ cm^{-2}$, potential in mV vs SCE.

Reference: A.W.M. Verkroost, M. Sluyters-Rehbach and J.H. Sluyters. J. Electroanal. Chem. 24 (1970) 1.

	$C/\mu F\ cm^{-2}$; p = 0.1 M							
x =	0.1	0.1	0.075	0.07	0.057	0.05	0.037	0.03
p - x =	-	-	0.025	0.03	0.043	0.05	0.063	0.07
(HF) =	-	-	0.083	0.077	0.19	0.18	0.43	0.42
K_2 =	4	5.5	4	5.5	4	5.5	4	5.5
E/mV								
+150	29.7		28.2		27.3		26.15	
+100	28.3		26.9		26.0		25.0	
+ 50	27.4		25.95		25.2		24.3	
0	26.5		25.45		24.7		24.0	
-100	25.1		25.1		24.8		24.3	
-200	23.6		24.45		24.7		24.65	
-300	22.0		22.6		22.9		23.6	
-400	20.5		20.55		20.7		21.3	
-500	19.5		19.5		19.6		19.8	
-600	18.8		18.8		18.8		18.8	

Differential capacity on mercury in $KF + KHF_2$ at 25°C.

Aqueous solution of small quantities of HF in p M KF (p = ion strength), resulting in a composition of $(F^-) = x$; $(HF_2^-) = p - x$ and

$$(HF) = \frac{1}{K_2} \cdot \frac{p-x}{x}; \quad K_2 = \frac{(HF_2^-)}{(HF)(F^-)} \quad ; \quad 4 < K_2 < 5.5.$$

x and p in mol l^{-1}.

Given are the capacity in μF cm^{-2}, potential in mV vs SCE.

Reference: A.W.M. Verkroost, M. Sluyters-Rehbach and J.H. Sluyters.
J. Electroanal. Chem. 24 (1970) 1.

	$C/\mu F\ cm^{-2}$; p = 1 M							
x =	0.94	0.93	0.68	0.66	0.35	0.30	0.14	0.11
p - x =	0.06	0.07	0.32	0.34	0.65	0.70	0.86	0.89
(HF) =	0.015	0.013	0.12	0.091	0.46	0.42	1.53	1.49
K_2 =	4	5.5	4	5.5	4	5.5	4	5.5
E/mV								
+200	36.3		35.65		32.05		28.3	
+150	34.1		33.0		29.9		26.7	
+100	32.2		31.5		28.25		25.5	
+ 50	31.3		30.0		26.9		24.7	
0	30.3		29.0		26.2		24.35	
-100	29.6		28.2		26.4		24.5	
-200	29.6		28.9		27.8		25.75	
-300	29.35		29.6		29.35		27.3	
-400	28.25		28.9		28.7		27.6	
-500	25.9		26.25		26.05		25.9	
-600	22.9		22.9		22.9		22.9	

Differential capacity on mercury in $KF + KHF_2$ at 25°C.
Aqueous solution of small quantities of HF in p M KF (p = ion strength), resulting in a composition of $(F^-) = x$; $(HF_2^-) = p - x$ and

$$(HF) = \frac{1}{K_2} \cdot \frac{p-x}{x} \; ; \; K_2 = \frac{(HF_2^-)}{(HF)(F^-)} \; ; \; 4 < K_2 < 5.5 .$$

x and p in mol l^{-1}.
Given are the capacity in $\mu F\ cm^{-2}$, potential in mV vs SCE.

Reference: A.W.M. Verkroost, M. Sluyters-Rehbach and J.H. Sluyters. J. Electroanal. Chem. 24 (1970) 1.

$C/\mu F\ cm^{-2}$ p = 4 M

x =	3.64	3.63	2.83	2.81	1.70	1.62	0.88	0.77
p - x =	0.36	0.37	1.17	1.19	2.30	2.38	3.12	3.23
(HF) =	0.025	0.019	0.10	0.077	0.34	0.27	0.89	0.76
K_2 =	4	5.5	4	5.5	4	5.5	4	5.5
E/mV								
+200	39.0		37.7		36.4		34.0	
+150	36.5		34.95		33.8		31.2	
+100	34.6		32.75		31.6		29.4	
+ 50	33.4		31.3		30.1		-	
0	32.75		30.2		29.05		27.1	
- 50	-		29.4		-		-	
-100	31.9		29.2		27.9		26.5	
-150	31.8		-		-		-	
-200	32.0		29.7		28.4		27.2	
-250	32.3		-		-		-	
-300	32.1		30.75		30.0		29.0	
-350	31.4		-		-		29.6	
-400	-		-		30.25		-	
-450	28.95		28.95		28.95		28.95	
-500	27.1		27.1		27.1		27.1	
-600	23.4		23.4		23.4		23.4	

Differential capacity on mercury in aqueous solutions of 0.1 mol l^{-1} KF + 0.05 mol l^{-1} KOH, at 25°C. Given are the potential in V, capacity in μF cm^{-2} as a function of surface charge σ .

Reference: P. Holmqvist. J. Electroanal. Chem. 78 (1977) 347.

σ /μ C cm^{-2}	-E/V	C/μ F cm^{-2}
-26	1.865	25.27
-24	1.779	21.77
-22	1.863	20.07
-20	1.580	18.66
-18	1.470	17.57
-16	1.353	16.63
-14	1.231	16.37
-12	1.109	16.08
-10	0.982	15.85
-8	0.858	16.35
-6	0.740	17.47
-4	0.629	18.77
-2	0.525	19.92
0	0.428	21.24
2	0.337	23.01
4	0.255	25.28
6	0.184	30.99
8	0.127	41.62
10	0.097	58.43

Differential capacity $C/\mu F\ cm^{-2}$ on mercury electrode at 125 Hz, 25°C for solutions of 1 mmol l^{-1} TlOH with various concentrations of $NaClO_4$

Reference: K. Takahashi and R. Tamamushi. Electrochim. Acta 16 (1971) 1157.

	$C/\mu F\ cm^{-2}$					
Concn. of $NaClO_4/mol\ l^{-1}$	0	5×10^{-6}	1×10^{-5}	2×10^{-5}	1×10^{-4}	1×10^{-3}
$-\xi/V$						
0.15		15.5	16.0	15.7	13.6	16.3
0.2	9.0	9.0	9.7	10.5	13.0	16.8
0.25	7.0	9.0	10.0	11.5	14.5	17.2
0.3	7.4	10.9	12.0	12.9	15.7	17.3
0.4	9.1	12.4	13.7	14.2	16.2	16.7
0.5	7.8	12.1	13.4	14.1	15.6	16.0
0.6	5.7	11.2	12.7	13.3	15.0	15.3
0.7	3.5	10.0	12.0	12.7	14.5	15.0
0.8	2.5	9.5	11.5	12.4	14.4	15.1
0.9	2.1	6.1	10.8	12.4	14.5	15.5
1.0	2.0	3.6	10.3	12.4	14.8	15.9
1.1	2.0	2.8	9.7	12.5	15.2	16.5
1.2	1.6	2.2	8.5	12.6	15.7	16.4
1.3	1.0	1.4	6.4	12.6	16.3	18.7
1.4	0.6	0.9	4.5	12.5	17.0	20.1

ξ : the electrode potential relative to the half-wave potential of Tl(I) reduction

Differential capacity in 0.5 M H_2SO_4 plus KBr = potassium bromide

Reference: R.J. Meakins, J. Appl. Chem. 15 (1965) 416.

	$C/\mu F\ cm^{-2}$			
	KBr			
$c/mol\ l^{-1}$	0	10^{-4}	10^{-3}	10^{-2}
E/V				
0.0	31.2	35.5	93	
0.1	29.1	31.5	41.0	121
0.2	30.7	31.9	34.5	52.4
0.3	33.5	34.6	35.4	43.2
0.4	32.7	33.7	34.3	41.2
0.5	27.2	27.8	28.2	34.0
0.6	22.2	22.4	22.6	24.8
0.7	19.0	19.1	19.1	19.6
0.8	17.3	17.2	17.3	17.1
0.9	16.3	16.1	16.2	15.95
1.0	15.6	15.5	15.5	15.45
1.1	15.4	15.5	15.3	15.4
1.2	15.7	15.5	15.5	15.5
1.3	15.9	15.7	15.6	15.8
1.4				16.4

Differential capacity of mercury in methanolic solutions of x mol l^{-1} CsCl + $(0.1 - x)$ mol l^{-1} LiCl. $T = 25^{\circ}C$

Potential with respect to a saturated calomel electrode in contact with the working solution.

Reference: R. Parsons and A. Stockton, unpublished.

	$C/\mu F\,cm^{-2}$							
x	0.10	0.050	0.020	0.010	0.0040	0.0020	0.0010	0.00
-E/volts								
0.21	45.75							
0.24	39.60							
0.27	37.80	32.60	32.70	31.85	31.65	31.75	31.40	29.30
0.30	27.95	27.30	27.45	27.00	26.90	27.00	26.75	25.30
0.33	23.15	23.00	23.40	22.80	22.70	22.85	22.45	21.20
0.36	19.25	19.30	19.60	19.20	19.20	19.15	18.95	18.00
0.39	16.15	16.50	16.50	16.40	16.30	16.25	16.15	15.40
0.42	14.15	13.35	14.15	14.15	14.15	14.05	14.20	13.50
0.45	12.70	12.50	12.80	12.80	12.80	12.70	12.65	12.25
0.48	11.65	11.60	11.65	11.75	11.75	11.70	11.75	11.30
0.51	10.85	10.85	10.95	10.95	10.95	10.85	10.90	10.70
0.54	10.55	10.25	10.30	10.30	10.40	10.45	10.40	10.10
0.57	10.10	9.90	9.90	9.95	10.00	9.95	10.00	9.75
0.60	9.75	9.65	9.60	9.60	9.60	9.60	9.60	9.40
0.63	9.50	9.40	9.40	9.35	9.40	9.30	9.30	9.20
0.66	9.30	9.10	9.10	9.10	9.20	9.10	9.10	8.95
0.69	9.25	9.05	9.00	8.90	8.95	8.90	8.90	8.80
0.72	9.25	8.95	8.85	8.80	8.85	8.80	8.80	8.70
0.75	9.25	8.95	8.75	8.75	8.75	8.70	8.65	8.60
0.78	9.30	8.95	8.75	8.70	8.65	8.65	8.60	8.55
0.81	9.50	9.05	8.75	8.70	8.65	8.65	8.60	8.55
0.84	9.75	9.10	8.85	8.75	8.65	8.65	8.60	8.55
0.87	10.10	9.40	9.00	8.75	8.75	8.65	8.60	8.55

Capacity of mercury in x mol l^{-1} CsCl + (0.1 - x) mol l^{-1} LiCl in methanol (continued)

	C/μF cm^{-2}							
x	0.10	0.050	0.020	0.010	0.0040	0.0020	0.0010	0.00
-E/volts								
0.90	10.55	9.65	9.10	8.85	8.75	8.70	8.65	8.60
0.93	10.95	10.00	9.30	8.95	8.85	8.75	8.70	8.70
0.96	11.55	10.40	9.60	9.20	8.95	8.85	8.80	8.80
0.99	12.25	10.95	10.00	9.40	9.20	9.00	8.95	8.85
1.02	13.05	11.60	10.30	9.70	9.40	9.15	9.05	9.05
1.05	13.90	12.40	10.95	10.00	9.60	9.35	9.20	9.15
1.08	14.85	13.25	11.65	10.40	9.90	9.60	9.40	9.30
1.11	15.90	14.20	12.40	10.95	10.20	9.80	9.60	9.45
1.14	16.90	15.25	13.20	11.65	10.60	10.15	9.85	9.65
1.17	17.95	16.35	14.35	12.25	11.15	10.45	10.15	9.80
1.20	19.00	17.65	15.35	13.10	11.65	10.85	10.40	10.10
1.23	20.00	18.75	16.60	14.25	12.35	11.40	10.70	10.25
1.26	21.05	19.90	18.05	15.45	13.10	12.00	11.10	10.50
1.29	21.90	21.00	19.50	16.70	14.25	12.70	11.65	10.75
1.32	22.70	22.20	20.95	18.35	15.35	13.45	12.15	11.00
1.35	23.45	23.15	22.20	19.80	16.60	14.60	12.85	11.35
1.38	24.25	24.00	23.50	21.15	18.15	15.70	13.70	11.70
1.41	25.05	24.85	24.75	22.60	19.60	16.95	14.60	12.15
1.44	25.80	25.55	25.80	24.00	21.15	18.40	15.65	12.65
1.47	26.60	26.25	25.60	25.35	22.70	19.75	16.80	13.25
1.50	27.20	26.85	27.35	26.40	24.25	21.20	17.80	13.80
1.53	27.85	27.55	28.95	27.20	25.45	22.65	19.05	14.40
1.56	28.50	28.15	28.55	28.05	26.70	24.00	20.20	15.00
1.59	29.10	28.75	29.20	28.85	27.65	25.25	21.20	15.60

Differential capacity of mercury in solutions of $NaHSO_4$ in formic acid.

T = 25°C

Potential with respect to a saturated calomel electrode in formic acid in contact with a 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

			C/μF cm^{-2}			
c/mol l^{-1}	0.020	0.050	0.10	0.20	0.50	1.00
-E/V						
0.01				24.67	25.12	
0.02		22.74		23.68	24.17	24.60
0.05		21.68		22.70	23.20	23.64
0.08		20.64		21.75	22.26	22.69
0.11		19.62		20.82	21.33	21.69
0.14		18.66		19.89	20.38	20.70
0.17	16.78	17.69	18.57	18.89	19.46	19.85
0.20	15.81	16.75	17.59	17.95	18.59	19.01
0.23	14.86	15.79	16.67	17.01	17.73	18.17
0.26	13.93	14.86	15.75	16.10	16.86	17.38
0.29	12.94	13.91	14.83	15.25	16.05	16.59
0.32	12.05	13.01	13.97	14.41	15.18	15.83
0.35	11.23	12.13	13.10	13.57	14.42	15.10
0.38	10.49	11.37	12.33	12.80	13.67	14.39
0.41	9.74	10.67	11.63	12.11	13.02	13.73
0.44	9.13	10.08	11.03	11.49	12.44	13.12
0.47	8.63	9.59	10.53	10.98	11.89	12.55
0.50	8.30	9.25	10.11	10.56	11.34	12.03
0.55	8.01	8.87	9.59	9.97	10.73	11.34
0.60	7.99	8.68	9.28	9.60	10.24	10.77
0.65	8.09	8.64	9.14	9.39	9.94	10.37
0.70	8.26	8.70	9.12	9.32	9.76	10.13
0.75	8.51	9.03	9.21	9.37	9.76	10.02
0.80	8.81	9.03		9.53	9.87	10.08
0.85		9.30		9.81	10.13	10.29
0.90		9.70		10.23	10.55	10.69
0.95					11.16	

Differential capacity for mercury in solutions of LiH_2PO_4 in formic acid.

T = 25°C

Potentials with respect to saturated calomel electrode in formic acid in contact with 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	$C/\mu F\ cm^{-2}$					
$c/mol\ l^{-1}$	0.020	0.050	0.10	0.20	0.50	1.00
$-E/V$						
0.02	18.52	19.37	20.50	21.20	22.41	23.74
0.05	17.61	18.42	19.48	20.17	21.37	22.74
0.08	16.66	17.56	18.46	19.20	20.37	21.75
0.11	15.97	16.75	17.57	18.31	19.43	20.79
0.14	15.29	15.93	16.69	17.43	18.53	19.83
0.17	14.64	15.18	15.85	16.60	17.64	18.95
0.20	13.96	14.48	15.10	15.82	16.82	18.10
0.23	13.22	13.82	14.37	15.09	16.04	17.31
0.26	12.55	13.17	13.69	14.43	15.32	16.48
0.29	11.91	12.54	13.05	13.76	14.62	15.71
0.32	11.24	11.94	12.45	13.15	13.96	15.01
0.35	10.61	11.34	11.86	12.58	13.34	14.34
0.38	9.98	10.76	11.32	12.03	12.78	13.72
0.41	9.49	10.28	10.85	11.56	12.28	13.17
0.44	8.93	9.83	10.42	11.13	11.82	12.67
0.47	8.52	9.46	10.05	10.76	11.42	12.23
0.50	8.18	9.16	9.76	10.47	11.07	11.84
0.55	7.92	8.83	9.39	10.05	10.60	11.31
0.60	7.93	8.69	9.17	9.75	10.26	10.90
0.65	8.08	8.69	9.08	9.58	10.03	10.61
0.70	8.31	8.78	9.09	9.53	9.92	10.44
0.75	8.57	8.94	9.19	9.58	9.92	10.39
0.80	8.84	9.17	9.38	9.72	10.04	10.48
0.85	9.13	9.45	9.66	9.98	10.29	10.71
0.90	9.44	9.85	10.01	10.37	10.66	11.09

Differential capacity for mercury in solutions of NaH_2PO_4 in formic acid, T = 25°C

Potentials with respect to saturated calomel electrode in formic acid in contact with 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1658.

c = 0.10 mol l^{-1}

E/V	C/μF cm^{-2}	E/V	C/μF cm^{-2}
0.22	29.70	-0.29	13.00
0.19	28.03	-0.32	12.42
0.16	26.71	-0.35	11.85
0.13	24.47	-0.38	11.32
0.10	24.21	-0.41	10.84
0.07	23.10	-0.44	10.40
0.04	21.96	-0.47	10.03
0.01	20.96	-0.50	9.73
-0.02	19.98	-0.55	9.35
-0.05	18.97	-0.60	9.10
-0.08	18.17	-0.65	8.99
-0.11	17.25	-0.70	8.99
-0.14	16.46	-0.75	9.09
-0.17	15.71	-0.80	9.27
-0.20	14.99	-0.85	9.54
-0.23	14.31	-0.90	9.92
-0.26	13.64	-0.95	10.43

Differential capacity in NaH_2PO_4 in formic acid
Reference: J. Lawrence and R. Parsons, Trans. Faraday Soc. 64 (1968) 1658.

	$C/\mu F\ cm^{-2}$				
$c/mol\ l^{-1}$	0.02	0.05	0.20	0.50	1.00
E/volts					
0.01	18.48				
-0.02	17.64	19.25			
-0.05	16.90	18.29			
-0.08	16.22	17.44		20.47	
-0.11	15.58	16.60		19.56	
-0.14	14.94	15.86		18.66	
-0.17	14.28	15.12		17.80	18.20
-0.20	13.63	14.44	15.81	16.97	17.37
-0.23	12.99	13.74	15.09	16.22	16.56
-0.26	12.33	13.13	14.40	15.48	15.80
-0.29	11.66	12.52	13.76	14.79	15.09
-0.32	11.01	11.93	13.15	14.15	14.42
-0.35	10.38	11.34	12.58	13.51	13.81
-0.38	9.78	10.79	12.04	12.96	13.24
-0.41	9.19	10.30	11.56	12.46	12.70
-0.44	8.67	9.86	11.14	12.00	12.26
-0.47	8.22	9.51	10.77	11.59	11.83
-0.50	8.12	9.20	10.45	11.25	11.47
-0.55	7.66	8.83	10.01	10.76	10.96
-0.60	7.66	8.68	9.72	10.41	10.60
-0.65	7.80	8.65	9.54	10.18	10.35
-0.70	8.03	8.71	9.48	10.08	10.22
-0.75	8.25	8.88	9.52	10.08	10.23
-0.80	8.53	9.08	9.68	10.24	10.37
-0.85	8.77	9.37		10.54	10.69
-0.90	9.04			10.99	
-0.95	9.34			11.64	
-1.00				12.46	

Differential capacity for mercury in solutions of KH_2PO_4 in formic acid.
T = 25°C

Potentials with respect to saturated calomel electrode in formic acid in contact with 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	$C/\mu F\,cm^{-2}$					
$c/mol\ l^{-1}$	0.020	0.050	0.10	0.20	0.50	1.00
$-E/V$						
0.02	18.26	19.52	20.67	21.50	22.87	24.20
0.05	17.37	18.56	19.65	20.47	21.79	23.14
0.08	16.68	17.63	18.77	19.51	20.82	22.13
0.11	15.96	16.82	17.82	18.57	19.82	21.09
0.14	15.21	16.04	16.94	17.69	18.91	20.11
0.17	14.49	15.31	16.04	16.86	18.03	19.17
0.20	13.74	14.58	15.26	16.08	17.19	18.26
0.23	13.06	13.91	14.58	15.32	16.38	17.41
0.26	12.42	13.26	13.92	14.60	15.60	16.60
0.29	11.77	12.65	13.26	13.92	14.90	15.83
0.32	11.17	12.05	12.61	13.28	14.23	15.11
0.35	10.53	11.49	12.03	12.71	13.68	14.45
0.38	9.93	10.97	11.52	12.18	13.09	13.86
0.41	9.45	10.49	11.07	11.71	12.57	13.31
0.44	9.04	10.08	10.65	11.27	12.11	12.81
0.47	8.66	9.71	10.25	10.88	11.70	12.37
0.50	8.29	9.41	9.94	10.55	11.34	11.98
0.55	8.00	9.07	9.55	10.11	10.83	11.44
0.60	7.98	8.90	9.28	9.82	10.48	11.04
0.65	8.07	8.87	9.16	9.64	10.24	10.78
0.70	8.26	8.94	9.18	9.60	10.13	10.66
0.75	8.51	9.12	9.30	9.67	10.17	10.69
0.80	8.82	9.39	9.52	9.91	10.39	10.87
0.85	9.15	9.70	9.86	10.27	10.77	11.24
0.90	9.51	10.13	10.36	10.77	11.30	11.84
0.95			11.03	11.50	12.06	12.68

Differential capacity for mercury in solutions of CsH_2PO_4 in formic acid.

T = 25°C

Potentials with respect to saturated calomel electrode in formic acid in contact with 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	C/μF cm^{-2}					
c/mol l^{-1}	0.020	0.050	0.10	0.20	0.50	1.00
-E/V						
0.02	18.31	19.41	20.40	21.32	22.80	24.41
0.05	17.51	18.48	19.43	20.32	21.77	23.24
0.08	16.75	17.58	18.52	19.38	20.79	22.20
0.11	15.97	16.79	17.66	18.48	19.79	21.23
0.14	15.22	16.02	16.85	17.62	18.92	20.26
0.17	14.50	15.28	16.09	16.82	18.08	19.39
0.20	13.84	14.59	15.35	16.05	17.26	18.62
0.23	13.20	13.84	14.65	15.34	16.48	17.86
0.26	12.55	13.25	14.00	14.66	15.76	17.12
0.29	11.90	12.64	13.37	14.02	15.07	16.35
0.32	11.25	12.06	12.78	13.40	14.42	15.66
0.35	10.61	11.48	12.18	12.83	13.80	14.95
0.38	9.96	10.91	11.65	12.30	13.25	14.37
0.41	9.96	10.41	11.16	11.81	12.76	13.82
0.44	8.93	9.99	10.74	11.40	12.31	13.28
0.47	8.56	9.61	10.37	11.04	11.94	12.86
0.50	8.19	9.33	10.09	10.73	11.63	12.37
0.55	7.97	9.03	9.75	10.37	11.25	11.97
0.60	8.03	8.94	9.60	10.18	11.04	11.75
0.65	8.29	9.02	9.64	10.18	11.01	11.71
0.70	8,65	9.28	9.83	10.34	11.17	11.89
0.75	9.08	9.66	10.20	10.69	11.54	12.25
0.80	9.66	10.19	10.73	11.24	12.10	12.80
0.85	10.34	10.87	11.46	11.99	12.87	13.61
0.90	11.05	11.73	12.39	12.91	13.82	

Differential capacity of mercury in solutions of Na_2SO_4 in formic acid.

T = 25°C

Potential with respect to saturated calomel electrode in formic acid in contact with a 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

			$C/\mu F\,cm^{-2}$			
$c/mol\ l^{-1}$	0.020	0.050	0.10	0.20	0.50	1.00
-E/V						
0.08	20.12		21.78			
0.11	19.10		20.80			
0.14	18.17		19.79		20.10	20.69
0.17	17.16	18.22	18.82	19.00	19.28	19.85
0.20	16.13	17.26	17.85	18.11	18.48	19.01
0.23	15.06	16.30	16.90	17.23	17.67	18.22
0.26	14.13	15.39	15.96	16.36	16.87	17.42
0.29	13.29	14.47	15.06	15.51	16.09	16.68
0.32	12.37	13.60	14.19	14.69	15.29	15.93
0.35	11.45	12.76	13.36	13.86	14.56	15.21
0.38	10.71	11.99	12.62	13.17	13.84	14.53
0.41	10.01	11.31	11.94	12.53	13.21	13.90
0.44	9.46	10.74	11.28	11.93	12.65	13.30
0.47	9.03	10.25	10.86	11.42	12.10	12.79
0.50	8.97	9.87	10.46	11.00	11.68	12.31
0.55	8.32	9.41	9.91	10.41	11.04	11.65
0.60	8.20	9.14	9.56	10.01	10.58	11.14
0.65	8.20	9.04	9.39	9.75	10.25	10.76
0.70	8.30	9.03	9.33	9.64	10.07	10.53
0.75	8.46	9.13	9.39	9.65	10.02	10.44
0.80	8.70	9.32	9.57	9.77	10.10	10.49
0.85	8.97		9.84	10.01	10.34	10.71
0.90	9.28		10.25			11.12
0.95	9.80		10.85			
1.00	10.40		11.66			
1.05	11.30					

Differential capacity for mercury in solutions of sodium formate in formic acid. T = 25°C

Potentials with respect to saturated calomel electrode in formic acid in contact with 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	$C/\mu F\,cm^{-2}$					
$c/mol\,l^{-1}$	0.020	0.050	0.10	0.20	0.50	1.00
$-E/V$						
0.05			18.00			
0.08			17.05			
0.11			16.39			17.18
0.14	14.92	15.20	15.71	15.79	16.29	16.47
0.17	14.28	14.59	15.08	15.20	15.65	15.84
0.20	13.65	14.01	14.47	14.63	15.07	15.25
0.23	13.02	13.43	13.86	14.07	14.52	14.71
0.26	12.40	12.86	13.26	13.55	13.98	14.16
0.29	11.79	12.31	12.73	13.04	13.48	13.65
0.32	11.17	11.79	12.23	12.54	13.01	13.18
0.35	10.56	11.26	11.71	12.06	12.54	12.72
0.38	9.96	10.73	11.17	11.64	12.14	12.33
0.41	9.43	10.27	10.75	11.24	11.77	11.97
0.44	8.92	9.86	10.42	10.88	11.44	11.66
0.47	8.49	9.50	10.07	10.56	11.14	11.37
0.50	8.14	9.20	9.75	10.28	10.87	11.12
0.55	7.84	8.84	9.41	9.90	10.49	10.78
0.60	7.79	8.65	9.15	9.62	10.21	10.53
0.65	7.89	8.60	9.05	9.45	10.04	10.37
0.70	8.09	8.64	9.03	9.38	9.96	10.31
0.75	8.32	8.77	9.07	9.42	9.98	10.36
0.80	8.57	8.96	9.26	9.57	10.15	10.55
0.85	8.84	9.22	9.51	9.85	10.44	10.91
0.90	9.13	9.58	9.96	10.25	10.88	11.46
0.95	9.50		10.54			
1.00			11.16			

Differential capacity for mercury in solutions of 0.1 M HCOONa in formic acid + water mixtures. T = 25°C

Potential with respect to mercury - mercurous formate electrode in the working solution.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

0 vol. % HCOOH

-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}
0.285	46.34	0.675	27.68	1.185	18.82
0.315	40.39	0.705	27.91	1.235	18.13
0.345	36.18	0.735	28.01	1.285	17.46
0.375	33.16	0.765	27.89	1.335	16.96
0.405	30.78	0.795	27.53	1.385	16.51
0.435	29.15	0.825	26.82	1.435	16.51
0.465	28.02	0.855	25.94	1.485	15.93
0.495	27.25	0.885	24.86	1.535	15.76
0.525	26.86	0.935	23.13	1.585	15.71
0.555	26.77	0.985	21.79	1.635	15.75
0.585	26.78	1.035	20.88	1.685	15.83
0.615	27.00	1.085	20.18	1.735	15.98
0.645	27.32	1.135	19.52	1.785	16.20

Capacity in formic acid + water mixtures (continued)
HCOONa 0.1 M

1 vol.% HCOOH

$-E$/volts	$C/\mu F\ cm^{-2}$	$-E$/volts	$C/\mu F\ cm^{-2}$	$-E$/volts	$C/\mu F\ cm^{-2}$
0.411	27.68	0.741	25.74	1.121	18.71
0.441	27.00	0.771	25.19	1.171	18.17
0.471	26.63	0.801	24.35	1.221	17.61
0.501	26.44	0.831	23.55	1.271	17.12
0.531	26.39	0.861	22.65	1.321	16.73
0.561	26.42	0.891	21.89	1.371	16.42
0.591	26.43	0.921	21.26	1.421	16.17
0.621	26.48	0.971	20.38	1.471	16.01
0.651	26.55	1.021	19.76	1.521	16.00
0.681	26.42	1.071	19.22	1.571	15.99
0.711	26.16				

5 vol.% HCOOH

$-E$/volts	$C/\mu F\ cm^{-2}$	$-E$/volts	$C/\mu F\ cm^{-2}$	$-E$/volts	$C/\mu F\ cm^{-2}$
0.413	26.31	0.743	22.15	1.173	17.97
0.443	25.44	0.773	21.82	1.223	17.69
0.473	24.75	0.803	21.46	1.273	17.35
0.503	24.22	0.833	21.05	1.323	17.04
0.533	23.73	0.863	20.57	1.373	16.76
0.563	23.46	0.893	20.10	1.423	16.53
0.593	23.23	0.923	19.71	1.473	16.35
0.623	22.99	0.973	19.08	1.523	16.27
0.653	22.73	1.023	18.75	1.573	16.24
0.683	22.59	1.073	18.50	1.623	16.25
0.713	22.38	1.123	18.24		

Capacity in formic acid + water mixtures (continued)

HCOONa 0.1 M.

10 vol. % HCOOH

-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}
0.454	24.70	0.784	20.31	1.184	18.00
0.484	23.85	0.814	20.05	1.234	17.91
0.514	23.19	0.844	19.77	1.284	17.75
0.544	22.64	0.874	19.47	1.334	17.59
0.574	22.19	0.904	19.10	1.384	17.41
0.604	21.80	0.934	18.78	1.434	17.22
0.634	21.50	0.984	18.33	1.484	17.04
0.664	21.23	1.034	18.10	1.534	16.93
0.694	21.01	1.084	18.05	1.584	16.86
0.724	20.78	1.134	18.05	1.634	16.81
0.754	20.55				

30 vol. % HCOOH

-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}
0.432	23.09	0.762	17.23	1.192	16.74
0.462	22.39	0.792	16.93	1.242	17.12
0.492	21.50	0.822	16.72	1.292	17.49
0.522	20.73	0.852	16.50	1.342	17.85
0.552	20.07	0.882	16.31	1.392	18.19
0.582	19.47	0.912	16.16	1.442	18.38
0.612	18.96	0.942	16.04	1.492	18.55
0.642	18.52	0.992	15.93	1.542	18.65
0.672	18.15	1.042	16.00	1.592	18.74
0.702	17.80	1.092	16.14	1.642	18.72
0.732	17.51	1.142	16.41		

Capacity in formic acid + water mixtures (continued)
HCOONa 0.1 M.

50 vol. % HCOOH

-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}
0.422	22.84	0.752	15.62	1.165	14.42
0.452	21.71	0.782	15.33	1.212	14.92
0.482	20.76	0.812	14.98	1.262	15.51
0.512	19.94	0.842	14.68	1.312	16.20
0.542	19.24	0.872	14.45	1.362	16.85
0.572	18.63	0.902	14.26	1.412	17.53
0.602	18.01	0.932	14.09	1.462	18.22
0.632	17.47	0.962	13.91	1.512	18.85
0.662	16.95	1.012	13.84	1.562	19.35
0.692	16.45	1.062	13.89	1.612	19.74
0.722	16.04	1.112	14.09		

70 vol. % HCOOH

-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}
0.485	20.90	0.785	14.46	1.095	12.32
0.515	19.67	0.815	14.05	1.145	12.51
0.545	18.83	0.845	13.69	1.195	12.80
0.575	18.13	0.875	13.36	1.245	13.27
0.605	17.45	0.905	13.09	1.295	13.85
0.635	16.84	0.935	12.85	1.345	14.59
0.665	16.25	0.965	12.67	1.395	15.44
0.695	15.72	0.995	12.49	1.445	16.33
0.725	15.32	1.045	12.34	1.495	17.24
0.755	14.86				

Capacity in formic acid + water (continued)

HCOONa 0.1 M.

90 vol. % HCOOH

-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}
0.522	19.77	0.792	13.90	1.082	10.40
0.552	18.93	0.822	13.42	1.132	10.22
0.582	18.15	0.852	12.96	1.182	10.19
0.612	17.39	0.882	12.48	1.232	10.29
0.642	16.66	0.912	12.09	1.282	10.53
0.672	16.08	0.942	11.73	1.332	10.90
0.702	15.51	0.972	11.38	1.382	11.41
0.732	14.93	1.002	11.05	1.432	12.06
0.762	14.44	1.032	10.79		

100 vol. % HCOOH

-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}	-E/volts	C/μF cm^{-2}
0.613	18.00	0.883	12.23	1.213	9.05
0.643	17.05	0.913	11.71	1.263	9.03
0.673	16.39	0.943	11.17	1.313	9.07
0.703	15.71	0.973	10.75	1.363	9.26
0.733	15.08	1.003	10.42	1.413	9.51
0.763	14.47	1.033	10.07	1.463	9.96
0.793	13.86	1.063	9.75	1.513	10.54
0.823	13.26	1.113	9.41	1.563	11.16
0.853	12.73	1.163	9.15		

Differential capacity on mercury in aqueous solutions of 0.1 M NaF with different concentrations of thiourea (c_{th} in moles/l). T = 25°C. Potential relative to normal calomel electrode.

Reference : F.W. Schapink, M. Oudeman, K.W. Leu and J.N. Helle,
Trans. Faraday Soc., 56, 415 (1960)

$\frac{E}{\text{Volts}}$ c_{th}	0	0.001	0.005	0.01	0.05	0.08	0.10	0.25	0.50
-0.40	21.0	24.4	29.5	31.8	32.7	32.7	32.7	32.7	32.6
-0.50	20.0	21.6	25.7	27.6	30.5	30.5	30.7	29.9	28.2
-0.55	19.7	21.2	24.2	25.8	28.4	28.7	28.3	28.5	26.5
-0.60	19.4	21.1	24.0	25.2	26.9	27.6	27.0	27.1	25.7
-0.65	19.0	20.8	23.9	25.2	25.8	26.2	26.0	26.2	25.0
-0.70	18.6	20.2	23.3	24.9	25.9	25.9	25.4	25.3	24.5
-0.75	18.2	19.6	22.6	24.2	27.1	26.3	26.0	24.9	24.2
-0.80	17.6	18.7	21.7	23.1	28.0	27.5	27.2	25.2	24.0
-0.85	17.1	18.1	20.2	21.6	27.6	28.4	28.6	26.5	24.5
-0.90	16.7	17.4	18.9	20.1	26.2	28.4	29.0	28.1	25.7
-0.95	16.4	16.7	17.9	18.8	24.2	27.5	28.2	29.2	27.4
-1.00	16.1	16.4	17.2	17.6	22.3	25.6	26.5	30.2	29.0
-1.05	15.9	16.1	16.8	17.1	20.2	23.5	24.3	29.0	30.5
-1.10	15.8	16.0	16.3	16.6	19.0	21.0	22.0	27.6	30.5
-1.15	15.8	15.9	16.2	16.5	18.1	19.4	20.0	25.2	29.0
-1.20	15.8	15.9	16.1	16.4	17.4	18.3	18.8	23.0	26.6
-1.30	16.1	16.1	16.3	16.5	17.0	17.4	17.6	19.9	22.7
-1.40	16.7	16.7	16.8	17.0	17.2	17.4	17.4	19.0	20.4
-1.50	17.5	17.5	17.5	17.5	17.5	17.7	17.7	18.3	18.9
-1.60	18.3	18.3	18.3	18.3	18.3	18.3	18.3	18.3	18.9
-1.70	19.2	19.2	19.2	19.2	19.2	19.2	19.2	19.2	19.2

Note. The data for pure NaF differs from that obtained by Grahame by more than an acceptable amount. The latter is to be preferred since it was obtained after a long series of studies of this system. However, Schapink et al.'s data is useful for comparison of his measurements in the presence of thiourea.

Differential capacity on mercury in methanol. Salt: potassium iodide, concentration in Mole/l. T = 25°C. Potential with respect to an aqueous saturated calomel electrode dipping into 0.1 M KCl in methanol.

Reference : J.D. Garnish and R. Parsons, Trans. Faraday Soc., 63, 1754 (1967)

conc. 0.02		0.05		0.1	
E	C	E	C	E	C
Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
-1.596	24.96	-1.605	26.18		
-1.546	23.18	-1.555	24.31	-1.596	26.81
-1.496	20.91	-1.505	22.23	-1.546	24.82
-1.446	18.94	-1.455	20.09	-1.496	22.98
-1.396	16.97	-1.405	18.03	-1.446	20.81
-1.346	15.08	-1.355	16.05	-1.396	18.60
-1.296	13.56	-1.305	14.36	-1.346	16.52
-1.246	12.20	-1.255	12.91	-1.296	14.69
-1.196	11.21	-1.205	11.74	-1.246	13.23
-1.146	10.49	-1.155	10.89	-1.196	12.02
-1.096	9.87	-1.105	10.18	-1.146	11.08
-1.046	9.39	-1.055	9.69	-1.096	10.35
-0.996	9.06	-1.005	9.31	-1.046	9.83
-0.956	8.89	-0.985	9.18	-0.996	9.49
-0.926	8.75	-0.955	9.08	-0.966	9.35
-0.896	8.71	-0.925	9.03	-0.946	9.33
-0.866	8.72	-0.895	9.07	-0.916	9.36
-0.836	8.80	-0.865	9.25	-0.886	9.60
-0.806	9.05	-0.835	9.69	-0.856	10.11
-0.776	9.53	-0.805	10.51	-0.826	11.11
-0.746	10.42	-0.775	12.02	-0.796	12.91
-0.716	12.09	-0.745	14.76	-0.766	16.01
-0.686	15.13	-0.715	19.50	-0.736	21.14
-0.656	20.46	-0.685	27.08	-0.706	28.65
-0.626	29.29	-0.655	37.64	-0.676	38.91
-0.596	41.52	-0.625	51.18	-0.646	51.14
-0.566	56.37	-0.595	65.60	-0.616	64.91
-0.536	74.20	-0.565	82.19	-0.586	78.95
-0.506	95.00	-0.535	100.0	-0.556	96.14

cont.

Capacity in KI-methanol (cont.)

conc. 0.2		0.3		0.5	
E	C	E	C	E	C
Volts	$\mu F/cm^2$	Volts	$\mu F/cm^2$	Volts	$\mu F/cm^2$
-1.593	27.96	-1.541	26.60	-1.583	29.32
-1.543	25.88	-1.491	24.47	-1.533	27.32
-1.493	23.83	-1.441	22.22	-1.483	25.19
-1.443	21.56	-1.391	19.85	-1.433	22.87
-1.393	19.26	-1.341	17.63	-1.383	20.40
-1.343	17.14	-1.291	15.63	-1.333	18.11
-1.293	15.19	-1.241	14.02	-1.283	16.00
-1.243	13.59	-1.191	12.63	-1.233	14.25
-1.193	12.32	-1.141	11.58	-1.183	12.87
-1.143	11.28	-1.091	10.81	-1.133	11.83
-1.093	10.54	-1.061	10.49	-1.083	11.08
-1.043	10.01	-1.031	10.24	-1.043	10.72
-0.993	9.68	-1.001	10.10	-1.013	10.61
-0.973	9.65	-0.971	10.13	-0.983	10.72
-0.943	9.71	-0.941	10.34	-0.953	11.07
-0.913	9.96	-0.911	10.87	-0.923	11.83
-0.883	10.55	-0.881	11.85	-0.893	13.20
-0.853	11.68	-0.851	13.70	-0.863	15.49
-0.823	13.63	-0.821	16.74	-0.833	19.00
-0.793	16.92	-0.791	21.26	-0.803	24.08
-0.763	22.18	-0.761	27.83	-0.773	30.68
-0.733	29.55	-0.731	36.22	-0.743	38.81
-0.703	39.11	-0.701	46.03	-0.713	48.05
-0.673	50.22	-0.671	56.81	-0.683	58.12
-0.643	62.69	-0.641	68.59	-0.653	69.00
-0.613	75.64	-0.611	81.20	-0.623	82.19
-0.583	90.14	-0.591	90.59	-0.593	95.99

Differential double layer capacity on mercury in methanol. Salt: KI + KF, concentration indicated. T = 25°C. Potential with respect to aqueous saturated calomel electrode. The last figures of E and C in this table are not significant.

Reference : J.D. Garnish and R. Parsons, Trans. Faraday Soc., 63, 1754 (1967)

salt	10^{-3} M KI + 0.3 M KF		$2\text{x}10^{-3}$ M KI+0.3 M KF		$5\text{x}10^{-3}$ M KI+0.3 M KF	
σ	E	C	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
-16	-1.5706	27.184	-1.5706	27.184	-1.5706	27.184
-15	-1.5327	25.661	-1.5329	25.892	-1.5329	25.835
-14	-1.4924	23.990	-1.4929	24.098	-1.4928	24.057
-13	-1.4490	22.079	-1.4496	22.125	-1.4495	22.138
-12	-1.4015	19.967	-1.4021	19.987	-1.4020	19.960
-11	-1.3482	17.628	-1.3490	17.648	-1.3488	17.650
-10	-1.2872	15.233	-1.2881	15.261	-1.2878	15.252
- 9	-1.2163	13.083	-1.2172	13.093	-1.2170	13.097
- 8	-1.1339	11.284	-1.1349	11.320	-1.1347	11.307
- 7	-1.0393	9.962	-1.0407	10.027	-1.0403	9.994
- 6	-0.9342	9.149	-0.9360	9.177	-0.9356	9.197
- 5	-0.8223	8.800	-0.8244	8.858	-0.8247	8.968
- 4	-0.7103	9.263	-0.7140	9.537	-0.7183	10.386
- 3	-0.6132	12.280	-0.6241	13.836	-0.6400	16.832
- 2	-0.5524	22.218	-0.5687	23.746	-0.5941	28.156
- 1	-0.5163	33.904	-0.5344	38.222	-0.5644	39.711
0	-0.4908	44.885	-0.5120	51.018	-0.5421	49.988
1	-0.4708	55.527	-0.4941	60.956	-0.5238	59.331
2	-0.4541	64.438	-0.4789	70.528	-0.5080	67.522
3	-0.4396	73.018			-0.4939	74.985
4	-0.4264	78.778			-0.4812	82.019
5					-0.4695	88.799

cont.

Capacity in KI + KF in methanol (cont.)

salt	0.01 M KI+0.29 M KF		0.02 M KI+0.28 M KF		0.05 M KI+0.25 M KF	
σ C/cm^2	E Volts	C μF/cm^2	E Volts	C μF/cm^2	E Volts	C μF/cm^2
-16	-1.5706	27.184	-1.5706	27.184	-1.5706	27.184
-15	-1.5328	25.813	-1.5329	25.860	-1.5329	25.831
-14	-1.4928	24.075	-1.4929	24.133	-1.4928	24.112
-13	-1.4494	22.097	-1.4497	22.189	-1.4496	22.185
-12	-1.4018	19.904	-1.4023	20.037	-1.4023	20.033
-11	-1.3484	17.555	-1.3493	17.681	-1.3492	17.656
-10	-1.2873	15.258	-1.2885	15.293	-1.2883	15.292
- 9	-1.2163	13.033	-1.2178	13.114	-1.2177	13.127
- 8	-1.1336	11.274	-1.1355	11.327	-1.1355	11.336
- 7	-1.0390	9.970	-1.0415	10.039	-1.0415	10.039
- 6	-0.9340	9.170	-0.9374	9.268	-0.9376	9.333
- 5	-0.8234	9.115	-0.8288	9.383	-0.8320	10.036
- 4	-0.7225	11.554	-0.7339	12.724	-0.7478	15.051
- 3	-0.6547	19.602	-0.6727	21.542	-0.6958	24.424
- 2	-0.6145	31.182	-0.6353	32.844	-0.6608	34.216
- 1	-0.5874	43.146	-0.6091	43.796	-0.6360	46.371
0	-0.5666	53.419	-0.5886	53.772	-0.6164	55.650
1	-0.5494	62.617	-0.5714	62.667	-0.5997	63.760
2			-0.5564	70.843	-0.5848	71.024
3			-0.5430	78.460	-0.5714	77.977
4			-0.5308	85.893	-0.5591	84.648
5			-0.5196	92.450	-0.5477	90.912

cont.

Capacity in KI + KF in methanol (cont.)

salt	0.1 M KI+0.2 M KF		0.2 M KI+0.1 M KF		0.3 M KI	
σ	E	C	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
-16	-1.5706	27.184	-1.5706	27.184	-1.5706	27.184
-15	-1.5329	25.837	-1.5329	25.831	-1.5330	25.991
-14	-1.4929	24.143	-1.4928	24.107	-1.4931	24.168
-13	-1.4497	22.188	-1.4495	22.111	-1.4500	22.222
-12	-1.4022	19.945	-1.4020	19.952	-1.4026	19.971
-11	-1.3488	17.553	-1.3486	17.543	-1.3493	17.603
-10	-1.2877	15.196	-1.2874	15.187	-1.2882	15.206
- 9	-1.2164	13.034	-1.2161	13.024	-1.2171	13.050
- 8	-1.1336	11.247	-1.1334	11.269	-1.1346	11.313
- 7	-1.0392	10.046	-1.0394	10.154	-1.0412	10.248
- 6	-0.9362	9.558	-0.9390	10.028	-0.9429	10.435
- 5	-0.8373	11.344	-0.8481	12.858	-0.8573	13.916
- 4	-0.7658	18.060	-0.7853	20.202	-0.7991	21.409
- 3	-0.7213	27.862	-0.7445	29.613	-0.7601	30.375
- 2	-0.6908	37.856	-0.7151	38.691	-0.7312	38.936
- 1	-0.6672	46.999	-0.6917	46.979	-0.7078	46.758
0	-0.6476	55.375	-0.6720	54.440	-0.6879	53.841
1	-0.6307	62.766	-0.6547	61.485	-0.6704	60.439
2	-0.6155	69.512	-0.6393	67.931	-0.6547	66.718
3	-0.6018	75.745	-0.6252	73.909	-0.6403	72.383
4	-0.5890	81.351	-0.6121	79.476	-0.6269	77.312
5	-0.5772	87.226	-0.6000	85.178	-0.6145	83.043

Differential capacity of mercury in methanol + 0.3 M KF + c_{th} mole/l thiourea. T = 25°C. Potential with respect to aqueous saturated calomel electrode. The last figures of E and C in this table are not significant.

Reference : J.D. Garnish and R. Parsons, Trans. Faraday Soc., 63, 1754 (1967)

c_{th} =	0		0.01		0.02	
σ C/cm^2	E Volts	C μ F/cm^2	E Volts	C μ F/cm^2	E Volts	C μ F/cm^2
-19	-1.6738	31.046				
-18	-1.6408	29.737	-1.6408	29.737	-1.6408	29.737
-17	-1.6065	28.491	-1.6064	28.486	-1.6065	28.570
-16	-1.5706	27.184	-1.5705	27.180	-1.5706	27.253
-15	-1.5328	25.723	-1.5327	25.718	-1.5330	25.816
-14	-1.4926	24.057	-1.4925	24.052	-1.4930	24.226
-13	-1.4493	22.131	-1.4492	22.126	-1.4500	22.274
-12	-1.4018	19.921	-1.4016	19.915	-1.4028	20.036
-11	-1.3484	17.595	-1.3483	17.589	-1.3499	17.806
-10	-1.2873	15.179	-1.2871	15.173	-1.2897	15.487
- 9	-1.2160	13.007	-1.2158	13.054	-1.2198	13.200
- 8	-1.1329	11.175	-1.1333	11.281	-1.1380	11.400
- 7	-1.0373	9.846	-1.0387	9.953	-1.0443	10.044
- 6	-0.9306	8.968	-0.9331	9.076	-0.9396	9.159
- 5	-0.8159	8.566	-0.8205	8.707	-0.8281	8.885
- 4	-0.6993	8.702	-0.7067	9.049	-0.7178	9.484
- 3	-0.5884	9.431	-0.6041	10.737	-0.6229	11.878
- 2	-0.4880	10.584	-0.5241	15.111	-0.5508	16.742
- 1	-0.3994	12.079	-0.4693	21.923	-0.5013	24.023
0	-0.3220	13.780	-0.4299	30.902	-0.4659	33.064
1	-0.2537	15.646			-0.4391	41.904
2	-0.1938	17.864			-0.4174	50.236
3	-0.1413	20.369			-0.3988	57.396
4	-0.0956	23.410			-0.3823	63.826
5	-0.0553	26.284			-0.3673	69.384

cont.

Capacity in KF in methanol + thiourea (cont.)

c_{th} =	0.05		0.1		0.2	
σ	E	C	E	C	E	C
C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
-18	-1.6408	29.737	-1.6408	29.737	-1.6260	29.250
-17	-1.6064	28.520	-1.6064	28.473	-1.5910	27.985
-16	-1.5706	27.268	-1.5705	27.210	-1.5545	26.785
-15	-1.5329	25.832	-1.5327	25.808	-1.5162	25.443
-14	-1.4929	24.167	-1.4928	24.215	-1.4757	23.936
-13	-1.4499	22.348	-1.4498	22.375	-1.4324	22.210
-12	-1.4030	20.254	-1.4032	20.505	-1.3852	20.207
-11	-1.3506	17.957	-1.3516	18.270	-1.3330	18.087
-10	-1.2910	15.634	-1.2931	15.953	-1.2740	15.850
- 9	-1.2219	13.398	-1.2258	13.832	-1.2063	13.793
- 8	-1.1414	11.568	-1.1477	11.915	-1.1286	12.068
- 7	-1.0491	10.210	-1.0582	10.525	-1.0411	10.897
- 6	-0.9465	9.382	-0.9588	9.714	-0.9469	10.482
- 5	-0.8383	9.245	-0.8552	9.800	-0.8541	11.430
- 4	-0.7347	10.383	-0.7591	11.500	-0.7758	14.835
- 3	-0.6509	14.107	-0.6851	16.347	-0.7194	21.603
- 2	-0.5916	20.443	-0.6341	23.613	-0.6805	30.346
- 1	-0.5503	28.706	-0.5980	32.399	-0.6514	38.899
0	-0.5202	38.075	-0.5705	40.717	-0.6278	45.795
1	-0.4968	47.628	-0.5478	47.260	-0.6073	51.320
2	-0.4771	53.949	-0.5278	52.513	-0.5885	54.722
3	-0.5493	58.380	-0.5095	56.485	-0.5706	56.460
4	-0.4428	63.203	-0.4922	59.490	-0.5530	57.111
5	-0.4274	66.653	-0.4759	62.837	-0.5354	56.503

cont.

Capacity in KF in methanol + thiourea (cont.)

c_{th} = 0.5

σ C/cm^2	E Volts	C μF/cm^2
-18	-1.6130	28.920
-17	-1.5775	27.573
-16	-1.5405	26.336
-15	-1.5014	24.851
-14	-1.4600	23.432
-13	-1.4159	22.044
-12	-1.3688	20.304
-11	-1.3171	18.402
-10	-1.2597	16.443
- 9	-1.1951	14.603
- 8	-1.1223	12.976
- 7	-1.0418	12.025
- 6	-0.9583	12.197
- 5	-0.8819	14.513
- 4	-0.8225	19.959
- 3	-0.7803	28.175
- 2	-0.7495	37.103
- 1	-0.7253	44.997
0	-0.7045	50.616
1	-0.6853	53.071
2	-0.6666	53.700
3	-0.6479	53.400
4	-0.6289	50.911
5	-0.6086	47.721

Differential capacity of mercury in 0.1 mol l^{-1} NH_4Cl in methanol.

Potential with respect to 0.1 mol l^{-1} NH_4Cl calomel electrode in methanol.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 11.

T/°C	25		0		-30	
	E/volts	C/µF cm^{-2}	E/volts	C/µF cm^{-2}	E/volts	C/µF cm^{-2}
	-1.801	38.93	-1.810	40.05	-1.820	41.30
	-1.763	34.20	-1.776	35.75	-1.788	37.68
	-1.705	30.65	-1.671	30.10	-1.737	34.51
	-1.624	27.31	-1.565	26.44	-1.692	32.51
	-1.532	24.12	-1.477	23.55	-1.642	30.48
	-1.456	21.61	-1.370	19.66	-1.587	28.86
	-1.357	18.45	-1.321	17.99	-1.529	27.25
	-1.251	15.06	-1.226	14.60	-1.478	25.64
	-1.200	13.65	-1.147	12.38	-1.426	23.90
	-1.126	12.01	-1.079	10.99	-1.356	20.98
	-1.052	10.81	-1.004	9.90	-1.278	17.38
	-1.005	10.23	-0.951	9.40	-1.217	14.85
	-0.946	9.68	-0.907	9.10	-1.144	12.29
	-0.895	9.38	-0.863	8.92	-1.082	10.77
	-0.850	9.23	-0.834	8.85	-1.000	9.42
	-0.820	9.17	-0.811	8.82	-0.933	8.76
	-0.797	9.16	-0.761	8.88	-0.871	8.41
	-0.773	9.17	-0.687	9.15	-0.805	8.26
	-0.738	9.24	-0.585	10.09	-0.789	8.26
	-0.726	9.27	-0.540	10.82	-0.688	8.57
	-0.709	9.34	-0.501	11.66	-0.584	9.52
	-0.676	9.49	-0.466	12.90	-0.508	10.83
	-0.643	9.71	-0.433	14.42	-0.437	13.12
	-0.598	10.13	-0.415	15.69	-0.405	14.92
	-0.551	10.76	-0.387	18.24	-0.381	16.81
	-0.517	11.37	-0.365	20.71	-0.337	21.83
	-0.482	12.33	-0.343	23.71	-0.294	28.41
	-0.449	13.70	-0.320	27.19	-0.271	32.06
	-0.407	16.46	-0.305	29.79	-0.232	39.30

Differential capacity of mercury in 0.1 mol l^{-1} NH_4NO_3 in methanol.

Potential with respect to 0.1 mol l^{-1} NH_4Cl calomel electrode in methanol.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 11.

T/°C	25		0		-30	
	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
	-1.587	26.03	-1.617	28.21	-1.642	30.26
	-1.561	25.13	-1.582	27.03	-1.617	29.42
	-1.527	23.97	-1.536	25.44	-1.585	28.44
	-1.462	21.71	-1.510	24.67	-1.521	26.55
	-1.392	19.39	-1.455	22.84	-1.434	23.58
	-1.319	17.14	-1.410	21.23	-1.364	20.59
	-1.243	14.75	-1.381	20.26	-1.317	18.55
	-1.176	13.06	-1.350	19.22	-1.227	14.69
	-1.189	13.37	-1.282	16.54	-1.151	12.10
	-1.149	12.45	-1.202	13.73	-1.061	10.10
	-1.071	11.05	-1.160	12.55	-0.966	8.88
	-1.028	10.47	-1.123	11.63	-0.885	8.35
	-0.970	9.86	-1.083	10.80	-0.796	8.19
	-0.937	9.59	-1.039	10.06	-0.697	8.48
	-0.844	9.30	-0.952	9.08	-0.595	9.32
	-0.842	9.15	-0.873	8.58	-0.499	10.65
	-0.787	9.09	-0.788	8.42	-0.418	12.51
	-0.752	9.12	-0.761	8.42	-0.331	16.46
	-0.702	9.26	-0.649	8.79	-0.273	20.59
	-0.662	9.44	-0.564	9.43	-0.237	23.38
	-0.618	9.70	-0.444	10.99	-0.190	26.39
	-0.572	10.04	-0.347	13.86	-0.147	28.48
	-0.529	10.48	-0.395	12.11	-0.089	30.66
	-0.487	10.95	-0.292	16.98	-0.032	32.57
	-0.445	11.59	-0.232	21.25	0.031	34.57
	-0.399	12.69	-0.173	25.36	0.065	35.57
	-0.366	13.82	-0.128	27.89	0.120	38.01
	-0.326	15.84	-0.098	29.41	0.175	41.47
	-0.295	17.75	-0.064	30.93	0.225	45.53

Capacity in 0.1 mol l^{-1} NH_4NO_3 in methanol (cont.)

$T/^{\circ}C$	25		0		-30	
	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$	E/volts	$C/\mu F\ cm^{-2}$
	-0.254	20.89	-0.023	32.68	0.283	51.47
	-0.221	23.22	0.021	34.26	0.343	63.00
	-0.197	24.72	0.056	35.46	0.312	55.90
	-0.176	26.11	0.102	37.18	0.372	80.78
	-0.137	28.05	0.143	39.13		
	-0.105	29.44	0.169	40.80		
	-0.075	30.73	0.210	44.58		
	-0.040	31.88	0.236	47.05		
	-0.004	33.22	0.283	51.13		
	0.022	34.16	0.315	55.66		
	0.054	35.41	0.353	63.92		
	0.087	36.75	0.376	73.79		
	0.118	38.24	0.391	87.05		
	0.151	40.33				
	0.173	42.22				
	0.180	42.92				
	0.203	46.20				
	0.190	44.21				

Differential capacity of mercury in aqueous 0.1 mol l^{-1} KNO_3 containing thiourea. T = 25°C

Potentials with respect to a saturated calomel electrode in contact with the working solution.

Reference: R. Parsons and P.C. Symons, Trans. Faraday Soc., 64 1077 (1968)

c/mol l^{-1}	0.0		0.00050		0.0010	
σ /μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-20	-1.5517	18.76	-1.5517	18.74	-1.5517	18.79
-19	-1.4975	18.20	-1.4975	18.16	-1.4975	18.22
-18	-1.4418	17.68	-1.4416	17.64	-1.4418	17.72
-17	-1.3844	17.21	-1.3842	17.17	-1.3847	17.30
-16	-1.3256	16.80	-1.3252	16.77	-1.3261	16.88
-15	-1.2655	16.48	-1.2650	16.44	-1.2663	16.57
-14	-1.2043	16.24	-1.2036	16.20	-1.2055	16.33
-13	-1.1424	16.10	-1.1416	16.08	-1.1440	16.22
-12	-1.0803	16.10	-1.0794	16.09	-1.0823	16.20
-11	-1.0184	16.24	-1.0175	16.24	-1.0208	16.36
-10	-0.9573	16.52	-0.9565	16.57	-0.9602	16.70
- 9	-0.8975	16.97	-0.8970	17.03	-0.9013	17.27
- 8	-0.8396	17.58	-0.8396	17.79	-0.8445	18.01
- 7	-0.7838	18.35	-0.7847	18.66	-0.7904	18.98
- 6	-0.7306	19.28	-0.7326	19.79	-0.7394	20.27
- 5	-0.6802	20.38	-0.6837	21.22	-0.6917	21.78
- 4	-0.6326	21.72	-0.6383	22.91	-0.6475	23.54
- 3	-0.5882	23.39	-0.5964	24.94	-0.6068	25.64
- 2	-0.5471	25.39	-0.5581	27.29	-0.5695	28.14
- 1	-0.5092	27.42	-0.5230	29.68	-0.5356	30.92
0	-0.4739	29.14	-0.4905	31.85	-0.5045	33.36
1	-0.4403	30.29	-0.4600	33.59	-0.4755	35.62
2	-0.4077	30.92	-0.4307	34.73	-0.4481	37.08
3	-0.3753	30.81	-0.4022	35.38	-0.4215	38.07
4	-0.3426	30.22	-0.3741	35.58	-0.3954	38.62
5	-0.3089	29.19	-0.3459	35.47	-0.3696	38.72
6	-0.2740	27.98	-0.3177	35.26	-0.3438	38.73
7	-0.2374	26.68	-0.2892	35.08	-0.3179	38.68
8	-0.1991	25.52	-0.2607	35.13	-0.2921	38.76
9	-0.1591	24.57	-0.2323	35.41	-0.2664	39.01
10	-0.1178	23.91	-0.2043	36.01	-0.2409	39.49
11	-0.0757	23.63	-0.1768	36.90	-0.2157	40.17
12	-0.0334	23.77	-0.1502	38.36		

Capacity of Hg in 0.1 mol l^{-1} KNO_3 + thiourea (cont.)

c/mol l^{-1}	0.0020		0.0050		0.010	
σ /μC cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}	E/volts	C/μF cm^{-2}
-20			-1.5517	18.67	-1.5517	18.72
-19			-1.4972	18.10	-1.4975	18.16
-18			-1.4412	17.61	-1.4416	17.65
-17	-1.3844	17.25	-1.3837	17.17	-1.3843	17.22
-16	-1.3258	16.87	-1.3248	16.79	-1.3255	16.86
-15	-1.2659	16.56	-1.2647	16.49	-1.2657	16.58
-14	-1.2050	16.33	-1.2036	16.28	-1.2051	16.43
-13	-1.1436	16.24	-1.1421	16.27	-1.1441	16.44
-12	-1.0820	16.27	-1.0809	16.42	-1.0837	16.67
-11	-1.0209	16.46	-1.0205	16.77	-1.0245	17.20
-10	-0.9608	16.89	-0.9619	17.42	-0.9677	18.10
- 9	-0.9027	17.53	-0.9060	18.39	-0.9142	19.35
- 8	-0.8470	18.45	-0.8534	19.69	-0.8646	21.05
- 7	-0.7944	19.60	-0.8046	21.26	-0.8191	22.93
- 6	-0.7450	21.05	-0.7594	22.97	-0.7772	24.85
- 5	-0.6993	22.79	-0.7175	24.86	-0.7385	26.83
- 4	-0.6572	24.77	-0.6787	26.85	-0.7024	28.70
- 3	-0.6186	27.12	-0.6429	29.07	-0.6687	30.69
- 2	-0.5834	29.87	-0.6100	31.79	-0.6373	33.07
- 1	-0.5515	32.71	-0.5799	34.84	-0.6082	35.73
0	-0.5221	35.45	-0.5523	37.57	-0.5813	38.56
1	-0.4948	37.61	-0.5266	40.15	-0.5562	41.18
2	-0.4688	39.34	-0.5022	41.83	-0.5325	42.93
3	-0.4438	40.48	-0.4787	43.12	-0.5095	44.06
4	-0.4193	41.23	-0.4557	43.82	-0.4869	44.43
5	-0.3952	41.72	-0.4330	44.06	-0.4644	44.26
6	-0.3713	41.80	-0.4102	43.78	-0.4417	43.65
7	-0.3474	41.88	-0.3873	43.48	-0.4185	42.77
8	-0.3235	41.85	-0.3642	42.95	-0.3949	42.08
9	-0.2996	41.89	-0.3408	42.45	-0.3709	41.09
10	-0.2758	41.94	-0.3170	41.87	-0.3463	40.30
11	-0.2520	42.12	-0.2930	41.44	-0.3212	39.60
12	-0.2283	42.44	-0.2688	41.05	-0.2958	39.15
13	-0.2049	42.93	-0.2444	40.97	-0.2702	39.12
14	-0.1818	43.76	-0.2200	41.23	-0.2448	39.62
15	-0.1593	45.08	-0.1960	42.44	-0.2200	40.87
20					-0.1961	43.65

Capacity of Hg in 0.1 mol l^{-1} KNO_3 + thiourea (cont.)

c/mol l^{-1}	0.025		0.050		0.10	
σ /µC cm^{-2}	E/volts	C/µF cm^{-2}	E/volts	C/µF cm^{-2}	E/volts	C/µF cm^{-2}
-20	-1.5517	18.76	-1.5517	18.84	-1.5517	18.88
-19	-1.4976	18.22	-1.4978	18.34	-1.4980	18.40
-18	-1.4419	17.73	-1.4426	17.89	-1.4431	18.03
-17	-1.3849	17.33	-1.3859	17.47	-1.3872	17.77
-16	-1.3266	17.02	-1.3283	17.25	-1.3307	17.67
-15	-1.2675	16.82	-1.2701	17.18	-1.2742	17.79
-14	-1.2080	16.81	-1.2121	17.34	-1.2186	18.24
-13	-1.1488	17.02	-1.1552	17.85	-1.1650	19.14
-12	-1.0908	17.54	-1.1005	18.79	-1.1144	20.52
-11	-1.0351	18.44	-1.0489	20.12	-1.0677	22.34
-10	-0.9827	19.82	-1.0013	21.97	-1.0249	24.44
- 9	-0.9344	21.63	-0.9578	24.08	-0.9857	26.64
- 8	-0.8903	23.74	-0.9181	26.23	-0.9495	28.66
- 7	-0.8499	25.84	-0.8813	28.24	-0.9156	30.26
- 6	-0.8126	27.75	-0.8469	29.77	-0.8832	31.35
- 5	-0.7776	29.44	-0.8140	30.89	-0.8516	31.91
- 4	-0.7444	30.82	-0.7820	31.69	-0.8204	31.98
- 3	-0.7127	32.21	-0.7508	32.35	-0.7892	32.22
- 2	-0.6824	33.92	-0.7203	33.30	-0.7584	32.63
- 1	-0.6538	36.03	-0.6908	34.69	-0.7280	33.22
0	-0.6269	38.22	-0.6627	36.46	-0.6984	34.51
1	-0.6014	40.30	-0.6359	38.19	-0.6700	35.86
2	-0.5771	41.72	-0.6102	39.36	-0.6425	36.83
3	-0.5534	42.58	-0.5850	40.07	-0.6155	37.26
4	-0.5300	42.76	-0.5601	40.02	-0.5887	37.24
5	-0.5066	42.90	-0.5350	39.69	-0.5616	36.63
6	-0.4832	42.29	-0.5097	39.37	-0.5342	36.16
7	-0.4592	40.99	-0.4841	38.54	-0.5064	35.74
8	-0.4345	40.09	-0.4577	37.39	-0.4781	34.78
9	-0.4093	39.41	-0.4307	36.80	-0.4489	33.90
10	-0.3837	38.53	-0.4032	35.68	-0.4191	33.25
11	-0.3574	37.57	-0.3748	35.01	-0.3889	32.93
12	-0.3305	36.99	-0.3462	34.79	-0.3585	33.00
13	-0.3034	36.92	-0.3174	34.94	-0.3284	33.56
14	-0.2765	37.55	-0.2891	35.64	-0.2990	34.59
15	-0.2505	39.57	-0.2616	37.30	-0.2709	36.79

Differential capacity of mercury in KF solutions in 99.8% methanol.
T = 25°C

Potentials with respect to aqueous 1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} NH_4F solution in methanol.

Reference: D.C. Grahame, Z. Elektrochem., 59 (1955) 740.

c/mol l^{-1}	C/uF cm^{-2}			
E/volts	0.001	0.01	0.1	0.446
-1.90	33.77		39.75	
-1.85	31.47	35.00	36.94	
-1.80	29.23	31.34	34.00	
-1.75	27.13	29.18	31.24	33.53
-1.70	25.05	27.40	29.10	31.01
-1.65	23.03	25.50	27.20	29.30
-1.60	21.07	23.50	25.25	27.50
-1.55	19.05	21.45	23.25	25.60
-1.50	17.12	19.48	21.12	23.41
-1.45	15.26	17.25	19.00	21.16
-1.40	13.70	15.38	16.92	18.88
-1.35	12.40	13.80	15.00	16.67
-1.30	11.34	12.43	13.43	14.72
-1.25	10.50	11.40	12.15	13.27
-1.20	9.92	10.60	11.14	12.10
-1.15	9.37	10.00	10.40	11.19
-1.10	8.93	9.50	9.84	10.38
-1.05	8.65	9.15	9.40	9.77
-1.00	8.47	8.88	9.09	9.35
-0.95	8.37	8.72	8.87	9.02
-0.90	8.42	8.62	8.74	8.76
-0.85	8.54	8.62	8.65	8.66
-0.80	8.72	8.74	8.64	8.61
-0.75	8.98	8.93	8.72	8.68
-0.70	9.22	9.21	8.91	8.86
-0.65	9.38	9.48	9.17	9.19
-0.60	9.47	9.68	9.50	9.62
-0.55	9.15	9.76	9.88	10.24
-0.50	8.25	9.60	10.29	11.02
-0.45	6.82	9.03	10.73	11.95
-0.40	4.92	8.30	11.20	13.01
-0.35	3.93	8.07	11.79	14.25
-0.30	4.58	8.88	12.71	15.78
-0.25	6.57	10.75	14.13	17.69
-0.20	9.55	13.05	16.20	20.19
-0.15	11.88	15.60	18.50	23.26
-0.10	14.30	18.65	22.15	26.95
-0.05	17.00	22.05	26.00	31.07
-0.00	20.35	22.40	29.95	35.72
0.05	24.05	29.93	34.50	
0.10	28.10	33.93		
0.15		37.45		
0.20		43.68		

Differential capacity on mercury electrode for KF in water + methanol mixtures with mole fraction of methanol.

$X_{MeOH} = 0.2 + 0.3 M$ KF

Adsorption of thiourea at various molar concentrations - given are the capacity in $\mu F\ cm^{-2}$, potential in V.

Reference electrode: saturated aqueous calomel electrode. T = 25°C.

Reference: S. Minc, J. Jastrzebska and M. Jurkiewicz-Herbich.
J. Electroanal. Chem. 65 (1975) 351.

	$C/\mu F\ cm^{-2}$					
$c/mol\ l^{-1}$	0	0.05	0.08	0.10	0.25	0.50
$-E/V$						
0.30	18.20	47.00	46.60			
0.35	16.55	44.85	41.61	40.45	37.75	39.00
0.40	15.12	43.35	41.00	38.55	34.68	32.80
0.45	13.88	41.25	40.43	38.40	34.20	31.72
0.50	12.72	36.80	38.65	38.28	34.88	31.53
0.55	12.08	31.28	35.42	37.07	35.90	32.02
0.60	11.45	25.90	31.00	33.58	35.78	32.61
0.65	11.15	21.40	25.55	29.08	34.37	33.10
0.70	10.95	17.27	20.80	24.60	31.40	33.42
0.80	11.30	13.62	15.38	18.00	23.08	30.85
0.90	12.00	13.21	14.00	15.35	17.80	24.46
1.00	13.48	14.05	14.60	15.65	16.52	20.13
1.10	15.80	16.20	16.41	17.08	17.70	19.18
1.20	18.60	18.78	19.00	19.40	19.60	20.25
1.30	21.18	21.35	21.50	21.50	21.62	21.91
1.40	22.95	23.00	23.05	23.08	23.10	23.22
1.50	24.20	24.20	24.20	24.22	24.25	24.27
1.60	25.20	25.12	25.10	25.05	25.00	24.88
1.70	26.18	26.03	26.00	25.70	25.45	25.40

Differential capacity on mercury for KF in water + methanol mixtures (continued)

$X_{MeOH} = 0.47 + 0.3\ M\ KF$

Containing thiourea at various molar concentrations.

	$C/\mu F\ cm^{-2}$					
$c/mol\ l^{-1}$	0	0.05	0.08	0.10	0.25	0.50
$-E/V$						
0.30	16.00				56.00	
0.35	14.62	54.50	51.00	48.03	41.62	40.62
0.40	13.44	50.11	49.23	46.16	41.10	34.08
0.45	12.53	41.02	46.18	45.73	42.63	33.61
0.50	11.71	31.00	39.65	42.98	43.81	35.36
0.55	11.05	22.52	31.00	36.82	44.00	38.71
0.60	10.44	16.95	22.52	28.63	39.03	41.93
0.65	10.11	13.82	16.78	21.42	30.06	41.02
0.70	9.85	11.88	13.95	16.33	23.18	35.62
0.80	9.65	10.40	11.12	12.05	14.61	21.41
0.90	10.05	10.42	10.89	11.38	12.17	13.67
1.00	11.00	11.20	11.46	11.86	12.44	13.62
1.10	12.62	12.83	13.02	13.43	14.22	15.08
1.20	15.12	15.41	15.83	16.20	17.03	17.68
1.30	19.08	19.22	19.66	19.73	20.48	20.95
1.40	23.42	23.59	23.65	24.05	24.16	24.70
1.50	26.83	26.85	26.86	26.90	26.95	27.01
1.60	29.60	29.22	29.15	29.05	29.00	28.85

Differential capacity on mercury for KF in water + methanol mixtures (continued)

$X_{MeOH} = 0.83 + 0.3$ KF

containing thiourea at various molar concentrations

	C/μF cm^{-2}				
c/mol l^{-1}	0	0.05	0.10	0.25	0.50
-E/V					
0.30	13.91	72.60			
0.35	12.62	68.26	54.86	46.11	
0.40	11.75	68.75	56.83	42.73	36.57
0.45	11.03	56.11	57.55	44.27	34.48
0.50	10.30	41.05	52.51	49.06	36.61
0.55	9.62	27.92	39.82	51.61	40.92
0.60	9.45	19.26	27.28	44.22	47.53
0.65	9.17	14.53	17.95	32.55	48.01
0.70	8.85	11.77	13.93	22.03	37.54
0.80	8.90	9.80	10.47	12.91	16.93
0.90	9.27	9.78	9.95	11.25	12.52
1.00	10.00	10.21	10.32	11.49	12.71
1.10	11.03	11.24	11.51	12.77	14.32
1.20	12.74	13.30	13.80	15.03	16.96
1.30	15.82	16.52	17.31	18.62	20.01
1.40	20.85	21.12	21.48	22.53	23.26
1.50	26.08	26.00	26.92	25.86	25.80

Differential capacity of mercury in 0.1M $NaClO_4$+ XCH_3OH,given are the capacity in µF cm^{-2} , the potential in mV.

Reference electrode: aqueous 1 mol l^{-1} NaCl calomel electrode.   T = 25°C

Reference: J. Taraszewska. J. Electroanal. Chem. 49 (1974) 443.

	C/µF cm^{-2}							
X	0	0.035	0.0735	0.167	0.4171	0.5680	0.8993	0.9959
-E/mV								
0	24.20	24.90	28.19	26.60	33.80	31.11	34.63	-
25	23.80	24.28	27.10	26.13	32.30	30.65	33.10	-
50	23.53	23.60	26.09	25.90	31.15	30.22	31.56	60.00
75	23.30	23.21	25.22	25.86	30.40	29.76	30.40	51.50
100	23.10	22.80	24.65	25.92	29.75	29.33	29.33	43.53
125	23.10	22.65	24.43	26.05	29.30	28.88	28.10	39.80
150	23.30	22.60	24.30	26.23	28.85	28.36	26.93	33.87
175	23.58	22.67	24.35	26.48	28.00	27.30	25.30	30.72
200	23.95	22.85	24.45	26.60	26.89	26.04	23.73	27.52
225	24.60	23.18	24.68	26.76	25.50	24.50	22.00	24.50
250	25.20	23.50	24.97	26.81	24.05	23.02	20.09	21.75
275	25.56	23.89	25.25	26.52	22.38	21.13	18.35	19.45
300	26.25	24.30	25.40	26.00	20.80	19.29	16.57	17.20
325	26.87	24.62	25.45	25.00	18.90	17.65	15.20	15.45
347.5								14.21
350	27.42	24.88	25.35	23.90	17.40	16.18	13.96	14.00
371						15.03		
375	27.98	24.90	24.95	22.48	15.95	14.82	13.02	13.00
376					15.90			
400	28.30	24.70	24.40	21.25	14.82	13.87	12.10	12.22
421				20.00				
425	28.33	24.15	23.28	19.80	13.85	13.07	11.76	11.65
450	28.10	23.35	22.12	18.40	13.05	12.44	11.29	11.20
463			21.53					
475	27.48	22.40	21.00	17.00	12.45	11.90	10.95	10.78
484		22.09						
500	26.80	21.55	19.90	15.73	11.86	11.47	10.70	10.48
504	26.99							
525	25.80	20.48	18.68	14.80	11.46	11.12	10.48	10.18
550	24.80	19.53	17.55	13.95	11.10	10.80	10.13	9.90
575	23.80	18.72	16.63	13.30	10.73	10.55	9.90	9.70

Capacity in 0.1 M $NaClO_4$ + XCH_3OH (continued)

	$C/\mu F\ cm^{-2}$							
X	0	0.035	0.0735	0.167	0.4171	0.5680	0.8993	0.9959
-E/mV								
600	22.93	18.06	16.00	12.75	10.47	10.31	9.69	9.50
625	22.03	17.55	15.40	12.40	10.22	10.06	9.49	9.32
650	21.30	17.00	14.90	12.02	10.03	9.87	9.33	9.18
675	20.60	16.57	14.51	11.77	9.88	9.68	9.13	9.02
700	20.08	16.20	14.20	11.53	9.69	9.51	8.98	8.92
725	19.60	15.90	14.00	11.43	9.58	9.38	8.88	8.81
750	19.20	15.68	13.86	11.38	9.50	9.29	8.80	8.70
775	18.75	15.50	13.75	11.36	9.40	9.22	8.73	8.61
800	18.40	15.30	13.69	11.38	9.37	9.20	8.71	8.57
825	18.00	15.17	13.70	11.41	9.37	9.20	8.71	8.55
850	17.65	15.05	13.72	11.50	9.40	9.20	8.71	8.52
875	17.32	14.98	13.77	11.61	9.46	9.22	8.71	8.53
900	17.05	14.90	13.82	11.75	9.50	9.29	8.71	8.55
925	16.80	14.88	13.95	11.95	9.62	9.38	8.78	8.60
950	16.60	14.83	14.02	12.18	9.75	9.51	8.89	8.63
975	16.40	14.81	14.18	12.42	9.89	9.63	8.95	8.72
1000	16.25	14.81	14.35	12.70	10.05	9.78	9.07	8.88
1025	16.16	14.80	14.52	13.00	10.25	9.95	9.20	8.98
1050	16.08	14.81	14.70	13.32	10.52	10.18	9.33	9.08
1075	15.99	14.83	14.93	13.71	10.80	10.40	9.48	9.20
1100	15.90	14.90	15.15	14.12	11.03	10.67	9.60	9.33
1125	15.89	15.01	15.31	14.51	11.40	10.94	9.78	9.49
1150	15.85	15.11	15.52	14.98	11.77	11.20	10.00	9.62
1175	15.83	15.25	15.70	15.39	12.15	11.55	10.19	9.80
1200	15.82	15.40	15.90	15.80	12.55	11.91	10.40	10.00
1225	15.89	15.48	16.09	16.23	13.08	12.27	10.62	10.20
1250	15.92	15.58	16.25	16.70	13.60	12.71	10.84	10.40
1275	16.00	15.70	16.43	17.20	14.10	13.15	11.13	10.65
1300	16.08	15.80	16.65	17.70	14.62	13.60	11.42	10.86
1325	16.15	15.90	16.80	18.11	15.30	14.10	11.78	11.18
1350	16.25	16.01	17.00	18.49	16.00	14.71	12.13	11.48
1375	16.35	16.15	17.16	18.91	16.70	15.38	12.43	11.80
1400	16.45	16.25	17.30	19.28	17.38	16.09	12.75	12.18

Differential capacity of mercury in 1 M $NaClO_4$ + X CH_3OH, given are the capacity in μF cm^{-2}, the potential in mV.

Reference electrode: aqueous 1 mol l^{-1} aqueous NaCl calomel electrode.

Reference: J. Taraszewska. J. Electroanal. Chem. 49 (1974) 443.

C/μF cm^{-2} T = 25°C

X	0	0.036	0.076	0.175	0.44	0.59	0.96
-E/mV							
0	24.02	25.20	25.65	26.65	28.25	27.56	29.80
25	23.78	24.82	25.33	26.36	28.00	27.28	29.10
50	23.50	24.50	25.02	26.15	27.80	27.02	28.48
75	23.33	24.28	24.70	26.00	27.75	26.98	28.00
100	23.25	24.13	24.45	25.93	27.72	26.96	27.62
125	23.20	24.05	24.30	25.90	27.72	26.90	27.20
150	23.20	24.00	24.18	25.87	27.72	26.84	26.70
175	23.28	24.01	24.20	25.90	27.68	26.60	26.17
200	23.40	24.05	24.22	25.95	27.52	26.23	25.58
225	23.69	24.18	24.30	26.05	27.30	25.80	24.85
250	24.00	24.38	24.45	26.20	26.92	25.26	24.00
275	24.32	24.62	24.70	26.30	26.32	24.47	22.78
300	24.70	25.00	25.00	26.35	25.69	23.68	21.55
325	25.20	25.32	25.30	26.25	24.82	22.60	20.10
350	25.85	25.70	25.60	26.03	23.82	21.55	18.75
375	26.45	26.13	25.80	25.70	22.65	20.28	17.50
378							17.36
400	27.00	26.60	25.90	25.23	21.40	19.19	16.35
401						19.13	
415					20.56		
425	27.60	26.90	25.83	24.50	20.05	17.90	15.26
450	28.13	27.15	25.70	23.60	18.85	16.70	14.25
464				22.92			
475	28.58	27.30	25.49	22.40	17.60	15.50	13.47
500	28.90	27.40	25.20	21.35	16.32	14.51	12.70
502			25.14				
521		27.23					
525	29.05	27.20	24.52	20.10	15.18	13.57	12.00
537	29.08						
550	29.12	26.90	23.70	18.90	14.22	12.75	11.40
575	29.00	26.45	22.83	17.60	13.30	11.98	10.72

Capacity in 1 M $NaClO_4$ + X CH_3OH (continued)

	$C/\mu F\ cm^{-2}$						
X	0	0.036	0.076	0.175	0.44	0.59	0.96
-E/mV							
600	28.73	25.80	21.98	16.60	12.60	11.41	10.12
625	28.10	24.93	21.00	15.60	12.00	10.83	9.78
650	27.38	24.03	20.07	14.83	11.45	10.50	9.45
675	26.50	23.20	19.13	14.18	11.03	10.17	9.20
700	25.60	22.30	18.22	13.55	10.72	9.96	8.92
725	24.58	21.48	17.52	13.07	10.50	9.73	8.85
750	23.60	20.60	16.90	12.71	10.33	9.59	8.81
775	22.52	19.86	16.47	12.49	10.20	9.45	8.80
800	21.47	19.20	16.12	12.36	10.12	9.35	8.80
825	20.58	18.67	15.92	12.32	10.10	9.38	8.80
850	19.70	18.20	15.72	12.35	10.11	9.41	8.82
875	18.98	17.81	15.60	12.40	10.18	9.46	8.85
900	18.37	17.50	15.55	12.55	10.25	9.53	8.89
925	17.85	17.30	15.58	12.80	10.38	9.62	8.95
950	17.43	17.12	15.68	13.05	10.52	9.77	9.02
975	17.03	16.97	15.80	13.32	10.72	9.90	9.12
1000	16.70	16.90	16.00	13.70	10.95	10.14	9.25
1025	16.50	16.82	16.20	14.10	11.18	10.32	9.40
1050	16.35	16.80	16.41	14.52	11.45	10.56	9.53
1075	16.22	16.78	16.60	14.97	11.80	10.83	9.70
1100	16.13	16.81	16.89	15.45	12.15	11.11	9.85
1125	16.10	16.83	17.05	15.98	12.56	11.48	10.02
1150	16.12	16.90	17.25	16.52	13.02	11.84	10.22
1175	16.16	16.95	17.48	17.10	13.55	12.25	10.43
1200	16.20	17.02	17.70	17.60	14.00	12.63	10.65
1225	16.25	17.18	17.88	18.10	14.65	13.12	10.90
1250	16.31	17.30	18.00	18.60	15.30	13.60	11.18
1275	16.41	17.45	18.22	19.10	15.92	14.14	11.45
1300	16.50	17.60	18.40	19.63	16.53	14.75	11.70
1325	16.62	17.75	18.60	20.02	17.35	15.40	12.10
1350	16.76	17.90	18.78	20.47	18.20	16.09	12.51
1375	16.89	18.00	18.94	20.90	19.03	16.96	12.95
1400	17.02	18.12	19.10	21.32	19.90	17.73	13.35

Differential capacity of mercury in 2 M $NaClO_4$ + X CH_3OH, given are the capacity in µF cm^{-2} , potential in mV.

T = 25°C

Reference electrode: aqueous 1 mol l^{-1} NaCl calomel electrode.

Reference: J. Taraszewska. J. Electroanal. Chem. 49 (1974) 443.

	$C/\mu F\ cm^{-2}$					
X	0	0.038	0.079	0.184	0.49	0.92
-E/mV						
0	23.80	26.96	26.84	27.20	27.81	29.96
25	23.61	26.63	26.40	26.90	27.60	29.32
50	23.40	26.35	25.98	26.59	27.32	28.71
75	23.25	26.01	25.62	26.30	27.10	28.17
100	23.12	25.74	25.32	26.05	26.96	27.64
125	23.02	25.55	25.03	25.83	26.81	27.15
150	22.95	25.38	24.83	25.68	26.72	26.76
175	22.95	25.22	24.62	25.46	26.60	26.23
200	23.03	25.14	24.53	25.38	26.44	25.78
225	23.13	25.05	24.50	25.28	26.27	25.25
250	23.30	25.01	24.53	25.26	26.05	24.62
275	23.55	25.07	24.57	25.24	25.70	23.88
300	23.85	25.26	24.65	25.23	25.38	22.93
325	24.20	25.41	24.78	25.20	24.73	22.10
350	24.60	25.69	24.89	25.20	24.29	21.33
375	25.09	25.93	25.00	25.02	23.52	20.25
400	25.60	26.17	25.10	24.67	22.74	19.20
414						18.58
425	26.13	26.40	25.20	24.26	21.80	18.10
446					20.92	
450	26.70	26.59	25.26	23.80	20.76	17.07
475	27.10	26.79	25.16	23.03	19.63	16.00
489				22.59		
500	27.67	26.96	24.83	22.25	18,52	15.02
525	28.05	27.00	24.44	21.40	17.45	14.10
527			24.39			
549		26.96				
550	28.35	26.90	23.98	20.46	16.33	13.30
556	28.40					
575	28.58	26.75	23.35	19.48	15.32	13.33
600	28.60	26.47	22.71	18.40	14.39	12.53

Differential capacity of 2 M NaClO$_4$ + XCH$_3$OH (continued)

	$C/\mu F\ cm^{-2}$					
X	0	0.038	0.079	0.184	0.49	0.92
-E/mV						
625	28.50	26.00	21.92	17.50	13.57	11.82
650	28.15	25.50	21.13	16.51	12.67	11.25
675	27.70	24.80	20.30	15.69	12.33	10.80
700	27.15	24.16	19.55	14.94	11.75	10.35
725	26.50	23.40	18.83	14.25	11.44	10.04
750	25.83	22.71	18.21	13.84	11.05	9.78
775	24.88	22.04	17.62	13.40	10.70	9.60
800	23.82	21.49	17.18	13.17	10.50	9.45
825	22.83	20.55	16.68	12.98	10.44	9.33
850	21.90	20.28	16.64	12.87	10.44	9.27
875	21.02	19.75	16.48	12.85	10.44	9.24
900	20.25	19.31	16.39	12.93	10.44	9.27
925	19.52	19.03	16.45	13.03	10.52	9.33
950	18.92	18.82	16.51	13.24	10.69	9.40
975	18.35	18.55	16.60	13.60	10.83	9.51
1000	17.85	18.40	16.70	13.80	11.05	9.61
1025	17.55	18.42	16.90	14.18	11.29	9.71
1050	17.28	18.46	17.12	14.51	11.34	9.88
1075	17.03	18.46	17.37	15.00	11.82	10.04
1100	16.88	18.46	17.61	15.48	12.20	10.20
1125	16.75	18.50	17.90	16.07	12.55	10.40
1150	16.70	18.38	18.15	16.64	12.00	10.60
1175	16.70	18.63	18.45	17.25	13.46	10.84
1200	16.70	18.70	18.70	17.70	13.96	11.10
1225	16.75	18.78	18.89	18.43	14.60	11.38
1250	16.80	18.82	19.12	19.06	15.18	11.68
1275	16.85	18.94	19.31	19.67	15.89	12.00
1300	16.98	19.06	19.55	20.28	16.67	12.48
1325	17.08	19.30	19.73	20.91	17.36	12.60
1350	17.18	19.55	19.91	21.55	18.21	13.30
1375	17.30	19.75	20.17	21.97	19.08	13.77
1400	17.42	19.91	20.40	22.28	20.03	14.22

Differential capacity of mercury in 3 M $NaClO_4$ + X CH_3OH, given are the capacity in μF cm^{-2} , the potential in mV.

Reference electrode: aqueous 1 mol l^{-1} NaCl calomel electrode. T = 25°C

Reference: J. Taraszewska. J. Electroanal. Chem. 49 (1974) 443.

	$C/\mu F\ cm^{-2}$					
X	0	0.039	0.083	0.195	0.505	0.90
-E/mV						
0	23.55	25.56	26.85	27.14	27.95	28.58
25	23.38	25.28	26.57	26.87	27.60	28.20
50	23.20	25.01	26.23	26.53	27.26	27.80
75	22.98	24.78	26.00	26.25	26.97	27.17
100	22.77	24.59	25.80	25.99	26.65	27.14
125	22.68	24.38	25.55	25.73	26.40	26.80
150	22.62	24.23	25.35	25.50	26.17	26.44
175	22.60	24.10	25.18	25.30	25.87	26.00
200	22.61	24.04	25.01	25.14	25.32	25.50
225	22.70	24.00	24.85	24.97	25.22	25.00
250	22.88	23.98	24.71	24.83	24.89	24.47
275	23.10	24.00	24.67	24.72	24.50	23.90
300	23.35	24.04	24.65	24.65	24.16	23.35
325	23.62	24.16	24.63	24.50	23.72	22.72
350	23.90	24.29	24.65	24.35	23.25	22.02
375	24.27	24.45	24.75	24.11	22.68	21.22
400	24.65	24.65	24.89	23.86	21.98	20.45
425	25.05	24.88	24.93	23.53	21.30	19.56
450	25.45	25.08	24.95	23.19	20.52	18.82
466					19.96	
475	25.88	25.29	24.58	22.63	19.65	18.82
500	26.26	25.44	24.71	22.04	18.72	17.01
506				21.88		
525	26.72	25.55	24.50	21.38	18.78	16.03
543			24.25			
550	26.12	25.62	24.16	20.64	18.82	15.26
561		25.60				
572	27.34					
575	27.38	25.59	23.68	19.85	15.90	14.36
600	27.50	25.50	23.19	19.06	15.06	13.65

Capacity in 3 M $NaClO_4$ + X CH_3OH (continued)

	C/µF cm^{-2}					
X	0	0.039	0.083	0.195	0.505	0.90
-E/mV						
625	27.60	25.23	22.59	18.08	14.20	12.95
650	27.50	24.89	21.98	17.24	13.48	12.34
675	27.25	24.50	21.36	16.40	12.80	11.80
700	26.92	24.10	20.76	15.79	12.26	11.29
725	26.50	23.60	20.05	15.07	11.73	10.86
750	26.02	23.07	19.43	14.57	11.35	10.52
775	25.35	22.50	18.90	14.07	11.08	10.26
800	24.55	21.98	18.46	13.72	11.81	10.03
825	23.85	21.45	19.05	13.46	10.65	9.87
850	23.10	20.95	17.73	13.30	10.54	9.77
875	22.25	20.48	17.45	13.19	10.49	9.70
900	21.40	20.04	17.30	13.17	10.44	9.67
925	20.67	19.73	17.32	13.23	10.50	9.70
950	20.00	19.49	17.43	13.42	10.60	9.79
975	19.33	19.20	17.52	13.68	10.72	9.88
1000	18.80	19.00	17.67	13.90	10.87	9.98
1025	18.40	18.90	17.83	14.25	11.07	10.12
1050	18.02	18.82	18.03	14.63	11.35	10.26
1075	17.70	18.77	18.30	15.08	11.62	10.45
1100	17.41	18.72	18.58	15.58	11.96	10.62
1125	17.28	18.70	18.87	16.10	12.35	10.85
1150	17.15	18.76	19.19	16.70	12.75	11.10
1175	17.10	18.80	19.50	17.38	13.28	11.34
1200	17.05	18.85	19.75	18.03	13.72	11.60
1225	17.03	18.90	20.00	18.70	14.30	11.93
1250	17.07	19.00	20.25	19.31	14.86	12.28
1275	17.10	19.08	20.50	20.00	15.65	12.68
1300	17.12	19.19	20.70	20.64	16.27	13.05
1325	17.15	19.30	20.92	21.28	17.03	13.60
1350	17.25	19.43	21.10	21.89	17.89	13.99
1375	17.36	19.55	21.30	22.45	18.85	14.57
1400	17.52	19.67	21.46	23.01	19.79	15.15

Differential capacity in μ F/cm^2 on mercury in aqueous ethanol of different concentrations (in mole/l). Salt: aqueous 0.1 M NaF; Temperature 25°C. 400 Hz. Potentials relative to a normal calomel electrode.
Reference: B.B. Damaskin, A.A. Survila, L.E. Rybalka,
Elektrokhimiya, 3, 146 (1967).

conc. of alcohol	0	0.5	1	2	4	8
E / volts						
-1.80	20.15	20.5	21.0	21.2	23.3	26.0
-1.75	-	-	-	-	-	26.9
-1.70	19.05	19.6	20.1	20.9	23.0	27.5
-1.60	18.10	18.8	19.3	20.6	23.6	30.1
-1.55	-	-	-	-	-	31.0
-1.50	17.30	18.0	18.9	20.8	25.1	31.0
-1.45	-	-	-	-	-	28.6
-1.41	-	-	-	-	26.4	-
-1.40	16.60	17.6	18.6	21.2	26.5	25.6
-1.39	-	-	-	-	26.55	-
-1.38	-	-	-	-	26.5	-
-1.35	-	-	18.5	-	26.3	21.6
-1.30	16.05	17.2	18.4	21.4	24.9	-
-1.25	15.80	-	18.3	21.2	22.6	14.9
-1.20	15.65	16.9	18.25	20.7	19.7	-
-1.15	15.60	-	18.1	19.5	16.8	11.2
-1.10	15.70	16.5	17.8	18.3	14.4	-
-1.05	15.80	-	17.4	19.6	12.6	9.25
-1.00	16.15	16.3	16.8	14.9	11.3	-
-0.95	16.45	-	16.1	13.2	10.2	8.1
-0.90	16.85	16.2	15.9	11.5	9.35	-

cont.

Double layer in water ethanol mixtures (cont.)

conc. of alcohol	0	0.5	1	2	4	8
$\frac{E}{\text{volts}}$						
-0.85	17.40	-	14.6	10.8	8.75	7.45
-0.80	18.05	16.1	13.8	10.2	8.35	7.2
-0.75	18.70	-	13.25	9.80	8.1	7.1
-0.70	19.55	16.2	13.0	9.65	7.95	7.1
-0.65	20.00	-	12.95	9.65	8.0	7.1
-0.60	20.40	16.8	13.3	9.8	8.2	7.25
-0.55	20.65	17.3	13.9	10.3	-	-
-0.50	20.80	18.0	14.9	11.1	8.85	7.8
-0.45	21.30	18.9	16.0	12.5	9.45	-
-0.40	22.25	-	17.6	14.3	10.4	8.95
-0.35	23.55	21.7	19.5	16.6	11.6	-
-0.30	24.75	23.7	22.0	19.2	12.9	10.8
-0.25	25.60	25.4	24.5	23.2	15.0	12.1
-0.20	26.15	26.6	26.8	29.0	17.7	14.1
-0.15	26.75	27.9	28.6	32.3	22.9	16.9
-0.10	27.25	28.7	29.9	35.6	31.1	22.0
-0.05	27.85	29.4	30.8	37.0	42.4	31.1
0	28.15	30.2	31.4	37.2	53.3	43.8
0.01	-	-	-	-	53.4	-
0.02	-	-	-	-	53.1	-
0.05	29.25	31.2	32.2	37.0	51.6	52.9
0.07	-	-	-	-	-	43.7
0.10	30.80	32.3	33.1	37.4	46.2	43.4
0.15	33.15	34.6	35.2	38.6	43.6	50.6
0.20	37.3	39.1	40.6	42.8	45.6	89.6
0.25	51.0	56.4	55.6	-	62.3	-

cont.

Differential capacity in μ F/cm^2 on mercury in aqueous n-propanol of different concentrations (in mole/l). Salt: aqueous 0.1 M NaF; 25°C; 400 Hz. Potentials relative to a normal calomel electrode.
Reference: B.B. Damaskin, A.A. Survila, L.E. Rybalka, Elektrokhimiya, 3, 146 (1967).

conc. of alcohol	0.4	0.8	1.2	1.6	2.0
E / volts					
-1.80	20.5	20.9	21.2	21.6	21.8
-1.75	20.0	-	-	21.3	21.6
-1.70	19.5	20.0	20.6	21.2	21.7
-1.65	19.2	-	-	21.3	-
-1.60	19.0	19.9	20.9	21.9	23.1
-1.55	18.8	20.0	-	22.9	-
-1.50	18.8	20.5	22.6	24.8	27.9
-1.45	18.9	-	22.4	28.1	33.8
-1.43	-	-	-	-	36.9
-1.40	19.1	22.6	27.6	33.9	42.1
-1.38	-	-	-	-	43.6
-1.37	-	-	-	-	44.3
-1.36	-	-	-	-	44.0
-1.35	19.6	24.7	32.5	40.5	43.0
-1.34	-	-	-	41.3	-
-1.33	-	-	-	41.7	-
-1.32	-	-	-	41.4	-
-1.31	-	-	36.5	40.7	-
-1.30	20.2	28.1	37.1	-	32.7
-1.29	-	-	37.3	-	-
-1.28	-	-	37.1	35.4	28.0
-1.25	21.3	32.0	34.4	28.2	21.3

cont.

Double layer in water-propanol mixtures (cont.)

conc. of alcohol	0.4	0.8	1.2	1.6	2.0
E / volts					
-1.23	-	32.9	31.5	24.3	-
-1.22	-	33.0	-	-	-
-1.21	-	33.0	-	-	-
-1.20	22.7	32.5	25.5	18.6	14.8
-1.19	-	-	-	-	-
-1.18	23.4	31.1	-	-	-
-1.17	23.8	-	-	-	-
-1.16	24.2	-	-	-	-
-1.15	24.5	27.0	17.9	13.3	10.9
-1.14	24.7	-	-	-	-
-1.13	25.0	24.0	-	-	-
-1.12	25.2	-	-	-	-
-1.11	25.4	-	-	-	-
-1.10	25.5	19.2	13.2	10.4	9.0
-1.09	25.7	-	-	-	-
-1.08	25.4	-	-	-	-
-1.07	25.7	-	-	-	-
-1.05	24.4	14.4	-	-	-
-1.00	20.3	11.1	8.75	7.60	6.85
-0.95	16.1	9.25	-	-	-
-0.90	12.5	8.05	7.00	6.40	5.80
-0.85	10.3	7.25	-	-	-
-0.80	8.85	6.70	6.15	5.80	5.35
-0.75	7.95	6.35	5.90	-	5.20
-0.70	7.40	6.15	5.75	5.50	5.15
-0.65	7.10	6.00	5.65	-	5.05

cont.

Double layer in water-propanol mixtures (cont.).

conc. of alcohol	0.4	0.8	1.2	1.6	2.0
$\frac{E}{\text{volts}}$					
-0.60	7.00	6.00	5.65	5.45	5.05
-0.55	7.15	6.10	5.70	-	5.15
-0.50	7.50	6.30	5.85	5.65	5.25
-0.45	8.15	-	-	-	-
-0.40	9.25	7.20	6.50	6.25	5.65
-0.30	13.85	9.20	7.85	7.50	6.60
-0.25	18.8	-	-	-	-
-0.20	27.3	13.8	10.6	9.90	8.05
-0.15	39.0	18.9	13.2	-	9.35
-0.12	43.5	-	-	-	-
-0.11	44.1	-	-	-	10.6
-0.10	44.0	29.5	17.7	15.1	-
-0.07	-	39.8	-	-	-
-0.05	40.5	-	-	21.0	14.2
-0.03	-	57.6	33.7	-	-
-0.02	-	60.9	-	-	-
-0.01	-	61.9	-	-	-
0	36.9	61.0	52.8	37.4	20.1
0.03	-	-	71.3	-	-
0.04	-	-	72.9	-	-
0.05	34.8	48.4	70.3	72.5	45.5
0.06	-	-	67.6	80.0	-
0.07	-	-	-	81.9	-
0.08	-	-	-	80.1	74.5
0.09	-	-	-	-	87.0
0.10	34.4	40.5	51.4	69.9	87.1
0.11	-	-	-	-	86.4
0.13	-	38.7	-	-	-
0.15	34.8	38.4	43.3	49.5	60.3
0.17	-	-	-	-	52.9
0.18	-	-	-	45.5	-
0.19	-	-	-	44.8	-
0.20	37.4	48.2	42.0	43.8	48.4
0.22	-	-	-	44.3	-
0.23	-	-	-	-	47.9
0.25	46.8	48.1	53.7	48.9	51.1

cont.

Differential capacity in μ F/cm^2 on mercury in aqueous n-butanol of different concentrations (in mole/l). Salt: aqueous 0.1 M NaF; Temp. 25°C; 400 Hz. Potentials relative to a normal calomel electrode.
Reference: B.B. Damaskin, A.A. Survila, L.E. Rybalka,
Elektrokhimiya, 3, 146 (1967).

conc. of alcohol	0.05	0.1	0.2	0.4	0.8
E / volts					
-1.80	20.6	30.3	20.5	20.6	20.0
-1.75	-	19.7	19.8	20.1	20.2
-1.70	19.3	19.2	19.4	19.8	20.4
-1.65	-	18.9	-	19.6	20.7
-1.60	18.3	18.5	18.8	19.6	21.3
-1.55	-	18.2	-	20.0	22.7
-1.50	17.7	18.1	18.8	20.8	26.0
-1.45	-	18.0	19.1	22.4	33.8
-1.43	-	-	-	-	39.5
-1.40	17.4	18.0	19.8	25.3	52.1
-1.39	-	-	-	-	55.2
-1.37	-	-	-	-	59.8
-1.36	-	-	-	-	59.2
-1.35	17.3	-	-	31.0	57.3
-1.33	-	-	-	34.7	50.4
-1.30	17.3	18.9	23.1	41.4	37.1
-1.28	-	-	-	45.1	30.0
-1.27	-	-	-	46.2	-
-1.26	-	-	-	46.8	-
-1.25	17.5	-	26.6	46.2	21.4
-1.22	-	-	-	40.0	-
-1.20	-	21.0	32.0	33.9	14.2
-1.18	-	-	34.8	28.3	-
-1.16	-	-	35.9	-	-
-1.15	18.3	-	36.0	21.2	10.7
-1.14	-	-	35.8	-	-
-1.12	-	-	34.0	-	-
-1.10	-	25.4	30.5	14.5	-
-1.06	-	27.0	-	-	-
-1.05	19.5	27.3	21.4	10.9	7.60
-1.04	-	27.5	-	-	-
-1.03	-	27.4	-	-	-
-1.02	-	27.2	-	-	-
-1.00	20.4	-	15.0	-	-
-0.95	21.1	21.8	11.4	7.75	6.30
-0.93	21.2	-	-	-	-
-0.92	21.0	-	-	-	-
-0.90	20.8	16.4	9.25	-	-
-0.85	19.0	12.6	-	6.35	5.65
-0.80	16.5	10.2	7.05	-	-
-0.75	13.8	8.65	-	5.60	5.24
-0.70	11.7	-	6.05	-	-
-0.65	-	7.10	-	5.25	5.05
-0.60	9.40	6.75	5.60	5.15	4.95

cont.

Double layer in water-n-butanol mixtures (cont.)

conc. of alcohol	0.05	0.1	0.2	0.4	0.8
$\frac{E}{\text{volts}}$					
-0.55	9.15	6.65	5.60	5.10	5.00
-0.50	9.50	6.75	5.65	5.15	-
-0.45	-	-	-	5.30	5.20
-0.40	12.6	7.80	6.15	-	-
-0.35	-	-	-	5.90	5.80
-0.30	22.2	11.2	7.60	-	-
-0.25	29.4	15.2	-	7.20	6.85
-0.20	33.7	23.8	11.4	-	-
-0.19	33.9	-	-	-	-
-0.18	34.2	-	-	-	-
-0.17	34.0	-	-	-	-
-0.15	33.7	41.2	15.5	10.1	8.65
-0.13	33.2	49.0	-	-	-
-0.12	-	51.5	-	-	-
-0.11	-	52.5	-	-	-
-0.10	32.4	52.1	25.5	12.9	-
-0.09	-	50.9	-	-	-
-0.07	31.8	46.3	38.3	-	-
-0.05	-	57.9	55.3	18.3	11.3
-0.04	31.4	-	66.2	-	-
-0.03	-	-	73.9	-	-
-0.02	-	-	80.4	-	-
-0.015	-	-	80.7	-	-
-0.01	-	-	79.9	-	-
0	31.2	36.1	74.3	33.8	14.2
0.01	-	-	66.9	-	-
0.02	-	-	60.2	-	-
0.03	-	-	-	61.7	-
0.05	31.1	33.8	46.0	96.4	20.5
0.06	-	-	-	109.9	-
0.07	-	-	-	113.9	-
0.08	-	-	-	105.8	29.0
0.09	-	-	-	88.0	-
0.10	31.3	33.1	38.3	72.7	44.2
0.12	-	-	-	-	69.8
0.13	-	-	-	52.4	94.8
0.14	-	-	-	-	129.4
0.15	32.4	33.8	37.0	46.3	155.8
0.16	-	-	-	-	156.3
0.17	-	-	-	43.8	122.8
0.18	-	-	-	43.0	93.8
0.19	-	-	-	-	75.2
0.20	34.4	36.5	38.3	42.7	63.3
0.22	-	-	-	-	54.0
0.23	-	-	-	45.4	-
0.25	41.7	45.1	46.5	50.1	54.5

cont.

Differential capacity in μF/cm^2 on mercury in aqueous n-pentanol of different concentrations (in mole/l). Salt: aqueous 0.1 M NaF; Temp. 25°C. 400 Hz. Potentials relative to a normal calomel electrode. Reference: B.B. Damaskin, A.A. Survila, L.E. Rybalka, Elektrokhimiya, 3, 146 (1967).

conc. of alcohol	0.025	0.05	0.10	0.15	0.2
E / volts					
-1.80	20.4	20.3	20.5	20.6	20.5
-1.75	-	-	-	-	19.9
-1.70	19.2	19.1	19.4	19.5	19.5
-1.65	-	-	-	-	19.3
-1.60	18.3	18.5	18.7	19.0	19.2
-1.50	17.7	18.0	18.8	19.5	20.5
-1.45	-	18.0	-	-	22.7
-1.40	17.4	18.3	20.7	23.6	28.8
-1.37	-	-	-	-	37.8
-1.35	-	-	23.4	31.2	51.7
-1.33	-	-	-	-	72.9
-1.32	-	-	-	43.1	81.5
-1.31	-	-	-	-	85.4
-1.30	17.9	20.3	30.5	56.2	79.7
-1.29	-	-	-	-	69.6
-1.28	-	-	-	68.9	-
-1.27	-	-	-	72.3	49.4
-1.265	-	-	-	72.4	-
-1.26	-	-	-	70.6	-
-1.25	-	-	47.7	65.6	33.7
-1.24	-	-	52.6	-	-
-1.23	-	-	57.6	48.2	24.6
-1.22	-	-	60.3	-	-
-1.21	-	-	60.5	-	-
-1.20	19.8	28.8	57.2	28.4	16.8
-1.19	-	-	53.0	-	-
-1.18	-	-	46.3	-	-
-1.17	-	-	-	-	-
-1.16	-	-	-	-	-
-1.15	-	39.5	29.7	15.4	10.9
-1.14	-	42.1	-	-	-
-1.13	-	44.0	-	-	-
-1.12	-	45.2	-	-	-
-1.11	-	44.9	-	-	-
-1.10	25.6	42.8	16.3	10.5	-
-1.08	-	37.0	-	-	-
-1.05	31.3	27.1	11.1	-	7.15
-1.04	32.3	-	-	-	-
-1.03	33.1	-	-	-	-
-1.02	33.2	-	-	-	-
-1.015	33.5	-	-	-	-
-1.01	33.2	-	-	-	-
-1.00	32.5	16.4	-	7.05	-
-0.99	31.5	-	-	-	-
-0.95	24.1	-	7.25	-	5.80

cont.

Double layer in water n-pentanol mixtures (cont.).

conc. of alcohol	0.025	0.05	0.10	0.15	0.20
E / volts					
-0.90	15.9	8.85	-	5.75	-
-0.85	-	-	5.85	-	5.15
-0.80	8.95	6.55	-	5.10	-
-0.75	-	-	5.10	-	4.75
-0.70	6.75	5.55	-	4.70	-
-0.65	-	-	4.75	-	4.55
-0.60	5.90	5.15	4.65	4.50	4.50
-0.55	5.70	5.05	4.60	-	4.45
-0.50	5.70	5.00	4.60	4.45	4.50
-0.45	5.80	5.05	-	-	-
-0.40	-	-	4.75	4.65	4.75
-0.35	6.80	5.55	-	-	-
-0.30	-	-	5.30	5.15	5.30
-0.25	10.15	6.95	-	-	-
-0.20	14.7	-	6.55	6.20	6.40
-0.15	28.3	11.3	7.75	-	-
-0.12	48.8	-	-	-	-
-0.11	59.0	-	-	-	-
-0.10	68.1	16.9	9.55	8.30	7.95
-0.09	70.1	-	-	-	-
-0.08	67.7	-	-	-	-
-0.07	60.3	26.1	-	-	-
-0.05	47.7	41.6	13.0	-	-
-0.04	-	53.7	-	-	-
-0.03	41.0	73.1	-	-	-
-0.015	-	108.4	-	-	-
-0.01	-	121.4	-	-	-
0	36.1	113.3	22.1	13.5	10.9
0.01	-	83.4	-	-	-
0.02	-	66.5	-	-	-
0.03	-	34.7	41.5	-	-
0.04	-	-	57.9	-	-
0.05	33.0	43.9	86.8	23.9	15.2
0.06	-	-	143.0	-	-
0.07	-	-	180.4	37.8	-
0.08	-	-	142.5	51.0	-
0.09	-	-	89.6	75.3	-
0.10	32.5	35.7	65.9	134.3	30.7
0.11	-	-	53.8	250.5	-
0.115	-	-	-	231.4	51.4
0.12	-	-	-	193.5	-
0.125	-	-	-	-	84.7
0.13	32.9	-	44.0	102.9	137.8
0.135	-	-	-	-	234.4
0.14	-	-	-	71.6	330.0
0.145	-	-	-	-	289.0
0.15	-	34.6	40.4	54.3	208.9
0.155	-	-	-	-	130.1
0.16	-	-	-	-	94.2
0.17	34.1	-	39.1	45.0	66.6
0.18	-	-	-	-	53.7
0.19	-	-	-	-	51.2
0.20	35.7	36.4	39.2	41.4	46.2
0.22	-	-	-	-	44.5
0.25	44.0	43.6	46.3	47.0	48.8

cont.

Differential capacity in μ F/cm^2 on mercury in aqueous n-hexanol of different concentrations (in mole/l). Salt: aqueous 0.1 M NaF; Temp. 25°C; 400 Hz. Potentials relative to a normal calomel electrode.
Reference: B.B. Damaskin, A.A. Survila, L.E. Rybalka,
Elektrokhimiya, 3, 146 (1967).

conc. of alcohol E / volts	0.002	0.004	0.008	0.016	0.03
-1.80	20.3	20.7	20.3	20.3	20.3
-1.75	-	-	-	19.7	-
-1.70	19.1	19.4	19.0	19.2	19.2
-1.60	18.1	18.3	18.1	18.3	18.4
-1.50	17.4	17.5	17.5	17.7	18.0
-1.40	16.8	16.9	17.1	17.7	18.9
-1.30	16.4	16.7	17.4	19.3	25.3
-1.27	-	-	-	-	32.2
-1.25	-	-	-	22.1	44.5
-1.24	-	-	-	-	48.7
-1.23	-	-	-	-	58.6
-1.22	-	-	-	29.4	67.7
-1.21	-	-	-	-	69.8
-1.20	16.4	17.3	19.2	-	62.5
-1.19	-	-	-	38.0	51.0
-1.18	-	-	-	40.8	39.8
-1.17	-	-	-	44.7	-
-1.16	-	-	-	47.5	-
-1.15	-	-	-	46.9	21.3
-1.14	-	-	-	43.1	-
-1.13	-	-	-	39.0	-
-1.00	17.0	19.3	28.0	-	11.2
-1.06	-	-	34.4	-	-
-1.05	-	21.1	34.5	-	8.0
-1.04	-	-	35.3	-	-
-1.03	-	-	34.6	-	-
-1.02	-	22.5	-	-	-
-1.00	18.1	23.5	28.1	10.4	6.50
-0.98	-	24.4	23.6	-	-
-0.97	18.6	24.8	-	-	-
-0.96	-	24.6	-	-	-
-0.95	18.7	-	17.2	9.20	-
-0.93	19.0	22.8	-	-	-
-0.92	19.1	-	-	-	-
-0.90	19.0	19.8	11.5	-	5.15
-0.85	18.0	14.2	-	5.85	-
-0.80	-	10.5	7.35	-	4.55
-0.75	12.8	-	-	5.00	-
-0.70	-	7.30	5.85	-	4.25
-0.65	9.15	-	-	4.55	-
-0.60	-	6.10	5.25	4.50	4.10
-0.55	7.75	-	5.10	4.45	-
-0.50	7.75	5.75	4.90	4.40	4.10

cont.

Double layer in water n-hexanol mixtures (cont.)

conc. of alcohol E volts	0.002	0.004	0.008	0.016	0.03
-0.40	9.00	6.20	5.30	4.55	4.20
-0.30	15.0	8.10	6.15	5.00	4.55
-0.25	22.0	-	-	-	-
-0.20	29.1	16.5	8.80	6.10	5.25
-0.17	30.3	24.3	-	-	-
-0.16	30.5	-	-	-	-
-0.15	30.3	33.1	12.4	-	-
-0.13	-	42.7	-	-	-
-0.12	-	44.0	-	-	-
-0.11	-	43.0	-	-	-
-0.10	29.7	40.7	25.0	9.30	6.70
-0.08	-	37.1	38.5	-	-
-0.07	-	-	52.0	-	-
-0.06	-	-	68.7	-	-
-0.05	29.5	33.7	78.0	14.2	-
-0.04	-	-	73.9	-	-
-0.03	-	-	60.8	18.5	-
-0.02	-	-	50.8	-	-
-0.01	-	-	-	29.2	-
0	29.6	31.2	39.8	41.2	11.3
0.01	-	-	-	65.7	-
0.02	-	-	35.8	108.6	-
0.03	-	-	-	119.9	-
0.04	-	-	-	76.8	-
0.05	29.8	30.5	33.2	57.1	23.1
0.06	-	-	-	46.9	-
0.07	-	-	-	-	48.2
0.075	-	-	-	-	63.6
0.08	-	-	-	-	104.2
0.085	-	-	-	-	165.5
0.09	-	-	-	-	212.7
0.095	-	-	-	-	166.6
0.10	30.6	31.0	32.2	36.2	104.8
0.105	-	-	-	-	78.2
0.11	-	-	-	-	62.5
0.12	-	-	-	-	47.5
0.13	-	-	-	-	43.1
0.15	32.1	32.5	33.1	34.5	38.5
0.17	-	-	-	-	37.3
0.20	35.2	35.2	35.6	37.0	37.9
0.25	43.4	42.7	42.8	44.1	43.7

Differential capacity on mercury electrode in 0.916 NaF - adsorption of 5-chloro-1 pentanol.

Given are the capacity in μF cm^{-2}, measured at 400 Hz and extrapolated to 0 Hz.

Potential in V vs Orion fluoride reversible electrode. $T = 25\ ^{o}C$.

Reference: K. Doblhofer and D.M. Mohilner. J. Phys. Chem. 75 (1971) 1968.

	$C/\mu F\ cm^{-2}$													
$c/mol\ l^{-1}$	0.0433		0.03436		0.02590		0.01290		0.004865		0.002662		0.001716	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-1.500	19.97	19.97	19.87	19.87	19.92	19.92	19.76	19.76	19.74	19.74	19.50	19.50	19.68	19.68
-1.475	19.79	19.79												
-1.450	19.70	19.70	19.47	19.53	19.55	19.55	19.33	19.33	19.23	19.23	19.05	19.05	19.17	19.17
-1.425	19.55	19.55												
-1.400	19.57	19.57	19.30	19.33	19.29	19.29	18.93	18.93	18.83	18.83	18.61	18.61	18.67	18.67
-1.375	19.60	19.60												
-1.350	19.77	19.77	19.28	19.28	19.17	19.17	18.61	18.61	18.39	19.39	18.21	18.21	18.28	18.28
-1.325	20.10	20.10												
-1.300	20.60	20.60	19.71	19.71	19.33	19.33	18.43	18.43	18.08	18.08	17.84	17.84	17.89	17.89
-1.275	21.47	21.47												
-1.250	22.97	23.00	21.14	21.30	20.08	20.10	18.53	18.53	17.84	17.84	17.60	17.60	17.55	17.55
-1.225	25.60	25.70												
-1.200	30.46	30.70	25.16	25.60	22.26	22.70	18.90	18.90	17.69	17.69	17.42	17.42	17.33	17.33
-1.175	40.83	45.50			24.53	25.53								
-1.16	50.25	59.70												
-1.15	55.39	64.20	36.85	38.45	28.68	31.00	19.81	19.85	17.72	17.72	17.45	17.50	17.09	17.12
-1.14	59.14	68.6												
-1.13	60.69	69.80	45.38	49.20										
-1.125					35.05	40.00								
-1.12	59.10	69.00												
-1.11	52.06	60.60	55.68	62.00										

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\,cm^{-2}$

c/mol l^{-1}	0.0433		0.03436		0.02590		0.01290		0.004865		0.002662		0.001716	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-1.10	45.65	52.80			44.02	52.00	21.95	22.26	18.01	18.10	17.51	17.70	17.03	17.08
-1.09	39.10	44.80	50.89	59.50	47.90	57.80								
-1.08	32.78	37.2			50.15	60.60								
-1.075					50.57	61.00								
-1.07	27.50	30.10	40.79	48.20	50.88	61.80	24.26	24.86						
-1.065					49.82	60.00								
-1.06	23.68	25.80			48.85	58.80								
-1.05	20.76	22.30	32.54	33.35	44.59	52.00			18.65	18.92	17.87	18.20	17.03	17.12
-1.04	18.57	19.70			40.82	48.00	28.03	29.58						
-1.03	16.69	17.60	21.62	23.45	34.69	39.80								
-1.025														
-1.02	15.13	15.80					31.66	36.52						
-1.015					27.54	31.20								
-1.01							33.57	38.30						
-1.00	13.11	13.50	15.91	16.85	22.78	24.80	35.32	40.38	19.68	19.98	18.44	19.00	17.11	17.35
-0.99							36.64	42.93						
-0.98	11.74	12.00					37.26	44.20						
-0.975					17.20	18.20								
-0.97							36.84	43.90	20.85	21.84				
-0.96	10.79	10.90					35.36	42.00						
-0.95			11.67	11.95	13.97	14.60	33.04	39.10			19.17	20.25	17.52	17.92
-0.94	10.13	10.18					30.28	35.65	22.10	23.25				
-0.93							27.20	30.92						

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\ cm^{-2}$

$c/mol\ l^{-1}$	0.0433		0.03436		0.02590		0.01290		0.004865		0.002662		0.001716	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-0.925					12.06	12.30								
-0.92	9.59	9.50					24.46	27.40						
-0.91							22.07	24.25	23.27	25.59				
-0.90	9.21	9.21	9.84	9.94	10.91	11.10	19.94	21.75	23.91	26.78	20.10	22.00	17.92	18.55
-0.89									24.05	27.32				
-0.88							16.66	17.75	24.21	27.74				
-0.875	8.91	8.91			10.17	10.31								
-0.87									24.48	28.06				
-0.86							14.44	15.17						
-0.85	8.67	8.67	9.03	9.05	9.61	9.68			23.55	26.85	20.48	23.40	18.26	19.14
-0.84							12.87	13.25						
-0.825					9.23	9.30			21.75	24.61				
-0.82							11.85	12.08						
-0.80	8.29	8.29	8.55	8.55	8.94	9.00	10.99	11.02	19.48	21.84	19.19	21.68	18.50	19.55
-0.775							10.31	10.35	17.10	19.03				
-0.77											17.75	19.95		
-0.76														
-0.75	8.05	8.05	8.26	8.26	8.60	8.62	9.82	9.85	15.10	16.48				
-0.73											15.45	17.25	18.33	19.70
-0.725									13.31	14.23				
-0.70	7.85	7.85	8.07	8.07	8.35	8.36	9.19	9.23	12.14	12.67	13.87	15.26	17.29	18.90
-0.68											13.13	14.02		
-0.675									11.21	11.47				

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\,cm^{-2}$

$c/mol\,l^{-1}$	0.0433		0.03436		0.02590		0.01290		0.004865		0.002662		0.001716	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-0.66											12.38	13.15		
-0.65	7.73	7.73	7.93	7.93	8.18	8.18	8.83	8.86	10.55	10.68			15.45	16.75
-0.64											11.80	12.40		
-0.625									10.05	10.11				
-0.62											11.31	11.65		
-0.60	7.61	7.61	7.83	7.84	8.08	8.08	8.63	8.63	9.74	9.80	11.01	11.31	13.48	14.45
-0.575									9.46	9.50				
-0.55	7.53	7.53	7.74	7.74	8.01	8.01	8.42	8.42	9.26	9.28	10.10	10.30	12.04	12.55
-0.50	7.49	7.49	7.74	7.74	7.96	7.96	8.38	8.38	8.95	8.95	9.61	9.70	11.00	11.32
-0.45	7.45	7.45	7.66	7.66	7.93	7.93	8.32	8.32	8.76	8.79	9.34	9.35	10.38	10.47
-0.40	7.41	7.41	7.63	7.63	7.89	7.89	8.29	8.29	8.69	8.69	9.25	9.25	10.06	10.06
-0.35	7.43	7.43	7.65	7.66	7.88	7.88	8.29	8.29	8.66	8.66	9.19	9.19	9.97	9.97
-0.30	7.49	7.49	7.68	7.68	7.93	7.93	8.35	8.35	8.71	8.71	9.35	9.35	10.12	10.12
-0.25	7.59	7.59	7.73	7.73	8.01	8.01	8.40	8.44	8.83	8.83	9.65	9.65	10.48	10.49
-0.20	7.72	7.72	7.85	7.85	8.10	8.10	8.58	8.58	9.08	9.11	10.22	10.22	11.24	11.25
-0.15	7.87	7.87	7.92	7.92	8.22	8.22	8.79	8.79	9.45	9.51	11.22	11.22	12.72	12.92
-0.125													14.06	14.45
-0.12											12.30	12.50		
-0.10	8.04	8.04	8.09	8.10	8.41	8.41	9.08	9.08	10.07	10.20	13.45	13.97	15.98	16.80
-0.08											14.82	15.72		
-0.075													18.40	20.05
-0.06											16.89	18.12		
-0.05	8.25	8.25	8.32	8.33	8.63	8.64	9.55	9.55	11.24	11.57			22.38	24.88

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\ cm^{-2}$

$c/mol\ l^{-1}$	0.0433		0.03436		0.02590		0.01290		0.004865		0.002662		0.001716	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-0.04											19.92	22.50		
-0.025													27.10	30.35
-0.02									12.98	14.65	24.37	28.38		
0.00	8.50	8.50			8.95	8.95	10.19	10.21	14.77	18.14	30.42	36.00	32.47	35.17
0.02									17.12	22.11	35.99	42.50		
0.025													31.48	36.50
0.04									20.27	27.07	37.88	43.20		
0.05	8.95	9.00	9.12	9.16	9.45	9.47	11.34	11.34					33.52	35.33
0.06									24.76	33.36	38.38	42.20		
0.075					9.83	9.87	12.26	12.27					33.64	35.15
0.08									32.12	42.50	37.04	39.20		
0.10	9.67	9.80	9.89	9.98	10.31	10.37	13.86	13.98	44.72	56.40	34.93	35.10	32.98	33.90
0.11									47.88	59.40				
0.12							15.80	15.95	50.98	62.00				
0.125					10.99	11.10							32.53	33.15
0.13									51.56	62.00				
0.14							19.39	20.45	50.41	57.76				
0.15	11.00	11.20	11.49	11.65	11.87	12.10	22.20	23.30	46.55	53.08	33.27	33.30	32.03	32.20
0.16							25.88	27.20	43.60	48.05				
0.17							32.62	37.45	41.00	43.40				
0.175	12.23	12.50	12.55	13.10	13.33	13.52								
0.18							43.25	52.00						
0.19							62.58	82.80	37.80	38.25				

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\ cm^{-2}$

$c/mol\ l^{-1}$	0.0433		0.03436		0.02590		0.01290		0.004865		0.002662		0.001716	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
0.20	14.22	14.70	15.54	15.89	15.85	16.30	86.95	120.00	36.69	36.96	32.70	32.70	31.92	32.15
0.21							107.72	158.00						
0.22	16.55	17.20					84.91	101.00						
0.225			20.41	21.10	20.91	21.70			35.27	35.40				
0.23							74.05	87.50						
0.24	20.35	21.20			26.50	28.50	63.25	69.30						
0.25	22.61	23.70	28.48	30.55			55.14	57.80	34.75	34.80	33.40	33.40	32.62	32.80
0.26	25.79	27.10			50.98	59.30	47.10	49.10						
0.27	28.31	30.80			77.14 (1000 Hz)									
0.275			45.60	48.50	112.02 (1000 Hz)									
0.28	33.61	35.30			142.19 (1000 Hz)									
0.285					132.32 (1000 Hz)									
0.29	40.61	42.50			114.30 (1000 Hz)									
0.30	52.43	59.00	118.94 (1000 Hz)		83.59 (1000 Hz)		44.95 (1000 Hz)		35.48	35.61	35.42	35.40	34.59	35.60
0.31	70.89	82.50												
0.32	131.90	166.00	108.55 (1000 Hz)		54.37 (1000 Hz)		46.21 (1000 Hz)		36.31	36.44				
0.33	191.90	252.00												
0.34					47.81 (1000 Hz)				38.01	38.14				
0.35							57.28 (1000 Hz)		39.50	39.57	46.00	46.00	41.75	43.80
0.36					48.98 (1000 Hz)									
0.38					58.61 (1000 Hz)									

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

C/μF cm^{-2}

c/mol l^{-1}	0.001322		0.001134		0.001016		0.0009144	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-1.50	19.63	19.63	19.60	19.60	19.62	19.62	19.62	19.62
-1.40	18.65	18.65	18.60	18.60	18.63	18.63	18.63	18.63
-1.30	17.82	17.82	17.79	17.79	17.80	17.80	17.76	17.76
-1.20	17.20	17.20	17.11	17.11	17.14	17.14	17.07	17.07
-1.10	16.79	16.79	16.69	16.69	16.69	16.69	16.65	16.65
-1.00	16.79	16.80	16.62	16.62	16.58	16.58	16.46	16.46
-0.95	16.95	16.95	16.72	16.73	16.66	16.66	16.53	16.53
-0.90	17.20	17.40	16.91	17.01	16.83	16.91	16.68	16.69
-0.85	17.52	17.92	17.19	17.43	17.06	17.27	16.88	17.01
-0.80	17.80	18.40	17.43	17.87	17.29	17.63	17.10	17.36
-0.75	17.92	18.80	17.55	18.20	17.46	18.00	17.27	17.66
-0.70	17.61	18.60	17.39	18.17	17.40	18.15	17.33	17.96
-0.65	16.66	17.67	16.79	17.54	16.97	17.72	17.10	17.70
-0.60	15.20	16.00	15.70	16.44	16.16	16.85	16.50	17.04
-0.55	13.64	14.30	14.35	14.92	15.00	15.55	15.69	16.09
-0.50	12.37	12.76	13.06	13.44	13.83	14.15	14.75	15.15
-0.45	11.49	11.62	12.15	12.30	12.98	13.15	13.88	14.13
-0.40	11.01	11.11	11.58	11.62	12.40	12.45	13.27	13.29
-0.35	10.81	10.82	11.38	11.38	12.23	12.23	13.17	13.19
-0.30	10.98	10.98	11.59	11.59	12.49	12.51	13.54	13.57
-0.25	11.48	11.50	12.22	12.24	13.31	13.35	14.61	14.63
-0.20	12.61	12.68	13.51	13.51	14.79	14.86	16.35	16.40
-0.15	14.74	15.06	15.96	16.22	17.95	18.32	19.60	20.02
-0.10	18.84	19.80	20.53	21.46	22.59	23.48	24.15	24.90
-0.05	25.50	27.30	26.49	28.03	27.56	28.90	28.04	29.30
0.00	30.80	33.00	30.36	31.96	30.41	31.75	30.12	31.24
0.05	32.17	33.50	31.41	32.37	31.13	31.95	30.68	31.32
0.10	31.67	31.90	31.13	31.49	30.94	31.22	30.54	30.72
0.15	31.18	31.24	30.79	30.79	30.76	30.79	30.46	30.48
0.20	31.11	31.11	30.93	30.93	30.94	30.95	30.73	30.73
0.25	31.76	31.79	31.71	31.75	31.74	31.74	31.65	31.65
0.30	33.11	34.01	33.52	34.02	33.39	34.09	33.42	33.54
0.35	34.27	36.27	38.34	39.46	38.40	39.75	39.05	40.85

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\ cm^{-2}$

$c/mol\ l^{-1}$	0.0008580		0.0007518		0.0006974		0.0006610	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-1.50	19.74	19.74	19.62	19.62	19.56	19.56	19.63	19.63
-1.40	18.73	18.73	18.61	18.61	18.57	18.57	18.63	18.63
-1.30	17.89	17.89	17.76	17.76	17.73	17.73	17.80	17.80
-1.20	17.23	17.23	17.10	17.10	17.05	17.05	17.13	17.13
-1.10	16.78	16.78	16.62	16.62	16.59	16.59	16.64	16.64
-1.00	16.57	16.62	16.42	16.42	16.37	16.37	16.39	16.39
-0.95	16.59	16.74	16.43	16.44	16.37	16.37	16.42	16.43
-0.90	16.79	16.95	16.52	16.54	16.49	16.49	16.49	16.51
-0.85	16.95	17.17	16.69	16.77	16.64	16.77	16.64	16.70
-0.80	17.20	17.48	16.88	17.06	16.82	17.01	16.87	17.04
-0.75	17.37	17.72	17.08	17.42	17.04	17.38	17.01	17.18
-0.70	17.40	17.82	17.19	17.59	17.19	17.61	17.22	17.48
-0.65	17.31	17.80	17.15	17.60	17.20	17.65	17.36	17.73
-0.60	17.00	17.65	16.80	17.50	17.13	17.71	17.33	17.71
-0.55	16.30	16.73	16.47	16.78	16.81	17.24	17.13	17.42
-0.50	15.53	15.83	15.91	16.13	16.35	16.67	16.87	17.05
-0.45	14.84	15.00	15.28	15.29	15.95	16.14	16.57	16.60
-0.40	14.36	14.40	14.93	14.93	15.68	15.71	16.43	16.44
-0.35	14.36	14.40	14.95	14.98	15.81	15.81	16.68	16.70
-0.30	14.92	14.92	15.53	15.53	16.54	16.55	17.47	17.47
-0.25	16.26	16.30	16.93	17.00	17.92	17.98	19.04	19.04
-0.20	18.71	18.88	19.06	19.06	20.12	20.25	21.23	21.38
-0.15	21.94	22.35	22.29	22.58	23.15	23.55	24.12	24.47
-0.10	25.90	26.70	25.83	26.52	26.25	26.92	26.80	27.33
-0.05	28.24	29.12	28.44	29.26	28.64	29.54	28.81	29.60
0.00	30.41	31.35	29.88	30.60	29.81	30.57	29.81	30.35
0.05	30.80	31.30	30.30	30.72	30.14	30.59	30.14	30.50
0.10	30.60	31.00	30.25	30.28	30.10	30.30	30.14	30.30
0.15	30.77	30.77	30.36	30.41	30.14	30.14	30.25	30.26
0.20	31.17	31.17	30.82	30.96	30.60	30.60	30.68	30.68
0.25	32.16	32.26	31.83	32.13	31.49	31.80	31.65	31.69
0.30	34.34	34.55	33.94	35.54	33.42	33.40	33.59	33.70
0.35	41.55	44.50	41.44	44.45	39.45	41.75	39.71	41.70

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\,cm^{-2}$

$c/mol\,l^{-1}$	0.0005670		0.0005008		0.0004572		0.0004290	
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.
-1.50	19.56	19.56	19.62	19.62	19.54	19.54	19.74	19.74
-1.40	18.57	18.57	18.55	18.55	18.55	18.55	18.75	18.75
-1.30	17.73	17.73	17.77	17.77	17.73	17.73	17.90	17.90
-1.20	17.04	17.04	17.07	17.07	17.02	17.02	17.21	17.21
-1.10	16.54	16.54	16.54	16.54	16.50	16.50	16.69	16.69
-1.00	16.28	16.28	16.29	16.29	16.21	16.21	16.39	16.39
-0.95	16.28	16.28	16.28	16.30	16.17	16.20	16.38	16.38
-0.90	16.36	16.36	16.31	16.31	16.20	16.20	16.38	16.38
-0.85	16.49	16.52	16.44	16.49	16.33	16.36	16.51	16.54
-0.80	16.66	16.77	16.64	16.74	16.50	16.57	16.61	16.70
-0.75	16.86	17.04	16.83	17.00	16.69	16.76	16.88	16.98
-0.70	17.06	17.31	17.05	17.26	16.95	17.08	17.14	17.25
-0.65	17.23	17.57	17.26	17.48	17.22	17.37	17.45	17.60
-0.60	17.26	17.54	17.41	17.63	17.50	17.65	17.76	17.95
-0.55	17.22	17.45	17.50	17.68	17.69	17.74	18.08	18.15
-0.50	17.07	17.17	17.55	17.68	17.98	18.03	18.42	18.46
-0.45	16.95	16.95	17.63	17.64	18.29	18.29	18.88	18.88
-0.40	17.00	17.01	17.87	17.90	18.71	18.75	19.50	19.50
-0.35	17.37	17.39	18.41	18.44	19.44	19.47	20.32	20.32
-0.30	18.24	18.28	19.40	19.44	20.56	20.56	21.44	21.47
-0.25	19.55	19.58	20.79	20.82	21.98	22.00	22.95	22.95
-0.20	21.80	21.86	22.87	22.93	23.80	23.83	24.70	24.80
-0.15	24.34	24.67	25.14	25.38	25.75	25.99	26.48	26.69
-0.10	26.80	27.26	27.21	27.60	27.40	27.70	27.81	28.23
-0.05	28.55	29.14	28.69	29.20	28.64	29.07	28.98	29.30
0.00	29.51	30.06	29.46	29.85	29.24	29.51	29.48	29.72
0.05	29.81	30.10	29.76	29.95	29.48	29.61	29.72	29.88
0.10	29.81	29.81	29.76	29.76	29.61	29.68	29.75	29.80
0.15	29.94	29.94	29.96	29.96	29.80	29.80	29.96	29.96
0.20	30.46	30.46	30.46	30.46	30.41	30.41	30.60	30.60
0.25	31.41	31.41	31.50	31.60	31.52	31.82	31.73	31.73
0.30	33.38	34.10	33.42	34.10	33.72	34.70	33.96	34.20
0.35	38.49	39.90	38.80	40.30	40.36	42.75	41.92	45.00

Capacity for 5-chloro-1-pentanol in 0.916 NaF (continued)

$C/\mu F\ cm^{-2}$

$c/mol\ l^{-1}$	0.0003305		0.0002145		0.0001395		
E/V	400 Hz	extr.	400 Hz	extr.	400 Hz	extr.	pure sat. NaF
-1.50	19.64	19.64	19.65	19.65	19.65	19.65	19.65
-1.40	18.63	18.63	18.60	18.60	18.66	18.66	18.76
-1.30	17.80	17.80	17.73	17.73	17.82	17.82	17.89
-1.20	17.09	17.09	17.04	17.04	17.07	17.07	17.16
-1.10	16.53	16.53	16.53	16.53	16.52	16.52	16.56
-1.00	16.22	16.22	16.18	16.18	16.21	16.21	16.19
-0.95	16.19	16.19	16.09	16.09	16.10	16.10	16.08
-0.90	16.19	16.19	16.11	16.11	16.10	16.10	16.05
-0.85	16.28	16.28	16.18	16.18	16.11	16.11	16.09
-0.80	16.46	16.46	16.38	16.39	16.33	16.33	16.24
-0.75	16.70	16.72	16.64	16.64	16.58	16.58	16.50
-0.70	17.04	17.11	17.01	17.01	16.91	16.91	16.89
-0.65	17.39	17.48	17.43	17.43	17.38	17.38	17.45
-0.60	17.77	17.85	17.98	17.98	17.96	17.96	18.11
-0.55	18.26	18.31	18.63	18.63	18.68	18.68	18.96
-0.50	18.79	18.85	19.40	19.40	19.54	19.54	19.99
-0.45	19.40	19.40	20.28	20.28	20.53	20.53	21.14
-0.40	20.18	20.21	21.33	21.37	21.66	21.66	22.42
-0.35	21.14	21.17	22.49	22.49	22.89	22.89	23.76
-0.30	22.38	22.38	23.78	23.78	24.18	24.18	25.10
-0.25	23.83	23.85	25.10	25.11	25.52	25.52	26.35
-0.20	25.37	25.41	26.35	26.36	26.73	26.73	27.32
-0.15	26.80	26.85	27.48	27.51	27.63	27.63	28.08
-0.10	27.95	28.03	28.26	28.29	28.35	28.35	28.53
-0.05	28.81	28.98	28.81	28.86	28.81	28.81	28.75
0.00	29.24	29.34	29.00	29.07	29.02	29.02	28.88
0.05	29.36	29.36	29.11	29.16	29.16	29.16	28.95
0.10	29.53	29.53	29.33	29.35	29.36	29.36	29.13
0.15	29.81	29.81	29.66	29.66	29.71	29.71	29.51
0.20	30.41	30.41	30.33	30.36	30.35	30.35	30.16
0.25	31.41	31.48	31.47	31.47	31.50	31.50	31.33
0.30	33.42	33.55	33.79	34.00	33.85	33.91	34.37
0.35	39.30	39.56	41.35	44.95	39.86	42.70	39.19

Capacity data for mercury in iso-pentanol in 0.1 M $HClO_4$ solution.
Reference electrode: SCE (NaCl solution); c/c_o as defined below. T = 25°C

Reference: K.G. Baikerikar, D.E. Broadhead and R.S. Hansen. J. Phys. Chem. 80 (1976) 370.

c = solute concentration
c_o its saturation concentration in the base electrolyte

	$C/\mu F\ cm^{-2}$									
c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.3000	0.3548	0.5000
$-E/V$										
1.100	15.56		16.46	16.95	18.12	19.42	22.92	30.98	29.56	19.93
1.075							24.14		24.59	16.63
1.050	15.72		16.67	17.32	18.88	20.81	25.73	31.46	20.24	14.21
1.025							27.09	28.74	16.89	
1.000	15.95	16.42	17.09	17.86	20.10	22.88	27.95	24.84	14.29	10.99
0.975							27.10	20.98		
0.950	16.14	16.62	17.38	18.41	21.27	24.35	25.12	17.65	10.96	9.01
0.925							22.29	14.97		
0.900	16.47	16.93	17.77	18.95	22.13	23.43	19.12	12.89	8.91	7.72
0.875							16.48			
0.850	16.74	17.08	17.90	19.19	20.99	19.34	14.18	10.08	7.61	6.82
0.800	17.16	17.28	17.89	18.74	17.81	14.77	10.89	8.34	6.71	6.17
0.750	17.58	17.28	17.31	17.16	14.14	11.49	8.92	7.23	6.09	5.68
0.700	18.25	17.35	16.33	14.85	11.30	9.34	7.63	6.47	5.62	5.33
0.650	19.11	17.42	15.09	12.66	9.45	8.05	6.84	5.97	5.31	5.05
0.600	20.55	18.02	14.21	11.13	8.37	7.27	6.33	5.63	5.07	4.87
0.550	22.54	19.43	14.29	10.58	7.87	6.90	6.07	5.45	4.95	4.76
0.500	25.12	22.19	16.19	11.24	7.98	6.92	6.03	5.40	4.91	4.72

Capacity data for iso-pentanol in 0.1 M $HClO_4$ (continued)

	$C/\mu F\ cm^{-2}$									
c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.2000	0.3548	0.5000
$-E/V$										
0.475	26.40		18.08							
0.466		24.54								
0.450	27.47	25.73	20.08	14.10	8.91	7.43	6.29	5.54	4.98	4.77
0.425		27.32	24.14							
0.400	28.84	28.66	27.66	21.32	11.62	8.96	7.05	5.96	5.21	4.95
0.375		29.36	30.19	26.95		10.52				
0.350	28.85	29.59	31.54	32.55	19.16	12.70	8.78	6.87	5.71	5.34
0.343			31.62							
0.338							9.48			
0.337			31.67							
0.332		29.47								
0.331			31.66							
0.325		29.40	31.61	35.15	26.85	16.57				
0.317				35.72						
0.315				35.75						
0.313				35.72						
0.306				35.60						
0.303				35.52						
0.300	27.94	28.90	30.99	35.34	37.80	23.36	12.98	8.83	6.72	6.08
0.297						24.31				
0.293								9.25		
0.287					42.81	28.45				
0.276					45.54					
0.275			29.96	33.67	45.72	34.59	17.18			

Capacity data for iso-pentanol in 0.1 M $HClO_4$ (continued)

$C/\mu F\ cm^{-2}$

c/c_0	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.2000	0.3548	0.5000
-E/V										
0.271					46.00					
0.270					46.15					
0.268									7.87	
0.266					46.18					
0.263					46.15					
0.262						42.29	20.40			
0.260					45.86					7.19
0.250	26.57	27.42	28.86	31.60	44.49	48.90	24.39	13.32	8.80	7.57
0.247					43.79					
0.242					42.70					
0.237					41.57	53.52	30.09			
0.232					40.44					
0.230						54.10				
0.225				29.84	38.90	53.84	36.91	17.35	10.52	
0.212						50.56	46.44			
0.200	25.26	25.94	26.87	28.33	34.19	46.36	55.82	24.23	13.08	10.54
0.187							63.36	29.34	14.77	
0.183							64.51			
0.180							64.94			
0.178							65.04			
0.175					30.92	38.17	64.78	35.86	16.72	12.93
0.170							63.66			
0.162							60.39	44.70	19.23	
0.150	24.19	24.72	25.36	26.22	28.74	32.99	53.46	54.63	22.15	16.31

Capacity data for iso-pentanol in 0.1 M $HClO_4$ (continued)

$C/\mu F\ cm^{-2}$

c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.2000	0.3548	0.5000
-E/V										
0.137							46.64	65.80	25.94	18.49
0.125						29.92	41.40	72.41	30.43	20.91
0.120								73.32		
0.112							37.19	72.00	36.50	24.06
0.100	23.52	23.96	24.41	24.93	26.30	28.03	34.37	65.61	43.65	27.52
0.087								56.43	52.87	31.84
0.075							30.68	48.94	62.64	36.73
0.062								42.69	72.63	43.20
0.050	23.21	23.54	23.91	24.28	25.13	26.03	28.54	38.49	77.68	50.43
0.045									78.24	
0.037								35.29	76.25	
0.025								33.11	69.44	59.26
0.012								31.38	59.93	67.91
0.000	23.30	23.55	23.83	24.11	24.73	25.27	26.60	30.27	51.96	75.59
-0.012									45.97	78.40
-0.025									41.05	75.86
-0.037									37.75	68.32
-0.050	23.66	23.84	24.07	24.30	24.75	25.15	25.96	27.88	35.20	60.21
-0.062										52.34
-0.075										46.51
-0.087									32.22	41.88
-0.100	24.29	24.41	24.62	24.78	25.13	25.43	25.99	27.20	30.63	38.82
-0.125										36.41
-0.150	25.08	25.18	25.34	25.46	25.73	25.95	26.39	27.23	29.28	31.93
-0.200	26.05		26.23	26.32	26.54	26.71	27.02	27.60	29.00	30.57

Capacity data for mercury in n-butanol in 0.1M $HClO_4$ solution
Reference electrode : SCE (NaCl solution); c/c_o defined below.

T = 25°C

c = solute concentration

c_o its saturation concentration
in the base electrolyte

Reference: K.G. Baikerikar, D.E. Broadhead and R.S. Hansen. J. Phys. Chem. 80 (1976) 370.

	$C/\mu F\ cm^{-2}$								
c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.2000	0.5000
-E/V									
1.100	15.74	16.36	17.07	17.68	19.58	21.77	25.99	28.49	14.65
1.075								26.16	
1.050	15.85	16.51	17.32	18.13	20.43	23.02	26.21	23.36	11.66
1.025						23.48		20.51	
1.000	16.01	16.72	17.66	18.57	21.39	23.63	23.53	17.81	9.80
0.975						23.15			
0.950	16.17	16.89	17.88	19.00	21.51	22.15	19.03	13.82	8.47
0.900	16.41	17.03	18.02	19.05	20.41	18.90	14.92	11.08	7.57
0.850	16.65	17.08	17.86	18.57	17.90	15.37	11.93	9.31	6.88
0.800	16.96	17.01	17.35	17.29	14.96	12.45	9.92	8.10	6.40
0.750	17.34	16.85	16.40	15.49	12.43	10.42	8.58	7.26	6.00
0.700	17.89	16.69	15.27	13.62	10.61	9.04	7.70	6.70	5.74
0.650	18.73	16.68	14.25	12.16	9.42	8.21	7.12	6.33	5.52
0.600	20.07	17.23	13.82	11.37	8.81	7.73	6.82	6.12	5.42
0.575	20.98								
0.550	22.09	18.74	14.43	11.50	8.71	7.66	6.73	6.07	5.37
0.525	23.31	19.96	15.31						
0.500	24.63	21.56	16.75	12.92	9.35	8.01	6.95	6.19	5.43

Capacity data for n-butanol in 0.1 M $HClO_4$ solution (continued)

	$C/\mu F\ cm^{-2}$								
c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.2000	0.5000
-E/V									
0.475			18.71	14.36					
0.470	26.16								
0.461		24.42							
0.450	27.05	25.26	21.35	16.54	11.03	9.08	7.56	6.57	5.59
0.446			21.79						
0.425		27.00	24.39	19.60					
0.400	28.52	28.41	27.50	23.54	15.01	11.49	8.91	7.37	5.96
0.387				25.81					
0.384			29.19						
0.375				28.00	18.39				
0.367			30.65						
0.362				30.27	20.71				
0.350	28.68	29 .65	31.59	32.13	23.29	16.55	11.67	8.94	6.62
0.337				33.68	26.37	18.62			
0.335			31.94						
0.325			32.02	34.64	29.58	20.95	13.99		
0.315			31.94						
0.312				35.16	33.13	23.91			
0.310			31.87						
0.305			31.79						
0.300	27.85	29.10	31.64	35.19	36.38	27.26	17.40	12.03	7.84
0.295			31.51						

Capacity data for n-butanol in 0.1 M $HClO_4$ solution (continued)

$C/\mu F\ cm^{-2}$

c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.2000	0.5000
-E/V									
0.287			31.25	34.83		31.03	19.68		
0.275			30.80	34.22	40.91	34.92	22.19	14.52	9.03
0.268					41.45				
0.262			30.25		41.72	39.09	25.45		
0.256					41.69				
0.250	26.54	27.70	29.72	32.60	41.48	42.50	28.96	18.03	10.08
0.237					40.41	45.12	33.31		
0.225			28.62	30.88	38.95	46.17	37.69	22.94	
0.220						46.21			
0.212						45.84	42.53	26.18	
0.200	25.24	26.18	27.59	29.36	35.66	44.50	46.65	29.68	14.09
0.187						42.31	49.69	33.88	
0.180							50.63		
0.175					32.62	40.09	50.92		17.03
0.170							51.01	38.24	
0.162						37.73	50.51	43.18	
0.150	24.21	24.96	25.93	29.96	30.30	35.74	48.70	47.67	20.77
0.137							45.90	51.70	
0.125					28.60	32.43	43.01	54.11	25.40
0.117								54.76	
0.112							40.06	54.81	

Capacity data for n-butanol in 0.1 M $HClO_4$ solution (continued)

$C/\mu\mathrm{F\,cm}^{-2}$

c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.2000	0.5000
-E/V									
0.107								54.58	
0.100	23.54	24.13	24.86	25.50	27.39	30.08	37.69	53.79	30.91
0.087							35.37	51.15	
0.075							33.68	48.10	37.33
0.062							32.15	44.59	
0.050	23.26	23.77	24.29	24.73	25.91	27.41	31.00	41.53	44.57
0.037								38.76	
0.025							29.31	36.60	51.43
0.000	23.32	23.73	24.18	24.46	25.33	26.25	28.24	33.28	55.98
-0.025									55.82
-0.050	23.65	24.02	24.40	24.62	25.27	25.93	27.17	29.92	51.69
-0.075									45.98
-0.100	24.25	24.55	24.88	25.06	27.39	26.08	26.97	28.69	41.28
-0.125									37.71
-0.150	25.04	25.29	25.57	25.70	26.13	26.55	27.20	28.43	35.51
-0.175									34.01
-0.200	26.01	26.21	26.43	26.51	26.88	27.19	27.69	28.64	33.16

Differential capacity in 0.5M H_2SO_4 containing n-Cctanol.

Reference: R.J. Meakins, J. Appl. Chem. 15 (1965) 416.

	$C/\mu F\ cm^{-2}$			
	n-Octanol			
$c/mol\ l^{-1}$	0	3×10^{-4}	10^{-3}	3×10^{-3}
E/V				
0.0	31.2	34.2	37.4	38.8
0.1	29.1	30.2	31.8	37.4
0.2	30.7	30.9	31.5	7.3
0.3	33.5	33.8	8.1	5.3
0.4	32.7	32.8	6.2	4.9
0.5	27.2	27.1	6.0	4.9
0.6	22.2	22.0	6.4	5.1
0.7	19.0	18.9	7.3	5.6
0.8	17.3	17.2	9.6	6.8
0.9	16.3	16.2	13.9	9.4
1.0	15.6	15.7	15.4	14.5
1.1	15.4	15.5	15.7	17.6
1.2	15.7	15.8	15.9	17.4
1.3	15.9			
1.4				

Differential capacity of mercury in 1 mol l^{-1} KNO_3 in water saturated with octyl alcohol. T = 25°C

Potential with respect to the potential of zero charge in the absence of octyl alcohol (i.e. -0.617 V with respect to a 0.1 mol l^{-1} KCl calomel dipping into 0.1 mol l^{-1} KNO_3).

Reference: D.C. Grahame, J. Amer. Chem. Soc., 68 (1946) 301.

			C/μF cm^{-2}		
f/Hz E/volts	240	480	1000	5000	10,000
0.80	43.0	42.8	42.3	42.2	41.7
0.50	27.2	27.2	27.0	26.8	26.5
0.475	28.9	28.6	27.8	27.6	27.0
0.45	56.5	45.5	41.5	31.8	29.3
0.40	11.0	10.6	10.5	10.1	10.0
0.30	5.21	5.21	5.20	5.15	5.14
0.00	4.00	4.03	4.00	4.03	4.05
-0.20	4.30	4.28	4.30	4.30	4.33
-0.50	8.36	8.08	7.88	7.71	7.66
-0.55	14.1	13.5	12.6	11.8	11.6
-0.60	49.6	41.4	35.5	28.1	26.4
-0.65	36.1	34.2	32.4	27.6	25.3
-0.70	22.7	22.4	22.2	21.8	21.4
-0.90	18.8	18.7	18.7	18.7	18.7

Differential capacity on mercury electrode in 0.1 M Na_2SO_4 at 25°C.
Adsorption of t-butanol - given are the capacity in $\mu F\ cm^{-2}$, potential in V.
Reference electrode: saturated calomel electrode.

Reference: A. de Battisti, H.B. Abd. El Nabey and S. Trasatti.
J. Chem. Soc. Farad. Trans. 1, 72 (1976) 2076.

	$C/\mu F\ cm^{-2}$							
$c/mol\ l^{-1}$	0	0.0097	0.0162	0.0216	0.0324	0.0432	0.0540	0.0594
-E/V								
0.0000	37.53	38.09	37.84	37.99	38.30	38.58	38.51	38.37
0.0500	37.53	38.07	38.21	38.38	38.50	38.80	38.82	38.83
0.0750								
0.0850								
0.0950								
0.1000	37.97	38.62	38.88	38.99	39.24	39.56	39.84	39.87
0.1100								
0.1200								
0.1250								
0.1258								
0.1300								
0.1350								
0.1400								
0.1430								
0.1500	38.16	38.94	39.27	39.47	39.70	40.31	40.51	40.88
0.1585								
0.1600								
0.1750		39.25	39.05		39.48	40.30	40.37	40.78
0.1796								
0.1987								
0.2000	37.10	37.86	38.20	38.50	38.77	39.46	40.04	40.16
0.2245								
0.2250					37.38	38.28	38.63	39.15
0.2500	34.34	34.94	35.23	35.60	35.60	36.59	36.93	37.50
0.2503								
0.2747								
0.2750		35.05	33.23		33.64	34.51	34.76	35.28
0.3000	30.75	31.11	31.34	31.45	31.55	31.99	32.48	
0.3062								
0.3250		29.26	29.45			29.90	30.43	37.40
0.3255								
0.3500	27.39	27.41	27.55	27.76	27.41	27.73	28.08	28.35
0.3505								
0.3704								

Differential capacity on mercury in 0.1 M Na_2SO_4 and tert.butanol (continued)

$c/mol\ l^{-1}$	0	0.0097	0.0162	0.0216	0.0324	0.0432	0.0540	0.0594
	$C/\mu F\ cm^{-2}$							
$-E/V$								
0.3750		26.01	26.03			25.69	25.77	26.16
0.3785								
0.3873								
0.3988								23.89
0.4000	24.57	24.45	24.48	24.35	23.98	23.89	23.69	23.94
0.4056							23.32	
0.4133						22.99		
0.4185					22.80			
0.4236				23.25				
0.4246			23.24					
0.4250		23.33						
0.4304	23.30							
0.4500	22.60	22.43	22.28	22.00	21.45	21.00	20.13	19.67
0.5000	21.29	21.84	20.83	20.54	19.74	19.00	17.42	16.49
0.5500	20.57	20.23	19.95	19.63	18.74	17.76	15.76	14.70
0.6000	19.93	19.60	19.29	18.98	18.09	17.17	15.46	14.56
0.6500	19.26	19.02	18.73	18.47	17.71	17.06	15.95	15.45
0.7000	18.54	18.41	18.31	18.06	17.47	17.15	16.71	16.59
0.7500	17.86	17.85	17.72	17.65	17.35	17.47	17.37	17.44
0.8000	17.28	17.31	17.32	17.31	17.21	17.52	17.66	17.90
0.8500	16.73	16.89	16.92	16.98	16.99	17.29	17.65	17.89
0.9000	16.31	16.47	16.59	16.71	16.81	17.16	17.53	17.76
0.9500	15.99	16.22	16.36	16.45	16.61	16.94	17.23	17.52
1.0000	15.76	16.03	16.16	16.23	16.39	16.73	17.02	17.24
1.0500	15.62	15.92	15.99	16.11	16.27	16.58	16.84	16.98
1.1000	15.59	15.82	15.95	16.02	16.16	16.44	16.62	16.82
1.1500	15.64	15.92	15.97	16.00	16.16	16.38	16.50	16.67
1.2000	15.75	15.92	16.05	16.06	16.20	16.38	16.46	16.63
1.2500	15.91	16.05	16.18	16.17	16.28	16.41	16.47	16.63
1.3000	16.13	16.27	16.31	16.35	16.43	16.56	16.59	16.71
1.3500	16.40	16.48	16.55	16.56	16.61	16.72	16.74	16.83
1.4000	16.70	16.79	16.81	16.82	16.86	16.95	16.96	17.02
1.4500	17.01	17.10	17.14	17.14	17.14	17.22	17.22	17.31
1.5000	17.39	17.45	17.49	17.48	17.48	17.54	17.53	17.62
1.5500	17.78	17.70	17.86	17.85	17.72	17.90	17.89	17.96
1.6000	18.18	18.28	18.29	18.28	18.26	18.30	18.29	18.36
p.z.c.	0.6201	0.6166	0.6182	0.6200	0.6221	0.6240	0.6343	0.6298

Differential capacity on mercury in 0.1 M Na_2SO_4 and tert.butanol (continued)

	$C/\mu F\ cm^{-2}$						
$c/mol\ l^{-1}$	0.0702	0.0756	0.0809	0.0917	0.1079	0.1214	0.1484
-E/V							
0.0000	38.65	38.80	39.05	38.90	39.32	39.47	39.70
0.0500	39.17	39.30	39.51	39.67	40.15	40.36	40.90
0.0750							
0.0850							
0.0950							
0.1000	40.15	40.40	40.81	41.04	41.90	42.37	43.50
0.1100							
0.1200							
0.1250							
0.1258							
0.1300							
0.1350							
0.1400							
0.1430							
0.1500	41.31	41.89	42.05	42.94	44.04	45.06	48.17
0.1585							
0.1600							
0.1750	41.34	42.07	42.22	43.21	45.17	47.08	52.99
0.1796							
0.1987							
0.2000	41.06	41.70	42.07	43.59	46.03	49.35	62.45
0.2245							
0.2250	39.94	40.91	41.28	43.34	47.65	54.35	90.07
0.2500	38.55	39.32	40.17	43.25	50.84	63.85	73.19
0.2503							
0.2747							31.61
0.2750	36.54	37.68	38.97	43.27	54.95	65.27	
0.3000	37.65	35.26	37.28	43.46	51.98	38.67	17.38
0.3062						33.20	
0.3250	32.26	33.99	35.49	40.87		21.60	12.04
0.3255					34.28		
0.3500	29.99	31.53	32.68		21.23	14.19	9.43
0.3505				33.37			
0.3704			29.06				

Differential capacity on mercury in 0.1 M Na_2SO_4 and tert.butanol

(continued)

	C/μF cm^{-2}						
c/mol l^{-1}	0.0702	0.0756	0.0809	0.0917	0.1079	0.1214	0.1484
-E/V							
0.3750	27.15	27.96	28.25	22.97	14.33	10.63	7.96
0.3785		27.36					
0.3873	25.82						
0.3998							
0.4000	24.19	24.91	22.68	16.34	10.96	8.85	7.12
0.4056							
0.4133							
0.4185							
0.4236							
0.4246							
0.4250		19.30		12.46	9.11		
0.4304							
0.4500	17.65	15.67	13.90	10.22	8.00	7.11	6.25
0.5000	13.25	11.51	10.33	8.32	7.10	6.55	5.93
0.5500	11.77	10.39	9.50	7.98	6.98	6.49	5.93
0.6000	12.11	11.08	10.78	8.61	7.40	6.84	6.17
0.6500	14.03	13.02	12.01	10.25	8.49	7.66	6.72
0.7000	16.30	15.81	15.29	12.49	10.64	9.27	7.73
0.7500	18.04	18.30	18.48	13.13	14.72	12.26	9.53
0.8000	18.72	19.35	19.95	17. 4	20.43	17.74	12.99
0.8500	18.72	19.30	19.86	21.	23.96	24.26	19.59
0.9000	18.38	18.87	19.35	21.70	23.22	25.65	27.49
0.9500	18.05	18.29	18.70	20.76	21.27	23.33	28.30
1.0000	17.60	17.84	18.11	19.43	19.77	20.95	24.29
1.0500	17. 7	17.49	17.71	18.83	18.85	19.55	21.36
1.1000	17. 4	17.19	17.38	18.1	18.08	18.54	21.36
1.1500	16.88	16.95	17.07	17.71	17.59	18.01	11.70
1.2000	16.77	16.82	16.93	17.36	17.27	17.60	18.08
1.2500	16.76	16.76	16.88	17.12	17.14	17.33	17.64
1.3000	16. 1	16.81	17.17	17.00	17.05	17.24	17.44
1.3500	16.9	16.90	16.99	17.00	17.08	17.22	17.40
1.4000	17.13	17.08	16.91	17.07	17.22	17.32	17.44
1.4500	17.35	17.33	16.73	17.22	16.78	17.44	17.56
1.5000	17.64	17.60	17.71	17.45	17.71	17.89	17.78
1.5500	17.98	17.93	18.07	17.73	18.00	17.99	18.09
1.6000	18.37	18.32	18.38	18.10	18.35	18.37	18.42
p.z.c.	0.6517	0.6587	0.6733	0.6910	0.7239	0.7343	0.7541

Differential capacity on mercury in 0.1 M Na_2SO_4 and tert.butanol (continued)

	C/μF cm^{-2}						
c/mol l^{-1}	0.1754	0.2159	0.2698	0.3238	0.4047	0.4857	0.5397
-E/V							
0.0000	40.13	40.73	41.55	42.63	44.04	45.80	46.83
0.0500	41.80	42.74	44.26	46.74	51.34	59.56	68.34
0.0750					65.24	101.57	194.40
0.8500					79.08	187.99	
0.9500					111.97	420.12	171.83
0.1000	45.06	48.15	54.37	67.33	149.86	319.57	100.39
0.1100					325.84		
0.1200					235.46		
0.1250			73.19	150.45		34.54	23.08
0.1258							
0.1300				204.95			
0.1350			92.36				18.32
0.1400			111.47	210.01	43.02		
0.1430						18.66	
0.1500	54.00	69.35	174.04	116.40	27.21	16.00	13.30
0.1585					20.78		
0.1600			226.68				
0.1750	66.02	136.47	62.26	26.81	14.68		
0.1796				23.21			
0.1987			26.06				
0.2000	108.11	92.13	24.73	14.57	10.46	8.76	8.16
0.2245		28.46					
0.2250	90.46	15.33	14.32				
0.2500			10.43	8.51	7.29	6.65	6.41
0.2503	31.21						
0.2747							
0.2750	16.71	10.97					
0.3000	11.60	8.75	7.32	6.57	6.03	5.89	5.55
0.3062							
0.3250	9.13						
0.3255							
0.3500	7.76	6.71	6.08	5.68	5.41	5.21	5.13
0.3505							
0.3704							

Differential capacity on mercury in 0.1 M Na_2SO_4 and tert.butanol (continued)

$C/\mu F\ cm^{-2}$

$c/mol\ l^{-1}$	0.1754	0.2159	0.2698	0.3238	0.4047	0.4857	0.5397
-E/V							
0.3750	6.91						
0.3785							
0.3873							
0.3998							
0.4000	6.38	5.86	5.52	5.26	5.09	4.95	4.90
0.4056							
0.4133							
0.4185							
0.4236							
0.4246							
0.4250							
0.4304							
0.4500	7.81	5.47	5.26	5.06	4.93	4.82	4.78
0.5000	5.61	5.34	5.16	4.99	4.87	4.77	4.73
0.5500	5.61	5.36	5.17	5.00	4.89	4.78	4.75
0.6000	5.00	5.51	5.29	5.10	4.98	4.84	4.81
0.6500	6.21	5.82	5.52	5.29	5.13	4.99	4.93
0.7000	6.92	6.37	5.30	5.80	5.37	5.20	5.12
0.7500	8.21	7.19	6.51	6.08	5.74	5.52	5.40
0.8000	10.49	8.67	7.53	6.82	6.31	5.98	5.80
0.8500	15.14	11.57	9.30	8.11	7.21	6.67	6.43
0.9000	23.84	17.31	9.30	10.35	8.72	7.80	7.39
0.9500	44.81	28.30	19.78	14.87	11.52	9.74	9.02
1.0000	30.52	34.68	32.92	24.63	17.35	13.55	12.06
1.0500	24.18	29.37	37.83	39.12	30.19	21.84	18.46
1.1000	21.24	24.09	29.74	36.96	43.71	38.85	33.26
1.1500	19.59		24.28	28.28	36.10	44.59	48.07
1.2000	13.65	19.66	21.29	23.15	27.35	33.22	38.29
1.2500	18.67		19.67	20.79	22.90	25.84	28.00
1.3000	17.76	18.25	18.81	19.49	20.72	22.11	23.26
1.3500	17.63		18.42	18.79	19.56	20.37	21.16
1.4000	17.63	17.90	18.18	18.46	19.00	19.61	19.90
1.4500	17.71		18.16	18.32	18.70	19.02	19.30
1.5000	17.90	18.03	18.19	18.36	18.62	18.81	18.90
1.5500	18.17		18.41	18.49	18.70	18.86	18.96
1.6000	18.49	18.52	18.67	18.75	18.87	18.96	19.56
p.z.c.	0.7644	0.7764	0.7879	0.7990	0.8047	0.8142	0.8145

Capacity data for mercury in benzene and toluene in 0.1 M KCl solution at 25°C . Reference electrode - SCE (NaCl solution).

c/c_0 is reduced concentration
c is adsorbate concentration
c_0 its saturation concentration in the base electrolyte solution

Reference: K.G. Baikerikar and R.S. Hansen. J. Colloid Interface Ci. 61 (1977) 239

$C/\mu F\ cm^{-2}$

c/c_0	Benzene		Toluene		
-E/V	0.40	1.00	0.40	0.80	1.00
1.600	19.12	19.22	-	19.19	19.17
1.550	18.57	18.68	-	18.64	18.58
1.500	18.06	18.19	18.12	18.11	18.12
1.450	17.62	17.76	17.66	17.66	17.67
1.400	17.23	17.39	17.28	17.28	17.32
1.350	16.93	17.14	16.96	17.00	17.04
1.300	16.71	16.99	16.70	16.70	16.88
1.250	16.58	17.02	16.55	16.73	16.90
1.200	16.58	17.23	16.51	16.89	17.19
1.150	16.73	17.70	16.62	17.32	17.82
1.125				17.67	
1.100	17.10	18.52	16.95	18.12	18.99
1.075		19.12		18.73	19.81
1.062					20.33
1.050	17.68	19.88	17.50	19.42	20.85
1.037					21.49
1.025	18.07	20.73		20.30	22.07
1.012					22.84
1.000	18.50	21.71	18.26	21.27	23.56
0.987					24.43
0.975	18.99	22.88	18.70	22.26	25.12
0.962					26.07
0.950	19.50	24.38	19.12	23.22	26.70
0.937					27.20
0.930					27.47
0.925	20.01	25.90	19.47	23.91	27.42
0.912					26.81
0.900	20.51	27.64	19.75	23.98	25.28
0.887				23.54	22.71

Capacity data for benzene and toluene in 0.1 M KCl solution (continued)

		$C/\mu F\,cm^{-2}$			
c/c_0	Benzene		Toluene		
-E/V	0.40	1.00	0.40	0.80	1.00
0.875	20.91	29.23	19.00	22.88	19.80
0.862		30.89		21.71	16.89
0.850	21.21	28.96	19.86	20.28	14.61
0.837		27.29		18.46	12.49
0.825	21.31	24.88	19.59	16.64	11.14
0.812		16.46		14.82	9.98
0.800	21.23	14.12	19.10	13.31	9.10
0.787		11.82		11.93	
0.775	20.87	10.74	18.37	10.87	7.86
0.762		9.36		9.96	
0.750	20.26	8.80	17.47	9.26	7.08
0.725	19.37	7.42	16.45	8.19	
0.700	18.32	6.95	15.36	7.46	6.10
0.675	17.14		14.31	6.95	
0.650	15.99	6.16	13.32	6.60	5.75
0.625	15.00		12.48		
0.600	14.20	5.80	11.68	6.21	5.54
0.575	13.73				
0.550	13.62	5.71	11.37	6.09	5.50
0.525	13.96				
0.500	14.90	5.86	11.98	6.23	5.60
0.487	15.69				
0.475	16.58		12.94		
0.462	17.85				
0.450	19.23	6.43	14.53	6.74	5.91
0.437	20.99		15.68		
0.425	22.87	6.88	16.99	7.30	
0.412	25.00		18.65		

Capacity data for benzene and toluene in 0.1 M KCl solution (continued)

$C/\mu F\,cm^{-2}$

c/c_0	Benzene		Toluene		
$-E/V$	0.40	1.00	0.40	0.80	1.00
0.400	27.32	8.15	20.44	8.84	6.72
0.387	29.74	8.78	22.54	9.01	
0.375	31.96	10.26	24.89	9.91	
0.362	34.13	11.60	27.41	11.27	
0.350	35.93	14.72	29.73	12.95	7.99
0.337	37.48	17.60	32.64	15.34	8.14
0.325	38.56	24.79	33.96	18.33	10.49
0.312	39.87	30.86	35.60	22.42	14.75
0.300	39.87	43.45	36.90	27.18	18.05
0.287	40.07	52.09	37.88	32.45	22.95
0.275	40.15	57.91	38.45	37.32	29.00
0.262		59.92	38.79	41.54	36.61
0.250	39.83	56.07	38.94	44.18	43.58
0.237		54.49		45.40	49.20
0.225	39.29	50.52	38.94	45.57	51.39
0.212		48.99		45.14	52.37
0.200	38.89	46.50	38.74	44.54	51.45
0.187		45.56		43.83	50.00
0.175	38.80	44.28	38.82	43.23	48.62
0.150	39.20	43.19	39.25	42.57	46.41
0.125	40.17	43.19	39.25	42.64	45.54
0.100	41.94	44.05	40.25	43.70	45.92
0.075	44.81	46.55	42.07	45.98	46.67
0.050	49.13	50.42	44.93	49.93	51.28
0.025	56.23	57.31	49.26	56.60	57.59
0.000	70.03		56.42	70.16	

Capacity data for mercury in 0.1 M KI solution and for this solution satured with benzene and with toluene. Reference electrode : SCE (NaCl solution).

T = 25°C

Reference: K.G. Baikerikar and Robert S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

0.1 M KI		+ benzene		+ toluene	
-E/V	C/µF cm^{-2}	-E/V	C/µF cm^{-2}	-E/V	C/µF cm^{-2}
1.600	19.20	1.600	19.22	1.600	19.21
1.550	18.63	1.550	18.65	1.550	18.67
1.500	18.12	1.500	18.14	1.500	18.16
1.450	17.63	1.450	17.71	1.450	17.73
1.400	17.22	1.400	17.39	1.400	17.38
1.350	16.87	1.350	17.14	1.350	17.12
1.300	16.56	1.300	17.02	1.300	17.01
1.250	16.32	1.250	17.06	1.250	17.13
1.200	16.19	1.200	17.35	1.200	17.55
1.150	16.15	1.150	18.00	1.150	18.51
1.100	16.31	1.100	19.16	1.125	19.26
1.050	16.81	1.075	19.96	1.100	20.27
1.025	17.26	1.050	21.02	1.075	21.58
1.000	17.93	1.025	22.36	1.062	22.41
0.975	18.93	1.000	24.01	1.050	23.51
0.950	20.34	0.975	26.06	1.037	24.55
0.925	22.32	0.962	27.37	1.025	25.57
0.900	25.05	0.950	28.65	1.012	26.94
0.875	28.58	0.937	30.19	1.000	28.40
0.850	32.94	0.925	31.64	0.987	31.43
0.825	37.91	0.912	33.27	0.975	32.63
0.800	43.19	0.900	34.67	0.962	33.05
0.775	48.31	0.887	35.75	0.950	33.42
0.750	52.79	0.875	36.06	0.937	31.56
0.725	56.62	0.862	35.28	0.925	27.89
0.700	59.72	0.850	33.47	0.912	
0.675	62.35	0.837	30.81	0.910	25.04
0.650	65.15	0.825	28.28	0.906	20.58
0.625	68.02	0.812	26.23	0.900	18.62

Capacity data for 0.1 M KI solution and with benzene and with toluene (continued)

0.1 M KI		+ benzene		+ toluene	
-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}	-E/V	C/μF cm^{-2}
0.600	72.35	0.800	25.15	0.887	15.38
0.575	77.58	0.787	24.83	0.875	13.25
0.550	84.92	0.775	24.81	0.862	11.74
0.525	94.88	0.762	26.98	0.850	10.86
0.500	107.72	0.750	29.27	0.837	10.30
0.475	124.71	0.737	33.20	0.825	10.03
0.450	145.76	0.725	37.30	0.800	10.29
		0.712	42.18	0.775	11.60
		0.700	46.55	0.762	12.77
		0.675	54.62	0.750	14.37
		0.650	59.79	0.737	16.52
		0.625	62.79	0.725	19.13
		0.600	65.26	0.712	22.91
		0.575	68.49	0.700	27.22
		0.550	73.50	0.687	32.67
		0.525	80.87	0.675	38.32
		0.500	91.78	0.650	50.70
		0.475	107.19	0.625	60.51
		0.450	129.15	0.600	66.10
				0.575	69.41
				0.550	73.45
				0.525	79.70
				0.500	90.01
				0.475	105.57
				0.450	128.99

Capacity data for mercury in benzene and toluene in 0.05M H_2SO_4 solution
reference electrode: SCE (NaCl solution). T = 25°C

c/c_o is reduced concentration

c is adsorbate concentration

c_o its saturation concentration in the base electrolyte concentration

Reference: K.G. Baikerikar and R.S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

$C/\mu F\ cm^{-2}$

c/c_o	Benzene		Toluene	c/c_o	Benzene		Toluene
	0.5	1.00	1.00		0.5	1.00	1.00
E/V				E/V			
-1.100	16.41	19.57	19.82	-0.850	19.78	12.68	9.51
-1.075		20.59	20.99	-0.825		9.38	7.96
-1.050	16.96	21.84	22.50	-0.800	19.87	7.87	7.03
-1.025		23.57	24.24	-0.775			6.42
-1.012		24.70		-0.750	19.21	6.31	6.01
-1.000	17.65	25.91	26.39	-0.725	18.27		
-0.987		27.61		-0.700	17.71	5.62	5.50
-0.975		29.47	28.44	-0.675	16.78		
-0.962		31.83	29.03	-0.650	15.80	5.29	5.23
-0.950	18.56	34.67	28.77	-0.625	14.84		
-0.937		38.39	26.91	-0.600	14.00	5.02	5.08
-0.932			25.65	-0.575	13.27		
-0.925		38.60	23.63	-0.550	12.72	4.96	5.02
-0.918			22.01	-0.525	12.40		
-0.912		36.07	20.27	-0.500	12.21	4.93	5.03
-0.906		33.72	18.66	-0.450	12.59	4.99	5.11
-0.900	19.21	30.73	17.12	-0.425	13.15		
-0.894			15.63	-0.412	13.58		
-0.887		23.57	14.30	-0.400	14.08	5.16	5.27
-0.875		18.96	12.25	-0.387	14.71		
-0.862		15.15	10.62	-0.375	15.40		

Capacity data for mercury in benzene and toluene in 0.05 M H_2SO_4 (continued)

	$C/\mu F\ cm^{-2}$						
	Benzene		Toluene		Benzene		Toluene
c/c_o	0.5	1.00	1.00	c/c_o	0.5	1.00	1.00
E/V				E/V			
-0.362	16.28			-0.125	36.07	61.56	27.25
-0.350	17.18	5.49	5.55	-0.118			31.04
-0.337	18.26			-0.112		58.60	35.94
-0.325	19.40			-0.108			38.49
-0.312	20.71			-0.100	36.49	54.85	44.12
-0.300	22.02	6.18	6.02	-0.094			47.73
-0.287	23.49			-0.087		51.01	50.81
-0.275	24.92			-0.075		48.22	58.01
-0.262	26.43			-0.062			53.05
-0.250	27.76	7.82	6.90	-0.050	36.95	44.62	51.70
-0.237	29.19			-0.037			49.94
-0.225	30.41	9.62	7.70	-0.032			49.16
-0.212	31.60			-0.025		42.72	48.38
-0.200	32.60	13.26	9.00	-0.012			47.06
-0.187	33.52	17.09	10.06	0.000	37.50	41.76	46.19
-0.175	34.26	22.37	11.40	0.050	38.47	41.64	44.39
-0.162	34.90	33.27	13.56	0.100	39.94	42.85	44.48
-0.150	35.39	46.24	16.42	0.150	41.89	44.72	45.35
-0.144			18.35	0.200	44.51	47.06	46.78
-0.137	35.79	58.07	21.11	0.250	48.25	50.14	49.38
-0.131			23.90	0.300		55.21	54.40
				0.350		67.29	66.86

Capacity data for mercury in benzene in 0.1M $HClO_4$ solution, reference electrode :SCE (NaCl solution) c/c_o is reduced concentration. T = 25°C

c is adsorbate concentration

c_o is its saturation concentration in the base electrolyte

Reference: K.G. Baikerikar and R.S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

$C/\mu F\ cm^{-2}$

c/c_o	0.10	0.20	0.30	0.40	0.60	0.80	1.00
E/V							
-1.100	15.50	15.68	15.91	16.21	16.74	17.39	19.15
-1.075						17.80	20.03
-1.050	15.58	15.92	16.25	16.64	17.43	18.39	21.20
-1.037							21.93
-1.025						18.93	22.68
-1.012							23.64
-1.000	15.93	16.30	16.70	17.21	18.28	19.65	24.67
-0.987							25.98
-0.975						20.36	27.48
-0.962							29.47
-0.950	16.23	16.76	17.26	17.91	19.34	21.23	31.76
-0.937							34.74
-0.925					19.91	22.13	37.31
-0.912							38.44
-0.900	16.75	17.32	17.88	18.65	20.41	23.06	36.40
-0.887							30.89
-0.875					19.24	23.98	24.83
-0.862							19.41
-0.850	17.23	17.89	18.48	19.32	21.30	24.66	15.75
-0.837							12.96
-0.825					21.54	24.74	11.20
-0.812							9.87
-0.800	17.91	18.45	18.98	19.74	21.52	23.46	8.99
-0.784						21.78	
-0.775					21.27		7.76
-0.767						19.36	

Capacity data for benzene in 0.1 M $HClO_4$ (continued)

	$C/\mu F\,cm^{-2}$						
c/c_o	0.10	0.20	0.30	0.40	0.60	0.80	1.00
E/V							
-0.750	18.49	18.97	19.34	19.81	20.53	16.93	7.01
-0.734						14.83	
-0.725					19.38		
-0.717						12.97	
-0.700	19.35	19.51	19.60	19.55	17.99	11.69	6.26
-0.675					16.40	10.15	
-0.650	20.29	20.23	19.98	19.29	15.07	9.30	6.01
-0.625					14.06	8.90	
-0.600	21.78	21.35	20.70	19.27	13.54	8.85	6.07
-0.550	23.56	22.90	21.87	19.74	13.54	9.40	6.50
-0.534						9.83	
-0.525					14.11		
-0.522				20.18			
-0.517						10.41	
-0.516			22.71				
-0.511		24.14					
-0.500	25.59	24.95	23.09	20.62	15.10	11.30	7.46
-0.497	25.63						
-0.487					15.65		
-0.484						11.97	
-0.475				21.17	16.25		8.25
-0.467						13.01	
-0.450	26.88	25.41	23.96	21.68	17.59	14.20	9.30
-0.446						14.47	
-0.434						15.39	
-0.425				22.13	18.97		10.79
-0.417						16.79	
-0.412							11.76
-0.400	27.20	25.56	24.23	22.48	20.24	18.34	12.76
-0.387							14.06
-0.384						19.50	
-0.375							15.41

Capacity data for benzene in 0.1 M $HClO_4$ (continued)

	$C/\mu F\ cm^{-2}$						
c/c_o	0.10	0.20	0.30	0.40	0.60	0.80	1.00
E/V							
-0.367						20.80	
-0.362							16.97
-0.350	26.53	25.03	23.92	22.76	21.95	21.86	18.50
-0.337							20.17
-0.325						22.93	21.69
-0.312							23.24
-0.300	25.43	24.09	23.32	22.58	22.48	23.42	24.51
-0.287							25.57
-0.275						23.49	26.29
-0.262							26.73
-0.250	24.20	23.08	22.58	22.09	22.20	23.25	26.84
-0.225						22.85	26.46
-0.200	23.19	22.26	21.92	21.58	21.71	22.44	25.63
-0.175							24.66
-0.150	22.47	21.76	21.53	21.27	21.27	21.75	23.77
-0.125							23.07
-0.100	22.19	21.64	21.49	21.29	21.21	21.42	22.56
-0.075							22.24
-0.050	22.29	21.89	21.82	21.65	21.50	21.56	22.06
-0.025							22.08
0.000	22.73	22.48	22.51	22.36	22.23	22.15	22.22
0.025						22.59	22.57
0.050	23.42	23.40	23.48	23.41	23.31	23.17	23.01
0.075						23.65	23.60
0.100	24.35	24.48	24.67	24.69	24.71	24.57	24.33
0.125				25.40	25.50	25.45	25.17
0.150	25.40	25.70	25.99	26.13	26.35	26,32	26.11
0.175				26.84		27.21	27.15
0.200	26.51	26.92	27.30	27.58	28.02	28.18	28.26

Capacity data for mercury in toluene in 0.1 M $HClO_4$ solution. T = 25°C
Reference electrode: SCE (NaCl solution).

c/c_0 is reduced concentration
c is adsorbed concentration
c_0 its saturation concentration in the base electrolyte solution

Reference: K.G. Baikerikar and R.S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

	$C/\mu F\ cm^{-2}$						
c/c_0	0.10	0.20	0.30	0.40	0.60	0.80	1.00
E/V							
-1.100	15.56	15.90	16.05	16.32	16.66	17.37	20.04
-1.087							20.68
-1.075				16.60	16.99	17.90	21.33
-1.062							22.11
-1.050	15.78	16.25	16.47	16.92	17.45	18.51	22.92
-1.037							23.93
-1.025				17.25	17.92	19.22	25.12
-1.012							26.35
-1.000	16.09	16.71	17.05	17.66	18.44	19.99	27.75
-0.987							29.14
-0.975				18.07	18.95	20.74	30.54
-0.962							31.35
-0.950	16.46	17.25	17.69	18.44	19.48	21.48	31.02
-0.937						21.84	25.67
-0.925				18.78	19.88	22.09	24.92
-0.912					20.05	22.32	20.52
-0.900	16.91	17.73	18.25	19.06	20.19	22.42	16.95
-0.887						22.34	14.95
-0.875				19.22	20.26	22.06	12.07
-0.862						21.53	10.52
-0.850	17.32	18.08	18.54	19.20	20.08	20.81	9.47
-0.837					19.86	19.69	8.61
-0.825				19.04	19.53	18.48	8.01
-0.812					19.09	17.01	7.51
-0.800	17.71	18.20	18.45	18.69	18.56	15.63	7,15

Capacity data for toluene in 0.1 M $HClO_4$ (continued)

	$C/\mu F\ cm^{-2}$						
c/c_o	0.10	0.20	0.30	0.40	0.60	0.80	1.00
E/V							
-0.325	24.99			19.91	18.92	18.00	15.08
-0.312							16.26
-0.300	24,55	22.27	21.31	20.38	19.79	19.34	17.35
-0.287							18.56
-0.275	24.08					20.36	19.59
-0.262							20.70
-0.250	23.60	21.91	21.36	20.87	20.80	21.11	21.59
-0.237							22.37
-0.225	23.16						22.97
-0.200	22.75	21.48	21.18	20.94	21.14	21.78	23.74
-0.175	21.41						24.00
-0.150	22.17	21.17	21.06	20.93	21.19	21.85	23.96
-0.125	22.03						23.73
-0.100	21.96	21.18	21.14	21.06	21.30	21.82	23.45
-0.075							23.21
-0.050	22.13	21.51	21.57	21.51	21.70	22.00	23.08
-0.025							23.13
0.000	22.63	22.53	22.36	22.33	22.51	22.73	23.28
0.025				22,87	23.03	23.22	23.59
0.050	23.39	23.26	23.46	23.51	23.69	23.87	24.07
0.075	23.86			24.22	24.42	24.65	24.74
0.100	24.38	24.51	24.23	25.01	25.27	25.52	25.57
0.125	24.91			25.77	26.14	26.45	26.58
0.150	25.45	25.79	26.26	26.56	26.99	27.44	27.74
0.175						28.39	28.98
0.200	26.57	27.07	27.60	28.00	28.62	29.34	30.24

Differential capacity of mercury in solutions of KPF_6 (concentration given in moles.l^{-1}) in sulpholane.
Temperature: 30°C. Potentials measured relative to aqueous saturated calomel electrode. (Absolute accuracy in region of hump 1,5 %, elsewhere within 1,0 %).
Reference: J. Lawrence, R. Parsons, Trans. Faraday Soc. 64 (1968) 751.

Given are Potential and Capacitance as a function of surface charge σ.

	Salt = 0.05		0.1		0.2		0.35	
σ	E	C	E	C	E	C	E	C
μ C/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2	Volts	μ F/cm^2
21			0.723	21.68	0.702	20.93		
20			0.677	22.36	0.655	22.01		
19			0.634	23.48	0.611	23.06	0.588	24.35
18			0.592	24.40	0.568	24.16	0.547	24.72
17			0.552	25.84	0.528	25.07	0.507	25.79
16			0.514	26.43	0.488	25.94	0.469	26.49
15			0.476	27.12	0.451	27.22	0.432	27.12
14	0.474	26.55	0.440	27.30	0.415	27.83	0.395	27.69
13	0.437	26.51	0.403	27.23	0.379	27.81	0.359	27.89
12	0.400	26.49	0.366	27.05	0.343	27.56	0.324	27.94
11	0.361	26.02	0.329	26.73	0.306	27.33	0.288	27.77
10	0.322	25.56	0.291	26.24	0.269	27.05	0.251	27.51
9	0.282	24.93	0.253	25.77	0.232	26.78	0.215	27.03
8	0.242	24.11	0.214	25.26	0.194	26.16	0.177	26.17
7	0.199	23.21	0.174	24.60	0.155	24.69	0.138	24.83
6	0.155	22.36	0.132	23.19	0.113	23.14	0.097	23.71
5	0.109	20.66	0.087	21.47	0.069	21.95	0.054	22.67
4	0.058	18.66	0.039	19.64	0.022	20.64	0.008	21.51
3	0.002	16.93	-0.015	17.92	-0.028	19.37	-0.039	20.56
2	-0.060	15.44	-0.072	16.91	-0.081	18.25	-0.089	19.44
1	-0.127	14.50	-0.133	15.84	-0.138	17.15	-0.143	18.19
0	-0.199	13.55	-0.198	15.00	-0.198	16.12	-0.199	16.99
-1	-0.274	13.22	-0.268	14.08	-0.262	15.17	-0.260	15.81
-2	-0.350	13.28	-0.341	13.22	-0.330	14.28	-0.326	14.52
-3	-0.428	12.47	-0.418	12.59	-0.403	13.17	-0.398	13.33
-4	-0.510	11.62	-0.501	11.55	-0.482	11.95	-0.477	12.07
-5	-0.601	10.57	-0.591	10.74	-0.570	10.86	-0.564	11.00
-6	-0.700	9.71	-0.688	9.79	-0.667	9.85	-0.660	9.92
-7	-0.807	9.01	-0.795	8.98	-0.772	9.05	-0.765	9.13
-8							-0.879	8.44
-9							-1.001	7.92

Differential capacity of mercury in solutions of KPF_6 in sulpholane (cont.)

	Salt = 0.5		0.75		1.00	
σ	E	C	E	C	E	C
$\mu C/cm^2$	Volts	$\mu F/cm^2$	Volts	$\mu F/cm^2$	Volts	$\mu F/cm^2$
22	0.712	20.55				
21	0.666	23.31	0.655	22.43		
20	0.624	24.00	0.611	23.00	0.587	25.32
19	0.583	24.25	0.568	23.89	0.547	25.39
18	0.542	24.75	0.527	24.80	0.508	26.04
17	0.502	25.54	0.488	25.69	0.470	27.09
16	0.464	26.43	0.449	26.59	0.434	27.95
15	0.427	27.22	0.412	27.29	0.399	28.57
14	0.390	27.84	0.376	27.88	0.364	28.99
13	0.354	28.01	0.341	28.32	0.330	29.12
12	0.319	28.10	0.305	28.33	0.295	28.99
11	0.283	27.99	0.270	28.03	0.260	28.49
10	0.247	27.47	0.234	27.59	0.225	27.85
9	0.211	27.02	0.197	27.00	0.189	27.23
8	0.173	26.14	0.160	26.32	0.151	26.55
7	0.134	25.20	0.121	25.65	0.113	25.95
6	0.093	24.27	0.082	24.69	0.074	25.28
5	0.052	23.36	0.040	23.72	0.034	24.57
4	0.008	22.08	-0.003	22.72	-0.007	23.55
3	-0.039	20.94	-0.048	21.62	-0.051	22.37
2	-0.088	19.75	-0.096	20.39	-0.097	21.10
1	-0.141	18.41	-0.146	19.08	-0.146	18.77
0	-0.197	17.13	-0.201	17.88	-0.200	18.40
-1	-0.258	15.73	-0.259	16.41	-0.257	16.75
-2	-0.324	14.41	-0.323	14.90	-0.319	15.26
-3	-0.397	13.06	-0.393	13.56	-0.388	13.84
-4	-0.477	11.84	-0.471	12.24	-0.464	12.52
-5	-0.566	10.78	-0.557	11.07	-0.548	11.29
-6	-0.663	9.83	-0.652	10.02	-0.642	10.14
-7	-0.770	8.95	-0.756	9.20	-0.745	9.27
-8	-0.886	8.33	-0.869	8.50	-0.857	8.59

Differential capacity of mercury in solutions of $NaClO_4$ (concentration given in moles.l^{-1}) in sulpholane.
Temperature: 30°C. Potentials measured relative to aqueous saturated calomel electrode. (Absolute accuracy in region of hump 1,5 %, elsewhere within 1,0 %).
Reference: J. Lawrence, R. Parsons, Trans Faraday Soc. 64 (1968) 751.

Given are Potential and Capacitance as a function of surface charge σ.

	Salt = 0.05		0.07		0.10	
σ μ C/cm^2	E Volts	C μ F/cm^2	E Volts	C μ F/cm^2	E Volts	C μ F/cm^2
24					0.582	33.04
23					0.552	33.26
22					0.522	33.82
21					0.493	34.31
20					0.464	34.75
19					0.435	35.18
18	0.448	34.25	0.427	34.83	0.407	35.61
17	0.419	34.43	0.398	35.06	0.379	36.01
16	0.390	34.60	0.369	35.33	0.351	36.57
15	0.361	34.81	0.341	35.72	0.324	36.86
14	0.332	34.97	0.313	35.73	0.297	37.03
13	0.304	35.02	0.285	35.63	0.270	37.08
12	0.275	34.98	0.257	35.43	0.243	37.02
11	0.247	34.96	0.229	35.17	0.216	36.67
10	0.218	34.79	0.200	34.90	0.188	35.87
9	0.189	33.61	0.171	33.69	0.160	35.11
8	0.158	32.07	0.141	32.17	0.131	34.09
7	0.126	30.17	0.109	30.81	0.101	31.47
6	0.092	28.35	0.076	29.03	0.068	29.45
5	0.056	26.45	0.040	27.05	0.033	27.69
4	0.016	24.76	0.002	25.00	-0.004	25.83
3	-0.026	22.26	-0.040	22.68	-0.045	23.44
2	-0.075	19.00	-0.087	20.41	-0.090	21.13
1	-0.132	16.02	-0.139	18.21	-0.140	18.60
0	-0.197	14.84	-0.198	15.81	-0.198	16.28
-1	-0.266	14,47	-0.264	14.70	-0.262	14.91
-2	-0.336	13.71	-0.334	13.81	-0.332	13.89
-3	-0.412	12.76	-0.409	12.84	-0.407	12.90
-4	-0.494	11.74	-0.490	11.66	-0.488	11.78
-5	-0.583	10.73	-0.580	10.74	-0.576	10.90
-6	-0.679	9.98	-0.678	9.74	-0.673	9.82
-7	-0.786	8.98	-0.785	9.00	-0.779	9.06
-8					-0.894	8.43

Differential capacity of mercury in solutions of $NaClO_4$ in sulpholane (cont.)

	Salt = 0.2		0.35		0.5	
σ $\mu C/cm^2$	E Volts	C $\mu F/cm^2$	E Volts	C $\mu F/cm^2$	E Volts	C $\mu F/cm^2$
24	0.548	34.81	0.527	36.40	0.494	38.53
23	0.519	34.95	0.500	36.50	0.468	38.54
22	0.491	35.23	0.472	36.65	0.442	39.01
21	0.462	35.62	0.445	36.88	0.417	39.78
20	0.434	36.07	0.418	37.20	0.392	40.53
19	0.407	36.52	0.391	37.61	0.367	41.29
18	0.380	36.90	0.365	38.07	0.343	42.00
17	0.353	37.34	0.339	38.53	0.320	42.17
16	0.326	37.88	0.313	39.01	0.296	42.29
15	0.300	38.32	0.287	39.47	0.272	42.37
14	0.274	38.53	0.262	39.55	0.249	42.34
13	0.248	38.48	0.237	39.52	0.225	42.15
12	0.222	38.22	0.212	39.37	0.201	41.47
11	0.195	37.82	0.186	38.89	0.177	40.46
10	0.169	37.36	0.160	37.95	0.151	38.80
9	0.142	36.40	0.133	36.56	0.125	36.51
8	0.114	34.49	0.105	34.32	0.097	34.22
7	0.084	32.83	0.075	31.74	0.067	32.68
6	0.052	30.08	0.042	30.01	0.035	31.21
5	0.018	28.22	0.008	28.81	0.002	29.47
4	-0.019	26.66	-0.027	27.59	-0.033	27.61
3	-0.058	24.87	-0.065	25.91	-0.070	25.84
2	-0.099	22.98	-0.105	23.57	-0.110	24.57
1	-0.145	20.27	-0.149	21.43	-0.152	22.61
0	-0.198	18.00	-0.199	19.30	-0.199	20.00
-1	-0.257	16.09	-0.255	16.83	-0.252	17.61
-2	-0.323	14.62	-0.318	15.01	-0.312	15.72
-3	-0.395	13.20	-0.388	13.35	-0.381	13.72
-4	-0.474	11.97	-0.467	12.03	-0.458	12.19
-5	-0.562	10.85	-0.555	10.82	-0.546	10.93
-6	-0.659	9.93	-0.652	9.86	-0.641	9.96
-7	-0.764	9.11	-0.758	9.06	-0.746	9.15
-8	-0.877	8.54	-0.872	8.43	-0.861	8.43

Differential capacity of mercury in solutions of 0.1 mol l^{-1} KPF_6 in sulpholane + water mixtures. T = 30°C

Composition is given in volume percent of sulpholane. Potentials with respect to an aqueous saturated calomel electrode in contact with the working solution.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 751.

$C/\mu F\ cm^{-2}$

Vol.% S	0	0.50	0.10	5.0	30	50	80	99
E/V								
0.31		27.34				52.63	41.92	27.10
0.28	26.50[1]	25.92	28.00	33.00		46.16	37.21	26.64
0.25	24.10	25.23	26.89	32.20		38.98	34.66	26.40
0.22		24.50	26.42	30.40		34.65	32.89	25.79
0.19	22.52	23.99	25.51	29.72		33.33	30.59	24.86
0.16	22.00	23.43	24.97	28.78	35.75	32.84	29.11	24.09
0.13	21.65	23.14	24.39	28.01	30.89	30.60	27.66	22.94
0.10	21.42	22.92	24.26	28.11	29.41	28.69	26.36	22.25
0.07	21.40	22.84	24.24	28.15	28.47	27.60	24.97	21.40
0.04	21.45	22.86	24.35	27.91	27.59	27.60	23.88	20.55
0.01	21.70	23.08	24.61	27.43	26.82	25.73	22.88	19.63
-0.02	21.84	23.36	24.96	27.09	26.03	24.79	21.82	18.94
-0.05	22.10	23.69	25.29	26.65	25.13	23.82	20.78	18.09
-0.08	22.64	24.09	25.66	26.01	24.13	22.87	19.71	17.40
-0.11	23.10	24.47	25.91	25.25	23.18	21.81	18.69	16.63
-0.14	23.75	24.79	26.10	24.38	22.20	20.73	17.67	15.93
-0.17	24.42	25.09	26.09	23.45	21.18	19.70	16.76	15.47
-0.20	25.27	25.30	25.87	22.57	20.08	18.75	15.98	15.17
-0.23	25.84	25.36	25.51	21.60	19.08	17.81	15.40	14.70
-0.26	26.19	25.23	24.79	20.61	18.12	16.94	14.85	14.47
-0.29	26.45	24.96	23.87	19.66	17.19	16.14	14.41	14.16
-0.32	26.46	24.36	22.95	18.62	16.34	15.43	13.98	13.93
-0.35	26.21	23.69	21.83	17.64	15.49	14.76	13.63	13.63
-0.38	25.79	22.92	20.68	16.81	14.76	14.17	13.29	13.16

[1] at E = 0.30V

Capacity in sulpholane + water (continued)

	C/μF cm^{-2}							
Vol.% S	0	0.50	0.10	5.0	30	50	80	99
E/V								
-0.41	24.97	21.87	19.57	15.88	14.13	13.70	12.96	12.86
-0.44	24.07	20.80	18.46	15.06	13.53	13.27	12.66	12.47
-0.47	23.16	19.78	17.36	14.32	13.05	12.88	12.36	12.16
-0.50	22.45	18.79	16.41	13.65	12.55	12.51	12.05	11.85
-0.55	21.60	17.27	14.97	12.66	11.95	12.01	11.52	11.39
-0.60	20.78	15.99	13.80	11.85	11.44	11.53	11.03	10.85
-0.65	20.01	14.95	12.85	11.18	10.97	11.06	10.56	10.39
-0.70	19.26	14.16	12.17	10.64	10.57	10.64	10.12	9.78
-0.75	18.53	13.65	11.64	10.25	10.25	10.26	9.72	9.39
-0.80	17.88	13.36	11.29	9.90	9.95	9.94	9.40	9.08
-0.85	17.36	13.33	11.09	9.65	9.68	9.64	9.08	8.78
-0.90	16.95	13.44	11.05	9.42	9.47	9.39	8.82	8.54
-0.95	16.66	13.86	11.21	9.30	9.28	9.17	8.61	8.39
-1.00	16.49	14.39	11.58	9.20	9.08	8.97	8.40	
-1.05	16.42	15.13	12.23	9.15	8.98	8.83	8.27	
-1.10	16.45	15.87	13.22	9.20	8.87	8.72	8.17	
-1.15	10.56	16.64	14.73	9.32	8.80	8.64	8.11	
-1.20	16.74	17.30	16.56	9.59	8.80	8.62	8.07	
-1.25	16.98	17.82	18.61	10.05	8.83	8.66	8.08	
-1.30	17.28	18.22	20.13	10.88	8.97	8.76	8.19	
-1.35	17.63	18.51	21.03	12.26	9.17	8.98	8.40	
-1.40	17.99	18.78	21.27	14.65	9.58	9.31	8.78	
-1.45	18.42	19.00	21.31		10.26	9.93	9.32	
-1.50			21.27		11.21	10.83		

Differential capacity on mercury in aqueous solutions of 0.5 M Na_2SO_4 with different concentrations of camphor (C_c in μ moles.l^{-1}).
T = 25 ± 1°C. Potentials measured with respect to a $Hg/Hg_2SO_4(S)$, Na_2SO_4(0,5 M) electrode. (this reference electrode has a potential of 0.634 Volt with respect to the standard hydrogen electrode).
Reference: K.G. Baikerikov, S.Sathyanorayana, J. Electroanal. and Interfac. Electrochem. 24 (1970) 333.

c_c = 0		c_c = 54.6		c_c = 81.8	
E	C	E	C	E	C
Volts	μ F/m^2	Volts	μ F/m^2	Volts	μ F/m^2
-0.202	46.33	-0.200	45.81	-0.210	47.49
-0.252	43.15	-0.250	43.85	-0.250	44.86
-0.303	40.33	-0.300	41.30	-0.300	42.28
-0.353	39.24	-0.350	40.10	-0.350	40.75
-0.402	39.21	-0.400	39.09	-0.404	40.12
-0.451	39.44	-0.450	39.06	-0.454	39.31
-0.502	39.75	-0.500	39.87	-0.504	39.39
-0.552	39.63	-0.550	39.90	-0.550	39.67
-0.602	38.84	-0.600	39.82	-0.601	38.49
-0.651	37.22	-0.650	36.27	-0.657	36.39
-0.701	34.02	-0.700	33.51	-0.703	32.84
-0.751	31.03	-0.750	30.15	-0.730	39.64
-0.803	29.16	-0.800	27.46	-0.753	6.21
-0.852	26.93	-0.825	25.36	-0.802	5.40
-0.902	24.46	-0.850	5.76	-0.852	5.20
-0.953	22.37	-0.900	4.84	-0.902	5.00
-1.002	21.38	-0.950	4.59	-1.000	4.50
-1.052	20.11	-1.000	4.53	-1.100	4.58
-1.102	19.20	-1.050	4.46	-1.150	4.58
-1.152	18.37	-1.100	4.54	-1.200	4.54
-1.202	17.57	-1.125	4.64	-1.243	4.89
-1.252	17.02	-1.150	18.37	-1.260	17.07
-1.302	16.72	-1.200	17.45	-1.300	16.59
-1.353	16.50	-1.250	17.46	-1.350	16.32
-1.401	16.00	-1.300	16.88	-1.397	16.12
-1.452	16.00	-1.350	16.75	-1.447	16.17
-1.500	16.11	-1.40	16.34	-1.497	16.29
-1.552	16.01	-1.45	16.55	-1.550	16.51
-1.651	16.36	-1.50	16.28	-1.600	16.21
-1.702	16.51	-1.55	16.27	-1.650	16.44
-1.752	16.88	-1.60	16.10	-1.694	16.49
-1.803	17.48	-1.70	16.70	-1.800	17.31
-1.852	17.68	-1.80	17.32	-1.903	18.03
-1.902	18.10	-1.85	17.70		
-2.000	19.28	-1.90	18.16		

Double layer in camphor (cont.)

c_c = 136		c_c = 218		c_c = 352	
E Volts	C $\mu F/m^2$	E Volts	C $\mu F/m^2$	E Volts	C $\mu F/m^2$
-0.205	44.48	-0.200	48.34	-0.204	44.25
-0.256	41.91	-0.250	46.72	-0.257	41.88
-0.304	40.03	-0.300	43.11	-0.304	39.94
-0.352	39.17	-0.350	42.07	-0.354	38.59
-0.403	38.64	-0.400	40.70	-0.404	38.68
-0.453	38.77	-0.450	40.21	-0.455	39.68
-0.502	39.34	-0.500	40.61	-0.505	40.71
-0.554	39.72	-0.550	41.16	-0.537	41.37
-0.603	39.02	-0.600	41.14	-0.564	8.95
-0.640	37.25	-0.638	7.74	-0.600	7.94
-0.651	7.02	-0.650	7.64	-0.649	6.71
-0.702	5.82	-0.698	5.88	-0.702	5.68
-0.754	5.27	-0.758	5.51	-0.751	5.33
-0.804	4.87	-0.804	5.01	-0.800	4.83
-0.856	4.61	-0.856	4.77	-0.902	4.43
-0.904	4.44	-0.900	4.60	-1.001	4.26
-0.954	4.28	-0.950	4.45	-1.102	4.07
-1.006	4.16	-1.00	4.29	-1.202	4.00
-1.056	4.07	-1.050	4.19	-1.303	3.95
-1.102	3.96	-1.100	4.10	-1.402	3.94
-1.155	3.92	-1.150	4.07	-1.456	3.95
-1.204	3.91	-1.200	4.01	-1.480	4.28
-1.254	3.91	-1.250	4.01	-1.488	17.03
-1.303	4.05	-1.300	4.01	-1.500	16.89
-1.330	4.40	-1.350	4.01	-1.555	16.33
-1.348	16.36	-1.400	4.38	-1.604	16.27
-1.401	16.21	-1.420	16.62	-1.655	16.14
-1.458	16.03	-1.450	16.44	-1.706	16.23
-1.510	16.43	-1.500	16.31	-1.801	16.78
-1.603	16.08	-1.550	16.38	-1.902	17.98
-1.705	16.49	-1.600	16.16		
-1.807	17.10	-1.700	16.46		
-1.908	17.92	-1.800	17.20		
		-1.900	17.95		

Double layer in camphor (cont.)

c_c = 880		c_c = 1761		c_c = saturated	
E	C	E	C	E	C
Volts	$\mu F/m^2$	Volts	$\mu F/m^2$	Volts	$\mu F/m^2$
-0.203	44.45	-0.250	42.07	-0.206	45.44
-0.253	41.88	-0.301	41.20	-0.253	43.22
-0.303	40.55	-0.351	39.25	-0.301	41.63
-0.354	39.61	-0.401	41.04	-0.352	41.05
-0.401	40.50	-0.420	42.20	-0.383	43.01
-0.451	41.18	-0.448	10.33	-0.3840	89.10
-0.470	41.95	-0.501	8.49	-0.3870	343.00
-0.503	9.17	-0.552	7.55	-0.3886	158.40
-0.553	7.82	-0.601	6.78	-0.3900	148.40
-0.603	6.91	-0.645	6.37	-0.396	55.38
-0.652	6.29	-0.701	5.72	-0.415	10.70
-0.703	5.74	-0.800	5.03	-0.425	10.30
-0.752	5.25	-0.900	4.60	-0.450	9.87
-0.800	5.09	-1.002	4.25	-0.475	9.31
-0.900	4.48	-1.103	4.05	-0.500	9.06
-1.003	4.19	-1.202	3.92	-0.550	8.08
-1.102	4.00	-1.302	3.90	-0.600	7.48
-1.202	3.95	-1.402	3.90	-0.652	6.61
-1.302	3.95	-1.504	3.90	-0.702	6.15
-1.401	3.95	-1.604	3.90	-0.752	5.82
-1.500	3.90	-1.652	3.90	-0.802	5.48
-1.551	3.96	-1.665	17.90	-0.851	5.09
-1.590	4.15	-1.700	17.52	-0.900	4.75
-1.602	17.21	-1.750	17.25	-0.950	4.54
-1.651	17.09	-1.800	17.39	-1.000	4.37
-1.700	16.93	-1.902	17.93	-1.050	4.28
-1.751	16.88	-2.000	19.20	-1.101	4.25
-1.802	17.20			-1.152	4.18
-1.905	18.20			-1.200	4.10
-2.000	19.30			-1.252	4.11
				-1.302	4.09
				-1.403	3.94
				-1.453	3.92
				-1.502	4.00
				-1.553	3.97
				-1.602	3.95
				-1.651	3.95
				-1.701	3.99
				-1.752	4.01
				-1.760	4.08
				-1.771	4.66
				-1.7743	32.45
				-1.7750	44.80
				-1.7770	72.76
				-1.780	18.91
				-1.801	18.46
				-1.851	17.68
				-1.902	18.33
				-2.00	19.03

Differential capacity on mercury in aqueous solutions of 0.5 *M* Na_2SO_4 with different concentrations of naphthalene (C_n in μ moles.l^{-1}). $T = 25 \pm 1°C$. Potentials measured with respect to a Hg/Hg_2SO_4 (S), Na_2SO_4 (0,5 M) electrode. (This reference electrode has a potential of 0.634 Volt with respect to the standard hydrogen electrode).

Reference: K.G. Baikerikov, S. Sathyanorayana, J. Electroanal. and Interfac. Electrochem. 24 (1970) 333.

$c_n = 16.4$		$c_n = 32.8$		c_n = saturated	
E	C	E	C	E	C
Volts	μ F/m^2	Volts	μ F/m^2	Volts	μ F/m^2
0.202	47.60	0.254	45.13	0.102	78.48
0.252	44.59	0.302	42.91	0.153	54.56
0.302	42.25	0.352	39.87	0.202	48.46
0.352	40.23	0.402	37.50	0.253	46.40
0.402	37.87	0.452	35.32	0.303	43.50
0.452	35.54	0.501	30.63	0.352	41.33
0.502	32.63	0.552	24.60	0.402	38.83
0.552	27.51	0.601	19.06	0.453	35.33
0.602	21.82	0.651	15.17	0.502	29.45
0.652	17.87	0.702	13.23	0.552	21.56
0.702	15.67	0.752	12.30	0.602	15.70
0.752	14.45	0.801	11.40	0.652	12.97
0.802	13.44	0.852	10.92	0.702	11.27
0.852	13.15	0.902	10.75	0.752	10.09
0.902	13.00	0.952	10.58	0.802	9.40
0.952	12.77	1.001	10.73	0.852	8.93
1.002	12.83	1.052	11.29	0.902	8.48
1.052	13.20	1.101	12.05	0.952	8.06
1.102	13.90	1.152	13.11	1.002	8.31
1.152	14.68	1.202	13.93	1.052	8.58
1.202	15.47	1.252	15.50	1.102	9.21
1.252	16.06	1.303	16.47	1.152	10.10
1.302	16.61	1.352	17.18	1.200	11.84
1.352	16.84	1.402	17.77	1.252	14.17
1.402	17.09	1.452	17.93	1.303	16.34
1.452	16.97	1.502	17.51	1.325	17.40
1.502	16.92	1.552	16.83	1.353	18.50
1.552	16.64	1.602	16.53	1.375	19.44
1.602	16.06	1.651	16.58	1.402	19.88
1.652	16.20	1.702	16.64	1.425	20.40
1.702	16.82	1.752	16.74	1.452	20.41
1.752	16.91	1.803	17.48	1.475	20.00
1.802	17.42	1.902	18.09	1.503	19.59
1.852	17.74			1.525	18.60
1.902	18.01			1.552	18.18
				1.574	17.72
				1.602	17.13
				1.653	16.87
				1.702	16.70
				1.752	16.88
				1.802	17.18
				1.852	17.62
				1.901	18.04
				2.00	18.99

Differential capacity of mercury in 0.1 mol l^{-1} salt in N-methyl acetamide
T = 30°C

Potential with respect to aqueous 1 mol l^{-1} calomel electrode.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 35.

Salt σ /μC cm^{-2}	KNO_3 E/volt	 C/μF cm^{-2}	$LiBF_4$ * E/volt	 C/μF cm^{-2}	Pr_4NBr E/volt	 C/μF cm^{-2}	$LiClO_4$ ** E/volt	 C/μF cm^{-2}
-16	-1.789	6.60						
-15	-1.634	6.49	-1.639	6.76	-1.720	6.85		
-14	-1.486	7.12	-1.496	7.37	-1.575	7.00	-1.496	7.52
-13	-1.357	8.53	-1.370	8.66	-1.437	7.55	-1.373	8.02
-12	-1.252	10.62	-1.265	10.58	-1.311	8.41	-1.271	10.90
-11	-1.166	12.78	-1.179	12.55	-1.198	9.37	-1.187	12.83
-10	-1.092	14.41	-1.103	14.01	-1.096	10.16	-1.113	14.21
- 9	-1.026	15.42	-1.034	14.85	-1.001	10.88	-1.045	14.97
- 8	-0.962	15.86	-0.968	15.16	-0.913	11.77	-0.979	15.26
- 7	-0.899	15.69	-0.902	15.07	-0.832	13.00	-0.913	15.13
- 6	-0.834	15.11	-0.835	14.64	-0.759	14.58	-0.846	14.66
- 5	-0.766	14.16	-0.765	13.85	-0.694	16.24	-0.776	13.86
- 4	-0.692	12.99	-0.690	12.79	-0.636	18.05	-0.701	12.75
- 3	-0.611	11.69	-0.607	11.55	-0.584	20.17	-0.619	11.56
- 2	-0.521	11.61	-0.516	10.44	-0.538	23.20	-0.528	10.48
- 1	-0.425	10.35	-0.417	9.99	-0.498	28.01	-0.430	10.15
0	-0.332	11.46	-0.320	10.78	-0.466	34.22	-0.335	11.19
1	-0.252	13.91	-0.234	12.83	-0.439	41.56	-0.253	13.54
2	-0.187	17.02	-0.164	15.67	-0.417	49.43	-0.187	16.72
3	-0.134	20.43	-0.106	18.87	-0.399	57.35	-0.133	20.37
4	-0.089	24.04	-0.056	20.93	-0.382	65.28	-0.087	24.12
5	-0.050	27.82	-0.011	25.56	-0.368	73.06	-0.049	27.87
6	-0.016	31.55	0.024	29.92	-0.355	80.68	-0.015	31.70
7	0.014	35.19	0.056	33.03	-0.342	89.47	0.014	35.31
8	0.041	38.69	0.085	35.61	-0.332	98.46	0.042	38.52
9	0.066	41.96	0.112	37.26	-0.322	107.5	0.066	40.92
10	0.089	44.98	0.138	37.27	-0.313	116.4	0.090	42.61
11	0.110	47.73	0.166	35.73	-0.305	125.3	0.113	43.68
12	0.131	50.17	0.195	32.69	-0.294	145.3	0.136	43.80
13	0.150	52.40	0.227	29.51	-0.287	162.9	0.159	43.24
14	0.169	54.74	0.263	27.12	-0.282	180.6	0.183	42.62
15	0.187	57.03	0.300	27.34	-0.276	198.1	0.206	42.08
16	0.204	59.18	0.336	29.14	-0.272	215.5	0.230	41.66
17	0.220	60.89	0.368	31.54	-0.267	232.6	0.254	41.32
18	0.237	62.38	0.399	33.89	-0.263	249.4	0.278	40.93
19	0.252	63.64			-0.259	265.9	0.302	40.53
20	0.268	64.65			-0.255	282.1	0.328	40.07

Capacity in 0.1 mol l^{-1} salts in N-methyl acetamide (cont.)
T = 30°C

salt	Et_4NClO_4		CsI		NH_4Cl	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-20			-1.860	50.66		
-19	-2.039	7.71	-1.834	36.81		
-18	-1.908	7.59	-1.803	28.07		
-17	-1.776	7.60	-1.760	20.22		
-16	-1.646	7.78	-1.699	13.18	-1.739	7.83
-15	-1.520	8.19	-1.603	8.75	-1.606	7.45
-14	-1.402	8.81	-1.479	7.85	-1.475	7.93
-13	-1.294	9.62	-1.358	8.94	-1.357	9.19
-12	-1.194	10.59	-1.257	10.92	-1.258	11.09
-11	-1.104	11.64	-1.173	12.99	-1.175	13.01
-10	-1.022	12.74	-1.101	14.51	-1.102	14.45
-9	-0.947	13.92	-1.034	15.47	-1.035	15.25
-8	-0.818	15.17	-0.971	16.07	-0.970	15.49
-7	-0.815	16.28	-0.910	16.67	-0.905	15.30
-6	-0.755	17.06	-0.852	17.87	-0.839	14.73
-5	-0.697	17.18	-0.799	20.29	-0.769	13.90
-4	-0.638	16.51	-0.754	24.22	-0.694	12.89
-3	-0.574	15.05	-0.716	29.79	-0.614	11.92
-2	-0.504	13.30	-0.686	35.25	-0.528	11.49
-1	-0.424	11.95	-0.660	41.62	-0.444	12.66
0	-0.340	12.15	-0.638	49.10	-0.374	16.40
1	-0.263	14.07	-0.619	57.00	-0.322	22.43
2	-0.198	17.07	-0.603	64.25	-0.283	29.81
3	-0.145	20.64	-0.588	71.17	-0.253	37.83
4	-0.100	24.46	-0.575	77.79	-0.229	46.90
5	-0.062	27.85	-0.562	84.17	-0.209	55.88
6	-0.028	31.16	-0.551	90.33	-0.191	67.34
7	0.002	34.45	-0.540	96.24	-0.178	79.81
8	0.030	37.55	-0.530	101.98	-0.166	92.21
9	0.055	40.36	-0.520	107.6	-0.156	104.36
10	0.080	42.35	-0.511	113.0		
11	0.103	43.78	-0.503	118.3		
12	0.125	44.16	-0.494	125.9		
13	0.148	43.11	-0.486	133.7		
14	0.171	42.27	-0.479	141.6		
15	0.195	41.65	-0.472	149.7		
16	0.219	41.33	-0.465	157.9		

* some time dependence of the pzc

** d.c. at E 0.6V but no polarisation resistance or pseudo-capacity

Differential capacity of 0.1 mol l^{-1} KPF_6 in various solvents. T = 25°C
Potential with respect to aqueous 0.1 mol l^{-1} calomel electrode.
(* T = 40°C for ethylene carbonate)

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 35.

Solvent	N-tert-butyl formamide		N-butyl acetamide		Valerolactone	
σ /μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-16	-1.748	20.26				
-15	-1.691	16.53				
-14	-1.627	14.70				
-13	-1.556	13.66				
-12	-1.481	13.15				
-11	-1.403	12.54				
-10	-1.320	11.56	-1.475	5.42		
- 9	-1.229	10.33	-1.284	5.22		
- 8	-1.126	9.20	-1.104	6.19	-1.028	7.18
- 7	-1.013	8.54	-0.969	9.17	-0.894	7.86
- 6	-0.894	8.39	-0.876	12.40	-0.774	8.89
- 5	-0.777	8.69	-0.800	13.45	-0.669	10.16
- 4	-0.666	9.48	-0.723	12.65	-0.577	11.55
- 3	-0.566	10.65	-0.640	11.09	-0.494	12.62
- 2	-0.476	11.57	-0.542	9.44	-0.418	13.46
- 1	-0.392	12.19	-0.428	8.37	-0.345	13.83
0	-0.313	13.20	-0.310	8.92	-0.273	14.18
1	-0.243	15.70	-0.209	11.27	-0.204	15.07
2	-0.186	20.03	-0.131	14.70	-0.141	16.60
3	-0.142	25.49	-0.071	18.37	-0.084	18.32
4	-0.106	31.52	-0.021	21.70	-0.032	19.92
5	-0.077	37.89	0.022	24.23	0.017	21.15
6	-0.053	44.29	0.062	25.35	0.063	22.19
7	-0.032	51.47	0.102	24.71	0.107	23.29
8	-0.013	58.72	0.144	22.93	0.149	24.42
9	0.003	66.48	0.189	21.98	0.189	25.43
10	0.017	75.19	0.234	22.84	0.228	26.23
11	0.030	84.00	0.276	25.17	0.266	26.79
12	0.041	92.80	0.316	24.73	0.303	27.13
13	0.054	105.3	0.356	24.80	0.340	27.12
14	0.063	119.7	0.396	25.25	0.377	26.81
15	0.070	134.1	0.435	25.58	0.414	26.22
16	0.078	148.6			0.453	25.44
17	0.084	162.8			0.493	24.53
18	0.090	176.9			0.535	23.53
19	0.095	190.8			0.578	22.56
20					0.623	21.75

Capacity of 0.1 mol l^{-1} KPF_6 in various solvents (cont.)

Solvent	Propylene carbonate		Ethylene carbonate*	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-16				
-15			-1.525	10.90
-14			-1.434	11.21
-13			-1.346	11.45
-12			-1.259	11.67
-11			-1.174	11.92
-10			-1.091	12.16
- 9			-1.010	12.39
- 8			-0.930	12.63
- 7			-0.851	12.85
- 6			-0.774	13.10
- 5			-0.698	13.24
- 4	-0.665	12.85	-0.623	13.37
- 3	-0.588	13.21	-0.549	13.55
- 2	-0.513	13.47	-0.476	13.79
- 1	-0.440	13.77	-0.404	14.10
0	-0.368	14.15	-0.334	14.64
1	-0.300	14.82	-0.267	15.38
2	-0.233	15.69	-0.204	16.35
3	-0.171	16.71	-0.145	17.41
4	-0.114	17.83	-0.089	18.46
5	-0.059	18.98	-0.037	19.55
6	-0.008	20.03	0.013	20.61
7	-0.041	21.13	0.061	21.53
8	0.087	22.12	0.106	22.35
9	0.131	22.84	0.150	23.09
10	0.174	23.33	0.193	23.66
11	0.217	23.69	0.235	24.08
12	0.259	24.00	0.276	24.33
13	0.301	24.13	0.317	24.43
14	0.342	24.33	0.358	24.42
15	0.383	24.40	0.399	24.44
16	0.424	24.31	0.440	24.65
17	0.465	24.12	0.480	24.66
18	0.507	23.91	0.521	24.40
19	0.549	23.67	0.562	24.15
20	0.591	23.46	0.604	23.90

Differential capacity of mercury in KPF_6 in N-methyl acetamide.

T = 30.5°C

Pottential with respect to aqueous 1 mol l^{-1} calomel electrode.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 35.

$c/mol\ l^{-1}$	0.010		0.10		1.0	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-15	-1.808	6.24				
-14	-1.649	6.49	-1.501	7.20		
-13	-1.504	7.43	-1.373	8.62	-1.352	9.09
-12	-1.381	9.12	-1.269	10.76	-1.252	10.89
-11	-1.283	11.26	-1.184	12.89	-1.167	12.67
-10	-1.201	13.10	-1.111	14.50	-1.092	14.13
- 9	-1.128	14.27	-1.045	15.52	-1.024	15.20
- 8	-1.059	14.74	-0.981	15.96	-0.960	15.86
- 7	-0.991	14.64	-0.919	15.83	-0.897	16.14
- 6	-0.922	14.01	-0.854	15.21	-0.835	15.99
- 5	-0.848	13.03	-0.786	14.24	-0.772	15.37
- 4	-0.767	11.76	-0.713	12.97	-0.705	14.34
- 3	-0.677	10.30	-0.632	11.58	-0.632	13.11
- 2	-0.572	8.93	-0.540	10.45	-0.552	12.05
- 1	-0.455	8.20	-0.442	10.03	-0.467	11.65
0	-0.333	8.50	-0.345	10.85	-0.383	12.38
1	-0.225	10.32	-0.260	12.83	-0.307	14.28
2	-0.139	13.12	-0.189	15.55	-0.243	17.01
3	-0.071	16.55	-0.130	18.62	-0.189	20.21
4	-0.016	19.99	-0.080	21.81	-0.143	23.67
5	0.030	23.38	-0.038	25.20	-0.104	27.37
6	0.070	26.84	-0.001	28.82	-0.070	31.15
7	0.105	30.26	0.032	32.08	-0.040	34.81
8	0.137	32.28	0.062	34.92	-0.012	38.14
9	0.167	33.07	0.090	36.92	0.013	40.78
10	0.197	32.25	0.116	37.28	0.037	42.25
11	0.229	29.86	0.144	35.89	0.060	41.97
12	0.266	26.31	0.172	33.67	0.085	39.93
13	0.306	23.95	0.205	29.42	0.111	36.51
14	0.349	23.00	0.241	26.14	0.140	32.62
15			0.282	23.96	0.173	28.97
16			0.324	23.31	0.209	26.67
17			0.366	23.41	0.248	25.51
18					0.287	24.93

Differential capacity of mercury in 0.1 mol l^{-1} KPF_6 in H_2O +
N-methyl acetamide. T = 25°C

Potential with respect to 1 mol l^{-1} KCl calomel electrode.

Reference R. Payne. Advances in Electrochemistry _7_ (1970) 37.

x/mol % NMA	0.405		2.97		14.07		56.1	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-20	-1.685	20.64						
-19	-1.636	20.12						
-18	-1.586	19.67						
-17	-1.534	19.29						
-16	-1.482	18.93			-1.705	9.38	-1.702	7.38
-15	-1.429	18.64			-1.594	8.99	-1.569	7.66
-14	-1.375	18.32			-1.486	9.47	-1.444	8.43
-13	-1.319	17.94			-1.385	10.44	-1.333	9.76
-12	-1.263	17.46			-1.294	11.78	-1.239	11.55
-11	-1.205	16.83	-1.283	13.64	-1.214	13.33	-1.158	13.44
-10	-1.144	16.03	-1.211	14.17	-1.143	14.89	-1.088	14.98
- 9	-1.080	15.16	-1.141	14.51	-1.079	16.04	-1.024	16.05
- 8	-1.012	14.49	-1.072	14.38	-1.018	16.50	-0.963	16.56
- 7	-0.942	14.05	-1.002	13.77	-0.957	16.23	-0.902	16.45
- 6	-0.870	13.90	-0.927	13.07	-0.894	15.35	-0.840	15.78
- 5	-0.798	13.98	-0.848	12.35	-0.826	14.17	-0.775	14.75
- 4	-0.727	14.25	-0.766	11.94	-0.752	13.02	-0.704	13.61
- 3	-0.659	14.84	-0.681	11.79	-0.673	12.27	-0.628	12.50
- 2	-0.593	15.80	-0.598	12.27	-0.591	12.13	-0.545	11.75
- 1	-0.532	16.99	-0.521	13.27	-0.509	12.58	-0.459	11.60
0	-0.475	18.10	-0.447	13.40	-0.432	13.43	-0.374	12.14
1	-0.421	18.91	-0.374	13.91	-0.360	14.47	-0.295	13.37
2	-0.369	19.37	-0.303	14.43	-0.294	15.50	-0.225	15.08
3	-0.318	19.53	-0.235	14.87	-0.231	16.39	-0.162	17.11
4	-0.266	19.51	-0.169	15.24	-0.171	17.14	-0.107	19.26
5	-0.215	19.40	-0.104	15.63	-0.114	17.73	-0.058	21.57
6	-0.163	19.30	-0.041	16.13	-0.058	18.29	-0.014	23.86
7	-0.111	19.27	0.020	16.87	-0.005	18.86	0.026	26.15
8	-0.060	19.34	0.078	17.90	0.048	19.56	0.062	28.41
9	-0.008	19.51	0.132	19.11	0.098	20.41	0.016	30.39
10	0.043	19.78	0.182	20.54	0.146	21.51	0.128	32.09
11	0.093	20.19	0.229	22.66	0.191	22.90	0.159	33.41
12	0.142	20.74	0.271	26.00	0.233	24.53	0.188	33.80
13	0.189	21.53			0.272	26.21	0.218	33.31
14	0.234	22.80					0.249	32.03
15	0.277	25.04					0.281	30.66
16							0.314	29.19
17							0.349	27.93

Differential capacity of mercury in 0.1 mol l^{-1} salt in N-ethyl formamide.
T = 25°C

Potential with respect to aqueous 1 mol l^{-1} KCl calomel electrode.

Reference : R. Payne. J. Physical Chem. 73 (1968) 3598.

Salt	CsI		KPF_6		$LiClO_4$	
σ / $\mu C\ cm^{-2}$	E/volt	C/$\mu F\ cm^{-2}$	E/volt	C/$\mu F\ cm^{-2}$	E/volt	C/$\mu F\ cm^{-2}$
-16	-1.599	16.58				
-15	-1.534	14.15	-1.591	11.70	-1.554	11.23
-14	-1.458	12.16	-1.498	10.08	-1.461	10.48
-13	-1.370	10.91	-1.395	9.51	-1.365	10.37
-12	-1.276	10.56	-1.290	9.80	-1.269	10.71
-11	-1.183	11.05	-1.192	10.69	-1.179	11.44
-10	-1.096	12.03	-1.104	11.92	-1.095	12.43
- 9	-1.017	13.25	-1.024	13.16	-1.018	13.42
- 8	-0.945	14.48	-0.951	14.10	-0.945	14.13
- 7	-0.879	16.02	-0.881	14.48	-0.875	14.39
- 6	-0.822	18.93	-0.812	14.33	-0.805	14.19
- 5	-0.775	23.70	-0.741	13.83	-0.734	13.68
- 4	-0.737	29.82	-0.667	13.24	-0.659	13.13
- 3	-0.707	36.57	-0.590	12.77	-0.582	12.73
- 2	-0.682	43.33	-0.511	12.59	-0.502	12.60
- 1	-0.661	50.29	-0.432	12.76	-0.424	12.86
0	-0.642	57.31	-0.355	13.43	-0.348	13.69
1	-0.626	64.37	-0.283	14.65	-0.278	15.11
2	-0.611	71.10	-0.219	16.54	-0.216	17.17
3	-0.597	77.62	-0.163	19.30	-0.162	19.85
4	-0.585	83.93	-0.115	23.18	-0.115	23.14
5	-0.574	90.07	-0.076	28.24	-0.075	27.02
6	-0.563	96.04	-0.043	34.21	-0.040	31.54
7	-0.553	101.87	-0.016	41.02	-0.011	36.80
8	-0.543	107.82	0.006	48.26	0.015	42.87
9	-0.534	113.71	0.026	56.03	0.037	49.51
10	-0.525	119.52	0.042	63.81	0.056	56.88
11	-0.517	125.3	0.058	72.62	0.072	65.21
12	-0.509	130.9	0.071	81.94	0.087	73.67
13	-0.502	136.6	0.082	91.28	0.100	82.12
14			0.093	100.54	0.113	98.37
15					0.122	112.65
16					0.131	126.88
17					0.138	140.9
18					0.145	154.7

Differential capacity of mercury in 0.1 mol l^{-1} KPF_6 in various solvents.

T = 25°C

Potentials with respect to aqueous 1.0 mol l^{-1} KCl calomel electrode.

Reference: R. Payne, J. Phys. Chem., 73 (1968) 3598.

Solvent	Formamide		Dimethyl formamide		N-methyl propionamide		N-methyl formamide	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-17	-1.542	13.65						
-16	-1.470	14.04					-1.417	10.99
-15	-1.400	14.57					-1.330	12.30
-14	-1.333	15.24			-1.797	6.54	-1.254	13.89
-13	-1.269	16.02			-1.627	5.60	-1.186	15.53
-12	-1.208	16.87	-1.494	6.91	-1.449	5.75	-1.124	16.93
-11	-1.150	17.66	-1.347	6.78	-1.287	6.79	-1.067	17.85
-10	-1.095	18.30	-1.202	7.01	-1.159	9.20	-1.011	18.13
-9	-1.041	18.68	-1.063	7.50	-1.065	12.19	-0.956	17.87
-8	-0.987	18.75	-0.936	8.38	-0.990	14.17	-0.899	17.16
-7	-0.934	18.49	-0.824	9.42	-0.922	14.99	-0.839	16.19
-6	-0.879	17.95	-0.724	10.54	-0.855	15.00	-0.775	15.18
-5	-0.822	17.17	-0.634	11.84	-0.788	14.46	-0.707	14.29
-4	-0.762	16.29	-0.554	13.02	-0.716	13.39	-0.635	13.61
-3	-0.699	15.43	-0.480	14.04	-0.637	11.89	-0.561	13.21
-2	-0.633	14.69	-0.411	14.76	-0.547	10.37	-0.485	13.16
-1	-0.563	14.13	-0.344	15.26	-0.445	9.50	-0.409	13.48
0	-0.492	14.10	-0.280	16.09	-0.342	10.16	-0.337	14.20
1			-0.220	17.20	-0.252	12.33	-0.269	15.31
2			-0.164	18.54	-0.179	15.36	-0.207	16.80
3			-0.112	20.20	-0.120	18.88	-0.150	18.56
4			-0.065	21.85	-0.072	22.68	-0.100	20.60
5			-0.020	23.34	-0.031	26.39	-0.053	23.03
6	-0.116	19.58	0.021	24.62	0.004	29.75	-0.012	25.79
7	-0.067	21.07	0.061	25.74	0.036	31.75	0.025	29.00
8	-0.021	22.61	0.099	26.72	0.068	31.98	0.058	32.73
9	0.022	24.20	0.136	27.64	0.100	30.07	0.087	37.19
10	0.062	25.80	0.172	28.50	0.135	26.65	0.112	42.12
11	0.099	27.41	0.206	29.32	0.176	22.95	0.134	47.44
12	0.135	29.03	0.240	30.04	0.222	21.19	0.154	52.93
13	0.166	30.51	0.273	30.74	0.269	21.79	0.172	58.56
14	0.198	32.08			0.314	22.82	0.189	64.21
15	0.228	33.74			0.356	23.70	0.204	70.00
16	0.257	35.43			0.398	24.19	0.217	76.06
17	0.285	37.00			0.439	24.45	0.230	82.16
18	0.311	38.52			0.480	24.55	0.242	88.24
19	0.337	40.03						

Differential capacity of mercury in 0.1 mol l^{-1} salts in N-ethyl acetamide. T = 25°C

Potential with respect to aqueous 1 mol l^{-1} KCl calomel electrode.

Reference : R. Payne. J. Physical Chem. 73 (1968) 3598.

Salt	KPF_6		CsI	
σ /μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-21			-1.842	51.26
-20			-1.821	44.85
-19			-1.796	39.42
-18			-1.770	36.58
-17			-1.741	33.23
-16			-1.709	29.23
-15			-1.672	24.37
-14			-1.626	18.77
-13	-1.681	5.78	-1.562	12.97
-12	-1.502	5.64	-1.465	8.54
-11	-1.333	6.35	-1.333	7.15
-10	-1.193	8.16	-1.201	8.38
-9	-1.086	10.79	-1.095	10.79
-8	-1.002	12.97	-1.012	13.00
-7	-0.929	14.12	-0.939	14.51
-6	-0.859	14.35	-0.874	16.11
-5	-0.788	13.88	-0.816	18.86
-4	-0.714	12.89	-0.769	23.22
-3	-0.632	11.53	-0.730	28.39
-2	-0.539	10.10	-0.698	34.50
-1	-0.435	9.38	-0.672	41.09
0	-0.331	10.05	-0.649	47.94
1	-0.240	12.26	-0.630	55.17
2	-0.167	15.46	-0.613	61.80
3	-0.109	19.14	-0.597	67.99
4	-0.062	22.75	-0.583	73.84
5	-0.021	25.81	-0.570	79.41
6	0.016	28.35	-0.558	84.75
7	0.051	30.71	-0.546	90.02
8	0.084	30.86	-0.535	95.31
9	0.105	30.80	-0.525	100.50
10	0.139	28.64	-0.516	105.6
11	0.175	26.20	-0.506	110.6
12	0.216	24.35	-0.496	118.1
13	0.257	24.49	-0.488	127.0
14	0.296	25.48	-0.480	136.4
15	0.338	22.71	-0.473	146.1
16	0.383	21.72	-0.467	155.9
17	0.428	21.17	-0.460	165.9
18			-0.454	175.8

Differential capacity of mercury in 0.1 mol l^{-1} KPF_6 in water + butyrolactone. T = 25°C

Potential with respect to 1 mol l^{-1} aqueous KCl calomel electrode.

Reference: R. Payne, Advances in Electrochemistry, 7 (1970) 49.

x/mol% lactone	3.20		0.40	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-17			-1.496	21.05
-16			-1.448	20.99
-15			-1.400	20.88
-14	-1.580	21.32	-1.352	20.51
-13	-1.522	14.52	-1.302	19.75
-12	-1.440	10.70	-1.250	18.51
-11	-1.337	9.11	-1.194	16.80
-10	-1.224	8.63	-1.131	14.93
- 9	-1.108	8.68	-1.060	13.33
- 8	-0.995	9.07	-0.981	12.38
- 7	-0.888	9.71	-0.899	12.14
- 6	-0.789	10.58	-0.818	12.47
- 5	-0.699	11.64	-0.740	13.24
- 4	-0.617	12.91	-0.667	14.34
- 3	-0.544	14.40	-0.600	15.66
- 2	-0.478	16.11	-0.539	17.23
- 1	-0.419	17.76	-0.484	18.85
0	-0.365	19.42	-0.433	20.45
1	-0.316	20.98	-0.386	21.86
2	-0.269	22.37	-0.341	23.02
3	-0.226	23.58	-0.299	23.91
4	-0.184	24.61	-0.257	24.53
5	-0.145	25.49	-0.217	24.95
6	-0.106	26.28	-0.177	25.22
7	-0.068	26.98	-0.138	25.35
8	-0.032	27.59	-0.098	25.33
9	0.004	28.11	-0.059	25.23
10	0.040	28.52	-0.019	25.03
11	0.074	28.80	0.021	24.77
12	0.109	28.95	0.062	24.53
13	0.144	28.96	0.103	24.32
14	0.178	28.88	0.144	24.30
15	0.213	28.84	0.185	24.50
16	0.247	29.04	0.226	25.07
17	0.282	29.87	0.265	26.25
18	0.315	32.57	0.302	28.37
19	0.343	38.30	0.335	32.28

Differential capacity of mercury in KPF_6 in dimethyl acetamide.

T = 25°C

Potential with respect to aqueous 1 mol l^{-1} calomel electrode.

Reference: R. Payne, Advances in Electrochemistry, 7 (1971) 37.

c/mol l^{-1}	0.010		0.10		Saturated	
σ /μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-12			-1.559	6.04		
-11			-1.392	6.02		
-10	-1.242	6.00	-1.229	6.31		
- 9	-1.081	6.59	-1.075	6.70		
- 8	-0.942	8.08	-0.933	7.51		
- 7	-0.824	8.38	-0.808	8.48		
- 6	-0.712	9.61	-0.699	10.03		
- 5	-0.615	11.08	-0.607	11.87		
- 4	-0.530	12.25	-0.526	12.66		
- 3	-0.451	13.05	-0.450	13.79	-0.487	14.99
- 2	-0.375	13.20	-0.380	14.80	-0.425	17.14
- 1	-0.296	11.68	-0.314	15.62	-0.370	19.21
0	-0.203	10.06	-0.252	16.59	-0.320	21.19
1	-0.112	13.04	-0.194	17.92	-0.275	23.10
2	-0.047	17.19	-0.140	19.43	-0.233	24.93
3	0.008	19.55	-0.091	21.02	-0.195	26.61
4	0.056	21.26	-0.045	22.27	-0.158	28.00
5	0.102	22.98	-0.001	23.03	-0.123	29.04
6	0.144	24.75	0.043	22.92	-0.089	29.73
7	0.183	25.85	0.087	22.31	-0.056	30.16
8	0.221	26.24	0.133	21.27	-0.023	30.33
9	0.260	26.17	0.180	21.10	0.011	30.30
10	0.298	26.31	0.227	22.13	0.044	30.22
11	0.336	26.37	0.270	23.56	0.077	30.33
12	0.374	26.33	0.311	24.93	0.109	31.45
13	0.412	26.21	0.351	25.91	0.140	32.88
14	0.450	26.32	0.389	26.50		
15			0.426	26.85		

Differential capacity of mercury in 0.1 mol l^{-1} KPF_6 in $H_2O + (CH_3)_2SO$

T = 25°C

Potentials with respect to 1 mol l^{-1} KCl calomel electrode.

Reference: R. Payne . J. Amer. Soc., 89 (1967) 489.

x/mol % $(CH_3)_2SO$	0		2.824		20.95		69.54	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20			-1.677	25.59				
-19			-1.637	24.29				
-18			-1.594	22.59	-1.852	24.90		
-17	-1.434	17.23	-1.548	20.53	-1.805	18.09		
-16	-1.375	16.81	-1.496	18.33	-1.739	13.17		
-15	-1.315	16.46	-1.438	16.10	-1.653	10.58		
-14	-1.254	16.19	-1.372	14.17	-1.552	9.47		
-13	-1.192	16.04	-1.297	12.73	-1.443	9.03	-1.557	7.55
-12	-1.129	16.00	-1.216	11.86	-1.332	8.93	-1.427	7.78
-11	-1.067	16.10	-1.129	11.48	-1.220	9.05	-1.300	8.07
-10	-1.005	16.33	-1.042	11.55	-1.112	9.38	-1.180	8.51
- 9	-0.944	16.75	-0.957	12.03	-1.008	9.98	-1.066	9.13
- 8	-0.886	17.29	-0.877	12.84	-0.912	10.86	-0.961	10.02
- 7	-0.829	17.94	-0.802	13.97	-0.824	12.07	-0.867	11.21
- 6	-0.774	18.65	-0.734	15.30	-0.746	13.50	-0.766	12.06
- 5	-0.722	19.36	-0.671	16.74	-0.676	15.06	-0.692	14.50
- 4	-0.671	20.03	-0.614	17.19	-0.612	16.58	-0.626	15.90
- 3	-0.622	20.64	-0.564	-	-0.555	18.01	-0.565	17.00
- 2	-0.574	21.30	-0.534	-	-0.499	18.44	-0.508	17.72
- 1	-0.528	22.23	-0.488	21.60	-0.448	20.56	-0.452	18.18
0	-0.484	23.37	-0.443	22.59	-0.401	21.69	-0.398	18.79
1	-0.442	24.60	-0.400	23.29	-0.356	22.67	-0.346	19.91
2	-0.403	25.46	-0.357	23.68	-0.313	23.48	-0.298	21.57
3	-0.364	25.78	-0.315	23.80	-0.271	24.10	-0.253	23.26
4	-0.325	25.61	-0.273	23.73	-0.230	24.54	-0.212	24.88
5	-0.285	25.03	-0.230	23.57	-0.189	14.86	-0.173	26.30
6	-0.245	24.19	-0.188	23.34	-0.149	25.12	-0.136	27.48
7	-0.202	23.27	-0.145	23.15	-0.109	25.31	-0.100	28.44
8	-0.159	22.41	-0.101	22.99	-0.070	25.69	-0.065	29.11
9	-0.113	21.69	-0.058	22.95	-0.032	26.04	-0.031	29.41
10	-0.067	21.19	-0.014	23.03	0.007	25.93	0.003	29.21
11	-0.019	20.95	0.029	23.23	0.045	26.21	0.038	28.61
12	0.029	20.98	0.072	23.55	0.083	26.64	0.073	27.61
13	0.076	21.29	0.114	23.99	0.120	27.25	0.110	26.23
14	0.123	21.94	0.155	24.54	0.157	28.16	0.150	24.69
15	0.167	23.04	0.195	25.42	0.191	29.76	0.191	23.32

Differential capacity of mercury in KPF_6 solutions in dimethyl sulphoxide.

T = 25°C

Potential with respect to an aqueous 0. mol l^{-1} KCl calomel electrode dipping in to a solution of 0.1 mol l^{-1} KAF_6 in DMSO.

Reference: R. Payne. J. Amer. Chem. Soc., 89 (1967) 489.

c/mol^{-1}	0.010		0.030		0.10		0.20		1.0	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-15					-1.742	7.25	-1.702	7.46	-1.616	7.38
-14					-1.603	7.26	-1.566	7.39	-1.481	7.46
-13					-1.467	7.44	-1.432	9.59	-1.348	7.64
-12					-1.335	7.72	-1.303	7.87	-1.220	7.95
-11					-1.208	8.09	-1.178	8.25	-1.097	8.31
-10					-1.088	8.55	-1.061	8.75	-0.980	8.85
- 9					-0.975	9.19	-0.951	9.47	-0.871	9.62
- 8			-0.915	9.96	-0.871	10.11	-0.850	10.53	-0.773	10.71
- 7	-0.928	9.45	-0.821	11.41	-0.778	11.32	-0.761	11.87	-0.685	12.14
- 6	-0.827	10.49	-0.737	12.38	-0.694	12.58	-0.681	13.15	-0.608	13.84
- 5	-0.737	11.59	-0.660	13.56	-0.618	13.96	-0.609	14.66	-0.540	15.79
- 4	-0.654	12.74	-0.589	14.74	-0.550	15.30	-0.544	16.22	-0.480	17.90
- 3	-0.579	13.55	-0.523	15.53	-0.487	16.40	-0.485	17.59	-0.427	19.87
- 2	-0.505	13.35	-0.459	15.58	-0.427	17.09	-0.430	18.71	-0.379	21.87
- 1	-0.427	11.81	-0.394	14.67	-0.370	17.44	-0.378	19.68	-0.336	23.81
0	-0.332	9.70	-0.323	13.94	-0.313	18.00	-0.328	20.68	-0.295	25.57
1	-0.238	12.25	-0.254	15.52	-0.259	19.14	-0.281	21.86	-0.257	27.14
2	-0.166	15.50	-0.195	18.12	-0.209	20.85	-0.237	23.26	-0.221	28.57
3	-0.107	18.28	-0.143	20.58	-0.163	22.70	-0.195	24.83	-0.187	29.85
4	-0.056	20.48	-0.097	22.91	-0.121	24.46	-0.156	26.33	-0.154	30.93
5	-0.009	22.13	-0.055	24.99	-0.081	26.06	-0.119	27.73	-0.122	31.78
6	0.035	23.06	-0.016	26.74	-0.044	27.39	-0.084	28.96	-0.091	32.41
7	0.078	23.43	0.020	27.98	-0.008	28.46	-0.050	29.93	-0.060	32.74
8	0.121	23.51	0.055	28.55	0.027	29.07	-0.017	30.47	-0.030	32.64
9	0.164	23.13	0.090	28.55	0.061	29.17	0.016	30.57	0.001	31.98
10	0.208	21.64	0.125	28.12	0.096	28.73	0.049	30.13	0.033	30.76
11			0.161	27.19	0.131	27.72	0.083	29.35	0.066	29.08
12			0.199	25.26	0.168	26.18	0.117	28.10	0.102	27.19
13			0.241	23.36	0.208	24.34	0.154	26.20	0.140	25.04
14					0.251	22.53	0.194	24.22	0.182	22.90
15							0.237	22.37	0.228	21.05

Differential capacity of mercury in 0.1 mol l^{-1} solutions of salts in dimethyl sulphoxide. T = 25°C

Potential with respect to an aqueous 1 mol l^{-1} KCl calomel electrode

Reference: R. Payne, J. Amer. Chem. Soc., 89 (1967) 489.

Salt	KNO_3		$LiClO_4$		NH_4Cl		KBr	
$\sigma/\mu C\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$	E/volt	$C/\mu F\ cm^{-2}$
-14							-1.482	7.47
-13	-1.429	7.59	-1.448	7.93	-1.468	8.16	-1.350	7.41
-12	-1.299	7.87	-1.323	8.15	-1.345	8.23	-1.223	8.04
-11	-1.175	8.25	-1.202	8.43	-1.225	8.41	-1.101	8.48
-10	-1.057	8.74	-1.086	8.89	-1.108	8.73	-0.986	9.07
- 9	-0.947	9.46	-0.978	9.48	-0.997	9.23	-0.882	9.96
- 8	-0.846	10.47	-0.876	10.28	-0.892	9.96	-0.788	11.62
- 7	-0.756	11.76	-0.784	11.37	-0.796	10.98	-0.710	14.05
- 6	-0.676	13.18	-0.700	12.61	-0.710	12.14	-0.649	19.25
- 5	-0.603	14.61	-0.625	13.98	-0.632	13.66	-0.605	27.28
- 4	-0.538	15.93	-0.556	15.26	-0.565	16.16	-0.573	35.73
- 3	-0.477	17.01	-0.493	16.35	-0.511	21.15	-0.548	44.06
- 2	-0.420	17.65	-0.433	17.01	-0.471	28.98	-0.527	52.14
- 1	-0.364	18.20	-0.375	17.40	-0.441	38.11	-0.509	59.33
0	-0.311	19.70	-0.319	18.15	-0.417	47.31	-0.493	65.90
1	-0.264	22.72	-0.266	19.25	-0.398	56.25	-0.479	72.01
2	-0.223	26.88	-0.218	22.01	-0.381	64.86	-0.465	77.75
3	-0.188	30.70	-0.175	24.40	-0.367	73.15	-0.453	83.20
4	-0.158	34.27	-0.136	26.68	-0.354	81.15		
5	-0.130	37.50	-0.100	28.76				
6	-0.104	40.44	-0.066	30.66				
7	-0.080	43.17	-0.034	32.35				
8	-0.058	45.74	-0.004	33.81				
9	-0.036	48.07	0.025	34.93				
10	-0.016	50.19	0.053	35.77				
11	0.004	52.11	0.081	36.29				
12	0.022	53.68	0.108	36.55				
13	0.041	55.01	0.136	36.57				
14	0.059	56.16	0.163	36.40				
15	0.076	37.17	0.191	36.09				
16	0.094	58.03	0.219	35.74				
17	0.111	58.68	0.247	35.43				
18	0.128	59.11						
19	0.145	59.36						
20	0.162	59.73						
21	0.178	60.25						
22	0.195	60.81						

Differential Capacity of mercury in solutions of $LiClO_4$ in N-methyl-formamide T = 25°C

Potential with respect to aqueous saturated calomel electrode

Reference: W.R. Fawcett and R.O. Loutfy, J. Electroanal. Chem. 39 (1972) 185

	$C/\mu F\ cm^{-2}$		
$c/mol\ l^{-1}$ E/volt	0.050	0.10	0.20
-0.20	14.57	14.98	15.66
-0.25	13.56	14.34	14.72
-0.30	12.83	13.41	13.86
-0.35	12.40	12.97	13.26
-0.40	12.16	12.63	13.01
-0.45	12.07	12.46	12.92
-0.50	12.18	12.54	12.83
-0.55	12.47	12.72	13.09
-0.60	12.81	13.06	13.35
-0.65	13.34	13.40	13.95
-0.70	14.00	14.26	14.38
-0.75	14.64	15.00	15.15
-0.80	15.36	15.64	15.83
-0.85	16.08	16.41	16.56
-0.90	16.57	16.78	16.82
-0.95	16.83	16.90	16.78
-1.00	16.67	16.69	16.44
-1.05	16.15	16.07	15.66
-1.10	15.35	15.03	14.55
-1.15	14.33	14.09	13.51
-1.20	13.36	13.14	12.66
-1.25	12.50	12.28	11.72
-1.30	11.68	11.51	11.12
-1.35	11.08	10.74	10.69
-1.40	10.56	10.31	10.18
-1.45	10.26	10.05	10.01
-1.50	9.95	9.88	9.67
-1.55	9.77	9.71	9.49
-1.60	9.60	9.45	9.41
-1.65	9.51	9.36	9.15
-1.70	9.45	8.57	9.07
-1.75	9.40		9.15
E^z/volt	-0.294	-0.290	-0.284

Differential Capacity of mercury in solutions of $NaClO_4$ in N-methyl-formamide

T = 25°C

Potential with respect to aqueous saturated calomel electrode

Reference: W.R. Fawcett and R.O. Loutfy, J. Electroanal. Chem., 39 (1972) 185

		$C/\mu F\ cm^{-2}$	
c/mol l^{-1}	0.050	0.10	0.20
E/volt			
-0.20	14.53	15.42	15.97
-0.25	13.60	14.26	14.92
-0.30	12.84	13.58	14.15
-0.35	12.32	12.85	13.49
-0.40	12.12	12.54	12.88
-0.45	12.09	12.48	12.88
-0.50	12.19	12.60	12.99
-0.55	12.50	12.73	13.21
-0.60	12.86	13.09	13.49
-0.65	13.41	13.70	14.04
-0.70	14.00	14.25	14.70
-0.75	14.74	15.12	15.41
-0.80	15.55	15.96	16.13
-0.85	16.17	16.58	16.63
-0.90	16.67	16.89	17.07
-0.95	16.93	17.07	17.02
-1.00	16.71	16.76	16.36
-1.05	16.03	15.96	15.53
-1.10	15.02	14.93	14.53
-1.15	13.94	13.70	13.43
-1.20	12.91	12.78	12.38
-1.25	11.96	11.81	11.34
-1.30	11.26	11.08	10.73
-1.35	10.54	10.47	10.12
-1.40	10.07	9.99	9.57
-1.45	9.56	9.50	9.30
-1.50	9.28	9.26	9.08
-1.55	9.01	9.02	8.97
-1.60	8.92	8.83	8.64
-1.65	8.78	8.77	8.53
-1.70	8.72	8.71	8.53
-1.75	8.78	8.92	9.11
E^z/volt	-0.297	-0.294	-0.291

Differential Capacity of mercury in solutions of $KClO_4$ in
N-methyl formamide T = 25°C

Potential with respect to aqueous saturated calomel electrode

Reference: W.R. Fawcett and R.C. Loutfy, J. Electroanal. Chem. 39 (1972)185

	$C/\mu F\ cm^{-2}$			
$c/mol\ l^{-1}$ E/volt	0.020	0.050	0.10	0.20
-0.20	13.30	14.54	15.37	16.08
-0.25	12.16	13.55	14.18	14.74
-0.30	11.50	12.83	13.36	13.96
-0.35	11.18	12.37	12.75	13.33
-0.40	11.08	12.11	12.51	12.98
-0.45	11.18	12.05	12.44	12.88
-0.50	11.39	12.17	12.54	12.98
-0.55	11.71	12.41	12.75	13.17
-0.60	12.12	12.79	13.16	13.62
-0.65	12.65	13.31	13.65	14.15
-0.70	13.35	13.92	14.32	14.81
-0.75	14.07	14.70	15.11	15.55
-0.80	14.84	15.47	15.80	16.24
-0.85	15.62	16.15	16.42	16.90
-0.90	16.34	16.74	16.91	17.14
-0.95	16.80	16.95	17.02	17.10
-1.00	16.75	16.77	16.48	16.40
-1.05	16.38	16.09	15.74	15.37
-1.10	15.56	15.05	14.65	14.34
-1.15	14.47	13.90	13.53	13.31
-1.20	13.28	12.74	12.35	
-1.25	12.20	11.76	11.46	
-1.30	11.30	10.96	10.63	
-1.35	10.57	10.26	10.13	
-1.40	9.94	9.67	9.56	
-1.45	9.45	9.23	9.26	
-1.50	9.11	8.97	9.06	
-1.55	8.82	8.70		
-1.60	8.60	8.51		
-1.65	8.48	8.38		
-1.70	8.44	8.33		
-1.75	8.42	8.38		
E^z/volt	-0.297	-0.299	-0.294	-0.296

Differential Capacity of mercury in solutions of C_sClO_4 in N-methyl-formamide

T = 25°C

Potential with respect to aqueous saturated calomel electrode

Reference: W.R. Fawcett and R.O. Loutfy, J. Electroanal. Chem. 39 (1972) 185

	$C/\mu F\ cm^{-2}$			$C/\mu F\ cm^{-2}$		
$c/mol\ l^{-1}$ E/volt	0.010	0.020	$c/mol\ l^{-1}$ E/volt	0.050	0.10	0.20
-0.260	11.93	12.53	-0.20	14.53	15.66	15.99
-0.310	10.82	11.53	-0.25	13.41	14.49	14.90
-0.360	10.56	11.17	-0.30	12.72	13.56	14.08
-0.410	10.60	11.18	-0.35	12.30	12.98	13.39
-0.460	10.95	11.39	-0.40	12.09	12.63	13.04
-0.510	11.45	11.69	-0.45	12.09	12.57	12.92
-0.560	11.90	12.06	-0.50	12.23	12.60	12.98
-0.610	12.47	12.55	-0.55	12.44	12.85	13.04
-0.660	13.00	13.13	-0.60	12.82	13.20	13.44
-0.710	13.62	13.80	-0.65	13.38	13.72	13.91
-0.760	14.36	14.53	-0.70	13.97	14.37	14.49
-0.810	15.12	15.42	-0.75	14.68	15.06	15.30
-0.860	15.87	16.07	-0.80	15.34	15.69	15.88
-0.910	16.37	16.46	-0.85	16.00	16.37	16.46
-0.960	16.75	16.83	-0.90	16.64	16.76	16.92
-1.010	16.60	16.66	-0.95	16.81	16.83	16.98
-1.060	15.75	15.80	-1.00	16.64	16.47	16.40
-1.110	14.71	14.88	-1.05	16.00	15.74	15.70
-1.160	13.68	13.86	-1.10	15.11	14.70	14.72
-1.210	12.67	12.82	-1.15	14.02	13.67	13.44
-1.260	11.76	11.90	-1.20	12.94	12.60	12.40
-1.310	11.05	11.22	-1.25	11.95	11.65	11.54
-1.360	10.49	10.60	-1.30	11.16	10.95	10.90
-1.410	10.20	10.18	-1.35	10.55	10.39	10.38
-1.460	9.98	9.98	-1.40	10.12	9.95	9.98
-1.510		9.70	-1.45	9.80	9.77	9.92
-1.560		9.72	-1.50	9.77	9.77	9.98
-1.610		9.89	-1.55	10.03	10.06	10.21
-1.660		10.29	-1.60	10.52	10.68	10.96
-1.710		10.82	-1.65	11.29	11.74	12.12
-1.760		12.39	-1.70	12.76	13.76	14.24
E^z/volt	-0.295	-0.294		-0.297	-0.296	-0.295

Differential capacity of mercury in 0.1 mol l^{-1} NaF in water containing formamide T = 25°

Potentials with respect to 0.1 mol l^{-1} calomel electrode

Reference: R. Payne, J. Electroanal. Chem. 47 (1973) 265.

c/mol l^{-1}	0.2584		0.5283		1.018		2.005	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-21	-1.734	18.54						
-20	-1.679	18.05						
-19	-1.623	17.60	-1.630	17.60	-1.637	17.58	-1.649	17.49
-18	-1.565	17.16	-1.573	17.18	-1.579	17.17	-1.591	17.09
-17	-1.506	16.76	-1.514	16.80	-1.520	16.81	-1.532	16.75
-16	-1.446	16.44	-1.453	16.48	-1.460	15.51	-1.472	16.49
-15	-1.385	16.17	-1.393	16.23	-1.399	16.29	-1.411	16.31
-14	-1.323	15.96	-1.331	16.04	-1.338	16.14	-1.349	16.22
-13	-1.260	15.87	-1.268	15.96	-1.275	16.08	-1.288	16.20
-12	-1.197	15.87	-1.205	15.96	-1.213	16.12	-1.226	16.27
-11	-1.134	16.00	-1.143	16.09	-1.151	16.22	-1.165	16.38
-10	-1.072	16.26	-1.081	16.31	-1.090	16.47	-1.104	16.62
-9	-1.011	16.64	-1.021	16.68	-1.030	16.79	-1.044	16.85
-8	-0.952	17.15	-0.961	17.11	-0.971	17.17	-0.986	17.17
-7	-0.894	17.72	-0.904	17.62	-0.913	17.60	-0.928	17.48
-6	-0.839	18.31	-0.848	18.14	-0.857	18.02	-0.871	17.75
-5	-0.785	18.89	-0.793	18.60	-0.802	18.29	-0.815	17.94
-4	-0.733	19.34	-0.740	18.96	-0.748	18.53	-0.759	18.03
-3	-0.681	19.61	-0.688	19.15	-0.694	18.61	-0.704	17.99
-2	-0.631	19.73	-0.636	19.22	-0.640	18.60	-0.648	17.87
-1	-0.580	19.83	-0.684	19.26	-0.587	18.59	-0.592	17.80
0	-0.530	20.17	-0.532	19.56	-0.533	18.81	-0.536	17.96
1	-0.481	20.85	-0.482	20.20	-0.481	19.41	-0.481	18.49
2	-0.434	21.81	-0.433	21.15	-0.430	20.30	-0.428	19.34
3	-0.390	22.87	-0.387	22.20	-0.382	21.33	-0.378	20.35
4	-0.347	23.83	-0.343	23.19	-0.336	22.45	-0.330	21.42
5	-0.305	24.68	-0.301	24.12	-0.293	23.44	-0.284	22.52
6	-0.266	25.42	-0.260	24.96	-0.251	24.42	-0.241	23.70
7	-0.227	26.07	-0.221	25.74	-0.211	25.41	-0.200	24.91
8	-0.189	26.71	-0.182	26.51	-0.172	26.39	-0.161	26.12
9	-0.152	27.29	-0.145	27.24	-0.135	27.37	-0.123	27.36
10	-0.116	27.89	-0.109	27.97	-0.099	28.23	-0.087	28.59
11	-0.080	28.47	-0.074	28.68	-0.064	29.27	-0.053	29.87
12	-0.045	29.11	-0.039	29.44	-0.031	30.27	-0.020	31.21
13	-0.011	29.84	-0.006	30.29				

Capacity of mercury in 0.1 mol l^{-1} NaF in water containing formamide (cont.)

c/mol l^{-1}	3.976		6.011		8.087	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-19						
-18	-1.602	16.70	-1.609	16.19		
-17	-1.542	16.43	-1.547	15.99	-1.553	15.61
-16	-1.481	16.25	-1.485	15.92	-1.489	15.62
-15	-1.419	16.17	-1.422	15.95	-1.425	15.75
-14	-1.357	16.17	-1.359	16.06	-1.362	15.93
-13	-1.295	16.26	-1.297	16.24	-1.300	16.19
-12	-1.234	16.40	-1.236	16.47	-1.238	16.49
-11	-1.173	16.59	-1.176	16.72	-1.178	16.83
-10	-1.114	16.83	-1.117	17.00	-1.119	17.12
-9	-1.054	17.06	-1.058	17.22	-1.061	17.35
-8	-0.996	17.28	-1.000	17.40	-1.004	17.50
-7	-0.939	17.48	-0.943	17.51	-0.947	17.57
-6	-0.882	17.60	-0.886	17.55	-0.890	17.53
-5	-0.825	17.64	-0.829	17.47	-0.833	17.37
-4	-0.768	17.56	-0.771	17.27	-0.775	17.10
-3	-0.711	17.40	-0.713	17.00	-0.716	16.75
-2	-0.653	17.17	-0.654	16.71	-0.655	16.41
-1	-0.595	17.03	-0.594	16.52	-0.594	16.16
0	-0.536	17.11	-0.533	15.59	-0.532	16.16
1	-0.478	17.54	-0.473	16.97	-0.471	16.48
2	-0.422	18.27	-0.415	17.64	-0.411	17.14
3	-0.369	19.21	-0.360	18.56	-0.354	18.02
4	-0.318	20.29	-0.308	19.66	-0.300	19.13
5	-0.270	21.52	-0.258	20.93	-0.250	20.46
6	-0.225	22.84	-0.212	22.38	-0.202	21.96
7	-0.183	24.31	-0.169	23.98	-0.158	23.65
8	-0.143	25.81	-0.129	25.68	-0.118	25.47
9	-0.105	27.37	-0.091	27.47	-0.080	27.43
10	-0.070	28.93	-0.056	29.33	-0.045	29.65
11	-0.036	30.60				

Differential capacity of mercury in 0.1 mol l^{-1} NaF in water containing N-methyl formamide T = 25°C

Potential with respect to 0.1 mol l^{-1} calomel electrode

Reference: R. Payne, J. Electroanal. Chem. 47 (1973) 265.

c/mol l^{-1}	0		0.1004		0.1823		0.4394	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20			-1.685	18.19				
-19	-1.663	17.57	-1.629	17.72	-1.631	17.80	-1.643	17.98
-18	-1.565	17.14	-1.572	17.29	-1.575	17.37	-1.587	17.49
-17	-1.506	16.74	-1.513	16.89	-1.516	16.97	-1.529	17.06
-16	-1.446	16.39	-1.453	16.56	-1.457	16.63	-1.470	16.70
-15	-1.384	16.10	-1.393	16.28	-1.396	16.37	-1.409	16.44
-14	-1.321	15.86	-1.331	16.07	-1.335	16.18	-1.348	16.25
-13	-1.258	15.72	-1.268	15.97	-1.272	16.07	-1.286	16.12
-12	-1.194	15.71	-1.206	15.93	-1.210	16.05	-1.224	16.12
-11	-1.131	15.82	-1.143	16.00	-1.148	16.11	-1.162	16.16
-10	-1.068	16.08	-1.081	16.19	-1.086	16.27	-1.100	16.26
-9	-1.007	16.48	-1.019	16.49	-1.025	16.51	-1.039	16.41
-8	-0.947	17.02	-0.959	16.88	-0.965	16.85	-0.979	16.59
-7	-0.889	17.65	-0.901	17.36	-0.907	17.23	-0.919	16.82
-6	-0.834	18.37	-0.844	17.86	-0.849	17.62	-0.860	17.05
-5	-0.780	19.00	-0.789	18.35	-0.793	17.99	-0.801	17.27
-4	-0.728	19.60	-0.735	18.73	-0.738	18.30	-0.744	17.44
-3	-0.678	19.97	-0.682	18.99	-0.683	18.49	-0.687	17.54
-2	-0.628	20.18	-0.630	19.13	-0.630	18.60	-0.630	17.60
-1	-0.579	20.34	-0.578	19.26	-0.576	18.73	-0.573	17.71
0	-0.530	20.74	-0.526	19.62	-0.523	19.09	-0.517	18.09
1	-0.482	21.39	-0.476	20.39	-0.472	19.86	-0.463	18.87
2	-0.437	22.53	-0.428	21.46	-0.422	20.96	-0.411	20.01
3	-0.394	23.46	-0.383	22.62	-0.376	22.16	-0.363	21.31
4	-0.352	24.46	-0.340	23.74	-0.332	23.39	-0.317	22.65
5	-0.312	25.22	-0.298	24.75	-0.290	24.52	-0.274	24.03
6	-0.272	25.85	-0.259	25.64	-0.251	25.58	-0.234	25.33
7	-0.234	26.33	-0.220	26.44	-0.212	26.56	-0.195	26.61
8	-0.196	26.83	-0.183	27.16	-0.175	27.50	-0.159	27.84
9	-0.159	27.18	-0.147	27.86	-0.139	28.40	-0.124	28.97
10	-0.123	27.67	-0.111	28.53	-0.105	29.22	-0.090	30.01
11	-0.089	28.21	-0.076	29.13	-0.071	29.92	-0.057	30.95
12			-0.042	29.76	-0.038	30.57	-0.025	31.85
13			-0.009	30.47	-0.005	31.35	+0.006	32.73

Capacity of mercury in 0.1 mol l^{-1} NaF in water containing N-methyl formamide (cont.)

c/mol l^{-1}	0.8631		1.701		3.386		6.586	
σ /μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-20								
-19								
-18	-1.603	17.30						
-17	-1.544	16.77	-1.562	15.58				
-16	-1.484	16.38	-1.497	15.33				
-15	-1.422	16.13	-1.432	15.30	-1.427	14.38	-1.406	13.77
-14	-1.360	16.00	-1.367	15.42	-1.358	14.89	-1.335	14.58
-13	-1.297	15.96	-1.302	15.60	-1.293	15.41	-1.269	15.39
-12	-1.235	15.97	-1.238	15.80	-1.229	15.84	-1.205	16.09
-11	-1.172	16.04	-1.175	15.99	-1.166	16.19	-1.144	16.61
-10	-1.110	16.13	-1.113	16.14	-1.105	16.40	-1.084	16.92
-9	-1.048	16.24	-1.051	16.23	-1.044	16.48	-1.026	17.01
-8	-0.987	16.36	-0.990	16.28	-0.983	16.47	-0.967	16.89
-7	-0.926	16.48	-0.928	16.29	-0.922	16.35	-0.907	16.62
-6	-0.865	16.59	-0.867	16.27	-0.861	16.17	-0.846	16.25
-5	-0.805	16.67	-0.806	16.23	-0.799	15.94	-0.784	15.82
-4	-0.745	16.74	-0.744	16.16	-0.735	15.74	-0.720	15.41
-3	-0.686	16.77	-0.682	16.10	-0.671	15.53	-0.654	15.10
-2	-0.626	16.78	-0.620	16.05	-0.607	15.42	-0.587	14.91
-1	-0.567	16.89	-0.557	16.14	-0.542	15.48	-0.520	14.94
0	-0.508	17.26	-0.496	16.51	-0.478	15.84	-0.454	15.31
1	-0.451	18.04	-0.437	17.27	-0.416	16.61	-0.390	16.09
2	-0.397	19.19	-0.380	18.41	-0.358	17.79	-0.330	17.31
3	-0.347	20.56	-0.328	19.86	-0.304	19.28	-0.275	18.91
4	-0.300	22.04	-0.280	21.46	-0.254	21.02	-0.224	20.86
5	-0.256	23.55	-0.235	23.16	-0.209	22.97	-0.179	23.10
6	-0.215	25.08	-0.193	24.96	-0.167	25.06	-0.138	25.57
7	-0.176	26.63	-0.155	26.80	-0.129	27.23	-0.101	28.28
8	-0.140	28.17	-0.119	28.66	-0.093	29.45	-0.067	31.05
9	-0.105	29.61	-0.085	30.42	-0.061	31.64	-0.036	33.81
10	-0.072	30.96	-0.053	32.14	-0.030	33.77	-0.007	36.58
11	-0.041	32.21	-0.022	33.75	-0.013	35.85		
12	-0.010	33.35	+0.007	35.25				

Differential capacity of mercury in KCl in formamide at 25°C.

Potential with respect to aqueous 0.1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} solution of KCl in formamide.

Reference: R. Payne, J. Electroanal. Chem. 41 (1973) 145.

c/mol l^{-1}	0.00813		0.01648		0.03913		0.08984	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
18					-0.099	157.8		
17			-0.082	130.9	-0.106	130.9	-0.132	128.6
16			-0.090	109.9	-0.115	105.4	-0.141	106.6
15	-0.078	95.12	-0.100	90.95	-0.125	94.72	-0.151	87.58
14	-0.090	80.92	-0.112	81.47	-0.137	81.25	-0.163	82.07
13	-0.103	69.13	-0.125	69.78	-0.150	69.92	-0.176	71.64
12	-0.118	62.71	-0.140	63.88	-0.165	64.07	-0.191	64.86
11	-0.135	56.68	-0.157	57.41	-0.181	57.39	-0.206	58.25
10	-0.154	51.17	-0.175	51.76	-0.199	52.40	-0.225	53.21
9	-0.174	46.13	-0.195	46.94	-0.220	47.49	-0.244	48.23
8	-0.197	41.58	-0.218	42.28	-0.242	42.89	-0.266	43.73
7	-0.223	37.37	-0.243	38.07	-0.266	38.73	-0.290	39.62
6	-0.251	33.45	-0.271	33.96	-0.293	34.80	-0.317	35.88
5	-0.283	29.48	-0.302	30.36	-0.324	31.22	-0.346	32.44
4	-0.319	25.86	-0.337	26.88	-0.358	27.94	-0.379	29.14
3	-0.361	22.22	-0.377	23.49	-0.396	24.78	-0.415	26.20
2	-0.409	18.85	-0.422	20.17	-0.439	21.71	-0.455	23.40
1	-0.468	15.60	-0.476	16.99	-0.488	18.90	-0.500	20.92
0	-0.538	13.14	-0.541	14.27	-0.545	16.44	-0.551	18.75
-1	-0.617	12.51	-0.615	13.15	-0.609	14.95	-0.607	17.13
-2	-0.695	13.28	-0.690	13.58	-0.677	14.70	-0.667	16.30
-3	-0.767	14.42	-0.761	14.58	-0.744	15.23	-0.729	16.32
-4	-0.834	15.58	-0.827	15.69	-0.808	16.11	-0.789	16.79
-5	-0.896	16.61	-0.889	16.71	-0.868	17.04	-0.847	17.51
-6	-0.955	17.49	-0.947	17.57	-0.926	17.86	-0.903	18.19
-7	-1.011	18.03	-1.003	18.20	-0.981	18.44	-0.958	18.69
-8	-1.066	18.42	-1.058	18.49	-1.035	18.71	-1.011	18.93
-9	-1.120	18.36	-1.112	18.49	-1.088	18.67	-1.064	18.86
-10	-1.175	18.05	-1.166	18.15	-1.142	18.30	-1.117	18.47
-11	-1.231	17.44	-1.222	17.57	-1.197	17.70	-1.172	17.85
-12	-1.290	16.80	-1.280	16.81	-1.255	16.92	-1.229	17.05
-13	-1.350	16.08	-1.341	16.02	-1.316	16.11	-1.290	16.21
-14			-1.405	15.33	-1.379	15.32	-1.353	15.39
-15					-1.446	14.27	-1.419	14.69

Capacity in KCl in formamide (cont.)

c/mol l^{-1}	0.1749		0.3596		0.7325	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
16	-0.161	105.2				
15	-0.172	90.01	-0.199	90.35		
14	-0.184	77.61	-0.211	78.22		
13	-0.197	71.70	-0.225	72.31	-0.251	72.13
12	-0.212	64.03	-0.239	64.86	-0.265	65.16
11	-0.228	58.75	-0.255	59.52	-0.281	59.77
10	-0.246	53.44	-0.273	54.17	-0.298	54.56
9	-0.265	48.56	-0.292	49.63	-0.318	50.05
8	-0.287	44.31	-0.313	45.22	-0.339	46.41
7	-0.311	40.32	-0.336	41.28	-0.361	42.59
6	-0.337	36.54	-0.362	37.65	-0.386	38.58
5	-0.365	33.14	-0.390	34.36	-0.413	35.51
4	-0.397	30.03	-0.420	31.42	-0.442	32.67
3	-0.432	27.23	-0.453	28.74	-0.474	30.25
2	-0.471	24.69	-0.490	26.40	-0.508	28.04
1	-0.513	22.40	-0.529	24.30	-0.546	26.17
0	-0.560	20.45	-0.572	22.55	-0.585	24.53
-1	-0.611	18.94	-0.618	21.10	-0.627	23.31
-2	-0.665	17.98	-0.667	20.03	-0.671	22.31
-3	-0.722	17.63	-0.717	19.40	-0.716	21.58
-4	-0.778	17.80	-0.769	19.17	-0.763	21.13
-5	-0.834	18.22	-0.821	19.25	-0.811	20.88
-6	-0.888	18.69	-0.873	19.46	-0.859	20.73
-7	-0.941	19.04	-0.924	19.61	-0.907	20.60
-8	-0.993	19.17	-0.975	19.60	-0.956	20.33
-9	-1.046	19.02	-1.027	19.33	-1.006	19.86
-10	-1.099	18.58	-1.079	18.80	-1.057	19.17
-11	-1.153	17.89	-1.133	18.06	-1.110	18.29
-12	-1.211	17.07	-1.190	17.19	-1.167	17.33
-13	-1.271	16.21	-1.250	16.28	-1.226	16.35
-14	-1.334	15.38	-1.313	15.42	-1.289	15.43
-15	-1.401	14.67	-1.379	14.68	-1.356	14.66
-16	-1.470	14.12	-1.449	14.08	-1.425	14.07

Differential capacity of mercury in KI in formamide at 25°C.

Potential with respect to aqueous 0.1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} KI solution in formamide.

Reference: R. Payne, J. Chem. Physics, 42 (1965) 3371.

c/mol l^{-1}	0.00932		0.02565		0.04463		0.08921	
σ /μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
25			-0.423	250.3	-0.460	239.2	-0.484	229.6
24			-0.427	238.2	-0.464	225.7	-0.488	216.2
23	-0.409		-0.432	225.3	-0.469	212.5	-0.493	203.2
22	-0.414		-0.436	203.5	-0.474	199.5	-0.498	190.7
21	-0.419		-0.442	189.1	-0.479	187.0	-0.503	179.2
20	-0.425		-0.447	175.3	-0.484	175.1	-0.509	172.2
19	-0.431		-0.453	162.4	-0.490	166.3	-0.515	162.0
18	-0.437		-0.459	150.8	-0.497	155.6	-0.521	152.2
17	-0.444		-0.466	146.0	-0.503	145.4	-0.528	142.8
16	-0.451		-0.473	137.4	-0.510	136.1	-0.535	135.7
15	-0.458		-0.481	128.9	-0.518	130.0	-0.543	127.6
14	-0.467		-0.489	120.7	-0.526	122.1	-0.551	119.9
13	-0.475		-0.497	113.4	-0.534	114.4	-0.560	112.4
12	-0.485		-0.506	106.0	-0.543	107.1	-0.569	106.1
11	-0.495		-0.516	98.79	-0.553	100.7	-0.579	99.46
10	-0.505		-0.527	92.49	-0.563	93.97	-0.589	93.07
9	-0.517		-0.538	86.08	-0.574	87.52	-0.600	87.31
8	-0.529		-0.550	79.89	-0.586	81.61	-0.612	81.49
7	-0.543		-0.563	74.23	-0.599	75.66	-0.625	75.93
6	-0.556	66.55	-0.577	68.47	-0.613	70.11	-0.638	70.77
5	-0.572	60.84	-0.592	63.01	-0.627	64.70	-0.653	65.67
4	-0.589	55.36	-0.608	57.44	-0.643	59.79	-0.669	60.91
3	-0.608	50.06	-0.627	51.90	-0.661	54.84	-0.686	56.47
2	-0.629	45.16	-0.647	46.54	-0.680	50.07	-0.704	52.17
1	-0.652	39.76	-0.670	41.32	-0.701	45.53	-0.724	48.03
0	-0.680	34.28	-0.696	36.30	-0.724	41.12	-0.746	44.08
-1	-0.712	28.94	-0.726	31.39	-0.750	36.65	-0.770	40.37
-2	-0.750	23.65	-0.760	26.18	-0.779	32.12	-0.796	36.63
-3	-0.797	19.48	-0.803	21.70	-0.812	28.07	-0.824	32.99
-4	-0.852	17.69	-0.852	18.90	-0.851	24.16	-0.857	29.35
-5	-0.909	17.55	-0.907	18.04	-0.895	21.35	-0.893	25.95
-6	-0.965	18.01	-0.962	18.10	-0.944	19.8	-0.934	23.17
-7	-1.020	18.52	-1.017	18.49	-0.995	19.30	-0.979	21.28
-8	-1.074	18.68	-1.071	18.71	-1.047	19.09	-1.028	20.14
-9	-1.128	18.62	-1.124	18.63	-1.100	18.84	-1.078	19.38
-10	-1.182	18.31	-1.178	18.29	-1.153	18.39	-1.131	18.68
-11	-1.237	17.74	-1.234	17.70	-1.209	17.73	-1.185	17.93

Capacity in KI in formamide (cont.)

$c/\text{mol l}^{-1}$	0.177		0.359		0.7411	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
20	-0.531	156.7	-0.555	159.2		
19	-0.538	150.9	-0.562	149.9		
18	-0.544	144.8	-0.569	141.0		
17	-0.551	138.5	-0.576	132.6		
16	-0.559	131.8	-0.584	124.9	-0.611	121.1
15	-0.567	123.4	-0.592	119.4	-0.620	114.4
14	-0.575	116.3	-0.600	112.7	-0.629	108.1
13	-0.584	109.4	-0.610	106.3	-0.638	102.9
12	-0.593	102.9	-0.619	100.8	-0.648	97.47
11	-0.603	97.36	-0.629	95.12	-0.659	92.35
10	-0.614	91.52	-0.640	89.70	-0.670	87.90
9	-0.625	86.25	-0.652	84.87	-0.682	83.26
8	-0.637	80.94	-0.664	79.88	-0.694	78.72
7	-0.650	75.76	-0.677	75.02	-0.707	74.34
6	-0.664	70.80	-0.691	70.36	-0.721	70.05
5	-0.678	65.99	-0.705	66.02	-0.736	65.94
4	-0.694	61.49	-0.721	61.93	-0.751	62.18
3	-0.711	57.24	-0.738	58.28	-0.768	58.89
2	-0.729	53.44	-0.755	54.81	-0.785	55.74
1	-0.748	49.83	-0.774	51.64	-0.804	52.85
0	-0.769	46.42	-0.794	48.51	-0.823	50.28
-1	-0.791	43.12	-0.815	45.68	-0.843	47.53
-2	-0.815	40.03	-0.838	43.01	-0.865	47.59
-3	-0.841	36.96	-0.862	40.46	-0.885	-
-4	-0.870	33.83	-0.887	37.82	-0.906	42.13
-5	-0.901	30.73	-0.915	35.25	-0.931	39.53
-6	-0.935	27.51	-0.944	32.37	-0.951	37.18
-7	-0.974	24.70	-0.977	29.31	-0.985	34.55
-8	-1.016	22.35	-1.013	26.38	-1.015	31.69
-9	-1.063	20.61	-1.053	23.53	-1.049	28.55
-10	-1.113	19.33	-1.098	21.16	-1.086	25.31
-11	-1.166	18.25	-1.147	19.26	-1.128	22.30
-12	-1.223	17.25	-1.202	17.76	-1.176	19.66
-13	-1.282	16.29	-1.260	16.52	-1.230	17.64
-14	-1.345	15.42	-1.322	15.52	-1.289	16.15
-15	-1.412	14.68	-1.389	14.72	-1.353	15.07
-16			-1.458	14.10	-1.421	14.30

Differential capacity of mercury in KNO_3 in formamide at 25°C.

Potential with respect to aqueous 0.1 mol l^{-1} KCl calomel electrode in contact with 0.1 mol l^{-1} KNO_3 solution in formamide.

Reference: R. Payne, J. Electroanal. Chem. 41 (1973) 145.

c/mol l^{-1}	0.009179		0.01633		0.02495		0.09028	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
11							-0.031	32.42
10					-0.048	29.41	-0.063	29.61
9			-0.063	27.32	-0.083	27.10	-0.098	27.23
8	-0.087	25.47	-0.101	25.22	-0.122	25.04	-0.136	25.18
7	-0.128	23.40	-0.143	23.29	-0.163	23.24	-0.178	23.40
6	-0.173	21.49	-0.187	21.62	-0.208	21.67	-0.222	21.86
5	-0.221	19.83	-0.235	20.10	-0.256	20.27	-0.269	20.55
4	-0.274	18.40	-0.287	18.64	-0.307	18.98	-0.319	19.38
3	-0.330	16.82	-0.343	17.18	-0.361	17.77	-0.372	18.33
2	-0.393	15.12	-0.404	15.65	-0.419	16.64	-0.428	17.40
1	-0.464	13.16	-0.471	13.99	-0.481	15.50	-0.487	16.59
0	-0.546	11.42	-0.547	12.61	-0.548	14.59	-0.549	15.84
-1	-0.634	11.60	-0.628	12.45	-0.618	14.21	-0.613	15.48
-2	-0.716	12.84	-0.706	13.28	-0.688	14.48	-0.677	15.57
-3	-0.790	14.17	-0.778	14.43	-0.755	15.20	-0.741	16.03
-4	-0.858	15.47	-0.845	15.64	-0.819	16.11	-0.802	16.71
-5	-0.899	17.09	-0.906	16.69	-0.880	17.04	-0.861	17.47
-6	-0.957	17.48	-0.965	17.59	-0.937	17.84	-0.917	18.15
-7	-1.014	17.95	-1.021	18.18	-0.992	18.43	-0.971	18.66
-8	-1.069	18.41	-1.075	18.51	-1.046	18.70	-1.024	18.89
-9	-1.123	18.52	-1.129	18.55	-1.099	18.66	-1.077	18.83
-10	-1.177	18.24	-1.183	18.16	-1.153	18.29	-1.131	18.45
-11	-1.233	17.73	-1.239	17.52	-1.209	17.69	-1.186	17.84
-12	-1.290	17.09	-1.297	16.81	-1.267	16.91	-1.243	17.04
-13	-1.350	16.26	-1.358	16.03	-1.327	16.09	-1.303	16.22
-14	-1.413	15.47	-1.422	15.25	-1.391	15.31	-1.366	15.42
-15			-1.489	14.62			-1.433	14.71
-16							-1.502	14.17
-17								

Capacity in KNO_3 in formamide (cont.)

$c/\text{mol l}^{-1}$	0.1835		0.4437		0.916	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
11			-0.053	30.26	-0.071	30.02
10	-0.072	28.49	-0.087	28.19	-0.105	28.02
9	-0.108	26.53	-0.124	26.31	-0.142	26.21
8	-0.147	24.73	-0.163	24.65	-0.182	24.61
7	-0.189	23.18	-0.205	23.21	-0.223	23.21
6	-0.233	21.82	-0.250	21.94	-0.268	22.03
5	-0.281	20.64	-0.296	20.88	-0.314	21.03
4	-0.330	19.63	-0.345	19.98	-0.363	20.21
3	-0.382	18.73	-0.396	19.25	-0.413	19.59
2	-0.437	17.98	-0.449	18.72	-0.465	19.20
1	-0.494	17.37	-0.503	18.34	-0.517	18.97
0	-0.552	16.89	-0.558	18.08	-0.570	18.87
-1	-0.612	16.63	-0.614	17.96	-0.623	18.97
-2	-0.672	16.61	-0.669	18.04	-0.675	19.18
-3	-0.732	16.88	-0.724	18.25	-0.727	19.47
-4	-0.790	17.39	-0.779	18.62	-0.778	19.82
-5	-0.847	18.00	-0.832	19.04	-0.828	20.14
-6	-0.901	18.57	-0.884	19.40	-0.877	20.36
-7	-0.955	19.00	-0.935	19.62	-0.926	20.41
-8	-1.007	19.15	-0.986	19.64	-0.976	20.25
-9	-1.059	19.01	-1.037	19.37	-1.026	19.83
-10	-1.113	18.59	-1.089	18.84	-1.077	19.16
-11	-1.167	17.91	-1.143	18.09	-1.130	18.30
-12	-1.224	17.09	-1.200	17.20	-1.186	17.32
-13	-1.285	16.21	-1.260	16.29	-1.246	16.35
-14	-1.348	15.39	-1.323	15.42	-1.309	15.44
-15	-1.414	14.67	-1.390	14.67	-1.375	14.67
-16	-1.484	14.08	-1.459	14.07	-1.445	14.06
-17	-1.556	13.69	-1.531	13.66		

Differential capacity of a mercury electrode in 0.1 M aqueous NaF containing various concentrations (c) of N,N'-dimethyl formamide T = 25°C

Potential with respect to aqueous 0.1 mol l^{-1} KCl calomel electrode

Reference: R. Payne, J. Electroanal. Chem. 47 (1973) 267.

$c/\text{mol l}^{-1}$	0		0.01118		0.02232		0.05449	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
14	0.018	29.85						
13	-0.016	29.14	-0.015	29.56	-0.013	29.91	-0.009	30.62
12	-0.051	28.55	-0.049	28.93	-0.047	29.29	-0.042	30.13
11	-0.086	28.04	-0.084	28.43	-0.082	28.80	-0.076	29.77
10	-0.122	27.62	-0.119	27.99	-0.117	28.37	-0.110	29.42
9	-0.158	27.20	-0.155	27.60	-0.152	27.96	-0.144	28.98
8	-0.195	26.80	-0.192	27.17	-0.188	27.51	-0.179	28.42
7	-0.233	26.37	-0.229	26.68	-0.225	26.99	-0.214	27.70
6	-0.271	25.90	-0.267	26.10	-0.262	26.33	-0.251	26.79
5	-0.310	25.30	-0.306	25.36	-0.301	25.48	-0.289	25.69
4	-0.350	24.55	-0.346	24.46	-0.341	24.46	-0.329	24.44
3	-0.392	23.60	-0.388	23.39	-0.383	23.29	-0.371	23.00
2	-0.435	22.56	-0.431	22.23	-0.427	22.01	-0.416	21.51
1	-0.481	21.54	-0.478	21.09	-0.474	20.78	-0.464	20.17
0	-0.528	20.80	-0.526	20.26	-0.523	19.89	-0.515	19.18
-1	-0.577	20.41	-0.576	19.83	-0.574	19.41	-0.568	18.58
-2	-0.626	20.25	-0.627	19.61	-0.626	19.13	-0.622	18.22
-3	-0.675	20.04	-0.678	19.37	-0.678	18.88	-0.678	17.87
-4	-0.726	19.66	-0.730	19.00	-0.732	18.51	-0.734	17.47
-5	-0.777	19.10	-0.783	18.50	-0.787	18.03	-0.792	17.03
-6	-0.831	18.42	-0.838	17.93	-0.843	17.53	-0.852	16.63
-7	-0.886	17.72	-0.895	17.36	-0.901	17.05	-0.913	16.33
-8	-0.944	17.07	-0.953	16.84	-0.960	16.64	-0.974	16.15
-9	-1.003	16.52	-1.014	16.43	-1.021	16.35	-1.036	16.09
-10	-1.065	16.11	-1.075	16.13	-1.082	16.15	-1.098	16.12
-11	-1.127	15.84	-1.138	15.95	-1.145	16.04	-1.160	16.21
-12	-1.191	15.71	-1.200	15.89	-1.207	16.03	-1.222	16.34
-13	-1.	15.71	-1.263	15.93	-1.269	16.11	-1.283	16.52
-14	-1.318	15.82	-1.326	16.06	-1.331	16.25	-1.343	16.70
-15	-1.380	16.06	-1.388	16.28	-1.392	16.45	-1.402	16.91
-16	-1.442	16.34	-1.449	16.55	-1.453	16.71	-1.461	17.14
-17	-1.503	16.70	-1.508	16.89	-1.512	17.04	-1.519	17.41
-18	-1.562	17.10	-1.567	17.28	-1.570	17.39	-1.576	17.72
-19	-1.620	17.54	-1.624	17.70	-1.626	17.79	-1.632	18.07
-20	-1.676	18.02	-1.680	18.15	-1.682	18.22		

Capacity of Hg in 0.1 M NaF (aq), containing DMF (cont.)

$c/\text{mol l}^{-1}$	0.109		0.1986		0.3941		0.8175	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
14			0.028	33.93	0.035	36.52	0.047	40.23
13	-0.006	31.82	-0.002	33.60	0.008	36.26	0.022	39.90
12	-0.038	31.50	-0.032	33.41	-0.020	36.07	-0.003	39.53
11	-0.096	31.26	-0.062	33.19	-0.048	35.69	-0.028	38.95
10	-0.107	30.96	-0.092	32.74	-0.076	35.06	-0.054	37.92
9	-0.134	30.27	-0.123	32.01	-0.105	33.97	-0.081	36.41
8	-0.168	29.60	-0.155	30.92	-0.135	32.50	-0.109	34.34
7	-0.202	28.61	-0.188	29.55	-0.167	30.63	-0.140	32.01
6	-0.238	27.40	-0.222	27.97	-0.200	28.62	-0.172	29.45
5	-0.275	25.99	-0.259	26.28	-0.237	26.52	-0.208	26.91
4	-0.315	24.47	-0.299	24.48	-0.276	24.41	-0.247	24.48
3	-0.357	22.83	-0.341	22.60	-0.319	22.33	-0.289	22.19
2	-0.403	21.16	-0.387	20.81	-0.366	20.42	-0.337	20.15
1	-0.452	19.68	-0.437	19.26	-0.417	18.78	-0.389	18.46
0	-0.504	18.58	-0.491	18.08	-0.472	17.55	-0.445	17.19
-1	-0.559	17.87	-0.548	17.32	-0.531	16.68	-0.505	16.29
-2	-0.616	17.41	-0.606	16.74	-0.592	16.01	-0.568	15.55
-3	-0.674	16.96	-0.667	16.15	-0.656	15.29	-0.634	14.74
-4	-0.734	16.46	-0.730	15.54	-0.723	14.52	-0.704	13.82
-5	-0.795	16.00	-0.796	14.96	-0.793	13.78	-0.779	12.89
-6	-0.859	15.65	-0.864	14.54	-0.868	13.19	-0.859	12.03
-7	-0.923	15.46	-0.933	14.37	-0.945	12.87	-0.945	11.42
-8	-0.988	15.48	-1.003	14.49	-1.022	12.93	-1.033	11.20
-9	-1.052	15.67	-1.071	14.89	-1.098	13.46	-1.122	11.52
-10	-1.115	15.97	-1.137	15.54	-1.170	14.41	-1.206	12.43
-11	-1.177	16.35	-1.199	16.30	-1.237	15.66	-1.282	13.95
-12	-1.238	16.72	-1.259	17.04	-1.298	16.95	-1.349	15.94
-13	-1.297	17.06	-1.317	17.67	-1.355	18.18	-1.408	18.07
-14	-1.355	17.35	-1.373	18.19	-1.409	19.17	-1.460	20.07
-15	-1.412	17.59	-1.427	18.57	-1.460	19.92	-1.508	21.69
-16	-1.469	17.82	-1.481	18.83	-1.510	-	-1.553	22.80
-17	-1.524	18.04	-1.533	19.02	-1.559	20.63	-1.596	23.42
-18	-1.580	18.28	-1.586	19.19	-1.607	20.73	-1.639	23.60
-19	-1.634	18.53	-1.638	19.34	-1.656	20.73	-1.681	23.52
- 20	-1.687	18.82	-1.689	19.52	-1.704	20.71	-1.724	23.24

Differential capacity on mercury in aqueous solutions of 0.5 M Na_2SO_4, Saturated with the organic compounds named n-nonylic acid, camphene, α-pinene.
T = 25 ± 1°C. Potentials measured with respect to a Hg/Hg_2SO_4(S), Na_2SO_4(0.5M) electrode. (This reference electrode has a potential of 0.634 Volt with respect to the standard hydrogen electrode.)
Reference: K.G. Baikerikov, S. Sathyanorayana, J. Electroanal. and Interfac. Electrochem. 24 (1970) 333.

n-nonylic acid		camphene		α-pinene	
E Volts	C μ F/cm²	E Volts	C μ F/cm²	E Volts	C μ F/cm²
0.154	134.4	0.205	45.11	0.204	45.48
0.200	109.0	0.252	43.67	0.252	42.61
0.251	96.4	0.302	40.74	0.302	41.54
0.300	83.73	0.352	40.33	0.352	40.13
0.352	75.86	0.401	40.05	0.402	16.14
0.350	63.90	0.452	39.32	0.452	11.45
0.375	55.01	0.502	40.46	0.502	9.49
0.402	48.82	0.551	40.88	0.552	7.79
0.453	36.93	0.601	6.61	0.602	6.98
0.501	29.97	0.651	4.91	0.651	6.56
0.551	26.81	0.702	4.51	0.701	6.23
0.576	24.27	0.801	4.29	0.802	6.16
0.601	18.20	0.902	4.22	0.902	5.76
0.626	11.77	1.001	4.34	1.002	5.67
0.653	6.407	1.102	4.56	1.102	5.66
0.702	2.668	1.152	4.93	1.202	5.52
0.750	1.873	1.201	6.22	1.301	5.70
0.802	1.702	1.251	16.79	1.401	5.81
0.851	1.675	1.300	16.27	1.502	6.05
0.900	1.690	1.400	15.93	1.602	6.41
0.952	1.690	1.502	15.89	1.701	7.79
1.001	1.755	1.601	15.85	1.752	8.96
1.052	1.827	1.701	16.43	1.802	16.98
1.102	1.931	1.801	17.24	1.852	17.30
1.150	2.160	1.902	18.00	1.901	17.98
1.201	2.480			2.00	19.13
1.250	2.952				
1.301	3.640				
1.350	4.099				
1.401	4.777				
1.425	5.359				
1.450	6.219				
1.475	8.03				
1.500	12.20				
1.526	35.51				
1.552	28.03				
1.602	18.59				
1.652	17.40				
1.702	16.90				
1.752	16.80				
1.802	17.40				
1.902	18.01				

Differential capacity of mercury in 0.1 mol l^{-1} salt solutions in ethylene carbonate. T = 40°C

Potential with respect to a calomel electrode in ethylene carbonate using a 0.1 mol l^{-1} tetraethylammonium perchlorate solution saturated with Hg_2Cl_2 and KCl.

Reference: W.R. Fawcett and M.D. Mackey, J. Chem. Soc., Faraday Transactions, 69 (1973) 634.

E/volt	C/μF cm^{-2}		
	$(C_2H_5)_4NBF_4$	$(C_2H_5)_4NBr$	$(C_2H_5)_4NI$
0.70	32.20		
0.65	32.15		
0.60	32.04		
0.55	31.30		
0.50	30.90		
0.45	29.58		
0.40	27.32		
0.35	25.00		
0.30	23.14		
0.25	21.69		
0.20	20.62		
0.15	19.82		
0.10	19.14		
0.05	18.44		
0.00	17.74		
-0.05	17.15		
-0.10	16.65		
-0.15	16.13		
-0.20	15.67		
-0.25	15.25		
-0.30	14.90		
-0.35	14.57		
-0.40	14.23		
-0.45	13.92	34.62	
-0.50	13.64	24.41	63.13
-0.55	13.37	18.57	44.54
-0.60	13.10	15.45	30.84
-0.65	12.81	13.81	21.85
-0.70	12.53	12.91	16.71
-0.75	12.24	12.36	14.03
-0.80	11.91	11.95	12.64
-0.85	11.60	11.60	11.87
-0.90	11.26	11.28	11.35
E^z/volt	-0.067	-0.363	-0.466

Differential capacity of mercury in 0.1 mol l^{-1} solutions of perchlorates in ethylene carbonate. T = 40°C

Potential with respect to a calomel electrode in ethylene carbonate using a 0.1 mol l^{-1} tetraethylammonium perchlorate solution saturated with Hg_2Cl_2 and KCl.

Reference: W.R. Fawcett and M.D. Mackey, J. Chem. Soc., Faraday Transactions, 69 (1973) 634.

E/volts	C/μF cm^{-2}						
Cation	Li^+	Na^+	K^+	Cs^+	$(CH_3)_4N^+$	$(C_2H_5)_4N^+$	$(nC_4H_9)_4N^+$
0.80	34.78	35.18	34.08	34.42	34.68	35.18	34.50
0.75	34.78	35.18	34.05	34.30	34.76	35.24	34.49
0.70	34.46	35.06	33.95	34.30	34.76	35.36	34.46
0.65	34.40	35.03	33.95	34.30	34.73	35.54	34.43
0.60	34.53	34.88	33.95	34.30	34.70	35.54	34.50
0.55	34.53	34.73	33.88	34.30	34.53	35.40	34.74
0.50	34.21	34.51	33.61	34.02	34.22	35.00	34.47
0.45	33.36	33.50	32.85	33.26	33.44	33.97	33.61
0.40	32.02	32.20	31.68	32.21	32.19	32.38	32.34
0.35	30.21	30.56	30.23	30.61	30.61	30.41	30.73
0.30	28.18	28.52	28.33	28.77	28.70	28.21	28.66
0.25	26.04	26.34	26.21	26.54	26.52	25.96	26.70
0.20	23.87	24.15	24.03	24.34	24.39	23.82	24.78
0.15	21.85	22.09	21.93	22.20	22.27	22.00	23.17
0.10	19.92	20.23	20.04	20.25	20.39	20.57	21.96
0.05	18.28	18.57	18.39	18.62	18.77	19.46	20.95
0.	16.90	17.18	17.04	17.32	17.45	18.53	20.05
-0.05	15.80	16.14	16.03	16.24	16.47	17.72	19.21
-0.10	15.06	15.33	15.31	15.53	15.76	17.06	18.21
-0.15	14.55	14.77	14.79	15.04	15.26	16.53	17.17
-0.20	14.22	14.40	14.43	14.70	14.92	16.01	16.06
-0.25	13.98	14.13	14.19	14.40	14.65	15.53	15.01
-0.30	13.76	13.92	13.99	14.19	14.44	15.12	14.02
-0.35	13.58	13.74	13.83	13.99	14.24	14.71	13.13
-0.40	13.37	13.54	13.67	13.80	14.04	14.31	12.35
-0.45	13.19	13.35	13.52	13.60	13.83	13.95	11.64
-0.50	13.01	13.18	13.37	13.40	13.64	13.63	10.99
-0.55	12.84	13.02	13.22	13.23	13.45	13.33	10.47
-0.60	12.70	12.85	13.07	13.05	13.29	13.03	10.06
-0.65	12.55	12.70	12.92	12.97	13.08	12.72	9.61
-0.70	12.38	12.54		12.83	12.73	12.41	9.18
-0.75		12.39		12.69		12.09	8.86
-0.80		12.23		12.55		11.76	8.63
-0.85		12.08		12.39			8.45
E^z/volt	-0.068	-0.067	-0.071	-0.071	-0.076	-0.067	-0.029

Differential capacity of mercury in solutions of KPF_6 in ethylene carbonate.

T = 40°C

Potential with respect to a calomel electrode in ethylene carbonate using a 0.1 mol l^{-1} tetraethylammonium perchlorate solution saturated with Hg_2Cl_2 and KCl.

Reference: W.R. Fawcett and M.D. Mackey, J. Chem. Soc., Faraday Trans., 69 (1973) 634.

	C/μF cm^{-2}			
c/mol l^{-1}	0.10	0.20	0.50	1.00
E/volt				
0.75	25.55		25.17	25.17
0.70	25.64	25.35	25.38	25.42
0.65	25.73	25.50	25.54	25.69
0.60	25.77	25.62	25.71	25.99
0.55	25.70	25.66	25.85	26.14
0.50	25.31	25.57	25.73	26.10
0.45	24.70	25.13	25.29	25.72
0.40	23.92	24.51	24.77	25.29
0.35	22.99	23.76	24.08	24.71
0.30	21.89	22.90	23.33	23.97
0.25	20.79	21.93	22.42	23.10
0.20	19.70	20.92	21.50	22.16
0.15	18.61	19.93	20.52	21.22
0.10	17.66	18.96	19.59	20.31
0.05	16.75	18.05	18.73	19.42
0.	15.98	17.25	17.96	18.62
-0.05	15.38	16.57	17.28	17.88
-0.10	14.96	16.00	16.62	17.22
-0.15	14.61	15.51	16.08	16.60
-0.20	14.38	15.11	15.58	16.05
-0.25	14.18	14.76	15.15	15.59
-0.30	14.03	14.47	14.77	15.16
-0.35	13.89	14.21	14.48	14.79
-0.40	13.76	14.01	14.25	14.46
-0.45	13.62	13.83	13.96	14.16
-0.50	13.48	13.65	13.67	13.90
-0.55	13.32	13.43	13.43	13.66
-0.60	13.20	13.24	13.23	13.43
-0.65	13.06	13.07	13.04	13.22
-0.70	12.91	12.90	12.86	13.01
-0.75	12.76	12.73	12.68	12.82
-0.80	12.62	12.57	12.51	12.63
-0.85	12.49	12.43	12.32	12.42
-0.90	12.34	12.26	12.15	12.23
-0.95	12.16	12.10	11.96	12.06
-1.00	12.01	11.95	11.84	11.89
-1.05	11.86	11.79	11.69	11.72
-1.10	11.71	11.63	11.54	11.57
-1.15	11.58	11.50	11.41	11.43
-1.20	11.45	11.38	11.29	11.32
-1.25	11.32	11.28		
E^z/volt	-0.071	-0.068	-0.064	-0.061

Differential capacity of mercury in propylene carbonate solutions of $NaClO_4$. Potential with respect to Ag 10.5 M $NaClO_4$, 0.05 M $AgClO_4$ in propylene carbonate.

Reference: R. Parsons and T. Biegler. J. Electroanal. Chem. 21 (1969) App.4.

	$C/\mu F\ cm^{-2}$					
$c/mol\ l^{-1}$	0.02	0.05	0.10	0.20	0.50	0.90
E/volts						
0.10		34.04	34.33	34.40		
0.13		33.87	34.21	34.53	35.36	35.85
0.16	32.88	33.45	34.00	34.40	35.36	35.97
0.19	32.96	33.20	33.53	33.75	34.82	35.71
0.22	33.05	33.24	33.53	33.61	34.52	35.33
0.25	33.24	33.49	33.61	33.69	34.56	35.25
0.28	33.51	33.79	33.84	33.88	34.77	35.25
0.31	33.78	34.09	34.11	34.15	35.02	35.29
0.34	34.20	34.48	34.38	34.42	35.29	35.21
0.37	34.47	34.86	34.69	34.80	35.52	35.17
0.40	34.77	35.25	34.99	35.11	35.83	35.32
0.43	34.97	35.63	35.26	35.46	36.29	35.67
0.46	34.89	35.67	35.42	35.69	36.79	36.06
0.49	34.77	35.36	35.50	35.88	36.94	36.56
0.52	34.12	34.98	35.15	35.88	36.86	36.59
0.55	33.35	34.25	34.61	35.38	36.40	36.17
0.58	32.47	33.33	33.72	34.57	35.52	35.44
0.61	31.20	32.18	32.72	33.50	34.60	34.56
0.64	29.86	30.79	31.56	32.31	33.49	33.48
0.67	28.24	29.30	30.22	30.97	32.27	32.38
0.70	26.59	27.61	28.60	29.67	31.04	31.44
0.73	24.78	25.84	27.09	28.21	29.85	30.48
0.76	22.75	23.92	25.51	26.95	28.74	29.44
0.79	20.71	22.15	23.97	25.68	27.51	28.36
0.82	18.58	20.50	22.62	24.38	26.36	27.17
0.85	16.87	19.24	21.39	23.11	25.02	25.94
0.88	15.50	18.20	20.27	21.85	23.80	24.70
0.91	14.58	17.34	19.17	20.62	22.53	23.36
0.94	13.89	16.51	18.09	19.47	21.27	22.09

Capacity in $NaClO_4$ in propylene carbonate (continued)

	$C/\mu F\ cm^{-2}$					
$c/mol\ l^{-1}$	0.02	0.05	0.10	0.20	0.50	0.90
E/volts						
0.97	13.33	15.72	17.17	18.48	20.12	20.97
1.00	12.87	15.09	16.42	17.57	19.08	19.95
1.03	12.54	14.61	15.78	16.77	18.14	19.01
1.06	12.45	14.25	15.24	16.12	17.40	18.18
1.09	12.49	13.98	14.79		16.75	17.45
1.12	12.58	13.79	14.47		16.15	16.76
1.15	12.68	13.65	14.20	14.79	15.63	16.18
1.18	12.79	13.54	13.99		15.19	15.64
1.21	12.89	13.44	13.79		14.78	15.18
1.25	12.99		13.55	13.82	14.30	
1.30	12.95	13.15	13.26	13.45	13.78	14.02
1.35	12.83	12.90	12.93	13.03	13.26	13.44
1.40	12.60	12.61	12.60	12.63	12.76	12.89
1.45	12.33	12.31	12.24	12.25	12.30	12.39
1.50	12.01	11.96	11.89	11.86	11.89	11.93
1.55	11.70	11.64	11.54	11.49	11.48	11.51
1.60	11.37	11.29	11.20	11.15	11.09	11.10
1.65	11.03	10.94	10.85	10.79	10.75	10.74
1.70	10.68	10.60	10.52	10.46	10.42	10.42
1.75	10.39	10.33	10.21	10.17	10.14	10.12
1.80	10.12	10.02	9.95	9.91	9.87	9.83
1.85	9.85	9.77	9.70	9.65	9.62	9.58
1.90	9.62	9.52	9.48	9.43	9.40	9.35
1.95	9.39	9.31	9.26	9.23	9.21	9.17
2.00	9.18	9.12	9.09	9.05	9.03	8.98
2.05	9.01	8.97	8.93	8.89	8.86	8.82
2.10	8.86	8.82	8.76	8.74	8.71	8.67
2.15		8.66	8.62	8.59	8.60	8.55
2.20		8.54	8.51	8.48	8.48	8.44
2.25			8.41	8.40		8.50
E_z/volts	-1.025	-1.018	-1.017	-0.1010	-1.009	-1.010

Differential capacity of KPF_6 and $LiBF_4$ in butyrolactone. T = 25°C

Potential with respect to Ag/AgCl(sat), LiCl(sat) in butyrolactone electrode (pzc possibly in error)

Reference: R. Payne, J. Phys. Chem., 71 (1967) 1548.

Salt	KPF_6		KPF_6		KPF_6		$LiBF_4$	
$c/\text{mol l}^{-1}$	0.010		0.10		1.0		0.10	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
-14	-1.489	10.32					-1.359	7.90
-13	-1.378	8.28			-1.237	8.07	-1.231	7.79
-12	-1.251	7.66			-1.109	7.68	-1.102	7.72
-11	-1.119	7.52	-1.024	7.57	-0.978	7.63	-0.972	7.73
-10	-0.987	7.65	-0.893	7.70	-0.848	7.78	-0.844	7.88
- 9	-0.858	7.95	-0.766	8.03	-0.722	8.12	-0.719	8.18
- 8	-0.736	8.46	-0.645	8.56	-0.603	8.67	-0.600	8.69
- 7	-0.623	9.22	-0.523	9.32	-0.492	9.47	-0.490	9.36
- 6	-0.519	10.11	-0.431	10.27	-0.392	10.53	-0.387	10.18
- 5	-0.425	11.14	-0.338	11.34	-0.302	11.80	-0.293	11.20
- 4	-0.339	12.04	-0.254	12.41	-0.222	13.21	-0.208	12.24
- 3	-0.258	12.63	-0.176	13.37	-0.150	14.68	-0.129	13.10
- 2	-0.179	12.62	-0.103	14.10	-0.085	16.19	-0.054	13.67
- 1	-0.097	11.48	-0.034	14.53	-0.026	17.69	0.018	13.95
0	-0.003	10.07	0.034	15.13	0.028	19.13	0.088	14.55
1	0.097	10.45	0.098	16.11	0.079	20.36	0.155	15.86
2	0.181	13.68	0.158	17.24	0.126	21.53	0.214	17.58
3	0.246	16.95	0.214	18.59	0.172	22.66	0.269	19.39
4	0.301	19.26	0.266	20.02	0.215	23.72	0.318	21.44
5	0.351	21.12	0.314	21.28	0.256	24.78	0.362	23.70
6	0.396	22.53	0.360	22.42	0.296	25.70	0.402	25.86
7	0.440	23.67	0.403	23.60	0.334	26.48	0.440	27.80
8	0.481	24.63	0.445	24.60	0.371	27.13	0.474	29.55
9	0.521	25.47	0.485	25.48	0.408	27.71	0.507	31.08
10	0.560	26.21	0.524	26.28	0.444	28.21	0.539	32.29
11	0.597	26.82	0.561	26.97	0.479	28.62	0.569	33.23
12	0.634	27.08	0.598	27.37	0.513	28.92	0.599	33.94
13	0.671	27.05	0.634	27.50	0.548	29.08	0.628	34.33
14	0.708	26.78	0.671	27.37	0.582	29.09	0.657	34.44
15	0.746	26.22	0.707	27.00	0.617	28.86	0.687	34.26
16	0.785	25.56	0.745	26.43	0.652	28.25	0.716	33.91
17	0.824	24.82	0.783	25.73	0.688	27.51	0.746	33.50
18	0.865	24.07	0.823	24.91	0.725	26.63	0.776	33.07
19			0.864	24.02	0.763	25.67	0.806	32.64
20					0.803	24.68	0.837	32.27
21					0.844	23.68	0.868	31.89
22					0.887	22.88	0.900	31.43

Differential capacity in 1 M KCl

Adsorption of 4-methyl-9-aminoacridine

Reference: R.J. Meakins, J. Appl. Chem. 17 (1967) 156.

$C/\mu F\ cm^{-2}$

	4-methyl-9-aminoacridine					
$c/mol\ l^{-1}$	0	10^{-5}	2×10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}
E/V						
0	268					
0.1	57.0	51.6	46.0	42.4	40.8	32.2
0.2	40.1	37.5	33.7	33.2	31.9	16.2
0.25		35.1				14.9
0.3	38.3	34.5	30.6	28.5	27.8	15.0
0.35		34.7	30.4			
0.4	41.5	35.1	30.6	24.3	24.2	15.4
0.5	39.8	32.8	28.3	21.3	21.4	13.6
0.6	30.9	26.0	23.9	19.2	19.0	11.1
0.7	22.9	20.4	19.2			18.3
0.8	19.0	16.9	16.5	16.9	14.3	13.8
0.9				14.7	10.6	10.6
1.0	16.2	14.4	14.5	12.9	9.35	
1.1					9.84	10.0
1.15			14.7	13.7	15.4	15.8
1.2	16.3	16.6	17.3	20.1	32.0	31.3
1.24						50.7
1.25		25.8	36.7	61.8	91.7	
1.3		20.9	23.5	28.9	38.2	19.4
1.35					16.0	
1.4	17.5	18.1	18.1	15.9	11.8	5.95
1.5						19.0
1.6	19.8	20.2	19.9	19.9	19.8	20.2
1.8	23.4	23.7				23.5

Differential capacity in 1 M KCl (continued)

Adsorption of acridine

			$C/\mu F\ cm^{-2}$		
			Acridine		
$c/mol\ l^{-1}$	0	5×10^{-5}	10^{-4}	10^{-3}	3×10^{-3}
E/V					
0	268	197	220	242	282
0.1	57.0	47.1	50.5	51.2	51.4
0.2	40.1	28.0	37.8	37.2	37.6
0.25			35.9	34.1	30.9
0.263		31.6			
0.3	38.3	9.82	34.4		10.1
0.35				3.85	
0.4	41.5	6.56	7.38	2.95	2.76
0.5	39.8	4.53	2.90		
0.6	30.9	3.73	2.42	1.14	0.810
0.7	22.9				
0.8	19.0	3.10	1.41	1.03	0.780
0.9					0.746
1.0	16.2	2.55	1.63	1.02	
1.1		3.00		1.28	~ 0.82
1.2	16.3	6.80	3.16		
1.25		10.5			readings irregular
1.3		16.4	13.7		
1.32				9.63	
1.35		17.6	17.0	14.2	8.84
1.4	17.5	17.8	17.8	15.6	11.8
1.5			18.5	17.9	14.7
1.6	19.8	19.8	19.7	19.7	17.7
1.8	23.4	24.1	23.8	23.9	23.7
1.9	36.6				30.9

Differential capacity in 1 M KCl (continued)

Adsorption of 9-aminoacridine

				C/μF cm^{-2}		
				9-aminoacridine		
c/mol l^{-1}	0	10^{-5}	2×10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}
E/V						
0	268	212	168	150		71.1
0.1	57.0	49.3	41.9	39.7	29.2	16.2
0.2	40.1	36.2	32.2	30.9	20.8	10.0
0.3	38.3	32.8	28.7	27.8	18.8	8.45
0.4	41.5	34.1	27.9	25.3	16.9	7.34
0.45		33.9	27.2			
0.5	39.8	32.4	25.9	23.7	16.3	5.89
0.6	30.9	26.5	22.2	20.2	15.2	5.42
0.7	22.9	20.3	19.1	19.1	13.7	5.29
0.8	19.0	17.5	17.3	16.0	7.89	7.61
0.9				12.3		
1.0	16.2	15.0	15.0	10.2	7.49	7.2
1.1						
1.15					8.40	
1.2	16.3	15.9	15.2	11.3		11.0
1.25				18.9	20.4	21.3
1.27						98.6
1.277		34.3			110	
1.28			49.6	104		
1.3		25.2	35.5	76.9	120	70.0
1.4	17.5	18.5	18.6	19.0	19.1	19.3
1.6	19.8	20.0	20.1	20.0	20.0	20.2
1.8	23.4	23.8	23.8	23.9	23.8	23.7

Differential capacity in 1 M KCl

Adsorption of isoquinoline hydrochloride

Reference: R.J. Meakins, J. Appl. Chem. 17 (1967) 156.

$C/\mu F\ cm^{-2}$

	isoquinoline hydrochloride				
$c/mol\ l^{-1}$	0	3×10^{-4}	10^{-3}	3×10^{-3}	3×10^{-2}
E/V					
0	268	261	412	309	364
0.1	57.0	54.3	66.3	63.1	64.1
0.2	40.1	37.6	43.5	43.2	43.9
0.25			39.4		
0.3	38.3	34.3	37.1	36.6	36.2
0.4	41.5	30.9	34.0	33.1	31.5
0.45		27.6			
0.5	39.8	23.1	30.7	29.1	27.0
0.6	30.9	17.0	25.8	24.6	22.7
0.7	22.9	14.9	21.5		
0.8	19.0	14.8	18.1	17.5	15.5
0.83					9.88
0.85			17.0	16.3	
0.9			15.9	12.6	6.97
0.95			12.0	7.69	
1.0	16.2	17.0	7.65	7.11	7.09
1.1		18.4	7.08	7.91	
1.15			6.79		
1.2	16.3	19.0	5.07	5.38	6.42
1.25		19.0	3.92		
1.3		18.6	21.3	4.11	
1.32			22.4		
1.4	17.5	18.9	17.2	2.91	7.41
1.5			11.7	2.78	
1.6	19.8	20.8	11.6	2.58	6.71
1.7			13.2	~3.2	
1.8	23.4	24.3	20.9	~4.4	
1.9	36.6		31.8	11.2	

Differential capacity in 1 M KCl (continued)
Adsorption of quinoline hydrochloride

$C/\mu F\ cm^{-2}$

	quinoline hydrochloride				
$c/mol\ l^{-1}$	0	3×10^{-4}	10^{-3}	3×10^{-3}	3×10^{-2}
E/V					
0	268	248	269	263	291
0.1	57.0	51.7	54.2	56.6	57.0
0.2	40.1	36.2	37.7	39.6	41.4
0.3	38.3	33.1	34.2	35.0	34.5
0.4	41.5	30.1	29.9	30.7	28.7
0.45		26.8			
0.5	39.8	22.8	22.2	25.2	23.9
0.6	30.9	17.0	16.9	20.5	20.7
0.65			16.0		
0.7	22.9	14.7	9.09	19.2	18.3
0.8	19.0	14.5	6.91	11.4	10.4
0.9				8.03	3.58
0.93			8.72		
0.95				3.34	
1.0	16.2	16.3	4.66	1.84	2.18
1.1		17.3	2.21	1.36	1.33
1.2	16.3	17.8	2.27	1.23	1.18
1.3		17.7			1.09
1.35			18.8	5.87	
1.4	17.5	18.0	18.2	10.4	
1.5				16.3	~ 0.77
1.6	19.8	19.7	19.8	19.7	
1.7					21.0
1.8	23.4	23.6	23.8	23.7	

Differential capacity on mercury electrode in guanidinium chloride at 25°C. Given are the capacity in μF cm^{-2}, potential in V vs. a 0.1 mol l^{-1} KCl calomel electrode dipping into a 0.1 mol l^{-1} guanidinium chloride solution.

Reference: L.M. Baugh and R. Parsons. J. Electroanal. Chem. 58 (1975) 229.

	C/μF cm^{-2}				
c/mol l^{-1}	1.0	0.63	0.40	0.25	0.16
-E/V					
0.09					81.42
0.10					73.83
0.11			80.63	80.79	68.40
0.12			73.04	73.07	62.96
0.13		73.21	67.50	67.21	58.66
0.14	72.43	68.28	63.19	61.34	54.45
0.15	67.69	63.44	59.50	57.43	50.97
0.16	63.68	60.46	56.42	54.03	48.81
0.17	60.29	57.27	53.65	50.64	46.86
0.18	57.30	54.70	50.99	43.17	44.81
0.19	55.04	32.24	49.04	46.11	43.28
0.20	52.78	50.18	47.19	44.67	41.75
0.21	50.72	48.23	45.45	42.92	
0.22	48.66	46.68	44.01	41.58	
0.23	46.91	44.83	42.47	40.34	38.76
0.24	45.47	43.29	41.24		
0.25	44.03	41.95	40.11		
0.26	42.80	40.82	39.19	37.67	36.92
0.27	41.56	39.90			
0.28	40.43	38.77			
0.29	39.40	38.04	36.62	35.92	35.69
0.32	36.73	35.78	34.98	34.89	35.06
0.35	34.77	34.14	33.85	34.48	35.07
0.38	33.33	33.01	33.44	34.38	35.38
0.41	32.30	32.49	33.44	34.89	36.20
0.44	31.58	32.29	33.75	35.61	
0.47	31.38	32.49	34.37	36.33	37.33
0.50	31.48	33.62	35.60	36.74	37.33
0.53	31.89	23.62	33.60	37.15	36.61
0.56	32.30	34.24	36.11	36.85	35.58
0.59	32.92	34.76	36.01	35.82	33.33
0.62	33.44	34.76	35.39	33.86	30.76

Differential capacity on mercury in guanidinium chloride (cont.)

	$C/\mu F\ cm^{-2}$				
$c/mol\ l^{-1}$	1.0	0.63	0.40	0.25	0.16
$-E/V$					
0.65	33.54	34.45	33.96	31.60	28.20
0.68	33.44	33.52	32.01	29.13	25.84
0.71	32.92	31.88	29.75	26.66	23.69
0.74	31.79	29.92	27.19	24.39	21.95
0.77	30.25	28.87	25.03	22.33	20.41
0.80	28.50	25.91	23.08	20.89	19.18
0.83	26.65	23.86	21.44	19.55	18.25
0.86	25.00	22.21	20.00	18.63	17.43
0.89	23.35	20.87	18.88	17.60	16.82
0.92	22.02	19.64	18.06	16.98	16.31
0.95	20.68	18.71	17.44	16.47	15.89
0.98	19.75	17.99	16.82	16.16	15.59
1.01	19.03	17.48	16.32	15.85	15.38
1.04	18.42	17.07	16.21	15.64	15.28
1.07	18.00	16.86	16.11	15.54	15.18
1.10	17.80	16.66	16.00	15.54	15.18
1.13	17.59	16.56	16.00	15.54	15.28
1.16	17.49	16.56	16.00	15.64	15.38
1.19	17.49	16.66	16.11	15.75	15.48
1.22	17.59	16.77	16.31	15.85	15.59
1.25	17.80	16.97	16.52	16.06	15.79
1.28	18.00	17.17	16.72	16.26	16.00
1.31	18.21	17.38	16.93	16.57	16.20
1.34	18.52	17.69	17.23	16.88	16.51
1.37	18.83	17.99	17.54	17.19	16.82
1.40	19.14	18.30	17.85	17.50	17.13
1.43	19.44	18.61	18.16	17.81	17.43
1.46	19.75	18.92	18.44	18.11	17.74
1.49	20.06	19.23	18.77	18.42	18.15
1.52	20.37	19.54	19.03	18.73	18.46
1.55	20.68	19.85	19.39	19.04	18.87
1.58	20.99	20.15	19.70	19.45	19.18
1.61		20.46	20.00	19.66	19.48
1.64			20.31	19.97	

Differential capacity for aqueous solutions of guanidinium chloride. Potentials with respect to a 0.1 mol l^{-1} KCl calomel electrode dipping into a 0.1 mol l^{-1} guanidinium chloride solution. T = 25°C

Reference: L.M. Baugh and R. Parsons. J. Electroanal. Chem. 58 (1975) 229.

	$C/\mu F\ cm^{-2}$				
c/mol l^{-1}	0.016	0.025	0.040	0.063	0.100
-E/V					
0.02	87.61				
0.03	71.63	87.32			
0.04	62.87	71.91			
0.05	57.61	63.69	77.74		
0.06	52.56	58.35	68.79	88.64	
0.07	50.19	53.83	62.40	76.71	88.64
0.08	47.82	51.78	57.15	69.31	78.02
0.11	42.26	44.28	48.60	55.01	60.19
0.14	39.17	40.58		46.48	49.57
0.17	37.10	38.32	39.23	41.44	43.39
0.20	35.55	36.68	37.38	38.66	40.09
0.23	35.55	36.06	36.14	36.81	37.83
0.26	35.55	35.75	36.04	35.99	36.59
0.29	35.55	35.54	36.14	35.99	35.87
0.32	35.55	35.96	36.25	36.30	35.97
0.35	35.56	36.37	36.56	36.71	36.38
0.38	35.66	36.37	37.07	37.33	37.00
0.41	34.01	35.54	36.76	37.63	37.62
0.44	32.47	33.80	35.94	37.33	38.13
0.47	29.37	31.64	34.50	36.50	38.03
0.50	26.28	29.18	32.44	35.06	37.21
0.53	23.91	26.61	29.76	32.90	35.56
0.56	22.06	24.83	27.39	30.44	33.39
0.59	21.23	22.98	25.23	28.07	30.92
0.62	20.61	21.99	23.58	25.71	28.24
0.65	19.99	20.92	22.04	23.75	25.87
0.68	19.48	20.14	21.01	22.31	23.71
0.71	19.07	19.31	20.08	20.77	22.16
0.74	18.55	18.70	19.26	19.84	20.72

Capacity on mercury in guanidinium chloride (cont.)

			C/μF cm^{-2}		
c/mol l^{-1}	0.016	0.025	0.040	0.063	0.100
-E/V					
0.77	18.04	18.80	18.43	18.82	19.58
0.80	17.53	17.46	17.71	18.10	18.55
0.83	17.01	16.95	17.09	17.38	17.73
0.86	16.59	16.54	16.68	16.86	17.11
0.89	16.18	16.13	16.27	15.35	16.49
0.92	15.87	15.82	15.86	15.94	16.08
0.95	15.56	15.51	15.55	15.63	15.77
0.98	15.36	15.31	15.34	15.42	15.46
1.01	15.15	15.10	15.14	15.22	15.25
1.04	15.05	15.00	15.03	15.12	15.15
1.07	14.94	14.90	14.93	15.01	15.05
1.10	14.84	14.80	14.93	15.01	15.05
1.13	14.84	14.80	14.93	15.01	15.15
1.16	14.84	14.90	14.93	15.01	15.25
1.19	14.84	15.00	15.03	15.12	15.36
1.22	14.94	15.10	15.14	15.32	15.46
1.25	15.05	15.20	15.34	15.33	15.56
1.28	15.25	15.31	15.55	15.73	15.77
1.31	15.46	15.51	15.75	15.94	15.98
1.34	15.67	15.82	15.96	16.14	16.18
1.37	15.98	16.13	16.27	16.45	16.49
1.40	16.28	16.44	16.58	16.76	16.80
1.43	16.59	16.75	16.89	17.07	17.11
1.46	16.90	17.05	17.20	17.38	17.52
1.49	17.21	17.36	17.51	17.69	17.93
1.52	17.52	17.77	17.92	18.10	18.35
1.55	17.93	18.08	18.33	18.51	18.65
1.58	18.35	18.39	18.74	18.82	18.96
1.61	18.76	18.70			
1.64	19.07				

Differential capacity on mercury electrode in mixture of x M guanidinium chloride + 3x M NaCl at 25°C - given are the capacity in μF cm^{-2}, potential in V vs 0.1 M KCl calomel electrode dipping into a 0.1 M guanidinium chloride solution.

Reference: L.M. Baugh and R. Parsons. J. Electroanal. Chem. 58 (1975) 229.

	C/μF cm^{-2}							
c/mol l^{-1}	0.4	0.25	0.16	0.10	0.063	0.025	0.016	0.01
-E/V								
0.02								80.71
0.03						79.28	75.73	68.67
0.04					79.39	69.46	67.42	63.18
0.05	79.62	80.03		77.64	71.46	62.84	61.21	57.37
0.06	73.21	73.21		70.49	65.28	58.71	56.95	53.43
0.07	68.35	67.73		64.68	60.92	54.88	53.94	50.42
0.08	64.12	63.07		59.81	56.67	51.58	50.52	48.13
0.09	60.80	59.46		56.49	53.55	48.89	48.55	45.96
0.10	57.90	56.35		53.18	50.44	46.72	46.06	44.50
0.11	55.32	53.67	53.98	50.27	48.15	45.17	44.19	42.95
0.12	52.94	51.29	50.56	48.10	46.18			
0.13	50.87	49.22	48.39	46.23	44.52			
0.14	48.91	47.15	46.12	44.47	43.07	40.93	40.26	39.32
0.15	46.94	45.39	44.67	42.91	41.93			
0.16	45.29	43.84	43.01	41.77	40.78			
0.17	43.95	42.60	41.67	40.74	39.75	38.35	38.07	37.55
0.18	42.70	41.26	40.43	39.70				
0.19	41.36	40.12	39.50	38.77				
0.20	40.22	39.29	38.57	38.15	37.57	37.11	36.83	36.41
0.23	37.53	36.81	36.40	36.59	36.32	36.38	36.52	36.00
0.26	35.26	34.95	35.05	35.45	35.80	36.38	36.62	36.00
0.29	33.71	33.81	34.23	35.24	35.80	36.80	36.82	36.21
0.32	32.58	33.19	34.02	35.24	36.32	37.31	37.24	36.83
0.35	31.95	32.98	34.12	35.97	36.95	37.93	37.66	36.93
0.38	31.64	33.09	34.54	36.49	37.67	38.24	37.76	36.52
0.41	31.74	33.61	35.26	37.32	38.29	37.83	37.03	35.58
0.44	32.16	34.12	35.98	37.83	38.29	36.80	35.48	33.61
0.47	32.57	34.95	36.50	37.83	37.46	34.73	33.30	30.81
0.50	33.09	35.36	36.50	37.11	36.22	32.56	30.60	28.32
0.53	33.61	35.67	36.09	35.87	34.14	29.97	28.01	25.52
0.56	34.02	35.47	34.74	33.90	31.44	27.49	25.62	23.65
0.59	33.92	34.64	32.78	31.30	28.85	25.12	23.86	22.30
0.62	33.50	33.19	30.50	28.40	26.36	23.36	22.51	21.37
0.65	32.47	31.23	28.23	25.91	24.28	21.91	21.27	20.54

Differential capacity on mercury - x M guanidinium chloride + 3x M NaCl (continued)

	$C/\mu F\ cm^{-2}$							
$c/mol\ l^{-1}$	0.4	0.25	0.16	0.10	0.063	0.025	0.016	0.01
-E/V								
0.68	31.02	29.16	25.64	23.84	22.31	20.67	20.33	19.92
0.71	29.37	26.88	23.78	21.87	20.86	19.64	19.50	19.30
0.74	27.50	24.82	21.92	21.42	19.61	18.60	18.78	18.67
0.77	25.54	22.96	20.37	19.28	18.38	17.95	18.15	18.05
0.80	23.89	21.40	19.23	18.24	17.85	17.36	17.53	17.64
0.83	22.44	20.15	18.30	17.52	17.12	16.85	17.01	17.12
0.86	21.08	19.03	17.47	16.79	16.60	16.43	16.49	16.70
0.89	20.06	18.30	16.85	16.38	16.19	16.02	16.18	16.29
0.92	19.34	17.68	16.44	16.07	15.88	15.71	15.87	15.98
0.95	18.61	17.27	16.13	15.86	15.67	15.50	15.66	15.66
0.98	18.20	16.96	15.92	15.65	15.46	15.30	15.46	15.46
1.01	17.78	16.75	15.82	15.55	15.36	15.19	15.25	15.25
1.04	17.58	16.54	15.82	15.55	15.36	15.09	15.15	15.15
1.07	17.47	16.44	15.82	15.55	15.36	14.99	15.04	15.04
1.10	17.37	16.44	15.92	15.55	15.36	14.99	15.04	15.14
1.13	17.37	16.54	16.03	15.65	15.36	14.99	15.04	15.04
1.16	17.37	16.65	16.12	15.76	15.46	15.09	15.04	15.04
1.19	17.47	16.75	16.23	15.86	15.57	15.19	15.15	15.04
1.22	17.58	16.85	16.34	16.07	15.77	15.30	15.25	15.15
1.25	17.68	17.06	16.54	16.27	15.98	15.50	15.35	15.25
1.28	17.89	17.27	16.75	16.48	16.19	15.71	15.56	15.35
1.31	18.10	17.47	16.96	16.69	16.40	15.92	15.77	15.56
1.34	18.30	17.68	17.16	16.90	16.60	16.12	15.98	15.77
1.37	18.51	17.89	17.37	17.10	16.81	16.33	16.18	15.98
1.40	18.72	18.10	17.68	17.31	17.02	16.54	16.39	16.18
1.43	18.42	18.30	17.89	17.62	17.23	16.74	16.60	16.49
1.46	19.13	18.61	18.20	17.93	17.54	17.05	16.81	16.70
1.49	19.34	18.82	18.41	18.14	17.85	17.26	17.12	17.01
1.52	19.54	19.03	18.72	18.35	18.06	17.57	17.43	17.22
1.55	19.75	19.34	18.92	18.66	18.37	17.88	17.74	17.53
1.58	19.95	19.54	19.13	18.97	18.58	18.09	18.05	17.84
1.61	20.16	19.75	19.34	19.18	18.78	18.40	18.36	18.15
1.64	20.38	19.96	19.65	19.38	19.10	18.60	18.57	18.36
1.67	20.47	20.16	19.85	19.59	19.41	18.92		18.67
1.70	20.68	20.37	19.06	19.90				
1.73	20.89	20.58						

Differential capacity on mercury electrode in mixture of x M guanidinium chloride + (1 - x) M NaCl at 25°C. Given are the capacity in μF cm^{-2}, potential in V vs a 0.1 M KCl calomel electrode dipping into a 0.1 M guanidinium chloride solution.

Reference: L.M. Baugh and R. Parsons. J. Electroanal. Chem. 58 (1975) 229.

	C/μF cm^{-2}									
c/mol l^{-1}	0.63	0.40	0.25	0.16	0.10	0.063	0.04	0.025	0.016	0.01
-E/V										
0.14	76.58	76.71	76.71	76.71	75.89	74.93	74.35	73.22	72.94	72.12
0.15	70.41	70.33	70.33	70.33	69.43	68.44	68.09	66.76	66.45	65.45
0.16	66.09	65.90	65.59	65.18	64.30	63.90	63.07	62.04	61.82	61.30
0.17	62.19	62.09	61.37	61.06	60.20	59.78	59.48	58.45	57.90	57.80
0.18	59.10	59.00	53.39	57.56	56.71	56.38	55.89	55.38	54.81	54.71
0.19	56.22	56.02	55.40	54.47	54.25	53.59	52.91	52.09	52.03	51.93
0.20	53.86	53.65	52.83	52.21	51.48	51.33	50.76	49.94	49.86	49.66
0.21	51.60	51.49	50.46	50.04	49.53	48.96		47.89	48.11	47.80
0.22	49.75	49.22	43.60	48.09	47.17	47.10		46.04	45.95	45.95
0.23	47.69	47.37	46.65	46.34	45.63	45.35	45.02	44.81	44.71	44.40
0.26	43.17	42.84	42.22	41.81	41.53	41.43	41.22	41.12	41.31	41.21
0.29	39.88	39.44	39.03	38.72	38.56	38.75	38.56	38.66	38.84	39.05
0.32	37.00	36.86	36.66	36.66	36.51	36.90	37.02	37.33	37.60	37.81
0.35	34.95	34.91	35.01	35.11	35.28	35.87	36.10	36.61	37.19	37.50
0.38	33.61	33.67	33.78	34.29	34.56	35.35	35.99	36.51	37.30	37.60
0.41	32.69	32.95	33.26	33.78	34.46	35.35	36.10	36.92	37.71	38.43
0.44	32.17	32.54	33.16	33.78	34.66	35.87	36.61	37.53	38.63	39.46
0.47	31.97	32.64	33.36	34.29	35.28	36.69	37.53	38.56	39.56	40.39
0.50	32.17	32.95	33.88	34.91	36.10	37.62	38.56	39.39	41.59	41.31
0.53	32.69	33.57	34.60	35.73	37.12	38.44	39.28	40.30	41.11	41.73
0.56	33.20	34.29	35.42	36.56	37.94	39.17	39.89	40,51	41.21	41.42
0.59	33.82	34.91	35.94	37.07	38.35	39.37	39.58	39.58	40.08	40.08
0.62	34.33	35.42	36.35	37.07	37.94	38.55	38.46	37.94	38.12	37.81
0.65	34.43	35.32	35.94	36.25	36.92	37.10	36.40	35.69	35.54	35.03
0.68	34.13	34.80	35.11	34.91	35.07	34.73	33.84	33.23	32.66	32.25
0.71	33.41	33.67	33.57	32.95	32.82	32.05	31.28	30.15	29.88	29.36
0.74	32.07	32.02	31.51	30.58	30.25	29.48	28.30	27.48	27.10	26.68
0.77	30.42	29.97	29.35	28.21	27.69	26.90	26.05	25.12	24.93	24.42
0.80	28.57	27.80	26.88	25.85	25.33	24.63	23.79	33.97	22.97	22.67
0.83	26.42	25.85	24.92	23.79	23.38	22.88	22.15	21.33	21.43	21.12

Differential capacity on mercury x M guanidinium chloride + (1 - x) M NaCl
(continued)

	$C/\mu F\ cm^{-2}$									
$c/mol\ l^{-1}$	0.63	0.40	0.25	0.16	0.10	0.063	0.04	0.025	0.016	0.01
-E/V										
0.86	24.69	23.89	23.07	22.14	21.64	21.13	20.61	20.10	20.09	19.99
0.89	23.02	22.35	21.52	20.70	20.30	19.89	19.59	19.07	19.06	18.96
0.92	21.59	20.90	20.18	19.56	19.18	18.96	18.66	18.25	18.34	18.13
0.95	20.35	19.77	19.15	18.64	18.36	18.14	17.84	17.64	17.62	17.62
0.98	19.32	18.84	18.33	17.92	17.74	17.62	17.33	17.13	17.21	17.10
1.01	18.71	18.23	17.81	17.30	17.23	17.11	16.92	16.72	16.79	16.79
1.04	18.09	17.71	17.30	16.99	16.82	16.80	16.61	16.41	16.48	16.48
1.07	17.68	17.30	16.99	16.68	16.51	16.49	16.41	16.20	16.28	16.28
1.10	17.47	17.09	16.68	16.48	16.31	16.39	16.20	16.10	16.18	16.18
1.13	17.27	16.89	16.58	16.37	16.31	16.29	16.10	16.00	16.07	16.07
1.16	17.17	16.78	16.58	16.37	16.31	16.18	16.10	16.00	16.07	16.07
1.19	17.17	16.78	16.58	16.37	16.31	16.18	16.10	16.00	16.07	16.07
1.22	17.17	16.78	16.58	16.37	16.31	16.18	16.10	16.00	16.07	16.07
1.25	17.27	16.89	16.68	16.48	16.31	16.28	16.20	16.10	16.07	16.07
1.28	17.47	17.09	16.78	16.58	16.41	16.39	16.31	16.20	16.18	16.18
1.31	17.68	17.30	16.89	16.68	16.51	16.49	16.41	16.31	16.28	16.28
1.34	17.88	17.51	17.09	16.78	16.61	16.59	16.51	16.41	16.38	16.38
1.37	18.09	17.71	17.30	16.99	16.82	16.80	16.61	16.51	16.59	16.59
1.40	18.40	17.92	17.51	17.20	17.02	17.01	16.82	16.72	16.79	16.69
1.43	18.71	18.12	17.71	17.40	17.23	17.21	17.02	16.92	17.00	16.90
1.46	19.02	18.43	17.92	17.61	17.43	17.42	17.23	17.13	17.10	17.10
1.49	19.32	18.74	18.12	17.81	17.64	17.62	17.43	17.33	17.31	17.31
1.52	19.63	18.95	18.43	18.02	17.84	17.83	17.64	17.54	17.51	17.51
1.55	19.94	19.26	18.64	18.23	18.05	18.04	17.84	17.74	17.72	17.72
1.58	20.15	19.46	18.84	18.54	18.25	18.24	18.05	17.94	17.93	17.93
1.61	20.35	19.77	19.05	18.74	18.46	18.45	18.25	18.15	18.13	18.13
1.64	20.56	19.98	19.26	18.95	18.66	18.65	18.46	18.36	18.44	18.44
1.67	20.65	20.18	19.56	19.15	18.87	18.86	18.77	18.66	18.75	18.75
1.70	20.76	20.39	19.77	19.46	19.18	19.07	19.07	18.87	18.96	19.06
1.73	21.14	20.59	19.98		19.48	19.38	19.38	19.18	19.27	19.37
1.76						19.69	19.69	19.48		19.68
1.79						19.99	20.00	19.79		19.99

Differential capacity on mercury electrode in 0.25 M NaF - adsorption of acetonitrile. Given are capacity in μF cm^{-2}, potential in V vs SCE.

Reference: A. de Battisti and S. Trasatti. J. Electroanal. Chem. 48 (1973) 213.

			C/μF cm^{-2}			T = 25°C
c/mol l^{-1}	0	0.031	0.076	0.153	0.252	0.382
-E/V						
0.0000	28.23	28.29	28.51	28.71	29.28	29.74
0.0500	28.69	27.74	27.94	28.17	28.81	29.16
0.1000	27.25	27.28	27.48	27.68	28.32	28.50
0.1500	26.89	26.90	27.04	27.19	27.72	27.70
0.2000	26.53	26.48	26.54	26.58	26.92	26.65
0.2500	26.03	25.95	25.90	25.81	25.94	25.39
0.3000	25.38	25.23	25.08	24.95	24.76	24.02
0.3500	24.59	24.37	24.16	23.91	23.55	22.65
0.4000	23.75	23.58	23.24	22.91	22.40	21.42
0.4500	22.94	22.66	22.38	22.00	21.41	20.38
0.5000	22.19	21.92	21.60	21.17	20.57	19.54
0.5500	21.41	19.54	20.85	20.41	19.82	18.82
0.6000	20.59	20.35	20.09	19.66	19.13	18.18
0.6500	19.74	19.54	19.32	18.96	18.51	17.64
0.7000	18.92	18.75	18.57	18.29	17.94	17.19
0.7500	18.13	18.02	17.89	17.67	17.44	16.83
0.8000	17.46	17.37	17.29	17.14	17.00	16.53
0.8500	16.89	16.82	16.79	16.68	16.66	16.31
0.9000	16.42	16.38	16.38	16.34	16.38	16.15
0.9500	16.09	16.06	16.10	16.09	16.17	16.05
1.0000	15.88	15.86	15.90	15.93	16.04	16.01
1.0500	15.76	15.74	15.81	15.85	15.99	16.04
1.1000	15.72	15.72	15.81	15.85	16.01	16.11
1.1500	15.76	15.77	15.87	15.91	16.07	16.22
1.2000	15.88	15.89	15.99	16.04	16.20	16.36
1.2500	16.06	16.06	16.15	16.22	16.37	16.53
1.3000	16.28	16.28	16.37	16.44	16.59	16.75
1.3500	16.55	16.55	16.62	16.70	16.84	16.99
1.4000	16.85	16.85	16.93	17.00	17.12	17.27
1.4500	17.20	17.23	17.27	17.33	17.45	17.59
1.5000	17.58	17.56	17.65	17.70	17.81	17.94
E(σ=0)/V	-0.4324	-0.4312	-0.4292	-0.4272	-0.4227	-0.4176

Capacity in 0.25 M NaF - adsorption of acetonitrile (continued)

	$C/\mu F\,cm^{-2}$					
$c/mol\ l^{-1}$	0.572	0.765	0.956	1.148	1.337	1.528
$-E/V$						
0.0000	30.79	31.90	32.73	34.02	35.24	36.70
0.0500	30.32	31.49	32.37	33.66	34.74	36.12
0.1000	29.65	30.73	31.28	32.30	32.93	33.80
0.1500	28.58	29.37	29.52	30.00	29.95	30.13
0.2000	27.17	27.47	27.19	27.09	26.55	26.21
0.2500	25.48	25.35	24.69	24.20	23.42	22.81
0.3000	23.71	23.23	22.40	21.70	20.86	20.19
0.3500	22.08	21.36	20.45	16.69	18.89	18.25
0.4000	20.68	19.85	18.91	18.15	17.41	16.83
0.4500	19.53	18.64	17.71	16.97	16.29	15.77
0.5000	18.62	17.70	16.77	16.07	15.43	14.95
0.5500	17.87	16.99	16.09	15.39	14.77	14.33
0.6000	17.29	16.43	15.55	14.88	14.27	13.86
0.6500	16.85	16.04	15.19	14.52	13.92	13.51
0.7000	16.52	15.80	14.98	14.32	13.71	13.29
0.7500	16.31	15.68	14.92	14.31	13.65	13.19
0.8000	16.18	15.68	14.99	14.41	13.71	13.24
0.8500	16.10	15.77	15.17	14.65	13.94	13.45
0.9000	16.10	15.91	15.44	15.02	14.35	13.82
0.9500	16.13	16.10	15.78	15.50	14.90	14.38
1.0000	16.19	16.29	16.13	15.99	15.58	15.13
1.0500	16.29	16.50	16.50	16.53	16.31	16.03
1.1000	16.42	16.68	16.81	17.02	17.05	16.97
1.1500	16.54	16.86	17.08	17.45	17.69	17.84
1.2000	16.68	17.02	17.30	17.77	18.13	18.51
1.2500	16.86	17.19	17.48	17.98	18.42	18.93
1.3000	17.05	17.38	17.66	18.15	18.60	19.17
1.3500	17.28	17.59	17.85	18.32	18.74	19.30
1.4000	17.53	17.84	18.07	18.51	18.89	19.38
1.4500	17.84	18.12	18.31	18.72	19.04	19.48
1.5000	18.18	18.43	18.60	18.96	19.25	19.60
$E(\sigma=0)/V$	-0.4086	-0.3995	-0.3904	-0.3801	-0.3712	-0.3608

Differential capacity on mercury electrode in 0.25 M NaF - adsorption of propionitrile. Given are the capacity in μF cm^{-2}, potential in V vs SCE.

Reference: B.A. Abd. el Nabey and S. Trasatti. J. Chem. Soc. Farad. Trans. 1 71 (1975) 1230.

				C/μF cm^{-2}					T = 25°C
c/mol l^{-1}	0	0.0283	0.0425	0.0623	0.0850	0.0963	0.1076	0.1190	0.1275
-E/V									
0.0000	28.43	28.68	28.78	29.03	29.27	29.47	29.63	29.72	29.87
0.0500	27.82	28.10	28.28	28.56	28.95	29.15	29.35	29.50	29.70
0.1000	27.36	27.72	27.90	28.21	28.66	28.92	29.11	29.35	29.56
0.1500	26.92	27.33	27.47	27.83	28.27	28.56	28.74	29.00	29.23
0.2000	26.49	26.85	26.94	27.26	27.53	27.89	28.04	28.25	28.45
0.2500	26.05	26.22	26.21	26.43	26.82	26.79	26.92	26.99	27.12
0.3000	25.38	25.33	25.28	25.31	25.54	25.38	25.45	25.36	25.42
0.3500	24.61	24.34	24.18	24.04	24.00	23.76	23.70	23.54	23.51
0.4000	23.78	23.32	23.12	22.70	22.52	22.22	22.06	21.85	21.67
0.4500	22.96	22.39	22.06	21.60	21.23	20.86	20.63	20.33	20.10
0.5000	22.21	21.50	21.16	20.60	20.10	19.69	19.41	19.04	18.79
0.5500	21.39	20.69	20.32	19.73	19.18	18.71	18.45	18.04	17.76
0.6000	20.62	19.92	19.55	18.97	18.43	17.99	17.69	17.30	16.98
0.6500	19.77	19.17	18.84	18.32	17.83	17.43	17.18	16.76	16.51
0.7000	18.95	18.46	18.22	17.79	17.40	17.06	16.83	16.48	16.25
0.7500	18.18	17.83	17.65	17.35	17.10	16.82	16.65	16.39	16.20
0.8000	17.49	17.29	17.18	17.01	16.89	16.67	16.56	16.39	16.27
0.8500	16.94	16.84	16.79	16.71	16.73	16.59	16.51	16.46	16.41
0.9000	16.47	16.49	16.49	16.50	16.62	16.52	16.53	16.53	16.54
0.9500	16.14	16.21	16.28	16.33	16.53	16.48	16.53	16.60	16.66
1.0000	15.90	16.05	16.12	16.23	16.47	16.44	16.53	16.62	16.71
1.0500	15.77	15.95	16.04	16.18	16.45	16.44	16.53	16.62	16.74
1.1000	15.74	15.95	16.04	16.20	16.46	16.45	16.53	16.65	16.77
1.1500	15.77	15.99	16.12	16.25	16.51	16.50	16.59	16.70	16.80
1.2000	15.87	16.12	16.21	16.36	16.61	16.57	16.67	16.77	16.87
1.2500	16.05	16.27	16.37	16.51	16.74	16.70	16.78	16.88	16.96
1.3000	16.25	16.47	16.57	16.71	16.92	16.86	16.93	17.02	17.08
1.3500	16.51	16.72	16.83	16.92	17.13	17.07	17.12	17.19	17.27
1.4000	16.83	17.00	17.11	17.21	17.39	17.31	17.36	17.42	17.48
1.4500	17.14	17.34	17.44	17.51	17.70	17.60	17.66	17.70	17.73
1.5000	17.60	17.69	17.79	17.86	18.01	17.88	17.95	17.95	18.02
1.5500	17.82	18.07	18.18	18.22	18.37	18.23	18.26	18.30	18.37
1.6000	18.35	18.46	18.01	18.64	18.80	18.64	18.69	18.69	18.72
$-E(\sigma=0)$/V	0.4298	0.4274	0.4247	0.4209	0.4158	0.4134	0.4107	0.4074	0.4049

Capacity in 0.25 M NaF - adsorption of propionitrile (continued)

	$C/\mu F\ cm^{-2}$							
$c/mol\ l^{-1}$	0.1473	0.1699	0.1926	0.2266	0.2832	0.3399	0.4532	0.5664
$-E/V$								
0.0000	30.21	30.47	30.69	31.38	32.29	33.67	36.25	39.89
0.0500	30.09	30.45	30.82	31.66	32.97	34.84	38.86	44.37
0.1000	30.00	30.42	30.93	31.89	33.56	35.86	40.81	45.49
0.1500	29.64	30.09	30.64	31.64	33.33	35.40	38.17	36.58
0.2000	28.78	29.14	29.60	30.36	31.47	32.36	30.94	26.46
0.2500	27.33	27.50	27.64	28.01	28.16	27.60	24.13	20.48
0.3000	25.37	25.26	25.24	25.15	24.39	23.04	19.63	17.09
0.3500	23.31	22.95	22.71	22.20	21.01	19.53	16.76	15.03
0.4000	21.37	20.87	20.42	19.72	18.39	17.04	14.91	13.69
0.4500	19.67	19.07	18.55	17.71	16.39	15.21	13.61	12.74
0.5000	18.30	17.57	17.02	16.11	14.87	13.90	12.66	11.99
0.5500	17.21	16.42	15.85	14.95	13.75	12.93	11.91	11.40
0.6000	16.44	15.63	15.01	14.08	12.93	12.18	11.32	10.89
0.6500	15.96	14.15	14.52	13.55	12.37	11.62	10.83	10.47
0.7000	15.76	14.97	14.41	13.32	12.03	11.24	10.44	10.10
0.7500	15.80	15.11	14.54	13.44	11.98	11.04	10.14	9.80
0.8000	16.02	15.47	15.00	13.93	12.22	11.03	9.94	9.54
0.8500	16.30	15.96	15.67	14.79	12.92	11.32	9.85	9.38
0.9000	16.58	16.47	16.33	15.88	14.16	12.06	9.93	9.28
0.9500	16.79	16.85	16.94	16.96	15.95	13.55	10.29	9.31
1.0000	16.92	17.11	17.32	17.78	17.91	16.09	11.15	9.54
1.0500	16.99	17.23	17.50	18.18	19.24	19.18	13.11	10.15
1.1000	17.02	17.27	17.56	18.26	19.64	21.42	17.42	11.59
1.1500	17.04	17.27	17.56	18.17	19.52	21.71	23.80	15.22
1.2000	17.08	17.29	17.54	18.06	19.15	20.99	27.06	23.86
1.2500	17.15	17.31	17.54	17.98	18.82	20.16	24.96	33.12
1.3000	17.27	17.40	17.58	17.94	18.59	19.55	22.49	29.04
1.3500	17.41	17.54	17.69	17.97	18.49	19.16	21.02	24.33
1.4000	17.62	17.71	17.84	18.01	18.47	18.98	20.18	21.99
1.4500	17.85	17.91	18.04	18.20	18.58	18.92	19.77	20.82
1.5000	18.15	18.16	18.22	18.40	18.72	18.99	19.57	20.28
1.5500	18.44	18.46	18.68	18.69	18.95	19.13	19.57	20.05
1.6000	18.79	18.84	19.04	18.99	19.20	19.35	19.66	20.04
$-E(\sigma=0)/V$	0.4000	0.3927	0.3841	0.3742	0.3568	0.3385	0.3075	0.2856

Differential capacity on mercury electrode in 0.25 M NaF - adsorption of succinonitrile. Given are the capacity in μF cm^{-2}, potential in V vs SCE.

Reference: A. de Battisti, V. Faggiano and S. Trasatti.
J. Electroanal. Chem. 73 (1976) 327.

C/μF cm^{-2} T = 25°C

c/mol l^{-1}	0	0.015	0.020	0.024	0.030	0.035	0.040	0.050
-E/V								
0.0000	28.49	28.92	29.79	28.96	29.32	29.54	29.19	29.30
0.0500	27.83	28.21	28.18	28.36	28.70	28.58	28.61	28.79
0.1000	27.32	27.66	27.78	27.90	28.17	28.02	28.20	28.40
0.1500	26.88	27.24	27.35	27.41	27.74	27.53	27.23	27.99
0.2000	26.43	26.76	26.82	26.85	27.19	26.98	27.13	27.38
0.2500	25.91	26.15	26.21	26.16	24.46	26.16	26.29	26.43
0.3000	25.27	25.30	25.34	25.30	25.47	25.19	25.21	25.23
0.3500	24.53	24.41	24.38	24.29	24.33	24.35	23.99	23.88
0.4000	23.74	23.45	23.36	23.28	23.19	22.95	22.79	22.57
0.4500	22.94	22.56	22.39	22.31	22.15	22.00	21.67	21.39
0.5000	22.21	21.70	21.50	21.40	21.26	21.06	20.71	20.38
0.5500	21.45	21.07	20.67	20.60	20.40	20.24	19.89	19.54
0.6000	20.69	20.13	19.86	19.81	19.64	19.42	19.16	18.79
0.6500	19.76	19.41	19.16	19.03	18.92	18.73	18.47	18.17
0.7000	18.91	18.66	18.46	18.36	18.27	18.12	17.98	17.64
0.7500	18.12	17.94	17.80	17.75	17.70	17.59	17.41	17.17
0.8000	17.50	17.36	17.25	17.22	17.19	17.11	16.96	16.82
0.8500	16.89	16.92	16.78	16.76	16.76	16.73	16.62	16.50
0.9000	16.48	16.42	16.41	16.40	16.45	16.42	16.32	16.25
0.9500	16.09	16.10	16.13	16.12	16.19	16.17	16.11	16.07
1.0000	15.88	15.92	15.93	15.94	16.02	15.99	15.95	15.94
1.0500	15.74	15.78	15.82	15.86	15.93	15.90	15.85	15.89
1.1000	15.70	15.74	15.79	15.79	15.89	15.88	15.81	15.88
1.1500	15.75	15.78	15.84	15.84	15.96	15.94	15.85	15.94
1.2000	15.87	15.89	15.94	15.94	16.07	16.05	15.98	16.06
1.2500	16.03	16.06	16.12	16.10	16.23	16.22	16.14	16.24
1.3000	16.26	16.34	16.35	16.30	16.44	16.44	16.35	16.44
1.3500	16.51	16.58	16.61	16.55	16.71	16.69	16.61	16.69
1.4000	16.81	16.82	16.93	16.88	17.00	16.99	16.91	16.98
1.4500	17.15	-	17.30	17.22	17.32	17.30	17.24	17.32
1.5000	17.52	17.55	17.64	17.58	17.70	-	17.63	17.69
1.5500	17.97	-	18.05	17.98	18.12	18.11	18.03	18.12
1.6000	18.44	18.41	18.48	18.43	18.56	18.55	18.47	18.35
E(σ=0)/V	-0.4320	-0.4274	-0.4270	-0.4250	-0.4240	-0.4221	-0.4209	-0.4169

Capacity in 0.25 M NaF - adsorption of succinonitrile (continued)

	$C/\mu F\ cm^{-2}$						
$c/mol\ l^{-1}$	0.063	0.075	0.090	0.200	0.225	0.450	0.500
$-E/V$							
0.0000	29.70	29.78	30.16	32.60	32.99	45.25	50.56
0.0500	29.23	29.43	29.85	33.26	33.94	47.01	46.27
0.1000	28.88	29.18	29.63	33.70	34.56	34.94	30.47
0.1500	28.50	28.82	29.24	32.70	32.83	25.25	23.06
0.2000	27.81	28.05	28.35	29.69	28.91	20.33	19.00
0.2500	26.76	26.84	26.92	25.79	24.64	17.79	16.98
0.3000	25.39	25.24	25.11	22.39	21.24	16.43	15.94
0.3500	23.85	23.58	23.24	19.82	18.86	15.43	15.05
0.4000	22.40	22.01	21.54	18.03	17.26	14.76	14.56
0.4500	21.01	20.64	20.21	16.81	16.21	14.34	14.04
0.5000	20.01	19.52	19.05	15.89	15.48	14.05	13.78
0.5500	19.31	18.65	18.19	15.45	14.99	13.89	13.62
0.6000	18.41	17.92	17.53	15.08	14.68	13.71	13.55
0.6500	17.82	17.41	17.02	14.82	14.45	13.61	13.52
0.7000	17.36	16.98	16.65	14.73	14.39	13.60	13.52
0.7500	16.98	16.68	16.40	14.68	14.32	13.58	13.50
0.8000	16.67	16.45	16.23	14.66	14.32	13.59	13.50
0.8500	16.42	16.26	16.11	14.72	14.36	13.59	13.50
0.9000	16.23	16.12	16.03	14.81	14.44	13.59	13.51
0.9500	16.08	16.00	16.00	14.95	14.56	13.63	13.53
1.0000	15.98	15.96	15.98	15.12	14.74	13.69	13.57
1.0500	15.93	15.93	15.99	15.34	14.96	13.74	13.62
1.1000	15.95	15.95	16.05	15.61	15.24	13.83	13.69
1.1500	16.01	16.03	16.11	15.94	15.58	13.96	13.77
1.2000	16.13	16.14	16.27	16.28	16.01	14.12	13.91
1.2500	16.28	16.32	16.43	16.66	16.44	14.39	14.12
1.3000	16.51	16.53	16.64	17.05	16.91	14.74	14.45
1.3500	16.78	16.80	16.90	17.45	17.37	15.38	14.92
1.4000	17.06	17.08	17.19	17.82	17.82	16.31	15.69
1.4500	17.39	17.42	17.53	18.21	18.26	17.66	16.88
1.5000	17.78	17.78	17.92	18.58	18.64	19.28	18.54
1.5500	18.16	18.20	18.32	18.95	19.03	20.83	20.53
1.6000	18.26	18.65	18.76	19.37	19.42	22.04	22.27
$E(\mu=0)/V$	-0.4130	-0.4033	-0.4033	-0.3652	-0.3587	-0.3220	-0.3187

Differential capacity (μF cm^{-2}) of Hg/x mol l^{-1} DL-norvaline +) in 0.5 mol l^{-1} NaF, given is potential in V. T = 25°C
Reference electrode: saturated calomel electrode.

Reference: T. Kakiuchi and M. Senda. J. Electroanal. Chem. 88 (1978) 219.

	C/μF cm^{-2}							
x	0.0	0.05	0.10	0.20	0.30	0.40	0.50	0.6
-E/V								
0.010	28.6	30.5	30.2	30.9	32.0	32.1	33.3	34.4
0.100	27.8	28.0	27.9	28.1	28.9	28.9	29.2	30.0
0.200	27.3	26.9	26.9	26.9	27.4	27.4	27.5	28.1
0.300	26.6	26.1	26.1	26.0	25.8	25.6	25.3	25.3
0.400	25.3	24.7	24.6	23.8	23.4	22.7	21.8	21.2
0.500	23.3	22.8	22.3	21.4	20.5	19.6	18.7	17.8
0.600	21.1	20.7	20.2	19.3	18.3	17.8	16.8	16.1
0.700	19.0	18.8	18.5	17.9	17.2	16.9	16.3	15.8
0.800	17.5	17.4	17.3	17.1	16.8	16.7	16.4	16.2
0.900	16.4	16.6	16.5	16.6	16.6	16.7	16.7	16.8
1.000	15.8	16.2	16.1	16.4	16.6	16.6	16.9	17.0
1.100	15.8		16.1	16.4		16.7	17.0	17.2
1.200	16.0		16.3	16.5		16.8	17.2	17.4
1.300	16.5		16.6	16.9		17.0	17.4	17.5
1.400	17.0		17.2	17.4		17.5	17.7	18.0
1.500	17.7		17.8	18.1		18.2	18.4	18.5
1.600	18.7		18.9	18.9		19.0	19.3	19.3
1.700	20.1		19.7	19.8		19.8	20.1	20.9
1.800	21.3		21.2	21.5		21.6	21.6	21.9

+) Norvaline is 2 amino pentanoic acid.

Differential capacity (μF cm^{-2}) of Hg/x mol l^{-1} DL-norleucine $^{+)}$ in 0.5 mol l^{-1} NaF, given is potential in V. T = 25°C.
Reference electrode: saturated calomel electrode.

Reference: T. Kakiuchi and M. Senda. J. Electroanal. Chem. 88 (1978) 219.

	C/μF cm^{-2}					
x	0.00	0.01	0.02	0.03	0.04	0.06
-E/V						
0.010	28.4	29.2	29.4	29.5	29.9	30.6
0.100	27.2	27.6	27.5	27.5	27.8	28.0
0.200	26.6	26.5	26.5	26.1	26.4	26.5
0.300	25.9	25.7	25.3	25.1	25.2	24.8
0.400	24.7	24.2	23.9	23.5	23.4	22.8
0.500	22.8	22.2	21.8	21.4	21.2	20.3
0.600	20.6	20.2	20.0	19.6	19.3	18.5
0.700	18.5	18.3	18.2	18.0	18.0	17.6
0.800	17.0	16.9	17.0	17.0	17.2	17.1
0.900	16.0	16.1	16.4	16.4	16.7	16.8
1.000	15.6	15.7	15.9	16.1	16.3	16.5
1.100	15.6	15.7	15.8	16.0	16.2	16.5
1.200	15.7	15.8	16.0	16.1	16.3	16.6
1.300	16.3	16.3	16.4	16.5	16.7	16.8
1.400	16.8	16.8	16.9	17.0	17.2	17.3
1.500	17.4	17.6	17.6	17.6	17.8	17.9
1.600	18.4	18.3	18.4	18.5	18.5	18.6

+) Norleucine is 2 amino hexanoic acid.

Differential capacity (μF cm^{-2}) of Hg/x mol l^{-1} DL-alanine in 0.5 mol l^{-1} NaF given is the potential in V. T = 25°C
Reference electrode: saturated calomel electrode.

Reference: T. Kakiuchi and M. Senda. J. Electroanal. Chem. 88 (1978) 219.

	C/μF cm^{-2}								
x	0.00	0.10	0.20	0.30	0.40	0.60	0.80	1.00	1.20
-E/V									
0.010	28.4	29.9	30.7	31.3	32.0	32.5	32.9	33.7	34.1
0.100	27.2	27.9	28.2	28.4	28.6	28.8	29.1	29.2	29.5
0.200	26.6	27.0	27.2	27.2	27.3	27.4	27.3	27.3	27.7
0.300	25.9	26.4	26.6	26.6	26.5	26.7	26.5	26.3	26.5
0.400	24.7	25.1	25.3	25.4	25.4	25.3	25.3	25.0	25.1
0.500	22.8	23.1	23.4	23.4	23.5	23.4	23.3	23.0	23.1
0.600	20.6	20.9	21.0	21.1	21.1	21.0	21.0	20.8	21.0
0.700	18.5	18.9	18.9	18.9	18.9	19.0	19.0	18.9	19.0
0.800	17.0	17.2	17.3	17.3	17.3	17.4	17.5	17.5	17.5
0.900	16.0	16.3	16.3	16.3	16.4	16.5	16.5	16.6	16.7
1.000	15.6	15.8	15.8	15.8	15.9	15.9	16.0	16.2	16.2
1.100	15.6	15.7	15.7	15.8	15.8	15.9	15.9	16.1	16.2
1.200	15.7	15.8	15.9	15.9	16.1	16.2	16.2	16.3	16.4
1.300	16.3	16.3	16.3	16.4	16.5	16.6	16.6	16.8	16.8
1.400	16.8	16.8	16.9	17.1	17.1	17.1	17.2	17.3	17.4
1.500	17.4	17.6	17.7	17.8	17.8	17.8	17.9	18.2	18.2
1.600	18.4	18.4	18.8	19.1	19.1	18.8	18.9		

Differential capacity on mercury in 0.9 M NaF at 25°C. Given are the capacity in μF cm^{-2}, potential in V vs NCE (normal calomel electrode). Adsorption of ortho- cresol.

Reference: S. Sathyanarayana and G. Manohar. J. Electroanal. Chem. 51 (1974) 151.

	C/μF cm^{-2}							
c/m mol l^{-1}	0	1.55	3.1	6.2	15.5	30.9	76.9	149.5
E/V								
0.012							27.77	
0.050		31.77	31.85	29.38		21.87	26.85	24.85
0.100	29.47	31.88	29.66	25.68	20.65	19.45	24.70	23.90
0.150	28.94	29.86	26.12	22.33	18.40	17.50	22.13	23.28
0.200	28.87	26.91	22.91	19.77	16.48	15.77	19.11	22.79
0.250	28.81	23.89	20.43	17.74	14.84	14.20	16.51	22.00
0.300	28.50	21.58	18.40	16.01	13.40	12.80	14.08	19.51
0.350	28.20	19.83	16.84	14.62	12.22	11.60	12.23	16.38
0.400	27.58	18.50	15.55	13.44	11.22	10.57	10.60	13.17
0.450	26.76	17.70	14.70	12.58	10.38	9.76	9.43	10.47
0.500	25.61	17.25	13.96	11.85	9.79	9.11	8.61	8.65
0.550	24.37	17.30	13.92	11.48	9.32	8.61	8.05	7.68
0.600	23.02	17.63	14.09	11.31	9.06	8.28	7.77	6.99
0.650	21.67	18.45	14.79	11.54	9.03	8.13	7.55	6.94
0.700	20.46	19.62	16.11	12.25	9.18	8.09	7.49	6.91
0.750	19.38	21.08	18.25	13.60	9.64	8.25	7.54	6.94
0.800	18.47	22.47	21.19	16.01	10.54	8.63	7.65	7.13
0.850	17.75	23.10	24.36	19.92	12.13	9.25	7.94	7.33
0.900	17.15	22.96	26.82	25.60	14.84	10.35	8.34	7.64
0.950	16.68	21.86	27.11	31.56	19.84	12.26	9.08	8.12
1.000	16.44	20.43	25.42	32.07	29.32	15.85	10.21	8.68
1.025					35.75			
1.050	16.23	18.87	22.84	30.63	38.97	23.37	12.07	9.68
1.075					43.91			
1.100	16.14	17.80	20.45	26.11	42.03	39.02	15.81	11.11
1.150	16.14	17.07	18.72	22.32	32.99	52.24	24.96	13.94
1.175					29.15			

Differential capacity on mercury in 0.9 M NaF in the presence of o-cresol (continued)

	$C/\mu F\ cm^{-2}$							
$c/m\ mol\ l^{-1}$	0	1.55	3.1	6.2	15.5	30.9	76.9	149.5
E/V								
1.200	16.24	16.80	17.67	19.83	25.90	39.28	50.05	20.49
1.225							73.76	
1.250	16.40	16.70	17.16	18.38	21.23	27.48	78.21	41.11
1.275							60.82	55.98
1.300	16.63	16.62	16.97	17.71	19.06	21.84	42.80	85.92
1.325							32.96	76.60
1.350	16.85	16.87	17.00	17.50	18.01	19.24	28.24	26.75
1.400	17.12	17.16	17.16	17.49	17.65	18.00	21.35	21.64
1.450	17.42	17.45	17.51	17.72	17.58	17.58	19.50	19.57
1.500	17.80	17.76	17.79	17.97	17.78	17.56	18.88	18.84
1.550	18.25	18.11	18.17	18.28	18.13	17.73	18.74	18.59
1.600	18.66	18.59	18.59	18.65	18.50	18.00	18.92	18.87
1.650	19.11		18.99	19.15	18.92	18.46	19.33	19.22
1.700	19.65		19.52	19.63	19.45	19.10		
1.750	20.24							
1.800	21.17							

Differential capacity on mercury electrode in 0.9 M NaF at 25°C. Adsorption of meta-cresol. Given are the capacity in μF cm^{-2}, potential in V vs NCE (normal calomel electrode).

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

	C/μF cm^{-2}						
c/m mol l^{-1}	1.5	3.0	6.0	14.9	19.7	73.9	137.9
E/V							
0.10	24.2	25.7	22.0	21.3	21.3	8.2	19.3
0.15	22.0	23.4	20.1	20.1	19.7	8.7	18.6
0.20	21.4	21.4	18.6	18.7	18.2	9.6	18.2
0.25	19.1	19.9	17.4	17.5	16.9	11.2	17.7
0.30	18.0	18.7	16.3	16.3	15.7	14.8	17.5
0.35	17.1	17.7	15.4	15.3	14.6	23.9	17.4
0.40	16.7	17.0	14.6	14.3	13.6	52.7	17.5
0.45	16.4	16.2	13.9	13.5	12.7	85.2	17.8
0.50	16.3	15.8	13.5	12.8	11.9	41.0	17.5
0.55	16.5	15.7	13.2	12.2	11.1	26.0	15.8
0.60	17.0	15.9	13.1	11.6	10.4	21.4	12.8
0.65	17.9	16.7	13.3	11.2	9.8	19.6	10.0
0.70	19.2	18.1	14.0	11.0	9.5	19.0	8.4
0.75	20.8	20.3	15.3	11.3	9.3	18.9	7.8
0.80	22.7	23.1	17.7	11.9	9.4	19.0	7.7
0.85	23.9	26.2	21.4	13.5	9.8	19.3	7.8
0.90	24.1	28.4	26.6	16.5	10.8	19.6	8.0
0.95	23.2	28.6	31.7	22.2	12.7	20.1	8.3
1.00	21.5	26.7	32.9	32.2	16.9	20.5	8.7
1.05	19.6	23.6	30.1	42.8	26.1	21.1	9.4
1.10	18.2	20.9	25.3	41.4	45.2	20.5	10.8
1.15	17.4	19.0	21.8	32.4	57.2	18.7	13.5
1.20	16.8	17.8	19.5	26.1	40.5	15.9	20.9
1.25	16.7	17.3	18.2	21.5	28.5	13.3	47.0
1.30	16.7	17.1	17.5	19.3	22.7	11.3	116.3
1.35	16.9	17.1	17.3	18.4	20.0	9.9	44.0
1.40	17.2	17.4	17.4	18.0	18.8	9.0	25.9
1.45	17.4	17.4	17.6	17.9	18.4	8.5	21.2
1.50	17.8	17.9	17.9	18.2	18.4	8.2	19.6
1.55		18.4	18.4	18.5	18.6	8.1	19.2
1.60	18.6	18.8	18.8	18.9	19.0	8.1	19.2

Differential capacity on mercury electrode in 0.9 M NaF at 25°C. Adsorption of para-cresol. Given are the capacity in μF cm^{-2}, potential in V vs NCE (normal calomel electrode).

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

	C/μF cm^{-2}						
c/m mol l^{-1}	1.4	2.9	5.8	14.5	28.9	72.0	128.7
E/V							
0.10	26.7	23.6	21.8	20.5	20.4	20.2	19.9
0.15	24.9	21.4	19.9	18.9	18.8	19.5	19.1
0.20	22.3	19.8	18.6	17.6	17.3	18.7	18.2
0.25	20.7	18.6	17.5	16.4	16.0	18.1	17.4
0.30	19.7	17.6	16.5	15.4	14.9	17.5	16.9
0.35	19.0	16.8	15.8	14.5	14.0	17.0	16.4
0.40	18.4	16.2	15.1	13.7	13.1	16.4	16.2
0.45	18.3	15.8	14.5	13.1	12.4	15.3	16.0
0.50	18.4	15.7	14.3	12.6	11.7	13.7	15.3
0.55	18.8	15.9	14.1	12.1	11.2	11.7	13.6
0.60	19.5	16.3	14.2	11.9	10.7	10.2	11.5
0.65	20.3	17.1	14.7	11.8	10.3	9.2	9.5
0.70	21.4	18.5	15.6	12.0	10.2	8.7	8.4
0.75	22.6	20.3	17.2	12.6	10.2	8.5	7.9
0.80	23.5	22.8	19.8	13.7	10.6	8.6	7.8
0.85	23.8	25.2	23.1	15.7	11.4	8.8	8.0
0.90	23.3	27.0	27.1	19.1	12.8	9.2	8.4
0.95	22.0	26.8	30.5	24.6	15.3	9.8	9.0
1.00	20.3	25.0	30.8	32.3	20.2	11.0	9.9
1.05	18.7	22.3	28.0	38.3	29.8	13.0	11.2
1.10	17.6	20.0	24.0	36.4	44.1	17.1	13.5
1.15	17.0	18.4	20.8	29.8	47.9	27.0	18.2
1.20	16.7	17.5	19.0	24.1	36.1	54.3	31.2
1.25	16.6	16.9	17.8	20.7	26.8	70.9	71.7
1.30	16.7	16.8	17.4	18.9	21.7	37.4	71.0
1.35	16.9	16.9	17.3	18.1	19.3	25.3	32.6
1.40	17.1	17.1	17.4	17.8	18.4	21.0	23.1
1.45	17.5	17.4	17.5	17.8	18.1	19.4	20.2
1.50	17.8	18.1	17.9	18.0	18.1	18.8	19.2
1.55	18.2		18.2	18.3	18.3	18.6	18.9
1.60	18.6	18.6	16.7	18.7	18.7	18.8	19.0

Capacity data for mercury in n-pentanoic acid in 0.1 M $HClO_4$ solution. Reference electrode : SCE (NaCl solution) ; c/c_o as defined below

T = 25°C

c = solute concentration

c_o its saturation concentration in the base electrolyte

Reference: K.G. Baikerikar, D.E. Broadhead and R.S. Hansen. J. Phys. Chem. 80 (1976) 370.

	$C/\mu F\ cm^{-2}$								
c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.1667	0.3333
-E/V									
1.100	15.77	16.30	16.92	17.47	18.83	20.18	23.59	28.94	38.28
1.075							25.42	32.67	26.84
1.050	15.98	16.64	17.42	18.11	20.01	22.14	27.97	36.95	19.04
1.035								38.32	
1.030								38.45	
1.025							31.00	38.31	14.29
1.020								37.95	
1.015								37.32	
1.012							32.54		
1.010								36.47	
1.000	16.28	17.09	18.07	19.06	21.82	25.32	33.94	33.90	11.28
0.996							33.76		
0.990								31.36	
0.987							33.97		
0.975					22.81	26.94	33.45	26.90	9.46
0.950	16.62	17.57	18.77	20.05	23.90	28.34	29.57	20.32	8.19
0.937						28.61			
0.925					24.87	28.46	23.94	15.60	
0.912					25.17	27.76			

Capacity date for n-pentanoic acid in 0.1 M $HClO_4$ solution (continued)

	$C/\mu F\ cm^{-2}$								
c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.1667	0.3333
$-E/V$									
0.900	17.00	18.06	19.46	21.09	25.31	26.52	18.50	12.36	6.60
0.887				21.28	25.12	24.99			
0.875			19.73	21.41	24.72	23.22	14.86		
0.862				21.52					
0.850	17.41	18.43	19.90	21.52	22.99	19.29	12.09	8.91	5.77
0.837			19.93	21.37					
0.825			19.91	21.22	20.40	15.88	10.19		
0.812			19.86						
0.800	17.80	18.59	19.75	20.45	17.53	13.09	8.78	7.15	5.26
0.787			19.54						
0.775			19.31	19.15	15.00				
0.750	18.13	18.41	18.60	17.55	12.89	9.81	7.26	6.24	4.97
0.725				15.86	11.23				
0.700	18.45	17.83	16.55	14.24	9.94	8.00	6.37	5.72	4.80
0.675				12.87					
0.650	18.85	17.01	14.45	11.76	8.42	7.11	5.96	5.47	4.73
0.600	19.53	16.43	12.93	10.35	7.69	6.69	5.76	5.37	4.73
0.550	20.79	16.50	12.45	9.89	7.52	6.66	5.80	5.43	4.79
0.500	22.73	17.83	13.21	10.42	7.89	6.94	6.01	5.63	4.93
0.475	24.12	19.65	14.17	11.08					
0.450	25.12	20.57	15.63	12.17	8.91	7.69	6.52	6.02	5.17
0.439				12.80					
0.425		22.37	17.62						

Capacity data for n-pentanoic acid in 0.1 M $HClO_4$ (continued)

c/c_0				$C/\mu F\ cm^{-2}$					
c/c_0	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.1667	0.3333
-E/V									
0.417				14.39					
0.411					10.54				
0.400	27.15	24.33	20.20	16.01	11.20	9.23	7.53	6.72	5.53
0.392						9.58			
0.384				17.72					
0.375		26.12	23.09		13.12	10.50			
0.367				19.95					
0.350	28.23	27.64	26.17	22.49	15.84	12.34	9.12	7.99	6.13
0.344								8.21	
0.337				24.60	17.61				
0.325		28.68	28.80	26.53	19.51	14.96			
0.312				28.59	21.89				
0.300	28.07	29.12	30.58	30.38	24.42	18.77	12.85	10.47	7.17
0.294									7.25
0.287			31.11	31.88	27.17				
0.284							14.81		
0.275		29.08		32.96	29.92	23.74			
0.267							17.36		
0.262			31.38	33.67	32.81				
0.250	27.14	28.68	31.22	33.93	35.26	30.09	20.56	15.63	9.10
0.247				33.93	37.10				
0.239							23.56		
0.237				33.85					
0.227					38.50				

Capacity data for n-pentanoic acid in 0.1 M $HClO_4$ (continued)

$C/\mu F\,cm^{-2}$

c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	0.1667	0.3333
-E/V									
0.225		28.07	30.53	33.49		36.87	26.66		10.71
0.215					39.26				
0.212				32.88			30.53		
0.210					39.32				
0.205					39.29				
0.200	25.91	27.41	29.55	32.23	39.22	41.99	34.11	26.08	13.10
0.187				31.48		43.18	38.70		
0.185						43.22			
0.184									15.09
0.180						43.33			
0.175		26.67		30.78	36.83	43.26	42.44	33.92	
0.167						43.03			17.82
0.162							45.88		
0.150	24.86	26.04	27.66	29.42	34.80	41.07	47.95	42.86	21.23
0.140							48.65		
0.137									24.37
0.135							48.70		
0.130							48.49		
0.125				28.27	32.44	37.50	48.11	50.16	27.68
0.112							46.11		31.82
0.110								52.05	
0.105								52.14	
0.100	24.08	25.00	26.19	27.28	30.54	34.22	44.26	52.01	36.10

Capacity data for n-pentanoic acid in 0.1 M $HClO_4$ (continued)

$C/\mu F\,cm^{-2}$

c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976	0.09090	0.1304	1.667	0.3333
-E/V									
0.087							41.52		40.95
0.075					29.05	31.77	39.21	48.05	45.66
0.062									50.85
0.050	23.65	24.35	25.23	25.93	27.94	29.96	35.29	42.32	54.94
0.037									58.43
0.025						28.71	32.53	37.41	59.90
0.020									60.04
0.015									59.90
0.012									59.49
0.000	23.59	24.12	24.78	25.25	26.64	27.80	30.66	34.01	57.42
-0.012									54.33
-0.025								31.79	50.39
-0.037									46.79
-0.050	23.85	24.23	24.74	25.06	26.04	26.79	28.50	30.29	43.46
-0.062									40.86
-0.075									38.45
-0.087									36.74
-0.100	24.39	24.65	25.06	25.23	26.01	26.49	27.56	28.65	35.27
-0.125									33.17
-0.150	25.75	25.32	25.64	25.72	26.30	26.60	27.32	28.02	31.86
-0.175									30.95
-0.200	26.08	26.19	26.45	26.45	26.91	27.08	27.57	28.01	30.44

Differential capacity on mercury electrode in 1 M $NaClO_4$ at 25°C. Adsorption of acetylacetone. Given are the capacity in $\mu F\cdot cm^{-2}$ potential in mV vs CE.

Reference: J. Dojlido, Z. Galus and L.J. Jeftic. J. Electroanal. Chem. 62 (1975) 433.

	pH = 2.8							pH = 11.5							
	C/μF cm^{-2}							C/μF cm^{-2}							
c/mol l^{-1}	0.01	0.05	0.1	0.2	0.4	0.8	1.0	0	0.001	0.01	0.4	0.5	0.7	0.8	1.0
-E/mV															
100	23.0	25.2	27.2	28.6	28.7	29.7	29.9	21.9							
200	23.3	25.4	26.0	28.0	26.6	25.2	24.9	22.5							
300	24.1	25.5	25.9	26.0	22.8	19.8	19.4	23.6	24.8	26.7					
400	25.5	24.9	24.0	22.7	17.5	16.2	15.8	25.8	26.3	28.0	27.7				34.2
500	26.3	23.0	21.0	18.9	15.6	14.0	13.7	28.0	27.8	29.1	28.5				24.5
600	26.0	20.3	17.7	16.3	13.6	12.6	12.4	28.3	28.0	29.1	23.3		19.5	10.8	12.5
700	23.4	17.9	15.7	14.1	12.4	11.7	11.6	25.9	25.3	26.6	15.6	10.3	14.8	9.5	9.8
800	20.7	16.8	14.3	12.5	10.4	10.8	10.7	22.0	21.1	23.1	14.7	10.8	13.2	9.6	9.4
900	17.9	17.1	14.1	11.7	9.9	10.1	9.9	18.5	18.4	19.4	17.2	11.8	14.5	11.7	10.1
1000	16.9	17.8	15.5	12.1	9.7	9.7	9.5	16.6	16.5	17.7	20.2	14.8	18.1	15.6	11.7
1100	16.2	18.2	18.6	13.9	9.9	9.7	9.5	16.0	15.8	16.7	21.6	20.5	22.4	22.4	15.7
1200	16.0	18.0	20.7	18.6	12.3	10.7	10.1	15.7	15.7	16.3	20.8	25.1	25.0	25.8	24.6
1300	16.2	17.8	20.4	24.0	19.8	13.5	12.1	16.0	15.9	17.2	19.4	24.3	24.1	24.4	29.0
1400	16.6		19.9		31.7	23.5	18.9	16.4	16.4	17.4	18.2	22.9	23.0	21.9	24.3
1500		18.8	20.1	23.9		38.7									21.3
1600								17.1	17.1	17.9	18.0	20.1	21.1	19.3	

Dependence of differential capacitance of mercury on potential against calomel electrode with 1M NaCl adsorption of 0.5 M pyridine in $NaClO_4$ + $HClO_4$. Given are the capacity in $\mu F\ cm^{-2}$, potential in mV.

T = 25°C, 500 Hz

Reference: Z. Galus, J. Dojlido and G. Chojnacko-Kalinowska. Electrochim. Acta 17 (1972) 265.

	$C/\mu F\ cm^{-2}$					
	0.6 M $HClO_4$ 0.4 M $NaClO_4$ pH = 1.8	0.52 M $HClO_4$ 0.48 M $NaClO_4$ pH = 1.95	0.512 M $HClO_4$ 0.488 M $NaClO_4$ pH = 2.9	0.5 M $HClO_4$ 0.5 M $NaClO_4$ pH = 3.56	0.4 M $HClO_4$ 0.6 M $NaClO_4$ pH = 4.42	0.35 M $HClO_4$ 0.65 M $NaClO_4$ pH = 4.70
E/mV						
250	-	22.7	-	22.6	22.9	23.2
300	21.8	21.9	22.3	21.7	22.0	22.5
350	21.4	-	21.6	21.0	21.6	22.0
400	20.9	20.8	21.2	20.6	21.1	21.6
450	20.6	20.5	20.8	-	20.5	-
500	20.4	20.3	20.5	19.9	20.4	21.0
550	20.2	20.2	-	19.7	20.1	20.7
600	20.2	20.1	20.3	19.5	19.7	20.4
650	20.2	20.1	20.3	19.5	19.3	19.9
700	20.1	20.1	20.2	19.5	18.9	19.1
750	20.0	19.9	20.2	19.4	18.3	18.1
800	19.9	19.8	20.0	19.3	17.6	16.9
850	19.6	19.5	19.7	19.1	16.7	15.6
900	19.3	19.5	19.4	18.8	15.9	14.5
950	18.6	18.8	19.3	18.6	15.2	13.7
1000	18.0	18.0	18.6	18.4	14.8	13.2
1050	17.4	17.4	17.7	17.8	14.8	13.1
1100	16.8	16.9		17.2	14.8	13.1
1150	16.5	16.5	16.8	16.9	15.0	13.4
1200	16.2	16.3	16.5	16.7	15.6	14.1
1250	16.2	16.3	16.5	16.6	16.6	15.3
1300	16.4	16.7	16.8	16.9	18.7	-

Capacity in $NaClO_4$ + $HClO_4$ containing pyridine (continued)

$C/\mu F\ cm^{-2}$

E/mV	0.3 M $HClO_4$ 0.7 M $NaClO_4$ pH = 4.91	0.2 M $HClO_4$ 0.8 M $NaClO_4$ pH = 5.06	0.1 M $HClO_4$ 0.9 M $NaClO_4$ pH = 5.46	0.05 M $HClO_4$ 0.9 M $NaClO_4$ pH = 5.85	0.05 M $HClO_4$ 0.95 M $NaClO_4$ pH = 6.81	1 M $NaClO_4$ pH = 8.53
250	23.6	24.3	25.8	26.9	28.0	27.7
300	22.9	23.8	25.4	26.4	-	27.4
350	22.5	23.5	25.5	26.7	28.1	27.9
400	22.2	23.5	26.1	27.4	28.8	28.7
450	22.0	23.7	26.6	27.9	29.1	29.0
500	21.8	23.8	26.4	27.7	27.5	27.4
550	21.7	23.6	25.0	24.1	-	-
600	21.3	22.6	21.5	20.0	19.1	18.6
650	20.5	20.4	17.8	16.0	15.1	15.0
700	19.2	17.6	14.4	13.3	11.5	12.2
750	16.7	15.4	11.9	10.9	10.4	10.2
800	15.6	12.8	10.3	10.2	8.9	8.7
850	14.0	11.4	9.3	8.5	7.9	7.8
900	12.8	10.6	8.7	7.9	7.3	7.1
950	12.1	10.2	8.4	7.6	6.8	6.6
1000	11.7	10.0	8.4	7.4	6.5	6.2
1050	11.7	10.2	8.5	7.5	6.3	5.9
1100	12.0	10.7	9.1	7.8	6.3	5.8
1150	12.3	11.1	9.6	8.3	6.4	5.6
1200	12.8	11.6	10.1	8.9	6.7	5.5
1250	13.8	12.4	10.8	9.5	7.0	5.4
1300	15.2	13.9	11.9	10.3	7.5	5.4
1350				11.0	8.3	5.4
1400						5.6
1500						6.0

Differential capacity curves of the mercury electrode on the potential in 0.2 M $NaClO_4$ at pH = 11 , containing the indicated molar concentration of benzoic acid. temp. 25°C, frequency - 500 Hz, reference electrode: calomel electrode with 1 mol l^{-1} NaCl.

Reference: J. Dojlido, M. Dmowska-Stanczak and Z. Galus.
J. Electroanal. Chem. 94 (1978) 107.

	$C/\mu F\ cm^{-2}$					
$c/mol\ l^{-1}$	0.005	0.01	0.015	0.02	0.03	0.04
-E/mV						
200	21.15	20.40	20.50	19.70	10.10	20.90
300	22.90	22.00	21.95	21.15	21.15	21.50
400	26.20	25.25	24.90	24.15	23.85	23.25
500	28.90	28.50	26.80	27.45	26.70	25.75
600	28.75	29.20	27.20	28.35	26.80	26.80
700	25.30	26.80	25.95	26.90	26.40	25.90
800	21.00	22.80	22.90	24.00	24.30	24.45
900	17.90	18.90	19.05	20.30	21.15	21.80
1000	16.25	16.70	16.90	17.45	18.10	19.00
1100	15.60	15.75	15.75	16.00	16.40	16.85
1200	15.20	15.50	15.35	15.65	15.80	16.05
1300	15.35	15.75	15.50	15.75	15.85	15.95
1400	16.15	16.20	16.00	16.15	16.20	16.25

Differential capacity of the mercury electrode on the potential in 0.2 M $NaClO_4$ containing the indicated molar concentration of benzoic acid.

temp. 25°C, frequency - 500 Hz, reference electrode: calomel with 1 M NaCl.

Reference: J. Dojlido, M. Dmowska-Stanczak and Z. Galus. J. Electroanal. Chem. 94 (1978) 107.

	$C/\mu F\ cm^{-2}$											
$c/mol\ l^{-1}$	0	0.001	0.002	0.003	0.005	0.006	0.008	0.01	0.012	0.015	0.02	saturated
-E/mV												
0	24.20	24.40	25.75	25.80	25.35	25.70	25.85	25.30	24.80	24.75	24.70	24.65
100	22.85	22.50	22.85	22.65	21.80	21.75	22.00	22.40	21.65	22.00	23.30	21.45
200	23.50	20.90	20.75	20.20	19.45	19.70	19.75	20.15	19.65	20.15	21.35	20.25
300	25.30	20.30	19.65	19.05	17.95	18.30	18.25	18.75	18.40	19.00	19.75	19.75
400	28.10	20.45	19.20	18.35	17.10	17.20	17.10	17.45	17.60	17.60	18.25	17.50
500	29.15	21.00	18.85	17.65	16.05	15.80	15.15	15.60	14.85	15.05	14.60	12.65
600	26.25	20.65	18.45	17.95	14.85	14.40	13.45	12.95	12.00	11.55	10.80	9.25
700	21.85	19.70	18.35	17.00	14.45	13.85	12.35	11.50	10.20	9.60	8.95	7.75
800	18.90	19.55	19.40	18.70	16.10	15.55	13.60	12.10	10.05	9.00	8.20	7.30
900	17.05	19.50	20.85	21.40	20.75	20.60	18.70	16.50	12.70	10.10	8.30	7.15
1000	16.10	18.40	20.25	21.90	24.55	24.90	25.85	26.60	23.35	17.65	10.50	7.45
1100	15.65	16.80	18.10	19.50	22.75	23.30	25.60	28.70	30.55	35.55	31.00	12.45
1200	15.65	15.95	16.65	17.30	18.55	19.30	20.80	23.05	23.95	26.50	34.20	50.40
1300	16.00	15.90	16.15	16.30	16.95	17.15	17.75	19.00	19.00	20.15	23.05	23.35
1400	16.35	16.20	16.35	16.35	16.60	16.70	16.90	17.75	17.35	17.85	19.45	18.35

Differential capacity of the mercury electrode on the potential in solution of 0.02 M benzoic acid in 0.2 M $NaClO_4$ at indicated pH. temp. 25°C, frequency - 500 Hz, reference electrode: calomel electrode with 1 M NaCl.

Reference: J. Dojlido, M. Dmowska-Syanczak and Z. Galus. J. Electroanal. Chem. 94 (1978) 107.

	$C/\mu F\,cm^{-2}$										
pH	1.2	2.81	3.29	4.00	4.33	4.45	4.91	5.41	5.85	10.73	11.6
-E/mV											
0	23.10	24.75	24.70	27.60	28.65	28.35	26.30	24.75	23.70		
100	20.35	21.60	23.30	22.70	23.75	23.80	22.70	22.05	21.15	21.05	
200	18.45	19.90	21.35	23.15		21.00	21.35	20.55	20.55	20.25	21.65
300	18.00	18.95	19.75	20.05	18.55	18.85	19.30	20.15	21.45	21.60	22.45
400	16.75	17.70	18.25	17.05	17.30	17.20	18.35	20.25	23.45	24.45	25.35
500	12.70	13.80	14.60	14.70	15.80	16.30	17.65	21.05	25.50	27.65	28.45
600	9.50	10.15	10.80	12.90	14.30	15.50	17.55	22.00	26.15	28.25	29.15
700	7.75	8.40	8.95	11.10	13.55	15.65	19.05	23.15	25.45	27.10	27.45
800	7.40	7.75	8.20	10.75	14.65	17.60	21.55	23.45	23.25	23.65	24.30
900	7.50	7.80	8.30	13.55	19.10	22.20	23.20	21.85	20.25	20.50	20.60
1000	8.55	9.65	10.50	23.25	25.75	25.05	21.30	19.30	17.50	17.65	17.70
1100	14.00	27.75	31.00	30.20	25.35	22.75	19.20	17.10	15.95	16.15	16.35
1200	35.10	32.20	34.20	23.45	20.55	18.90	17.10	16.00	15.50	15.70	16.00
1300	22.70	21.65	23.05	18.75	17.65	17.10	16.30	15.90	15.60	15.85	16.05
1400	18.70	18.00	19.45	17.10	16.80	16.65	16.35	16.20	16.00	16.25	16.30

Differential capacity on mercury electrode in 1 M NaOH at 25°C - given are the capacity in $\mu F\ cm^{-2}$, potential in V vs mercury oxide electrode. Adsorption of Ortho-cresol.

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 31 (1971) 301.

	$C/\mu F\ cm^{-2}$										
$c/m\ mol\ l^{-1}$	0	1.55	3.1	6.2	15.5	30.9	76.9	152.5	227.0	478.0	908.0
E/V											
0.036			65.64								
0.039											48.17
0.045				53.44		51.82					
0.049		46.49									
0.0502					44.79						
0.0543	77.54										
0.057							39.01				
0.058									38.75		
0.059								39.11			
0.100	48.42	36.81	37.97	36.46	34.84	36.48	33.90	34.12	33.83	34.70	44.28
0.150	35.97	35.82	35.88	35.31	34.54	34.19	33.70	33.45	33.27	33.89	35.22
0.200	30.47	35.44	35.81	35.32	34.76	34.14	33.43	33.04	32.81	32.96	26.00
0.250	28.77	35.10	35.42	34.81	34.17	33.39	32.53	31.78	31.47	30.76	20.48
0.300	27.29	33.82	34.40	33.78	32.86	31.95	30.79	29.81	29.17	27.57	16.63
0.350	25.82	32.59	32.85	32.26	31.18	30.01	28.74	27.55	26.74	24.41	15.73
0.400	24.46	30.90	31.42	30.95	29.69	28.46	26.96	25.57	24.75	22.11	15.02
0.450	22.98	29.24	30.06	30.09	28.76	27.41	25.72	24.29	23.39	21.00	14.32
0.500	21.67	27.46	28.65	29.33	28.47	26.99	25.28	23.79	22.87	20.85	13.87
0.550	20.47	25.46	27.08	28.44	28.65	27.18	25.20	23.61	22.10	20.21	13.66
0.600	19.27	23.42	25.01	27.02	28.59	27.92	25.89	24.08	22.90	20.68	13.75

Differential capacity on mercury in 1 M NaOH in the presence of o-cresol (continued)

	$C/\mu F\ cm^{-2}$										
$c/m\ mol\ l^{-1}$	0	1.55	3.1	6.2	15.5	30.9	76.9	152.5	227.0	478.0	908.0
E/V											
0.650	18.41	21.36	22.83	25.01	27.90	28.32	27.05	25.23	24.07	21.39	14.54
0.700	17.71	19.60	20.77	22.62	26.19	28.11	28.26	26.72	25.63	22.73	16.12
0.750	17.22	18.38	19.11	20.62	23.89	26.74	28.85	28.44	27.52	24.55	18.76
0.800	16.97	17.64	18.02	19.06	21.55	24.47	28.14		29.04	26.68	23.21
0.850	16.69	17.03	17.37	17.95	19.71	21.91	25.98	28.71	29.88	28.96	28.97
0.900	16.49	16.54	16.71	17.11	18.22	19.66	23.08	26.46	28.38	29.82	34.45
0.950	16.32	16.36	16.46	16.67	17.34	18.23	20.56	23.64	25.78	29.08	37.06
1.000	16.35	16.31	16.27	16.51	16.86	17.41	18.99	21.14	23.08	26.83	36.40
1.050	16.46	16.40	16.40	16.47	16.71	17.03	17.97	19.41	20.78	24.11	32.73
1.100	16.61	16.51	16.50	16.53	16.69	16.85	17.45	18.29	19.30	21.77	28.33
1.150	16.83		16.72	16.72	16.78	16.86	17.25	17.80	18.39	20.04	24.65
1.200	17.07	17.01	16.96	16.97	17.03	17.02	17.33	17.63	17.80	19.05	22.12
1.250	17.34	17.31	17.22	17.22	17.20	17.23	17.42	17.58	17.80	18.51	20.43
1.300	17.71	17.66	17.61	17.54	17.57	17.56	17.69	17.52	17.90	18.34	19.60
1.350	18.01	18.00	18.06	17.90	17.83	17.68	17.98	18.06		18.35	19.16
1.400	18.44	18.42	18.52	18.33	18.25	18.25	18.32	18.40		18.53	19.04
1.450	18.84	18.80	18.89	18.75	18.67	18.69	18.73			18.80	19.04
1.500	19.26	19.28	19.37	19.25	19.16	19.15	19.17	19.27		19.22	19.27
1.550	19.78			19.82	19.77		19.74	19.78		19.67	19.67
1.600	20.46	20.82		20.39	20.30		20.37	20.37			20.11
1.650	21.83				21.50		21.62				21.70
1.700	23.20				27.04		27.43	26.79			22.90

Differential capacity on mercury in 1 M NaOH at 25°C.
Adsorption of meta-cresol - given are the capacity in µF cm^{-2}, potential in V vs mercury oxide electrode.

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 31 (1971) 301.

	C/µF cm^{-2}							
c/m mol l^{-1}	1.5	3.0	6.0	14.9	29.7	73.9	146.7	218.0
E/V								
0.10	44.6	36.7	37.2	33.0	33.9	33.3	33.4	33.6
0.15	39.7	37.1	36.3	35.0	34.9	34.1	34.5	34.9
0.20	37.8	38.3	37.6	36.5	36.8	35.3	35.2	35.4
0.25	37.2	38.4	38.0	37.0	36.3	35.1	34.6	34.3
0.30	35.4	37.1	36.8	36.0	35.1	33.4	32.5	31.8
0.35	32.9	34.9	34.8	34.0	33.7	31.1	29.8	29.0
0.40	30.4	32.4	32.5	31.6	30.7	28.9	27.6	26.6
0.45	27.9	30.1	30.6	29.9	28.9	27.4	26.0	25.1
0.50	25.8	28.0	29.1	28.9	27.9	26.3	25.1	24.4
0.55	23.6	25.9	27.6	28.4	27.5	25.8	24.6	24.2
0.60	21.8	23.8	26.0	27.9	27.6	26.0	24.8	24.5
0.65	20.1	21.8	23.9	27.1	27.6	26.6	25.6	25.1
0.70	18.7	20.0	21.8	25.5	27.1	27.4	26.7	26.2
0.75	17.8	18.7	20.0	23.3	25.9	27.6	27.8	27.6
0.80	17.1	17.8	18.7	21.4	23.9	26.6	28.4	28.7
0.85	16.8	17.3	17.8	19.6	21.5	24.9	27.8	28.9
0.90	16.4	16.7	17.0	18.0	19.5	22.4	25.6	27.6
0.95	16.3	16.4	16.6	17.3	18.2	20.3	23.1	25.1
1.00	16.3	16.4	16.4	16.8	17.5	18.7	20.8	22.5
1.05	16.4	16.4	16.5	16.7	17.1	17.7	19.1	20.5
1.10	16.5	16.6	16.6	16.6	16.9	17.2	18.0	19.0
1.15	16.7	16.9	16.7	16.8	17.0	17.1	17.6	18.2
1.20	17.0	17.0	17.0	17.0	17.1	17.2	17.5	18.0
1.25	17.3	17.3	17.5	17.3	17.2	17.4	17.5	17.9
1.30	17.6	17.7	17.6	17.7	17.6	17.6	17.7	18.0
1.35	18.0	18.0	18.0	18.1	17.9	18.0	18.0	18.3
1.40	18.4	18.4	18.3	18.4	18.4	18.3	18.3	18.7
1.45	18.9	18.9	18.8	18.8	18.8	18.7	18.7	19.1
1.50	19.3	19.3	19.2	19.3	19.2	19.2	19.2	19.5

Differential capacity on mercury electrode in 1 M NaOH at 25°C. Adsorption of para-cresol - given are the capacity in μF cm^{-2}, potential in V vs mercury oxide electrode.

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

	C/μF cm^{-2}							
c/m mol l^{-1}	1.4	2.9	5.8	14.5	28.9	72.0	142.9	212.7
E/V								
0.10	44.9	35.5	34.8	33.6	32.1	31.7	31.9	32.5
0.15	37.6	37.3	36.7	35.5	34.6	33.8	33.6	34.2
0.20	39.3	39,3	38.9	37.7	36.7	35.6	35.0	35.2
0.25	39.2	39.6	39.9	38.8	38.0	36.3	34.8	34.4
0.30	38.0	38.5	39.1	38.5	37.4	35.1	33.2	32.2
0.35	35.1	35.8	36.7	36.2	35.5	33.1	30.8	29.7
0.40	31.8	32.9	34.1	33.9	33.2	31.1	28.9	27.5
0.45	28.9	29.9	31.5	31.6	31.0	29.3	27.4	26.5
0.50	26.2	27.3	28.9	29.9	29.5	28.1	26.6	25.8
0.55	23.8	24.9	26.7	28.6	28.4	27.4	26.3	25.7
0.60	21.7	22.6	24.5	27.4	27.9	27.2	26.3	25.9
0.65	20.1	20.8	22.4	25.5	27.1	27.3	26.7	26.5
0.70	18.8	19.4	20.5	23.4	25.7	27.1	27.3	27.1
0.75	17.7	18.2	19.1	21.4	24.0	26.5	27.5	27.6
0.80	17.2	17.6	18.1	19.8	22.0	25.2	27.1	27.9
0.85	16.9	17.1	17.4	18.5	20.1	23.1	25.6	26.9
0.90	16.5	16.6	16.8	17.5	18.6	20.9	23.3	25.2
0.95	16.3	16.4	16.6	17.0	17.7	19.3	21.2	23.0
1.00	16.3	16.4	16.5	16.6	17.2	18.3	19.5	21.0
1.05	16.4	16.5	16.5	16.6	17.0	17.7	18.2	19.4
1.10	16.5	16.6	16.6	16.6	17.0	17.3	17.7	18.4
1.15		16.8	16.8	16.8	17.0	17.3	17.6	18.1
1.20	17.0	17.0	17.1	17.1	17.3	17.3	17.5	17.8
1.25			17.4	17.3	17.5	17.5	17.8	17.9
1.30	17.6	17.6	17.6	17.6	17.9	17.8	17.9	18.0
1.35			18.0	18.0	18.3	18.1	18.1	18.2
1.40	18.3	18.4	18.4	18.4	18.5	18.4	18.5	18.5
1.45						18.9	18.9	18.9
1.50	19.2	19.2	19.3	19.2	19.4	19.3	19.4	19.4

The electrical double layer differential capacity on mercury electrode 0.5 M H_2SO_4 media. Given are the capacity in μF cm^{-2}, potential in V vs. SCE.

Reference: M.K. Nauruzbaev, M.M. Ibraev and V.P. Gladyshev.
Bulletin des Ecoles Supérieures de Chimie et Technologie chimique URSS, 19 (1976) 1221.

E/V	C/μF cm^{-2} TBA = 2.10^{-5} mol l^{-1}	TBA = 5.10^{-4} mol l^{-1}	TBA = 5.10^{-3} mol l^{-1}
+0.10	45.5	42.5	42.3
+0.15	37.9	39.0	39.0
+0.10	35.1	37.0	37.1
+0.05	33.1	35.6	35.8
0.00	32.3	35.0	35.1
-0.05	30.5	34.0	34.0
-0.10	30.7	33.3	33.5
-0.15	30.3	32.8	40.2
-0.20	29.9	32.5	58.0
-0.25	30.0	32.2	31.5
-0.30	30.7	32.3	10.1
-0.35	30.3	26.2	8.0
-0.40	25.4	17.4	7.7
-0.45	20.2	10.8	6.7
-0.50	17.0	8.7	6.4
-0.55	14.5	8.1	6.2
-0.60	13.3	7.3	6.0
-0.65	12.5	6.7	6.0
-0.70	11.7	6.7	6.0
-0.75	11.3	6.7	6.0
-0.80	11.3	6.8	6.0
-0.85	10.9	6.8	6.0
-0.90	10.9	7.5	6.0
-0.95	10.9	8.1	6.1
-1.00	11.3	8.9	6.4
-1.05	11.7	9.3	6.5
-1.10	12.5	9.7	7.3
-1.15	13.3	10.9	8.2
-1.20	13.7	12.9	8.9

TBA is tert-butylamine

Differential capacity on mercury electrode in 0.5 M K_2SO_4 at 25°C. Adsorption of pyridine - given are the capacity in $\mu F\ cm^{-2}$, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas. Unpublished results.

	$C/\mu F\ cm^{-2}$							
$c/mol\ l^{-1}$	0.003	0.01	0.02	0.03	0.05	0.1	0.2	0.3
-E/V								
0.350	21.3							
0.400	36.7	41.4	61.8	62.4	75.5	73.2	58.0	54.2
0.425			59.8	60.2	69.1	66.6	53.7	49.9
0.450	44.3	46.4	57.4	57.5	63.2	60.6	49.7	45.9
0.500	46.5	46.9	51.6	51.3	52.7	49.8	42.4	38.7
0.550	45.2	44.4	45.3	44.5	43.8	40.9	36.2	32.6
0.600	41.8	40.2	38.9	37.9	36.4	33.6	30.8	27.5
0.625	39.7	37.7	35.8	34.7	33.2	30.5	28.4	25.3
0.650	37.5	35.2	32.8	31.8	30.4	27.7	26.2	23.2
0.675	35.3	32.8	30.0	29.1	27.8	25.3	24.2	21.4
0.700	33.1	30.3	27.5	26.6	25.5	23.2	22.3	19.8
0.725	31.1	28.0	25.1	24.4	23.5	21.3	20.7	18.3
0.750	29.2	25.9	22.9	22.4	21.7	19.7	19.2	17.0
0.775	27.4	24.0	21.0	20.7	20.2	18.3	17.8	15.9
0.800	25.9	22.3	19.4	19.2	18.9	17.2	16.6	14.9
0.825	24.6	20.8	18.0	18.0	17.7	16.2	15.5	14.0
0.850	23.5	19.6	16.8	16.9	16.8	15.5	14.5	13.2
0.875	22.6	18.5	15.8	16.1	16.0	14.8	13.7	12.6
0.900	21.9	17.7	15.1	15.4	15.3	14.3	12.9	12.0
0.950	20.9	16.7	14.1	14.5	14.4	13.7	11.7	11.2
1.000	20.4	16.3	13.9	14.0	13.8	13.3	10.8	10.5
1.050	20.1	16.4	14.1	13.7	13.4	13.0	10.1	10.1
1.100	19.7	16.8	14.7	13.4	13.1	12.7	9.5	9.7
1.150	19.2	17.2	15.5	13.2	12.7	12.2	9.0	9.3
1.200	18.5	17.7	16.4	13.0	12.2	11.3	8.6	8.8
1.250	17.7	18.3	17.3	13.1	11.2	9.8	8.0	8.0
1.300	17.1	19.1	18.3	13.7	9.8	7.6	7.3	7.0
1.350	17.3	20.4	19.2	15.4	7.7	4.6	6.4	5.6
1.400	19.0	22.8	20.5	18.9	4.9	0.5	5.2	3.8
1.450	23.4	27.1	22.2	25.1	1.2	-4.8	3.6	1.3
1.500	31.9	34.3	24.8	35.3	-3.6	-11.6	1.5	-1.8

Differential capacity on mercury electrode in 1 M Na_2SO_4; pH adjusted to 11.6 at 25°C. Adsorption of furfurylalcohol - given are the capacity in $\mu F\ cm^{-2}$, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas and J.M. Sedlak. Electrochimica Acta, 16 (1971) 2091.

	$C/\mu F\ cm^{-2}$								
$c/mol\ l^{-1}$	0	0.01	0.02	0.03	0.04	0.05	0.075	0.1	0.15
-E/V									
0.350				105.7	117.0				
0.375				96.4	104.7				
0.400	126.3			87.7	93.4				
0.425		81.6	93.6	79.6	83.1	96.2	97.2	81.3	63.8
0.450	95.4	75.0	83.5	72.0	73.7	83.9	81.2	68.8	54.1
0.500	72.3	63.0	65.9	58.4	57.4	63.2	55.7	48.6	38.5
0.550	55.4	52.5	51.5	46.9	44.1	46.8	37.5	33.9	27.3
0.600	43.7	43.5	40.0	37.2	33.6	34.2	25.2	23.5	19.6
0.650	35.8	35.9	30.9	29.2	25.4	24.8	17.4	16.6	14.7
0.675		32.7	27.2	25.8	22.1	21.1	14.9	14.2	13.0
0.700	30.9	29.8	24.1	22.8	19.2	18.0	13.0	12.4	11.7
0.725		27.2	21.4	20.2	16.8	15.4	11.7	11.0	10.6
0.750	28.1	24.9	19.1	17.9	14.8	13.3	11.0	10.0	10.2
0.775		22.9	17.3	15.9	13.2	11.6	10.5	9.4	9.8
0.800	26.7	21.3	15.8	14.3	11.9	10.3	10.4	8.9	9.6
0.825	26.3	19.9	14.7	12.9	10.9	9.3	10.5	8.7	9.5
0.850	26.0	18.7	13.8	11.8	10.2	8.6	10.6	8.6	9.5
0.875	25.8	17.8	13.3	11.0	9.7	8.1	10.9	8.6	9.5
0.900	25.6	17.2	13.0	10.5	9.5	7.9	11.1	8.7	9.5
0.950	25.1	16.5	13.1	10.0	9.6	8.0	11.5	8.9	9.4
1.000	24.2	16.6	13.8	10.4	10.3	8.5	11.4	8.9	9.1
1.050	22.9	17.2	15.0	11.4	11.6	9.5	10.9	8.8	8.5
1.100	21.3	18.2	16.6	13.0	13.1	10.8	10.0	8.5	7.7
1.150	19.3	19.3	18.2	15.0	14.8	12.3	8.9	8.2	6.8
1.200	17.4	20.5	19.8	17.4	16.7	14.2	7.9	8.2	6.0
1.250	15.9	21.4	21.3	19.9	18.6	16.4	7.7	8.7	5.6
1.300	15.3	21.9	22.6	22.5	20.5	19.1	8.8	10.4	6.1
1.350	16.3	21.7	23.4	25.1	22.4	22.5	12.2	13.6	7.9
1.400	19.8	20.4	23.8	27.5	24.2	26.8	18.8	19.2	11.7
1.450	26.5	17.9	23.7	29.7	26.1	32.5	29.7	28.0	18.2
1.500	37.5	13.9	23.0	31.5	28.0	39.7	46.3	40.7	28.1
1.550		8.0	21.8	32.8	29.9	49.0	69.9	58.5	42.3

Differential capacity on mercury electrode in 1 M Na_2SO_4 at 25°C. Adsorption of furfurylamine - given are the capacity in $\mu F\ cm^{-2}$, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas and J.M. Sedlak. Electrochimica Acta, 16 (1971) 2091.

	$C/\mu F\ cm^{-2}$							
$c/mol\ l^{-1}$	0.004	0.01	0.02	0.03	0.05	0.1	0.2	0.3
-E/V								
0.350	51.3	64.2	86.7	107.5	173.9	166.4	57.8	41.0
0.400	55.9	71.6	87.6	100.0	138.5	130.3	50.7	
0.450	56.4	71.4	81.6	88.0	106.4	99.7	44.1	32.1
0.500	54.1	66.1	71.6	73.7	79.5	74.2	38.3	24.8
0.550	50.0	58.0	59.6	59.1	57.7	53.5	28.2	18.9
0.600	44.9	48.5	47.5	45.4	40.6	37.1	20.3	14.4
0.625	42.2	43.7		39.2	33.7	30.4	17.1	12.5
0.650	39.5	39.1	36.3	33.5	27.8	24.5	14.3	10.9
0.675	36.8	34.6	31.3	28.4	22.9	19.5	11.9	9.6
0.700	34.1	30.5	26.9	23.9	18.8	15.3	10.0	8.4
0.725	31.6	26.7	22.9	20.0	15.4	11.9	8.4	7.6
0.750	29.2	23.4	19.6	16.7	12.8	9.1	7.1	6.9
0.775	27.0	20.4	16.7	14.1	10.9	6.9	6.2	6.3
0.800	25.0	18.0	14.5	12.0	9.5	5.4	5.5	6.0
0.825	23.3	15.9	12.7	10.5	8.5	4.3	5.1	5.8
0.850	21.7	14.3	11.4	9.5	8.0	3.7	4.9	5.7
0.875	20.3	13.1	10.6	8.8	7.8	3.6	4.9	5.8
0.900	19.2	12.3	10.1	8.6	7.8	3.8	5.1	5.9
0.950	17.5	11.7	10.1	8.9	8.4	5.0	5.9	6.4
1.000	16.5	12.2	10.9	9.8	9.2	6.9	7.0	7.0
1.050	16.2	13.4	12.0	11.0	10.0	9.2	8.3	7.7
1.100	16.4	14.8	13.0	11.8	10.3	11.4	9.5	8.3
1.150	16.9	16.2	13.6	12.1	10.1	13.0	10.3	8.6
1.200	17.6	17.3	13.7	11.8	9.6	13.7	10.6	8.5
1.250	18.4	18.0	13.4	11.1	8.9	12.8	10.1	7.9
1.300	19.3	18.4	13.3	10.6	8.7	10.2	8.7	6.6
1.350	20.2	18.9	14.1	11.1	9.7	5.2	6.0	4.6
1.400	21.2	19.9	17.3	14.0	12.9	-2.6	1.9	1.6
1.450	22.5	22.6	24.5	21.1	19.8	-13.5	-3.8	-2.5
1.500	24.5	28.1	38.2	34.9	32.0	-28.1	-10.6	

Differential capacity on mercury electrode in 1 M Na_2SO_4 at 25°C. Adsorption of tetrahydrofuran - given are the capacity in $\mu F\ cm^{-2}$, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas and J.M. Sedlak. Electrochimica Acta, 16 (1971) 2091.

	$C/\mu F\ cm^{-2}$							
$c/mol\ l^{-1}$	0.01	0.03	0.05	0.075	0.10	0.15	0.20	0.30
-E/V								
0.375	16.9	-11.2	-48.9	-8.8	78.7	230.9	235.8	195.8
0.400	27.7	11.6	-5.2	27.9	90.1	194.7	195.5	159.8
0.450	41.8	42.2	49.9	69.6	97.0	134.8	130.1	102.0
0.500	48.3	57.2	73.3	82.6	89.4	89.6	82.0	60.6
0.550	49.5	61.1	76.0	77.6	73.9	56.7	48.2	32.5
0.600	47.2	57.9	66.8	71.2	55.4	33.6	25.7	14.8
0.650	42.9	54.6	59.8	63.2	37.4	18.5	12.0	5.1
0.675	40.3	50.5	52.1	54.4	29.4	13.4	7.8	2.5
0.700	37.7	46.0	44.2	45.4	22.3	9.6	4.9	1.1
0.725	35.1	41.3	36.6	36.7				
0.750	32.5	36.6	29.4	28.5	11.1	5.1	2.5	1.0
0.775	30.1	31.9	23.1	21.3				
0.800	27.9	27.6	17.7	15.1		3.8	2.9	3.1
0.825	25.9	23.6	13.3	10.1	4.4			
0.850	24.1	20.0	9.9	6.4	1.6	4.4	4.9	6.1
0.875	22.6	16.9	7.5	2.3				
0.900	21.4	12.1	5.3	2.3	2.0	5.9	7.3	8.9
0.950	19.6	9.3	5.9	4.9	4.4	7.6	9.3	10.9
1.000	18.7	8.3	7.8	8.5	7.5	6.9	10.2	11.6
1.050	18.5	8.7	9.7	11.7	10.0	9.5	10.0	10.8
1.100	18.6	10.0	10.5	13.0	11.1	9.1	8.5	8.6
1.150	18.9	11.8	9.5	11.6	10.1	8.0	6.0	5.6
1.200	19.0	13.8	6.9	7.8	7.1	6.2	3.2	2.3
1.250	18.9	16.0	3.8	2.6	3.1	4.4	1.0	0.2
1.300	18.4	18.3	2.6	-1.5	-0.2	3.1	0.4	0.5
1.350	17.6	21.3	7.4	-0.2	0.4	3.4	3.0	2.8
1.400	16.9	25.8	24.0	12.7	9.3	6.2	10.4	11.8
1.450	16.6	33.3	60.2	45.5	32.6	13.0	24.7	28.5
1.500	17.5	45.7	126.4	109.0	77.9	25.2	48.2	55.4

Differential capacity on mercury electrode in 1 M Na_2SO_4 at 25°C. Adsorption of tetrahydrofurfurylamine. Given are the capacity in μFcm^{-2}, potential in V vs internal Hg/Hg_2SO_4 ref electrode. Reference: R.G. Barradas, J.M. Sedlak, Electrochemica Acta. 16 (1971) 2091.

	$C/\mu F\ cm^{-2}$									
$c/mol\ l^{-1}$ $-E/V$	0	0.004	0.01	0.02	0.03	0.04	0.05	0.075	0.1	0.15
0.300	22.5									
0.350	38.7									
0.375		48.7	32.5	61.2						
0.400	47.8	49.3	39.4	64.7	103.0	122.5	136.1	163.2	184.9	
0.425					96.7	112.0	122.7	143.0	158.4	154.5
0.450	51.5	50.3	48.8	67.7	90.4	102.1	110.2	124.8	135.0	131.6
0.475						92.8		108.5	114.3	111.3
0.500	51.5	49.6	53.3	66.4	77.7		88.0	93.9	96.1	93.4
0.525						75.7	78.1	81.0	80.2	77.9
0.550	48.9	47.8	53.9	61.9	65.5		69.1	69.5	66.5	64.4
0.575						60.8	60.8	59.4	54.8	52.8
0.600	44.9	45.1	51.7	55.4	54.0		53.3	50.6	44.8	42.9
0.625			49.8	51.6	48.6	47.9	46.5	43.0	36.4	34.6
0.650	40.3	41.9	47.4	47.6	43.5	42.3	40.4	36.4	29.4	27.7
0.675		40.1	44.8	43.5	38.7	37.1	35.0	30.7	23.7	22.1
0.700	35.6	38.3	41.9	39.4	34.3	32.4	30.1	25.9	19.2	17.6
0.725		36.4	38.9	35.4	30.1					
0.750	31.3	34.5	35.8	31.5	26.4	25.4	22.1	18.6	12.8	11.3
0.775		32.6	32.7	27.7	23.0					
0.800	27.6	30.8	29.7	24.2	19.9	18.3	16.2	13.7	9.4	7.9
0.825		29.0	26.7	21.0	17.3					
0.850	25.1	27.2	24.0	18.1	15.0	13.7	12.1	10.6	8.0	6.6
0.875		25.5	21.4							
0.900	23.3	23.9	19.0	13.4	11.6	9.6	9.5	8.9	7.9	6.7
0.950	22.4	21.0	15.1	10.2	9.5	8.3	8.1	8.1	8.5	7.5
1.000	22.0	18.6	12.3	8.6	8.6	8.0	7.8	7.9	9.2	8.5
1.050	22.2	16.8	10.8	8.4	8.8	8.4	8.2	7.9	9.8	9.3
1.100	22.5	15.4	10.5	9.3	9.7	9.3	9.1	8.1	9.9	9.8
1.150	22.7	14.7	11.1	11.1	11.0	10.3	10.1	8.2	9.6	9.7
1.200	22.6	14.5	12.5	13.1	12.4	11.1	11.1	8.3	8.8	9.1
1.250	21.8	14.8	14.3	14.9	13.4	11.5	11.7	8.3	7.8	8.1
1.300	20.3	15.6	16.1	15.5	13.4	11.0	11.6	8.3	6.8	6.8
1.350	17.7	16.7	17.2	14.3	11.9	9.2	10.7	8.6	6.3	5.8
1.400	14.2	18.1	17.0	10.0	8.4	5.9	8.6	9.5	6.9	5.4
1.450	9.6	19.5	14.8	1.6	2.0	0.5	5.1	11.1	9.2	6.3
1.500	4.2	20.9	9.7	-12.1	-8.0	-7.3	-0.2	14.1	14.2	9.2
1.550								18.8	22.8	14.9
1.600								25.8	36.2	24.5

Differential capacity for mercury in solutions of 0.1 M sodium formate in formic acid containing di-n-butyl ether. T = 25°C

Potential with respect to a saturated calomel electrode in formic acid.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	$C/\mu F\ cm^{-2}$							
$c/mol\ l^{-1}$	0.0099	0.0176	0.0277	0.0328	0.0360	0.0473	0.0665	0.0792
E/V								
0.07					25.03			
0.04					21.77	21.73	21.89	22.32
0.01					20.69	20.72	21.03	21.35
-0.02	19.24	19.33	18.86	18.95	19.77	19.73	20.32	20.64
-0.05	18.35	18.51	18.05	18.21	18.98	19.02	19.78	20.07
-0.08	17.52	17.72	17.39	17.52	18.24	18.38	19.36	19.69
-0.11	16.80	16.99	16.76	16.86	17.66	17.81	19.06	19.57
-0.14	16.13	16.33	16.14	16.32	17.13	17.44	18.92	19.63
-0.17	15.49	15.75	15.59	15.92	16.63	17.26	18.93	19.93
-0.20	14.90	15.19	15.16	15.61	16.29	17.30	19.11	20.45
-0.23	14.34	14.65	14.59	15.06	15.80	16.65	18.56	19.76
-0.26	13.74	14.10	14.08	14.47	15.32	16.03	17.77	18.30
-0.29	13.20	13.55	13.58	13.97	14.86	15.43	16.40	16.37
-0.32	12.63	13.03	13.09	13.49	14.36	14.78	14.82	14.54
-0.35	12.10	12.48	12.57	12.96	13.80	14.03	12.91	12.51
-0.38	11.56	11.94	12.04	12.38	13.09	12.96	11.49	11.12
-0.41	11.06	11.43	11.51	11.82	12.30	11.92	10.40	9.87
-0.44	10.66	10.98	11.00	11.19	11.50	10.88	9.44	9.02
-0.47	10.24	10.52	10.47	10.59	10.69	10.01	8.68	8.34
-0.50	9.86	10.10	9.95	9.98	10.00	9.22	8.05	7.75
-0.55	9.37	9.44	9.24	9.07	8.94	8.17	7.30	7.07
-0.60	8.98	8.95	8.54	8.25	8.04	7.41	6.72	6.53
-0.65	8.71	8.56	7.93	7.58	7.37	6.82	6.28	6.15
-0.70	8.53	8.19	7.40	7.05	6.85	6.38	5.98	5.86
-0.75	8.42	7.90	6.96	6.62	6.45	6.07	5.75	5.64
-0.80	8.41	7.67	6.66	6.33	6.18	5.84	5.58	5.48
-0.85	8.52	7.57	6.47	6.17	6.02	5.72	5.48	5.38
-0.90	8.79	7.66	6.45	6.15	5.98	5.69	5.48	

Differential capacity of mercury in 1.06 mol l^{-1} NaOH in water containing methionine. T = 25°C

Potentials with respect to hydrogen electrode in the base solution.

R.PARSONS and R. PAYNE unpublished work.

c/mol l^{-1}	0		0.002292		0.00578		0.01009	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-26			-1.694	22.04	-1.694	22.04	-1.694	22.04
-25			-1.648	21.28	-1.648	21.30	-1.648	21.27
-24	-1.600	20.52	-1.600	20.58	-1.600	20.60	-1.600	20.57
-23	-1.551	19.89	-1.551	19.94	-1.551	19.96	-1.551	19.94
-22	-1.500	19.34	-1.500	19.39	-1.500	19.41	-1.500	19.40
-21	-1.447	18.82	-1.448	18.88	-1.448	18.92	-1.448	18.91
-20	-1.393	18.34	-1.394	18.41	-1.394	18.44	-1.394	18.45
-19	-1.338	17.90	-1.339	17.97	-1.339	18.03	-1.339	18.05
-18	-1.282	17.48	-1.283	17.56	-1.283	17.64	-1.283	17.69
-17	-1.224	17.09	-1.225	17.19	-1.226	17.31	-1.226	17.40
-16	-1.165	16.75	-1.166	16.89	-1.168	17.05	-1.168	17.21
-15	-1.104	16.48	-1.107	16.66	-1.109	16.89	-1.110	17.12
-14	-1.043	16.26	-1.046	16.51	-1.050	16.84	-1.052	17.15
-13	-0.981	16.15	-0.986	16.49	-0.990	16.91	-0.994	17.32
-12	-0.919	16.14	-0.925	16.60	-0.931	17.12	-0.936	17.61
-11	-0.858	16.26	-0.865	16.87	-0.874	17.48	-0.880	18.00
-10	-0.797	16.56	-0.807	17.28	-0.817	17.93	-0.825	18.43
- 9	-0.737	17.04	-0.750	17.79	-0.762	18.42	-0.772	18.85
- 8	-0.679	17.72	-0.694	18.46	-0.709	18.96	-0.719	19.26
- 7	-0.624	18.59	-0.641	19.20	-0.657	19.52	-0.668	19.63
- 6	-0.572	19.62	-0.590	20.01	-0.606	20.11	-0.617	20.04
- 5	-0.522	20.80	-0.541	20.88	-0.557	20.75	-0.568	20.49
- 4	-0.476	22.05	-0.495	21.83	-0.510	21.48	-0.520	21.05
- 3	-0.432	23.31	-0.450	22.82	-0.464	22.28	-0.472	21.74
- 2	-0.390	24.54	-0.407	23.84	-0.420	23.18	-0.428	22.59
- 1	-0.350	25.74	-0.366	24.86	-0.378	24.15	-0.384	23.55
0	-0.312	26.86	-0.326	25.85	-0.337	25.17	-0.343	24.62
1	-0.275	27.90	-0.288	26.85	-0.298	26.20	-0.303	25.74
2	-0.240	29.06	-0.252	27.88	-0.261	27.30	-0.265	26.91
3	-0.206	30.60	-0.217	29.07	-0.225	28.51	-0.229	28.30
4	-0.175	32.81	-0.183	30.81	-0.191	30.15	-0.194	29.89
5	-0.146	36.41	-0.152	33.32	-0.159	32.31	-0.162	31.99
6			-0.124	37.34	-0.129	35.57	-0.132	34.95

Capacity of mercury in 1.06 mol l^{-1} NaOH in water containing methionine (cont.)

$c/\text{mol l}^{-1}$	0.01678		0.02661		0.05155		0.1000	
$\sigma/\mu\text{C cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$	E/volt	$C/\mu\text{F cm}^{-2}$
-26	-1.694	22.04	-1.694	22.04	-1.694	22.04	-1.694	22.04
-25	-1.648	21.31	-1.648	21.30	-1.648	21.34	-1.648	21.34
-24	-1.600	20.62	-1.600	20.62	-1.600	20.68	-1.600	20.71
-23	-1.551	19.99	-1.551	20.00	-1.551	20.06	-1.551	20.16
-22	-1.500	19.47	-1.500	19.47	-1.501	19.58	-1.501	19.73
-21	-1.448	18.99	-1.448	19.02	-1.449	19.15	-1.450	19.40
-20	-1.395	18.55	-1.395	18.60	-1.396	18.82	-1.398	19.20
-19	-1.340	18.19	-1.341	18.28	-1.343	18.57	-1.346	19.10
-18	-1.285	17.87	-1.286	18.02	-1.289	18.44	-1.294	19.14
-17	-1.229	17.65	-1.230	17.86	-1.235	18.42	-1.242	19.28
-16	-1.172	17.52	-1.174	17.83	-1.180	18.54	-1.190	19.54
-15	-1.115	17.49	-1.118	17.92	-1.127	18.76	-1.139	19.83
-14	-1.058	17.61	-1.062	18.14	-1.074	19.07	-1.089	20.11
-13	-1.001	17.86	-1.008	18.44	-1.022	19.37	-1.040	20.28
-12	-0.946	18.20	-0.954	18.79	-0.970	19.61	-0.990	20.33
-11	-0.892	18.58	-0.901	19.09	-0.920	19.73	-0.941	20.21
-10	-0.838	18.94	-0.849	19.33	-0.899	19.74	-0.891	19.95
- 9	-0.786	19.23	-0.798	19.46	-0.818	19.62	-0.841	19.59
- 8	-0.734	19.45	-0.746	19.51	-0.767	19.42	-0.789	19.19
- 7	-0.683	19.65	-0.695	19.54	-0.715	19.21	-0.737	18.80
- 6	-0.632	19.86	-0.644	19.59	-0.663	19.07	-0.683	18.49
- 5	-0.582	20.06	-0.593	19.74	-0.610	19.05	-0.629	18.35
- 4	-0.532	20.26	-0.543	20.06	-0.558	19.25	-0.574	18.45
- 3	-0.484	21.20	-0.493	20.57	-0.507	19.69	-0.520	18.84
- 2	-0.438	21.97	-0.446	21.33	-0.457	20.40	-0.468	19.52
- 1	-0.393	22.95	-0.400	22.31	-0.409	21.43	-0.418	20.62
0	-0.350	24.06	-0.356	23.51	-0.364	22.80	-0.371	22.15
1	-0.310	25.28	-0.315	24.81	-0.321	24.28	-0.328	23.78
2	-0.271	26.56	-0.275	26.21	-0.281	25.82	-0.287	25.45
3	-0.235	28.01	-0.238	27.72	-0.243	27.47	-0.249	27.32
4	-0.200	29.64	-0.205	29.47	-0.208	29.37	-0.214	29.02
5	-0.167	31.73	-0.171	31.57	-0.176	31.60	-0.181	31.73
6	-0.137	34.45	-0.140	34.29	-0.145	34.36	-0.150	34.60
7							-0.123	38.35

Differential capacity of mercury in 0.9845 mol l^{-1} $HClO_4$ in water containing methionine.
T = 25°C

Potential with respect to a hydrogen electrode in the base solution.

R.PARSONS and R.PAYNE unpublished work

c/mol l^{-1}	0		0.001803		0.0050		0.010	
σ/μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-9	-0.638	18.59						
-8	-0.587	20.18	-0.598	20.08	-0.606	20.24	-0.617	20.16
-7	-0.539	22.02	-0.550	21.81	-0.559	21.82	-0.569	21.66
-6	-0.495	23.93	-0.506	23.60	-0.515	23.48	-0.525	23.21
-5	-0.455	25.67	-0.465	25.17	-0.473	24.87	-0.483	24.46
-4	-0.417	27.15	-0.426	26.47	-0.434	25.99	-0.443	25.42
-3	-0.381	28.27	-0.389	27.42	-0.396	26.72	-0.404	25.99
-2	-0.346	28.99	-0.353	27.97	-0.359	27.08	-0.366	26.24
-1	-0.312	29.30	-0.318	28.15	-0.322	27.11	-0.327	26.18
0	-0.278	29.24	-0.282	27.98	-0.285	26.86	-0.289	25.84
1	-0.244	28.85	-0.246	27.54	-0.247	26.37	-0.250	25.37
2	-0.209	28.21	-0.209	26.90	-0.209	25.78	-0.210	24.80
3	-0.173	27.42	-0.172	26.16	-0.170	25.11	-0.169	24.21
4	-0.136	26.51	-0.133	25.38	-0.129	24.44	-0.128	23.66
5	-0.097	25.60	-0.093	24.64	-0.088	23.87	-0.085	23.18
6	-0.057	24.77	-0.052	24.00	-0.046	23.40	-0.041	22.86
7	-0.016	24.09	-0.010	23.52	-0.003	23.10	0.003	22.70
8	0.026	23.61	0.033	23.23	0.041	22.97	0.047	22.69
9	0.068	23.32	0.076	23.11	0.084	23.02	0.091	22.88
10	0.111	23.24	0.120	23.17	0.128	23.26	0.134	23.22
11	0.154	23.34	0.163	23.41	0.170	23.65	0.177	23.71
12	0.197	23.60	0.205	23.79	0.212	24.13	0.218	24.30
13	0.239	23.96	0.247	24.25	0.253	24.74	0.259	24.98
14	0.280	24.45	0.287	24.81	0.293	25.40	0.298	25.77
15	0.321	25.00	0.327	25.46	0.332	26.16	0.337	26.66
16	0.360	25.61	0.366	26.16	0.369	27.01	0.373	27.65
17	0.399	26.31	0.403	27.00	0.406	27.98	0.409	28.78
18	0.436	27.14						
19	0.472	28.09						
20	0.507	29.43						
21	0.540	31.56						
22	0.571	33.31						

Capacity of mercury in 0.9845 mol l^{-1} $HClO_4$ in water containing methionine (cont.)

c/mol l^{-1}	0.020		0.040		0.060		0.100	
σ /μC cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}	E/volt	C/μF cm^{-2}
-7	-0.582	21.49	-0.595	21.15	-0.602	20.95	-0.611	20.67
-6	-0.537	22.85	-0.549	22.37	-0.556	22.03	-0.564	21.64
-5	-0.494	23.97	-0.505	23.33	-0.511	22.93	-0.519	22.45
-4	-0.453	24.75	-0.463	24.02	-0.468	23.53	-0.475	22.97
-3	-0.413	25.20	-0.421	24.37	-0.426	23.84	-0.432	23.24
-2	-0.373	25.35	-0.380	24.45	-0.384	23.90	-0.389	23.28
-1	-0.334	25.22	-0.339	24.28	-0.342	23.73	-0.345	23.13
0	-0.294	24.88	-0.298	23.97	-0.300	23.42	-0.302	22.85
1	-0.253	24.44	-0.256	23.54	-0.257	23.03	-0.258	22.51
2	-0.212	23.93	-0.213	23.09	-0.213	22.62	-0.213	22.17
3	-0.170	23.40	-0.169	22.67	-0.169	22.26	-0.168	21.87
4	-0.127	22.95	-0.125	22.32	-0.123	21.97	-0.122	21.66
5	-0.083	22.61	-0.080	22.09	-0.078	21.81	-0.075	21.58
6	-0.038	22.41	-0.034	22.01	-0.032	21.82	-0.029	21.64
7	0.006	22.38	0.011	22.09	0.014	21.95	0.017	21.88
8	0.051	22.51	0.056	22.33	0.059	22.27	0.062	22.28
9	0.095	22.81	0.100	22.77	0.104	22.74	0.107	22.80
10	0.139	23.26	0.144	23.33	0.147	23.36	0.150	23.52
11	0.181	23.86	0.186	24.01	0.189	24.11	0.192	24.32
12	0.222	24.57	0.227	24.83	0.230	24.98	0.232	25.24
13	0.262	25.37	0.267	25.75	0.269	25.97	0.271	26.28
14	0.301	26.28	0.305	26.80	0.307	27.06	0.308	27.44
15	0.338	27.31	0.341	27.95	0.343	28.28	0.344	28.74
16	0.374	28.49	0.376	29.27	0.378	29.70	0.378	30.22
17	0.409	29.82			0.411	31.30	0.410	31.92

Differential capacity on mercury electrode in 0.2 M HCl - adsorption of hyamine+). Given are the capacity in $\mu F\,cm^{-2}$, potential in mV/SCE. Temp. 25 °C.

Reference: B. Lovrecek, N. Ban , Lj. Duic and M. Bravar. Tenside 8 (1971) 140.

	$C/\mu F\,cm^{-2}$				
$c/mol\,l^{-1}$	0	10^{-3}	10^{-4}	10^{-5}	10^{-6}
-E/V					
1	98.00				118.00
10	81.30				86.20
50	58.20			91.00	60.50
100	50.40		101.00	73.40	54.00
150	44.60	94.50	81.30	54.10	47.00
200	41.80	56.40	46.50	45.75	42.80
250	41.00	28.40	26.80	35.00	42.20
275	44.60	22.20	21.30	30.50	43.25
300	44.60	18.10	16.90	26.25	43.25
325	45.00	15.20	14.60	22.70	44.00
350	45.75	12.90	13.25	19.90	42.20
375	45.75	11.60	11.80	16.20	41.80
400	45.00	10.85	10.40	13.60	40.50
450	41.00	10.00	9.40	11.50	38.30
500	36.40	10.40	8.85	9.40	34.00
600	27.10	12.60	10.60	8.85	23.40
700	21.60	14.96	13.60	9.40	19.70
800	19.90	16.00	15.20	10.00	17.40
900	18.10	18.00	16.00	11.30	16.50
1000	17.90	19.60	16.70	11.90	15.80
1100	17.70	18.60	16.70	12.25	16.20
1200	17.70	17.60	16.40	13.60	16.50

+) Hyamine is $(CH_3)_3CCH_2C(CH_3)_2C_6H_4(OC_2H_4)_2N^+(CH_3)_2CH_2C_6H_5\ Cl^-$

The electrical double layer differential capacity on mercury electrode in 0.5 M H_2SO_4 media (the dependence on potential); given are the capacity in μF cm^{-2}, potential in V vs SCE.

Reference: M.K. Nauruzbaev, M.M. Ibraev and V.P. Gladyshev. Elektrokhimiya 12 (1967) 1464.

	C/μF cm^{-2}		
E/V	ANP = 10^{-3} wt %	ANP = 5.10^{-3} wt %	ANP = 10^{-2} wt %
+0.20	44.5	45.0	45.0
+0.15	39.7	41.7	43.2
+0.10	38.0	40.3	42.6
+0.05	37.5	41.0	45.1
0.00	38.6	43.0	49.8
-0.05	40.2	44.8	37.3
-0.10	41.6	32.0	30.0
-0.15	34.7	18.8	11.2
-0.20	27.5	14.3	10.0
-0.25	24.8	11.4	9.1
-0.30	22.5	10.0	8.5
-0.35	20.5	9.0	7.4
-0.40	20.0	8.1	7.0
-0.45	17.8	7.5	6.8
-0.50	17.4	7.4	6.5
-0.55	15.9	7.3	6.5
-0.60	15.3	7.2	6.5
-0.65	14.5	7.2	6.5
-0.70	14.1	7.2	6.5
-0.75	13.3	7.3	6.5
-0.80	13.6	7.5	6.5
-0.85	13.8	7.7	6.8
-0.90	14.0	8.0	7.2
-0.95	14.1	8.8	8.0
-1.00	14.2	9.5	9.2
-1.05	14.3	11.6	11.5
-1.10	14.5	13.5	13.5
-1.15	14.8	14.9	14.9
-1.20	15.0	15.2	15.0

ANP is an aliphatic amine of the type $R-NH_2$ where $R = CH_3(CH_2)_m-CH-(CH_2)_nCH_3$ and m+ n is between 10 and 16.

Differential capacity on mercury electrode in 0.5 M Na_2SO_4 + x mg l^{-1} PAA ($\bar{M}$ = 1.83×10^5) at 20°C. PAA = polyacrylic acid.

Reference electrode: saturated calomel electrode. Given are the capacity in μF cm^{-2}, potential in V.
Reference: T. Osaka, Y. Sawada and T. Yoshido. Denki Kagaku Ayobi,Kogyo Butsuri kagaku, 42 (1974) 477.

	C/μF cm^{-2}			
(PAA)/mg l^{-1}	0	30	50	75
E/V				
0.20	44.98	-	-	-
0.15	41.03	-	-	-
0.10	39.30	-	-	-
0.05	39.05	-	-	-
0.00	38.80	-	-	-
-0.05	38.56	39.55	-	-
-0.10	38.06	33.00	30.89	29.66
-0.15	37.57	28.50	23.97	22.24
-0.20	36.58	25.46	20.39	18.54
-0.30	31.14	23.11	17.80	13.59
-0.40	26.82	20.76	16.07	11.74
-0.50	23.85	18.54	14.83	10.13
-0.60	21.26	16.93	13.59	9.39
-0.70	-	15.32	12.72	8.90
-0.80	16.93	14.34	12.11	9.14
-0.90	-	13.47	11.62	9.64
-1.00	15.57	13.47	11.67	10.26
-1.05	-	13.59	12.11	11.62
-1.10	-	14.83	13.30	14.09
-1.125	-	18.04	-	18.29
-1.15	-	19.77	21.38	22.24
-1.175	-	15.69	25.33	26.69
-1.20	15.32	15.57	21.82	32.13
-1.225	-	-	16.07	-
-1.25	-	16.93	16.31	17.30
-1.30	-	17.18	-	15.82
-1.40	16.68	-	-	-

Differential capacity on mercury electrode in 0.5 M Na_2SO_4 + x NaOH with 30 mg l^{-1} PAA ($\bar{M} = 1.83 \times 10^5$) at 20°C. PAA = polyacrylic acid.

Reference electrode: saturated calomel electrode. Given are the capacity in μF cm^{-2}, potential in V.

Reference: T. Osaka, Y. Sawada and T. Yoshida. Denki Kagaku Ayobi/Kogyo Butsuri kagaku, 42 (1974) 477.

	$C/\mu F\ cm^{-2}$		
-E/V	pH = 5.85	pH = 6.98	pH = 9.82
0.05	-	74.08	-
0.075	-	-	65.48
0.10	-	39.54	45.96
0.125	-	-	34.84
0.15	-	28.42	32.50
0.175	-	-	32.20
0.20	-	25.95	32.50
0.225	-	-	32.64
0.250	-	-	32.64
0.275	-	-	32.99
0.30	21.00	23.85	36.33
0.325	-	-	39.04
0.40	19.27	22.86	45.22
0.425	-	-	38.80
0.45	-	23.10	25.33
0.50	17.54	23.97	22.49
0.525	-	24.34	-
0.55	-	24.22	-
0.60	16.06	22.49	19.77
0.65	-	20.76	-
0.70	15.07	18.29	17.79
0.80	13.84	16.19	16.31
0.85	14.33	-	-
0.875	14.83	-	-
0.90	15.82	15.32	15.20
0.925	16.06	-	-
0.95	16.19	-	-
0.975	15.57	-	-
1.00	15.32	15.20	15.20
1.05	15.20	-	-
1.10	15.44	15.44	-
1.15	-	-	-
1.20	15.69	15.57	-
1.30	15.94	15.69	-

Differential capacity on mercury electrode in 0.5 M H_2SO_4 at 20°C.

Reference electrode: saturated calomel electrode. Given are the capacity in μF cm^{-2}, potential in V.
Reference: T. Osaka, Y. Sawada and T. Yoshida. Denki Kagaku Ayobi, Kogyo Butsuri kagaku, 42 (1974) 477.

PAA = polyacrylic acid ($\bar{M} = 1.83 \times 10^5$)

	C/μF cm^{-2}					
(PAA)/mg l^{-1}	0	10	30	50	75	150
E/V						
0.15	51.14	45.12	56.58	65.90	-	-
0.125	-	-	50.57	61.16	-	-
0.10	37.10	38.11	44.89	57.21	59.32	-
0.05	-	-	-	47.86	50.47	57.95
0.00	31.33	30.96	30.86	37.84	42.15	43.00
-0.10	29.49	28.09	25.08	24.48	27.12	27.45
-0.15	29.60	-	-	-	-	-
-0.20	30.56	27.76	21.74	17.40	16.23	15.26
-0.25	32.16	-	-	-	-	-
-0.30	32.67	28.76	21.74	13.46	11.22	10.86
-0.40	31.20	27.49	19.00	11.49	9.49	8.95
-0.50	25.82	23.25	14.30	11.22	8.45	8.05
-0.60	21.31	19.57	13.16	10.32	7.98	7.72
-0.70	18.74	-	12.63	10.22	7.82	7.55
-0.80	17.40	15.80	12.16	10.22	7.88	7.52
-0.90	15.97	-	12.16	10.25	8.22	7.92
-1.00	15.36	14.63	12.22	10.69	9.19	-

Differential capacity on mercury electrode in 0.5 M H_2SO_4 + xM NaOH with 30 mg l^{-1} PAA ($\bar{M}$ = 1.83×10^5) at 20 °C. Reference electrode: saturated calomel electrode. Given are the capacity in μF cm^{-2}, potential in V.

Reference: T. Osaka, Y. Sawada and T. Yoshida. Denki Kagaku Ayobi, Kogyo Butsuri Kagaku, 42 (1974) 477.

PAA = polyacrylic acid.

	$C/\mu F\ cm^{-2}$		
E/V	pH = 4.63	pH = 4.90	pH = 5.28
0.15	47.20	47.45	-
0.10	48.92	-	-
0.075	-	52.39	-
0.07	52.14	-	-
0.00	-	-	-
-0.025	51.15	59.31	-
-0.05	38.06	40.28	-
-0.10	29.41	30.64	30.64
-0.20	24.22	24.22	24.46
-0.30	21.25	20.76	21.25
-0.40	19.27	18.78	19.52
-0.50	17.42	17.54	17.79
-0.60	15.69	15.82	16.19
-0.70	13.96	14.41	14.70
-0.80	13.22	13.59	13.67
-0.85	-	-	-
-0.90	12.85	12.97	13.10
-0.95	12.92	12.97	13.12
-1.00	-	12.97	13.34
-1.05	13.05	13.10	15.14
-1.075	-	-	17.79
-1.10	13.47	14.09	15.82
-1.125	-	17.30	
-1.15	14.95	21.38	15.69
-1.175	-	16.31	
-1.20	21.13	15.69	15.57
-1.25	15.57	15.69	-
-1.30	15.69	15.94	16.06
-1.40	16.80	14.33	-

Differential capacity at a dropping mercury electrode in M KCl at 20°C. Given are the capacity in μF cm^{-2}, potential in V vs SCE.

Reference: R.J. Meakins. J. Appl. Chem. 17 (1967) 156.

C_{12}TMABr = n-dodecyltrimethylammonium bromide.

	C/μF cm^{-2}				
	C_{12}TMABr				
c/mol l^{-1}	0	5 x 10^{-5}	10^{-4}	3 x 10^{-4}	10^{-3}
E/V					
0.0	373	390	369	725	
0.1	58	61	70	131	
0.2	40.5	48	51	60	
0.3	39.2	33.7	33.1	39.5	
0.4	43.2	19.8	17.5	22.5	
0.5	39.7	11.9	10.5	12.5	
0.6	30.4	8.7	8.7	9.8	
0.7	22.6	7.9	8.2	9.1	
0.8	19.0	7.7	8.2	8.7	
0.9	17.3	7.9	8.3	8.7	
1.0	16.6	8.4	8.8	9.3	11.2
1.1	16.55	9.6	9.8	10.6	12.5
1.2	16.7	12.8	13.0	12.7	14.4
1.3	17.5	23.3	21.1	15.5	15.6
1.4	18.2	peak 28.4 23.5	32.1	17.7	17.0
1.5	19.2	19.4	21.3	19.2	18.2
1.6	20.6	20.4	20.4	23.1	19.4
1.7	22.2	22.1	21.9	27.4	20.9
1.8	25.4	24.5	24.3	22.7	24.2

Differential capacity in M KCl (continued)

C_{16}TMABr = n-hexadecyltrimethylammonium bromide

	$C/\mu F\ cm^{-2}$					
	C_{16}TMABr					
$c/mol\ l^{-1}$	0	1.5×10^{-4}	2×10^{-4}	3×10^{-4}	5×10^{-4}	10^{-3}
E/V						
0.0	373	329	342	262	209	
0.1	58	77	86	68	70	79
0.2	40.5	31.8	33.1	32.5	34.4	34.0
0.3	39.2	22.0	23.5	24.5	24.5	24.4
0.4	43.2	15.8	16.5	16.4	16.5	16.4
0.5	39.7	11.2	11.5	11.3	11.3	11.0
0.6	30.4	8.6	8.7	8.3	8.3	8.1
0.7	22.6	7.4	7.6	7.4	7.9	7.5
0.8	19.0	7.5	8.0	8.6	9.7	9.6
0.9	17.3	8.8	9.9	10.3	10.8	10.4
1.0	16.6	10.5	10.7	10.2	10.9	10.7
1.1	16.55	15.4	17.2	11.6	13.7	13.5
1.2	16.7	16.3	16.1	16.9	18.7	18.4
1.3	17.5	19.8	19.3	17.2	20.0	19.3
1.4	18.2	23.7	22.6	18.4	18.8	19.4
1.5	19.2	24.5	25.9	19.6	19.3	20.1
1.6	20.6	24.8	25.9	22.7	20.2	21.2
1.7	22.2	25.4	26.5	26.5	22.3	20.8
1.8	25.4	26.2	28.2	27.0	26.0	20.2

Differential capacity in M KCl (continued)

C_{12}TEABr = n-dodecyltriethylammonium bromide

	$C/\mu F\ cm^{-2}$					
	C_{12}TEABr					
$c/mol\ l^{-1}$	0	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-4}
E/V						
0.0	373	370	389	376	487	161
0.1	58	78	92	92	100	107
0.2	40.5	56	57	55	55	56
0.3	39.2	30.8	29.7	28.3	32.3	31.9
0.4	43.2	11.8	11.6	11.4	12.2	13.4
0.5	39.7	7.3	7.6	7.7	8.0	8.2
0.6	30.4	6.7	7.0	7.0	7.2	7.3
0.7	22.6	6.5	6.9	6.8	7.1	7.3
0.8	19.0	6.6	6.8	6.8	7.3	8.0
0.9	17.3	6.7	7.1	7.1	8.1	8.6
1.0	16.6	7.0	7.5	7.9	8.8	9.3
1.1	16.55	7.7	8.4	8.9	9.8	10.2
1.2	16.7	8.9	9.7	10.3	10.9	11.1
1.3	17.5	10.9	11.5	11.5	11.6	11.7
1.4	18.2	20.8	13.4	12.2	11.9	11.7
		peak 33.4				
1.5	19.2	31.2	23.0	13.0	11.9	11.6
			peak 29.4			
1.6	20.6	21.3	27.1	14.3	11.9	11.7
1.7	22.2	22.5	22.7	peak 26.9	12.9	11.8
1.8	25.4	25.1	25.3	24.5	26.0	14.1

Differential capacity in M KCl (continued)

$C_{16}TEABr$ = n-hexadecyltriethylammonium bromide

	$C/\mu F\ cm^{-2}$				
	$C_{16}TEABr$				
$c/mol\ l^{-1}$	0	6×10^{-5}	8×10^{-5}	1.5×10^{-4}	3×10^{-4}
E/V					
0.0	373	458	492	538	567
0.1	58	87	89	86	91
0.2	40.5	57	47	39.2	42
0.3	39.2	22.6	25.3	26.4	26.7
0.4	43.2	9.3	10.4	14.5	15.9
0.5	39.7	7.0	7.3	7.9	8.2
0.6	30.4	6.5	6.7	6.9	7.4
0.7	22.6	6.3	6.5	7.6	7.9
0.8	19.0	6.3	6.5	7.9	8.3
0.9	17.3	6.4	6.7	8.3	9.2
1.0	16.6	6.9	7.5	9.4	10.4
1.1	16.55	8.2	8.8	10.6	11.6
1.2	16.7	16.6	11.8	12.1	12.2
1.3	17.5	38.2	37.5	13.3	12.4
1.4	18.2	25.9	28.5	14.4	12.3
1.5	19.2	22.2	23.8	21.7	12.0
1.6	20.6	21.5	20.8	22.8	11.8
1.7	22.2	22.4	23.0	23.0	12.3
1.8	25.4	24.9	25.7	25.5	20.4

Differential capacity on mercury electrode in 1 M KCl.
Adsorption of C_8TPABr. Given are the capacity in μF cm^{-2}, potential in V vs SCE = C_8 TPABr = n octyltripropylammonium bromide. T = 20 °C.

Reference: R. J. Meakins, unpublished results.

	C/μF cm^{-2}							
	C_8TPABr							
c/mol l^{-1}	0	2 x 10^{-4}	5 x 10^{-4}	10^{-3}	2 x 10^{-3}	5 x 10^{-3}	10^{-2}	3 x 10^{-2}
E/V								
0	127	318	353	328	300	447	472	492
0.1	49.5	63.3	76.4	106	118	170	248	296
0.15		57.1	79.7	83.4			115	283
0.2	39.5	57.6	57.3	55.2	56.9	71.3	81.3	104.7
0.25		44.1	35.2	31.8	34.5	47.6	57.7	77.5
0.3	39.5	28.3	19.7	16.2	18.7	28.3	37.4	53.3
0.4	41.1	9.35	7.52	7.01	7.36	8.94	11.2	18.9
0.5	35.4	6.08	5.88					7.92
0.6	26.1	5.62	5.71	5.91	6.03	6.19	6.29	6.63
0.7	20.6							
0.8	18.1	5.69	5.85	6.07	6.09	6.24	6.31	6.86
1.0	16.1	5.94	6.03	6.24	6.32	6.48	6.70	7.48
1.2	16.3	6.52	6.46	6.69	6.80	7.15	7.75	8.41
1.4	17.6	9.05	7.97	7.78	7.75	8.51	9.04	9.37
1.45		17.3	8.89					
1.475		21.1						
1.5		20.2	12.7	9.51	9.10	9.61		
1.51			25.1					
1.52			23.6					
1.54								
1.55				36.5				
1.585								
1.59					59.8			
1.60	20.1	20.5	20.3	21.0	28.0	10.9	10.3	10.2
1.624								
1.64								

Differential capacity on mercury in 1 M KCl (continued)
Adsorption of C_8TPABr (continued)

	$C/\mu F\,cm^{-2}$							
	C_8TPABr							
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
E/V								
1.645						95.8		
1.65						56.6	10.8	
1.66						25.8		
1.666								
1.68								
1.69								
1.696							96.1	
1.70		21.8	21.9	21.9	22.1	22.1	73.0	10.4
1.71							37.6	
1.73								
1.75								10.3
1.77								
1.80	23.5	24.0	24.1	24.3	24.4		24.2	33.0
1.812								48.3
1.82								36.0
1.832								
1.84								27.8
1.86								

Differential capacity on mercury in 1 M KCl (continued)

Adsorption of C_8TBABr = n octyltributhylammonium bromide.

$C/\mu F\ cm^{-2}$

	C_8TBABr							
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
E/V								
0	127	295	310	276	289			
0.1	49.5	98.8	104	103	112	147	213	
0.15		76.1	75.4	75.6		83.8	98.7	
0.2	39.5	50.2	49.3	51.9	56.5	63.9	70.9	
0.25		29.4	29.2	32.6	37.6	47.4		
0.3	39.5	14.8	14.9	16.9	20.5	29.2	39.9	
0.4	41.1	6.14	6.24	6.61	6.95	8.48	11.3	
0.5	35.4							
0.6	26.1	5.33	5.37	5.48	5.64	5.84	5.88	
0.7	20.6							
0.8	18.1	5.38	5.41	5.57	5.78	6.10	6.50	
1.0	16.1	5.54	5.64	5.79	6.13	6.81	7.44	
1.2	16.3	5.95	6.05	6.32	6.83	7.62	8.29	
1.4	17.6	6.74	6.79	7.11	7.73	8.30	8.64	
1.45								
1.475								
1.5		7.86						
1.51								
1.52								
1.54		30.4						
1.55		28.6	7.96					
1.585								
1.59			40.9					
1.60	20.1	20.3	27.2	8.37	8.43	8.48	8.28	
1.624				50.4				
1.64					8.60			

Differential capacity on mercury in 1 M KCl (continued)
Adsorption of C_8TBABr (continued)

	$C/\mu F\ cm^{-2}$						
	C_8TBABr						
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}
E/V							
1.645							
1.65				23.1			8.14
1.66							
1.666					52.2		
1.68					31.8		
1.69							
1.696							
1.70		21.6	21.9	21.7	24.2	8.39	8.09
1.71							
1.73						24.5	
1.75						29.7	
1.77						29.7	
1.80	23.5	23.9	24.1	23.9	24.1	25.1	7.38
1.812							
1.82							36.5
1.832							
1.84							
1.86							29.6

Differential capacity in 1 M KCl (continued)
Adsorption of C_8TMABr = n octyltrimethylammonium bromide.

	$C/\mu F\ cm^{-2}$							
	C_8TMABr							
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
E/V								
0	127	282	274	262	412	358	466	737
0.05								
0.1	49.5	60.5	59.7	62.1	72.3	106	170	357
0.15					61.8	77.5		138
0.2	39.5	42.7	45.3	51.2	61.3	68.1	71.1	83.1
0.25		41.8	46.0	53.5	60.2			
0.3	39.5	42.6	47.6	50.5	53.5	52.1	53.1	56.4
0.4	41.1	37.0	35.1	34.0	33.6	33.1	33.5	35.5
0.5	35.4	22.3	20.3	19.3	18.6	18.2	18.3	19.6
0.6	26.1	13.9	12.8	12.1	11.8	11.5	11.5	12.1
0.7	20.6	10.7	9.98	9.6				
0.8	18.1	9.56	9.08	8.82	8.55	8.58	8.6	8.90
1.0	16.1	9.91	9.47	9.20	8.93	8.80	8.85	9.06
1.2	16.3	12.7	11.7	11.3	10.8	10.5	10.3	10.0
1.3		16.9	14.8	13.7				
1.4	17.6	19.0	20.9	20.2	20.0	16.1	14.9	15.1
1.45			21.5	24.5	28.2	20.3		
1.475					31.9			
1.5		19.5	21.0	24.8	31.5	30.2	21.6	18.9
1.525								
1.55						48.7	32.2	21.0
1.575							42.9	
1.6	20.1	20.2	20.3	21.0	21.6	31.7	57.9	23.8
1.61								
1.65							35.7	30.5
1.67								
1.675								
1.7		21.8	21.8	21.7	22.0	23.0	24.4	51.9
1.72								
1.722								69.3
1.75								51.2
1.775								
1.80	23.5	24.4	24.1	23.8	23.8	24.4	24.6	26.0

Differential capacity in 1 M KCl (continued)
Adsorption of C_8TEABr

$C/\mu F\ cm^{-2}$

	C_8TEABr							
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
E/V								
0	127	291	349	354	351	321		
0.05							338	
0.1	49.5	61.3	66.1	72.2	90.0	133	217	366
0.15			55.7					159
0.2	39.5	46.3	56.1	69.9	73.8	75.0	81.8	98.1
0.25		47.1	56.9	63.0	68.9	62.4	66.7	
0.3	39.5	45.6	51.5	52.5	50.3	50.5	51.8	60.4
0.4	41.1	26.3	26.4	25.2	23.9	24.1	25.1	29.8
0.5	35.4	11.0	10.6	10.1	9.75	9.73	10.0	12.2
0.6	26.1	7.16	7.07	6.90	6.79	6.86	7.02	7.67
0.7	20.6							
0.8	18.1	6.17	6.25	6.17	6.27	6.42	6.60	6.91
1.0	16.1	6.46	6.53	6.54	6.55	6.71	6.93	7.21
1.2	16.3	7.55	7.35	7.34	7.43	7.50	7.78	8.07
1.3		8.64	8.24					
1.4	17.6	11.4	9.92	9.33	9.19	9.14	9.28	9.94
1.45								
1.475								
1.5		21.2	17.0	12.5				
1.525			25.0					
1.55				19.2	13.4			
1.575				31.9				
1.6	20.1	20.5	21.4	23.9	28.8	14.2	12.8	12.4
1.61					43.7			
1.65					22.9	24.1	15.1	
1.67						87.7		
1.675						48.4		
1.7		21.8	22.0	22.1	22.4	24.6	25.9	14.3
1.72							117	
1.722							80.1	
1.75							24.8	
1.775								24.4
1.80	23.5	24.0	24.7	24.3	24.5	24.5	24.7	23.8

Differential capacity in 1 M KCl (continued)

Adsorption of C_{12}TPABr = n dodecyltripropylammonium bromide

	$C/\mu F\ cm^{-2}$							
	C_{12}TPABr							
$c/mol\ l^{-1}$	0	10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}	10^{-2}
E/V								
0	127	260	258	261	328	413	371	205
0.1	49.5	58.0	65.1	85.7	93.1	90.8	124	
0.2	39.5	42.0	45.0	47.3	40.0	37.4	35.8	78.3
0.25		37.8	35.8					
0.3	39.5	35.5	28.4	16.1	14.4	15.8	16.6	31.5
0.32		35.5						
0.35								
0.40	41.1			6.86	6.87	7.26	7.44	12.7
0.42		35.7						
0.45								
0.5	35.4	33.0	15.6					7.98
0.6	26.1	24.9	10.5	5.59	5.71	5.77	5.89	7.32
0.7	20.6	18.9						
0.75			9.5					
0.8	18.1	16.1		5.72	5.65	5.85	6.61	7.65
1.0	16.1	14.8	10.9	6.19	6.03	6.77	7.45	8.88
1.2	16.3	16.2	15.0	8.18	7.16	8.15	8.74	11.3
1.3			17.7	18.2		9.11	9.59	13.1
1.4	17.6	18.4	19.9	21.5	9.58	9.79	10.0	13.9
1.43		18.8						
1.45		18.6		27.3	22.0			
1.465			22.6					
1.475								
1.48				29.2				
1.49								

Differential capacity in 1 M KCl (continued)
Adsorption of C_{12}TPABr (continued)

	C/μF cm^{-2}							
	C_{12}TPABr							
c/mol l^{-1}	0	10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}	10^{-2}
E/V								
1.50		18.8	21.2	27.8		10.5	10.4	13.0
1.504					41.2			
1.52								
1.55				21.2	30.3			
1.555								
1.6	20.1	19.9	19.8	19.6	22.3	11.6	10.7	11.2
1.645						30.8		
1.65								
1.68						24.7		
1.7						22.9	11.9	10.6
1.728							27.0	
1.75							23.7	
1.80	23.5	23.3	23.5	23.4	23.5	23.3	22.4	12.5
1.85								29.2
1.867								36.0
1.88								31.9
1.90								

Differential capacity in 1 M KCl (continued)

Adsorption of C_{12}TBABr = n dodecyltributylammonium bromide

	$C/\mu F\ cm^{-2}$						
	C_{12}TBABr						
$c/mol\ l^{-1}$	0	10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}
E/V							
0	127	271	290	328	366	239	
0.1	49.5	64.9	106	90.9	90.1	101	174
0.2	39.5	41.3	41.5	40.2	42.8	45.8	49.5
0.25							
0.3	39.5	31.1	11.0	12.9	14.4	17.8	21.0
0.32		30.0					
0.35							
0.40	41.1	29.3	5.52	5.78	5.93	6.47	7.62
0.42							
0.45							
0.5	35.4	29.0					
0.6	26.1	27.3					
0.7	20.6	21.2	5.06	5.14	5.08	5.44	6.42
0.75		16.7					
0.8	18.1	14.2	5.22	5.40	5.65	6.48	7.43
1.0	16.1	13.4	5.61	6.15	6.85	7.82	9.03
1.2	16.3	16.0	6.71	7.35	8.24	9.22	10.2
1.3			16.1		8.68	9.44	10.3
1.4	17.6	18.5	21.9	8.33	8.64	9.38	9.91
1.43							
1.45		20.2	29.1				
1.465							
1.475		21.0					
1.48							
1.49			33.7				

Differential capacity in 1 M KCl (continued)
Adsorption of C_{12}TBABr (continued)

	$C/\mu F\ cm^{-2}$						
	C_{12}TBABr						
$c/mol\ l^{-1}$	0	10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}
E/V							
1.50		21.0		8.62	8.29	8.62	9.07
1.504							
1.52				34.4			
1.55			28.7				
1.555				39.4			
1.6	20.1	19.9	24.7	32.1	7.72	7.94	8.16
1.645							
1.65				26.1			
1.68							
1.7			21.9	23.5	7.27	7.16	7.26
1.728					24.5		
1.75					22.8		
1.80	23.5	23.5	23.6	23.6	22. 9	6.80	6.53
1.85						8.12	
1.867						25.0	
1.88						27.7	
1.90						35.2	20.1

Differential capacity in 1 M KCl (continued)

Adsorption of C_8PyBr = n octylpyridinium bromide

	$C/\mu F\ cm^{-2}$		
	C_8PyBr		
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}
E/V			
0	127		
0.1	49.5	61.2	65.1
0.15		50.0	56.2
0.2	39.5	46.0	51.1
0.25		42.8	45.7
0.3	39.5	38.6	40.3
0.4	41.1	29.5	29.5
0.5	35.4	20.9	20.9
0.6	26.1	14.8	14.7
0.7	20.6		
0.8	18.1	10.0	10.0
1.0	16.1	9.76	9.52
1.1			
1.15			
1.2	16.3	11.9	11.7
1.23			
1.25			
1.268			
1.3		16.0	15.0
1.32			
1.35			
1.39		18.7	
1.4	17.6		20.6
1.45		18.7	19.3
1.5		19.0	19.3
1.6	20.1	20.1	20.3
1.8	23.5		

Differential capacity in 1 M KCl (continued)

Adsorption of $C_{12}PyBr$ = n dodecylpyridinium bromide

	$C/\mu F\ cm^{-2}$							
	$C_{12}PyBr$							
$c/mol\ l^{-1}$	0	10^{-5}	3×10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}
E/V								
0	127							
0.1	49.5	58.9	66.7					
0.15				56.9	56.7	56.0	57.6	65.5
0.2	39.5	41.8	43.9	46.0	45.0	43.2	43.6	47.6
0.25		37.0						
0.3	39.5	35.2	30.2	29.4	30.4	29.4	28.3	29.4
0.4	41.1	36.4	22.4	20.2	19.6	19.8	20.1	21.2
0.5	35.4	29.0	16.6	13.7	12.6	12.9	14.2	15.9
0.6	26.1	22.0	12.0	10.3	10.0	10.3	11.5	13.5
0.7	20.6	14.7						
0.8	18.1		9.05	8.71	9.46	10.7	12.7	13.7
1.0	16.1	13.7	9.87	9.04	11.9	14.1	15.5	16.9
1.1					13.8	14.8	15.6	15.8
1.15					13.5	14.5		
1.2	16.3	17.2	17.3	16.8	16.6	16.4	16.3	15.6
1.23		18.1	20.9					
1.25		17.9	18.9	21.6	25.1			
1.268						27.9		
1.3		17.5	17.5	17.7	19.3	22.2	31.1	28.1
1.32								31.7
1.35							18.1	19.1
1.39								
1.4	17.6	18.1	18.3	17.8	18.2	18.0	18.2	18.4
1.45								
1.5								
1.6	20.1	20.0	20.3	20.5	20.2	20.2	20.3	20.6
1.8	23.5	23.6	23.9	23.5	23.8	24.0	23.8	23.9

Differential capacity in 1 M KCl (continued)

Adsorption of $C_{10}iQBr$ = n decylisoquinolinium bromide.

	$C/\mu F\ cm^{-2}$						
	$C_{10}iQBr$						
$c/mol\ l^{-1}$	0	10^{-5}	2×10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	10^{-3}
E/V							
0	127						
0.1	49.5	58.2	61.5				
0.15		46.7	49.0	53.3	60.4	80.6	62.4
0.2	39.5	40.0	40.9	44.6	48.4	52.2	57.7
0.25		37.1	35.7				83.7
0.3	39.5	35.5	31.2	31.4	31.3	33.1	41.3
0.4	41.1	34.4	24.8	21.8	21.5	21.8	23.0
0.5	35.4	30.1	19.4	15.2	14.5	14.6	15.2
0.6	26.1	22.9	14.8	11.0	10.3	10.2	10.6
0.7	20.6	17.4	12.2				
0.8	18.1	14.8	10.9	8.48	8.68	8.94	10.6
0.9					8.11	8.29	9.53
1.0	16.1	13.9	11.9	10.0	9.81	8.82	7.97
1.1			15.1	13.9	14.7	14.2	9.88
1.18							19.9
1.2	16.3	16.5	16.6	15.9	16.8	16.9	19.2
1.3				17.1	17.5	17.4	17.7
1.4	17.6	18.0	17.9	17.7	18.0	18.1	18.5
1.5		18.6	18.7	18.8			
1.6	20.1	20.0	20.2	19.6	20.2	20.0	20.1
1.8	23.5	23.3	23.7	22.9	24.1	23.6	23.9

Differential capacity in 1 M KCl (continued)

Adsorption of C_{16}TEABr = n hexadecyltriethylammonium bromide

	$C/\mu F\ cm^{-2}$							
	C_{16}TEABr							
$c/mol\ l^{-1}$	0	5×10^{-6}	10^{-5}	3×10^{-5}	10^{-4}	1.5×10^{-4}	3×10^{-4}	10^{-3}
E/V								
0	268	272	266	272				
0.1	57.0	61.7	62.9	75.0	77.2	77.3	79.4	87.1
0.15								
0.2	40.1	42.9	42.7	48.5	38.1	38.8	39.7	43.9
0.25								
0.3	38.3	37.3	36.6	30.0	26.6	26.0	26.4	27.5
0.35								
0.4	41.5	37.9	35.8	23.9	12.1	15.3	16.2	17.4
0.45		37.5	34.8					
0.5	39.8	34.9	31.9	21.0	7.70	8.06	8.20	8.54
0.6	30.9	26.3	23.9	16.4	6.77	7.10	7.17	7.27
0.7	22.9	20.0	18.4	11.8				
0.8	19.0	16.7	15.6	9.42	7.57	8.05	8.37	8.38
0.9				9.08				
1.0	16.2	15.3	14.4	9.86	8.55	9.45	10.1	10.4
1.1		15.5	15.3	12.7	9.91			
1.15								
1.18								
1.2	16.3	17.9	18.3	21.8	11.8	11.9	12.0	12.0
1.25								
1.29			21.6					
1.295				30.9				
1.3		20.1			14.0	12.4	12.2	
1.33								
1.34								
1.35			19.8	25.0	14.6			
1.38								

Differential capacity in 1 M KCl (continued)

Adsorption of C_{16}TEABr (continued)

	$C/\mu F\ cm^{-2}$							
	C_{16}TEABr							
$c/mol\ l^{-1}$	0	5×10^{-6}	10^{-5}	3×10^{-5}	10^{-4}	1.5×10^{-4}	3×10^{-4}	10^{-3}
E/V								
1.40	17.5	19.4	19.8	23.1		12.8	12.1	12.1
1.48					25.5			
1.50		19.3	19.5	20.7		13.2	11.9	
1.51								
1.53					23.0			
1.55						13.5		
1.6	19.8	20.3	20.5	20.7	21.8		11.5	11.6
1.70					22.7	22.7	11.6	
1.75						23.6		
1.8	23.4	24.1	24.1	24.3	24.5	24.7	12.7	11.9
1.82								
1.85								
1.90	36.6							

Differential capacity in 1 M KCl (continued)

Adsorption of C_{16}TMABr = n hexadecyltrimethylammonium bromide

	$C/\mu F\ cm^{-2}$							
	C_{16}TMABr							
$c/mol\ l^{-1}$	0	5×10^{-6}	10^{-5}	3×10^{-5}	10^{-4}	1.5×10^{-4}	3×10^{-4}	10^{-3}
E/V								
0	268	273	266	235				
0.1	57.0	60.9	63.5	71.4	80.3	59.5	62.3	72.8
0.15					39.9			
0.2	40.1	42.9	45.9	62.8	31.4	31.3	32.5	33.8
0.25				38.9				
0.3	38.3	36.9	35.2	29.7	22.5	23.4	24.3	25.1
0.35		37.2						
0.4	41.5	38.0	33.8	20.4	16.0	16.5	16.8	17.0
0.45		37.6	32.2					
0.5	39.8	35.1	29.6		11.3	11.4	11.5	11.6
0.6	30.9	26.4	22.3	14.8	8.67	9.27	8.33	8.44
0.7	22.9	19.9	16.8	7.63		7.12	7.09	7.18
0.8	19.0	16.5	14.1	8.03	7.13	7.50	8.47	8.51
0.9		15.1	13.3				10.1	10.3
1.0	16.2	14.7	13.0	9.01	9.92	10.0	10.0	10.3
1.1		15.2	14.3	11.9	15.8	11.0	11.4	11.9
1.15			16.7	18.4	14.9			
1.18						21.2		
1.2	16.3	18.8	21.6	28.9	15.2	14.8	16.0	16.4
1.25		20.4	24.1	36.5	16.9		18.0	
1.29								
1.295								
1.3		19.3	21.1	25.7	18.7	16.7	18.8	18.7
1.33							18.3	
1.34								
1.35							17.1	18.8
1.38								

Differential capacity in 1 M KCl (continued)

Adsorption of $C_{16}TMABr$ (continued)

	$C/\mu F\ cm^{-2}$							
	$C_{16}TMABr$							
$c/mol\ l^{-1}$	0	5×10^{-6}	10^{-5}	3×10^{-5}	10^{-4}	1.5×10^{-4}	3×10^{-4}	10^{-3}
E/V								
1.40	17.5	18.4	19.0	21.0	22.8	19.2	17.4	18.7
1.48								
1.50						22.4	17.9	
1.51								
1.53								
1.55		19.9	20.2	21.2	23.6	24.5	18.7	20.1
1.6	19.8					25.0		
1.70								
1.75								
1.8	23.4	23.5		24.0	24.6	25.2	22.7	18.6
1.82								
1.85								
1.90	36.6				36.6		35.5	29.0

Differential capacity in 1 M KCl (continued)

Adsorption of C_{16}TPABr = n hexadecyltripropylammonium bromide

	$C/\mu F\ cm^{-2}$							
	C_{16}TPABr							
c/mol l^{-1}	0	5×10^{-6}	10^{-5}	3×10^{-5}	10^{-4}	1.5×10^{-4}	3×10^{-4}	10^{-3}
E/V								
0	268	335	363					
0.1	57.0	62.4	70.5	78.4	86.4	84.6	87.5	104
0.15								
0.2	40.1	43.5	45.9	49.3	36.8	39.8	42.3	48.9
0.25								
0.3	38.3	37.4	32.7	26.6	15.8	19.0	21.3	24.8
0.35								
0.4	41.5	38.4	30.9		7.08	7.80	8.42	9.28
0.45		38.6	30.9					
0.5	39.8	36.4	29.1		5.71	5.74	5.96	6.00
0.6	30.9	27.8	21.4		5.43	5.84	6.37	6.47
0.7	22.9	20.7	16.1	11.9				
0.8	19.0	17.5	14.1	9.8	5.92	6.70	7.04	7.13
0.9				9.3				
1.0	16.2	15.7	13.8	10.6	6.76	7.79	8.86	8.75
1.1		16.4	16.2	15.9	7.80			
1.15			17.8	19.2				
1.18								
1.2	16.3	17.2	17.7		9.02	9.47	10.8	10.4
1.25			17.8	19.1				
1.29								
1.295								
1.3		18.9	20.3	24.1	11.5	10.5	11.7	
1.33								
1.34		19.6						
1.35			22.5	29.0				
1.38								

Differential capacity in 1 M KCl (continued)
Adsorption of C_{16}TPABr (continued)

	$C/\mu F\ cm^{-2}$							
	C_{16}TPABr							
$c/mol\ l^{-1}$	0	5×10^{-6}	10^{-5}	3×10^{-5}	10^{-4}	1.5×10^{-4}	3×10^{-4}	10^{-3}
E/V								
1.40	17.5	19.2	20.7	23.9	34.0	11.6	12.2	12.4
1.48								
1.50		19.6			26.7		10.9	
1.51						29.4		
1.53								
1.55								
1.6	19.8	20.7	21.0	22.0	26.4	28.9	10.3	9.94
1.70							9.97	9.40
1.75								
1.8	23.4	24.2	24.0	24.1		24.5	24.8	11.0
1.82							26.8	
1.85							36.0	21.9
1.90	36.6							60.1

Differential capacity at a dropping mercury electrode in 0.5 M H_2SO_4 at 20°C. Given are the capacity in $\mu F\ cm^{-2}$, potential in V vs SCE.

Reference: R. J. Meakins. J. Appl. Chem. 15 (1965) 416.

C_8TMABr = n-octyltrimethylammonium bromide

$C/\mu F\ cm^{-2}$

	C_8TMABr			
$c/mol\ l^{-1}$	0	10^{-3}	3×10^{-8}	10^{-2}
E/V				
0.0	31.2	127	253	388
0.1	29.1	49	91	159
0.2	30.7	47	61	71
0.3	33.5	40	46	51
0.4	32.7	24	28	33
0.5	27.2	13.9	15.5	18.5
0.6	22.2	10.2	10.5	11.4
0.7	19.0	9.1	8.9	9.2
0.8	17.3	8.8	8.6	8.5
0.9	16.3	9.0	8.8	8.5
1.0	15.6	9.5	9.2	8.9
1.1	15.4	10.3	9.9	9.5
1.2	15.7	11.6	11.2	10.7
1.3	15.9	13.1	12.7	12.4
1.4		14.8	14.3	14.0

Differential capacity in 0.5 M H_2SO_4 (continued)

C_{12}TMABr = n-dodecyltrimethylammonium bromide

	$C/\mu F\ cm^{-2}$			
	C_{12}TMABr			
$c/mol\ l^{-1}$	0	10^{-4}	3×10^{-4}	10^{-3}
E/V				
0.0	31.2	39	67	97
0.1	29.1	34	37	49
0.2	30.7	23	24	31
0.3	33.5	14	14.1	18.1
0.4	32.7	9.3	9.3	10.5
0.5	27.2	8.1	8.2	9.0
0.6	22.2	7.6	8.2	8.8
0.7	19.0	7.7	8.3	8.7
0.8	17.3	7.8	8.4	8.7
0.9	16.3	8.1	8.6	9.1
1.0	15.6	8.7	9.1	10.4
1.1	15.4	9.8	10.4	12.2
1.2	15.7	12.8	13.6	14.3
1.3	15.9	17.8	17,3	16.0
1.4		22.0	18.8	17.0

Differential capacity in 0.5 M H_2SO_4 (continued)

C_{16}TMABr = n-hexadecyltrimethylammonium bromide

	$C/\mu F\ cm^{-2}$			
	C_{16}TMABr			
$c/mol\ l^{-1}$	0	10^{-4}	3×10^{-4}	10^{-3}
E/V				
0.0	31.2	63	75	
0.1	29.1	29	31	45
0.2	30.7	19.2	22	28
0.3	33.5	13.6	16.1	20.8
0.4	32.7	10.1	11.2	14.4
0.5	27.2	8.3	8.1	9.4
0.6	22.2	7.6	6.8	7.1
0.7	19.0	6.9	6.7	6.8
0.8	17.3	7.1	8.5	8.8
0.9	16.3	8.2	10.0	10.3
1.0	15.6	10.4	10.3	10.5
1.1	15.4	13.8	12.1	12.3
1.2	15.7	15.9		16.8
1.3	15.9	17.4	16.5	17.0
1.4		18.3	17.1	17.5

Differential capacity in 0.5 M H_2SO_4 (continued)

$(C_{12}TMA)_2SO_4$ = n-dodecyltrimethylammonium sulphate

	$C/\mu F\ cm^{-2}$			
	$(C_{12}TMA)_2SO_4$			
$c/mol\ l^{-1}$	0	10^{-4}	10^{-3}	10^{-2}
E/V				
0.0	31.2	32.7	56	47
0.1	29.1	31.3	29.8	25.8
0.2	30.7	19.6	13.4	14.1
0.3	33.5	12.2	8.9	9.9
0.4	32.7	9.0	8.5	9.5
0.5	27.2	8.0	8.6	
0.6	22.2	7.7	8.8	9.2
0.7	19.0	7.7	8.8	8.5
0.8	17.3	7.9	8.8	8.8
0.9	16.3	8.2	9.3	10.6
1.0	15.6	9.0	10.8	12.1
1.1	15.4	10.2	12.5	13.3
1.2	15.7	13.5	14.6	14.7
1.3	15.9	19.3	16.4	16.1
1.4		25.3	17.4	17.2
1.5		29.5	20.1	20.1

Differential capacity on mercury electrode in 0.5 M H_2SO_4.

Adsorption of C_8TBABr. Reference electrode: saturated calomel electrode.

C_8TBABr = n octyltributylammonium bromide.

Reference: R.J. Meakins, unpublished results.

				$C/\mu F\ cm^{-2}$			
				C_8TBABr			
$c/mol\ l^{-1}$	0	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}
E/V							
+0.1						38.5	100
0	29.7	30.9	30.9	33.7	43.4	33.0	77.6
0.05				30.9			
0.1	28.2	26.0	28.3	25.4	19.2	25.5	48.8
0.115							
0.15		23.7	24.2				
0.16							
0.19							
0.2	29.7	22.4	21.6	10.1	13.4	18.1	31.1
0.3	32.9	21.0	19.7	6.02	6.68	7.39	9.57
0.35	33.5						
0.4	32.9	19.5	16.7	4.96	5.06	5.10	5.47
0.5	27.3	15.1	13.6				
0.6	22.3	12.7	11.2	4.72	4.83	4.88	5.13
0.7	18.9	11.1					
0.8	17.1	10.4	9.28	4.81	4.91	5.04	5.29
1.0	15.3	10.5	9.42	5.15	5.14	5.30	5.55
1.2	15.4	11.5	10.0	6.03	5.66	5.79	6.11

Differential capacity in 0.5 M H_2SO_4 (continued)
Adsorption of C_8TEABr = n-octyltriethylammonium bromide.

	$C/\mu F\ cm^{-2}$							
	C_8TEABr							
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
E/V								
+0.1								
0	29.7	40.5	65.1					
0.05								
0.1	28.2	32.3	42.3	57.6	75.8		160	
0.115								
0.15				51.5				
0.16			41.4					
0.19		31.8						
0.2	29.7		39.2	46.4	55.0	67.7	77.6	87.2
0.3	32.9	21.0	27.5	32.3	38.5	46.6	52.3	57.0
0.35	33.5							
0.4	32.9	10.8	13.3	15.3	17.9	24.6	27.9	31.8
0.5	27.3	7.31	7.42	7.74	8.01	9.74	10.7	12.9
0.6	22.3	6.39	6.32	6.32	6.28	6.50	6.71	7.37
0.7	18.9							
0.8	17.1	6.16	5.98	5.99	6.00	6.10	6.25	6.56
1.0	15.3	6.78	6.30	6.34	6.29	6.44	6.58	6.92
1.2	15.4	8.22	7.15	7.31	7.15	7.25	7.53	7.67
1.3						8.40	8.50	

Differential capacity in 0.5 M H_2SO_4 (continued)
Adsorption of C_8TPABr = n-octyltripropylammonium bromide.

	C/μF cm^{-2}							
					C_8TPABr			
c/mol l^{-1}	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
E/V								
+0.1								
0	29.7	44.8	67.1		175	254		
0.05		38.5	53.3	88.6				
0.1	28.2		45.9	65.3	87.0	145		
0.115		37.3						
0.15			39.5					
0.16								
0.19		23.3	32.6	43.2	53.7	70.8	73.4	
0.2	29.7	11.8	15.8	19.5	22.2	31.0	40.8	
0.3	32.9							
0.35	33.5	6.59	6.93	7.36	7.44	9.31	12.1	15.4
0.4	32.9							6.80
0.5	27.3	5.38	5.37	5.43	5.52	5.79	5.98	5.79
0.6	22.3							
0.7	18.9	5.41	5.47	5.52	5.66	5.88	6.01	6.05
0.8	17.1	5.71	5.71	5.86	5.98	6.22	6.37	6.82
1.0	15.3	6.58	6.45	6.52	6.48	6.79	7.34	7.65

Differential capacity in 0.5 M H_2SO_4 (continued)
Adsorption of $(C_8TBA)_2SO_4$ = n octyltributylammonium bromide

	$C/\mu F\ cm^{-2}$						
	$(C_8TBA)_2SO_4$						
$c/mol\ l^{-1}$	0	5×10^{-5}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}
E/V							
+0.083							114.6
+0.045						85.9	
+0.014					71.1		
0	29.7	31.0	31.9	40.3			
0.023				49.0			
0.05			31.6				
0.057		30.3					
	28.2	22.0	12.5	8.28	7.56	7.37	6.19
0.2	29.7	6.33	5.82	5.66	5.76	5.94	5.85
0.3	32.9						
0.35	33.5						
0.4	32.9	4.96	5.01	5.17	5.39	5.57	5.56
0.5	27.3						
0.6	22.3	4.80	4.94	5.19	5.45	5.52	5.63
0.7	18.9						
0.8	17.1	4.93	5.09	5.34	5.57	5.76	6.36
1.0	15.3	5.21	5.35	5.60	5.96	6.43	7.49
1.2	15.4	5.89	5.94	6.17	6.55	7.43	8.43

Differential capacity in 0.5 M H_2SO_4 (continued)
Adsorption of $(C_8TEA)_2SO_4$ = n octyltriethylammonium sulfate.

	$C/\mu F\ cm^{-2}$							
	$(C_8TEA)_2SO_4$							
$c/mol\ l^{-1}$	0	$1x10^{-4}$	$2.5x10^{-4}$	$5x10^{-4}$	$1x10^{-3}$	$2.5x10^{-3}$	$5x10^{-2}$	$1.5x10^{-2}$
E/V								
+0.05								48.0
+0.03								
0	29.7	30.5	31.1	31.1	32.3	35.8	44.3	
0.01								86.0
0.023								
0.05		29.7				55.6		
0.057							85.9	
0.074								
0.095						83.9		
0.10	28.2	29.7	32.8	37.6	54.8		63.9	31.6
0.12								
0.14					66.5			
0.15						59.8		
0.155								
0.165				50.4				
0.185			39.6					
0.2	29.7			45.9	48.7	38.2	22.0	11.8
0.21		32.8						
0.23								
0.25			30.4					
0.28								
0.3	32.9	20.4	20.0	18.5	17.5	13.3	10.5	
0.35	33.5							
0.4	32.9	10.3	9.52	9.16	8.85	7.94	7.51	7.31
0.5	27.3	7.38	7.08	6.99				
0.6	22.3	6.51	6.37	6.36	6.30	6.36	6.40	6.85
0.7	18.9							
0.8	17.1	6.14	6.10	6.06	6.14	6.20	6.34	6.79
1.0	15.3	6.39	6.40	6.38	6.42	6.54	6.63	7.18
1.2	15.4	7.29	7.32	7.08	7.12	7.12	7.16	7.89

Differential capacity in 0.5 M H_2SO_4 (continued)

Adsorption of $(C_8TMA)_2SO_4$ = n octyltrimethylammonium sulfate

				C/μF cm^{-2}				
				$(C_8TMA)_2SO_4$				
c/mol l^{-1}	0	1x10^{-4}	2.5x10^{-4}	5x10^{-4}	1x10^{-3}	2.5x10^{-3}	5x10^{-3}	1.5x 10^{-2}
E/V								
0	29.7	30.1	30.2	30.7	31.0	31.4	34.0	52.2
0.01								
0.023								
0.05								
0.057								72.2
0.074							50.4	
0.095								
0.10	28.2	28.7	29.6	30.3	32.7	44.4		51.5
0.12							61.1	
0.14								
0.15					40.9			
0.155						62.1		
0.165								
0.185								
0.2	29.7	31.3	34.3	39.8	49.3	50.1	39.0	23.7
0.21								
0.23				41.6				
0.25			36.2		41.8			
0.28		33.6						
0.3	32.9	33.0	32.5	33.0	30.4	25.1	20.5	13.1
0.35	33.5							
0.4	32.9	22.9	20.2	18.9	17.1	14.2	12.3	9.29
0.5	27.3	14.9	13.1	12.3	11.3	10.2	9.41	
0.6	22.3	11.3	10.2	9.86	9.13	8.80	8.41	8.09
0.7	18.9							
0.8	17.1	9.51	8.91	8.62	8.35	8.27	8.12	8.31
1.0	15.3	10.1	9.64	9.25	8.94	8.89	8.84	8.79
1.2	15.4	11.9	11.5	11.4	11.2	10.7	10.6	10.5
1.3			13.4	13.1	13.1	12.5	12.4	

Differential capacity in 0.5 M H_2SO_4 (continued)

Adsorption of $(C_8TPA)_2SO_4$ = n octyltripropylammonium sulfate

				$C/\mu F\ cm^{-2}$				
				$(C_8TPA)_2SO_4$				
$c/mol\ l^{-1}$	0	$1x10^{-4}$	$2.5x10^{-4}$	$5x10^{-4}$	$1x10^{-3}$	$2.5x10^{-3}$	$5x10^{-3}$	$1.5x10^{-2}$
E/V								
+0.014								
+0.07								129
+0.05								
+0.03							96.9	
0	29.7	30.7	31.3	33.5	37.1	82.6	36.7	74.3
0.01					68.4			
0.023						29.1		
0.05								
0.057								
0.074			56.5					
0.095								
0.10	28.2	32.2	40.3	42.8	15.9	11.7	9.09	8.66
0.12		32.6						
0.2	29.7	16.4	12.9	8.98	7.02	6.76	6.45	6.61
0.3	32.9	7.52	6.81	6.21	5.99	6.01		
0.35	33.5							
0.4	32.9	6.05	5.89	5.78	5.67	5.88	5.99	6.18
0.5	27.3							
0.6	22.3	5.47	5.48	5.48	5.61	5.88	6.05	6.10
0.7	18.9							
0.8	17.1	5.47	5.53	5.63	5.82	6.02	6.15	6.59
1.0	15.3	5.75	5.75	5.94	6.05	6.38	6.43	7.37
1.2	15.4	6.61	6.48	6.55	6.62	6.87	7.34	8.37

Differential capacity in 0.5 M H_2SO_4 (continued)
Adsorption of $(C_{16}TMA)_2SO_4$= n hexadecyltrimethylammonium sulfate

	$C/\mu F\ cm^{-2}$							
	$(C_{16}TMA)_2SO_4$							
$c/mol\ l^{-1}$	0	$2.5x10^{-5}$	$3.5x10^{-5}$	$5x10^{-5}$	$7.5x10^{-5}$	$1x10^{-4}$	$3x10^{-4}$	$5x10^{-4}$
E/V								
+0.09		53.8	61.1			103	109	116
+0.08					79.3			
+0.06				64.1				
0	29.7	46.3		26.3	23.0	22.6	23.4	21.8
0.05			23.8					
0.1	28.2	35.5		16.6	15.9	16.4	16.6	16.6
0.15		14.7	20.1					
0.2	29.7	9.08	11.0	13.2	13.0	13.3	13.8	13.6
0.3	32.9							
0.35	33.5							
0.4	32.9	7.92	8.20	8.89	8.80	8.80	8.89	9.05
0.5	27.3							
0.6	22.3	8.15	8.09	7.05	6.67	6.29	6.37	6.36
0.7	18.9							
0.8	17.1	7.80	7.34	7.03	7.28	8.30	8.22	8.25
1.0	15.3	9.14	10.1	10.2	9.99	9.92	10.1	10.0
1.1				13.8	11.4			
1.13					19.8			
1.15					15.8			
1.2	15.4	18.2	16.3	14.7	14.5	16.2	16.3	16.4
1.25		22.6	18.1					
1.255						18.1		
1.275							18.3	
1.28						16.7		
1.3		21.6	18.1	17.1			17.1	18.3
1.35						17.0		17.4
1.4			19.4	18.7	17.8	17.2	17.3	16.9

Differential capacity on mercury electrode in 0.5 M Na_2SO_4 .

Adsorption of $(C_{16}TMA)_2SO_4$. Reference electrode: saturated calomel electrode.
$(C_{16}TMA)_2SO_4$ = n hexadecyltrimethylammonium sulfate.
Reference: R.J. Meakins, unpublished results.

	$C/\mu F\ cm^{-2}$							
	$(C_{16}TMA)_2SO_4$							
$c/mol\ l^{-1}$	0	$2.5x10^{-5}$	$3.5x10^{-5}$	$5x10^{-5}$	$7.5x10^{-5}$	$1.5x10^{-4}$	$2.5x10^{-4}$	$3x10^{-4}$
E/V								
+0.075								121
+0.065						109		
+0.06					90.0		113	
0	39.2	61.2	58.3	56.6	31.8	32.4	30.6	31.8
0.028		86.5						
0.035			81.4	78.6	28.5			
0.1	36.8	28.2	15.2	15.7	27.2	23.0	22.1	22.5
0.2	36.0	17.0	7.40	7.09	8.70	13.3	13.2	13.7
0.3	31.0	9.10				9.04	8.87	
0.4	26.0	7.34	7.45	7.93	8.35	7.15	6.83	7.01
0.5	22.4							
0.6	20.0	6.83	7.86	7.88	6.52	6.06	6.00	5.87
0.7	17.8							
0.8	16.4	7.04	7.69	7.42	7.24	8.57	8.68	8.52
1.0	15.1	8.76	9.01	10.3	14.8	11.5	11.1	11.2
1.015								17.8
1.05						14.0	14.2	14.6
1.1		12.2	12.0	13.4	13.3	13.3	13.4	
1.2	15.8	22.3	20.3	19.6	14.8	15.0	15.1	14.9
1.255			29.3	29.4				
1.26		30.9						
1.3		26.6	22.8	18.9				
1.4	17.1	22.4	22.6	21.9	19.0	17.0	16.9	16.7
1.6	18.5	20.1	21.0	22.3	22.8	18.9	18.3	18.7
1.7		20.9						
1.8	20.6	21.7	21.6	22.2	22.7	20.3	18.1	17.8
1.9	38.5	43.7	45.1	44.9	46.5	47.3	47.8	48.4

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_6TMA)_2SO_4$= hexatrimethylammonium sulfate

$C/\mu F\ cm^{-2}$

	$(C_6TMA)_2SO_4$							
$c/mol\ l^{-1}$	0	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
E/V								
0	39.2	38.3	38.1	38.6	38.6	38.7	38.9	42.7
0.1	36.8	38.9	39.1	38.4	40.3	41.3	43.6	65.1
0.14								77.3
0.15		38.9						
0.18			39.8					
0.2	36.0	39.0			42.7	49.4	60.7	50.7
0.21				40.3			61.9	
0.235						53.0		
0.25					43.6			
0.3	31.0	33.8	34.8	37.4	40.6	42.4	34.6	19.9
0.4	26.0	27.6	26.9	26.9	26.3	22.2	17.6	12.6
0.5	22.4	21.4	19.6	18.1	16.6	14.2	12.2	
0.6	20.0	16.6	14.7	13.6	12.5	11.1	10.1	8.87
0.7	17.8							
0.8	16.4	12.6	11.4	10.7	10.1	9.31	8.94	8.43
1.0	15.1	12.3	11.3	10.7	10.2	9.78	9.28	9.20
1.2	15.8	14.5	13.3	12.6	11.9	11.2	10.9	10.4
1.4	17.1	17.3	17.7	17.9	17.5	15.5	14.1	13.3
1.43			18.2					
1.47				19.7				
1.5		18.1	18.8	19.9	22.3	22.6	18.5	
1.57						28.1		
1.6	18.5	18.8	19.2	19.9	21.7	27.7	31.8	19.4
1.63							34.6	
1.7			19.9	19.9	20.4	22.1	27.1	38.7
1.74								49.9
1.8	20.6					21.7	22.7	31.8
1.9	38.5							

Differential capacity in 0.5 M Na_2SO_4 (continued)

Adsorption of $(C_{10}TMA)_2SO_4$ = n decyltrimethylammonium sulfate

	$C/\mu F\ cm^{-2}$							
	$(C_{10}TMA)_2SO_4$							
$c/mol\ l^{-1}$	0	$2.5x10^{-5}$	$5x10^{-5}$	$1x10^{-4}$	$2.5x10^{-4}$	$5x10^{-4}$	$1.5x10^{-3}$	$5x10^{-3}$
E/V								
+0.08								113
0	39.2	38.8	38.5	38.8	39.9	43.6	91.3	66.9
0.015							95.3	
0.065						63.2		
0.085					50.1			
0.1	36.8	39.9				53.1	43.2	26.6
0.11				43.4				
0.12			40.9					
0.15		40.1						
0.2	36.0	35.1	30.0	28.1	24.3	22.1	16.1	12.5
0.3	31.0	17.9	15.2	14.2	12.5	11.4	9.41	
0.4	26.0	11.7	10.3	9.84	9.19	8.72	8.00	8.07
0.5	22.4							
0.6	20.0	8.50	7.83	7.78	7.58	7.56	7.60	8.44
0.7	17.8							
0.8	16.4	8.10	7.73	7.61	7.71	7.74	8.13	9.05
1.0	15.1	8.87	8.44	8.46	8.50	8.51	8.92	11.0
1.2	15.8	11.1	10.7	10.7	10.9	11.2	12.2	14.0
1.3		14.1	13.3	13.3	13.8	13.9	15.8	15.4
1.4	17.1	17.6	18.4	19.0	18.9	18.9	17.9	16.3
1.43			19.6					
1.46				22.1				
1.5		18.3	18.6	20.5	25.7	22.3	18.5	
1.53					26.7	25.9		
1.6	18.5	18.8	19.1	19.2	21.1		19.1	17.9
1.63						26.3		
1.7		19.7			19.9	21.4	20.8	
1.8	20.6					21.0	26.7	18.6
1.85							28.5	
1.9	38.5						51.1	60.5

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_{12}TMA)_2SO_4$ = n dodecyltrimethylammonium sulfate

	$C/\mu F\ cm^{-2}$							
	$(C_{12}TMA)_2SO_4$							
$c/mol\ l^{-1}$	0	1.5×10^{-5}	2.5×10^{-5}	5×10^{-5}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	5×10^{-3}
E/V								
+0.11								168
+0.09							91.0	
+0.06		42.3						
+0.05			40.6					
+0.04						77.3		
0	39.2	36.9	38.9	42.1	73.5	67.7	59.1	57.1
0.03				48.0				
0.05		34.6			46.4	46.7		
0.055			39.4					
0.1	36.8	31.7	34.3	31.2	25.9	23.6	25.1	22.7
0.2	36.0	22.7	16.4	13.6	10.5	9.88	10.3	9.40
0.3	31.0	9.72	9.60	8.57				
0.4	26.0	7.73	7.72	7.47	7.55	7.94	8.14	8.48
0.5	22.4	6.94						
0.6	20.0	6.82	7.08	7.43	8.02	8.16	8.04	8.18
0.7	17.8							
0.8	16.4	7.06	7.38	7.89	8.20	8.28	8.57	8.61
1.0	15.1	8.09	8.30	8.72	9.69	10.7	11.0	11.1
1.2	15.8	12.0	12.4	13.1	13.9	13.8	13.7	14.2
1.3		18.8	19.2	18.4	15.9	15.1		
1.4	17.1	25.4	27.9	20.3	16.7	16.1	16.5	16.2
1.45			26.0					
1.5		23.3	23.2	22.4	17.7	17.0	16.9	17.0
1.565				24.7				
1.6	18.5	19.7	19.5	23.7	19.2	17.7	17.6	17.7
1.7				20.3	23.6	18.4		
1.745					24.9			
1.8	20.6				24.0	19.2	18.1	18.5
1.85					27.1	22.3		
1.9	38.5					61.8	59.6	58.8

Differential capacity in 0.5 M Na_2SO_4 (continued)

Adsorption of $(C_{14}TMA)_2SO_4$ = n tetradecyltrimethylammonium sulfate

	C/μF cm^{-2}							
	$(C_{14}TMA)_2SO_4$							
c/mol l^{-1}	0	2.5x10^{-5}	3.5x10^{-5}	5x10^{-5}	1x10^{-4}	2.5x10^{-4}	5x10^{-4}	1x10^{-3}
E/V								
+0.09					81.7			
+0.06						79.0		
+0.05							75.4	
+0.045								73.2
+0.035				68.6				
0	39.2	71.3	70.3	58.1	43.7	38.4	37.1	36.9
0.005		74.1						
0.015					39.2			
0.03								33.2
0.04					41.2			
0.05		39.9	42.7			33.2	32.7	
0.06								35.1
0.1	36.8	22.1	16.0	16.1	24.4	30.0	24.7	24.2
0.15						19.9	14.5	
0.2	36.0	9.1	7.60	7.39	8.85	12.9	10.8	9.84
0.3	31.0							
0.4	26.0	7.24	9.57	7.88	8.14	8.31	8.10	8.02
0.5	22.4							
0.6	20.0	7.83	8.00	7.88	6.96	6.66	6.60	6.64
0.7	17.8							
0.8	16.4	7.86	8.00	7.69	8.09	8.72	9.22	8.80
1.0	15.1	8.93	9.77	10.5	11.2	11.5	11.8	11.6
1.1		11.1	12.1	12.7				
1.2	15.8	15.9	15.4	14.2	14.1	14.0	14.1	14.1
1.3		22.7	17.0	15.8				
1.35		22.5						
1.4	17.1	22.6	18.3	17.1	16.3	16.0	16.1	16.2
1.5		23.3	22.0	18.3	17.1	16.9	17.0	16.8
1.51		23.5						
1.6	18.5	20.9	22.4	20.4	18.0	17.8	17.9	17.6
1.7			22.5					
1.8	20.6	22.1	22.4	22.8	19.0	17.4	17.8	17.5
1.9	38.5	43.8	44.4	44.4	38.5	35.4	35.7	35.4

Differential capacity in 0.5 M Na_2SO_4 (continued)

Adsorption of $(C_8TMA)_2SO_4$ = n octyltrimethylammonium sulfate

	$C/\mu F\ cm^{-2}$							
	$(C_8TMA)_2SO_4$							
$c/mol\ l^{-1}$	0	$1x10^{-4}$	$2.5x10^{-4}$	$5x10^{-4}$	10^{-3}	$2.5x10^{-3}$	$5x10^{-3}$	$1.5x10^{-2}$
E/V								
+0.08								
+0.06								
+0.04								
0	39.2	38.5	38.4	38.8	39.0	40.6	43.0	63.7
0.027								
0.037								84.9
0.05								
0.06								
0.094							89.4	
0.1	36.8	39.6	41.0	42.7	48.5	74.0		52.1
0.125						84.3		
0.15						73.1	56.6	30.0
0.155					64.9			
0.17				52.2				
0.18			45.8					
0.19		40.9						
0.2	36.0			48.8	51.6	42.9	31.9	18.2
0.25		35.9	35.0					
0.3	31.0	27.7	24.5	23.2	21.2	17.1	14.6	11.0
0.4	26.0	16.1	14.3	13.4	12.6	11.1	10.2	8.84
0.5	22.4	11.8	10.7	10.1	9.70	9.05		
0.6	20.0	10.0	9.27	8.81	8.56	8.23	8.11	7.96
0.7	17.8	9.14						
0.8	16.4	9.04	8.49	8.27	8.12	7.99	8.04	8.18
1.0	15.1	9.57	9.27	9.04	8.77	8.68	8.76	8.84
1.2	15.8	11.5	11.0	10.8	10.6	10.4	10.4	10.6
1.3		13.8						

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_8TMA)_2SO_4$ (continued)

	$C/\mu F\ cm^{-2}$							
	$(C_8TMA)_2SO_4$							
$c/mol\ l^{-1}$	0	$1x10^{-4}$	$2.5x10^{-4}$	$5x10^{-4}$	10^{-3}	$2.5x10^{-3}$	$5x10^{-3}$	$1.5x10^{-2}$
E/V								
1.4	17.1	17.4	16.5	16.0	15.2	14.5	14.3	14.8
1.45		18.3	19.0					
1.49			20.3					
1.5		18.9		22.5	21.8	18.4		17.6
1.53				23.7				
1.55			20.1					
1.575					28.3			
1.59								
1.6	18.5	19.1	19.7	21.6	27.7	28.9	22.2	19.8
1.635								
1.65						38.0	27.8	
1.68								
1.70					20.7	23.1	39.4	21.8
1.726							43.7	
1.74								
1.75								
1.80	20.6	21.0	20.8	21.4	21.1	21.7	24.4	29.2
1.82								
1.875								59.0
1.89								
1.90	38.5							53.0

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_8TPA)_2SO_4$ = n octyltripropylammonium sulfate

	$C/\mu F\ cm^{-2}$							
	$(C_8TPA)_2SO_4$							
c/mol l^{-1}	0	10^{-4}	2.5×10^{-4}	5×10^{-4}	10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
E/V								
+0.08								170
+0.06							131	
+0.04						117		
0	39.2	40.3	49.0	55.5	106	14.4	15.2	12.5
0.027				77.3				
0.037			52.9					
0.05				34.7				
0.06		40.8						
0.094								
0.1	36.8	21.4	9.88	14.0	9.04	8.22	7.72	6.79
0.125								
0.15								
0.155								
0.17								
0.18								
0.19								
0.2	36.0	7.13	6.16	5.92	5.99	5.97	6.04	6.09
0.25								
0.3	31.0							
0.4	26.0	5.33	5.31	5.50	5.41	5.65	5.87	5.99
0.5	22.4							
0.6	20.0	5.11	5.35	5.64	5.55	5.81	6.02	6.20
0.7	17.8							
0.8	16.4	5.33	5.69	5.84	5.88	6.02	6.25	6.67
1.0	15.1	5.66	5.96	6.10	6.16	6.45	6.95	7.47
1.2	15.8	6.26	6.36	6.50	6.73	7.34	8.05	8.43
1.3			6.77	6.87	7.29	8.05	8.63	9.16

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_8TPA)_2SO_4$ (continued)

	$C/\mu F\ cm^{-2}$							
	$(C_8TPA)_2SO_4$							
$c/mol\ l^{-1}$	0	10^{-4}	2.5×10^{-4}	5×10^{-4}	10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
E/V								
1.4	17.1	8.04	7.63	7.62	7.98	8.95	9.39	9.94
1.45								
1.49								
1.5		11.1	9.31	8.92	8.98	9.87	10.1	10.6
1.53								
1.55		20.3	11.1					
1.575								
1.59			25.9					
1.6	18.5	19.2		11.9	10.8	10.8	11.1	11.1
1.635				36.7				
1.65				24.6	13.3			
1.68					48.8			
1.70		19.4	19.4	19.6	24.6	12.6	11.4	11.4
1.726								
1.74						61.8		
1.75							11.2	
1.80	20.6				21.3	21.6	65.2	8.84
1.82							30.0	
1.875								
1.89								36.4
1.90	38.5							35.6

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_8TBA)_2SO_4$ = n octyltributylammonium sulfate

	C/µF cm^{-2}					
	$(C_8TBA)_2SO_4$					
c/mol l^{-1}	0	1 x 10^{-4}	2.5 x 10^{-4}	5 x 10^{-4}	1 x 10^{-3}	2.5 x 10^{-3}
E/V						
+0.1						173
+0.07					137	
+0.04				97.0		
+0.02			83.7			
+0.01		50.6				
0	39.2	45.5		10.1		
0.037						
0.04						
0.05		8.78			8.08	7.79
0.07						
0.094						
0.1	36.8	7.57	5.99	6.15	6.61	6.33
0.117						
0.125						
0.13						
0.14						
0.15						
0.2	36.0	5.25	5.34	5.49	5.64	5.83
0.3	31.0					
0.4	26.0	4.90	5.14	5.33	5.62	5.72
0.5	22.4					
0.6	20.0	4.98	5.23	5.41	5.66	5.94
0.7	17.8					
0.8	16.4	5.05	5.32	5.63	5.85	6.51
1.0	15.1	5.28	5.53	5.81	6.46	7.20

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_8TBA)_2SO_4$ (continued)

	$C/\mu F\ cm^{-2}$					
	$(C_8TBA)_2SO_4$					
$c/mol\ l^{-1}$	0	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}
E/V						
1.2	15.8	5.70	5.98	6.44	7.24	7.77
1.3		5.91	6.32	6.83	7.56	8.15
1.4	17.1	6.32	6.57	7.19	7.67	8.20
1.5		7.08	7.00	7.37	7.94	8.08
1.55		7.66				
1.6	18.5		7.63	7.65	8.06	7.51
1.161		29.7				
1.638						
1.65		21.3	8.09			
1.672			37.4			
1.685						
1.7		19.9	22.1	8.27	8.18	6.61
1.718				36.1		
1.75			20.7	21.7	8.51	
1.767					33.2	
1.795						
1.80	20.6			21.2	23.9	6.02
1.833						5.97
1.87						
1.884						29.3
1.9	38.5					

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_8TEA)_2SO_4$ = n octyltriethylammonium sulfate

	$C/\mu F\ cm^{-2}$							
	$(C_8TEA)_2SO_4$							
$c/mol\ l^{-1}$	0	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
E/V								
+0.1								
+0.07								
+0.04								
+0.02								
+0.01								
0	39.2	38.4	38.6	39.0	40.3	44.1	56.5	121
0.037								
0.04							11.7	
0.05								42.0
0.07						113		
0.094								
0.1	36.8	41.7	47.6	62.4	91.0	85.8	53.2	20.1
0.117				66.0				
0.125								
0.13			51.6					
0.14		43.8						
0.15				57.3	56.0			
0.2	36.0	31.8	30.6	29.9	24.1	16.9	12.2	9.32
0.3	31.0	12.3	11.7	10.8	9.73	8.53	7.70	7.19
0.4	26.0	8.06	7.78	7.45	7.15	6.74	6.56	6.70
0.5	22.4							
0.6	20.0	6.20	6.16	6.00	6.03	6.00	6.13	6.45
0.7	17.8							
0.8	16.4	5.94	5.95	5.95	6.03	6.09	6.19	6.55
1.0	15.1	6.39	6.34	6.33	6.37	6.49	6.83	7.08

Differential capacity in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_8TEA)_2SO_4$ (continued)

	C/μF cm^{-2}							
	$(C_8TEA)_2SO_4$							
c/mol l^{-1}	0	1 x 10^{-4}	2.5 x10^{-4}	5 x 10^{-4}	10^{-3}	2.5 x10^{-3}	5 x 10^{-3}	1.5 x10^{-2}
E/V								
1.2	15.8	7.82	7.44	7.34	7.37	7.51	7.72	7.99
1.3		8.91						
1.4	17.1	11.1	10.3	9.69	9.59	9.39	9.58	10.0
1.5		15.8	13.7	12.6			11.4	11.3
1.55		19.6	17.0					
1.6	18.5	19.9	22.7	20.2	16.2	14.3	13.7	13.4
1.61								
1.638				28.0				
1.65					23.7			
1.672								
1.685					39.2			
1.7			19.8	21.0	31.5	22.4	17.4	14.9
1.718								
1.75					21.2	68.6	24.0	
1.767								
1.795							99.4	
1.80	20.6	21.1	20.8	21.0		22.6		18.5
1.833								
1.85							24.3	
1.87								
1.884								17.1
1.9								67.0

Electrocapillary data on mercury in aqueous KCl. T = 25°

Potential measured with respect to sat. calomel electrode in contact with working solution.

Given is the interfacial tension x 10^3 in N.m^{-1}:

Reference: M.A.V. Devanathan and M. Peries, Trans. Faraday Soc. 50, 1236 (1954).

c/mole l^{-1}	0.01	0.03	0.1	0.3	1.0	3.0
$\frac{E}{\text{Volts}}$						
-0.0	368.6	369.8	361.6	359.2	353.3	337.9
-0.05	378.9	381.5	376.6	374.3	368.8	359.7
-0.1	388.0	391.0	387.2	385.3	381.3	373.6
-0.15	396.2	398.7	395.8	394.3	390.7	384.0
-0.2	402.8	405.3	402.7	401.6	398.7	392.7
-0.25	408.6	410.9	408.9	407.7	405.4	400.2
-0.3	413.3	415.7	414.2	413.2	411.1	406.2
-0.35	417.3	419.4	418.7	417.7	415.9	411.1
-0.4	420.5	422.2	422.0	421.2	419.7	415.1
-0.45	423.3	424.2	424.3	423.5	422.3	418.2
-0.5	425.3	425.7	425.6	424.8	423.7	420.4
-0.55	426.4	426.3	426.2	425.2	424.2	421.7
-0.6	426.8	426.4	425.8	424.8	423.7	421.8
-0.65	426.7	425.8	425.0	423.7	422.3	420.7
-0.7	426.0	424.7	423.6	422.0	420.3	418.8
-0.75	424.6	422.9	421.6	419.5	417.9	416.1
-0.8	422.8	420.7	419.0	416.4	414.7	412.7
-0.85	420.6	417.8	415.9	412.9	410.8	409.0
-0.9	417.8	414.4	412.6	408.9	406.5	405.7
-0.95	414.8	410.7	408.7	404.7	401.8	399.7
-1.00	411.2	406.7	404.4	400.1	396.7	394.3
-1.05	407.4	402.3	399.5	395.0	391.2	388.5
-1.1	403.2	397.6	394.4	389.6	385.3	382.2
-1.15	398.6	392.5	388.9	383.7	318.7	375.2
-1.2	393.4	387.3	383.1	377.5	371.7	367.8
-1.25	387.9	381.5	376.7	370.6	364.2	359.9
-1.3	382.2	275.4	370.0	363.3	356.5	351.8
-1.35	375.9	368.4	362.8	355.7	349.4	343.9
-1.4	369.0	360.7	355.1		339.8	335.4

Electrocapillary data on mercury in aqueous KBr. T = 25°

Potential measured with respect to sat. $Hg/Hg_2Br_2/$ electrode in contact with working solution.

Given is the interfacial tension x 10^3 in $N.m^{-1}$

Reference: M.A.V. Devanathan and P. Peries, Trans. Faraday Soc. 50, 1236 (1954)

c/mole l^{-1}	0.01	0.03	0.1	0.3	1.0	3.0
E/Volts						
-0.0	365.9	362.7	-	332.7		
-0.05	381.4	379.7	370.2	363.4	351.6	340.3
-0.1	393.4	391.4	385.7	379.7	368.5	359.1
-0.15	401.6	400.8	395.8	390.7	381.8	372.2
-0.2	408.7	408.0	403.8	399.5	391.2	382.6
-0.25	414.2	413.7	410.3	406.5	398.8	391.1
-0.3	418.6	418.4	415.6	412.6	405.1	397.8
-0.35	421.9	421.8	419.7	417.2	410.0	402.9
-0.4	-	424.2	422.6	420.1	413.7	406.9
-0.45	425.6	425.6	424.2	422.3	416.1	409.8
-0.5	426.3	426.0	424.4	422.2	417.3	411.3
-0.55	426.1	425.3	423.7	421.6	417.3	412.0
-0.6	425.4	423.8	422.3	420.0	416.5	411.3
-0.65	524.1	421.8	420.1	417.9	414.6	409.9
-0.7	422.2	419.2	417.4	414.8	411.7	407.3
-0.75	419.7	416.3	414.2	411.2	408.1	403.7
-0.8	416.9	412.8	410.6	407.2	403.8	399.5
-0.85	413.6	409.3	406.8	402.8	399.2	394.8
-0.9	409.9	405.4	402.3	398.0	393.8	389.4
-0.95	405.8	400.9	397.1	392.7	388.0	383.3
-1.0	401.5	396.3	391.7	386.9	381.8	376.7
-1.05	396.5	391.6	386.2	380.7	375.2	369.5
-1.1	391.2	386.3	380.1	374.3	368.1	362.2
-1.15	385.8	380.4	373.1	367.4	360.7	354.2
-1.2	379.8	374.2	366.8	360.4	352.8	346.1
-1.25	373.5	367.3	359.6	352.7	344.7	336.7
-1.3	366.5	360.0	351.4	344.5	336.1	327.6
-1.35	359.3	352.5	342.0	336.1	326.8	318.3
-1.4	350.9	344.1	330.5	326.5	316.9	

Electrocapillary data on mercury in aqueous NaH_2PO_4. T = 25°

Potential measured with respect to sat. calomel electrode.

Given is the interfacial tension x 10^3 in $N.m^{-1}$.

Reference: R. Parsons and F.G.R. Zobel, J. Electroanal. Chem. 9, 333 (1965)

$\frac{E}{\text{Volts}}$	$\frac{c}{\text{Mole.l}^{-1}}$ =					
	0.01045	0.02008	0.04945	0.09896	0.1989	0.4709
-1.70	293.8	289.0	282.9	279.0	274.4	272.0
-1.65	302.5	299.1	293.5	289.6	285.4	283.0
-1.60	311.4	309.2	303.5	300.2	296.3	293.9
-1.55	321.3	318.6	313.3	310.2	306.3	304.5
-1.50	329.8	327.6	322.5	319.5	316.3	314.3
-1.45	338.2	336.2	331.4	328.7	325.5	323.9
-1.40	345.9	344.5	340.1	337.6	335.0	333.5
-1.35	353.7	352.5	348.7	346.4	343.6	342.0
-1.30	361.0	359.9	356.0	354.4	351.5	350.2
-1.25	368.0	367.2	363.5	361.9	359.3	358.3
-1.20	374.7	373.6	370.4	368.9	366.6	365.7
-1.15	381.4	380.1	376.8	375.6	373.3	372.5
-1.10	387.1	386.2	383.3	381.7	379.8	379.0
-1.05	392.3	391.7	389.1	387.7	385.9	385.3
-1.00	397.5	396.9	394.6	393.2	391.6	391.2
-0.95	402.4	401.8	399.5	398.3	397.0	396.4
-0.90	406.5	406.2	404.2	403.2	401.8	401.3
-0.85	410.2	410.0	408.3	407.3	406.1	406.1
-0.80	413.7	413.8	412.1	411.4	410.3	410.2
-0.75	416.7	416.8	415.5	414.9	413.8	414.0
-0.70	419.4	419.4	418.2	417.8	416.8	417.1
-0.65	421.7	421.0	420.6	420.4	419.5	419.7
-0.60	423.3	423.2	422.9	422.3	421.8	422.0
-0.55	424.3	424.4	424.0	423.8	423.4	423.8
-0.50	425.0	425.2	425.0	424.8	424.5	425.0
-0.45	425.1	425.4	425.3	425.1	425.0	425.6
-0.40	424.7	425.6	424.9	424.9	424.9	425.4
-0.35	423.8	424.0	424.4	424.1	424.3	424.5
-0.30	422.3	422.6	423.1	422.8	423.0	423.1
-0.25	420.3	420.4	421.1	420.7	421.1	420.9
-0.20	417.6	417.7	418.5	418.0	418.4	417.4
-0.15	414.2	414.3	415.0	414.3	414.7	413.2
-0.10	409.7	410.0	410.9	410.0	410.1	408.1
-0.05	404.9	404.8	405.9	404.6	404.4	401.8
-0.00	399.0	399.0	399.6	398.4	398.1	394.7
+0.05	391.9	391.3	392.8	390.4	389.8	386.5
+0.10	383.1	384.1	387.4	382.8	382.1	376.4
+0.15	375.0	375.5	376.2	374.0	372.4	366.0
+0.20	367.6	365.7	366.9	363.7	361.6	354.5
+0.25		355.3	355.9	352.2	349.6	343.1
+0.30				338.9	336.0	

Electrocapillary data on mercury in aqueous NaH_2PO_4 (cont.)

$\frac{E}{\text{Volts}}$	$\frac{c}{\text{mole.l}^{-1}}$ = 0.9796	1.9444	3.9779	5.7432
-1.70	270.4	269.0		
-1.65	282.0	280.7	288.1	
-1.60	293.0	292.0	296.6	
-1.55	303.3	299.5	305.8	
-1.50	313.6	312.9	314.7	
-1.45	323.1	322.7	325.0	324.0
-1.40	330.4	332.9	335.1	334.2
-1.35	231.5	341.5	344.0	343.6
-1.30	349.7	350.1	352.7	352.5
-1.25	357.7	358.0	360.7	361.1
-1.20	365.2	365.7	368.6	368.6
-1.15	372.4	372.7	375.8	376.3
-1.10	379.0	379.7	382.4	383.6
-1.05	385.3	386.0	389.3	390.1
-1.00	391.3	391.8	394.9	396.4
-0.95	396.9	397.6	400.8	402.2
-0.90	401.8	402.6	405.7	407.4
-0.85	406.6	407.5	410.8	412.6
-0.80	410.9	411.7	415.0	417.0
-0.75	414.5	415.6	418.7	421.0
-0.70	417.9	418.9	421.6	424.7
-0.65	420.3	421.9	424.8	427.6
-0.60	423.0	424.2	427.0	429.5
-0.55	424.9	425.8	428.1	430.9
-0.50	425.8	426.8	428.5	431.6
-0.45	426.3	427.1	428.3	430.6
-0.40	426.0	426.6	427.9	429.0
-0.35	425.2	425.5	424.9	426.0
-0.30	423.4	423.2	422.1	422.8
-0.25	420.4	419.7	417.3	418.4
-0.20	416.8	415.2	413.0	412.9
-0.15	412.0	410.5	406.0	406.0
-0.10	406.3	404.2	398.4	399.2
-0.05	399.4	396.3	391.3	391.1
-0.00	391.5	387.6	382.7	382.8
+0.05	383.1	378.7	373.2	373.3
+0.10	373.3	368.4	362.7	364.2
+0.15	362.5	357.5	351.2	352.2
+0.20	350.8	345.5	338.7	340.2
+0.25	337.5	337.1	324.8	327.4

Electrocapillary data on mercury in $HClO_4$ solutions at various temperatures.

Potential measured with respect to saturated calomel electrode.

Given is the interfacial tension $\gamma \times 10^3$ in $N.m^{-1}$.

Reference: K.M. Joshi, Thesis, University of Bristol, 1959.

0.02m $HClO_4$

20°		30°		40°		50°	
E Volts		E Volts		E Volts		E Volts	
0.443	323.9	0.002	398.1	0.444	323.8	0.448	322.7
0.348	344.4	0.122	381.5	0.344	344.2	0.342	344.7
0.251	362.7	0.217	366.8	0.244	362.9	0.245	362.4
0.153	378.1	0.318	350.2	0.143	379.0	0.140	379.0
0.153	392.7	0.421	328.9	0.040	392.9	0.044	391.9
-0.058	405.6	-0.102	408.6	-0.039	401.7	-0.049	402.1
-0.160	414.7	-0.201	414.7	-0.138	411.2	-0.143	410.6
-0.266	421.7	-0.308	421.2	-0.240	418.2	-0.238	416.6
-0.359	425.7	-0.410	425.0	-0.335	422.3	-0.340	420.4
-0.457	427.3	-0.503	426.1	-0.435	424.5	-0.435	422.8
-0.563	427.2	-0.605	424.6	-0.529	424.6	-0.538	423.2
-0.664	424.5	-0.703	421.3	-0.635	422.4	-0.637	421.3
-0.761	419.9	-0.809	415.6	-0.735	418.5	-0.737	417.4
-0.857	413.7	-0.908	408.8	-0.835	412.9	-0.839	411.9
-0.966	405.1	-1.018	399.1	-0.933	405.7	-0.937	404.6
-1.068	395.1	-1.122	388.4	-1.045	395.9	-1.035	396.6
-1.189	377.5	-1.220	377.2	-1.144	385.5	-1.129	386.7
-1.265	371.9	-1.321	365.1	-1.242	374.3	-1.245	374.2

0.05m $HClO_4$

0.459	318.9	0.463	316.0	0.451	318.4	0.454	317.8
0.357	341.2	0.364	338.0	0.356	339.2	0.355	339.0
0.258	360.6	0.262	357.8	0.257	358.4	0.258	357.3
0.158	377.6	0.162	374.8	0.155	375.7	0.157	374.8
0.054	391.3	0.060	389.7	0.055	390.6	0.056	381.0
-0.048	402.8	-0.046	402.3	-0.044	401.7	-0.046	400.9
-0.144	412.0	-0.145	411.9	-0.144	411.4	-0.148	410.3
-0.236	419.5	-0.247	419.2	-0.244	418.4	-0.248	416.8
-0.340	426.0	-0.345	423.6	-0.341	422.8	-0.346	420.8
-0.437	426.8	-0.440	425.8	-0.432	424.7	-0.451	423.3
-0.535	427.2	-0.538	425.6	-0.530	424.6	-0.549	422.6
-0.631	425.0	-0.638	422.8	-0.631	422.1	-0.644	420.0
-0.729	421.4	-0.735	418.4	-0.729	417.8	-0.743	415.6
-0.838	413.9	-0.835	412.0	-0.826	411.7	-0.841	409.5
-0.937	405.9	-0.936	403.6	-0.923	404.3	-0.939	402.0
-1.046	395.5	-1.036	394.2	-1.038	393.6	-1.042	392.5
-1.145	384.1	-1.133	383.3	-1.140	382.4	-1.146	381.0
-1.255	370.1	-1.257	368.1	-1.242	370.2		
-1.352	357.5						

Electrocapillary data on mercury in $HClO_4$ (cont.)

0.1m $HClO_4$

20°		30°		40°		50°	
$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ
0.451	314.6	0.460	314.4	0.456	313.3	0.464	311.0
0.351	337.4	0.355	335.3	0.362	334.3	0.363	333.6
0.249	357.8	0.249	356.7	0.269	353.4	0.263	353.5
0.149	375.0	0.149	374.1	0.166	371.3	0.164	370.9
0.048	389.7	0.055	388.1	0.062	387.4	0.065	385.6
-0.055	402.2	-0.050	401.1	-0.047	400.1	-0.047	399.1
-0.154	411.8	-0.148	410.9	-0.148	410.0	-0.148	409.0
-0.260	419.4	-0.248	418.9	-0.244	417.0	-0.281	417.7
-0.361	425.3	-0.351	424.5	-0.341	421.8	-0.344	420.5
-0.461	425.9	-0.448	426.6	-0.443	423.9	-0.449	422.5
-0.558	425.0	-0.552	426.0	-0.543	423.7	-0.550	422.8
-0.664	421.5	-0.651	423.0	-0.641	420.9	-0.650	419.1
-0.767	415.9	-0.750	417.7	-0.740	416.1	-0.748	414.4
-0.868	408.5	-0.848	410.8	-0.840	409.4	-0.843	408.1
-0.968	399.7	-0.945	402.4	-0.937	401.5	-0.944	400.2
-1.069	389.4	-1.051	391.6	-1.035	392.2	-1.042	390.6
-1.173	377.4	-1.153	379.7	-1.133	382.0	-1.142	379.7
-1.245	369.1	-1.250	366.8	-1.235	369.0	-1.240	368.2

0.2m $HClO_4$

20°		30°		40°		50°	
$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ
0.450	318.6	0.453	310.3	0.446	311.0	0.449	310.9
0.351	334.3	0.350	333.7	0.347	334.1	0.354	332.2
0.252	354.7	0.250	354.2	0.247	354.1	0.254	352.4
0.155	371.6	0.157	370.7	0.147	371.9	0.152	370.3
0.052	387.2	0.060	385.6	0.045	386.8	0.054	384.8
-0.051	400.2	-0.045	399.0	-0.045	397.7	-0.045	397.2
-0.154	410.7	-0.145	405.6	-0.140	407.6	-0.149	407.7
-0.255	418.0	-0.249	417.5	-0.242	415.6	-0.253	215.4
-0.350	423.6	-0.341	422.5	-0.341	421.0	-0.347	420.0
-0.450	426.6	-0.436	425.1	-0.437	423.6	-0.442	422.3
-0.553	425.8	-0.539	425.1	-0.534	423.6	-0.537	422.1
-0.648	422.9	-0.633	422.5	-0.638	420.8	-0.634	419.7
-0.748	417.4	-0.727	418.0	-0.740	416.0	-0.734	415.2
-0.847	410.5	-0.827	411.5	-0.843	409.0	-0.835	408.6
-0.949	401.7	-0.927	403.2	-0.943	400.6	-0.929	401.0
-1.049	391.4	-1.026	393.5	-1.045	390.3	-1.027	391.6
-1.150	386.3	-1.127	381.8	-1.144	379.0	-1.142	378.8
-1.258	364.7	-1.224	368.8	-1.240	367.4	-1.251	365.3
		-1.334	354.7				

Electrocapillary data on mercury in $HClO_4$ (cont.)

0.5m $HClO_4$

20°		30°		40°		50°	
E Volts	γ	E Volts	γ	E Volts	γ	E Volts	γ
0.449	305.7	0.370	325.8	0.351	329.4	0.352	328.3
0.340	331.6	0.267	347.7	0.254	349.6	0.244	350.2
0.230	354.4	0.170	365.7	0.156	367.2	0.144	367.7
0.131	372.1	0.073	380.6	0.057	382.1	0.047	382.4
0.014	389.2	-0.038	395.0	-0.061	397.2	-0.045	394.3
-0.086	401.1	-0.133	406.0	-0.161	407.5	-0.146	404.9
-0.184	410.6	-0.234	414.0	-0.259	414.8	-0.245	413.1
-0.286	418.2	-0.334	421.0	-0.353	419.9	-0.344	418.8
-0.378	422.7	-0.432	423.5	-0.457	422.8	-0.444	421.8
-0.478	425.1	-0.529	424.8	-0.559	422.5	-0.539	421.8
-0.592	424.1	-0.627	421.9	-0.654	418.9	-0.638	419.5
-0.697	420.0	-0.724	417.3	-0.750	414.0	-0.736	414.6
-0.798	413.7	-0.824	410.5	-0.851	407.3	-0.832	407.9
-0.900	405.1	-0.922	402.2	-0.951	398.3	-0.930	399.7
-0.978	397.2	-1.016	392.7	-1.050	387.9	-1.013	391.4
-1.073	386.7	-1.114	380.9	-1.149	375.9	-1.116	379.7
-1.173	374.0	-1.204	368.6				

1.0m $HClO_4$

20°		30°		40°		50°	
E Volts	γ	E Volts	γ	E Volts	γ	E Volts	γ
0.345	321.4	0.352	323.6	0.339	325.2	0.335	327.0
0.247	342.9	0.255	344.8	0.240	347.9	0.233	347.3
0.148	361.4	0.156	363.5	0.143	365.2	0.132	365.4
0.050	377.4	0.059	378.3	0.042	380.7	0.033	380.6
-0.053	391.4	-0.057	393.7	-0.042	391.6	-0.048	391.1
-0.151	402.6	-0.155	404.0	-0.140	402.5	-0.148	398.0
-0.248	411.2	-0.254	412.4	-0.238	410.8	-0.246	410.0
-0.345	420.1	-0.355	418.5	-0.338	417.9	-0.346	416.0
-0.444	422.0	-0.456	421.8	-0.439	421.0	-0.441	419.3
-0.542	423.5	-0.550	422.5	-0.538	421.9	-0.536	420.4
-0.641	422.1	-0.651	420.3	-0.626	420.8	-0.638	418.4
-0.737	418.0	-0.751	415.5	-0.726	417.4	-0.741	413.9
-0.834	411.7	-0.846	408.5	-0.824	410.5	-0.839	407.4
-0.931	403.2	-0.942	400.0	-0.923	402.1	-0.933	399.0
-1.030	392.8	-1.042	389.2	-1.020	392.2	-1.032	389.2
-1.129	380.8	-1.139	377.5	-1.121	381.1	-1.131	378.0
-1.222	370.7						

Electrocapillary data on mercury in HClO$_4$ (cont.)

2m HClO$_4$

20°		30°		40°		50°	
$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ	$\frac{E}{\text{Volts}}$	γ
0.335	319.7	0.324	320.5	0.346	316.8	0.315	322.7
0.235	345.2	0.223	341.9	0.252	336.9	0.210	344.6
0.133	360.2	0.130	359.0	0.150	356.3	0.108	369.9
0.033	376.2	0.031	374.9	0.045	372.7	0.027	374.9
-0.048	386.6	-0.058	386.7	-0.053	387.2	-0.059	386.6
-0.143	397.5	-0.155	397.9	-0.155	399.3	-0.162	397.7
-0.245	407.3	-0.257	407.2	-0.262	406.8	-0.262	406.3
-0.337	414.2	-0.361	414.3	-0.359	413.1	-0.360	412.8
-0.439	419.1	-0.465	418.7	-0.458	417.4	-0.454	416.6
-0.533	421.1	-0.559	420.3	-0.557	419.2	-0.547	418.3
-0.628	421.1	-0.653	419.7	-0.656	418.3	-0.645	417.4
-0.582	421.4	-0.754	415.9	-0.755	414.6	-0.741	414.3
-0.719	418.4	-0.861	409.2	-0.855	408.4	-0.844	408.1
-0.813	413.1	-0.963	400.1	-0.955	399.9	-0.938	400.3
-0.911	404.7	-1.060	389.8	-1.055	389.2	-1.048	389.0
-1.005	395.8	-1.135	380.3	-1.150		-1.148	
-1.115	382.5						

3m HClO$_4$

0.332	309.1	0.294	317.5	0.410	288.0	0.406	292.6
0.230	332.2	0.185	340.0	0.331	308.0	0.313	314.8
0.129	351.5	0.088	358.0	0.236	329.4	0.316	335.9
0.020	368.9	-0.025	374.9	0.134	349.0	0.114	354.4
-0.074	382.5	-0.127	388.3	0.033	366.1	0.017	369.9
-0.169	394.0	-0.226	398.8	-0.075	381.4	-0.086	383.3
-0.275	404.2	-0.326	407.5	-0.180	393.6	-0.181	394.4
-0.381	412.2	-0.430	413.9	-0.273	402.3	-0.269	402.0
-0.480	417.3	-0.527	417.8	-0.374	409.6	-0.365	409.0
-0.581	419.9	-0.625	419.1	-0.478	415.2	-0.467	413.7
-0.690	419.9	-0.720	417.9	-0.577	417.7	-0.569	416.8
-0.787	416.6	-0.818	413.8	-0.673	417.7	-0.662	416.0
-0.881	411.0	-0.923	406.5	-0.769	415.1	-0.762	412.5
-0.979	402.6	-1.018	396.8	-0.870	409.9	-0.862	407.3
-1.076	392.5	-1.120	385.8	-0.972	401.6	-0.962	399.5
-1.180	379.9			-1.073	391.4	-1.059	389.8

Electrocapillary data on mercury in aqueous solutions of LiCl of indicated concentration in mole.l^{-1}. T = 25°.

Potential measured with respect to a calomel electrode in the working solution.

Given is the interfacial tension x 10^3 in N.m^{-1}.

Reference: Z. Kovac, Dissertation, University of Pennsylvania, Philadelphia, 1964.

$\frac{E}{\text{Volts}}$	c = 3.0	1.0	0.30	0.10	0.03	0.01
-0.000		354.9	359.7	363.4	363.4	363.5
-0.050	361.7	367.8	372.5	375.3	373.8	373.9
-0.100	372.7	379.2	383.6	385.7	383.1	382.9
-0.150	382.5	389.1	393.1	394.8	391.4	390.9
-0.200	391.2	397.6	401.1	402.5	398.7	397.9
-0.250	398.7	404.9	407.8	409.0	405.1	404.1
-0.300	405.2	410.8	413.4	414.3	410.6	409.4
-0.350	410.5	415.6	417.7	418.6	415.2	413.9
-0.400	414.9	419.3	421.1	421.8	419.0	417.6
-0.450	418.1	421.9	423.4	424.1	421.9	420.7
-0.500	420.4	423.5	424.8	425.6	424.1	423.1
-0.550	421.6	424.2	425.4	426.2	425.6	424.9
-0.600	421.8	423.9	425.2	426.1	426.3	426.0
-0.650	421.0	422.9	424.2	425.3	426.3	426.8
-0.700	419.4	421.0	422.5	424.0	425.6	426.5
-0.750	416.9	418.4	420.2	422.0	424.3	425.9
-0.800	413.6	415.2	417.3	419.5	422.4	424.7
-0.850	409.7	411.4	413.8	416.5	420.0	423.0
-0.900	405.1	407.0	409.9	413.1	417.0	420.7
-0.950	399.9	402.2	405.4	409.2	413.5	418.0
-1.000	394.3	396.9	400.6	405.0	409.6	414.8
-1.050	383.2	391.3	395.4	400.4	405.1	411.1
-1.100	381.8	385.3	389.8	395.4	400.4	406.9
-1.150	375.1	378.9	383.9	390.0	395.2	402.2
-1.200	367.0	372.2	377.7	384.3	389.7	397.4
-1.250	360.5	365.2	371.1	378.3	383.9	392.0
-1.300	352.7	357.8	364.1	371.9	377.7	386.3
-1.350	344.2	350.0	356.7	365.1	371.2	380.2
-1.400	335.1	341.6	348.8	357.9	364.3	373.8
-1.450		332.5	340.3	350.3	357.1	367.0
-1.500				342.2	349.4	359.9
-1.550						352.4
-1.600						349.5

Electrocapillary data on mercury in aqueous solutions of RbCl of indicated concentration in mole.l^{-1}. T = 25°.

Potential measured with respect to a calomel electrode in the working solution.

Given is the interfacial tension x 10^3 in N.m^{-1}.

Reference: Z. Kovac, Dissertation, University of Pennsylvania,
Philadelphia, 1964.

E Volts	c = 3.0	1.0	0.3	0.1	0.03	0.01
-0.000		349.3	360.1	363.7	368.0	365.6
-0.050	360.0	365.3	372.9	375.6	379.6	377.0
-0.100	373.2	378.4	384.0	386.0	389.7	386.9
-0.150	383.6	389.2	393.5	395.0	398.2	395.4
-0.200	392.2	393.0	401.4	402.7	405.5	402.6
-0.250	399.5	405.2	408.0	409.2	411.5	408.7
-0.300	405.8	411.0	413.4	414.5	416.4	413.7
-0.350	411.0	415.6	417.7	418.7	420.2	417.8
-0.400	414.9	419.2	420.9	421.9	423.1	421.1
-0.450	418.5	421.7	423.1	424.2	425.1	423.5
-0.500	420.7	423.3	424.4	425.6	426.3	425.3
-0.550	421.7	424.0	424.9	426.6	426.7	426.3
-0.600	421.6	423.7	424.6	426.0	426.6	426.8
-0.650	420.6	422.7	423.5	425.1	425.9	426.7
-0.700	419.0	420.9	421.7	423.6	424.6	426.1
-0.750	416.0	418.4	419.3	421.5	422.8	424.9
-0.800	413.1	415.2	416.4	418.9	420.5	423.3
-0.850	409.0	411.3	412.9	415.7	416.8	421.2
-0.900	404.7	407.0	408.9	412.1	414.6	418.7
-0.950	399.6	402.1	464.5	408.1	411.0	415.7
-1.000	393.8	396.8	399.7	403.8	407.0	412.3
-1.050	387.6	391.1	394.5	399.0	402.6	408.4
-1.100	381.4	385.1	388.9	393.9	397.8	404.1
-1.150	374.5	378.6	383.0	388.4	392.5	399.4
-1.200	368.1	371.7	376.7	382.4	386.7	394.3
-1.250	359.4	364.2	370.0	376.0	380.4	388.6
-1.300	352.5	355.8	362.7	367.0	373.7	382.6
-1.350			354.8	361.2	366.4	376.0

Electrocapillary data on mercury in aqueous solutions of CsCl of indicated concentration in mole.l^{-1}. T = 25°.
Potential measured with respect to a calomel electrode in the working solution.
Given is the interfacial tension x 10^{3} in N.m^{-1}.
Reference: Z. Kovac, Dissertation, University of Pennsylvania,
Philadelphia, 1964.

E / Volts	c = 3.0	1.0	0.3	0.1	0.03	0.01
-0.000	340.4	345.0	358.2	372.2	371.1	368.3
-0.050	355.6	362.3	371.3	382.8	381.7	376.9
-0.100	368.6	376.4	382.4	390.0	390.2	385.3
-0.150	379.9	387.8	391.9	398.0	398.2	392.3
-0.200	388.5	397.1	399.9	405.5	405.1	398.8
-0.250	398.6	404.6	406.7	411.5	410.5	404.5
-0.300	404.2	410.5	412.2	416.1	415.3	409.8
-0.350	409.5	415.2	416.7	419.9	419.4	414.5
-0.400	413.6	418.7	420.0	423.2	422.4	418.6
-0.450	415.5	421.1	422.5	425.1	424.4	422.0
-0.500	417.4	422.6	424.0	426.2	425.8	424.5
-0.550	419.2	423.2	424.7	426.3	426.5	426.2
-0.600	419.1	422.9	424.5	426.0	426.5	426.8
-0.650	418.1	421.8	423.5	424.8	425.7	426.7
-0.700	416.3	419.9	421.8	423.1	424.4	426.1
-0.750	412.6	417.3	419.4	420.7	422.6	425.0
-0.800	410.3	414.0	416.4	418.0	420.2	423.2
-0.850	406.1	410.1	412.7	414.6	417.3	420.8
-0.900	401.4	405.6	408.6	411.0	414.1	417.9
-0.950	396.0	400.6	403.9	407.1	410.6	414.5
-1.000	390.0	395.1	393.9	402.9	406.4	410.8
-1.050	383.2	389.2	393.4	397.6	401.6	406.7
-1.100	376.6	382.9	387.5	392.4	396.7	402.1
-1.150	369.3	376.3	381.3	386.4	391.0	396.8
-1.200	361.6	369.2	374.7	379.9	385.1	391.3
-1.250	353.7	361.7	367.6	373.5	379.0	395.3
-1.300	344.6	353.6	359.9	367.0	371.9	378.9
-1.350	337.4		351.6		369.2	372.2

Electrocapillary data on mercury in aqueous solutions of CsI of indicated concentration in mole.l^{-1}. T = 25^{o}.

Potential measured with respect to a calomel electrode in the working solution.

Given is the interfacial tension x 10^3 in N.m^{-1}.

Reference: Z. Kovac, Dissertation, University of Pennsylvania, Philadelphia, 1964.

E/Volts	c = 1.0	0.3	0.1	0.03	0.01
0.050	296.9	289.3	311.0	322.0	337.9
0.000	325.7	324.0	340.9	352.1	363.2
-0.050	347.5	350.1	363.6	374.3	382.3
-0.100	363.9	369.4	380.4	390.4	396.5
-0.150	375.8	383.4	392.7	401.8	406.7
-0.200	384.3	393.2	401.2	409.3	413.6
-0.250	390.2	399.9	406.8	414.1	418.1
-0.300	393.9	404.2	410.1	416.6	420.5
-0.350	395.9	406.4	411.6	417.4	421.4
-0.400	396.5	406.9	411.7	417.0	421.1
-0.450	395.8	406.1	410.5	415.5	419.8
-0.500	394.1	404.8	408.3	413.2	417.7
-0.550	391.3	401.5	405.4	410.2	415.0
-0.600	387.6	397.8	401.6	406.6	411.7
-0.650	382.9	393.3	397.2	402.6	407.9
-0.700	377.3	388.1	392.2	397.6	403.6
-0.750	370.8	382.2	386.6	392.4	398.9
-0.800	363.6	375.7	380.4	386.7	393.8
-0.850	355.7	368.5	373.6	380.5	388.1
-0.900	347.4	360.7	366.4	373.8	382.1
-0.950	338.9	352.4	358.7	366.8	375.5
-1.000	330.6	343.6	350.6	359.3	368.5
-1.050		334.4	342.1	351.5	361.1
-1.100		324.8	333.2	343.3	353.2
-1.150		314.9	324.0	334.7	345.0
-1.200		304.7	314.5	325.5	336.3
-1.250		294.2	304.6	315.7	327.3
-1.300		283.3	294.4	305.0	318.0
-1.350		271.9	283.7		308.2
-1.400		259.6	272.4		298.1
-1.450			260.2		287.6
-1.500			246.8		276.6

Electrocapillary data on mercury in aqueous solutions of NaSCN of indicated concentration in mole.l^{-1}. $T = 25^{\circ}$.

Potential measured with respect to a sodium responsive glass-electrode dipping into the working solution.

Given is the interfacial tension x 10^{3} in N.m^{-1}.

Reference: Z. Kovac, Dissertation, University of Pennsylvania, Philadelphia, 1964.

E Volts	c = 3.0	1.0	0.3	0.1	0.03	0.01
-0.200			327.3	357.5	377.2	390.0
-0.250		316.4	343.9	370.0	387.9	398.3
-0.300	307.5	332.9	358.4	380.9	396.9	405.3
-0.350	322.8	347.4	370.8	390.3	404.3	411.2
-0.400	336.5	360.0	381.4	398.3	410.5	416.0
-0.450	348.7	370.9	390.3	405.0	415.3	419.7
-0.500	359.5	380.2	397.7	410.3	419.0	422.4
-0.550	368.9	388.0	403.8	414.5	421.7	424.3
-0.600	377.0	394.4	408.5	417.6	423.4	425.3
-0.650	383.9	399.6	412.1	419.5	424.3	425.5
-0.700	389.6	403.6	414.6	420.5	424.3	425.1
-0.750	394.2	406.5	416.0	420.6	423.6	423.9
-0.800	397.7	408.4	416.6	419.9	422.3	422.1
-0.850	400.1	409.3	416.2	418.3	420.4	419.8
-0.900	401.7	409.4	415.1	416.1	417.9	416.9
-0.950	402.3	408.4	413.1	413.2	414.8	413.6
-1.000	402.0	406.7	410.5	409.8	411.3	409.7
-1.050	400.8	404.2	407.1	405.8	407.2	405.5
-1.100	398.9	401.0	403.0	401.3	402.7	400.8
-1.150	396.2	397.0	398.3	396.3	397.7	395.8
-1.200	392.6	392.3	393.1	390.9	392.3	390.4
-1.250	388.4	387.0	387.3	385.2	386.4	384.7
-1.300	383.4	381.0	380.9	379.0	380.1	378.6
-1.350	377.8	374.4	374.1	372.4	373.4	373.6
-1.400	371.4	367.2	367.0	365,4	366.3	365.6
-1.450	364.4	359.6	359.4	358.1	358.9	358.6
-1.500	356.8	351.6	351.6	350.4	351.2	351.4
-1.550	348.6	343.4	343.7	342.2	343.3	343.9
-1.600	339.9	335.0	335.6	332.6	335.1	336.1
E' Volts	0.236	0.206	0.177	0.153	0.126	0.102

E' is the potential of this glass electrode with respect to a saturated calomel electrode.

Electrocapillary data on mercury in aqueous solutions of $NaBrO_3$ of indicated concentration in mole.l^{-1}. T = 25°.
Potential measured with respect to a sodium responsive glass electrode dipping into the working solution.
Given is the interfacial tension x 10^3 in N.m^{-1}.
Reference: Z. Kovac, Dissertation, University of Pennsylvania, Philadelphia, 1964.

E/Volts	c = 1.0	0.3	0.1	0.03	0.01
-0.200		390.7	397.0	402.1	407.2
-0.250	386.3	396.6	402.4	406.6	411.2
-0.300	393.1	402.5	407.7	411.0	414.9
-0.350	399.2	407.7	412.5	415.2	418.4
-0.400	404.5	412.3	416.6	418.9	421.4
-0.450	409.1	415.7	419.8	421.6	423.5
-0.500	412.9	418.5	422.2	423.7	425.2
-0.550	415.9	420.5	423.9	425.1	426.2
-0.600	418.1	421.9	424.9	425.9	426.6
-0.650	419.5	422.5	425.2	426.1	426.7
-0.700	420.1	422.6	424.8	425.5	425.9
-0.750	419.9	422.0	423.9	424.4	424.6
-0.800	418.9	420.7	422.2	422.9	422.8
-0.850	417.4	418.7	420.0	420.7	420.5
-0.900	415.2	416.1	417.2	417.9	417.6
-0.950	412.6	413.1	414.0	414.7	414.2
-1.000	409.5	409.5	410.2	411.0	410.2
-1.050	405.9	405.9	406.2	407.0	406.3
-1.100	401.9	401.4	401.7	402.5	401.8
-1.150		396.3	396.5	397.6	396.6
-1.200		391.2	391.2	392.5	391.4
-1.250		385.6	385.6	387.0	385.6
-1.300		379.5	379.9	381.0	380.2
-1.350		373.1	373.7	374.8	374.0
-1.400		366.1	367.3	368.3	366.8
-1.450		359.4	360.6	361.8	360.2
E'/Volts	0.209	0.183	0.159	0.130	0.104

E' is the potential of this glass electrode with respect to a saturated calomel electrode.

Electrocapillary data on mercury in aqueous solutions of $NaClO_3$ of indicated concentration in $mole.l^{-1}$. T = 25^o. Potential measured with respect to a sodium responsive glass electrode dipping into the working solution. Given is the interfacial tension x 10^3 in $N.m^{-1}$.

Reference: Z. Kovac, Dissertation, University of Pennsylvania,
Philadelphia, 1964.

E Volts	c = 3.0	1.0	0.3	0.1	0.03	0.01
-0.100		376.6	379.6	383.6	387.9	391.3
-0.150	358.0	384.3	386.6	389.8	394.3	397.0
-0.200	370.5	391.3	393.2	395.8	400.3	402.5
-0.250	380.0	397.5	399.2	401.5	405.8	407.5
-0.300	387.9	403.0	404.6	406.8	410.6	412.1
-0.350	394.0	407.8	409.5	411.6	414.8	416.1
-0.400	399.5	412.0	413.7	415.7	418.5	419.5
-0.450	404.4	415.4	417.2	419.2	421.5	422.2
-0.500	408.3	418.2	420.1	421.9	423.8	424.3
-0.550	411.7	420.3	422.2	424.0	425.5	425.8
-0.600	414.8	421.8	423.7	425.3	426.4	426.6
-0.650	417.2	422.5	424.4	426.0	426.6	426.8
-0.700	418.6	422.6	424.5	425.9	426.4	426.4
-0.750	419.2	421.1	423.8	425.3	425.4	425.4
-0.800	419.1	420.8	422.5	424.0	423.8	423.9
-0.850	418.4	419.0	420.6	422.0	421.6	421.9
-0.900	417.0	416.5	418.0	419.6	418.8	419.4
-0.950	414.7	413.5	414.9	416.6	415.5	416.4
-1.000	412.2	409.8	411.2	413.1	411.8	413.0
-1.050	409.2	405.6	407.1	409.1	407.5	409.1
-1.100	405.6	400.9	402.4	409.7	403.0	404.8
-1.150	401.3	395.7	397.4	399.8	397.9	400.1
-1.200	396.1	390.1	392.0	394.5	392.5	395.0
-1.250	390.4	384.0	386.1	388.9	386.7	389.4
-1.300	384.1	377.4	380.0	382.8	380.6	383.5
-1.350	377.3	370.5	373.5	376.3	374.1	377.1
-1.400	370.9	363.2	366.6	369.5	367.4	370.3
-1.450	363.6	355.6	359.3	362.3	360.4	363.1
-1.500	356.0	347.5	351.6	354.7	353.2	355.5
-1.550	348.3	339.1	343.2	396.7	345.6	347.5
-1.600	340.0	330.1	334.2	338.4	337.1	339.4
E' Volts	0.234	0.208	0.183	0.159	0.130	0.102

E' is the potential of this glass electrode with respect to a saturated calomel electrode.

Electrocapillary data on mercury in aqueous solutions of NaCN of indicated concentration in mole.l^{-1}. T = 25°.

Potential measured with respect to a sodium responsive glass electrode dipping into the working solution.

Given is the interfacial tension x 10^{3} in N.m^{-1}.

Reference: Z. Kovac, Dissertation, University of Pennsylvania,
Philadelphia, 1964.

E / Volts	c = 3.0	1.0	0.3	0.1	0.03	0.01
-0.200					403.5	406.7
-0.250				400.9	407.0	410.6
-0.300		383.8	396.0	404.7	410.6	414.2
-0.350		390.1	401.0	408.7	414.1	417.5
-0.400		396.0	405.9	412.6	417.3	420.7
-0.450	394.8	401.6	410.5	416.3	420.2	423.1
-0.500	399.9	406.5	414.5	419.5	422.5	424.8
-0.550	404.5	410.8	417.8	422.0	424.2	425.9
-0.600	408.5	414.3	420.5	423.9	425.3	426.3
-0.650	411.8	416.9	422.2	424.9	425.7	426.0
-0.700	414.3	418.5	423.2	425.1	425.4	425.2
-0.750	416.1	419.3	423.3	424.5	424.5	423.8
-0.800	417.0	419.0	422.3	423.0	422.8	422.0
-0.850	417.1	418.9	420.5	420.8	420.5	419.6
-0.900	416.3	415.8	417.9	417.8	417.6	416.8
-0.950	414.6	413.0	414.6	414,2	414.1	413.5
-1.000	412.1	409.4	410.6	410.1	410.1	409.8
-1.050	408.8	405.2	406.0	405.5	405.6	405.6
-1.100	409.7	400.3	401.0	400.5	400.8	410.1
-1.150	399.8	395.0	395.7	395.2	395.6	396.2
-1.200	394.4	389.3	389.8	389.7	390.1	390.8
-1.250	388.3	383.3	383.9	383.9	384.4	385.1
-1.300	381.7	376.9	377.7	378.0	378.4	378.9
-1.350	374.6	370.3	371.4	371.8	372.1	372.4
-1.400	367.2	363.4	364.7	365.2	365.4	365.5
-1.450	359.4	356.1	357.7	358.3	358.4	358.3
-1.500	351.3	348.4	350.1	350.6	350.7	350.6
-1.550	343.1	340.0	341.5	342.1	342.4	342.6
-1.600	334.5	330.8	331.7	332.2	333.0	334.3
E' / Volts	0.218	0.188	0.168	0.144	0.131	0.108

E' is the potential of this glass electrode with respect to a saturated calomel electrode.

Electrocapillary data on mercury in aqueous solutions of $NaClO_4$ of indicated concentration in $mole.l^{-1}$. T = 25°.
Potential measured with respect to a sodium responsive glass electrode dipping into the working solution.
Given is the interfacial tension x 10^3 in $N.m^{-1}$.
Reference: Z. Kovac, Dissertation, University of Pennsylvania, Philadelphia, 1964.

E / Volts	c = 3.0	1.0	0.3	0.1	0.03	0.01
-0.200	371.0	384.1	390.9	398.7	402.1	405.5
-0.250	378.7	390.7	396.8	404.6	408.3	410.6
-0.300	385.9	396.7	402.3	409.8	412.8	414.7
-0.350	392.3	402.3	407.3	404.2	416.9	418.1
-0.400	398.1	407.3	411.7	417.9	420.3	420.8
-0.450	403.2	411.6	415.5	420.8	423.0	423.1
-0.500	407.6	415.2	418.6	423.1	424.9	424.8
-0.550	411.3	418.1	421.1	424.6	426.1	426.0
-0.600	414.3	420.4	422.8	425.6	426.7	426.7
-0.650	416.6	421.8	423.7	425.7	426.6	426.8
-0.700	418.1	422.5	423.9	425.2	426.0	426.5
-0.750	419.0	422.4	423.4	424.2	424.8	425.5
-0.800	419.1	421.7	422.2	422.5	423.0	423.9
-0.850	418.5	420.2	420.4	420.2	420.7	421.6
-0.900	417.3	418.0	417.9	417.4	417.9	418.7
-0.950	415.3	415.3	414.8	414.1	414.6	415.3
-1.000	412.7	411.9	411.1	410.1	410.9	411.3
-1.050	409.4	407.9	407.0	405.9	406.8	406.9
-1.100	405.6	403.3	402.4	401.1	402.3	402.2
-1.150	401.1	398.3	397.4	395.9	397.4	397.1
-1.200	396.2	392.8	391.9	390.4	392.1	391.7
-1.250	390.7	386.8	386.2	389.5	386.4	386.2
-1.300	384.7	380.4	380.1	378.2	380.4	380.4
-1.350	378.3	373.6	373.8	371.8	374.1	374.5
-1.400	371.4	366.4	367.1	365.1	367.3	
-1.450	364.1	358.9	360.2	358.2	360.2	361.4
-1.500	356.4	351.0	359.0	351.2	352.7	353.9
-1.550	398.2	392.7	396.0	344.1	344.7	345.3
-1.600	339.6	334.1	338.2	336.9	336.2	335.0
E' / Volts	0.245	0.217	0.188	0.165	0.135	0.111

E' is the potential of this glass electrode with respect to a saturated calomel electrode.

Electrocapillary data on mercury in aqueous $LiNO_3$ of various concentrations (mole-l^{-1}). Potential E_+ relative to reference electrode, reversible to the cation in the solution. Temperature 25°C.

Given is the interfacial tension γ in mN.m^{-1}.

Reference: B.B. Damaskin, V.F. Ivanov, N.I. Melekhova, L.F. Mayorova, Elektrokhimiya, 4, 1342 (1968).

conc. 0.05		0.20		0.45		1.10		3.20		4.80		7.45		9.33	
E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ
0.298	364.5	0.264	360.9	0.245	357.3	0.215	351.1	0.162	341.4	0.130	339.0	0.090	330.9	0.067	331.4
0.198	383.7	0.164	380.5	0.145	377.3	0.115	371.3	0.062	361.2	0.030	357.4	0.010	349.6	-0.033	350.0
0.098	397.2	0.064	394.3	0.045	391.7	0.015	383.9	-0.038	376.3	-0.070	373.1	-0.110	366.5	-0.133	366.7
-0.002	408.5	-0.036	405.1	-0.055	402.5	-0.085	395.7	-0.138	388.5	-0.170	385.3	-0.210	380.5	-0.233	380.3
-0.102	417.1	-0.136	413.9	-0.155	411.9	-0.185	407.7	-0.238	399.3	-0.270	394.1	-0.310	391.9	-0.333	392.1
-0.202	421.9	-0.236	420.3	-0.255	418.8	-0.285	415.3	-0.338	407.9	-0.370	404.5	-0.410	401.4	-0.433	401.3
-0.302	425.7	-0.336	424.5	-0.355	423.3	-0.385	420.7	-0.438	414.1	-0.470	411.5	-0.510	408.9	-0.533	408.7
-0.402	426.5	-0.436	425.1	-0.355	424.5	-0.485	422.5	-0.538	417.5	-0.570	416.2	-0.610	414.3	-0.633	413.5
-0.502	424.9	-0.536	423.7	-0.555	423.1	-0.585	422.1	-0.638	418.7	-0.670	417.7	-0.710	416.5	-0.733	415.9
-0.602	420.9	-0.636	420.3	-0.655	419.7	-0.685	419.1	-0.738	416.7	-0.770	415.9	-0.810	416.1	-0.833	415.5
-0.702	414.7	-0.736	413.5	-0.755	413.1	-0.785	412.5	-0.838	411.5	-0.870	411.4	-0.910	411.7	-0.933	411.9
-0.802	407.0	-0.836	405.5	-0.855	404.8	-0.885	403.9	-0.938	403.5	-0.970	403.5	-1.010	404.7	-1.033	405.1
-0.902	397.4	-0.936	395.4	-0.955	394.3	-0.985	393.3	-1.038	393.1	-1.070	393.5	-1.110	395.2	-1.133	396.5
-1.002	386.5	-1.036	383.9	-1.055	382.7	-1.085	381.5	-1.138	381.0	-1.170	381.7	-1.210	384.5	-1.233	385.5
-1.102	373.9	-1.136	370.9	-1.155	369.7	-1.185	367.9	-1.238	367.3	-1.270	368.1	-1.310	370.9	-1.333	372.1
-1.202	359.9	-1.236	355.8	-1.255	354.3	-1.285	352.7	-1.338	351.9	-1.370	352.3	-1.410	355.3	-1.433	356.3
-1.302	343.9	-1.336	339.4	-1.355	337.7	-1.385	335.5	-1.438	334.8	-1.470	335.4	-1.510	336.4	-1.533	337.6

Electrocapillary data on mercury in aqueous KF of various concentration (mole-l^{-1}). Potential E_+ relative to reference electrode, reversible to the cation in the solution. Temperature 25°C. Given is the interfacial tension γ in mN.m^{-1}.

Reference: N.I. Melekhova, V.F. Ivanov, B.B. Damaskin, Elektrokhimiya, 5 (1969).

conc.

0.1		1.0		2.0		4.0		6.0		8.0		10.75		12.60	
E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ
2.190	404.5	2.132	396.9	2.115	400.0	2.090	401.5	2.073	405.6	2.060	414.2	2.048	422.0	2.040	420.0
2.090	413.8	2.032	409.6	2.015	411.7	1.990	413.9	1.973	420.3	1.960	427.5	1.948	435.5	1.940	436.5
1.990	421.0	1.932	419.2	1.915	420.4	1.890	423.4	1.873	430.3	1.860	437.6	1.848	446.0	1.840	449.8
1.890	425.1	1.832	425.3	1.815	426.0	1.790	429.9	1.773	436.7	1.760	440.0	1.748	452.0	1.740	458.7
1.790	423.6	1.732	427.6	1.715	428.6	1.690	432.4	1.673	439.6	1.660	446.6	1.648	455.7	1.640	463.2
1.690	426.1	1.632	427.2	1.615	428.3	1.590	432.1	1.573	439.6	1.560	446.5	1.548	456.3	1.540	463.8
1.590	423.8	1.532	424.3	1.515	425.4	1.490	429.2	1.473	436.8	1.460	443.2	1.448	453.5	1.440	461.4
1.490	419.3	1.432	419.2	1.415	420.5	1.390	423.6	1.373	431.3	1.360	437.3	1.348	447.2	1.340	456.0
1.390	412.8	1.332	412.0	1.315	413.2	1.290	415.5	1.273	423.5	1.260	429.1	1.248	438.4	1.240	448.3
1.290	404.6	1.232	403.0	1.215	404.6	1.190	406.4	1.173	413.8	1.160	419.0	1.148	428.3	1.140	438.3
1.190	394.8	1.132	392.3	1.115	393.8	1.090	395.3	1.073	401.8	1.060	407.1	1.048	416.2	1.040	426.3
1.090	383.3	1.032	380.3	1.015	381.3	0.990	382.3	0.973	388.2	0.960	393.2	0.948	402.0	0.940	412.7
0.990	370.3	0.932	367.1	0.915	367.4	0.890	367.0	0.873	374.0	0.860	377.6	0.848	386.1	0.840	396.7
0.890	355.4	0.832	351.3	0.815	352.0	0.790	350.4	0.773	357.1	0.760	360.0	0.748	368.1	0.740	378.9
0.790	339.3	0.732	333.9	0.715	334.1	0.690	332.2	0.673	338.7	0.660	341.4	0.648	348.9	0.640	349.5
0.690	321.7	0.632	315.4	0.615	414.9	0.590	312.9	0.573	317.4	0.560	324.5	0.548	328.2	0.540	339.8
0.590	300.8	0.532	294.3	0.515	294.0	0.490	291.5	0.473	295.9	0.460	297.3	0.448	305.2	0.440	

Electrocapillary data on mercury in aqueous $NaClO_4$ of various concentrations ($mole\text{-}l^{-1}$). Potential E_+ relative to reference electrode, reversible to the cation in the solution. Temperature 25°C. Given is the interfacial tension γ in $mN.m^{-1}$.

Reference: B.B. Damaskin, A.N. Frumkin, V.F. Ivanov, N.I. Melekhova, V.F. Khonina, Elektrokhimiya, 4, 1336, (1968)

conc.

0.1		1.0		3.3		5.5		7.6		9.2	
E_+ volts	γ	E_+ volts	γ	E_+ volts	γ	E_+ volts	γ	E_+ volts	γ	E_+ volts	γ
2.454	344.6	2.405	334.9	2.370	328.4	2.346	327.0	2.333	323.8	2.323	320.7
2.354	364.9	2.305	356.0	2.270	349.2	2.246	346.8	2.235	344.7	2.223	342.0
2.254	381.4	2.205	372.9	2.170	366.0	2.146	363.0	2.135	360.9	2.123	359.8
2.154	394.7	2.105	387.0	2.070	380.8	2.046	377.6	2.035	376.5	2.023	374.2
2.054	404.9	2.005	399.2	1.970	393.0	1.946	391.0	1.935	389.8	1.923	386.8
1.954	413.3	1.905	409.0	1.870	404.0	1.846	402.0	1.835	399.5	1.823	397.9
1.845	420.6	1.805	416.8	1.770	411.4	1.746	409.6	1.735	408.8	1.723	407.6
1.754	424.2	1.705	421.0	1.670	417.0	1.646	415.8	1.635	415.2	1.623	414.2
1.654	425.2	1.605	422.8	1.570	420.4	1.546	418.6	1.535	418.5	1.523	418.6
1.554	423.2	1.505	421.5	1.470	420.0	1.446	419.0	1.435	420.3	1.423	420.3
1.454	419.2	1.405	417.6	1.370	417.0	1.346	417.0	1.335	418.6	1.323	420.0
1.354	412.8	1.305	410.7	1.270	411.2	1.246	411.2	1.235	414.5	1.223	417.6
1.254	404.7	1.205	402.2	1.170	403.2	1.146	403.6	1.135	407.7	1.123	411.0
1.154	394.8	1.105	391.4	1.070	392.8	1.046	393.9	1.035	397.6	1.023	401.5
1.054	382.8	1.005	378.8	0.970	380.2	0.946	381.8	0.935	386.8	0.923	390.2
0.954	369.6	0.905	364.8	0.870	365.8	0.846	367.8	0.835	379.1	0.823	376.9
0.854	355.6	0.805	349.5	0.770	349.6	0.746	351.5	0.735	356.3	0.723	359.8
0.754	339.0	0.705	332.5	0.670	332.3	0.646	333.0	0.635	337.7	0.623	340.7
0.654	320.8	0.605	313.4	0.570	312.0	0.546	313.0	0.535	316.9	0.523	319.5
0.554	300.9	0.505	292.4	0.470	290.0	0.446	290.0	0.435	293.3	0.423	296.2

Electrocapillary data on mercury in aqueous NH_4NO_3 of various concentrations ($mole\text{-}l^{-1}$). Potential E_+ relative to reference electrode, reversible to the cation in the solution. Temperature 25°C.

Given is the interfacial tension γ in $mN.m^{-1}$.

Reference: B.B. Damaskin, V.F. Ivanov, N.I. Melekhova, L.F. Mayorova, Elektrokhimiya, 4, 1342 (1968).

conc. 0.05		0.20		0.45		1.13		3.65		4.45		7.6		10.4	
E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ	E_+/volts	γ
0.288	368.0	0.257	362.0	0.237	357.6	0.216	350.7	0.190	341.8	0.186	341.0	0.171	338.0	0.163	332.4
0.188	383.4	0.157	378.5	0.137	374.8	0.116	368.0	0.090	360.6	0.086	360.0	0.071	356.8	0.063	354.8
0.088	396.0	0.057	392.5	0.037	388.8	0.016	383.2	-0.010	376.8	-0.014	375.8	-0.029	372.8	-0.037	371.2
-0.012	407.8	-0.043	303.0	-0.063	401.0	-0.084	396.4	-0.110	390.2	-0.114	389.4	-0.129	385.7	-0.137	384.4
-0.112	416.4	-0.143	413.4	-0.163	410.9	-0.184	406.1	-0.210	401.0	-0.214	399.9	-0.229	396.6	-0.237	395.0
-0.212	422.4	-0.243	420.5	-0.263	418.1	-0.284	414.3	-0.310	409.0	-0.314	408.2	-0.329	405.0	-0.337	403.8
-0.312	426.0	-0.343	424.1	-0.363	422.2	-0.384	418.7	-0.410	414.8	-0.414	414.0	-0.429	410.8	-0.437	409.8
-0.412	426.4	-0.443	424.6	-0.463	423.2	-0.484	420.8	-0.510	417.2	-0.514	416.8	-0.529	414.8	-0.537	413.6
-0.512	425.0	-0.543	423.8	-0.563	422.8	-0.584	420.9	-0.610	416.9	-0.614	416.4	-0.629	414.8	-0.637	414.0
-0.612	420.4	-0.643	419.6	-0.663	418.8	-0.684	417.5	-0.710	414.5	-0.714	414.0	-0.729	413.6	-0.737	414.2
-0.712	414.6	-0.743	413.8	-0.763	411.6	-0.784	410.9	-0.810	409.6	-0.814	409.6	-0.829	410.0	-0.837	410.8
-0.812	406.9	-0.843	404.0	-0.863	403.0	-0.884	401.8	-0.910	401.8	-0.914	402.0	-0.929	402.1	-0.937	402.7
-0.812	406.9	-0.843	404.0	-0.863	403.0	-0.884	401.8	-0.910	401.8	-0.914	402.0	-0.929	402.1	-0.937	402.7
-0.912	397.3	-0.943	395.0	-0.963	392.8	-0.984	391.5	-1.010	390.9	-1.014	391.2	-1.029	392.0	-1.037	393.4
-1.012	385.8	-1.043	383.4	-1.063	381.6	-1.084	379.8	-1.110	378.8	-1.114	378.9	-1.129	379.4	-1.137	380.3
-1.112	372.6	-1.143	369.8	-1.163	367.0	-1.184	365.8	-1.210	364.6	-1.214	364.8	-1.229	365.5	-1.237	366.0
-1.212	357.4	-1.243	354.0	-1.263	351.0	-1.284	349.0	-1.310	348.9	-1.314	349.1	-1.329	349.9	-1.337	351.4
-1.312	340.8	-1.343	336.7	-1.363	334.4	-1.384	331.8	-1.410	332.8	-1.414	332.9	-1.429	333.6	-1.437	334.2

Electrocapillary data on mercury in aqueous solutions of KNO_3 and $HClO_4$ at 19° and 20°C respectively.

Potentials measured with respect to a normal calomel electrode.

Given is the interfacial tension in mN.m^{-1}.

Reference: S. Bordi, G. Papeschi, J. Electroanal. Chem. and Interfacial Electrochem. 20 (1969) 297.

(N.B. These solutions are the base solutions for the electrocapillary data of pages 22C-222.)

electrolyte :	KNO_3	$HClO_4$
concentration:	1 mole.l^{-1}	0.1 mole.l^{-1}
E / Volts		
0.1	-	381.07
0.05	-	387.95
0.0	375.49	394.00
-0.05	383.25	399.71
-0.1	389.94	404.70
-0.15	396.10	409.47
-0.2	401.58	413.35
-0.25	406.67	416.58
-0.3	410.78	419.52
-0.35	414.43	421.70
-0.4	416.94	423.34
-0.45	419.07	424.23
-0.5	420.59	424.52
-0.55	421.43	424.17
-0.6	421.05	423.05
-0.65	419.99	421.29
-0.7	418.24	419.05
-0.75	415.65	416.41
-0.8	412.46	413.29
-0.85	408.81	409.70
-0.9	404.70	405.70
-0.95	399.98	401.35
-1.0	394.58	396.33
-1.05	388.80	392.00
-1.1	382.41	386.95
-1.15	375.64	-
-1.2	368.65	-
-1.25	361.19	-
-1.3	353.13	-

Electrocapillary data on mercury in aqueous solutions of NaN_3. T = 25°C

Potential with respect to the silver/silver azide electrode in the working solution.

Reference: C.V. D'Alkaine, E.R. Gonzalez and R. Parsons.
J. Electroanal. Chem. 32 (1971) 57.

c/mole l^{-1}	0.00157		0.00392		0.00583	
	E/volts	γ /mN m^{-1}	E/volts	γ /mN m^{-1}	E/volts	γ /mN m^{-1}
	-0.908	422.40	-1.158	404.50	-1.401	377.25
	-0.850	423.80	-1.108	408.80	-1.349	383.30
	-0.804	424.65	-1.061	412.00	-1.296	389.10
	-0.702	425.20	-1.001	415.75	-1.267	392.40
	-0.695	425.45	-0.951	418.65	-1.197	399.30
	-0.657	424.70	-0.900	420.55	-1.145	403.50
	-0.605	423.20	-0.845	422.70	-1.109	406.65
	-0.603	423.20	-0.795	423.20	-1.050	411.00
	-0.551	421.70	-0.755	424.75	-1.002	414.30
	-0.950	416.05	-0.700	425.25	-0.949	417.30
	-0.400	411.50	-0.648	425.05	-0.899	419.70
	-0.351	406.85	-0.651	425.10	-0.893	419.80
	-0.502	418.50	-0.599	424.10	-0.848	421.45
	-0.307	400.20	-0.550	422.55	-0.803	422.85
	-0.550	421.65	-0.549	422.35	-0.751	424.20
			-0.501	420.25	-0.700	424.90
			-0.451	416.65	-0.651	425.10
			-0.399	412.60	-0.601	424.50
					-0.548	423.15
					-0.501	421.30
					-0.454	418.75
					-0.406	415.20
					-0.349	410.25
					-0.296	404.20
					-0.800	423.15

Mercury in aqueous NaN_3 (cont.)

c/mol l^{-1}	0.00968		0.0199		0.0407	
	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}
	-1.404	373.45	-1.403	368.10	-1.405	363.70
	-1.354	379.35	-1.354	374.90	-1.354	370.30
	-1.306	385.35	-1.303	381.50	-1.304	376.95
	-1.255	390.90	-1.254	387.15	-1.252	383.20
	-1.204	396.25	-1.204	392.65	-1.204	388.80
	-1.155	400.90	-1.154	397.60	-1.153	394.30
	-1.105	405.40	-1.104	402.35	-1.104	399.25
	-1.053	409.55	-1.053	406.65	-1.052	403.85
	-0.954	416.20	-1.004	410.50	-1.003	408.05
	-0.905	418.70	-0.954	413.90	-0.951	411.95
	-0.850	416.20	-0.904	417.10	-0.901	415.20
	-0.807	422.70	-0.854	419.60	-0.855	417.95
	-0.755	424.15	-0.804	421.65	-0.804	420.40
	-0.707	424.90	-0.775	423.25	-0.750	422.40
	-0.655	425.10	-0.604	424.85	-0.703	423.60
	-0.600	424.85	-0.553	423.90	-0.654	424.20
	-0.554	423.95	-0.505	422.20	-0.605	424.30
	-1.004	412.85	-0.452	419.40	-0.554	423.50
			-0.404	416.05	-0.503	421.60
			-0.354	411.45	-0.454	419.05
			-0.305	406,20	-0.404	415.30
			-0.253	399.80	-0.355	410.95
			-0.204	392.80	-0.304	405.55
			-1.155	397.55	-0.254	399.30
			-0.705	424.35	-0.204	391.95
			-0.655	424.85	-0.152	383.00

Mercury in aqueous NaN_3 (cont.)

$c/mol\ l^{-1}$	0.0567		0.0787		0.0956	
	E/volts	$\gamma/mN\ m^{-1}$	E/volts	$\gamma/mN\ m^{-1}$	E/volts	$\gamma/mN\ m^{-1}$
	-1.406	360.75	-1.403	359.00	-1.406	356.90
	-1.356	367.80	-1.352	366.40	-1.354	364.55
	-1.303	374.55	-1.301	373.35	-1.302	371.95
	-1.255	381.20	-1.254	379.45	-1.250	378.70
	-1.202	387.25	-1.204	385.45	-1.201	384.65
	-1.154	392.60	-1.152	391.20	-1.150	390.40
	-1.103	397.70	-1.103	396.55	-1.100	395.70
	-1.054	402.45	-1.052	401.50	-1.050	400.55
	-1.003	406.60	-1.002	405.85	-1.000	405.20
	-0.956	410.50	-0.954	409.90	-0.950	409.25
	-0.906	414.00	-0.902	413.50	-0.902	413.00
	-0.852	417.25	-0.854	416.65	-0.851	416.15
	-0.805	419.75	-0.802	419.30	-0.801	418.85
	-0.753	421.70	-0.754	421.35	-0.750	421.10
	-0.704	423.15	-0.703	422.90	-0.700	422.60
	-0.654	424.10	-0.650	423.75	-0.649	423.40
	-0.603	424.10	-0.603	423.85	-0.600	429.40
	-0.553	423.25	-0.554	423.05	-0.551	422.50
	-0.506	421.45	-0.502	421.15	-0.501	420.65
	-0.454	418.75	-0.455	418.55	-0.450	417.75
	-0.404	415.05	-0.404	414.75	-0.401	414.20
	-0.354	410.55	-0.352	410.10	-0.351	409.45
	-0.304	405.20	-0.304	404.85	-0.301	404.05
	-0.254	398.75	-0.254	398.45	-0.250	397.45
	-0.203	391.20	-0.203	390.60	-0.198	389.50
	-0.154	382.70	-0.154	382.15	-0.151	381.05
			-1.104	396.45	-1.405	357.05
					-1.354	364.55

Mercury in aqueous NaN_3 (cont.)

c/mol l^{-1}	0.199		0.395		0.627	
	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}
	-1.404	352.40	-1.405	348.35	-1.394	347.05
	-1.354	360.40	-1.356	355.95	-1.355	353.45
	-1.304	367.60	-1.306	363.45	-1.306	361.10
	-1.254	374.25	-1.255	370.65	-1.254	368.40
	-1.204	380.40	-1.201	377.95	-1.199	375.85
	-1.155	386.10	-1.152	384.00	-1.159	381.15
	-1.105	391.65	-1.104	389.50	-1.094	389.00
	-1.054	397.10	-1.053	395.05	-1.054	393.35
	-1.002	402.15	-1.004	400.10	-1.004	398.30
	-0.954	406.35	-0.954	404.65	-0.954	403.20
	-0.904	410.30	-0.904	408.90	-0.904	407.40
	-0.855	414.05	-0.846	412.90	-0.855	411.20
	-0.804	417.10	-0.804	415.60	-0.804	414.40
	-0.754	419.40	-0.754	418.10	-0.755	416.90
	-0.704	421.10	-0.704	419.80	-0.703	418.70
	-0.65	422.00	-0.654	420.85	-0.654	419.45
	-0.603	422.15	-0.604	420.80	-0.604	419.40
	-0.552	421.30	-0.553	419.80	-0.557	418.40
	-0.503	419.55	-0.504	417.75	-0.505	416.25
	-0.453	416.75	-0.454	414.85	-0.455	413.20
	-0.404	413.10	-0.404	411.00	-0.406	409.50
	-0.353	408.20	-0.354	406.30	-0.355	404.55
	-0.303	402.30	-0.304	400.50	-0.305	398.90
	-0.252	396.05	-0.254	393.80	-0.254	391.90
	-0.203	388.05	-0.203	385.70	-0.205	383.85
	-0.153	379.35	-0.153	376.15	-0.155	374.20
	-0.103	368.10	-0.103	364.35	-0.103	361.55

Mercury in aqueous NaN_3 (cont.)

c/mol l^{-1}	0.814		1.006	
	E/volts	Y/mN m^{-1}	E/volts	Y/mN m^{-1}
	-1.398	344.90	-1.400	343.20
	-1.354	352.15	-1.350	351.20
	-1.302	360.25	-1.300	359.05
	-1.253	367.35	-1.251	366.40
	-1.203	374.35	-1.202	373.25
	-1.154	380.85	-1.151	379.95
	-1.105	386.85	-1.102	386.10
	-1.053	392.70	-1.050	391.90
	-1.006	397.70	-1.000	397.30
	-0.955	402.30	-0.950	402.15
	-0.905	406.75	-0.901	406.35
	-0.854	410.55	-0.851	410.15
	-0.804	414.05	-0.801	413.20
	-0.752	416.55	-0.750	415.60
	-0.704	418.10	-0.700	417.20
	-0.654	418.85	-0.652	417.80
	-0.604	418.65	-0.600	417.60
	-0.553	417.60	-0.551	416.60
	-0.505	415.45	-0.502	414.25
	-0.453	412.15	-0.450	411.05
	-0.404	408.20	-0.400	406.90
	-0.355	403.40	-0.350	401.95
	-0.304	397.40	-0.300	395.80
	-0.254	390.45	-0.250	388.80
	-0.203	382.00	-0.201	380.20
	-0.153	371.95	-0.150	370.20
	-0.105	359.85	-0.100	356.90
			-0.050	339.35

Electrocapillary data for mercury in aqueous sodium formate. T = 25°C

Potential with respect to 0.1 M KCl calomel electrode in contact with 0.1 M solution of salt.

Reference: J. Lawrence and R. Parsons (unpublished)

				γ/mN m^{-1}				
c/mol l^{-1}	0.020	0.050	0.10	0.20	0.50	1.0	2.0	5.0
E/volts								
0.00	374.8	377.0	377.6	377.7	377.7	375.4	373.6	370.3
-0.10	387.5	390.4	391.2	392.4	392.7	390.0	389.0	387.1
-0.20	400.8	403.1	403.9	404.7	404.6	400.6	400.2	400.0
-0.30	411.0	413.9	413.6	414.7	413.9	409.8	409.2	409.0
-0.40	419.8	419.8	420.2	421.3	420.6	417.4	417.2	417.3
-0.45	422.1	423.0	422.4	423.3	422.7	421.1		
-0.50	423.8	425.1	424.0	424.2	423.7	423.3	421.8	421.7
-0.55	425.1	425.6	424.8	424.7	424.4	424.0	423.0	422.6
-0.60	425.0	425.6	424.7	424.7	424.6	424.0	423.4	422.9
-0.65						432.2	422.6	422.6
-0.70	423.4	423.6	422.8	423.0	422.2	421.7	421.1	421.7
-0.80	420.1	419.7	418.6	418.2	417.3	417.1	416.3	417.2
-0.90	414.7	413.7	412.1	411.7	410.4	410.2	409.4	410.6
-1.00	407.1	405.8	404.2	403.1	401.8	401.3	400.5	402.1
-1.10	398.2	396.0	394.6	393.3	391.6	390.9	389.9	391.6
-1.20	387.8	385.0	383.3	381.7	379.5	382.5	377.6	379.4
-1.30	375.9	372.7	370.5	368.4	366.0	365.2	363.8	365.4
-1.40	362.5	358.4	356.1	353.7	350.4	349.8	348.1	349.7
-1.50	347.1	342.8	345.6	337.0	333.6	332.4	330.4	331.9

Electrocapillary data for mercury in aqueous sodium acetate. $T = 25^{\circ}C$

Potential with respect to 0.1 M KCl calomel electrode in contact with 0.1 M solution of salt.

Reference: J. Lawrence and R. Parsons (unpublished)

	γ/mN m^{-1}							
c/mol l^{-1} E/volts	0.020	0.050	0.10	0.20	0.50	1.0	2.0	5.0
0.00	377.0	377.8	373.0	375.9	420.1	371.1	371.4	362.2
-0.10	392.0	392.4	385.6	389.6	391.2	388.9	388.3	382.6
-0.20	404.7	404.9	398.5	400.4	403.9	402.6	401.3	396.1
-0.30	414.3	415.0	408.6	411.6	413.2	412.0	411.0	404.9
-0.40	421.5	421.8	419.1	421.0	420.2	419.0	418.1	412.2
-0.45	423.6	424.0	423.8	423.1				
-0.50	425.3	425.6	425.4	424.2	425.0	422.8	422.1	416.6
-0.55	425.9	426.0	425.6	424.8	425.2	423.6	422.1	417.7
-0.60	425.6	425.5	425.5	424.4	425.2	424.0	423.2	418.0
-0.65		424.8			424.3	423.1	422.7	417.7
-0.70	423.6	423.3	423.3	422.3	423.0	421.4	421.7	416.5
-0.80	419.8	419.5	418.5	417.4	418.1	417.2	417.2	412.7
-0.90	414.2	413.6	412.2	411.2	411.0	410.5	410.5	406.9
-1.00	406.7	405.7	403.6	402.6	402.4	401.5	402.1	399.1
-1.10	406.7	395.3	393.7	392.9	391.9	390.8	391.7	389.6
-1.20	386.4	384.0	382.6	380.8	379.9	378.8	379.5	378.2
-1.30	374.0	371.0	369.4	367.5	366.1	365.1	365.4	365.0
-1.40	360.1	356.9	355.1	352.7	350.8	350.4	349.8	349.9
-1.50	344.8	341.0	338.8	336.3	334.0	332.7	332.5	332.6

Electrocapillary data for mercury in sodium propionate solutions in water.
T = 25°C

Potential with respect to 0.1 M KCl calomel electrode in contact with 0.1 M solution of salt.

Reference: J. Lawrence and R. Parsons (unpublished)

				γ /mN m^{-1}				
c/mol l^{-1}	0.020	0.050	0.10	0.20	0.50	1.0	2.0	4.0
E/volts								
0.00	3772.	376.9	375.0	376.1	372.5	370.3	365.0	360.9
-0.05			384.4					
-0.10	389.1	392.3	391.2	391.9	389.2	387.3	383.1	380.2
-0.20	398.7	405.2	404.9	404.7	401.8	399.9	396.4	393.9
-0.30	412.1	415.0	414.1	414.3	411.7	408.7	403.8	403.1
-0.40	421.3	421.2	420.9	420.8	417.6	415.7	410.7	409.2
-0.45	423.9	423.4	423.2	422.8				
-0.50	425.4	425.2	424.7	424.1	421.5	419.8	415.7	412.7
-0.55	426.1	425.6	425.2	424.6	422.4	420.9	416.9	413.1
-0.60	425.9	425.4	425.0	424.5	422.4	421.3	417.2	413.1
-0.65					421.7	420.8	416.7	412.8
-0.70	424.0	423.6	422.9	422.4	420.6	419.4	415.7	411.6
-0.80	420.2	419.5	418.7	418.2	416.5	415.2	412.1	408.3
-0.90	414.0	412.9	412.2	411.6	410.1	409.0	406.2	403.0
-1.00	406.5	405.1	404.1	403.5	401.5	400.9	398.9	396.0
-1.10	397.6	395.4	394.6	393.5	391.4	390.9	389.6	387.3
-1.20	386.5	384.1	383.2	381.9	379.7	379.2	378.5	376.6
-1.30	374.2	371.4	370.1	368.9	366.1	365.7	365.7	364.3
-1.40	360.2	356.9	355.6	354.2	351.2	350.6	350.6	349.8
-1.50	344.8	341.1	339.3	337.9	334.4	333.6	333.8	333.2

Electrocapillary data for mercury in aqueous solutions of $MgSO_4$

$T = 25^{\circ}C$

Potentials with respect to a mercury - mercurous sulphate electrode in 1.52 mol l^{-1} $MgSO_4$

Reference: J.A. Harrison, J.E.B. Randles and D.J. Schiffrin, J. Electroanal. Chem., 25 (1970) 197.

c/mol l^{-1}	0.1	0.5	1.0	1.5	2.0	2.5
E/volts			γ/mN m^{-1}			
-0.40		390.9	389.5	388.9	387.9	387.2
-0.45	400.3	398.8	397.7	397.4	396.5	395.7
-0.50	406.6	405.6	404.8	404.5	404.1	403.5
-0.55	411.7	411.5	411.0	410.9	409.5	410.5
-0.60	416.2	416.3	415.9	416.1	416.1	416.4
-0.65	419.6	419.8	419.7	420.3	420.6	421.3
-0.70	422.5	422.5	422.7	423.3	423.8	424.7
-0.75	424.1	424.5	424.8	425.4	425.7	427.2
-0.80	425.0	425.6	425.9	426.5	426.9	428.5
-0.85	425.5	425.9	426.1	427.0	427.5	428.9
-0.90	425.4	425.6	426.0	426.6	427.3	428.8
-0.95	424.8	424.6	424.9	425.5	426.4	427.8
-1.00	423.6	423.0	423.4	424.1	424.9	426.1
-1.05	421.7	421.0	421.1	421.7	422.6	423.8
-1.10	419.3	418.5	418.5	418.9	419.7	421.0
-1.15	416.4	415.4	415.2	415.7	416.5	417.7
-1.20	412.9	411.8	411.7	411.9	412.7	413.8
-1.25	409.2	407.8	407.5	407.8	408.5	409.5
-1.30	404.9	403.3	403.0	403.1	403.8	404.8
-1.35	400.3	398.5	398.1	398.1	398.8	399.5
-1.40	395.2	393.2	392.8	392.7	393.1	394.0
-1.45	389.8	387.5	386.9	387.0	387.2	387.9
-1.50	383.8	381.4	380.8	380.6	381.0	381.6
-1.55	377.5	375.0	374.2	373.9	374.2	374.8
-1.60	370.8	368.0	367.3	366.9	367.2	367.7
-1.65	363.6	360.7	359.7	359.4	359.6	360.2
-1.75	348.8	345.0	343.9	343.3	343.4	343.5
-1.85	331.9	327.5	326.2	325.6	325.4	325.3
-1.90	322.8	318.1	316.7	315.9	315.7	315.9

Electrocapillary data for mercury in aqueous solutions of LiCl

T = 25°C

Potentials with respect to a silver-silver chloride electrode in 4 mol l^{-1} KCl with a bridge of 3.05 mol l^{-1} LiCl to the working solution.

Reference: J.A. Harrison, J.E.B. Randles and D.J. Schiffrin, J. Electroanal. Chem., 25 (1970) 197.

c/mol l^{-1}	0.972	2.08	3.05	4.44	5.90	6.94	8.94
E/volts	γ /mN m^{-1}						
-0.44	422.5	420.6	419.2	417.6	415.7	414.9	413.3
-0.48	423.2	421.4	420.7	419.3	417.6	416.9	415.6
-0.52	423.2	421.7	421.4	420.2	418.9	418.2	417.3
-0.56	422.8	421.5	421.4	420.4	419.4	419.0	418.2
-0.60	422.1	420.9	420.7	420.0	419.2	419.1	418.6
-0.64	420.8	419.7	419.3	419.0	418.5	418.5	418.3
-0.68	418.9	417.8	417.4	417.4	417.2	417.2	417.4
-0.72	416.6	415.4	415.1	415.2	415.2	415.4	415.9
-0.76	413.8	412.7	412.4	412.5	412.6	413.1	413.9
-0.80	410.8	409.6	409.2	409.4	409.7	410.3	411.3
-0.84	407.4	406.2	405.7	405.9	406.4	407.0	408.2
-0.88	403.8	402.4	402.0	402.1	402.5	403.2	404.6
-0.92	400.0	398.3	397.7	397.8	398.3	399.1	400.5
-0.96	395.7	393.7	393.1	393.1	393.7	394.4	395.8
-1.00	391.0	389.0	388.1	388.2	388.8	389.4	391.3
-1.04	386.3	384.2	383.1	383.1	383.5	384.3	386.2
-1.08	381.6	379.1	377.8	377.7	377.8	378.6	380.4
-1.12	376.5	373.7	372.1	371.9	372.0	372.7	374.4
-1.15	372.3	369.4	367.9	367.5	367.5	367.9	369.5
-1.22					356.1	356.4	358.0
-1.32					338.5	338.5	339.3

Electrocapillary data for mercury in aqueous 0.795 mol l^{-1} NaF at various temperatures.

Potentials with respect to a cation reversible electrode in the working solution whose potential at 25 °C is taken as equal to that of a 1 mol l^{-1} KCl calomel electrode.

Reference: J..A. Harrison, J.E.B. Randles and D.J. Schiffrin, J. Electroanal. Chem. , 48 (1973) 359.

T/°C	0	15.9	25	38	45
E/volts	γ/mN m^{-1}				
-0.10	382.5		382.2		380.6
-0.05	390.1		389.8	388.3	388.3
0.00	397.1		396.7	396.3	394.8
0.05	403.2		402.7	401.8	400.8
0.10	408.9		408.1	407.5	406.1
0.15	413.8	414.1	412.8	411.4	410.4
0.20	418.1	417.8	416.7	415.4	414.2
0.25	420.7	421.4	420.0	418.2	417.3
0.30	424.6	424.1	422.7	420.8	419.7
0.35	426.7	426.1	424.6	422.7	421.6
0.40	428.3	427.2	425.9	424.0	422.7
0.45	429.0	427.9	426.4	424.4	423.1
0.50	429.1	427.9	426.4	424.3	423.1
0.55	428.5	427.3	425.7	423.9	422.5
0.60	427.2	426.2	424.5	422.6	421.3
0.65	425.3	424.5	422.7	420.9	419.5
0.70	422.9	422.0	420.4	418.6	417.5
0.75	420.0	419.2	417.6	416.1	414.7
0.80	416.6	416.0	414.3	412.5	411.5
0.85	412.6	412.1	410.6	408.8	407.9
0.90	408.4	407.7	406.4	404.9	404.1
0.95	403.6	402.9	401.8	400.1	399.7
1.00	398.5	398.0	396.9	395.1	395.2
1.05	393.1	392.3	391.6	390.6	390.2
1.10	387.2	386.1	387.0	384.5	384.4
1.15	380.8	380.5	380.9		378.5
1.20	374.3	374.1	374.4		372.2
1.25	367.0	366.2			365.4
1.30	359.4	360.1			358.3
1.35	351.6	352.6			350.7

Electrocapillary data for mercury in aqueous 0.1 mol l^{-1} NaF at various temperatures.

Potentials with respect to a cation reversible electrode in the working solution whose potential at 25°C is taken as equal to that of a 1 mol l^{-1} KCl calomel electrode.

Reference: J.A. Harrison, J.E.B. Randles and D.J. Schiffrin, J. Electroanal. Chem, 48 (1973) 359

T/°C	1.5	14.1	25.2	32.0	36.0	37.3	52.0	65.2
E/volts				γ/mN m^{-1}				
-0.15	413.1	412.0	412.3					
-0.20	417.2	416.7	416.0					
-0.25	420.8	420.3	419.3	418.7	413.5		413.8	413.6
-0.30	423.8	423.0	421.8	420.8	417.2	419.3	416.3	414.6
-0.35	425.7	425.1	423.6	423.1	420.7		418.6	415.5
-0.40	427.1	426.2	424.9	424.0	422.8	423.4	419.9	415.7
-0.45	428.0	427.0	425.5	424.8	423.9		420.5	415.9
-0.50	428.0	427.1	425.5	424.2	424.3	424.2	420.4	415.6
-0.55	427.6	426.6	425.0	424.0	423.8		420.0	414.6
-0.60	426.4	425.6	423.9	422.7	422.7	422.7	419.0	412.5
-0.65	424.8	423.9	422.3	421.5	421.2		417.5	410.6
-0.70	422.7	421.7	420.3	419.4	419.2	419.2	415.6	408.3
-0.75	419.9	419.1	417.6	417.0	416.6		413.2	405.7
-0.80	416.7	415.9	414.6	413.7	413.7	413.8	410.5	402.4
-0.85	413.1	412.3	411.2	410.5	410.3		407.3	399.3
-0.90	408.9	408.3	407.3	406.4	406.4		403.7	395.4
-0.95	404.3	403.8	402.9	402.0	402.1		399.7	392.8
-1.00	399.4	399.0	398.3		397.6		395.3	388.6
-1.05	394.0	393.8	393.1		392.6		390.6	383.6
-1.10	387.8	388.1	387.7		387.2		385.1	378.5
-1.15	382.0	382.1	381.8		381.2		379.5	373.3

Electrocapillary data for mercury in aqueous solutions of KI   T = 5°C

Potential with respect to a silver-silver iodide electrode reversible to the working solution

Reference: W. Anderson and R. Parsons, Proc. II^nd Int. Congr. Surf. Act. III 45 (1957)

$c/\text{mol l}^{-1}$	0.00232		0.00663		0.0117		0.0358	
$\sigma\ /\mu\text{C cm}^{-2}$	E/volts	$\gamma/\text{mN m}^{-1}$	E/volts	$\gamma/\text{mN m}^{-1}$	E/volts	$\gamma/\text{mN m}^{-1}$	E/volts	$\gamma/\text{mN m}^{-1}$
-10					-0.710	401.2	-0.679	399.0
- 9			-0.739	401.1	-0.655	406.5	-0.624	404.2
- 8	-0.763	405.3	-0.661	407.7	-0.601	411.1	-0.567	409.2
- 7	-0.685	410.7	-0.578	414.0	-0.546	415.2	-0.512	413.2
- 6	-0.611	415.8	-0.503	418.6	-0.492	418.5	-0.468	416.1
- 5	-0.546	419.3	-0.452	421.4	-0.450	420.8	-0.439	418.0
- 4	-0.481	422.2	-0.413	423.2	-0.420	422.2	-0.415	418.8
- 3	-0.435	423.8	-0.386	424.1	-0.398	422.9	-0.395	419.6
- 2	-0.393	425.0	-0.365	424.6	-0.378	423.6	-0.376	420.0
-1	-0.355	425.4	-0.342	425.0	-0.359	423.8	-0.357	420.2
0	-0.328	425.6	-0.325	425.1	-0.340	424.0	-0.346	420.3
1	-0.300	425.4	-0.300	424.9	-0.318	423.8	-0.321	420.2
2	-0.277	425.1	-0.279	424.5	-0.298	423.5	-0.304	419.8
3	-0.259	424.1	-0.257	423.9	-0.278	423.1	-0.287	419.3
4	-0.243	424.2	-0.236	423.1	-0.258	422.3	-0.270	418.8
5	-0.228	423.5	-0.215	422.1	-0.238	421.3	-0.254	418.1
6	-0.215	422.8	-0.193	420.8	-0.218	420.2	-0.239	417.3
7	-0.204	422.0	-0.171	419.5	-0.198	418.9	-0.220	416.3
8	-0.192	421.0	-0.150	418.0	-0.178	417.4	-0.214	415.4
9	-0.180	420.0	-0.129	416.3	-0.158	415.8	-0.202	414.3
10	-0.170	419.2	-0.106	414.1	-0.138	413.9	-0.190	413.2
12	-0.150	417.3	-0.081	411.3			-0.168	410.7
14	-0.130	414.5	-0.065	409.3			-0.147	407.6
16	-0.117	412.2	-0.056	407.7			-0.128	405.0
18	-0.106	410.3					-0.112	402.1
20							-0.096	399.2

Electrocapillary data for Hg in KI (cont.) $T = 5^{\circ}C$

c/mol l^{-1}	0.903		0.958		2.15	
σ/μC cm^{-2}	E/volts	γ /mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ /mN m^{-1}
-10	-0.652	388.7	-0.635	389.9	-0.659	384.2
- 9	-0.620	391.8	-0.607	392.6	-0.622	387.6
- 8	-0.589	394.4	-0.579	395.0	-0.591	390.1
- 7	-0.558	396.8	-0.547	397.3	-0.567	392.0
- 6	-0.527	398.7	-0.510	399.8	-0.545	393.3
- 5	-0.495	400.5	-0.488	401.0	-0.525	394.5
- 4	-0.472	401.6	-0.470	401.8	-0.505	395.3
- 3	-0.453	402.5	-0.454	402.3	-0.485	396.0
- 2	-0.434	402.8	-0.438	402.6	-0.465	396.5
- 1	-0.416	402.9	-0.423	402.7	-0.448	396.6
0	-0.399	403.0	-0.406	402.8	-0.430	396.7
1	-0.381	402.8	-0.390	402.7	-0.412	396.6
2	-0.363	402.6	-0.375	402.5	-0.395	396.3
3	-0.347	402.0	-0.360	402.2	-0.378	396.0
4	-0.330	401.3	-0.344	401.6	-0.361	395.3
5	-0.315	400.5	-0.329	400.9	-0.345	394.7
6	-0.299	399.5	-0.313	400.0	-0.330	393.8
7	-0.284	398.7	-0.297	398.8	-0.314	392.7
8	-0.268	397.5	-0.282	397.8	-0.298	391.5
9	-0.254	396.4	-0.265	396.2	-0.279	389.9
10	-0.244	395.4	-0.251	394.8	-0.263	388.4
12	-0.227	393.5	-0.229	392.3	-0.235	385.2
14	-0.214	391.9	-0.210	389.8	-0.211	382.2
16	-0.202	390.2	-0.195	387.5	-0.194	379.5
18	-0.196	389.1	-0.183	385.4	-0.180	377.2
20	-0.190	388.0	-0.170	383.0	-0.174	375.0

Electrocapillary data for Hg in KI (cont.) T = 5°C

c/mol l^{-1}	0.0503		0.0775		0.0928		0.180	
σ/μC cm^{-2}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}
-10	-0.602	404.5	-0.635	400.7	-0.643	398.8	-0.633	397.7
- 9	-0.563	408.3	-0.604	403.6	-0.604	402.8	-0.597	401.2
- 8	-0.528	411.3	-0.564	407.0	-0.565	406.1	-0.562	404.2
- 7	-0.497	413.6	-0.514	410.7	-0.526	409.1	-0.526	406.8
- 6	-0.470	415.3	-0.489	412.3	-0.485	411.3	-0.485	409.6
- 5	-0.443	416.8	-0.464	413.7	-0.453	413.5	-0.459	411.0
- 4	-0.420	417.8	-0.440	414.7	-0.435	414.3	-0.443	411.7
- 3	-0.398	418.5	-0.418	415.5	-0.417	415.0	-0.428	412.2
- 2	-0.380	418.8	-0.395	416.0	-0.399	415.4	-0.412	412.7
- 1	-0.363	419.2	-0.370	416.5	-0.382	415.6	-0.396	412.9
0	-0.343	419.3	-0.350	416.6	-0.363	415.7	-0.380	413.0
1	-0.328	419.2	-0.336	416.5	-0.349	415.6	-0.364	412.9
2	-0.313	418.9	-0.320	416.3	-0.332	415.5	-0.347	412.6
3	-0.297	418.4	-0.305	415.8	-0.316	414.9	-0.332	412.2
4	-0.281	417.8	-0.289	415.2	-0.302	414.1	-0.317	411.6
5	-0.271	417.4	-0.274	414.6	-0.285	413.6	-0.301	410.8
6	-0.251	416.2	-0.259	413.8	-0.271	412.8	-0.285	410.0
7	-0.237	415.3	-0.243	412.7	-0.257	411.8	-0.269	409.0
8	-0.223	414.3	-0.230	411.8	-0.242	410.6	-0.253	407.8
9	-0.209	412.9	-0.220	410.9	-0.229	409.4	-0.240	406.7
10	-0.192	411.3	-0.210	409.9	-0.215	408.0	-0.228	405.5
12	-0.165	408.3	-0.189	407.6	-0.189	405.1	-0.204	402.7
14	-0.143	405.7	-0.167	404.8	-0.165	402.0	-0.180	399.6
16	-0.124	402.7	-0.148	401.9	-0.147	399.0	-0.160	396.6
18	-0.110	400.3	-0.134	399.4	-0.126	395.7	-0.140	393.0
20	-0.098	398.0	-0.123	397.2	-0.114	393.2	-0.122	389.5

Electrocapillary data for mercury in aqueous solutions of KI (cont.)

$T = 25^{\circ}C$

$c/\mathrm{mol\ l^{-1}}$	0.00232		0.00663		0.0117		0.0139	
$\sigma/\mu\mathrm{C\ cm^{-2}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$
-10			-0.776	398.6	-0.760	396.7		
- 9			-0.729	403.0	-0.685	403.8	-0.691	403.3
- 8			-0.678	407.3	-0.611	410.0	-0.634	408.3
- 7	-0.673	410.6	-0.622	411.5	-0.560	413.7	-0.587	411.7
- 6	-0.613	414.7	-0.558	415.5	-0.506	417.2	-0.533	415.2
- 5	-0.553	417.7	-0.505	418.5	-0.460	419.7	-0.469	418.7
- 4	-0.491	420.5	-0.459	420.7	-0.435	420.8	-0.426	420.7
- 3	-0.427	422.7	-0.417	422.1	-0.402	421.9	-0.400	421.5
- 2	-0.379	424.0	-0.382	423.0	-0.382	422.4	-0.375	422.1
- 1	-0.358	424.4	-0.350	423.4	-0.364	422.6	-0.360	422.3
0	-0.337	424.6	-0.331	423.5	-0.340	422.7	-0.350	422.4
1	-0.314	424.4	-0.303	423.3	-0.324	422.6	-0.331	422.3
2	-0.290	424.0	-0.285	423.0	-0.305	422.4	-0.308	421.9
3	-0.266	423.3	-0.270	422.6	-0.285	421.8	-0.289	421.4
4	-0.240	422.4	-0.255	422.0	-0.266	421.2	-0.277	421.1
5	-0.212	421.1	-0.239	421.3	-0.247	420.2	-0.265	420.4
6	-0.180	419.2	-0.224	420.5	-0.226	419.0	-0.240	419.0
7	-0.150	417.3	-0.209	419.4	-0.207	417.7	-0.220	417.7
8	-0.115	414.7	-0.193	418.2	-0.188	416.3	-0.200	416.3
9			-0.177	416.9	-0.169	414.8	-0.188	415.1
10			-0.162	415.3			-0.174	413.7
12			-0.132	411.9			-0.154	411.6
14							-0.135	409.1
16							-0.118	406.3
18							-0.100	402.3
20							-0.082	399.7

Electrocapillary data for mercury in aqueous KI (cont.)

T = 25°C

$c/\mathrm{mol\ l^{-1}}$	0.0358		0.0503		0.0775		0.0928	
$\sigma/\mu\mathrm{C\ cm^{-2}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$
-10	-0.676	399.3	-0.663	399.3	-0.670	396.3	-0.664	396.7
- 9	-0.625	404.1	-0.614	404.0	-0.625	400.7	-0.598	402.7
- 8	-0.576	408.3	-0.564	408.2	-0.569	405.5	-0.549	406.8
- 7	-0.525	412.2	-0.514	411.3	-0.495	411.1	-0.507	409.8
- 6	-0.476	415.4	-0.477	414.2	-0.469	412.8	-0.475	411.8
- 5	-0.446	417.1	-0.450	415.7	-0.447	414.0	-0.449	413.2
- 4	-0.424	418.0	-0.431	416.6	-0.429	414.8	-0.425	414.1
- 3	-0.405	418.6	-0.412	417.3	-0.410	415.3	-0.408	414.6
- 2	-0.385	419.0	-0.392	417.7	-0.394	415.8	-0.393	415.0
- 1	-0.367	419.3	-0.373	418.0	-0.380	416.0	-0.378	415.3
0	-0.350	419.4	-0.350	418.1	-0.365	416.1	-0.363	415.4
1	-0.333	419.3	-0.337	418.0	-0.354	416.0	-0.349	415.3
2	-0.315	419.0	-0.319	417.7	-0.344	415.8	-0.333	415.0
3	-0.298	418.6	-0.302	417.3	-0.332	415.5	-0.318	414.7
4	-0.282	418.1	-0.285	416.7	-0.319	415.0	-0.304	414.0
5	-0.266	417.4	-0.269	416.0	-0.304	414.3	-0.289	413.4
6	-0.253	416.7	-0.254	415.1	-0.287	413.3	-0.275	412.6
7	-0.240	415.7	-0.239	414.2	-0.269	412.0	-0.259	411.7
8	-0.228	414.8	-0.225	413.1	-0.249	410.5	-0.245	410.7
9	-0.215	413.6	-0.214	412.2	-0.231	409.1	-0.230	409.3
10	-0.202	412.3	-0.203	411.1	-0.215	407.5	-0.215	408.0
12	-0.180	409.9	-0.183	408.8	-0.183	403.8	-0.185	404.7
14	-0.157	406.8	-0.164	406.2	-0.156	400.3	-0.161	401.6
16	-0.137	403.8	-0.147	403.3	-0.135	396.9	-0.143	398.7
18	-0.120	400.8	-0.132	400.8	-0.120	394.2	-0.130	396.5
20	-0.104	397.6	-0.117	398.0	-0.110	392.	-0.122	394.7

Electrocapillary data for mercury in aqueous KI (cont.)

T = 25°C

c/mol l^{-1}	0.180		0.455		0.463		0.903	
σ /μC cm^{-2}	E/volts	γ/mN m^{-1}	E/volts	γ /mN m^{-1}	E/volts	γ /mN m^{-1}	E/volts	γ/mN m^{-1}
-10	-0.649	395.8	-0.638	393.1	-0.634	393.6	-0.617	391.0
- 9	-0.607	400.1	-0.605	396.2	-0.606	396.2	-0.592	393.4
- 8	-0.564	403.6	-0.570	399.0	-0.568	399.5	-0.569	395.3
- 7	-0.522	406.8	-0.535	401.7	-0.535	402.0	-0.544	397.0
- 6	-0.488	409.0	-0.500	404.0	-0.510	403.6	-0.521	398.5
- 5	-0.469	410.1	-0.480	405.2	-0.485	405.0	-0.499	399.8
- 4	-0.450	411.0	-0.465	405.7	-0.465	406.0	-0.476	400.8
- 3	-0.431	411.7	-0.449	406.3	-0.448	406.5	-0.456	401.5
- 2	-0.414	412.0	-0.432	406.7	-0.432	406.9	-0.437	402.0
- 1	-0.397	412.1	-0.415	406.9	-0.414	407.1	-0.417	402.1
0	-0.380	412.2	-0.400	407.0	-0.394	407.2	-0.399	402.2
1	-0.364	412.1	-0.383	406.9	-0.380	407.1	-0.380	402.1
2	-0.348	411.9	-0.367	406.7	-0.363	406.9	-0.363	401.8
3	-0.332	411.5	-0.351	406.4	-0.346	406.5	-0.346	401.5
4	-0.317	411.0	-0.335	405.8	-0.330	405.9	-0.331	401.0
5	-0.302	410.2	-0.320	405.2	-0.314	405.2	-0.315	400.2
6	-0.287	409.4	-0.303	404.3	-0.295	404.1	-0.300	399.3
7	-0.273	408.5	-0.287	403.2	-0.280	403.0	-0.287	398.4
8	-0.259	407.5	-0.272	402.0	-0.262	401.6	-0.273	397.4
9	-0.245	406.3	-0.255	400.5	-0.245	400.0	-0.260	396.3
10	-0.232	405.2	-0.239	399.0	-0.229	398.5	-0.248	395.2
12	-0.206	402.3	-0.205	395.2	-0.202	395.6	-0.223	392.5
14	-0.185	399.5	-0.180	391.9	-0.184	393.0	-0.201	389.8
16	-0.169	397.0	-0.170	390.2	-0.170	391.0		
18	-0.156	394.8	-0.158	388.2	-0.160	389.2		
20	-0.148	393.1	-0.146	385.8	-0.150	387.3		

Electrocapillary data for mercury in aqueous KI (cont.)

$T = 25^{\circ}C$

$^{c}/$mol l^{-1}	0.958		0.960		2.15	
$^{\sigma}/_{\mu}$C cm^{-2}	E/volts	γ/mN m^{-1}	E/volts	γ /mN m^{-1}	E/volts	γ /mN m^{-1}
-10	-0.640	388.9	-0.647	387.9	-0.670	382.9
- 9	-0.610	391.3	-0.604	391.8	-0.635	386.2
- 8	-0.585	394.0	-0.567	395.0	-0.605	388.8
- 7	-0.550	396.6	-0.538	397.0	-0.579	390.7
- 6	-0.515	398.8	-0.514	398.6	-0.556	392.2
- 5	-0.497	399.8	-0.492	399.8	-0.535	393.2
- 4	-0.480	400.5	-0.475	400.6	-0.514	394.3
- 3	-0.463	401.0	-0.455	401.2	-0.493	395.0
- 2	-0.446	401.4	-0.437	401.6	-0.473	395.4
- 1	-0.428	401.6	-0.420	401.7	-0.453	395.6
0	-0.413	401.7	-0.402	401.8	-0.434	395.7
1	-0.395	401.6	-0.384	401.7	-0.414	395.6
2	-0.378	401.4	-0.367	401.4	-0.395	395.3
3	-0.361	401.2	-0.351	400.9	-0.377	394.8
4	-0.345	400.5	-0.336	400.3	-0.360	394.2
5	-0.328	399.8	-0.320	399.7	-0.342	393.1
6	-0.310	398.8	-0.305	398.8	-0.325	392.0
7	-0.294	397.7	-0.290	397.7	-0.308	390.8
8	-0.277	396.5	-0.275	396.5	-0.292	389.6
9	-0.260	395.0	-0.261	395.2	-0.276	388.3
10	-0.245	393.4	-0.248	393.9	-0.260	386.7
12	-0.221	390.7	-0.220	391.0	-0.232	383.7
14	-0.200	388.0	-0.195	387.5	-0.207	380.5
16	-0.185	385.7	-0.172	384.3	-0.188	377.7
18	-0.169	383.2	-0.148	380.3	-0.177	375.8
20	-0.151	378.9	-0.135	377.7	-0.168	374.0

Electrocapillary data for mercury in aqueous solutions of KI (contd.)

$T = 45^{\circ}C$

c/mol l^{-1}	0.00232		0.00584		0.0117		0.0139	
$\sigma/_{\mu}$C cm^{-2}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}
-10			-0.749	399.7	-0.782	394.9	-0.628	405.4
- 9	-0.783	402.2	-0.715	403.1	-0.725	400.5	-0.607	407.2
- 8	-0.743	405.6	-0.679	406.1	-0.670	405.0	-0.583	409.3
- 7	-0.699	408.8	-0.638	409.3	-0.613	409.3	-0.554	411.4
- 6	-0.650	412.0	-0.590	412.2	-0.558	412.9	-0.520	413.7
- 5	-0.594	415.0	-0.536	415.2	-0.500	416.1	-0.476	416.0
- 4	-0.528	417.8	-0.469	418.2	-0.446	418.5	-0.422	418.4
- 3	-0.453	420.5	-0.410	420.2	-0.410	419.8	-0.384	419.8
- 2	-0.400	421.8	-0.380	421.0	-0.386	420.3	-0.364	420.3
- 1	-0.370	422.3	-0.357	421.3	-0.367	420.7	-0.340	420.6
0	-0.346	422.4	-0.342	421.4	-0.347	420.8	-0.312	420.8
1	-0.318	422.3	-0.312	421.2	-0.329	420.7	-0.300	420.7
2	-0.291	421.8	-0.290	420.7	-0.310	420.4	-0.280	420.4
3	-0.260	421.0	-0.267	419.9	-0.290	420.0	-0.260	419.8
4	-0.230	419.8	-0.245	419.0	-0.270	419.3	-0.243	419.2
5	-0.196	418.3	-0.223	418.0	-0.250	418.2	-0.225	418.3
6			-0.200	416.8	-0.230	417.1	-0.209	417.3
7			-0.177	415.4	-0.211	415.8	-0.192	416.2
8			-0.155	413.8	-0.192	414.4	-0.177	415.2
9			-0.132	411.8	-0.173	412.8	-0.162	414.0
10			-0.110	409.5	-0.159	411.4	-0.148	412.7
12					-0.135	408.8	-0.129	410.7
14					-0.118	406.3	-0.115	408.8
16					-0.105	404.3	-0.107	407.2
18					-0.095	402.5	-0.099	406.0
20					-0.090	401.4	-0.095	405.0

Electrocapillary data for mercury in aqueous KI (cont.)

T = 45°C

$c/\text{mol l}^{-1}$	0.0358		0.0503		0.0775		0.0928	
$\sigma/\mu\text{C cm}^{-2}$	E/volts	$\gamma/\text{mN m}^{-1}$	E/volts	$\gamma/\text{mN m}^{-1}$	E/volts	$\gamma/\text{mN m}^{-1}$	E/volts	$\gamma/\text{mN m}^{-1}$
-10	-0.708	396.3	-0.710	394.5	-0.678	395.9	-0.693	392.8
- 9	-0.663	400.5	-0.658	399.4	-0.639	399.7	-0.632	398.7
- 8	-0.610	405.1	-0.596	404.7	-0.583	404.6	-0.572	403.8
- 7	-0.553	409.4	-0.533	409.3	-0.511	409.8	-0.511	408.3
- 6	-0.507	412.5	-0.480	412.7	-0.470	412.5	-0.482	410.3
- 5	-0.475	414.2	-0.454	413.8	-0.448	413.8	-0.462	411.4
- 4	-0.449	415.3	-0.431	415.0	-0.432	414.5	-0.442	412.2
- 3	-0.425	416.2	-0.412	415.7	-0.417	415.0	-0.423	412.8
- 2	-0.400	416.9	-0.396	416.1	-0.400	415.4	-0.404	413.3
- 1	-0.378	417.2	-0.379	416.4	-0.385	415.7	-0.382	413.6
0	-0.357	417.3	-0.359	416.5	-0.370	415.8	-0.360	413.7
1	-0.333	417.2	-0.347	416.4	-0.350	415.7	-0.348	413.6
2	-0.310	416.8	-0.330	416.2	-0.335	415.4	-0.333	413.4
3	-0.294	416.4	-0.313	415.7	-0.322	415.1	-0.316	412.9
4	-0.280	415.8	-0.297	415.1	-0.312	414.7	-0.300	412.3
5	-0.268	415.3	-0.280	414.4	-0.304	414.3	-0.285	411.8
6	-0.255	414.6	-0.264	413.5	-0.295	413.8	-0.270	410.8
7	-0.242	413.8	-0.248	412.4	-0.285	413.2	-0.255	409.8
8	-0.229	412.7	-0.230	411.0	-0.270	412.0	-0.240	408.6
9	-0.217	411.7	-0.214	409.6	-0.247	410.0	-0.228	407.5
10	-0.204	410.3	-0.200	408.3	-0.225	408.0	-0.216	406.3
12	-0.180	407.6	-0.180	406.0	-0.181	404.0	-0.194	403.8
14	-0.159	405.0	-0.163	403.9	-0.164	400.5	-0.173	401.1
16	-0.138	402.0	-0.149	401.6	-0.143	397.4	-0.154	398.2
18	-0.118	398.7	-0.136	399.4	-0.129	394.9	-0.136	395.2
20	-0.099	394.8	-0.125	397.3	-0.125	392.5	-0.121	392.3

Electrocapillary data for mercury in aqueous KI (cont.)

$T = 45^{\circ}C$

c/mol l^{-1}	0.180		0.455		0.463		0.903	
σ/μC cm^{-2}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}	E/volts	γ/mN m^{-1}
-10	-0.662	395.2	-0.660	391.0	-0.656	391.0	-0.652	388.3
- 9	-0.627	398.3	-0.624	394.3	-0.610	395.3	-0.634	389.8
- 8	-0.587	401.3	-0.587	397.5	-0.574	397.3	-0.610	391.9
- 7	-0.546	404.9	-0.551	400.0	-0.542	400.7	-0.580	394.2
- 6	-0.505	407.5	-0.515	402.4	-0.515	402.5	-0.543	396.6
- 5	-0.482	408.9	-0.487	404.0	-0.490	403.8	-0.510	398.4
- 4	-0.465	409.5	-0.469	404.8	-0.470	404.7	-0.489	399.5
- 3	-0.448	410.1	-0.450	405.5	-0.450	405.4	-0.470	400.0
- 2	-0.430	410.4	-0.434	405.7	-0.431	405.8	-0.450	400.4
- 1	-0.413	410.7	-0.417	406.0	-0.413	406.1	-0.431	400.6
0	-0.397	410.8	-0.398	406.1	-0.395	406.2	-0.412	400.7
1	-0.381	410.7	-0.382	406.0	-0.379	406.1	-0.394	400.6
2	-0.365	410.4	-0.365	405.8	-0.362	405.8	-0.375	400.3
3	-0.350	410.0	-0.349	405.5	-0.345	405.3	-0.357	399.9
4	-0.336	409.5	-0.332	404.8	-0.330	404.8	-0.342	399.3
5	-0.321	408.8	-0.315	403.9	-0.315	404.0	-0.325	398.5
6	-0.308	408.0	-0.298	402.9	-0.299	403.2	-0.309	397.5
7	-0.293	407.1	-0.280	401.5	-0.284	403.3	-0.292	396.5
8	-0.280	406.1	-0.264	400.3	-0.270	401.2	-0.278	395.5
9	-0.266	405.0	-0.247	398.9	-0.255	400.0	-0.262	394.1
10	-0.250	404.0	-0.229	397.3	-0.242	398.6	-0.247	392.7
12	-0.223	400.3	-0.200	393.9	-0.217	395.7	-0.225	390.3
14	-0.200	397.5	-0.181	391.5	-0.197	393.2	-0.209	388.2
16	-0.182	394.6	-0.166	389.3	-0.180	390.7	-0.195	386.1
18	-0.161	392.4	-0.152	387.0	-0.167	388.0	-0.184	384.2
20	-0.155	389.5	-0.140	384.5	-0.154	385.0	-0.175	382.5

Electrocapillary data for mercury in aqueous KI (cont.)

T = 45°C

$c/\mathrm{mol\ l^{-1}}$	0.958		2.15	
$\sigma/\mu\mathrm{C\ cm^{-2}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$	E/volts	$\gamma/\mathrm{mN\ m^{-1}}$
-10	-0.632	389.8	-0.671	382.0
- 9	-0.619	391.1	-0.642	384.8
- 8	-0.585	394.1	-0.607	387.7
- 7	-0.560	396.0	-0.568	390.6
- 6	-0.539	397.3	-0.543	392.3
- 5	-0.516	398.6	-0.525	393.2
- 4	-0.495	399.5	-0.510	393.8
- 3	-0.475	400.2	-0.494	394.4
- 2	-0.455	400.7	-0.479	394.7
- 1	-0.435	401.1	-0.464	395.0
0	-0.418	401.2	-0.448	395.1
1	-0.396	401.1	-0.430	395.0
2	-0.376	400.5	-0.414	394.8
3	-0.358	400.0	-0.398	394.3
4	-0.342	399.3	-0.381	393.8
5	-0.329	398.7	-0.365	393.0
6	-0.315	398.0	-0.348	392.0
7	-0.300	397.1	-0.328	390.8
8	-0.289	396.3	-0.310	389.3
9	-0.275	395.0	-0.290	387.6
10	-0.265	394.0	-0.272	385.8
12	-0.242	391.5	-0.245	382.8
14	-0.221	388.9	-0.225	380.3
16	-0.205	386.3	-0.209	377.9
18	-0.190	383.9	-0.192	375.0
20	-0.178	381.5	-0.177	372.3

Electrocapillary data on mercury in HCl in water: at 0°C [(a)] and 25°C, 40°C and 60°C [(b)] - given are interfacial tension in $mN\ m^{-1}$, potential E in Volts.
The reference electrode SCE aq. was constantly at the same temperature as the system studied.

References: (a) S. Minc and M. Jurkiewicz-Herbich. J. Electroanal. Chem. 40 (1972) 229.
(b) S. Minc and M. Jurkiewicz-Herbich. J. Electroanal. Chem. 34 (1972) 351.

	$\gamma/mN\ m^{-1}$															
$c/mol\ l^{-1}$	0.01 M				0.1 M				0.2 M				0.4 M			
T/°C	0	25	40	60	0	25	40	60	0	25	40	60	0	25	40	60
-E/V																
0.200	426.5	419.0	410.1	409.4	418.6	413.0	408.4	407.4	411.5	408.3	407.8	406.8	407.5	404.6	403.9	402.1
0.250																
0.300	430.8	424.9	419.0	417.7	426.0	421.2	417.3	415.6	417.0	417.7	416.4	414.7	418.1	415.4	413.7	411.7
0.350	431.8	426.7	421.8	420.2	428.4	423.9	420.1	418.3	422.0	420.9	419.3	416.2	422.3	419.1	417.1	415.0
0.400	432.0	426.8	423.6	421.7	429.7	426.4	422.0	419.9	425.7	423.0	421.2	418.8	425.5	421.8	419.4	417.3
0.450	431.6	426.3	424.5	422.4	430.0	426.5	423.0	420.8	428.3	424.3	422.1	419.6	427.8	423.4	420.8	418.5
0.500	430.7	425.7	424.6	422.5	429.2	426.4	423.2	420.9	429.6	424.5	422.2	419.7	429.0	423.9	421.1	418.7
0.550	429.3	424.7	424.0	422.0	427.3	425.8	422.7	420.3	430.0	423.8	421.6	419.1	429.1	423.5	420.6	418.3
0.600	427.6	422.3	422.9	420.9	424.7	424.6	421.6	419.1	429.8	422.7	420.4	418.0	428.4	422.4	419.4	416.9
0.650	425.4	420.4					419.9						426.8	420.5	417.5	415.1
0.700	422.7	417.1	418.9	417.4	417.7	420.4	417.9	415.3	422.5	417.5	416.3	414.2	424.6	418.0	415.0	413.4
0.800	415.3	412.0	413.0	412.1	408.6	414.1	412.2	409.7	415.0	410.7	410.5	408.0	417.9	411.4	408.6	407.3
0.900	407.1		405.5	405.2	398.6	406.0	404.7	402.3	405.6	401.8	402.6	401.0	409.6	402.9	400.5	401.6
1.000				396.9	386.8		395.6	393.3	394.5	381.3			399.7	392.6	390.8	390.1

Electrocapillary data in HCl (continued) γ/mN m^{-1}

					γ/mN m^{-1}							
c/mol l^{-1}	0.6 M				1.0 M				1.4 M			
T/^{o}C	0	25	40	60	0	25	40	60	0	25	40	60
-E/V												
0.200	403.5	401.2	401.3	399.1	403.8	397.7	397.4	397.4	398.6	397.0	395.2	391.8
0.250												
0.300	416.2	412.4	411.7	410.5	415.2	409.7	407.8	408.0	410.4	408.5	406.8	403.8
0.350	421.0	415.5	415.4	414.3	419.6	413.3	413.0	412.0	415.0	413.2		408.3
0.400	424.7	419.5	418.0	416.0	423.1	416.8	416.2	415.0	420.8	416.6	414.5	411.7
0.450	427.2	421.5	419.5	417.5	425.7	419.5	418.1	417.0	423.8	419.0	416.8	414.2
0.500	428.7	421.9	420.0	417.9	427.3	420.6	419.3	417.5	425.8	420.3	418.3	415.5
0.550	429.0	421.7	419.7	417.5	427.8	421.1	419.2	417.3	426.9	420.8	418.7	415.9
0.600	428.4	421.0	418.5	416.4	427.3	420.2	418.4	416.2	426.9	420.3	418.3	415.4
0.650					425.7	418.7			425.9	418.9	416.8	414.0
0.700	424.6	417.3	414.3	412.2	423.2	416.4	414.2	413.0	423.9	417.6	414.6	411.9
0.800	418.2	411.2	408.0	405.6	416.0	410.1	407.0	406.5	417.8	410.7	407.7	405.6
0.900	409.7	404.0	399.7	397.2	406.7	402.0	400.2	399.0	409.2	401.3	398.2	397.1
1.000	399.4	393.2	389.7	387.0	395.6	393.5	384.3		398.5	389.8	386.8	386.6
1.100				375.2								

Electrocapillary data on mercury electrode in 0.001 M HCl at different temperatures - given are the interfacial tension in mN m^{-1}, potential in V vs RCE.

Reference: E.W. Hermann. Ph.D. Dissertation, University of Carleton, 1970.

	γ /mN m^{-1}								
T/°C	0.4	4.9	10.1	15.2	19.8	24.9	30.0	35.0	39.9
-E/V									
0.050	378.1	377.7	376.2	374.8	373.4	372.7	371.9	370.6	369.1
0.100	386.5	386.8	384.8	383.6	382.4	381.7	380.8	379.8	378.4
0.150	393.9	394.6	392.2	391.0	390.0	389.3	388.3	387.6	386.3
0.200	400.7	401.6	298.9	397.8	396.8	396.0	395.0	394.5	393.2
0.250	406.7	407.8	405.0	403.9	402.9	402.0	401.1	400.6	399.3
0.300	411.9	413.1	410.2	409.2	408.2	407.3	406.3	405.8	404.4
0.350	416.2	417.5	414.6	413.5	412.6	411.6	410.7	410.1	408.6
0.400	419.7	420.9	418.0	417.0	416.1	415.1	414.1	413.4	411.9
0.450	422.4	423.6	420.7	419.7	418.8	417.8	416.8	415.9	414.5
0.500	424.6	425.6	422.8	421.8	421.0	419.9	418.8	417.9	416.5
0.550	426.3	427.1	424.5	423.5	422.6	421.6	420.4	419.5	418.1
0.600	427.8	428.3	426.0	424.9	423.9	423.0	421.7	420.8	419.5
0.650	428.9	429.2	427.1	426.1	425.0	424.1	422.9	421.9	420.7
0.675	429.3	429.5	427.6	426.6	425.5	424.6	423.4	422.5	421.2
0.700	429.7	429.7	428.0	427.0	425.9	425.1	423.9	422.9	421.7
0.725	429.9	429.9	428.3	427.3	426.3	425.5	424.3	423.3	422.0
0.750	430.0	429.9	428.5	427.6	426.5	425.8	424.6	423.7	422.4
0.775	429.9	429.8	428.6	427.6	426.5	425.9	424.8	423.9	422.6
0.800	429.7	429.6	428.4	427.6	426.7	426.0	424.9	423.9	422.6
0.825	429.3	429.2	428.2	427.3	426.5	425.8	424.8	423.9	422.6
0.850	428.7	428.7	427.7	426.9	426.2	425.5	424.6	423.6	422.3
0.900	427.1	427.2	426.2	425.6	425.0	424.4	423.6	422.6	421.4
0.950	424.7	425.1	424.0	423.5	423.1	422.5	421.8	420.9	419.8
1.000	421.8	422.5	421.2	420.8	420.5	420.0	419.4	418.6	417.5
1.050	418.5	419.4	418.1	417.8	417.4	417.1	416.6	415.9	414.9
1.100	415.0	416.1	414.8	414.5	414.3	414.1	413.6	413.0	412.0
1.150	411.1	412.5	411.2	411.0	411.2	410.8	410.4	410.0	409.0
1.200	406.6	408.2	406.8	406.7	508.3	406.6	406.4	406.2	405.2

(continued) in 0.002 M HCl

	γ /mN m^{-1}								
T/°C	0.1	5.0	9.9	15.1	20.0	25.1	30.1	35.0	40.0
-E/V									
0.050	.0	.0	.0	371.8	371.2	369.5	367.7	366.4	365.7
0.100	383.5	382.7	382.1	382.0	380.8	379.9	378.4	377.4	376.9
0.150	391.8	391.2	390.4	390.2	389.5	388.3	386.9	386.0	385.4
0.200	399.4	398.8	398.0	397.5	397.1	395.7	394.3	393.4	392.9
0.250	406.0	405.4	404.6	404.1	403.6	402.3	400.9	400.1	399.5
0.300	411.6	411.0	410.2	409.7	409.1	408.0	406.7	405.9	405.3
0.350	416.2	415.7	414.9	414.4	413.6	412.7	411.4	410.6	410.0
0.400	420.1	419.4	418.7	418.2	417.3	416.5	415.2	414.4	413.7
0.450	423.2	422.5	421.7	421.1	420.3	419.3	418.1	417.2	416.5
0.500	425.6	424.9	424.1	423.3	422.6	421.5	420.3	419.3	418.6
0.550	427.6	426.8	425.9	425.0	424.5	423.2	422.0	421.0	420.2
0.600	429.0	428.2	427.3	426.3	425.9	424.6	423.3	422.3	421.5
0.625	429.6	428.7	427.8	426.8	426.4	425.2	423.9	422.8	422.1
0.650	430.0	429.1	428.2	427.3	426.9	425.6	424.3	423.3	422.6
0.675	430.3	429.4	428.5	427.6	427.2	426.0	424.7	423.7	423.0
0.700	430.5	429.6	428.7	427.8	427.4	426.3	425.0	424.0	423.3
0.725	430.5	429.7	428.8	428.0	427.5	426.5	425.3	424.3	423.6
0.750	430.4	429.6	428.7	428.0	427.5	426.6	425.4	424.4	423.7
0.800	429.9	429.1	428.3	427.6	427.1	426.3	425.2	424.3	423.6
0.850	428.8	428.0	427.3	426.7	426.2	425.5	424.5	423.6	422.9
0.900	427.1	426.4	425.8	425.1	424.8	424.1	423.3	422.4	421.7
0.950	424.9	424.2	423.7	423.0	422.8	422.1	421.4	420.6	419.9
1.000	422.1	421.6	421.2	420.4	420.4	419.7	419.1	418.4	417.8
1.050	418.9	418.4	418.1	417.4	417.5	417.0	416.4	415.7	415.3
1.100	415.2	414.9	414.6	414.1	414.1	413.9	413.3	412.8	412.4
1.150	411.2	411.0	410.7	410.4	410.4	410.3	409.9	409.4	409.1
1.200	406.9	406.7	406.6	406.2	406.4	406.3	406.1	405.6	405.3

(continued) in 0.005 M HCl

	γ /mN m^{-1}								
T/°C	0.6	5.0	10.3	15.0	20.1	25.2	29.8	35.0	40.1
-E/V									
0.000	368.5	367.1	366.0	365.2	364.1	363.1	361.4	360.3	358.7
0.050	379.1	379.8	376.8	376.6	375.5	374.1	373.6	372.1	371.0
0.100	387.9	389.1	386.0	385.6	384.5	383.5	382.6	381.4	380.4
0.150	395.3	396.7	393.4	393.0	392.0	391.6	390.3	389.4	388.4
0.200	401.9	403.4	400.0	399.7	398.7	398.4	397.2	396.4	395.2
0.250	407.9	409.5	406.0	405.8	404.8	404.2	403.3	402.4	401.1
0.300	413.2	414.7	411.4	411.0	410.2	409.1	408.5	407.5	406.2
0.350	417.7	419.0	415.9	415.4	414.5	413.3	412.7	411.6	410.5
0.400	421.2	422.4	419.3	418.8	418.0	416.7	416.0	414.8	413.9
0.450	423.8	425.0	422.0	421.5	420.5	419.5	418.6	417.4	416.5
0.500	425.9	427.0	424.0	423.5	422.5	421.8	420.7	419.5	418.4
0.550	427.5	428.4	425.6	425.1	424.0	423.5	422.3	421.2	419.9
0.600	428.7	429.2	426.8	426.4	425.3	424.7	423.5	422.4	421.1
0.625	429.2	429.5	427.3	426.9	425.8	425.1	424.0	422.9	421.6
0.650	429.5	429.6	427.7	427.2	426.1	425.4	424.4	423.3	422.0
0.675	429.6	429.6	427.9	427.4	426.4	425.5	424.6	423.5	422.4
0.700	429.5	429.5	427.9	427.5	426.5	425.5	424.7	423.6	422.6
0.725	429.3	429.2	427.8	427.4	426.4	425.3	424.6	423.6	422.6
0.750	428.8	428.7	427.5	427.1	426.1	425.1	424.4	423.4	422.5
0.800	427.4	427.3	426.2	425.9	425.0	424.1	423.5	422.5	421.6
0.850	425.2	425.3	424.3	424.0	423.2	422.5	421.9	420.9	420.0
0.900	422.6	422.8	421.8	421.6	420.8	420.4	419.7	418.8	417.8
0.950	419.7	419.9	418.9	418.7	418.1	417.8	417.0	416.3	415.4
1.000	416.3	416.6	415.7	415.5	415.0	414.7	414.0	413.4	412.6
1.050	412.4	412.8	412.0	412.0	411.5	411.1	410.6	410.2	409.4
1.100	408.1	408.7	407.9	407.9	407.6	407.2	406.9	406.5	405.7
1.150	403.4	404.0	403.4	403.3	403.1	402.9	402.7	402.3	401.6
1.200	398.5	399.0	398.6	398.6	398.4	398.4	398.0	397.7	397.5

(continued) in 0.01 M HCl

	γ /mN m^{-1}								
T/°C	0	5.0	10.0	15.0	20.0	25.1	30.0	34.9	40.0
-E/V									
0.000	369.2	368.1	367.4	365.9	364.6	364.1	364.6	363.0	361.0
0.050	381.7	380.6	379.8	378.5	377.3	376.9	376.9	375.7	374.2
0.100	390.8	389.8	388.9	387.8	386.8	386.4	386.3	385.2	383.9
0.150	398.4	397.4	396.6	395.5	394.7	394.3	394.1	393.0	391.9
0.200	405.1	404.1	403.4	402.4	401.6	401.2	401.1	399.8	398.8
0.250	411.2	410.2	409.5	408.5	407.8	407.2	407.1	405.9	404.9
0.300	416.4	415.5	414.8	413.8	413.0	412.4	412.3	411.0	410.0
0.350	420.7	419.8	419.0	418.1	417.3	416.6	416.3	415.1	414.1
0.400	424.1	423.2	422.4	421.5	420.6	419.8	419.5	418.3	417.3
0.450	426.8	425.8	425.0	424.1	423.1	422.3	421.8	420.7	419.7
0.500	428.7	427.7	426.8	426.0	425.0	424.2	423.6	422.4	421.4
0.550	430.0	429.0	428.1	427.3	426.3	425.4	424.8	423.7	422.7
0.575	430.5	429.4	428.6	427.7	426.8	425.9	425.2	424.1	423.1
0.600	430.7	429.7	428.9	428.1	427.1	426.2	425.5	424.4	423.5
0.625	430.9	429.9	429.1	428.3	427.3	426.4	425.7	424.6	423.7
0.650	430.8	429.9	429.1	428.3	427.4	426.5	425.7	424.7	423.8
0.675	430.6	429.7	429.0	428.2	427.3	426.4	425.6	424.6	423.8
0.700	430.3	429.4	428.7	427.9	427.1	426.2	425.3	414.4	423.6
0.750	429.1	428.2	427.6	426.9	426.1	425.4	424.4	423.6	422.8
0.800	427.3	426.5	426.0	425.3	424.6	423.9	422.8	422.1	421.5
0.850	424.9	424.1	423.8	423.2	422.5	421.9	420.8	420.2	419.6
0.900	422.1	421.4	421.1	420.6	420.0	419.4	418.4	417.9	417.3
0.950	418.9	418.2	418.0	417.7	417.1	416.4	415.6	415.2	414.7
1.000	415.2	414.7	414.5	414.3	413.7	413.2	412.4	412.1	411.7
1.050	411.1	410.7	410.6	410.4	410.0	409.5	408.8	408.6	408.2
1.100	406.6	406.3	406.3	406.1	405.8	405.5	404.7	404.5	404.3
1.150	401.7	401.5	401.5	401.4	401.3	400.9	400.3	400.1	400.0
1.200	396.5	396.3	396.4	396.4	396.3	396.0	395.4	395.4	395.4

(continued) in 0.02 M HCl

	γ /mN m^{-1}								
T/°C	0	5.0	10.0	15.0	19.8	25.0	30.1	35.0	39.9
-E/V									
0.000	369.4	369.6	367.0	365.9	364.7	363.4	362.1	360.8	359.6
0.050	380.4	381.7	378.3	377.4	376.3	374.9	374.2	373.0	371.9
0.100	389.1	390.8	387.2	386.4	385.3	384.2	383.2	382.3	381.2
0.150	396.5	398.4	394.7	394.0	392.8	391.9	390.8	389.9	389.0
0.200	403.1	405.2	401.4	400.7	399.6	398.6	397.6	396.7	395.9
0.250	409.1	411.2	407.5	406.7	405.6	404.5	403.7	402.8	401.9
0.300	414.4	416.4	412.7	412.0	410.7	409.7	409.0	408.1	407.1
0.350	418.3	420.8	417.1	416.4	415.0	414.1	413.3	412.4	411.4
0.400	422.4	424.2	420.6	419.8	418.5	417.5	416.7	415.8	414.8
0.450	425.1	426.7	423.3	422.5	421.1	420.2	419.3	418.4	417.4
0.500	427.2	428.5	425.3	424.5	423.2	422.2	421.3	420.4	419.3
0.550	428.6	429.6	426.8	426.0	424.7	423.7	422.8	421.8	420.7
0.575	429.1	429.8	427.3	426.5	425.2	424.2	423.3	422.3	421.2
0.600	429.4	430.0	427.7	426.9	425.6	424.6	423.8	422.7	421.5
0.625	429.5	429.9	427.9	427.1	425.8	424.8	424.0	423.0	421.8
0.650	429.5	429.7	427.9	427.1	425.9	424.9	424.1	423.1	421.9
0.675	429.3	429.4	427.7	427.0	425.8	424.8	424.1	423.1	421.9
0.700	428.8	428.8	427.3	426.6	425.5	424.5	423.8	422.9	421.7
0.750	427.4	427.3	426.0	425.4	424.3	423.4	422.8	421.9	420.8
0.800	425.2	425.2	423.9	423.3	422.5	421.6	421.0	420.1	419.2
0.850	422.5	422.5	421.3	420.8	420.1	419.2	418.7	417.8	417.0
0.900	419.3	419.5	418.3	417.8	417.2	416.4	415.9	415.1	414.4
0.950	415.8	416.0	414.9	414.4	413.9	413.1	412.8	412.0	411.3
1.000	411.8	412.1	411.1	410.7	410.2	409.6	409.3	408.6	407.9
1.050	407.4	407.8	406.8	406.6	406.0	405.5	405.3	404.7	404.0
1.100	402.5	403.1	402.1	402.0	401.5	401.0	400.9	400.4	399.6
1.150	397.4	397.9	397.1	397.1	396.7	396.2	396.2	395.6	395.1
1.200	391.7	392.4	391.7	391.6	391.2	391.2	391.1	390.6	390.2

(continued) in 0.05 M HCl

	$\gamma/\mathrm{mN\ m^{-1}}$								
$T/^{\circ}C$	0	5.0	10.2	15.0	20.1	25.0	30.0	34.9	40.0
$-E/V$									
0.000	369.6	368.3	367.5	366.2	365.5	364.1	362.3	361.2	359.8
0.050	382.1	381.3	380.7	379.6	378.9	377.4	376.3	375.1	374.3
0.100	391.5	390.4	390.0	389.1	388.4	387.4	385.9	385.3	384.2
0.150	399.1	398.1	397.7	396.9	396.2	395.4	393.9	393.4	392.2
0.200	405.9	405.0	404.6	403.8	403.2	402.3	401.0	400.2	399.3
0.250	411.9	411.2	410.7	409.9	409.3	408.3	407.2	406.2	405.5
0.300	417.1	416.4	416.0	415.1	414.5	413.4	412.4	411.3	410.7
0.350	421.5	420.7	420.3	419.4	418.8	417.7	416.6	415.6	414.8
0.400	425.0	424.1	423.6	422.7	422.1	421.1	419.9	419.0	418.0
0.450	427.6	426.6	426.1	425.2	424.5	423.7	422.3	421.5	420.4
0.500	429.3	428.3	427.8	426.9	426.2	425.3	423.9	423.1	422.0
0.525	429.9	428.9	428.3	427.4	426.7	425.8	424.5	423.6	422.5
0.550	430.2	429.2	428.7	427.8	427.0	426.1	424.8	424.0	422.9
0.575	430.3	429.4	428.9	428.0	427.2	426.3	425.0	424.1	423.1
0.600	430.2	429.4	428.8	427.9	427.2	426.3	425.0	424.1	423.1
0.625	430.0	429.1	428.6	427.8	427.0	426.1	424.9	423.9	423.0
0.650	429.6	428.7	428.3	427.4	426.7	425.7	424.5	423.6	422.7
0.700	428.2	427.4	427.0	426.2	425.5	424.6	423.5	422.5	421.7
0.750	426.3	425.5	425.1	424.4	423.7	422.9	421.8	421.0	420.2
0.800	423.8	423.0	422.7	422.0	421.5	420.8	419.7	419.0	418.1
0.850	420.7	420.1	419.9	419.2	418.7	418.2	417.1	416.4	415.6
0.900	417.3	416.7	416.6	415.9	415.6	415.1	414.1	413.4	412.6
0.950	413.3	412.9	412.0	412.3	412.0	411.5	410.6	410.0	409.3
1.000	409.0	408.7	408.8	408.2	408.0	407.5	406.8	406.1	405.6
1.050	404.3.	404.1	404.2	403.8	403.6	403.2	402.5	401.8	401.5
1.100	399.2	399.1	399.3	398.9	398.7	398.5	397.7	397.3	397.0
1.150	393.7	393.7	393.9	393.7	393.6	393.5	392.7	392.5	392.0
1.200	388.0	387.9	388.2	387.9	388.1	387.9	387.3	387.1	386.7

(continued) in 0.1 M HCl

	γ /mN m^{-1}								
T/°C	0	5.0	10.2	15.0	20.2	25.1	30.1	35.0	40.1
-E/V									
0.000	368.1	367.3	364.5	364.2	363.3	362.3	360.6	359.3	358.4
0.050	381.2	380.3	378.2	378.1	376.8	376.8	375.5	374.3	373.8
0.100	390.3	390.3	387.1	387.2	386.5	386.8	384.5	383.5	382.9
0.150	397.7	398.5	394.6	394.7	394.2	395.0	392.0	391.1	390.6
0.200	404.2	405.4	401.3	401.5	400.8	402.1	399.0	398.0	397.6
0.250	410.1	411.4	407.4	407.5	406.7	408.2	405.2	404.2	403.9
0.300	415.3	416.5	412.7	412.8	411.9	413.4	410.4	409.5	409.0
0.350	419.8	420.9	417.1	417.1	416.4	417.6	414.6	413.7	413.2
0.400	423.5	424.4	420.7	420.6	420.0	421.0	418.0	417.1	416.6
0.450	426.3	427.0	423.4	423.4	422.8	423.4	420.7	419.7	419.2
0.500	428.2	428.6	425.3	425.3	424.6	425.0	422.7	421.6	421.0
0.525	428.8	429.1	426.0	425.9	425.2	425.5	423.4	422.3	421.7
0.550	429.2	429.3	426.4	426.4	425.5	425.8	423.9	422.8	422.2
0.575	429.4	429.4	426.7	426.6	425.7	425.8	424.2	423.1	422.4
0.600	429.3	429.2	426.7	426.6	425.6	425.7	424.3	423.2	422.5
0.625	429.1	428.8	426.4	426.4	425.4	425.4	424.1	423.1	422.4
0.650	428.6	428.2	426.0	426.0	425.0	424.9	423.8	422.7	422.0
0.700	427.1	426.5	424.5	424.5	423.7	423.4	422.5	421.5	420.8
0.750	424.8	424.3	422.3	422.3	421.8	421.4	420.5	419.5	419.0
0.800	421.3	421.5	419.6	419.6	419.2	418.9	418.0	417.0	416.6
0.850	418.6	418.3	416.4	416.5	416.2	415.9	415.0	414.1	413.8
0.900	414.8	414.7	412.8	413.0	412.6	412.6	411.7	410.9	410.6
0.950	410.7	410.6	408.8	409.1	408.6	408.7	408.0	407.3	406.9
1.000	406.1	406.2	404.4	404.8	404.2	404.4	403.8	403.2	402.9
1.050	401.2	401.2	399.6	400.0	399.5	399.7	399.2	398.7	398.4
1.100	395.7	395.8	394.4	394.8	394.6	394.7	394.3	393.8	393.6
1.150	389.9	390.1	388.9	389.3	389.3	389.5	388.9	388.6	388.4
1.200	383.8	384.1	382.9	383.4	383.2	383.7	383.2	382.8	382.9

(continued) in 0.2 M HCl

	γ /mN m^{-1}								
T/°C	0.1	5.1	10.0	15.0	20.0	25.0	30.0	35.0	40.0
-E/V									
0.000	365.5	365.4	364.1	363.3	360.9	360.8	359.1	357.9	356.6
0.050	378.7	378.9	377.9	377.2	375.1	375.1	374.0	373.0	372.2
0.100	388.6	388.9	388.1	387.4	385.5	385.6	384.7	383.8	382.8
0.150	396.7	397.1	396.3	395.7	393.8	394.0	393.1	392.3	391.3
0.200	403.7	404.1	403.3	402.8	400.8	401.1	400.2	399.4	398.5
0.250	409.8	410.2	409.3	408.9	406.9	407.2	406.3	405.5	404.8
0.300	415.2	415.5	414.7	414.2	412.1	412.5	411.6	410.8	410.0
0.350	419.7	420.0	419.1	418.7	416.6	416.9	416.0	415.2	414.4
0.400	423.2	423.6	422.7	422.2	420.1	420.4	419.5	418.6	417.7
0.450	425.8	426.2	425.3	424.7	422.6	422.9	422.0	421.1	420.1
0.475	426.7	427.1	426.2	425.6	423.5	423.8	422.9	422.0	421.0
0.500	427.4	427.8	426.9	426.2	424.1	424.4	423.6	422.7	421.7
0.525	427.8	428.2	427.3	426.6	424.5	424.9	424.0	423.1	422.1
0.550	428.0	428.4	427.5	426.8	424.7	425.1	424.2	423.3	422.3
0.575	428.0	428.3	427.5	426.8	424.7	425.0	424.1	423.3	422.3
0.600	427.8	428.1	427.2	426.6	424.5	424.8	423.9	423.1	422.2
0.650	426.6	427.0	426.2	425.6	423.4	423.9	422.9	422.2	421.3
0.700	424.8	425.2	424.4	423.9	421.8	422.2	421.3	420.6	419.8
0.750	422.2	422.8	422.0	421.5	419.5	420.0	419.1	418.5	417.7
0.800	419.1	419.8	419.1	418.7	416.7	417.3	416.5	415.8	415.1
0.850	415.6	416.3	415.7	415.3	413.5	414.1	413.3	412.8	412.0
0.900	411.6	412.4	411.9	411.6	409.8	410.5	409.8	409.2	408.6
0.950	407.3	408.1	407.6	407.4	405.7	406.4	405.8	405.3	404.7
1.000	402.6	403.4	403.0	402.8	401.2	402.0	401.4	400.9	400.4
1.050	397.5	398.3	397.9	397.9	396.3	397.1	396.6	396.2	395.8
1.100	391.9	392.8	392.5	392.5	391.0	391.9	391.4	391.2	390.8
1.150	385.9	386.9	386.8	386.8	385.4	386.3	385.9	385.8	385.4
1.200	379.5	380.7	380.6	380.6	379.4	380.2	379.9	379.8	379.5

(continued) in 0.51 M HCl

	γ /mN m^{-1}								
T/°C	0	5.1	10.0	15.2	20.0	25.0	30.0	35.1	40.0
-E/V									
0.000	362.4	361.3	360.2	359.3	357.5	357.0	355.5	353.9	352.5
0.050	376.2	375.4	374.7	373.9	372.4	372.4	371.0	369.6	368.7
0.100	386.6	385.9	385.3	384.7	383.4	382.8	381.8	381.0	380.2
0.150	395.0	394.3	393.8	393.3	392.1	391.5	390.5	389.8	389.0
0.200	402.1	401.5	401.0	400.6	399.4	399.1	398.0	397.2	396.4
0.250	408.5	407.9	407.4	407.0	405.8	405.6	404.5	403.5	402.8
0.300	414.1	413.5	413.0	412.5	411.3	410.9	409.9	408.9	408.2
0.350	418.8	418.3	417.7	417.2	415.9	415.2	414.4	413.4	412.8
0.400	422.7	422.1	421.5	421.0	419.6	418.7	417.9	417.0	416.4
0.450	425.5	424.8	424.2	423.7	422.3	421.4	420.5	419.6	419.1
0.475	426.6	425.8	425.2	424.7	423.3	422.5	421.5	420.6	420.0
0.500	427.3	426.6	425.9	425.4	424.0	423.2	422.2	421.3	420.7
0.525	427.8	427.0	426.4	425.9	424.4	423.8	422.7	421.7	421.1
0.550	428.0	427.2	426.6	426.1	424.6	424.1	423.0	421.9	421.3
0.575	428.0	427.2	426.5	426.0	424.6	424.1	423.0	421.9	421.3
0.600	427.7	426.9	426.3	425.8	424.4	423.9	422.8	421.7	421.0
0.650	426.5	425.7	425.1	424.6	423.2	422.8	421.8	420.7	419.9
0.700	424.4	423.7	423.1	422.7	421.4	420.8	419.9	418.9	418.2
0.750	421.7	421.1	420.5	420.1	418.8	418.2	417.4	416.5	415.8
0.800	418.3	417.8	417.3	416.9	415.7	415.2	414.4	413.6	412.9
0.850	414.5	414.1	413.6	413.2	412.1	411.7	410.9	410.2	409.5
0.900	410.3	409.9	409.5	409.1	408.1	407.7	407.0	406.3	405.6
0.950	405.6	405.2	404.9	404.6	403.6	403.3	402.7	402.0	401.4
1.000	400.6	400.2	399.9	399.6	398.7	398.5	397.9	397.3	396.7
1.050	395.1	394.8	394.5	394.3	393.4	393.3	392.6	392.2	391.7
1.100	389.2	389.0	388.7	388.6	387.8	387.7	386.9	386.7	386.3
1.150	382.9	382.7	382.5	382.4	381.8	381.6	380.9	380.8	380.4

(continued) in 1.03 M HCl

	γ /mN m^{-1}								
T/°C	0.1	5.3	10.0	15.0	20.0	25.0	30.1	35.0	40.0
-E/V									
0.000	358.5	357.6	356.2	355.2	354.0	352.9	351.7	350.2	348.9
0.050	372.6	371.9	371.0	370.4	369.4	368.3	367.5	366.3	365.4
0.100	383.2	382.6	381.8	381.1	380.5	379.6	378.8	377.9	377.0
0.150	391.8	391.2	390.4	389.8	389.3	388.5	387.8	387.0	386.1
0.200	399.0	398.5	397.8	397.4	396.7	396.0	395.3	394.6	393.8
0.250	405.4	405.0	404.3	404.0	403.2	402.5	401.9	401.1	400.4
0.300	411.1	410.6	410.0	409.6	408.9	408.2	407.6	406.8	406.2
0.350	415.9	415.5	414.9	414.4	413.8	413.0	412.4	411.5	411.0
0.400	419.9	419.5	418.8	418.2	417.7	416.9	416.3	415.3	414.8
0.450	423.0	422.4	421.8	421.1	420.6	419.7	419.1	418.1	417.5
0.475	424.1	423.5	422.8	422.2	421.6	420.8	420.1	419.1	418.4
0.500	425.0	424.4	423.6	423.0	422.4	421.5	420.8	419.9	419.1
0.525	425.6	424.9	424.2	423.6	422.9	422.0	421.3	420.4	419.6
0.550	425.9	425.2	424.5	424.0	423.2	422.3	421.6	420.6	419.8
0.575	426.0	425.3	424.5	424.0	423.2	422.3	421.6	420.6	419.8
0.600	425.8	425.1	424.3	423.8	423.0	422.1	421.3	420.4	419.5
0.650	424.7	424.0	423.2	422.7	421.9	421.0	420.2	419.3	418.5
0.700	422.6	422.0	421.3	420.7	420.0	419.2	418.3	417.5	416.7
0.750	419.8	419.3	418.7	418.0	417.4	416.6	415.8	415.0	414.3
0.800	416.4	415.9	415.4	414.7	414.2	413.4	412.7	411.9	411.3
0.850	412.5	412.0	411.5	411.0	410.5	409.8	409.1	408.4	407.8
0.900	408.1	407.6	407.1	406.8	406.3	405.6	405.0	404.3	403.7
0.950	403.1	402.8	402.3	402.2	401.6	401.0	400.5	399.8	399.3
1.000	397.9	397.6	397.2	397.0	396.5	395.9	395.5	394.9	394.4
1.050	392.2	392.0	391.7	391.4	391.0	390.5	390.1	389.5	389.2
1.100	386.1	385.9	385.9	385.4	385.1	384.7	384.3	383.9	383.5

(continued) in 2.09 M HCl

	γ /mN m^{-1}								
T/°C	0.3	5.2	10.0	15.0	19.9	25.1	30.1	35.0	40.0
-E/V									
0.000	355.9	354.1	352.6	351.4	349.5	348.1	347.5	346.1	344.7
0.050	369.8	368.6	367.8	367.0	365.6	364.2	363.9	362.9	361.8
0.100	380.3	379.3	378.4	377.6	376.5	375.2	375.4	374.5	373.6
0.150	388.8	387.9	387.2	386.4	385.4	384.2	384.4	383.6	382.8
0.200	396.0	395.2	394.9	394.1	393.1	392.1	392.0	391.3	390.5
0.250	402.5	401.7	401.5	400.8	399.8	398.8	398.7	398.0	397.3
0.300	408.2	407.5	407.1	406.5	405.5	404.6	404.6	404.0	403.2
0.350	413.2	412.5	411.9	411.3	410.4	409.4	409.6	409.0	408.2
0.400	417.3	416.6	416.0	415.3	414.4	413.4	413.7	413.0	412.3
0.450	420.6	419.8	419.2	418.5	417.5	416.5	416.7	416.0	415.2
0.475	421.8	421.0	420.5	419.8	418.8	417.7	417.9	417.1	416.3
0.500	422.8	422.0	421.5	420.8	419.8	418.7	418.7	418.0	417.2
0.525	423.6	422.8	422.3	421.5	420.5	419.4	419.4	418.6	417.8
0.550	424.1	423.2	422.8	422.0	421.0	419.8	419.7	419.0	418.1
0.575	424.3	423.4	423.0	422.3	421.2	420.0	419.9	419.1	418.3
0.600	424.2	423.4	423.0	422.2	421.1	420.0	419.8	419.0	418.2
0.650	423.4	422.6	422.1	421.4	420.2	419.1	418.9	418.2	417.4
0.700	421.5	420.8	420.3	419.6	418.4	417.3	417.3	416.5	415.7
0.750	418.8	418.2	417.7	417.0	415.9	414.8	414.8	414.1	413.3
0.800	415.3	414.8	414.4	413.7	412.7	411.7	411.7	411.0	410.2
0.850	411.3	410.8	410.4	409.9	408.9	407.9	407.9	407.2	406.5
0.900	406.7	406.3	406.0	405.5	404.6	403.7	403.6	403.0	402.4
0.950	401.7	401.4	401.1	400.7	399.8	398.9	398.9	398.4	397.8
1.000	396.3	396.1	395.8	395.4	394.6	393.7	393.8	393.4	392.9
1.050	390.3	390.1	.0	.0	.0	.0	388.3	387.5	387.3

(continued) in 3.19 M HCl

	γ /mN m^{-1}								
T/°C	0	5.1	10.0	15.0	20.0	25.0	30.1	35.0	40.0
-E/V									
0.000	349.7	348.0	347.2	345.7	344.4	343.3	342.3	341.1	339.8
0.050	365.0	363.3	362.6	361.8	361.1	360.1	359.2	358.1	357.2
0.100	375.4	374.1	373.8	372.0	372.0	371.3	370.9	369.9	369.2
0.150	384.1	382.8	382.7	381.7	381.1	380.4	379.9	379.1	378.5
0.200	391.7	390.3	390.2	389.3	389.9	388.3	387.6	386.9	386.4
0.250	398.4	396.9	396.8	396.0	395.7	395.2	394.4	393.8	393.3
0.300	404.1	402.8	402.6	402.0	401.4	401.0	400.3	399.8	399.3
0.350	409.0	407.8	407.6	407.1	406.3	405.9	405.4	404.9	404.4
0.400	413.1	412.0	411.8	411.3	410.5	410.1	409.6	409.0	408.5
0.450	416.6	415.4	415.2	414.6	413.9	413.4	412.9	412.2	411.7
0.500	419.3	418.0	417.7	417.0	416.5	415.9	415.3	414.6	414.0
0.525	420.3	418.9	418.6	417.9	417.4	416.8	416.1	415.4	414.8
0.550	421.1	419.6	419.3	418.5	418.1	417.5	416.7	416.0	415.4
0.575	421.6	420.1	419.7	419.0	418.5	417.8	417.1	416.4	415.8
0.600	421.9	420.4	419.9	419.2	418.7	418.0	417.3	416.6	415.9
0.625	421.9	420.4	419.9	419.2	418.6	417.9	417.2	416.5	415.8
0.650	421.6	420.2	419.6	418.9	418.3	417.6	416.9	416.2	415.5
0.700	420.2	419.0	418.3	417.7	417.0	416.2	415.7	415.0	414.3
0.750	418.0	416.7	416.1	415.5	414.7	414.1	413.5	412.8	412.1
0.800	414.9	413.5	413.1	412.5	411.7	411.2	410.6	409.9	409.2
0.850	411.1	409.6	409.3	408.7	408.1	407.6	406.9	406.2	405.5
0.900	406.6	405.2	404.9	404.4	403.9	403.3	402.8	401.9	401.3
0.950	401.6	400.4	400.0	399.6	399.1	398.5	398.3	396.7	396.1

Electrocapillary data on mercury electrode in 0.001 M LiCl at different temperatures - given are the interfacial tension in mN m^{-1}, potential in mV vs CE.

Reference: E.W. Hermann. M.Sc. Thesis, Toronto University 1967.

	γ/mN m^{-1}								
T/$^{\circ}$C	1.00	10.00	15.00	20.00	25.01	30.00	35.00	40.00*	45.02*
-E/mV									
575	433.0	427.6							
600	433.6	428.6	426.8	425.8	423.7	422.2			
650	433.3	430.0	428.5	420.5	425.7	424.5	423.3	422.1	421.1
700	432.4	430.5	429.1	429.6	426.8	425.8	424.6	423.6	422.7
750	431.3	430.3	428.8	429.5	426.9	426.0	424.9	424.1	423.3
800	429.7	429.6	428.0	428.8	426.3	425.5	424.5	423.7	422.9
850	427.8	428.5	426.9	427.9	425.4	424.5	423.5	422.9	422.1
900		427.7	425.7	426.7	424.4	423.3	422.4	421.9	421.1
950		425.3	423.7	424.9	422.8	421.9	421.1	420.6	419.9
1000						420.0	419.0	418.8	418.1

* For T = 40.00°C and T = 45.02°C in this table the potential values, E, given are 5 mV too low. Thus -650 mV should read -645 mV, -700 mV should read -695 mV etc.

(continued) in 0.002 M LiCl

					γ /mN m^{-1}				
T/°C	0.70	10.00	15.00	20.00	25.00	30.00	35.00	40.00*	44.92*
-E/mV									
475						416.7	415.8	414.9	413.7
500					419.4	418.5	417.6	416.7	415.5
550		426.0	426.2	425.2	422.9	421.8	420.8	419.8	418.7
600	429.7	429.1	420.3	427.5	425.3	424.2	423.2	422.1	421.1
650	430.7	430.1	429.2	428.7	426.0	425.8	424.7	423.6	422.7
700	430.5	430.0	429.0	428.8	427.3	426.4	425.3	424.2	423.6
750	429.7	429.3	428.2	428.0	427.0	426.2	425.1	424.2	423.3
800	428.6	428.4	427.1	426.8	426.1	425.3	424.3	423.5	422.6
850	427.2	426.9	425.5	425.5	424.7	423.9	423.1	422.3	421.5
900	425.2	424.9	423.2	423.9	423.0	422.2	421.6	420.7	419.9
950	422.7	422.4		421.4	421.0	420.2	419.7	418.7	418.1
1000					418.4	417.7	417.3	416.4	416.0
1050					415.1	414.6	414.1	413.7	413.4
1100								410.3	409.9

* For T = 40.00°C and T = 45.02°C in this table the potential values, E, given are 5 mV too low. Thus -650 mV should read -645 mV, -700 mV should read -695 mV etc.

(continued) in 0.005 M LiCl

	γ /mN m^{-1}								
T/°C	0.81	10.00	15.00	20.00	25.00	30.00	35.00	40.00	45.00
-E/mV									
450					418.0	416.7	416.2	415.0	413.9
500					421.3	420.2	419.4	418.4	417.2
550	431.2	428.8		425.8	424.2	423.0	422.1	421.0	419.9
600	432.9	431.2	428.6	427.9	426.1	424.8	423.8	422.7	421.9
650	433.3	432.2	429.3	428.7	426.7	425.6	424.6	423.6	422.7
700	432.8	431.9	428.8	428.5	426.4	425.6	424.5	423.7	422.6
750	431.9	430.8	427,8	427.5	425.6	424.8	423.7	423.1	421.9
800	430.6	429.4	426.4	426.2	424.6	423.6	422.6	421.9	420.8
850	428.7	427.8	424.7	424.6	423.6	421.9	421.2	420.3	419.5
900	425.7	425.3	422.5	422.1	422.5	419.9	419.6	418.3	417.8
950			419.7		420.2	417.6	417.4	416.0	415.3
1000					415.3	414.8	413.8	413.3	410.8
1050						411.0		410.3	
1100								406.7	

(continued) in 0.01 M LiCl

	γ /mN m^{-1}								
T/°C	0.72	10.12	15.24	20.00	25.00	30.00	35.00	40.00	44.90
-E/mV									
300					408.7	408.7	406.8	405.6	404.4
350		417.6	416.1	414.8	413.6	412.7	411.2	410.4	409.1
400	424.0	421.8	420.3	418.9	417.8	416.7	415.4	414.6	413.4
450	427.1	425.1	423.5	422.4	421.2	420.1	419.1	418.1	417.0
500	429.6	427.7	426.0	425.1	423.9	422.9	421.9	420.9	419.9
550	431.8	429.6	427.7	426.8	425.7	424.8	423.0	422.9	421.9
600	433.1	430.5	428.5	427.8	426.7	425.8	424.9	424.1	423.0
650	433.0	430.6	428.5	427.9	427.0	426.1	425.1	424.4	423.3
700	432.1	430.0	427.8	427.4	426.5	425.6	424.7	424.0	422.9
750	431.0	428.9	426.6	426.3	425.3	424.6	423.7	423.0	421.9
800	429.7	427.3	425.0	424.7	423.7	423.1	422.2	421.5	420.5
850	427.8	425.3	423.1	422.6	421.7	421.1	420.3	419.6	418.7
900	424.4	422.6	420.8	420.1	419.3	418.8	418.1	417.4	416.5
950	419.6	419.1	417.8	417.2	416.5	416.0	415.5	414.9	414.0
1000	417.6	416.0	414.1	413.8	413.2	412.0	412.4	411.9	411.1
1050			409.8	409.9	409.5	409.1	408.8	408.3	407.6
1100			405.7	405.5	405.2	404.9	404.5	404.2	403.6
1150				400.8	400.6	400.3	399.9	399.6	399.2
1200				395.9	395.8	395.6	395.4	395.2	394.7

(continued) in 0.02 M LiCl

	γ /mN m^{-1}								
T/^{o}C	0.70	10.10	15.00	20.00	25.00	30.06	35.10	40.00	44.94
-E/mV									
150								388.0	
200								395.2	
250						403.4		401.5	
300					409.6	408.6	407.6	407.0	406.4
350				415.3	414.2	413.4	412.3	411.9	410.9
400		421.0	420.2	419.1	418.5	417.6	416.5	416.0	415.0
450	425.0	424.6	423.7	422.8	421.9	420.9	419.9	419.2	418.3
500	428.8	427.3	426.4	425.5	424.4	423.3	422.4	421.6	420.7
550	430.5	428.8	427.9	427.0	425.9	424.8	423.9	423.1	422.1
600	430.7	429.2	428.4	427.5	426.4	425.5	424.5	423.7	422.7
650	430.2	428.7	427.9	427.2	426.2	425.3	424.3	423.5	422.6
700	429.1	427.7	426.9	426.2	425.3	424.5	423.5	422.7	421.9
750	427.6	426.1	425.4	424.0	423.9	423.1	422.2	421.4	420.7
800	425.5	424.2	423.5	422.9	422.2	421.3	420.4	419.7	419.2
850	422.9	421.8	421.2	420.6	420.0	419.1	418.3	417.6	417.2
900	419.8	418.9	418.4	417.9	417.3	416.5	415.8	415.2	414.7
950	415.2	415.6	415.1	414.6	414.3	413.5	412.9	412.4	411.8
1000		411.7	411.3	411.0	410.7	410.0	409.5	409.2	408.5
1050		407.5	407.2	406.9	400.5	406.1	405.6	405.3	404.7
1100		402.8	402.7	402.5	401.9	401.6	401.3	400.9	400.4
1150		397.5	397.9	397.6	397.0	396.7	396.6	396.1	395.9
1200		390.8	392.2	392.1	392.0	391.7	391.8	391.5	391.1

(continued) in 0.05 M LiCl

	γ/mN m^{-1}								
T/°C	0.75	10.00	14.80	20.00	25.10	29.90	35.00	40.00	45.00
-E/mV									
50								374.8	373.8
100				387.2	386.1	386.9	384.4	383.6	382.9
150			396.3	395.3	394.3	393.9	392.7	391.7	390.9
200			403.3	402.5	401.5	400.6	400.0	399.0	398.1
250		409.9	409.5	408.9	407.7	407.0	406.3	405.4	404.3
300	417.0	415.3	414.8	414.3	413.1	412.4	411.6	410.8	409.6
350	421.4	419.9	419.3	418.7	417.6	416.9	416.0	415.2	414.0
400	425.2	423.7	422.9	422.2	421.1	420.4	419.5	418.6	417.4
450	428.0	426.4	425.5	424.7	423.7	423.0	422.0	421.2	420.0
500	429.7	428.1	427.2	426.4	425.4	424.6	423.6	422.8	421.7
550	430.5	428.9	428.0	427.3	426.2	425.4	424.5	423.7	422.6
600	430.3	428.9	428.0	427.3	426.3	425.4	424.6	423.8	422.8
650	429.5	428.1	427.3	426.8	425.6	424.8	424.0	423.3	422.3
700	428.0	426.0	426.0	425.6	424.4	423.7	422.9	422.2	421.3
750	426.1	424.9	424.1	423.0	422.7	422.0	421.3	420.8	419.7
800	423.7	422.6	421.8	421.6	420.5	419.9	419.2	418.6	417.7
850	420.9	419.8	419.0	418.9	417.9	417.3	416.8	416.1	415.3
900	417.6	416.6	415.9	415.8	414.9	414.4	413.9	413.2	412.5
950	413.8	413.0	412.4	412.2	411.5	411.0	410.6	409.9	409.2
1000	409.5	408.8	408.4	408.3	407.6	407.1	406.8	406.2	405.6
1050	404.6	404.1	403.9	403.9	403.2	402.8	402.5	401.9	401.4
1100		398.9	398.8	399.0	398.2	398.0	397.8	397.3	396.8
1150		393.4	393.3	393.7	392.9	392.8	392.6	392.3	391.8
1200		387.9	387.4	388.0	387.5	387.3		387.0	386.6
1250			381.7	382.0					

(continued) in 0.10 M LiCl

	γ /mN m^{-1}								
T/°C	0.70	10.03	15.01	20.00	25.00	30.01	34.90	40.00	45.00
-E/mV									
0		366.2	365.2	363.9	362.7	361.3	360.3	358.7	357.5
50		379.4	378.6	377.6	376.7	375.5	374.6	373.5	372.4
100		389.6	388.9	388.1	387.3	386.3	385.4	384.6	383.6
150	399.0	398.0	397.3	396.6	395.8	394.9	394.1	393.3	392.4
200	406.1	405.0	404.4	403.7	402.9	402.1	401.2	400.5	399.6
250	412.3	411.1	410.5	409.8	409.0	408.2	407.3	406.6	405.7
300	417.6	416.3	415.7	415.0	414.2	413.3	412.4	411.7	410.8
350	422.1	420.7	420.0	419.3	418.5	417.6	416.7	415.9	415.0
400	425.6	424.2	423.4	422.8	421.9	421.0	420.0	419.2	418.3
450	428.2	426.7	426.0	425.2	424.4	423.4	422.5	421.6	420.7
500	429.8	428.3	427.5	426.8	425.9	424.9	424.0	423.1	422.2
550	430.4	429.0	428.2	427.4	426.5	425.5	424.6	423.7	422.8
600	430.2	428.8	428.0	427.2	426.3	425.3	424.4	423.5	422.6
650	429.2	427.8	427.0	426.3	425.4	424.4	423.5	422.7	421.8
700	427.5	426.2	425.4	424.7	423.0	422.9	422.0	421.2	420.3
750	425.3	423.9	423.2	422.6	421.7	420.9	420.0	419.3	418.5
800	422.4	421.2	420.6	420.0	419.2	418.4	417.7	417.0	416.2
850	419.2	418.1	417.5	417.0	416.3	415.6	414.9	414.3	413.5
900	415.5	414.5	414.0	413.7	413.1	412.4	411.8	411.2	410.5
950	411.2	410.6	410.2	409.9	409.4	408.7	408.2	407.7	407.1
1000	406.8	406.2	405.9	405.6	405.2	404.8	404.2	403.6	402.3
1050	401.8	401.3	401.1	400.8	400.5	400.0	399.6	399.1	398.8
1100	396.3	395.9	395.7	395.6	395.2	394.8	394.4	394.0	393.8
1150	390.4	390.1	389.9	389.9	389.6	389.2	388.9	388.6	388.4
1200	384.4	384.3	383.7	384.2	383.9	383.6	383.4	383.3	383.0

(continued) in 0.20 M LiCl

	γ /mN m^{-1}								
T/°C	0.80	10.00	15.10	20.00	25.04	30.04	35.00	40.04	45.04
-E/mV									
0	366.4	364.4	363.5	362.2	361.0	359.9	358.4	356.5	355.9
50	379.6	378.2	377.3	376.4	375.5	374.3	373.1	371.8	371.2
100	389.8	388.7	387.9	387.2	386.4	385.3	384.3	383.1	382.7
150	398.2	397.2	396.4	395.8	395.0	394.0	393.1	391.9	391.6
200	405.4	404.4	403.6	403.0	402.2	401.4	400.5	399.1	399.0
250	411.7	410.6	409.9	409.2	408.4	407.6	406.8	405.3	405.1
300	417.2	416.0	415.2	414.5	413.7	412.9	412.1	410.6	410.3
350	421.7	420.5	419.7	418.9	418.1	417.3	416.4	415.0	414.6
400	425.4	424.0	423.2	422.3	421.5	420.7	419.9	418.5	417.9
450	428.0	426.6	425.7	424.9	424.0	423.1	422.3	421.1	420.4
500	429.5	428.1	427.3	426.4	425.5	424.6	423.8	422.6	421.8
550	430.2	428.7	427.8	426.9	426.1	425.1	424.3	423.1	422.4
600	429.8	428.3	427.5	426.6	425.8	424.8	424.0	422.8	422.1
650	428.6	427.1	426.3	425.5	424.6	423.7	423.0	421.7	421.1
700	426.6	425.2	424.5	423.7	422.9	422.0	421.3	420.0	419.4
750	424.1	422.8	422.1	421.3	410.5	419.7	419.0	417.8	417.3
800	421.0	419.9	419.2	418.5	417.8	417.1	416.4	415.2	414.7
850	417.6	416.5	416.0	415.3	414.6	414.0	413.4	412.3	411.8
900	413.7	412.8	412.3	411.7	411.1	410.5	410.0	409.0	408.5
950	409.4	408.7	408.2	407.7	407.1	406.6	406.1	405.3	404.8
1000	404.6	404.1	403.6	403.2	402.7	402.2	401.8	401.1	400.6
1050	399.4	398.9	398.4	398.1	397.7	397.3	396.9	396.2	395.9
1100	393.6	393.2	392.8	392.6	392.2	391.8	391.5	390.8	390.6
1150	387.5	387.1	386.8	396.6	386.3	385.9	385.7	385.0	384.9
1200	381.3	381.1	380.8	380.5	380.4	380.2	380.0	379.5	379.4

(continued) in 0.51 M LiCl

	γ /mN m^{-1}								
T/°C	0.95	10.05	15.06	20.01	25.02	30.06	35.08	40.00	45.00
-E/mV									
0	362.9	361.0	359.7	358.6	357.4	356.5	354.9	353.7	352.7
50	376.8	375.2	374.3	373.2	372.2	371.4	370.3	369.3	368.4
100	387.3	386.0	385.3	384.3	383.5	382.8	381.9	381.0	380.2
150	395.9	394.7	394.1	393.3	392.5	391.9	391.1	390.3	389.5
200	403.3	402.2	401.5	400.9	400.1	399.5	398.7	398.0	397.1
250	409.7	408.6	407.9	407.4	406.6	406.0	405.2	404.5	403.6
300	415.4	414.2	413.5	412.9	412.2	411.6	410.7	410.0	409.1
350	420.2	419.0	418.2	417.6	416.8	416.3	415.3	414.6	413.7
400	424.0	422.7	422.9	421.3	420.5	419.8	419.0	418.2	417.3
450	426.8	425.5	424.7	424.0	423.2	422.4	421.6	420.7	419.8
500	428.5	427.1	426.4	425.6	424.8	424.0	423.2	422.3	421.4
550	429.1	427.7	426.9	426.2	425.3	424.6	423.7	422.8	421.9
600	428.7	427.3	426.5	425.8	424.9	424.2	423.4	422.4	421.5
650	427.3	426.9	425.2	424.5	423.7	423.0	422.1	421.2	420.4
700	425.2	423.9	423.2	422.5	421.6	421.0	420.2	419.3	418.5
750	422.4	421.2	420.5	419.9	419.0	418.4	417.7	416.8	416.0
800	419.0	417.9	417.3	416.7	415.9	415.4	414.7	413.9	413.1
850	415.2	414.3	413.7	413.1	412.4	411.9	411.3	410.5	409.8
900	411.0	410.2	409.7	409.2	408.4	408.1	407.5	406.8	406.1
950	406.4	405.7	405.2	404.8	404.1	403.8	403.3	402.6	402.0
1000	401.4	400.7	400.3	400.0	399.3	399.0	398.6	398.0	397.4
1050	395.8	395.2	394.9	394.6	394.0	393.8	393.3	392.8	392.3
1100	389.8	389.2	389.0	388.8	388.2	388.0	387.6	387.2	386.7
1150	383.3	382.9	382.7	382.6	382.0	381.9	381.5	381.1	380.7
1200	376.8	376.6	376.5	376.4	375.8	375.8	375.6	375.2	374.8

(continued) in 1.02 M LiCl

	γ /mN m^{-1}								
T/°C	1.00	10.00	15.00	20.00	25.00	30.00	35.05	40.00	45.00
-E/mV									
0	358.9	357.6	356.6	355.5	354.2	352.6			
50	373.1	372.0	371.4	370.5	369.4	368.4	367.7	366.9	365.6
100	384.0	383.1	382.6	381.9	381.0	380.2	379.2	378.4	377.2
150	392.9	392.1	391.7	391.1	390.2	389.6	388.6	387.8	386.8
200	400.6	399.7	399.3	398.8	398.0	397.3	396.5	295.7	394.8
250	407.3	406.4	406.0	405.5	404.8	404.0	403.4	402.5	401.7
300	413.2	412.3	411.8	411.3	410.6	409.7	409.2	408.3	407.5
350	418.2	417.3	416.7	416.2	415.4	414.5	414.0	413.2	412.3
400	422.3	421.3	420.7	420.1	419.3	418.4	417.8	416.9	416.1
450	425.3	424.2	423.6	423.0	422.2	421.2	420.6	419.7	418.7
500	427.2	426.0	425.4	424.7	423.9	423.0	422.3	421.3	420.4
550	427.9	426.8	426.1	425.4	424.6	423.6	422.9	421.9	421.0
600	427.6	426.4	425.8	425.1	424.2	423.3	422.6	421.6	420.7
650	426.3	425.1	424.5	423.8	423.0	422.0	421.4	420.3	419.5
700	424.1	422.9	422.4	421.7	420.9	420.0	419.4	418.4	417.6
750	421.1	420.1	419.6	418.9	418.1	417.3	416.7	415.7	415.0
800	417.6	416.6	416.2	415.5	414.9	414.0	413.5	412.6	411.9
850	413.5	412.7	412.3	411.7	411.1	410.3	409.8	409.0	408.3
900	409.0	408.3	407.9	407.4	406.9	406.2	405.7	404.9	404.3
950	404.1	403.5	403.2	402.8	402.3	401.7	401.2	400.5	399.9
1000	398.8	398.3	398.0	397.6	397.3	396.7	396.2	395.6	395.0
1050	393.0	392.6	392.3	392.0	391.7	391.2	390.8	390.2	389.7
1100	386.7	386.4	386.2	385.9	385.6	385.2	384.9	384.3	383.9
1150	380.1	379.9	379.7	379.5	379.2	378.8	378.6	378.1	377.7
1200	373.4	373.2	373.1	373.1	372.8	272.5	372.1	371.7	371.3

(continued) in 1.55 M LiCl

T/°C	γ/mN m^{-1} 1.00	10.00	15.00	20.00	25.00	30.03	35.02	40.00	44.98
-E/mV									
25	365.4	363.9	362.9	362.2	360.0	359.8	358.6	357.6	356.5
50	371.6	370.2	369.4	368.7	367.5	366.5	365.5	364.5	363.5
100	382.4	381.1	380.4	379.8	378.8	377.8	377.2	376.2	375.4
150	391.3	390.2	389.6	389.0	388.2	387.3	386.8	385.8	385.1
200	399.1	398.1	397.5	396.9	396.2	395.4	394.8	394.0	393.3
250	405.9	405.0	404.3	403.8	403.1	402.3	401.8	401.0	400.3
300	411.9	411.0	410.3	409.8	409.1	408.3	407.7	407.0	406.2
350	417.1	416.1	415.4	414.9	414.1	413.4	412.6	411.9	411.2
400	421.2	420.2	419.4	418.9	418.2	417.3	416.6	415.9	415.1
450	424.4	423.3	422.5	421.9	421.1	420.3	419.5	418.7	417.9
500	426.4	425.2	424.4	423.8	423.0	422.1	421.3	420.6	419.7
550	427.4	426.1	425.3	424.6	423.8	422.9	422.1	421.3	420.4
600	427.2	425.9	425.1	424.4	423.5	422.7	421.9	421.1	420.1
650	426.0	424.7	424.0	433.2	422.3	421.5	420.7	420.0	419.0
700	423.9	422.6	421.9	421.1	420.2	419.5	418.7	418.0	417.0
750	420.9	419.7	419.1	418.3	417.5	416.8	416.0	415.3	414.4
800	417.3	416.1	415.6	414.9	414.2	413.5	412.7	412.1	411.3
850	413.1	412.0	411.6	410.9	410.3	409.6	408.9	408.3	407.6
900	408.4	407.5	407.1	406.5	406.0	405.3	404.6	404.0	403.5
950	403.3	402.6	402.1	401.6	401.3	400.6	399.9	399.3	398.9
1000	397.9	397.2	396.8	396.4	396.1	395.5	394.8	394.3	393.0
1050	392.0	391.5	391.0	390.7	390.5	389.8	389.2	388.7	388.2
1100	385.7	385.2	384.8	384.5	384.3	383.8	383.2	382.8	382.2
1150	378.9	378.2	370.1	377.9	377.6	377.2	376.7	376.4	373.8
1200	371.7	370.6	371.2	370.9	370.5	370.4	369.9	369.8	369.4

(continued) in 2.08 M LiCl

	γ /mN m^{-1}								
T/°C	1.40	10.04	15.08	20.00	25.00	30.10	35.08	40.10	45.00
-E/mV									
25	363.1	361.8	361.0	360.1	359.9	357.6	356.0	355.3	354.3
50	369.5	367.4	367.5	366.7	365.7	364.4	363.8	362.4	361.4
100	380.4	378.0	378.6	377.9	377.1	375.9	375.5	374,2	373,5
150	389.4	387.5	387.8	387.3	386.6	385.5	385.1	384.0	383.3
200	397.3	396.0	395.8	395.3	394.6	393.6	393.3	392.3	391.6
250	404.2	403.4	402.8	402.3	401.6	400.7	400.4	399.4	398.7
300	410.3	409.8	408.8	408.4	407.7	406.8	406.4	405.5	404.8
350	415.5	415.0	414.0	413.5	412.8	412.0	411.6	410.6	409.9
400	419.8	419.2	418.2	417.7	416.9	416.2	415.7	414.6	414.0
450	423.1	422.2	421.4	420.8	420.0	419.3	418.8	417.7	417.0
500	425.3	424.2	423.5	422.9	422.1	421.3	420.8	419.6	418.9
550	426.4	425.2	424.5	423.8	423.0	422.2	421.7	420.5	419.8
600	426.4	425.1	424.4	423.7	422.9	422.1	421.6	420.3	419.6
650	425.3	424.1	423.3	422.6	421.8	421.0	420.5	419.2	418.5
700	423.3	422.2	421.3	420.6	419.8	419.0	418.5	417.3	416.6
750	420.4	419.4	418.5	417.9	417.0	416.3	415.8	414.6	414.0
800	416.8	415.9	415.0	414.4	413.6	412.9	412.5	411.3	410.8
850	412.6	411.7	410.9	410.4	409.7	409.0	408.6	407.5	407.1
900	407.9	407.0	406.4	405.8	405.2	404.7	404.2	403.3	402.9
950	402.7	401.8	401.4	400.9	400.4	399.8	399.4	398.5	398.2
1000	397.1	396.2	395.9	395.5	395.1	394.5	394.2	393.4	393.1
1050	391.1	390.3	390.1	389.7	389.3	388.8	388.6	387.7	387.4
1100	384.6	384.1	383.5	383.5	383.1	382.6	382.5	381.7	381.3
1150	377.8	377.4	377.0	376.8	376.5	376.0	375.9	375.1	374.8
1200	370.6	370.0	369.8	369.8	369.5	369.1	369.1	368.4	368.2

(continued) in 3.19 M LiCl

	γ /mN m^{-1}								
T/°C	1.30	10.00	14.98	19.98	25.00	30.03	35.03	39.97	45.00
-E/mV									
75	372.4	371.5	371.0	369.7	369.3	368.2	367.7	366.5	
100	377.7	377.0	376.5	375.3	374.8	373.8	373.3	372.2	371.2
150	387.0	386.5	385.9	385.0	384.4	383.5	383.0	382.1	381.2
200	394.9	394.5	394.0	393.2	392.6	391.8	391.3	390.6	389.7
250	402.0	401.5	401.0	400.2	399.8	399.1	398.6	398.0	397.1
300	408.2	407.7	407.2	406.4	406.0	405.3	404.8	404.2	403.4
350	413.6	413.0	412.5	411.8	411.3	410.6	410.1	409.5	408.7
400	418.1	417.5	416.9	416.2	415.7	415.0	414.4	413.8	413.0
450	421.7	421.0	420.3	419.6	419.0	418.3	417.7	417.0	416.2
500	424.2	423.4	422.7	422.0	421.3	420.5	419.9	419.1	418.3
550	425.6	424.7	423.9	423.3	422.5	421.7	421.0	420.2	419.4
600	425.9	424.8	424.1	423.4	422.7	421.8	421.1	420.3	419.4
650	425.0	423.9	423.2	422.5	421.8	420.9	420.2	419.4	418.5
700	423.2	422.1	421.4	420.7	419.9	419.1	418.4	417.7	416.7
750	420.5	419.3	418.7	418.0	417.3	416.6	415.8	415.1	414.2
800	417.0	415.8	415.3	414.6	413.9	413.3	412.5	411.9	411.0
850	412.8	411.7	411.2	410.6	409.9	409.3	408.7	408.0	407.2
900	408.0	407.0	406.6	406.0	405.4	404.9	404.3	403.6	402.8
950	402.7	401.9	401.5	401.0	400.4	399.9	399.4	398.8	398.0
1000	397.0	396.3	395.9	395.5	394.9	394.4	394.0	393.4	392.7
1050	390.9	390.2	389.9	389.5	389.0	388.5	388.2	387.6	387.0
1100	384.2	383.7	383.4	383.0	382.6	382.1	381.9	381.4	380.8
1150	377.1	376.7	376.4	376.1	375.7	375.3	375.1	374.7	373.2
1175	373.4	373.1	372.8	372.5	372.1	371.8	371.6	371.1	370.8

Electrocapillary data on mercury electrode in LiCl - given are the interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: J. Jastrzebska. Rocz. Chem. 44 (1970) 1779.

			γ /mN m^{-1}			
c/mol l^{-1}	0.05	0.1	0.5	0.8	1.0	2.0
-E/V						
0.1	405.5	398.0	386.8	385.0	384.2	382.6
0.2	415.7	410.6	404.2	401.6	399.8	398.0
0.3	421.6	418.5	415.0	412.8	412.0	409.1
0.4	424.9	423.3	421.5	420.1	419.4	416.6
0.5	424.7	424.6	423.8	423.4	422.6	421.0
0.6	422.9	422.7	422.3	422.0	421.8	421.1
0.7	419.2	418.9	418.4	418.1	418.0	417.2
0.8	413.1	413.0	411.8	411.5	411.3	410.5
0.9	406.9	406.3	403.5	403.1	403.0	401.6
1.0	396.6	396.4	393.8	392.7	392.4	391.3

Electrocapillary data on mercury electrode in KCl in water - given are the interfacial tension in mN m^{-1}, potential in V, the reference electrode NCE aq. was constantly at the same temperature as the system studied.

Reference: S. Minc and J. Andrzejczak. Roczniki Chem., 40 (1966) 1547.

	γ /mN m^{-1}				
T/°C	20			40	60
c/mol l^{-1}	1.5	2.5	3.5	3.5	3.5
E/V					
-0.10	374.2	372.4	368.4	366.8	363.6
-0.20	394.8	390.2	386.2	384.9	381.4
-0.30	406.5	403.8	400.1	398.5	396.1
-0.40	415.3	412.6	410.2	408.7	405.6
-0.50	419.9	418.2	416.3	419.9	411.9
-0.55	421.2	419.6	417.8	416.4	413.6
-0.60	421.0	419.4	418.1	417.1	414.1
-0.65	419.6	418.9	417.6	416.9	414.0
-0.70	417.7	417.2	415.8	415.2	412.3
-0.80	411.5	411.3	409.9	410.1	407.4
-0.90	401.3	400.8	400,7	400.8	399.6

Electrocapillary data on mercury electrode in 0.1 M NaCl at 25°C - given are the interfacial tension in mN m^{-1}, tension in V vs (Orion fluoride reversible electrode)

Reference: K. Doblhofer, D.M. Mohilner. J. Phys. Chem. 75 (1971) 1698.

-E/V	γ/mN m^{-1}	-E/V	γ/mN m^{-1}	-E/V	γ/mN m^{-1}
1.6015	311.1	1.0000	399.8	0.4300	425.3
1.5801	315.3	0.9803	401.8	0.4134	424.8
1.5600	319.1	0.9603	403.7	0.3967	424.2
1.5401	322.9	0.9401	405.5	0.3809	423.5
1.5200	326.6	0.9202	407.2	0.3644	422.7
1.5001	330.2	0.9001	408.9	0.3505	421.8
1.4800	333.8	0.8800	410.6	0.3338	420.9
1.4601	337.3	0.8601	412.1	0.3192	419.9
1.4402	340.6	0.8402	413.6	0.3028	418.7
1.4200	344.1	0.8202	415.0	0.2885	417.5
1.4000	347.4	0.8001	416.3	0.2734	416.3
1.3804	350.5	0.7800	417.6	0.2589	415.0
1.3601	353.7	0.7615	418.7	0.2439	413.6
1.3400	356.9	0.7400	419.8	0.2292	412.1
1.3200	359.9	0.7208	420.9	0.2143	410.6
1.3001	362.9	0.7001	421.8	0.1997	408.9
1.2801	365.8	0.6803	422.7	0.1847	407.2
1.2601	368.6	0.6600	423.5	0.1704	405.5
1.2401	371.4	0.6401	424.2	0.1562	403.6
1.2203	374.1	0.6200	424.8	0.1424	401.8
1.2000	376.8	0.6001	425.3	0.1288	399.8
1.1800	379.4	0.5800	425.8	0.1147	397.7
1.1600	381.9	0.5601	426.1	0.1014	395.7
1.1401	384.3	0.5400	426.2	0.0878	393.6
1.1204	386.7	0.5200	426.4	0.0742	391.4
1.1001	389.1	0.5081	426.4	0.0611	389.1
1.0805	391.3	0.5052	426.5	0.0482	386.7
1.0600	393.6	0.4773	426.2	0.0358	384.3
1.0400	395.7	0.4629	426.1	0.0232	381.9
1.0201	397.8	0.4461	425.8	0.0108	379.3

Electrocapillary data on mercury electrode in LiBr at 25°C - given are the interfacial tension in mN m^{-1}, potential in V vs Ag/AgBr electrode.

Reference: E.W. Hermann. Ph.D. Dissertation, Carleton University, 1970.

c/mol l^{-1}					γ /mN m^{-1}						
-E/V	.001	.002	.005	.010	.020	.050	.100	.200	.510	1.030	2.120
-.025				379.6	378.3	375.5	373.2	369.1	361.7	356.0	350.5
.000			389.4	386.9	385.8	383.5	381.4	377.5	370.4	365.1	359.6
.025			394.1	392.7	391.7	389.8	387.8	384.2	377.6	372.6	367.1
.050			398.2	397.6	396.6	395.0	393.2	389.8	383.5	378.7	373.4
.075		396.6	401.9	401.7	400.8	399.4	397.8	394.6	388.6	384.0	378.9
.100		400.5	405.3	405.4	404.7	403.3	401.8	398.8	393.1	388.7	383.7
.125		403.9	408.4	408.7	408.2	406.9	405.5	402.6	397.1	392.9	388.1
.150		406.9	411.2	411.8	411.3	410.1	408.8	406.1	400.8	396.7	392.1
.175		409.6	413.8	414.5	414.2	413.0	411.8	409.2	404.1	400.2	395.7
.200		412.1	416.2	417.0	416.8	415.6	414.5	412.1	407.1	403.4	399.0
.225		414.4	418.3	419.2	419.1	418.0	417.0	414.6	409.8	406.2	402.0
.250		416.6	420.3	421.1	421.0	420.0	419.1	416.8	412.2	408.8	404.8
.275		418.6	422.0	422.7	422.6	421.8	420.9	418.7	414.3	411.1	407.2
.300		420.5	423.4	424.1	424.0	423.3	422.4	420.3	416.1	413.0	409.3
.325	418.4	422.1	424.7	425.2	425.0	424.3	423.5	421.5	417.5	414.7	411.1
.350	421.6	423.3	425.7	426.0	425.8	425.2	424.4	422.4	418.6	416.0	412.6
.375	423.2	424.8	426.4	426.6	426.3	425.7	424.9	423.1	419.4	417.0	413.8
.400	424.3	425.8	427.0	427.0	426.6	426.0	425.2	423.4	419.9	417.7	414.7
.425	425.3	426.5	427.3	427.2	426.7	426.0	425.2	423.4	420.1	418.0	415.2
.450	426.3	427.0	427.4	427.2	426.6	425.8	425.0	423.2	420.1	418.1	415.5
.475	427.0	427.4	427.3	427.0	426.3	425.4	424.5	422.8	419.7	417.9	415.5
.500	427.3	427.5	427.1	426.7	425.9	424.8	423.8	422.1	419.1	417.4	415.2
.525	427.3	427.4	426.6	426.2	425.3	424.0	423.0	421.2	418.2	416.7	414.7
.550	427.1	427.2	426.1	425.6	424.6	423.1	422.0	420.1	417.1	415.7	413.9
.575	426.8	426.8	425.4	424.9	423.7	422.1	420.8	418.9	415.9	414.5	412.8
.600		426.3	424.7	424.1	422.7	420.9	419.5	417.5	414.4	413.1	411.5

Electrocapillary data on mercury electrode in LiBr at 25°C (continued)

$c/\text{mol l}^{-1}$					$\gamma/\text{mN m}^{-1}$						
$-E/V$	.001	.002	.005	.010	.020	.050	.100	.200	.510	1.030	2.120
.625		425.7	423.8	423.2	421.6	419.6	418.1	416.0	412.8	411.5	410.0
.650		424.9	422.9	422.1	420.4	418.2	416.6	414.3	411.0	409.7	408.3
.675		424.1	421.8	420.9	419.1	416.6	414.9	412.5	409.1	407.8	406.4
.700		423.2	420.7	419.7	417.6	415.0	413.1	410.6	407.1	405.7	404.4
.725		422.2	419.4	418.3	416.1	413.2	411.3	408.5	404.9	403.4	402.1
.750		421.1	418.1	416.8	414.4	411.4	409.3	406.4	402.6	401.1	399.7
.775		419.9	416.6	415.1	412.7	409.4	407.2	404.2	400.2	398.5	397.1
.800		418.6	415.1	413.4	410.8	407.3	404.9	401.8	397.7	395.9	394.5
.825		417.1	413.4	411.6	408.8	405.1	402.6	399.3	395.1	393.2	391.6
.850		415.6	411.5	409.7	406.7	402.8	400.2	396.8	392.3	390.3	388.7
.875		414.0	409.5	407.6	404.5	400.4	397.6	394.1	389.5	387.4	385.6
.900		412.2	407.4	405.5	402.2	397.9	395.0	391.3	386.6	384.3	382.5
.925		410.2	405.2	403.3	399.8	395.3	392.2	388.5	383.6	381.2	379.2
.950		408.1	402.9	400.9	397.3	392.6	389.4	385.5	380.5	378.0	375.8
.975		405.8	400.6	398.5	394.7	389.8	386.5	382.5	377.3	374.6	372.4
1.000			398.2	396.0	392.0	386.9	383.4	379.4	374.0	371.2	368.8
1.025			395.6	393.3	389.2	384.0	380.3	376.1	370.6	367.7	365.1
1.050			393.0	390.6	386.3	380.9	377.1	372.8	367.1	364.1	361.3
1.075				387.3	383.3	377.8	373.9	369.4	363.6	360.4	357.4
1.100				384.8	380.2	374.5	370.5	365.9	359.9	356.6	353.4
1.125				381.7	377.0	371.2	367.0	362.3	356.1	352.7	349.3
1.150				378.5	373.7	367.8	363.5	358.5	352.2	348.7	345.0
1.175				375.2	370.4	364.2	359.8	354.7	348.3	344.6	340.7
1.200				371.3	366.9	360.5	356.0	350.8	344.3	340.4	336.3
1.225				368.3	363.3	356.8	352.2	346.8	340.0	336.1	331.7
1.250				364.8	359.7	352.9	348.2	342.7	335.7	331.7	327.1
1.275				361.1	355.9	348.9	344.1	338.5	331.3	327.1	322.4
1.300				357.4	352.0	344.8	339.9	334.1	326.8	322.5	317.6
1.325				353.7	348.0	340.7	335.6	329.7	322.2	317.8	312.7

Electrocapillary data on mercury electrode in LiI at 25°C - given are the interfacial tension in mN m^{-1}, potential in V vs Ag/AgI electrode.

Reference: E.W. Hermann. Ph.D.Dissertation, Carleton University, 1970.

c/mol l^{-1}					γ /mN m^{-1}						
-E/V	0.001	0.002	0.005	0.010	0.020	0.050	0.100	0.200	0.510	2.040	2.170
0.000				373.3	369.4	362.6	357.8	352.3	344.8	338.8	331.7
0.025			388.4	382.9	379.9	373.3	368.5	363.0	355.4	349.4	342.0
0.050			395.6	392.2	388.4	382.1	377.4	372.0	364.5	358.5	351.0
0.075		405.1	401.7	398.8	395.2	389.4	384.8	379.6	372.2	366.2	358.6
0.100	411.6	409.5	406.8	404.1	400.8	395.4	391.1	386.1	378.8	372.9	365.2
0.125	415.1	413.3	411.0	408.5	405.5	400.4	396.3	391.6	384.5	378.5	370.9
0.150	418.0	416.4	414.5	412.1	409.4	404.5	400.7	396.2	389.2	383.3	375.8
0.175	420.4	419.0	417.3	415.1	412.6	407.9	404.4	400.0	393.3	387.4	379.9
0.200	422.4	421.1	419.5	417.5	415.2	410.8	407.4	403.2	396.6	390.9	383.5
0.225	423.9	422.8	421.2	419.5	417.3	413.0	409.9	405.7	399.4	393.7	386.5
0.250	425.0	424.0	422.5	420.9	418.9	414.8	411.8	407.8	401.7	396.1	389.0
0.275	425.8	424.8	423.4	422.0	420.0	416.2	413.3	409.4	403.5	398.0	391.0
0.300	426.3	425.3	423.9	422.6	420.7	417.1	414.4	410.6	404.8	399.5	392.7
0.325	426.6	425.6	424.1	422.9	421.1	417.7	415.0	411.3	405.8	400.6	393.9
0.350	426.6	425.5	424.1	422.9	421.1	417.9	415.3	411.7	406.4	401.4	394.9
0.375	426.4	425.3	423.8	422.6	420.8	417.7	415.2	411.7	406.6	401.8	395.4
0.400	426.0	424.9	423.2	422.0	420.2	417.2	414.7	411.4	406.5	401.9	395.7
0.425	425.5	424.3	422.5	421.2	419.3	416.4	414.0	410.8	406.0	401.6	395.6
0.450	424.9	423.5	421.6	420.1	418.3	415.3	413.0	409.9	405.3	401.1	395.2
0.475	424.2	422.7	420.6	418.9	417.0	414.0	411.8	408.7	404.3	400.3	394.6
0.500	423.3	421.7	419.4	417.6	415.5	412.5	410.3	407.3	403.0	399.2	393.6
0.525	422.3	420.6	418.1	416.1	414.0	410.7	408.6	405.7	401.4	397.8	392.4
0.550	421.3	419.3	416.6	414.5	412.2	409.3	406.6	403.8	399.6	396.1	390.9
0.575	420.1	412.0	415.0	412.7	410.4	406.8	404.6	401.7	397.6	394.2	389.1
0.600	418.8	416.6	413.4	410.9	408.4	404.7	402.3	399.5	395.4	392.1	387.1

Electrocapillary data on mercury electrode in LiI at 25°C (continued)

c/mol l^{-1}	γ/mN m^{-1}										
-E/V	.001	.002	.005	.010	.020	.050	.100	.200	.510	2.040	2.170
.625	417.4	415.0	411.6	409.0	406.3	402.4	400.0	397.0	393.0	389.7	384.8
.650	416.0	413.4	409.8	406.9	404.1	400.0	397.5	394.4	390.3	387.2	382.4
.675	414.4	411.6	407.8	404.8	401.9	397.6	394.8	391.7	387.6	384.4	379.7
.700	412.7	409.7	405.7	402.6	399.5	395.0	392.1	388.9	384.6	381.4	376.7
.725	410.9	407.8	403.6	400.3	397.0	392.4	389.3	385.9	381.5	378.3.	373.6
.750	409.0	405.7	401.3	397.8	394.4	389.7	386.4	382.9	378.3	375.0	370.3
.775	407.0	403.6	398.9	395.3	391.7	386.8	383.4	379.7	375.0	371.5	366.9
.800				392.7	389.0	383.9	380.3	376.4	371.5	368.0	363.3
.825				389.9	386.1	380.9	377.1	373.1	368.0	364.3	359.5
.850				387.1	383.2	377.8	373.9	369.7	364.3	360.5	355.6
.875				384.2	380.1	374.5	370.6	366.2	360.6	356.6	351.6
.900				381.2	377.0	371.2	367.1	362.6	356.8	352.6	347.5
.925				378.1	373.8	367.7	363.6	358.9	352.9	348.5	343.2
.950				374.9	370.4	364.2	360.0	355.2	348.9	344.4	338.8
.975				371.7	367.0	360.6	356.2	351.3	344.9	340.2	334.4
1.000				368.3	363.5	356.9	352.4	347.4	340.7	335.8	329.8
1.025				364.9	359.9	353.1	348.5	343.4	336.5	331.4	325.2
1.050				361.3	356.2	349.2	344.4	339.2	332.2	326.9	320.4
1.075				357.7	352.4	345.2	340.3	335.0	327.8	322.2	315.5
1.100				353.9	348.4	341.1	336.0	330.6	323.2	317.5	310.5
1.125				350.1	344.4	337.0	331.7	326.1	318.6	312.6	305.4
1.150				346.2	340.3	332.7	327.2	321.5	313.8	307.6	300.2
1.175				342.2	336.1	323.3	322.7	316.8	308.9	302.5	294.8
1.200					331.8	323.8	318.0	312.0	303.9	297.2	289.3
1.225					327.5	319.2	313.2	307.0	298.7	291.8	283.7
1.250					323.0	314.4	308.4	302.0	293.4	286.2	277.9
1.275					313.5	309.6	303.4	296.8	288.0	280.6	272.0
1.300					313.8	304.6	298.3	291.5	282.4	274.9	266.0

Electrocapillary data on mercury electrode in KI at various concentrations at 18°C - given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: O. Essin and B. Markov. Acta Physicochimica URSS, 10 (1939) 35

c/mol l^{-1}	γ /mN m^{-1}								
-E/V	0.01	0.05	0.1	0.2	0.5	0.7	1.0	1.5	3.0
1.40	348.9	341.7	340.0	336.3	334.3	333.0	332.2	331.3	329.4
1.30	363.2	358.1	356.8	353.1	351.2	350.5	349.8	348.9	347.1
1.20	377.1	372.6	369.5	367.7	366.4	366.1	365.5	364.7	362.3
1.10	389.0	385.4	387.5	381.7	380.2	379.9	378.9	378.3	374.8
1.00	399.3	396.6	395.4	393.3	391.9	391.3	390.6	389.2	385.3
0.90	408.4	406.4	405.3	403.0	400.6	399.6	398.0	395.9	390.8
0.88	-	-	-	-	-	-	-	-	391.3
0.86	411.6	409.9	408.5	405.3	403.1	402.0	399.8	397.7	391.7
0.84	412.9	411.0	410.0	407.7	403.8	402.7	400.5	397.8	391.7
0.82	414.3	412.3	410.9	408.7	404.4	403.4	400.8	398.0	391.7
0.80	415.7	413.1	412.2	409.6	405.2	403.8	400.8	397.9	391.0
0.78	417.0	414.8	413.2	410.2	405.4	403.8	400.8	397.5	
0.76	417.8	416.0	414.0	410.9	405.4	403.8	400.7	397.1	
0.74	419.0	416.8	414.7	411.0	405.4	403.7	400.4	396.4	
0.72	420.1	417.3	414.8	411.0	404.9	402.7	396.3	394.7	
0.70	420.5	417.9	414.8	410.9	404.3	461.5	398.4	394.3	
0.68	421.4	418.0	414.8	410.6	403.3	401.1	397.3	-	
0.66	421.9	418.0	414.7	409.7	401.9	-	395.8	-	
0.64	422.0	417.7	414.0	408.9	-	-	-		
0.62	422.3	417.3	413.2	-					
0.60	422.3	416.5	412.1	-					
0.58	422.3	415.6	410.5	-					
0.56	422.2	414.1	-						
0.54	421.6	-							
0.52	420.5	-							
0.50	419.4	-							

Electrocapillary data on mercury electrode in 0.5 M KF - aqueous solutions. Given are the interfacial tension in mN m^{-1}, potential in V vs SCE aq. T = 25°C

Reference: S. Minc and M. Brzostowska. Roczniki Chem. 40 (1966) 1759.

-E/V vs SCE aq.	γ/mN m^{-1}
0.10	405.4
0.20	418.5
0.30	424.0
0.40	426.5
0.45	427.4
0.50	426.3
0.60	425.2
0.70	419.1
0.80	412.8
0.90	402.5
1.00	395.2
1.10	380.1
1.20	367.9
1.30	351.3
1.40	332.7
1.50	315.1

Electrocapillary data on mercury electrode in NaF at different molality and temperature - given are the interfacial tension in mN m^{-1}, potential difference in mV vs NCE.

Reference: Nguyen Huu Cuong, C.V. d'Alkaine, A. Jenard and H.D. Hurwitz. J. Electroanal. Chem. 51 (1974) 377.

	γ/mN m^{-1}								
c/mol l^{-1}	10^{-1}			10^{-2}			10^{-3}		
T°C	15	25	35	15	25	35	15	25	35
-E/mV									
000	400.2	399.5	398.5	405.9	405.0	404.2	411.9	410.6	409.9
100	410.9	410.0	408.9	415.1	414.1	412.9	419.1	418.1	416.9
200	418.8	417.7	416.5	421.6	420.5	419.1	424.5	423.0	421.6
300	424.1	422.9	421.5	425.7	424.4	422.8	427.3	425.7	424.1
400	426.9	425.6	424.1	427.6	426.3	424.6	428.4	426.8	425.2
450	427.5	426.1	424.6	427.9	426.6	424.9	428.5	427.0	425.3
500	427.4	426.0	424.6	427.9	426.6	424.9	428.5	427.0	425.3
600	425.8	424.5	423.0	426.8	425.5	423.9	427.9	426.4	424.8
700	422.0	420.9	419.6	424.0	422.8	421.3	426.0	424.6	423.1
800	416.3	415.3	414.2	419.3	418.3	416.9	422.3	421.2	419.8
900	408.8	408.0	407.0	412.7	411.9	410.8	416.8	415.9	414.7
1000	399.5	399.0	398.2	404.4	403.8	403.0	409.4	408.8	407.9
1100	388.6	388.3	387.7	394.4	394.1	393.5	400.4	400.0	399.4
1200	376.1	376.1	375.7	382.8	382.8	382.5	389.7	389.6	389.3
1300	362.0	362.3	362.0	369.6	369.9	369.9	377.5	377.6	377.6
1400	346.3	346.8	346.8	354.7	355.3	355.7	363.7	364.0	364.3
1500	328.9	329.7	329.9	338.2	339.2	339.9	348.3	348.8	349.5

Electrocapillary data for 0.1 N $HClO_4$ solution.
Reference electrode: SCE (NaCl solution) T = 25°C

Reference: K.G. Baikerikar and Robert S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

E/V	γ /mN m^{-1}
-1.10	385.9
-1.05	392.1
-1.00	397.3
-0.95	402.0
-0.90	406.4
-0.85	410.4
-0.80	413.9
-0.75	417.3
-0.70	419.9
-0.65	422.2
-0.60	423.9
-0.55	425.1
-0.50	425.5
-0.45	425.4
-0.40	424.6
-0.35	423.1
-0.30	420.8
-0.25	417.9
-0.20	414.2
-0.15	410.2
-0.10	405.3
-0.05	400.0
0.0	394.0
0.05	387.7
0.10	380.4
0.15	372.7
0.20	364.0

Electrocapillary data for mercury electrode in aqueous solutions of various salts at 18°C - given are the interfacial tension in mN m^{-1}, potential in V vs 1 mol l^{-1} KCl calomel electrode.

Reference: G. Gouy. Ann. Chim. phys. 7 (1903) 145.

γ/mN m^{-1}	-E/V						
	1 mol l^{-1} KOH	1 mol l^{-1} NaOH	1 mol l^{-1} CsOH	1 mol l^{-1} RbOH	1 mol l^{-1} LiOH	0.06 mol l^{-1} $Sr(OH)_2$	0.22 mol l^{-1} $Ba(OH)_2$
350.0							
367.1							
384.1							
401.2	1.067					0.991	1.032
418.3	1.171	1.150	1.191	1.171	1.136	1.121	1.144
426.7							
435.2							
435.2							
426.7							
418.3	1.635	1.636	1.648	1.643	1.636	1.633	1.641
401.2	1.847	1.848	1.854	1.853	1.845	1.847	1.846
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.127	2.126	2.120	2.128	2.127	2.127	2.126
350.0	2.237	2.237	2.227	2.234	2.243	2.238	2.234
320.1	2.404	2.404	2.388	2.402	2.413	2.405	2.402
298.8	2.511	2.510	2.480	2.508	2.520	2.511	2.506
277.4		2.607	2.583	2.604	2.619	2.607	
256.1		2.696	2.667	2.695	2.711	2.695	
234.7						2.776	
213.4						2.854	
$-E_z$/V	1.38					1.36	1.37
γ_z/mN m^{-1}	425.9	426.3	425.9	426.3	426.5	426.3	426.3

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	-E/V						
	0.01 mol l^{-1} KNO_3	0.1 mol l^{-1} KNO_3	1 mol l^{-1} KNO_3	0.33 mol l^{-1} $Ba(NO_3)_2$	0.005 mol l^{-1} $Sr(NO_3)_2$	0.05 mol l^{-1} $Sr(NO_3)_2$	0.5 mol l^{-1} $Sr(NO_3)_2$
350.0			0.740				0.744
367.1		0.752	0.830		0.701	0.767	0.836
384.1	0.789	0.851	0.941	0.937	0.804	0.868	0.948
401.2	0.909	0.982	1.082	1.072	0.928	1.000	1.090
418.3	1.089	1.172	1.302	1.294	-	-	-
426.7							
435.2							
435.2							
426.7							
418.3	1.619	1.623	1.623	1.616	-	-	-
401.2	1.842	1.844	1.849	1.846	1.847	1.850	1.846
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.133	2.129	2.126	2.125	2.126	2.127	2.123
350.0		2.241	2.237	2.234			2.234
320.1				2.400			
298.8							
277.4							
256.1							
234.7							
213.4							
$-E_z$/V	1.34	1.39	1.46	1.45	1.37	1.41	1.46
γ_z/mN m^{-1}	426.5	425.4	422.2	422.5	426.5	426.5	421.6

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	-E/V						
	0.5 mol l^{-1} $Ca(NO_3)_2$	5.2 mol l^{-1} sat $Ca(NO_3)_2$	14.6 mol l^{-1} KNO_2	1 mol l^{-1} $NaNO_2$	0.5 mol l^{-1} $Ca(NO_2)_2$	sat $Na_2B_4O_7$	0.5 mol l^{-1} K_2CO_3
350.0	0.743	0.830					
367.1	0.837	0.934		0.838	0.848	0.800	
384.1	0.951	1.061	0.935	0.926	0.930	0.874	0.910
401.2	1.099	1.235	1.028	1.049	1.057	0.969	0.997
418.3	1.336		1.181	1.231	1.254	1.144	1.133
426.7							
435.2							
435.2							
426.7							
418.3	1.614		1.695	1.638	1.631	1.631	1.644
401.2	1.847	1.838	1.870	1.848	1.850	1.845	1.848
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.124		2.111	2.124	2.128	2.131	2.125
350.0	2.235		2.209		2.230	2.241	2.236
320.1	2.404		2.355		2.409		2.403
298.8							2.507
277.4							
256.1							
234.7							
213.4							
$-E_z$/V	1.48	1.54	1.43	1.43	1.44	1.38	1.37
γ_z/mN m^{-1}	421.2	412.6	428.4	424.0	423.7	425.8	427.3

Electrocapillary data on mercury in various salts (continued)

γ/mN m^{-1}	-E/V						
	5.9 mol l^{-1} sat K_2CO_3	0.5 mol l^{-1} Na_2CO_3	sat $KHCO_3$	1 mol l^{-1} KCNO	sat KCNO	0.1 mol l^{-1} K_2HPO_4	1 mol l^{-1} K_2HPO_4
350.0			0.766	0.832	0.897		
367.1			0.836	0.862	0.926		0.842
384.1	0.907	0.912	0.931	0.933	0.999	0.856	0.905
401.2	0.962	1.004	1.052	1.053	1.123	0.946	0.990
418.3	1.038	1.148	1.223	1.240		1.102	1.131
426.7	1.086						
435.2	1.148						
435.2	1.562						
426.7	1.659						
418.3	1.747	1.643	1.671	1.641		1.631	1.652
401.2	1.883	1.843	1.856	1.851	1.856	1.846	1.852
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.103	2.126	2.119	2.125	2.117	2.128	2.124
350.0	2.198	2.236	2.223	2.233	2.219	2.239	2.233
320.1	2.345	2.403	2.387	2.399	2.374	2.407	2.397
298.8	2.438	2.510	2.490	2.502	2.475	2.152	2.500
277.4		2.605					
256.1		2.694					
234.7							
213.4							
$-E_z$/V	1.33	1.38	1.44	1.44	1.48	1.35	1.38
γ_z/mN m^{-1}	445.5	427.1	426.1	424.1	421.2	428.0	427.8

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	-E/V						
	6.4 mol l^{-1} K_2HPO_4	1 mol l^{-1} KH_2PO_4	1 mol l^{-1} Na_3PO_4	sat Na_2HPO_4	sat $K_4P_2O_7$	0.5 mol l^{-1} K_2HPO_3	0.5 mol l^{-1} Na_2HPO_3
350.0		0.733			0.866	0.761	0.746
367.1	0.805	0.807	0.898		0.904	0.818	0.811
384.1	0.855	0.892	0.934	0.891	0.962	0.901	0.893
401.2	0.914	0.997	0.999	0.975	1.049	1.016	1.005
418.3	0.992	1.145	1.123	1.118	1.179	1.198	1.180
426.7	1.040				1.277		
435.2	1.104						
435.2	1.563						
426.7	1.662				1.585		
418.3	1.742	1.645	1.649	1.645	1.704	1.637	1.639
401.2	1.880	1.850	1.851	1.850	1.869	1.848	1.848
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.105	2.125	2.124	2.127	2.112	2.124	2.124
350.0	2.198	2.236	2.233	2.238	2.210	2.234	2.236
320.1	2.347	2.402	2.397	2.407	2.366	2.400	2.404
298.8			2.502	2.511	2.464	2.505	2.510
277.4							
256.1							
234.7							
213.4							
$-E_z$/V	1.31	1.39	1.37	1.36	1.43	1.41	1.40
γ_z/mN m^{-1}	445.0	427.1	428.0	427.1	431.0	425.6	426.3

Electrocapillary data on mercury in various salts (continued)

γ/mN m^{-1}	-E/V						
	1 mol l^{-1} NaH_2PO_2	sat NaH_2PO_2	sat $Ba(H_2PO_2)_2$	0.5 mol l^{-1} $Ca(H_2PO_2)_2$	sat $Ca(H_2PO_2)_2$	0.5 mol l^{-1} $Mg(H_2PO_2)_2$	0.01 mol l^{-1} K_3AsO_4
350.0							
367.1	0.828	0.998		0.784	0.818	0.780	
384.1	0.909	1.080	0.965	0.886	0.922	0.882	0.873
401.2	1.026	1.240	1.095	1.020	1.063	1.019	0.951
418.3	1.222		1.357	1.239	1.350	1.241	
426.7							
435.2							
435.2							
426.7							
418.3	1.607		1.512	1.591	1.518	1.591	
401.2	1.842	1.791	1.832	1.838	1.833	1.837	1.842
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.128	2.149	2.135	2.130	2.138	2.132	
350.0	2.240	2.276	2.249	2.243	2.251	2.245	
320.1	2.407	2.453	2.428	2.413	2.421	2.418	
298.8	2.512	2.564	2.527	2.518	2.529	2.522	
277.4							
256.1							
234.7							
213.4							
$-E_z$/V	1.41	1.50	1.43	1.41	1.43	1.41	1.33
γ_z/mN m^{-1}	423.1	423.1	418.8	422.4	419.0	422.4	426.8

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	-E/V						
	0.1 mol l^{-1} K_3AsO_4	1 mol l^{-1} K_3AsO_4	0.01 mol l^{-1} K_2HAsO_4	0.1 mol l^{-1} K_2HAsO_4	1 mol l^{-1} K_2HAsO_4	0.01 mol l^{-1} KH_2AsO_4	0.1 mol l^{-1} KH_2AsO_4
277.4							
298.8							
320.1							
350.0						0.584	0.685
367.1		0.942	0.711	0.817	0.876	0.654	0.759
384.1	0.940	0.981	0.807	0.887	0.945	0.742	0.846
401.2	1.013	1.048	0.905	0.978	1.030	0.855	0.955
418.3		1.145			1.151		
418.3		1.663			1.659		
401.2	1.840	1.855	1.845	1.849	1.854	1.834	1.847
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.127	2.119	2.130	2.128	2.121	2.131	2.128
350.0		2.223		2.240	2.229	2.247	2.241
320.1		2.385			2.391		
298.8		2.488			2.495		
277.4							
256.1							
$-E_z$/V	1.38	1.38	1.35	1.34	1.38	1.31	1.37
γ_z/m Nm^{-1}	426.9	428.7	426.8	426.9	427.9	426.8	426.7

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	-E/V						
	1 mol l^{-1} KH_2AsO_4	0.33 mol l^{-1} Na_2HAsO_4	0.5 mol l^{-1} K_2SO_4	0.0005 mol l^{-1} Na_2SO_4	0.05 mol l^{-1} Na_2SO_4	0.5 mol l^{-1} Na_2SO_4	0.005 mol l^{-1} Li_2SO_4
277.4						0.519	
298.8					0.503	0.544	
320.1			0.603		0.545	0.606	
350.0		0.825	0.714		0.654	0.714	
367.1	0.866	0.856	0.788	0.650	0.728	0.786	
384.1	0.956	0.923	0.873	0.743	0.817	0.871	
401.2	1.066	1.007	0.978	0.863	0.928	0.975	0.886
418.3	1.217	1.133	1.133	1.035	1.089	1.130	1.047
418.3	1.645	1.645	1.645	1.610	1.626	1.638	1.619
401.2	1.850	1.848	1.852	1.833	1.843	1.848	1.841
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.127	2.125	2.125	2.122	2.129	2.124	
350.0	2.236	2.236	2.236	2.237	2.242	2.237	
320.1		2.404	2.402		2.412	2.405	
298.8		2.509	2.506		2.520	2.510	
277.4		2.604					
256.1							
$-E_z$/V	1.42	1.37	1.37	1.31	1.34	1.37	1.32
γ_z/mN m^{-1}	426.3	427.0	427.4	426.6	426.8	427.4	426.6

Electrocapillary data on mercury in various salts (continued)

γ/mN m^{-1}	-E/V					
	0.5 mol l^{-1} Li$_2$SO$_4$	0.005 mol l^{-1} MgSO$_4$	0.05 mol l^{-1} MgSO$_4$	0.5 mol l^{-1} MgSO$_4$	dens.1.21 MgSO$_4$	0.00166 mol l^{-1} Al$_2$(SO$_4$)$_3$
298.8	0.545				0.553	
320.1	0.606				0.622	0.522
350.0	0.714			0.720	0.735	0.631
367.1	0.789		0.757	0.794	0.811	0.709
384.1	0.874	0.795	0.831	0.880	0.897	0.799
401.2	0.978	0.899	0.936	0.989	0.999	0.911
418.3	1.128	1.067	1.098	1.130	1.133	1.076
418.3	1.642	1.630	1.633	1.640	1.658	1.639
401.2	1.847	1.844	1.848	1.849	1.854	1.848
384.1	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.129	2.128	2.128	2.129	2.124	2.131
350.0	2.241	2.244	2.239	2.241	2.232	2.242
320.1	2.410		2.409	2.408	2.398	2.417
298.8	2.520		2.517	2.515	2.503	
277.4	2.617			2.611	2.599	
256.1				2.700	2.687	
$-E_z$/V	1.37	1.34	1.35	1.37	1.38	1.35
γ_z/mN m^{-1}	427.1	426.6	426.8	427.1	428.8	426.9

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	-E/V						
	0.0166 mol l^{-1} $Al_2(SO_4)_3$	0.166 mol l^{-1} $Al_2(SO_4)_3$	sat $Al_2(SO_4)_3$	0.5 mol l^{-1} $BeSO_4$	0.5 mol l^{-1} $KOHSO_4$	sat $Na_2S_2O_6$	0.5 mol l^{-1} Na_2SO_3
298.8	0.500		0.528	0.518	0.526		
320.1	0.563	0.605	0.601	0.584	0.592		
350.0	0.672	0.716	0.721	0.695	0.704		
367.1	0.746	0.792	0.801	0.772	0.782	0.805	
384.1	0.832	0.878	0.893	0.859	0.873	0.897	1.006
401.2	0.940	0.982	1.000	0.965	0.987	1.011	1.097
418.3	1.101	1.133	1.141	1.119	1.155	1.179	1.215
418.3	1.640	1.642	1.677	1.636	1.641	1.649	1.642
401.2	1.850	1.849	1.855	1.848	1.848	1.852	1.847
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.125	2.127	2.120	2.126	2.130	2.127	2.128
350.0	2.237	2.238		2.238		2.235	2.239
320.1	2.406	2.407		2.410		2.404	
298.8	2.515			2.517		2.509	2.513
277.4							
256.1							2.697
$-E_z$/V	1.37	1.37	1.39	1.36	1.39	1.40	1.41
γ_z/mN m^{-1}	427.0	427.1	430.5	427.1	426.3	426.7	425.4

Electrocapillary data on mercury in various salts (continued)

γ/mN m^{-1}				-E/V			
	0.5 mol l^{-1} $Na_2S_2O_3$	sat $Na_2S_2O_3$	0.5 mol l^{-1} K_2SeO_4	conc $Na_8SiW_{12}O_{42}$	0.01 mol l^{-1} Na acetate	0.1 mol l^{-1} Na acetate	1 mol l^{-1} Na acetate
298.8			0.582	0.838			
320.1			0.642	0.900			
350.0			0.750	0.991			0.721
367.1			0.826	1.063		0.722	0.790
384.1	1.213	1.301	0.914	1.154	0.756	0.812	0.885
401.2	1.297	1.405	1.024	1.266	0.859	0.931	1.008
418.3	1.428		1.179	1.466	1.040	1.108	1.191
418.3	1.633		1.638	1.634	1.619	1.628	1.635
401.2	1.848	1.835	1.847	1.848	1.842	1.846	1.848
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.128	2.126	2.124	2.127	2.130	2.128	2.125
350.0	2.237	2.234	2.232		2.244	2.241	2.236
320.0	2.407					2.412	2.404
298.8	2.512					2.519	2.510
277.4							2.605
$-E_z$/V	1.52	1.60	1.39	1.55	1.32	1.36	1.41
γ_z/mN m^{-1}	422.0	413.5	427.3	410.4	426.7	426.5	424.6

Electrocapillary data on mercury in various salts (continued)

γ/mN m^{-1}	$-E$/V					
	0.5 mol l^{-1} Mg-acetate	0.005 mol l^{-1} K-oxalate	0.05 mol l^{-1} K-oxalate	0.5 mol l^{-1} K-oxalate	0.5 mol l^{-1} KH-oxalate	0.005 mol l^{-1} K-succinate
298.8						
320.1				0.711		
350.0				0.752	0.716	
367.1	0.779		0.757	0.819	0.801	0.682
384.1	0.875	0.786	0.844	0.905	0.908	0.778
401.2	1.001	0.899	0.953	1.011	1.046	0.895
418.3	1.186	1.071	1.112	1.159	1.248	
418.3	1.628	1.619	1.627	1.649	1.629	
401.2	1.845	1.840	1.846	1.852	1.847	1.844
384.1	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.130	2.127	2.127	2.126	2.128	2.132
350.0	2.240	2.240	2.240	2.236		
320.0	2.409		2.408	2.402		
298.8	2.516		2.517	2.507		
277.4				2.601		
$-E_z$/V	1.40	1.33	1.35	1.39	1.44	1.34
γ_z/mN m^{-1}	424.6	426.5	426.7	427.1	423.9	426.7

Electrocapillary data on mercury in various salts (continued)

γ/mN m^{-1}	-E/V						
	0.05 mol l^{-1} K-succinate	0.5 mol l^{-1} K-succinate	sat K-succinate	sat $KC_2H_5SO_4$	0.1 mol l^{-1} KSCNO	1 mol l^{-1} KSCNO	0.01 mol l^{-1} NaSCNO
298.8							
320.1		0.701					
350.0	0.692	0.751	0.784	0.809	0.955	1.118	
367.1	0.754	0.817	0.841	0.910	1.013	1.190	0.838
384.1	0.845	0.909	0.928	1.018	1.092	1.289	0.913
401.2	0.962	1.026	1.045	1.275	1.205	1.450	1.016
418.3		1.197	1.216		1.427		
418.3		1.643	1.671		1.584		
401.2	1.845	1.850	1.861	1.697	1.841	1.820	1.842
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.128	2.124	2.114	2.155	2.129	2.131	2.131
350.0	2.241	2.234	2.215	2.290	2.242	2.242	
320.1		2.400	2.369	2.469	2.411	2.408	
298.8		2.504	2.467	2.573	2.516	2.513	
277.4						2.608	
256.1						2.695	
$-E_z$/V	1.38	1.41	1.44	1.48	1.50	1.63	1.40
γ_z/mN m^{-1}	426.6	426.1	426.9	402.8	419.8	407.9	425.6

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	-E/V					
	0.1 mol l^{-1} NaSCNO	1 mol l^{-1} NaSCNO	0.5 mol l^{-1} Ba(SCNC)$_2$	0.5 mol l^{-1} Mg(SCNC)$_2$	0.0025 mol l^{-1} $K_4Fe(CN)_6$	0.25 mol l^{-1} $K_4Fe(CN)_6$
298.8						
320.1		1.013	1.045	1.019		
350.0	0.940	1.097	1.134	1.113		
367.1	0.996	1.170	1.210	1.189		0.907
384.1	1.073	1.267	1.313	1.290	0.839	0.977
401.2	1.181	1.422	1.490	1.453	0.942	1.067
418.3					1.089	1.137
418.3					1.615	1.641
401.2	1.846	1.826	1.805	1.818	1.840	1.849
384.1	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.129	2.129	2.131	2.130		2.127
350.0	2.242	2.241	2.243	2.242		2.238
320.1		2.407	2.410	2.411		
298.8		2.512	2.514	2.517		
277.4		2.608				
256.1		2.698				
$-E_z$/V	1.49	1.62	1.65	1.63	1.33	1.40
γ_z/mN m^{-1}	420.8	408.8	405.8	407.5	426.6	426.9

Electrocapillary data on mercury in various salts (continued)

γ /mN m^{-1}	$-E$/V					
	0.166 mol l^{-1} $K_6Co(CN)_{12}$	0.05 mol l^{-1} $K_2Pt(CN)_4$	0.5 mol l^{-1} $K_2Pt(CN)_4$	0.05 mol l^{-1} $BaPt(CN)_6$	0.005 mol l^{-1} $MgPt(CN)_4$	0.05 mol l^{-1} $MgPt(CN)_4$
298.8						
320.1			0.847	0.755		0.741
350.0		0.870	1.022	0.921	0.825	0.905
367.1		0.985	1.156	1.044	0.936	1.026
384.1	1.004	1.125	1.334	1.200	1.066	1.177
401.2	1.157	1.320		1.445	1.237	1.396
418.3	1.423					
418.3	1.589					
401.2	1.849	1.837		1.802	1.844	1.823
384.1	2.000	2.000	2.000	2.000	2.000	2.000
367.1		2.128	2.148	2.130	2.127	2.127
350.0		2.240	2.262	2.239	2.237	2.240
320.1		2.408	2.429	2.408		2.408
298.8		2.514	2.530	2.513		2.516
277.4						
256.1						
$-E_z$/V	1.51	1.59	1.68	1.63	1.55	1.62
γ_z/mN m^{-1}	409.6	411.3	398.1	405.4	416.5	407.5

Electrocapillary data on mercury in various salts (continued)

γ/mN m^{-1}	-E/V						
	0.5 mol l^{-1} $MgPt(CN)_4$	sat $MgPt(CN)_4$	1 mol l^{-1} KCl	0.01 mol l^{-1} NaCl	0.1 mol l^{-1} NaCl	1 mol l^{-1} NaCl	1 mol l^{-1} LiCl
298.8	0.738	0.812					
320.1	0.835	0.920					
350.0	1.019	1.117	0.899			0.891	0.889
367.1	1.158	1.269	0.943		0.825	0.938	0.933
384.1	1.345	1.485	1.018	0.788	0.898	1.016	1.012
401.2			1.125	0.898	1.004	1.123	1.120
418.3			1.292			1.287	1.286
418.3			1.639			1.642	1.636
401.2			1.851	1.843	1.844	1.850	1.849
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.148	2.176	2.125	2.129	2.128	2.126	2.127
350.0	2.262	2.298	2.233		2.240	2.235	2.237
320.1	2.431	2.466	2.398		2.410	2.401	2.406
298.8	2.535	2.570	2.503		2.518	2.507	2.511
277.4							
256.1							
$-E_z$/V	1.69	1.75	1.46	1.33	1.39	1.46	1.45
γ_z/mN m^{-1}	388.4	392.1	424.1	426.8	426,4	424.3	424.1

Electrocapillary data for mercury in various salts (continued)

γ/mN m^{-1}	-E/V						
	1 mol l^{-1} RbCl	1 mol l^{-1} CsCl	0.5 mol l^{-1} $SrCl_2$	0.5 mol l^{-1} $BaCl_2$	0.5 mol l^{-1} $CaCl_2$	4.6 mol l^{-1} $CaCl_2$	0.5 mol l^{-1} $MgCl_2$
298.8							
320.1						0.908	
350.0		0.911	0.898		0.929	0.989	0.832
367.1	0.936	0.949	0.948		0.952	1.065	0.937
384.1	1.011	1.021	1.028	1.040	1.033	1.166	1.017
401.2	1.117	1.130	1.137		1.142	1.310	1.124
418.3	1.283	1.300	1.307	1.317	1.310		1.291
418.3	1.638	1.627	1.636	1.641	1.639		1.635
401.2	1.851	1.847	1.849		1.850	1.849	1.848
384.1	2.000	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.124	2.122	2.123		2.128	2.117	2.125
350.0	2.233	2.227	2.233		2.237	2.216	2.234
320.1	2.398	2.394	2.399		2.402	2.368	2.402
298.8		2.496	2.503		2.507	2.463	2.507
277.4					2.602		2.604
256.1					2.689		2.691
$-E_z$/V	1.45	1.46	1.46	1.47	1.47	1.58	1.45
γ_z/mN m^{-1}	424.1	422.9	423.2	423.1	423.2	413.9	424.1

Electrocapillary data for mercury in various salts (continued)

γ/mN m^{-1}	-E/V					
	1 mol l^{-1} KBr	1 mol l^{-1} NaBr	6.9 mol l^{-1} sat NaBr	1 mol l^{-1} RbBr	1 mol l^{-1} LiBr	0.5 mol l^{-1} $SrBr_2$
298.8						
320.1	1.030	1.033	1.099		1.012	
350.0	1.091	1.084	1.175	1.073	1.065	1.082
367.1	1.126	1.134	1.244	1.118	1.111	1.132
384.1	1.196	1.208	1.343	1.183	1.183	1.209
401.2	1.303	1.320	1.506	1.286	1.290	1.321
418.3						
418.3						
401.2	1.845	1.847	1.819	1.849	1.846	1.847
384.1	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.123	2.127	2.125	2.124	2.126	2.126
350.0	2.234	2.236	2.229	2.231	2.239	2.235
320.1	2.400	2.403	2.383	2.394	2.402	2.401
298.8	2.504	2.506	2.479		2.513	2.505
277.4	2.599	2.603				
256.1	2.687					
$-E_z$/V	1.55	1.57	1.66	1.55	1.55	1.57
γ_z/mN m^{-1}	417.7	417.7	405.4	418.2	417.3	416.0

Electrocapillary data for mercury in various salts (continued)

γ/mN m^{-1}	-E/V				
	0.5 mol l^{-1} $BaBr_2$	0.5 mol l^{-1} $MgBr_2$	0.005 mol l^{-1} $CaBr_2$	0.05 mol l^{-1} $CaBr_2$	0.5 mol l^{-1} $CaBr_2$
298.8					
320.1					1.028
350.0	1.084	1.072	0.883	0.983	1.080
367.1	1.134	1.120	0.917	1.017	1.130
384.1	1.210	1.194	1.963	1.075	1.205
401.2	1.324	1.305	1.051	1.164	1.318
418.3					
418.3					
401.2	1.846	1.846	1.849	1.850	1.847
384.2	2.000	2.000	2.000	2.000	2.000
367.1	2.125	2.127	2.127	2.127	2.126
350.0	2.232	2.237	2.237	2.237	2.237
320.1	2.399	2.405			2.401
298.8	2.504	2.508			2.506
$-E_z$/V	1.57	1.56	1.40	1.47	1.57
γ_z/mN m^{-1}	414.3	416.9	425.8	423.3	416.0

Electrocapillary data for mercury in various salts (continued)

γ/mN m^{-1}	$-E$/V					
	0.01 mol l^{-1} KI	0.1 mol l^{-1} KI	1 mol l^{-1} KI	1 mol l^{-1} NaI	1 mol l^{-1} LiI	0.5 mol l^{-1} CaI_2
298.8			1.282		1.274	
320.1		1.182	1.311	1.329	1.310	
350.0		1.229	1.369	1.393	1.391	
367.1		1.266	1.423	1.452	1.442	1.458
384.1	1.163	1.318	1.505	1.543	1.518	1.558
401.2	1.229	1.400				
418.3						
418.3						
401.2	1.844	1.844				
384.1	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.126	2.127	2.131	2.135	2.132	2.135
350.0		2.238	2.241	2.249	2.244	2.250
320.1		2.406	2.405	2.415	2.413	
298.8			2.509	2.518	2.520	
$-E_z$/V	1.48	1.59	1.73	1.75	1.74	1.76
γ_z/mN m^{-1}	422.6	414.5	401.1	399.8	400.2	398.5

Electrocapillary data for mercury electrode in various acids at 18°C. Given are the interfacial tension in mN m^{-1}, potential in V vs 1 mol l^{-1} KCl calomel electrode.

Reference: G. Gouy. Ann. Chim. Phys. 7 (1903) 145.

γ/mN m^{-1}	1 mol l^{-1} HNO_3	1 mol l^{-1} H_3PO_4	0.5 mol l^{-1} H_3PO_4	1 mol l^{-1} H_3PO_2	0.005 mol l^{-1} H_2SO_4	0.05 mol l^{-1} H_2SO_4
277.4	0.513					0.428
298.8	0.571				0.407	0.463
320.1	0.650				0.463	0.523
350.0	0.786		0.701		0.573	0.633
367.1	0.884	0.752	0.783	0.797	0.649	0.710
384.1	1.000	0.857	0.877	0.889	0.739	0.803
401.2	1.149	0.982	1.001	1.024	0.853	0.916
418.3	1.406	1.155	1.203	1.284	1.017	1.083
418.3	1.650	1.625	1.604	1.519	1.614	1.620
401.2	1.899	1.844	1.840	1.829	1.837	1.841
384.1		2.000	2.000	2.000	2.000	2.000
367.1		2.132	2.131	2.139	2.135	2.133
350.0		2.255		2.253		
320.1						
298.8						
277.4						
$-E_z$/V	1.53	1.38	1.40	1.40	1.30	1.34
γ_z/mN m^{-1}	420.3	425.8	425.0	419.9	426.6	426.5

Electrocapillary data for mercury in various acids (continued)

γ/mN m^{-1}	-E/V						
	0.5 mol l^{-1} H_2SO_4	1 mol l^{-1} H_2SO_4	2 mol l^{-1} H_2SO_4	5 mol l^{-1} H_2SO_4	1 mol l^{-1} oxalic acid	0.5 mol l^{-1} oxalic acid	sat succinic acid
277.4	0.465	0.464	0.460	0.478			
298.8	0.508	0.513	0.515	0.550			0.553
320.1	0.574	0.584	0.590	0.636			0.616
350.0	0.692	0.706	0.717	0.779		0.790	0.743
367.1	0.772	0.789	0.804	0.882	0.597	0.883	0.846
384.1	0.871	0.889	0.908	1.000	0.698	1.000	0.996
401.2	0.989	1.013	1.036	1.151	0.841	1.154	1.271
418.3	1.164	1.191	1.222	1.382	1.178	1.421	
418.3	1.633	1.632	1.650	1.730	1.429	1.688	
401.2	1.847	1.847	1.855	1.935	1.816	1.957	1.708
384.1	2.000	2.000	2.000		2.000		2.000
367.1	2.132	2.123			2.143		2.182
350.0		2.221					
320.1							
298.8							
277.4							
$-E_z$/V	1.39	1.40	1.43	1.56	1.30	1.55	1.49
γ_z/mN m^{-1}	426.3	425.6	425.4	422.8	419.0	421.2	416.9

Electrocapillary data for mercury in various acids (continued)

γ/mN m^{-1}	$-E$/V				
	1 mol l^{-1} HCNS	0.05 mol l^{-1} $H_2Pt(CN)_4$	0.01 mol l^{-1} HCl	0.1 mol l^{-1} HCl	1 mol l^{-1} HCl
277.4					
298.8					
320.1	0.992	0.707			
350.0	1.104	0.871			0.888
367.1	1.184	0.987	0.720	0.824	0.937
384.1	1.293	1.135	0.795	0.897	1.018
401.2	1.464	1.344	0.903	1.003	1.130
418.3			1.061		1.308
418.3			1.617		1.627
401.2	1.724	1.837	1.840	1.843	1.846
384.1	2.000	2.000	2.000	2.000	2.000
367.1			2.126	2.130	2.128
350.0				2.244	
320.1					
298.8					
277.4					
$-E_z$/V	1.64	1.61	1.32	1.38	1.46
γ_z/mN m^{-1}	406.6	410.1	426.5	426.1	422.9

Electrocapillary data for mercury in various acids (continued)

γ/mN m^{-1}	-E/V				
	2 mol l^{-1} HCl	0.01 mol l^{-1} HBr	0.1 mol l^{-1} HBr	1 mol l^{-1} HBr	0.01 mol l^{-1} HI
277.4					
298.8					1.035
320.1				1.022	1.057
350.0		0.843	0.952	1.073	1.102
367.1		0.873	0.985	1.121	1.130
384.1	1.065	0.928	1.039	1.196	1.168
401.2	1.184	1.002	1.130	1.307	1.231
418.3	1.391	1.148	1.286		1.346
418.3	1.606	1.623	1.619		1.618
401.2	1.846	1.844	1.842	1.845	1.841
384.1	2.000	2.000	2.000	2.000	2.000
367.1		2.125	2.129	2.129	2.126
350.0					
298.8					
277.4					
$-E_z$/V	1.50	1.37	1.44	1.56	1.46
γ_z/mN m^{-1}	420.5	425.8	423.7	416.1	422.9

Electrocapillary data on mercury electrode in various salts, bases and acids at 18°C. Given are the interfacial tension in $mN\ m^{-1}$, potential in V vs KCl calomel electrode.

Reference: G. Gouy. Ann. Chim. Phys. 8 (9)(1906) 75.

	-E/V					
γ /mN m^{-1}	1 mol l^{-1} NH_4OH	10 mol l^{-1} NH_4OH	0.5 mol l^{-1} $(NH_4)_2SO_4$	1 mol l^{-1} $(NH_4)_2HPO_4$	sat NH_4NO_3	1 mol l^{-1} NH_4NO_2
298.8			0.549		0.618	
320.1			0.612		0.699	
350.0			0.721	0.762	0.835	
367.1			0.795	0.821	0.933	
384.1			0.880	0.895	1.056	0.927
401.2			0.986	0.986	1.237	1.053
418.3	1.091		1.145	1.121	-	1.252
418.3	1.619	-	1.640	1.656	-	1.628
401.2	1.842	1.793	1.849	1.853	1.830	1.847
384.1	2.000	2.000	2.000	2.000	2.000	2.000
367.1	2.133	2.163	2.124	2.123	2.125	2.127
350.0	2.250	2.300	2.232	2.232	2.230	2.237
320.1		2.500	2.397	2.393	-	2.403
298.3		2.567	2.501	2.496	-	
277.4			2.594			
256.1			2.681			
E_z /V	1.35	-	1.38	1.38	1.54	1.44
γ_z/mN m^{-1}	425.0	414.3	427.6	427.6	412.0	423.3

Electrocapillary data on mercury in various salts, bases and acids

(continued)

	-E/V				
$\gamma/\mathrm{mN\ m^{-1}}$	0.166 mol $\mathrm{l^{-1}}$ $(NH_4)_4H_2(CO_3)_3$	sat $(NH_4)_4H_2(CO_3)_3$	1 mol $\mathrm{l^{-1}}$ NH_4CNS	1 mol $\mathrm{l^{-1}}$ NH_4Br	sat $(NH_4)HC_2O_4$
298.8					
320.1			1.032	1.040	-
350.0		0.761	1.124	1.090	0.747
367.1	0.799	0.845	1.199	1.137	0.816
384.1	0.887	0.942	1.300	1.208	0.905
401.2	1.012	1.075	1.468	1.317	1.013
418.3	1.178	1.254	-	-	1.164
418.3	1.642	1.652	-	-	1.642
401.2	1.850	1.853	1.803	1.847	1.849
384.1	2.000	2.000	2.000	2.000	2.000
367.1	2.127	2.123	2.121	2.126	2.127
350.0	2.235	2.229	2.230	2.234	2.236
320.1	2.401	2.391	2.396	2.398	2.402
298.3	2.503	2.492	2.498	2.499	2.505
277.4				2.592	2.598
256.1					
E_z/V	1.40	1.45	1.63	1.56	1.39
$\gamma_z/\mathrm{mN\ m^{-1}}$	426.3	425.2	406.6	416.7	426.8

Electrocapillary data on mercury electrode in $NaBF_4$ at 25°C.
Given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: A. de Battisti, and S. Trasatti. J. Electroanal. Chem. 59 (1975) 137.

c/mol l^{-1}	γ /mN m^{-1} 0.01	0.02	0.04	0.07	0.15	0.27	0.50	0.65	1.0
E/V									
.20	373.5	373.0	372.9	372.8	371.5	371.6	370.6	374.0	368.0
.15	381.2	380.2	380.6	380.5	378.6	378.6	378.4	376.9	375.6
.10	388.0	387.0	387.8	386.4	385.5	385.8	385.7	383.4	382.8
.05	394.7	394.5	394.7	394.0	391.9	392.4	392.3	390.7	389.2
.00	400.3	399.9	399.9	399.8	398.5	398.3	398.1	396.4	395.1
-.05	404.3	405.0	404.9	405.4	404.0	403.6	403.2	401.7	400.5
-.10	409.6	408.7	409.6	409.3	408.6	408.3	408.1	406.4	405.3
-.15	413.4	412.5	413.8	413.3	412.4	412.4	412.3	410.7	410.2
-.20	416.7	415.9	417.1	417.0	416.6	416.2	415.8	414.6	413.0
-.25	419.5	419.2	419.6	419.8	419.4	419.5	419.2	417.3	416.4
-.30	421.2	421.6	422.0	422.1	421.8	421.9	421.8	419.4	419.0
-.35	422.7	423.0	423.9	423.7	423.1	423.6	423.6	421.7	420.8
-.40	423.9	424.0	424.7	424.7	423.8	424.6	424.4	422.4	421.8
-.45	424.4	424.3	425.4	424.9	424.1	424.8	424.7	422.8	422.1
-.50	424.2	424.2	425.0	424.6	424.0	424.3	424.3	422.4	421.8
-.55	424.0	423.9	424.3	424.1	423.5	423.3	423.0	421.8	420.7
-.60	423.2	423.2	422.8	422.6	421.7	421.8	421.5	419.8	418.9
-.65	422.1	421.7	421.3	420.4	419.7	419.7	419.1	417.6	416.9
-.70	420.1	419.8	419.6	418.0	417.3	417.1	416.6	414.7	414.3
-.75	418.3	417.3	416.6	415.2	414.3	413.6	413.5	412.0	410.8
-.80	415.8	414.3	413.6	412.5	411.1	409.9	410.1	407.9	407.0
-.85	412.5	411.0	410.0	408.7	407.1	406.0	406.1	404.4	402.6
-.90	408.8	407.2	406.3	405.0	402.6	401.6	401.4	399.8	398.4
-.95	404.9	403.2	401.8	400.1	397.9	396.7	396.4	394.6	393.4
-1.00	400.4	398.5	396.9	395.5	393.0	391.4	390.9	389.2	387.9
-1.05	395.7	393.3	391.8	390.1	387.5	386.0	384.9	383.4	382.1
-1.10	390.4	388.3	386.4	384.3	382.7	380.0	378.6	377.3	376.0
-1.15	384.4	382.4	380.6	378.1	375.7	373.6	372.5	370.9	369.1
-1.20	387.7	376.5	374.2	371.6	369.2	366.6	365.8	363.7	362.1
-1.25	372.6	369.8	367.4	364.8	361.9	359.4	358.5	356.4	354.8
-1.30	366.1	362.6	360.5	357.7	354.4	352.0	351.1	348.7	347.0
-1.35	358.1	355.4	352.8	350.1	346.7	343.9	342.8	340.9	338.5
-1.40	351.0	347.9	345.1	342.2	338.3	335.8	334.1	332.4	330.0
-1.45	343.4	339.6	336.7	333.7	329.6	327.3	325.0	324.1	320.7
-1.50	333.9	331.0	327.9	322.8	320.9	317.4	315.6	313.6	311.1

Electrocapillary data on mercury in aqueous solutions of sodium maleate.
T = 25°C

Potential with respect to saturated calomel electrode in contact with 0.1 M solution of salt.

Values on positive branch in dilute solution are unreliable owing to 'dewetting'.

Reference: R. Parsons and J.T. Reilly (unpublished)

	γ /mN m^{-1}					
c/mol l^{-1} E/volts	0.010	0.016	0.025	0.040	0.063	0.1000
0.15	366.2	364.4	361.3	360.3	362.3	360.9
0.10	376.0	375.0	375.2	375.0	373.9	372.2
0.05	384.3	384.6	383.9	384.5	383.5	381.8
0.00	391.5	391.3	390.8	391.5	392.6	392.0
-0.05	398.2	397.6	397.7	397.9	399.1	398.6
-0.10	403.8	404.0	403.7	404.6	405.2	405.0
-0.15	409.0	409.4	409.4	410.1	410.5	410.6
-0.20	413.5	413.7	413.9	414.4	415.1	415.1
-0.25	417.8	417.6	417.6	418.7	418.7	418.7
-0.30	420.9	420.8	420.8	421.7	421.6	421.5
-0.35	423.4	422.9	423.0	423.5	423.5	423.5
-0.40	424.9	424.2	424.6	424.8	424.6	425.1
-0.425	425.3	424.8	424.8	425.1	425.0	425.3
-0.45	425.5	424.9	424.8	425.4	425.2	425.4
-0.475	425.6	424.9	424.9	425.4	425.2	425.5
-0.50	425.3	424.6	424.6	425.0	425.0	425.3
-0.525	424.9	424.3	424.4	424.7	424.5	425.0
-0.550	424.5	423.9	423.9	424.2	424.3	424.4
-0.575	424.1	423.1	423.4	423.4	423.6	423.9
-0.60	423.4	422.4	422.5	422.8	422.9	423.2
-0.65	421.6	420.7	420.8	421.1	420.6	421.3
-0.70	419.6	418.8	418.7	418.2	418.5	418.6
-0.75	417.1	415.7	415.6	415.7	415.4	415.6
-0.80	413.7	412.5	412.6	412.7	412.1	412.1
-0.85	410.3	408.9	408.7	408.5	408.1	408.2
-0.90	406.5	405.1	404.6	405.5	403.7	404.1
-0.95	401.8	400.4	400.0	399.6	399.1	399.1
-1.00	397.3	395.8	394.9	395.2	394.1	393.9
-1.05	392.2	390.2	389.8	389.6	388.6	388.8
-1.10	386.0	384.7	383.9	383.9	382.8	382.8
-1.15	380.2	378.9	377.8	378.3	376.3	376.7
-1.20	374.3	371.9	367.5	371.4	369.7	369.1
-1.25	367.3	365.3	364.5	363.6	363.5	362.0

Electrocapillary data for aqueous sodium maleate (cont.)

	γ/mN m^{-1}				
c/mol l^{-1}	0.160	0.250	0.400	0.630	1.000
E/volts					
0.15		355.2	352.3	349.7	
0.10	369.8	367.9	365.5	363.8	360.5
0.050	380.6	378.2	376.1	374.3	371.9
0.00	388.9	387.3	385.6	383.8	380.7
-0.05	396.3	395.0	393.1	391.6	388.3
-0.10	403.1	401.5	399.8	398.0	395.4
-0.15	408.6	407.2	405.6	404.0	401.1
-0.20	413.3	411.8	410.4	409.0	406.5
-0.25	417.0	415.5	414.2	413.1	409.7
-0.30	420.1	418.5	417.2	416.0	413.5
-0.35	422.4	420.9	419.8	418.3	416.1
-0.40	423.6	422.5	421.4	420.2	418.2
-0.425	424.2	423.3	422.2	420.8	418.9
-0.45	424.2	423.6	422.4	421.0	419.3
-0.475	424.4	423.6	422.6	421.4	419.7
-0.50	424.6	423.6	422.7	421.8	420.0
-0.525	424.3	423.3	422.6	421.7	420.0
-0.550	423.8	423.0	422.4	421.5	419.9
-0.575	423.2	422.5	421.9	421.2	419.7
-0.60	422.6	421.8	421.1	420.8	419.5
-0.65	420.4	419.9	419.7	419.0	417.8
-0.70	417.7	417.2	417.1	416.8	415.7
-0.75	414.8	414.4	414.1	413.9	412.9
-0.80	411.5	411.1	410.6	410.5	409.8
-0.85	407.6	406.9	406.7	406.7	405.9
-0.90	402.8	402.5	402.4	402.2	401.6
-0.95	398.7	397.6	397.2	397.4	396.9
-1.00	393.2	392.7	392.3	392.1	391.6
-1.05	387.4	386.8	386.6	386.3	385.9
-1.10	381.3	380.6	380.3	380.2	379.8
-1.15	375.1	374.1	373.7	373.7	373.4
-1.20	367.6	368.0	366.8	366.4	366.8
-1.25	360.8	359.8	359.5	359.4	359.3

Electrocapillary data on mercury in aqueous solutions of sodium fumarate.
T = 25°C

Potential with respect to saturated calomel electrode in contact with 0.1 M solution of salt.

Values on positive branch in dilute solution are unreliable owing to 'dewetting'.

Reference: R. Parsons and J.T. Reilly (unpublished)

	γ /mN m^{-1}					
c/mol l^{-1}	0.010	0.016	0.025	0.040	0.063	0.100
E/volts						
0.20	353.0					
0.15	363.3	367.5		362.8	362.5	360.3
0.10	372.5	375.9		372.2	372.0	369.6
0.05	382.2	383.5		380.8	379.3	377.5
0.00	388.9	389.7	389.9	388.5	387.4	385.4
-0.05	394.9	396.1	395.9	394.7	394.2	392.1
-0.10	401.7	402.2	402.4	401.3	400.4	398.7
-0.15	406.9	407.7	407.8	406.9	406.1	403.0
-0.20	411.8	412.9	412.8	411.8	410.9	409.9
-0.25	416.2	417.0	417.0	416.1	415.6	414.4
-0.30	419.7	420.2	420.1	419.8	419.2	419.6
-0.35	422.2	422.6	422.3	422.2	422.1	420.9
-0.40	424.1	423.9	423.9	424.1	424.0	423.3
-0.425	424.7	424.9	424.7	424.7	424.7	424.1
-0.45	425.1	425.1	424.8	425.0	425.0	424.6
-0.475	425.1	425.3	424.9	425.0	425.0	424.7
-0.50	425.1	425.1	424.8	424.8	425.2	424.7
-0.525	424.9	424.8	424.5	424.6	425.0	424.5
-0.55	424.4	424.3	424.0	424.1	424.5	424.1
-0.575	423.9	423.8	423.3	423.6	423.9	423.0
-0.60	423.2	422.8	422.3	422.8	423.0	422.7
-0.65	421.7	421.1	420.6	420.9	421.2	420.7
-0.70	419.4	418.8	418.3	418.5	418.6	418.3
-0.75	416.8	416.1	415.4	415.7	415.6	415.2
-0.80	413.8	413.0	412.4	412.2	412.2	412.1
-0.85	410.7	409.6	408.4	408.5	408.4	407.8
-0.90	406.3	405.4	404.2	404.3	404.0	403.4
-0.95	402.4	400.8	399.8	399.1	399.3	398.9
-1.00	397.1	395.4	394.8	394.6	394.1	393.5
-1.05	393.5	390.6	389.1	388.9	388.4	387.7
-1.10	386.3	384.8	383.1	383.2	382.8	381.9
-1.15	380.2	379.0	377.4	376.8	376.6	375.9
-1.20	374.1	372.4	371.0	370.3	369.8	369.4
-1.25		365.6	364.4		362.7	361.9

Electrocapillary data for aqueous sodium fumarate (cont.)

			γ /mN m^{-1}		
c/mol l^{-1} E/volts	0.160	0.250	0.400	0.630	1.000
0.15	356.7	354.5	351.2		
0.10	366.6	364.6	361.5	358.7	355.0
0.05	374.9	374.5	369.8	367.9	364.6
0.00	382.1	380.8	377.9	375.6	374.0
-0.05	389.3	387.8	385.0	382.6	380.6
-0.10	396.1	394.5	391.7	389.7	387.8
-0.15	402.0	400.8	397.5	395.8	394.0
-0.20	407.5	406.2	403.2	401.5	399.9
-0.25	412.0	411.0	408.4	406.7	405.2
-0.30	416.1	415.0	412.8	411.2	409.9
-0.35	419.2	418.1	416.3	414.9	413.6
-0.40	421.9	420.9	419.3	417.8	416.7
-0.425	422.5	421.9	420.3	419.1	418.2
-0.450	423.0	422.4	421.1	420.0	419.4
-0.475	423.4	422.9	421.7	420.7	420.6
-0.50	423.5	423.1	422.2	421.2	420.7
-0.525	423.3	423.2	422.2	421.4	421.2
-0.55	423.1	423.0	422.0	421.4	421.4
-0.575	422.5	422.5	421.7	421.3	421.4
-0.60	421.8	421.8	420.9	420.8	421.0
-0.65	419.9	420.0	419.4	419.5	419.9
-0.70	417.4	417.4	416.9	417.1	417.8
-0.75	414.3	414.4	413.9	414.2	415.1
-0.80	410.8	410.8	410.2	410.8	411.8
-0.85	406.8	406.8	406.2	406.7	407.9
-0.90	402.3	402.4	402.4	402.1	403.4
-0.95	397.4	397.3	396.6	397.1	398.6
-1.00	392.9	392.0	391.2	391.4	393.2
-1.05	386.1	386.4	385.4	385.7	387.2
-1.10	380.6	380.1	379.2	379.9	380.9
-1.15	373.3	373.3	372.6	372.8	374.3
-1.20	367.3	366.8	365.7	365.9	367.0
-1.25			358.2	358.3	359.7

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 at 12°C. Given are the interfacial tension in mN m^{-1}, potential in V vs NCE.

Reference: J.A.V. Butler and C. Ockrent. J. Phys. Chem. 34 (1930) 2286.

	γ/mN m^{-1}						
E/V	0.5 M sodium sulphate (A)	0.5 M sodium sulphate (B)	0.01 M benzoic acid	0.002 M salicylic acid	0.002 M o-toluic acid	0.002 M m-toluic acid	sat. p-toluic acid
+0.1	368.7	364.5	366.2	364.7	367.5	366.6	366.7
+0.0	388.4	386.6	377.9	376.2	384.5	382.0	382.0
-0.1	403.2	402.4	387.0	386.6	394.9	390.8	390.9
-0.2	414.1	414.0	393.4	395.4	401.0	396.7	396.6
-0.3	422.3	422.2	397.7	402.4	405.0	401.0	400.4
-0.4	425.4	425.7	400.9	408.4	407.6	403.7	400.4
-0.5	426.2	426.2	401.7	412.1	410.8	405.6	407.6
-0.6	424.9	424.6	402.8	413.5	412.5	405.8	409.1
-0.7	420.8	420.5	401.7	413.3	414.2	406.7	409.2
-0.8	414.2	413.9	399.5	411.6	412.5	406.8	412.3
-0.9	406.0	405.4	395.4	404.4	404.7	403.8	404.4
-1.0	396.5	395.6	392.1	395.0	395.0	394.3	394.3
-1.1	384.1	384.0	382.8	383.6	382.9	383.6	382.8
-1.2	371.5	370.7	370.7	370.7	370.6	370.6	370.4
-1.3	356.6	355.9	356.3	356.1	356.1	356.3	355.8
-1.4	340.2	339.4	340.0	339.3	339.9	340.1	339.9
-1.5	322.1	320.9	322.3	321.7	322.0	321.8	321.9
-1.6	301.8	301.0	302.2	301.7	301.9	301.7	301.7
-1.7	280.3	279.7	280.5	280.2	280.2	280.1	279.8
-1.8	257.0	255.9	257.0	256.9	257.0	256.7	256.6
-1.9	231.2	230.2					
-2.0	203.3	202.3					

Electrocapillary data on mercury in 0.5 M Na_2SO_4 (continued)

E/V	γ/mN m^{-1} 0.1 M trichlor-acetic acid	0.1 M sodium o-toluate	0.1 M sodium m-toluate	0.1 M sodium p-toluate	0.1 M sodium hydrogen phthalate	0.1 M di-sodium phthalate	0.05 M di-sodium phthalate
+0.2	-	-	-	-	328.1	-	-
+0.1	-	355.4	-	346.0	350.1	350.2	349.8
0.0	381.6	370.1	363.6	363.8	364.0	370.4	373.6
-0.1	394.3	382.6	376.1	375.7	374.9	387.4	389.4
-0.2	404.1	392.2	385.3	385.3	383.1	400.4	401.6
-0.3	411.9	399.9	393.5	392.7	389.4	410.8	368.7
-0.4	417.4	405.0	399.2	397.9	394.6	417.5	417.9
-0.5	420.1	408.7	403.3	400.8	397.9	420.9	421.4
-0.6	420.0	409.8	405.0	402.5	399.8	421.2	421.4
-0.7	417.4	409.8	405.0	402.9	400.7	417.9	419.5
-0.8	412.4	407.1	402.1	401.2	399.4	412.9	412.9
-0.9	405.0	401.7	398.3	398.3	396.5	405.0	362.1
-1.0	395.3	393.8	391.8	392.4	391.0	395.4	395.0
-1.1	384.0	383.0	382.5	382.1	382.4	383.5	383.5
-1.2	371.4	370.8	370.1	370.8	370.3	370.5	370.1
-1.3	356.8	356.1	356.3	355.8	356.5	356.1	355.6
-1.4	340.2	339.9	339.8	339.4	339.9	339.4	339.1
-1.5	322.7	321.5	321.5	321.0	321.8	321.3	321.0
-1.6	302.7	301.3	301.2	301.4	301.8	300.9	300.9
-1.7	281.1	279.8	279.6	279.5	280.1	278.8	279.2
-1.8	257.1	256.0	255.9	256.0	256.5	255.3	255.9

Electrocapillary data on mercury in 0.5 M Na_2SO_4 (continued)

E/V	γ/mN m^{-1} 0.1 M sodium terephthalate	0.1 M sodium trichloracetate	0.1 M sodium trichloracetate	0.1 M potassium benzene sulphonate	0.05 M potassium benzene sulphonate
+0.2	-	-	-	-	-
+0.1	-	-	367.4	356.5	361.7
0.0	359.0	-	384.3	371.1	376.2
-0.1	376.8	396.1	397.1	382.3	387.0
-0.2	388.7	406.3	407.6	391.3	396.4
-0.3	398.0	414.6	415.1	398.3	403.1
-0.4	406.0	419.7	420.1	404.6	407.9
-0.5	412.0	422.1	422.5	409.0	411.6
-0.6	414.6	421.6	422.4	411.0	414.0
-0.7	414.4	418.4	419.3	410.8	413.0
-0.8	410.4	412.6	413.3	408.4	411.1
-0.9	403.8	405.2	405.6	402.5	404.3
-1.0		395.4	395.7	394.0	395.2
-1.1	383.3		384.2	382.8	383.5
-1.2	369.9		371.3	370.3	371.0
-1.3	356.1		356.6	356.0	355.8
-1.4	339.3		340.2	339.4	339.7
-1.5	322.1		322.1	321.4	321.7
-1.6	302.7		302.1	301.8	301.7
-1.7	-		280.4	279.7	279.8
-1.8	-		257.2	256.6	256.6

Electrocapillary data on mercury in 0.5 M Na_2SO_4 (continued)

	γ/mN m^{-1}					
E/V	0.1 M sodium maleate	0.1 M sodium fumarate	0.1 M maleic acid	0.1 M p-toluene sulphonic acid	0.1 M sodium p-toluene sulphonate	0.05 M p-toluene sulphonic acid
+0.4	-	-	-	280.2	-	277.6
+0.3	-	-	317.7	317.0	-	317.2
+0.2	-	-	345.2	323.7	339.1	341.8
+0.1	-	-	366.6	354.7	354.5	358.4
0.0	384.3	379.1	381.7	367.9	368.2	372.2
-0.1	399.4	393.5	392.4	378.8	378.8	381.7
-0.2	409.5	405.1	400.7	388.3	388.0	392.0
-0.3	417.5	414.7	407.3	395.4	395.4	399.9
-0.4	422.7	421.6	411.0	401.4	401.2	405.8
-0.5	424.6	423.9	413.3	406.2	406.3	409.8
-0.6	424.2	424.0	413.8	407.9	408.0	411.9
-0.7	420.6	420.8	412.6	408.3	408.4	411.6
-0.8	414.2	414.5	410.4	406.2	406.3	408.4
-0.9	406.2	406.2	403.8	400.9	400.8	402.9
-1.0	396.1	396.1	395.0	393.4	393.0	394.3
-1.1	383.4	384.5	383.7	382.8	382.7	383.6
-1.2	371.2	370.7	371.5	370.0	369.8	370.8
-1.3	356.1	355.8	357.4	356.1	355.2	355.9
-1.4	340.0	-	340.1	339.4	339.0	339.5
-1.5	321.8	321.3	322.2	321.4	320.9	321.6
-1.6	301.8	301.4	302.0	301.8	301.1	301.8
-1.7	279.9	280.0	280.7	280.3	279.2	280.3
-1.8	256.2	255.9	256.9	256.9	255.8	256.9

Electrocapillary data on mercury in 0.5 M Na_2SO_4 (continued)

	γ/mN m^{-1}				
E/V	0.1 M sodium benzoate	0.01 M sodium benzoate	0.1 M sodium cinnamate	0.1 M sodium salicylate	0.05 M sodium salicylate
+0.4	-	-	-	-	-
+0.3	-	-	-	-	-
+0.2	-	-	-	320.2	-
+0.1	355.0	363.2	-	339.2	343.2
0.0	367.8	379.7	356.5	354.1	357.8
-0.1	380.3	391.1	368.0	366.7	369.5
-0.2	391.0	401.2	377.9	376.1	384.0
-0.3	399.2	409.4	385.2	385.4	388.3
-0.4	405.0	415.0	389.9	392.4	395.1
-0.5	410.0	418.4	393.3	397.3	399.9
-0.6	412.1	419.1	394.8	400.5	403.1
-0.7	411.3	416.7	394.8	402.3	404.4
-0.8	408.1	411.3	392.9	401.2	403.2
-0.9	402.5	404.1	391.4	398.1	399.2
-1.0	394.3	394.5	387.2	391.6	391.8
-1.1	383.7	384.0	379.1	382.5	382.0
-1.2	371.1	371.0	369.1	370.7	369.4
-1.3	356.1	355.8	355.5	356.1	366.9
-1.4	339.4	340.0	339.2	339.6	338.9
-1.5	321.2	321.9	320.7	321.4	320.6
-1.6	301.4	301.8	301.1	301.0	300.7
-1.7	279.6	280.2	279.4	279.3	278.9
-1.8	255.8	256.7	255.9	255.8	297.3

Electrocapillary data on mercury in HCl, $BaCl_2$ and HCl + $BaCl_2$ mixtures. T = 25°. Potential measured with respect to saturated calomel electrode.

Given is the interfacial tension x 10^3 in $N.m^{-1}$.

Reference: K.M. Joshi and R. Parsons, Electrochim. Acta 4, 129 (1961)

E / Volts	0.2mHCl	E / Volts	0.2mHCl + 0.01mBaCl$_2$	E / Volts	0.2mHCl + 0.02mBaCl$_2$
0.019	375.8	0.000	379.5	0.000	378.4
0.014	376.9	-0.056	389.7	-0.056	389.4
-0.053	389.8	-0.104	397.6	-0.115	398.0
-0.102	397.4	-0.155	404.7	-0.158	404.3
-0.147	404.1	-0.207	410.8	-0.208	409.6
-0.200	410.5	-0.257	415.9	-0.259	414.4
-0.246	415.3	-0.304	419.7	-0.305	418.5
-0.304	419.9	-0.355	423.0	-0.365	422.6
-0.352	422.8	-0.405	424.9	-0.405	424.4
-0.402	424.9	-0.450	425.8	-0.467	425.7
-0.452	425.9	-0.504	426.2	-0.511	425.7
-0.500	426.0	-0.551	425.6	-0.558	425.1
-0.557	425.2	-0.605	424.0	-0.609	423.6
-0.604	424.0	-0.657	421.9	-0.656	421.6
-0.657	421.8	-0.705	419.5	-0.711	418.9
-0.715	418.7	-0.759	416.2	-0.814	412.1
-0.767	415.5	-0.808	412.7	-0.916	404.1
-0.803	413.4	-0.852	409.1	-1.026	392.7
-0.870	407.8	-0.905	404.5	-1.105	382.3
-0.906	404.6	-0.954	399.9	-1.210	370.5
-0.964	399.4	-1.006	394.2		
-1.013	394.0	-1.047	389.7		
-1.073	387.7	-1.105	382.7		
-1.113	382.5	-1.152	376.8		
-1.155	377.4	-1.228	366.7		
-1.202	371.4				

Electrocapillary data on mercury in aqueous HCl + $BaCl_2$ (cont.)

E / Volts	0.2m HCl + 0.05m$BaCl_2$	E / Volts	0.21mHCl + 0.1m$BaCl_2$	E / Volts	0.24mHCl + 0.12m$BaCl_2$
0.004	375.3	-0.006	374.3	-0.023	377.5
-0.073	388.6	-0.050	383.9	-0.069	386.5
-0.131	404.3	-0.110	394.2	-0.106	393.2
-0.233	412.1	-0.160	401.5	-0.156	400.3
-0.288	416.9	-0.208	407.8	-0.206	407.1
-0.330	420.0	-0.258	413.3	-0.255	412.7
-0.428	424.2	-0.304	417.5	-0.307	416.6
-0.481	425.1	-0.353	420.9	-0.352	420.7
-0.520	425.1	-0.408	423.6	-0.404	423.5
-0.578	424.1	-0.450	424.8	-0.454	425.1
-0.643	421.8	-0.505	425.4	-0.507	425.6
-0.727	418.0	-0.552	425.0	-0.550	425.1
-0.833	410.0	-0.604	423.6	-0.611	423.6
-0.938	401.4	-0.654	422.6	-0.650	421.9
-1.038	390.0	-0.712	418.8	-0.704	419.2
-1.133	378.0	-0.751	416.2	-0.752	416.4
-1.215	367.9	-0.810	412.1	-0.808	412.2
		-0.861	407.8	-0.862	407.7
		-0.904	404.1	-0.907	403.6
		-0.957	398.9	-0.957	398.7
		-1.005	393.4	-1.000	394.3
		-1.051	388.3	-1.058	386.9
		-1.110	380.8	-1.110	380.2
		-1.221	365.8	-1.154	374.5

Electrocapillary data on mercury in aqueous HCl + $BaCl_2$ (cont.)

E/Volts	0.05m$BaCl_2$	E/Volts	0.05m$BaCl_2$ + 0.05m HCl	E/Volts	0.05m$BaCl_2$ +0.1mHCl
-0.003	387.2	-0.012	382.2	-0.009	381.0
-0.059	395.1	-0.068	390.7	-0.060	389.7
-0.106	401.5	-0.109	396.6	-0.112	398.1
-0.155	407.4	-0.165	403.4	-0.165	404.9
-0.207	413.8	-0.212	409.1	-0.211	410.5
-0.256	418.3	-0.248	415.0	-0.265	415.2
-0.307	422.1	-0.315	420.2	-0.316	419.7
-0.354	424.4	-0.359	423.5	-0.367	423.3
-0.405	425.6	-0.414	425.3	-0.418	425.3
-0.456	426.4	-0.469	426.1	-0.472	426.2
-0.505	425.6	-0.513	426.2	-0.532	426.2
-0.562	424.4	-0.564	425.5	-0.586	425.1
-0.613	423.8	-0.617	424.1	-0.619	423.7
-0.670	421.2	-0.675	422.1	-0.664	421.6
-0.712	418.8	-0.715	419.6	-0.711	419.1
-0.762	415.6	-0.763	416.5	-0.760	416.6
-0.809	411.7	-0.818	413.0	-0.814	412.2
-0.860	408.1	-0.867	408.9	-0.888	408.3
-0.915	403.1	-0.908	403.9	-0.907	404.2
-0.960	396.6	-0.962	399.3	-0.952	399.7
-1.010	392.9	-1.010	394.0	-1.009	393.6
-1.056	387.4	-1.056	389.1	-1.053	388.3
-1.119	379.6	-1.115	382.0	-1.114	380.7
-1.166	373.6	-1.164	374.5	-1.150	376.4
-1.229	364.3	-1.254	364.1	-1.236	364.0

Electrocapillary data on mercury in aqueous HCl + $BaCl_2$ (cont.)

$\frac{E}{\text{Volts}}$	0.057m $BaCl_2$ + 0.3m HCl	$\frac{E}{\text{Volts}}$	0.065m $BaCl_2$ + 0.5m HCl
-0.011	375.7	-0.014	372.1
-0.056	385.3	-0.053	380.2
-0.108	394.1	-0.109	390.8
-0.150	400.3	-0.149	397.1
-0.203	407.2	-0.207	405.2
-0.247	412.2	-0.249	410.2
-0.308	417.8	-0.307	415.6
-0.354	420.7	-0.354	419.4
-0.409	423.6	-0.404	422.2
-0.449	424.7	-0.454	424.3
-0.505	425.4	-0.504	425.2
-0.546	425.2	-0.554	424.9
-0.609	423.7	-0.608	423.9
-0.650	422.1	-0.652	422.2
-0.707	419.5	-0.705	419.7
-0.752	416.5	-0.752	416.8
-0.806	412.5	-0.811	422.7
-0.860	408.4	-0.854	409.2
-0.908	403.8	-0.907	404.5
-0.956	399.3	-0.951	400.2
-1.008	393.2	-1.024	392.6
-1.051	388.7	-1.069	387.4
-1.111	381.6	-1.105	382.4
-1.155	376.0	-1.151	376.2
-1.240	363.4		

Electrocapillary data on mercury in aqueous 0,1 m $KClO_4$ at various temperatures.

Potential measured with respect to saturated calomel electrode.

Given is the interfacial tension $\gamma \times 10^3$ in $N.m^{-1}$.

Reference: K.M. Joshi, Thesis, University of Bristol, 1959.

20°		30°		40°		50°	
E / Volts	γ	E / Volts	γ	E / Volts	γ	E / Volts	γ
0.158	365.7	0.174	360.8	0.141	336.3	0.151	364.5
0.058	383.5	0.115	373.3	0.084	376.1	-0.003	390.4
-0.046	398.3	-0.004	392.1	-0.035	393.4	-0.102	401.7
-0.143	409.0	-0.110	404.7	-0.130	404.2	-0.206	412.3
-0.239	417.6	-0.209	414.3	-0.228	413.1	-0.307	418.6
-0.341	423.2	-0.313	420.6	-0.328	418.9	-0.410	422.6
-0.443	426.0	-0.411	424.0	-0.431	421.9	-0.459	422.7
-0.499	426.0	-0.478	424.3	-0.533	421.5	-0.576	423.0
-0.545	425.5	-0.528	424.4	-0.628	421.6	-0.620	420.5
-0.649	422.3	-0.630	421.8	-0.728	417.8	-0.720	
-0.748	417.2	-0.727	417.3	-0.819	411.4	-0.824	409.7
-0.845	410.5	-0.825	410.8	-0.924	402.4	-0.923	401.6
-0.942	401.5	-0.925	402.6	-1.021	392.6	-1.020	392.1
-1.040	391.2	-1.022	392.6	-1.123	380.2	-1.118	381.1
-1.143	378.6	-1.122	381.1	-1.226	365.9	-1.216	368.5
-1.241	366.4	-1.223	367.8	-1.327	351.2	-1.316	354.1
-1.327	351.4	-1.318	353.5	-1.420	334.5	-1.426	336.1
-1.431	333.9	-1.430	335.0	-1.534	314.8	-1.531	317.5
-1.526	316.6	-1.524	317.4	-1.647	291.9	-1.629	298.1
-1.643	292.3	-1.617	298.7	-1.758	273.3	-1.734	275.8
-1.732	271.8	-1.706	279.4	-1.858	255.7	-1.833	251.9
-1.871	244.1	-1.802	255.3				

Electrocapillary data on mercury in aqueous solutions of tetra ethyl ammonium perchlorate $(C_2H_5)_4N\ ClO_4$ (TEA). Concentration c_{TEA} in mole.l^{-1} indicated. Various temperatures. Potential measured with respect to saturated calomel electrode. Given is the interfacial tension $\gamma \times 10^3$ in N.m.$^{-1}$.

Reference: K.M. Joshi, Thesis, University of Bristol, 1959.

$c_{TEA} = 0.1$

20°		30°		40°		50°	
E / Volts	γ	E / Volts	γ	E / Volts	γ	E / Volts	γ
0.175	357.2	0.150	364.8	0.150	363.4	0.170	357.9
0.075	379.8	0.049	382.3	0.054	380.7	0.072	376.6
-0.083	400.0	-0.035	393.6	-0.052	394.6	-0.060	394.6
-0.190	408.8	-0.136	403.7	-0.152	407.5	-0.164	404.3
-0.286	412.4	-0.234	409.6	-0.252	409.9	-0.270	409.7
-0.390	412.4	-0.264	410.5	-0.348	411.4	-0.340	410.7
-0.492	410.0	-0.330	411.9	-0.400	410.9	-0.447	409.7
-0.528	408.8	-0.383	411.9	-0.491	408.7	-0.551	405.6
-0.594	402.1	-0.437	410.5	-0.595	403.4	-0.656	398.5
-0.694	394.4	-0.530	407.2	-0.698	396.3	-0.757	390.8
-0.785	386.7	-0.619	401.7	-0.798	388.4	-0.859	381.2
-0.885	376.2	-0.724	393.4	-0.885	379.4	-0.951	371.8
-0.986	367.5	-0.817	385.1	-0.987	367.2	-1.048	360.2
-1.085	353.6	-0.912	375.2	-1.078	356.9	-1.148	348.2
-1.199	339.8	-1.004	364.5	-1.174	344.8	-1.242	335.9
-1.289	326.3	-1.099	353.3	-1.263	332.9	-1.342	322.3
-1.388	312.1	-1.180	342.7	-1.363	318.8	-1.457	305.5
-1.490	296.4	-1.286	328.7	-1.469	303.0	-1.552	290.7
-1.592	280.2	-1.393	313.6	-1.587	284.5	-1.651	274.7
-1.696	262.8	-1.485	299.2	-1.682	268.6	-1.756	256.7
-1.804	243.9	-1.579	284.5	-1.782	251.1	-1.855	238.7
-1.908	223.8	-1.668	269.6	-1.881	233.3		
		-1.866	235.7				

Electrocapillary data on mercury in aqueous T.E.A. (cont.)

0.05m TEA

20°		30°		40°		50°	
E Volts	γ	E Volts	γ	E Volts	γ	E Volts	γ
0.134	379.2	0.170	372.7	0.180	372.5	0.174	375.7
0.075	387.2	0.066	389.7	0.068	389.9	0.068	391.6
-0.026	400.4	-0.033	401.6	-0.056	403.8	-0.031	401.7
-0.121	409.1	-0.133	409.9	-0.156	411.1	-0.115	409.2
-0.217	414.1	-0.234	414.2	-0.253	414.3	-0.215	413.3
-0.262	415.1	-0.329	414.7	-0.313	414.2	-0.314	413.3
-0.358	415.1	-0.438	412.1	-0.418	411.8	-0.414	410.4
-0.412	413.7	-0.535	406.7	-0.475	409.4	-0.509	405.7
-0.516	410.6	-0.637	400.4	-0.581	402.9	-0.613	398.0
-0.616	403.5	-0.730	392.2	-0.678	395.0	-0.706	391.2
-0.716	395.1	-0.828	382.5	-0.775	386.6	-0.799	381.6
-0.813	385.9	-0.924	371.6	-0.864	377.3	-0.887	372.2
-0.919	374.3	-1.020	360.1	-0.962	366.3	-0.976	361.9
-1.021	362.2	-1.122	347.1	-1.061	354.4	-1.071	349.4
-1.117	350.1	-1.240	331.4	-1.160	341.2	-1.170	336.9
-1.215	337.1	-1.335	317.4	-1.286	323.8	-1.272	322.6
-1.310	323.3	-1.449	300.1	-1.345	315.1	-1.372	308.1
-1.457	301.1	-1.540	285.3	-1.463	297.1	-1.489	289.5
-1.555	284.9	-1.640	268.6	-1.589	276.6	-1.590	272.6
-1.661	267.2	-1.735	251.7	-1.687	259.5	-1.686	256.0
-1.760	249.0	-1.878	224.0	-1.785	241.7	-1.807	233.7
-1.882	225.4			-1.889	219.8		

0.02m TEA

0.127	371.4	0.174	358.5	0.166	360.2	0.142	362.5
-0.087	401.4	0.068	380.2	0.062	378.1	0.078	374.3
-0.187	411.1	-0.026	392.5	-0.059	394.8	-0.067	394.7
-0.268	416.2	-0.121	404.3	-0.158	405.1	-0.169	405.2
-0.301	417.8	-0.280	415.7	-0.253	412.4	-0.269	412.2
-0.344	419.0	-0.322	417.3	-0.354	416.0	-0.370	415.5
-0.399	419.3	-0.376	418.1	-0.456	415.7	-0.466	414.9
-0.446	418.4	-0.630	417.4	-0.556	414.2	-0.571	410.8
-0.541	415.6	-0.532	413.9	-0.659	405.9	-0.669	404.6
-0.648	408.4	-0.633	407.9	-0.763	398.0	-0.781	396.3
-0.803	395.6	-0.734	400.0	-0.857	389.4	-0.879	387.4
-0.746	400.8	-0.837	390.6	-0.958	379.1	-0.977	376.9
-0.844	391.6	-0.937	380.3	-1.053	368.2	-1.075	366.0
-0.944	380.5	-1.037	369.1	-1.148	356.7	-1.173	353.9
-1.043	369.7	-1.148	355.3	-1.247	347.8	-1.266	341.9
-1.142	357.3	-1.293	336.1	-1.348	329.9	-1.367	327.8
-1.243	343.9	-1.360	326.3	-1.458	313.7	-1.483	310.3
-1.346	329.5	-1.476	309.4	-1.560	297.9	-1.581	295.6
-1.537	300.7	-1.572	291.6	-1.663	281.1	-1.681	279.1
-1.458	313.1	-1.672	277.9	-1.756	265.2	-1.815	255.8
-1.634	281.2	-1.771	260.0	-1.859	246.5	-1.888	242.7
-1.758	262.9	-1.865	244.2				

Electrocapillary data on mercury in aqueous T.E.A. (cont.)

0.01m TEA

20°		30°		40°		50°	
E/Volts	γ	E/Volts	γ	E/Volts	γ	E/Volts	γ
0.134	371.1	0.139	360.6	0.166	361.7	0.169	356.4
0.040	384.4	0.077	369.1	0.063	377.5	0.067	375.8
-0.074	399.1	-0.027	390.9	-0.040	391.5	-0.020	387.1
-0.171	409.7	-0.131	403.3	-0.145	403.5	-0.116	398.6
-0.274	417.0	-0.222	411.0	-0.248	411.9	-0.214	407.7
-0.372	420.0	-0.319	415.5	-0.349	416.2	-0.314	414.6
-0.473	418.8	-0.415	417.2	-0.452	416.4	-0.416	416.4
-0.577	414.2	-0.520	414.8	-0.554	413.2	-0.514	414.5
-0.674	407.5	-0.619	409.3	-0.663	407.7	-0.606	410.3
-0.774	399.3	-0.720	401.8	-0.762	400.0	-0.707	403.7
-0.882	389.9	-0.820	392.8	-0.858	391.1	-0.829	393.4
-0.973	386.3	-0.920	383.4	-0.963	380.3	-0.927	384.1
-1.069	367.9	-1.019	372.6	-1.061	369.1	-1.027	373.5
-1.164	355.7	-1.120	360.5	-1.158	357.4	-1.120	362.2
-1.264	342.4	-1.221	347.5	-1.261	343.8	-1.216	350.3
-1.359	329.3	-1.354	329.0	-1.359	330.4	-1.327	335.5
-1.454	315.1	-1.470	311.9	-1.476	315.8	-1.441	319.5
-1.547	300.8	-1.569	296.5	-1.582	297.0	-1.544	303.8
-1.646	284.4	-1.669	280.2	-1.679	281.2	-1.643	287.8
-1.761	264.5	-1.766	263.4	-1.778	263.9	-1.747	270.5
		-1.871	244.2	-1.876	346.4	-1.862	244.6

Electrocapillary data on mercury in aqueous mixtures of tetraethylammonium perchlorate $(C_2H_5)_4N\ ClO_4$ (TEA) and $KClO_4$ at various temperatures.

Potential measured with respect to saturated calomel electrode.

Given is the interfacial tension $\gamma \times 10^3$ in $N.m^{-1}$

Reference: K.M. Joshi, Thesis, University of Bristol, 1959.

0.05m TEA + 0.05m $KClO_4$

20°		30°		40°		50°	
E/Volts	γ	E/Volts	γ	E/Volts	γ	E/Volts	γ
0.165	367.6	0.162	361.9	0.147	365.6	0.189	358.1
0.059	384.8	0.064	382.4	0.041	384.4	0.084	377.7
-0.037	397.8	-0.033	395.7	-0.041	396.6	-0.082	399.3
-0.135	407.5	-0.135	405.9	-0.156	406.6	-0.176	407.4
-0.237	413.5	-0.235	411.5	-0.258	413.0	-0.283	412.6
-0.346	414.9	-0.346	413.8	-0.362	414.5	-0.381	413.9
-0.417	414.2	-0.451	412.8	-0.460	414.1	-0.482	412.1
-0.521	412.1	-0.551	409.3	-0.559	409.9	-0.578	408.0
-0.624	405.7	-0.657	403.0	-0.659	404.0	-0.683	401.4
-0.724	398.9	-0.757	394.8	-0.760	395.8	-0.781	393.1
-0.827	390.4	-0.864	384.6	-0.855	387.0	-0.885	381.1
-0.927	380.5	-0.985	372.4	-0.952	376.8	-0.983	373.3
-1.028	368.6	-1.079	360.3	-1.043	366.2	-1.079	362.4
-1.124	356.3	-1.176	347.8	-1.151	353.0	-1.179	349.7
-1.120	343.5	-1.280	335.3	-1.253	339.8	-1.281	336.3
-1.320	329.4	-1.378	320.7	-1.347	326.8	-1.379	322.4
-1.437	312.1	-1.483	305.0	-1.469	308.3	-1.469	308.9
-1.534	296.5	-1.576	289.9	-1.554	293.8	-1.569	293.2
-1.626	281.2	-1.684	272.7	-1.658	278.2	-1.666	277.3
-1.778	253.5	-1.777	255.2	-1.767	258.8	-1.769	259.0
		-1.884	232.4	-1.868	238.8	-1.874	238.4

Electrocapillary data on mercury in aqueous mixtures of tetraethyl ammonium perchlorate $(C_2H_5)_4N\ ClO_4$ (TEA) (cont.)

0.02m TEA + 0.08m $KClO_4$

20°		30°		40°		50°	
E Volts	γ	E Volts	γ	E Volts	γ	E Volts	γ
0.148	370.0	0.163	364.6	0.155	365.1	0.162	361.4
0.043	387.4	0.094	378.4	0.054	383.6	0.060	380.3
-0.049	399.1	-0.074	400.9	-0.037	395.8	-0.025	392.3
-0.157	410.8	-0.174	411.1	-0.151	408.0	-0.125	403.6
-0.257	417.2	-0.280	417.9	-0.239	414.6	-0.225	412.0
-0.359	419.8	-0.377	419.7	-0.393	418.7	-0.325	417.0
-0.413	419.8	-0.481	418.8	-0.463	417.9	-0.428	418.2
-0.510	417.9	-0.580	415.1	-0.570	414.6	-0.531	415.7
-0.610	413.3	-0.678	409.3	-0.678	408.0	-0.634	410.6
-0.706	407.1	-0.778	401.9	-0.778	401.3	-0.731	403.8
-0.805	398.8	-0.878	392.7	-0.880	390.9	-0.832	395.4
-0.906	388.9	-0.977	382.1	-0.986	379.6	-0.933	385.5
-0.999	378.9	-1.076	370.5	-1.067	370.3	-1.033	374.5
-1.082	368.9	-1.177	357.8	-1.167	358.2	-1.150	360.8
-1.177	357.1	-1.277	344.6	-1.266	345.0	-1.252	347.1
-1.273	344.2	-1.376	330.6	-1.362	331.4	-1.356	332.9
-1.374	329.8	-1.482	314.5	-1.460	316.8	-1.464	315.2
-1.479	313.7	-1.577	298.9	-1.559	301.1	-1.567	300.8
-1.578	297.5	-1.675	281.8	-1.663	283.3	-1.663	284.9
-1.680	279.3	-1.784	260.9	-1.762	264.7	-1.765	264.4
-1.777	260.8	-1.881	246.1	-1.857	244.5		
-1.888	235.2						

0.01m TEA + 0.09m $KClO_4$

0.170	362.2	0.163	361.1	0.166	361.5	0.153	363.2
0.079	379.2	0.097	373.9	0.053	381.3	0.053	381.6
-0.073	400.4	-0.023	391.2	-0.043	394.8	-0.021	392.0
-0.175	411.2	-0.127	402.7	-0.149	404.7	-0.127	403.6
-0.275	417.7	-0.233	410.8	-0.257	414.3	-0.230	412.2
-0.377	420.8	-0.338	416.9	-0.357	418.3	-0.330	417.0
-0.477	420.4	-0.434	416.9	-0.477	416.0	-0.456	418.6
-0.578	416.9	-0.539	414.9	-0.576	414.8	-0.558	415.7
-0.684	411.1	-0.638	410.0	-0.678	409.5	-0.659	410.4
-0.779	403.5	-0.739	402.8	-0.780	401.9	-0.760	403.1
-0.884	394.0	-0.840	394.1	-0.871	394.1	-0.862	394.6
-0.989	382.7	-0.939	384.1	-0.967	383.4	-0.962	384.6
-1.078	372.6	-1.028	376.2	-1.058	372.7	-1.058	374.1
-1.186	358.9	-1.124	364.0	-1.161	360.0	-1.155	361.7
-1.286	345.5	-1.222	351.6	-1.258	349.2	-1.262	348.4
-1.387	331.0	-1.320	337.9	-1.350	335.1	-1.360	335.2
-1.513	311.1	-1.439	319.7	-1.479	316.2	-1.462	319.3
-1.614	294.1	-1.539	303.9	-1.568	301.2	-1.560	303.9
-1.711	276.0	-1.638	286.7	-1.666	284.1	-1.657	286.5
-1.813	253.8	-1.734	267.9	-1.760	265.1	-1.757	267.5
-1.931	223.9	-1.842	244.0	-1.858	242.3	-1.862	243.7

Electrocapillary data on mercury in aqueous mixtures of tetraethyl ammonium perchlorate $(C_2H_5)_4N\ ClO_4$ (TEA) (cont.)

0.005m TEA + 0.1m $KClO_4$

20°		30°		40°		50°	
E/Volts	γ	E/Volts	γ	E/Volts	γ	E/Volts	γ
0.064	378.2	0.144	365.8	0.166	355.4	0.131	355.8
-0.055	394.2	0.050	381.7	0.074	373.5	0.034	375.1
-0.151	404.4	-0.057	395.7	-0.030	389.7	-0.068	391.9
-0.258	412.7	-0.149	406.8	-0.128	398.9	-0.159	403.1
-0.352	416.8	-0.249	414.1	-0.225	407.2	-0.263	413.1
-0.443	417.7	-0.348	420.6	-0.324	415.6	-0.366	420.0
-0.543	415.9	-0.444	422.4	-0.444	420.2	-0.418	422.1
-0.641	411.4	-0.495	422.4	-0.496	420.6	-0.518	423.4
-0.723	406.3	-0.604	419.4	-0.595	419.2	-0.620	421.7
-0.821	398.4	-0.694	414.6	-0.697	413.7	-0.718	417.0
-0.923	388.4	-0.795	406.7	-0.787	407.5	-0.822	410.1
-1.031	377.3	-0.886	398.3	-0.883	399.3	-0.919	401.6
-1.134	368.2	-0.979	389.0	-0.987	389.7	-1.028	390.9
-1.236	354.6	-1.073	377.9	-1.088	378.7	-1.129	380.1
-1.332	342.5	-1.173	366.4	-1.188	376.2	-1.225	368.3
-1.450	323.6	-1.271	353.6	-1.284	354.4	-1.328	354.9
-1.555	304.7	-1.368	339.9	-1.386	340.4	-1.425	341.6
-1.652	278.2	-1.466	324.7	-1.479	326.9	-1.526	326.7
-1.748	264.5	-1.577	306.7	-1.576	310.4	-1.629	309.3
-1.852	239.5	-1.669	289.0	-1.666	294.3	-1.727	291.5
		-1.770	268.1	-1.760	274.3	-1.825	271.1
		-1.868	244.4	-1.884	245.6		

Electrocapillary data on mercury in aqueous NH_4NO_3 solutions at three temperatures containing 2 mM $Hg_2(NO_3)_2$ and 2 mM HNO_3. Potentials measured with respect to an Hg/Hg_2NO_3 (2mM) electrode. Given is the interfacial tension γ in $mN.m^{-1}$.

Reference: L.H. Hansen and J.W. Williams, J. Phys. Chem., 39, 439 (1935).

Conc. of NH_4NO_3 (mole l^{-1})	1.0	1.0	1.0	0.2	0.2	0.2	0.02	0.02	0.02
Temperature	25°C	50°C	75°C	25°C	50°C	75°C	25°C	50°C	75°C
E volts									
-0.000	312.57	314.78	315.61	306.27	306.99	307.74	310.46	311.32	310.64
-0.025	319.26	321.52	322.20	312.73	313.93	313.98	317.03	317.59	316.97
-0.050	325.41	327.72	328.17	319.18	319.93	320.22	322.67	323.98	322.59
-0.075	331.43	333.21	333.80	325.37	325.92	326.08	329.12	329.39	328.41
-0.100	337.11	338.82	339.11	331.47	331.77	331.86	334.72	334.89	333.80
-0.125	342.47	343.96	344.31	337.34	337.22	337.02	339.73	340.07	339.07
-0.150	347.48	348.99	349.12	342.27	342.52	342.10	344.82	345.13	344.19
-0.175	352.33	353.62	353.89	347.13	347.39	347.18	350.41	349.92	349.15
-0.200	356.95	358.22	358.42	351.94	352.22	351.95	355.30	354.56	353.96
-0.250	365.75	366.52	366.64	361.41	360.99	360.67	363.95	363.29	361.91
-0.300	373.85	374.43	374.17	369.90	369.32	368.47	372.29	371.47	369.79
-0.350	380.94	381.64	381.07	377.37	376.96	375.84	379.88	378.83	376.81
-0.400	387.71	388.10	387.32	384.77	383.89	382.70	386.96	385.88	383.36
-0.450	393.85	393.91	392.94	391.19	390.32	388.48	393.38	392.04	389.68
-0.500	399.45	399.28	397.98	397.45	396.17	394.41	399.33	397.72	394.84
-0.550	404.34	403.88	402.44	403.08	401.43	399.30	404.54	402.44	399.53
-0.600	408.65	407.74	405.93	407.82	405.94	403.60	409.31	406.92	403.64
-0.650	412.48	411.48	409.23	412.48	410.62	407.17	413.65	410.85	406.90
-0.675	414.16	412.68	410.27	414.63	412.22	409.03	415.57	412.41	408.49
-0.700	415.65	414.40	411.40	416.24	413.85	410.58	417.65	413.93	409.77
-0.725	416.82	415.37	412.37	417.80	415.45	411.71	419.17	415.33	411.01
-0.750	418.00	416.35	413.11	419.25	416.81	412.68	420.78	416.50	411.98
-0.775	419.05	417.24	413.53	420.54	417.75	413.49	421.99	417.63	412.91

cont.

Electrocapillary data on mercury in NH_4NO_3. (cont.).

conc. of NH_4NO_3(mole l^{-1})	1.0	1.0	1.0	0.2	0.2	0.2	0.02	0.02	0.02
Temperature	25°C	50°C	75°C	25°C	50°C	75°C	25°C	50°C	75°C
$\frac{E}{\text{volts}}$									
-0.800	420.07	417.71	414.08	421.72	418.53	414.42	422.85	418.45	414.00
-0.825	420.66	417.98	414.39	422.81	419.35	414.93	423.44	419.15	414.58
-0.850	421.09	418.10	414.42	423.59	420.01	415.39	424.30	419.81	415.39
-0.875	421.21	418.10	414.27	424.06	420.36	415.63	424.69	420.36	415.98
-0.900	421.21	418.02	413.80	424.30	420.52	415.78	425.04	420.79	416.40
-0.925	421.05	417.75	413.34	424.49	420.59	415.78	425.39	421.18	416.52
-0.950	420.62	417.36	421.72	424.49	420.52	415.70	425.47	421.37	416.52
-0.975	419.92	416.39	411.90	424.34	420.36	415.55	425.47	421.22	416.44
-1.000	419.09	415.37	410.78	424.02	419.97	415.00	525.36	421.06	416.25
-1.025	418.15	414.52	409.62	423.28	419.11	414.23	425.28	420.79	413.05
-1.050	417.14	413.23	408.52	422.41	418.06	413.45	424.85	420.40	415.70
-1.075	416.16	411.59	407.13	421.17	416.58	412.49	423.91	419.62	415.04
-1.100	414.40	410.31	405.58	419.99	415.65	411.44	423.05	418.84	414.35
-1.125	412.68	408.40	404.22	418.70	414.13	410.35	422.07	418.02	413.53
-1.150	410.91	406.53	402.56	417.18	412.68	408.88	421.17	416.89	412.37
-1.175	408.92	404.66	400.77	415.65	411.17	407.64	419.88	415.61	411.13
-1.200	406.88	402.91	398.91	414.08	409.30	405.85	418.43	414.28	409.93
-1.225	404.65	401.23	396.82	412.32	407.66	404.65	416.94	413.00	408.80
-1.250	402.27	398.62	394.57	410.33	406.06	402.79	415.38	411.67	407.68
-1.275	399.76	396.13	392.32	408.25	404.31	401.01	413.42	410.23	406.47
-1.300	397.37	393.56	389.68	406.26	401.78	399.18	411.58	408.36	404.88
-1.350	391.93	388.88	384.33	401.52	397.69	394.76	408.14	404.39	401.94
-1.400	385.87	382.65	378.63	396.55	392.70	389.95	403.67	400.69	397.44
-1.450	397.45	376.49	372.97	391.58	387.75	385.03	399.17	396.05	392.86
-1.500	372.52	369.71	365.99	385.13	382.22	379.13	393.62	399.99	388.21
-1.550	365.52	362.78	359.20	378.59	375.95	373.20	388.02	385.65	383.48
-1.600	357.81	355.03	351.83	371.86	369.21	367.11	381.68	383.74	377.58

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of formic acid (in mole.l^{-1}).
Potential relative to saturated calomel electrode. Temperature 33.5°C.
Given is the interfacial tension in mN.m^{-1}.
Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{mole \cdot l^{-1}}$	0	0.203	0.500	0.981	1.472	1.970	3.019	3.489	3.890
$\frac{E}{volts}$									
0.0	385.3	384.2	384.0	384.0	383.6	383.6	382.6	382.6	381.3
-0.050	393.8	391.8	391.8	391.8	391.8	391.7	390.4	390.4	389.5
-0.100	402.1	399.2	399.2	399.2	398.1	398.1	395.9	395.8	395.8
-0.150	408.8	405.3	405.4	405.3	404.6	404.0	402.9	402.9	402.3
-0.200	414.3	410.5	410.6	410.3	409.5	408.6	407.6	407.5	407.0
-0.250	418.5	414.9	414.6	414.4	413.1	412.3	411.1	410.7	410.3
-0.300	422.1	418.0	417.9	417.2	415.9	415.1	414.0	413.6	412.7
-0.350	424.5	420.4	420.2	419.2	417.9	416.9	415.8	415.3	414.5
-0.400	426.2	421.8	421.5	420.4	419.0	418.0	417.1	416.2	415.6
-0.450	426.8	422.4	422.1	420.7	419.5	418.4	417.7	416.6	416.1
-0.500	426.7	422.3	422.0	420.6	419.4	418.3	417.7	416.6	416.2
-0.550	426.1	421.7	421.5	420.1	419.0	417.9	417.3	416.2	415.9
-0.600	424.7	420.6	420.4	418.9	417.9	416.8	416.4	415.4	414.6
-0.650	422.6	418.7	418.6	417.0	416.3	415.1	414.8	414.2	413.4
-0.700	420.4	416.3	416.1	414.8	414.0	412.9	412.7	412.0	411.8
-0.750	417.1	413.4	413.3	411.9	410.9	410.1	409.7	409.7	408.9
-0.800	413.6	410.1	409.9	408.9	407.8	407.0	407.0	406.3	406.2
-0.850	409.7	406.8	406.6	405.1	404.2	403.6	403.6	402.7	402.4
-0.900	405.1	402.6	402.6	400.9	400.1	399.6	399.6	398.6	398.6
-0.950	400.6	397.7	397.7	396.5	395.6	395.0	395.1	393.8	393.7
-1.000	395.0	392.4	392.3	391.0	390.8	390.1	390.1	388.9	388.9
-1.050	389.0	387.1	386.8	385.5	385.1	384.7	384.6	383.6	383.6
-1.100	383.4	381.2	381.2	379.4	379.1	379.1	379.0	378.6	378.5

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of acetic acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.

Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{mole\cdot l^{-1}}$	0.533	1.00	1.992	2.456	2.975	3.522	4.048
$\frac{E}{volts}$							
0.0	382.2	381.7	381.2	381.0	379.5	379.2	378.7
-0.050	391.1	389.6	389.5	389.5	387.6	387.6	387.0
-0.100	398.7	396.9	396.6	396.2	395.0	394.6	393.4
-0.150	405.0	403.2	402.1	401.4	400.2	399.9	398.7
-0.200	410.1	408.2	406.6	406.6	404.4	404.0	402.5
-0.250	414.0	411.9	409.9	409.1	407.4	406.3	405.7
-0.300	416.8	414.8	412.4	411.4	409.5	408.2	407.2
-0.350	418.9	416.7	413.5	412.3	410.6	409.4	408.1
-0.400	420.2	417.7	414.2	412.9	411.2	409.9	408.4
-0.450	420.7	418.1	414.1	413.0	411.2	409.8	408.4
-0.500	420.6	417.9	414.0	412.8	411.1	409.6	408.1
-0.550	420.2	417.6	413.5	412.5	410.6	408.9	407.5
-0.600	419.2	416.6	412.7	411.6	409.7	407.9	406.7
-0.650	414.7	415.1	411.4	410.2	408.2	406.4	405.3
-0.700	415.3	413.3	409.6	408.6	406.4	404.3	403.7
-0.750	412.3	409.9	407.0	406.1	404.2	402.2	401.5
-0.800	409.1	406.9	404.0	403.3	402.0	399.8	399.1
-0.850	405.1	404.1	400.9	400.3	399.1	396.8	396.4
-0.900	400.6	400.2	397.2	396.7	395.9	393.9	393.5
-0.950	395.8	395.8	393.4	393.2	392.2	390.3	390.2
-1.000	390.9	390.9	388.9	388.9	388.4	386.0	386.0
-1.050	386.0	385.9	383.0	383.0	383.8	381.6	381.6
-1.100	380.9	379.9	379.4	378.9	378.7	376.5	376.5

Electrocapillary data on mercury in aqueous solutions of 1.992 M acetic acid and varying concentrations of KCl (in mole l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in $mN.m^{-1}$.
Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{salt}}{mole \cdot l^{-1}}$ $\frac{E}{volts}$	0.01	0.05	0.1	0.5
0.0	391.8	386.5	382.9	370.9
-0.100	403.5	400.5	398.1	390.7
-0.200	411.0	409.6	408.3	403.8
-0.300	414.8	414.5	413.9	411.5
-0.350	415.6	415.6	415.2	413.9
-0.400	416.2	416.2	415.9	415.1
-0.450	416.4	416.3	416.1	415.6
-0.500	416.2	416.0	415.9	415.4
-0.550	415.9	415.6	415.4	414.9
-0.600	415.1	414.8	414.6	414.1
-0.650	413.8	412.9	412.8	412.3
-0.700	412.2	411.8	411.2	410.2
-0.800	407.8	406.6	405.7	404.9
-0.900	401.9	400.0	399.0	397.8
-1.000	394.4	392.1	391.0	389.0
-1.100	384.9	381.9	380.4	377.9

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of monochloroacetic acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.
Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{mole \cdot l^{-1}}$ $\frac{E}{volts}$	0.078	0.109	0.139	0.172	0.201	0.227	0.256	0.321
0.0	384.0	383.5	383.4	383.5	383.5	383.4	383.4	382.9
-0.050	392.8	392.0	392.0	392.1	392.1	392.1	392.1	391.7
-0.100	400.5	399.5	399.4	399.4	399.4	399.4	399.2	398.6
-0.150	407.0	405.9	405.5	405.5	404.9	404.9	404.5	404.2
-0.200	411.9	410.9	410.8	410.5	409.8	409.8	409.8	408.7
-0.250	415.6	415.0	414.4	413.8	413.0	413.0	413.0	411.9
-0.300	418.6	418.0	417.3	416.8	416.1	416.0	415.6	414.1
-0.350	420.8	420.1	419.3	418.5	418.1	417.5	417.3	415.7
-0.400	422.2	421.2	420.9	419.7	419.3	418.7	418.3	416.5
-0.450	422.8	421.6	420.9	420.0	419.7	419.2	418.6	416.7
-0.500	422.8	421.7	421.0	419.9	419.7	419.3	418.6	416.7
-0.550	422.4	421.4	420.6	419.6	419.2	419.0	418.1	416.3
-0.600	421.3	420.3	419.6	418.7	418.6	418.0	417.3	415.5
-0.650	419.3	418.2	417.6	417.1	416.8	416.2	415.9	414.2
-0.700	416.9	415.8	415.5	414.8	414.9	414.0	414.0	412.4
-0.750	414.2	412.9	412.9	412.0	411.9	411.1	411.1	410.4
-0.800	411.0	409.8	409.7	409.0	409.2	409.1	408.8	407.5
-0.850	407.4	406.3	406.0	405.6	405.9	405.6	405.4	404.4
-0.900	403.4	402.4	402.2	401.7	402.0	401.7	401.6	400.9
-0.950	398.3	397.8	397.8	397.4	397.5	397.5	397.3	396.9
-1.000	393.0	392.6	392.6	392.6	392.6	392.6	392.6	392.4
-1.050	387.8	387.8	387.6	387.4	387.5	387.5	387.5	
-1.100	381.9	381.1	381.1	381.1				

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of dichloroacetic acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.

Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{mole \cdot l^{-1}}$	0.099	0.120	0.129	0.157	0.177	0.202
$\frac{E}{volts}$						
0.0	383.8	383.8	383.7	382.2	382.2	381.9
-0.050	389.6	389.3	389.3	389.3	389.3	389.2
-0.100	397.7	396.4	396.4	396.4	396.4	396.4
-0.150	404.1	402.8	402.8	402.5	402.3	401.4
-0.200	408.1	407.7	407.7	407.1	406.2	405.1
-0.250	412.5	411.8	411.0	410.2	409.2	408.1
-0.300	415.4	414.9	414.0	412.9	411.8	410.3
-0.350	417.7	417.4	417.0	414.9	413.7	412.2
-0.400	419.6	419.4	418.1	416.1	414.8	413.4
-0.450	420.4	420.1	419.0	416.9	415.4	414.3
-0.500	420.6	420.3	419.4	417.2	415.6	414.7
-0.550	420.4	420.1	419.4	417.2	415.6	414.6
-0.600	419.8	419.5	418.8	416.9	415.1	414.8
-0.650	418.6	418.5	417.6	416.0	414.3	414.0
-0.700	417.1	416.5	415.6	414.3	413.1	412.8
-0.750	414.7	414.3	413.5	411.9	411.4	410.2
-0.800	411.9	411.2	410.5	409.3	409.2	
-0.850	408.1	408.0	407.2	406.0	405.9	
-0.900	403.3	403.3	403.0	401.9		
-0.950						
-1.000						
-1.050						
-1.100						

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of trichloroacetic acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C.

Given is the interfacial tension in mN.m^{-1}.

Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{mole \cdot l^{-1}}$	0.094	0.108	0.122	0.139	0.150	0.172	0.187
$\frac{E}{volts}$							
0.0	381.8	381.8	380.9	380.9	380.9	380.8	380.8
-0.050	389.7	389.7	388.4	388.4	388.1	388.0	388.0
-0.100	397.0	397.0	395.8	395.7	394.6	394.5	394.5
-0.150	401.4	401.5	401.5	401.5	400.4	400.4	400.4
-0.200	408.2	407.9	406.2	406.1	404.8	404.7	404.5
-0.250	411.8	411.9	410.0	409.9	408.4	408.3	407.4
-0.300	415.1	414.9	412.9	412.8	411.2	411.0	410.0
-0.350	417.5	417.2	415.2	415.0	413.5	413.1	412.0
-0.400	419.2	419.0	417.2	416.5	415.3	414.7	413.5
-0.450	420.3	420.4	418.2	417.5	416.5	415.8	414.6
-0.500	420.7	420.4	418.5	418.1	417.1	416.8	415.3
-0.550	420.7	420.4	418.6	418.3	417.2	416.7	415.4
-0.600	419.8	419.7	418.1	417.9	417.0	416.3	415.1
-0.650	417.8	417.3	417.2	417.2	416.1	415.6	414.1
-0.700	415.6	415.6	415.6	415.6	414.5	414.3	412.6
-0.750	413.2	412.9	412.8	412.7	412.4	412.0	410.6
-0.800	410.0	409.7	409.7	409.7	409.7	409.2	407.8
-0.850	406.6	406.1	406.0	406.0	406.0	406.0	404.5
-0.900	402.4	401.6	401.6	401.6	401.6	401.6	400.4
-0.950	397.6	397.3					
-1.000	392.2	392.1					
-1.050	386.5	386.5					
-1.100	380.6						

Electrocapillary data on mercury in aqueous solutions of 1.992 M propionic acid and varying concentrations of KCl (in mole.l^{-1}).
Potential relative to saturated calomel electrode. Temperature 33.5°C.
Given is the interfacial tension in mN.m^{-1}.
Reference: S.W. Dhawale, Thesis Bombay, 1968.

c_{salt} (mole.l^{-1})	0.01	0.05	0.1	0.5	1.0
$\frac{E}{\text{volts}}$					
0.0	391.9	387.1	383.2	371.8	364.1
-0.100	404.2	401.3	398.9	391.5	386.4
-0.200	410.8	410.1	408.8	404.8	401.5
-0.300	413.9	413.4	413.2	411.7	409.7
-0.350	414.5	414.4	414.2	413.4	411.9
-0.400	414.8	414.5	414.4	414.1	412.9
-0.450	414.8	414.5	414.4	414.1	412.9
-0.500	414.4	414.2	413.8	413.6	412.5
-0.550	413.8	413.4	413.1	412.9	411.9
-0.600	413.1	412.3	412.1	411.8	410.9
-0.650	411.6	410.8	410.0	409.8	409.1
-0.700	410.0	409.0	408.5	408.0	407.0
-0.800	406.6	404.5	403.5	403.2	402.6
-0.900	410.0	398.8	397.5	396.5	396.2
-1.000	394.2	391.2	389.2	388.1	387.8
-1.100	385.2	381.4	378.7	377.1	377.0

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of propionic acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.

Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{\text{mole}\cdot\text{l}^{-1}}$	0.105	0.525	1.070	1.479	2.011	3.004	3.542	3.871
$\frac{E}{\text{volts}}$								
0.0	383.7	382.3	380.7	379.5	377.3	377.1	377.1	376.7
-0.050	392.1	390.7	389.1	387.3	385.1	384.8	383.8	383.4
-0.100	399.5	397.8	395.8	393.6	391.3	389.5	388.8	388.1
-0.150	400.3	403.5	400.9	397.7	395.1	392.7	391.8	391.3
-0.200	411.5	408.3	404.2	399.8	397.2	394.6	393.6	393.0
-0.250	415.8	411.1	406.1	401.3	398.4	395.4	394.5	393.8
-0.300	418.6	413.0	407.3	401.5	398.8	395.6	394.7	394.1
-0.350	420.8	414.0	408.2	402.3	399.4	396.3	395.0	394.8
-0.400	422.1	414.3	408.2	402.2	399.2	396.2	394.9	394.5
-0.450	422.3	414.1	407.7	401.9	398.9	395.8	394.6	394.0
-0.500	422.0	413.7	407.2	401.6	398.5	395.1	394.2	393.7
-0.550	421.5	413.0	406.4	400.7	397.2	394.3	393.4	392.8
-0.600	420.5	412.0	405.2	399.6	396.5	393.3	392.2	391.9
-0.650	419.1	410.6	403.9	398.4	395.0	392.1	391.1	391.0
-0.700	416.7	408.6	402.3	397.1	393.7	390.8	389.7	389.4
-0.750	412.6	406.3	400.1	395.2	391.8	388.4	387.7	387.2
-0.800	409.8	403.8	398.0	392.5	390.0	386.5	385.6	385.5
-0.850	405.9	400.8	395.6	391.3	387.8	384.2	383.8	383.6
-0.900	401.6	397.6	393.2	388.6	385.6	382.2	381.6	380.9
-0.950	397.0	393.6	389.5	386.1	383.1	379.3	378.7	378.5
-1.000	392.3	389.6	386.5	383.2	380.1	376.2	376.0	375.8
-1.050	386.6	384.9	382.7	379.6	377.0	373.1	373.1	372.9
-1.100	380.3	379.5	377.9	375.7	373.4	369.3	369.3	369.1

Electrocapillary data on mercury in aqueous solutions of 1.992 M butyric acid and varying concentrations of KCl (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in $mN.m^{-1}$.
Reference: S.W. Dhawale, Thesis Bombay, 1968.

c_{salt} (mole l^{-1}). $\frac{E}{volts}$	0.025	0.1	0.5	1.0	2.00
0.0	386.9	384.6	372.2	-	-
-0.100	400.9	400.1	392.2	387.6	380.8
-0.200	409.0	408.0	405.6	402.3	397.8
-0.300	412.9	412.9	411.7	410.2	407.2
-0.350	413.3	412.9	412.4	411.1	408.5
-0.400	413.3	412.9	412.6	411.1	408.5
-0.450	413.1	412.7	412.4	411.0	408.3
-0.500	412.7	412.2	411.8	410.5	407.9
-0.550	411.9	411.4	410.9	409.9	406.9
-0.600	411.1	410.5	410.1	408.5	405.9
-0.650	409.4	409.1	408.4	407.1	404.3
-0.700	408.1	407.6	406.9	405.5	403.1
-0.800	404.8	403.6	402.7	401.6	399.5
-0.900	399.9	398.8	397.3	396.6	394.7
-1.000	393.0	391.3	389.4	388.7	388.5
-1.100	383.2	381.2	378.4	377.8	377.7

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of butyric acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.

Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{\text{mole}\cdot\text{l}^{-1}}$	0.058	0.078	0.09	0.109	0.181	0.311	0.394	0.500
$\frac{E}{\text{volts}}$								
0.0	383.5	382.5	382.5	382.5	382.2	381.5	381.5	381.2
-0.050	391.8	391.3	391.3	391.2	390.4	389.5	389.5	388.8
-0.100	399.7	398.3	398.3	398.7	397.6	395.8	395.7	394.2
-0.150	405.7	405.1	404.7	404.7	403.2	400.4	399.1	397.2
-0.200	411.0	409.9	409.5	-	404.5	402.7	400.8	398.5
-0.250	414.7	413.6	413.1	413.1	408.3	404.1	401.6	399.3
-0.300	418.0	416.1	414.9	414.3	409.2	404.7	401.9	399.5
-0.350	419.8	417.4	416.1	415.3	409.4	405.0	402.1	399.5
-0.400	420.2	417.9	416.7	415.6	409.4	404.8	402.0	399.1
-0.450	420.3	418.0	416.6	415.6	409.0	404.5	401.5	398.9
-0.500	419.9	417.8	416.0	415.3	408.2	403.8	401.0	398.3
-0.550	419.5	417.2	415.4	414.5	407.3	403.0	400.3	397.5
-0.600	418.5	416.1	414.5	413.3	406.6	401.9	399.4	396.6
-0.650	417.0	414.6	412.8	412.3	405.4	400.5	398.2	395.2
-0.700	415.2	412.5	411.5	410.1	403.5	399.0	396.9	393.9
-0.750	412.7	410.1	408.5	408.2	402.0	397.2	395.2	392.3
-0.800	408.8	407.1	405.9	405.8	400.0	395.3	393.4	390.5
-0.850	406.1	403.6	402.6	402.8	397.6	393.1	391.5	388.7
-0.900	402.1	400.2	399.4	399.4	395.1	390.5	389.3	386.4
-0.950	397.7	396.2	395.1	395.1	391.5	388.0	386.8	383.9
-1.000	392.6	391.3	390.1	390.2	387.9	384.7	384.6	381.1
-1.050	387.3	385.7	384.9	384.9	383.0	380.8	380.8	378.0
-1.100	381.3	380.1	379.0	379.0	377.5	376.3	376.3	374.8

Electrocapillary data on mercury in aqueous solutions of 1.992 M isobutyric acid and varying concentrations of KCl (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.
Reference: S.W. Dhawale, Thesis Bombay, 1968.

c_{salt} (mole.l^{-1})	0.01	0.1	0.5	1.0	2.0
$\frac{E}{\text{volts}}$					
0.0	393.6	384.4	372.6	364.5	353.0
-0.100	405.7	400.0	392.4	387.2	381.0
-0.200	412.0	409.7	405.9	402.5	398.1
-0.300	413.1	412.7	411.7	409.5	407.1
-0.350	413.1	412.7	411.9	410.1	407.5
-0.400	413.3	412.4	411.6	410.0	407.4
-0.450	412.7	412.1	411.2	409.6	406.8
-0.500	412.2	411.2	410.4	409.0	406.0
-0.550	411.3	410.4	409.4	408.0	404.8
-0.600	410.3	409.1	408.2	406.7	403.7
-0.650	408.7	407.7	406.4	405.0	402.4
-0.700	407.2	406.1	404.8	403.6	400.6
-0.800	403.8	402.6	401.3	399.8	397.2
-0.900	399.9	397.9	396.4	395.4	392.9
-1.000	395.8	391.1	389.3	388.8	387.1
-1.100		381.0	378.8	378.7	377.6

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of isobutyric acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.

Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{mole \cdot l^{-1}}$	0.084	0.098	0.154	0.202	0.254	0.312	0.424
$\frac{E}{volts}$							
0.0	383.6	382.9	382.9	382.8	382.4	382.1	381.9
-0.050	391.7	391.6	391.6	391.1	391.1	391.1	391.0
-0.100	400.0	399.0	399.0	398.7	398.7	398.6	397.6
-0.150	406.4	405.4	405.1	404.4	404.4	403.9	401.5
-0.200	411.4	410.8	409.9	408.3	407.3	406.3	403.6
-0.250	415.5	415.0	412.7	410.3	408.4	407.4	404.3
-0.300	418.1	417.1	414.1	411.3	408.8	407.8	404.4
-0.350	419.4	418.2	414.6	411.6	408.6	407.7	404.1
-0.400	419.8	418.3	414.5	411.4	408.2	407.2	403.6
-0.450	419.6	418.2	413.9	410.8	407.5	406.5	402.8
-0.500	419.2	417.8	413.2	410.0	406.7	405.8	402.0
-0.550	418.4	416.9	412.1	409.0	405.8	404.6	400.9
-0.600	417.4	415.7	410.8	408.0	404.9	403.4	399.7
-0.650	416.0	414.1	409.4	406.6	403.7	402.1	398.4
-0.700	414.1	412.2	407.8	405.1	402.4	400.7	397.0
-0.750	411.5	409.9	405.9	403.4	400.9	399.1	395.4
-0.800	408.3	406.9	403.6	401.4	399.1	397.3	393.6
-0.850	404.7	403.9	401.1	399.2	397.0	395.4	391.5
-0.900	400.8	400.6	398.3	396.5	394.8	393.2	389.4
-0.950	396.1	396.1	394.5	393.4	392.2	390.8	386.8
-1.000	391.0	391.1	390.0	390.1	389.2	388.0	384.2
-1.050	385.3	385.3	385.2	385.2	385.2	384.7	381.0
-1.100	380.0	379.9	379.2	379.2	379.1	378.9	377.1

Electrocapillary data on mercury in aqueous solutions of 0.1 M KCl and varying concentrations of n-valeric acid (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.

Reference: S.W. Dhawale, Thesis Bombay, 1968.

$\frac{c_{acid}}{mole\cdot l^{-1}}$	0.014	0.031	0.049	0.072	0.103
$\frac{E}{volts}$					
0.0	383.4	383.4	383.2	383.2	383.1
-0.050	392.1	392.1	392.1	392.1	392.1
-0.100	399.7	399.7	399.2	399.2	398.2
-0.150	406.8	406.8	405.1	403.2	401.0
-0.200	411.4	411.3	407.6	404.6	402.0
-0.250	415.0	413.5	408.9	405.4	402.5
-0.300	417.9	414.4	409.5	405.9	402.7
-0.350	419.7	414.8	409.8	406.0	402.7
-0.400	420.0	414.9	409.7	405.9	402.4
-0.450	420.1	414.7	409.4	405.4	402.1
-0.500	420.0	414.4	408.8	404.9	401.4
-0.550	419.3	413.8	408.1	404.2	400.7
-0.600	418.0	413.0	407.2	403.3	399.9
-0.650	417.3	412.0	406.3	402.3	398.9
-0.700	415.3	410.7	405.1	401.3	397.7
-0.750	412.8	409.1	403.8	400.1	396.7
-0.800	410.0	407.1	402.2	398.6	394.6
-0.850	406.7	404.6	400.3	396.8	392.8
-0.900	403.0	401.5	397.8	394.9	390.6
-0.950	398.5	397.8	395.0	392.4	388.4
-1.000	393.4	393.2	391.6	389.8	386.0
-1.050	387.9	387.9	386.8	386.6	384.4
-1.100	382.1	382.0	381.1	381.0	380.4

Electrocapillary data on mercury in aqueous solutions of 1.992 M n-valeric acid and varying concentrations of KCl (in mole.l^{-1}). Potential relative to saturated calomel electrode. Temperature 33.5°C. Given is the interfacial tension in mN.m^{-1}.
Reference: S.W. Dhawale, Thesis Bombay, 1968.

c_{salt} (mole.l^{-1}) E/volts	0.01	0.1	0.5	1.0	2.0
0.0	394.1	385.0	373.4	365.1	353.7
-0.100	406.0	400.9	392.8	388.1	381.5
-0.200	410.1	409.2	405.7	402.9	398.1
-0.300	411.6	410.6	409.6	408.4	404.4
-0.350	411.8	410.9	409.9	408.8	405.8
-0.400	411.6	410.9	409.9	408.7	405.8
-0.450	411.4	410.6	409.6	408.4	405.4
-0.500	411.3	410.2	409.2	407.6	404.7
-0.550	410.4	409.5	408.4	407.0	403.9
-0.600	409.2	408.6	407.4	406.0	402.9
-0.650	408.4	407.3	406.2	404.7	401.7
-0.700	407.7	405.8	405.0	403.4	400.4
-0.800	405.5	403.2	402.1	400.7	397.7
-0.900	402.8	399.3	397.8	396.6	393.9
-1.000	399.5	392.7	390.6	390.0	388.6
-1.100	386.6	382.5	379.9	379.2	378.1

Electrocapillary data for mercury in aqueous solutions of
x mol l^{-1} CsCl + (1-x) mol l^{-1} LiCl. T = 25°C

Potential with respect to saturated calomel electrode in contact with the working solution.

Reference: R. Parsons and A. Stockton. J. Electroanal. Chem. 25 (1970) App. 10

	γ /mN m^{-1}				
x	1.00	0.50	0.20	0.10	0.00
E/volts					
-0.10	386.6	386.3	387.5	387.5	388.0
-0.15	394.8	394.8	395.8	395.7	396.1
-0.20	402.1	402.3	402.8	403.0	402.3
-0.25	408.4	408.0	409.0	409.0	408.4
-0.30	413.6	412.4	414.2	414.3	413.9
-0.35	417.0	416.9	418.4	418.3	417.9
-0.40	420.5	421.2	421.5	421.5	421.0
-0.45	422.3	423.2	423.4	423.5	423.7
-0.50	422.9	424.4	424.4	424.5	424.3
-0.55	422.7	424.2	424.1	424.5	424.2
-0.60	422.0	422.8	423.1	422.9	423.0
-0.65	421.2	420.8	421.2	421.3	420.9
-0.70	418.1	418.4	418.7	418.8	419.1
-0.75	414.5	415.2	415.4	415.4	415.9
-0.80	411.0	411.4	411.7	411.6	412.1
-0.85	406.7	406.7	407.4	407.4	407.9
-0.90	401.9	402.0	402.6	402.8	403.4
-0.95	396.6	396.8	397.6	397.5	398.0
-1.00	390.8	391.1	391.9	392.0	392.4
-1.05	384.7	385.0	385.9	386.1	386.7
-1.10	377.9	378.4	379.5	379.8	380.4
-1.15	370.4	371.4	372.4	373.0	373.7
-1.20	362.8	363.9	365.2	365.7	366.6
-1.25	354.7	355.8	357.3	358.1	359.1
-1.30	346.4	347.5	349.0	350.2	350.8

Electrocapillary data on mercury in aqueous solutions of 0.1 N $HClO_4$ and various concentrations of ortho-hydroxybenzoic acid. Potentials measured with respect to a normal calomel electrode. Given is the interfacial tension in $mN.m^{-1}$.

Temperature 20.0 ± 0.5°C.

Reference: S. Bordi, G. Papeschi, J.Electroanal.Chem. and Interfacial Electrochem., 20 (1969), 297.

conc. of hydr. benz. acid in $mmole.l^{-1}$. / E volts	0.01	0.05	0.1	0.2	0.4	0.7	1	2	4	7	10
0.1	381.07	381.13	380.66	380.36	379.30	378.13	377.19	374.42	371.54	368.49	366.02
	387.95	388.30	387.77	387.24	385.83	384.30	383.60	380.13	376.95	373.95	371.13
0.0	394.00	394.18	393.59	393.06	391.53	389.77	388.77	385.48	381.83	378.66	375.66
	399.71	399.71	398.88	398.06	396.18	394.53	393.53	389.89	386.36	382.71	379.60
-0.1	404.70	404.47	403.53	402.76	400.41	398.65	397.41	393.89	390.12	386.59	383.66
	409.47	409.00	407.82	406.82	404.47	402.35	401.00	397.24	393.59	389.89	387.24
-0.2	413.35	412.88	411.70	410.47	407.76	405.70	404.12	400.00	396.30	392.65	390.53
	416.58	416.11	414.82	413.23	410.82	408.06	406.53	402.71	398.65	395.06	392.89
-0.3	419.52	418.76	417.23	415.82	413.29	410.47	408.70	404.94	400.82	397.18	394.77
	421.70	420.76	418.99	417.64	414.99	412.41	410.48	406.76	402.41	398.83	396.41
-0.4	423.34	422.17	420.40	419.05	416.23	413.76	411.88	408.17	403.82	400.24	397.65
	424.23	422.93	421.46	419.70	417.05	414.64	412.58	409.11	404.70	401.06	398.47
-0.5	424.52	423.23	421.76	420.11	417.64	415.23	413.17	409.59	405.41	401.71	399.00
	424.17	422.76	421.35	419.99	417.52	415.52	413.47	410.05	405.76	402.06	399.53
-0.6	423.05	421.93	420.58	419.40	417.23	415.29	413.64	409.88	405.88	402.00	399.35
	421.29	420.40	419.23	418.41	416.46	414.58	413.23	409.47	405.53	401.77	399.06
-0.7	419.05	418.46	417.52	416.82	415.11	413.35	412.29	408.64	404.82	401.18	398.59
	416.41	415.94	415.17	414.76	413.23	411.76	410.82	407.29	403.94	400.35	397.88
-0.8	413.29	412.88	412.47	412.11	410.94	409.59	408.76	405.70	402.59	399.30	396.88
	409.70	409.29	409.00	408.76	407.88	406.94	406.23	403.47	400.88	397.94	395.59
-0.9	405.70	405.29	405.00	404.94	404.28	403.71	403.23	400.88	398.71	396.18	394.00
	401.35	401.06	400.77	400.77	400.24	399.88	399.65	397.83	396.12	393.94	392.06
-1.0	396.83	396.41	396.24	396.24	395.94	395.71	395.53	394.24	392.89	391.18	389.59
	392.00	391.65	391.48	391.59	391.36	391.18	391.00	390.06	389.12	387.83	386.59
-1.1	386.95	386.65	386.54	386.60	386.48	386.30	386.30	385.48	384.77	383.77	382.83

Electrocapillary curves for mercury in aqueous KBr + 0.1314 mol l^{-1}
Isobutyric Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}						
c/mol l^{-1}	0.01	0.025	0.05	0.10	0.50	1.00	2.00
E/volts							
-0.10		383.5	380.8	374.6	350.1	333.5	
-0.20		402.6	401.7	397.4	384.4	375.3	363.4
-0.30		413.5	413.1	410.0	402.0	396.1	387.2
-0.35		416.5	415.6	413.5	408.1	402.9	395.5
-0.40		417.1	416.9	415.1	412.0	408.3	401.6
-0.45		417.0	416.9	415.2	413.9	411.3	406.2
-0.50		416.9	416.2	415.0	414.2	412.6	409.2
-0.55		416.3	415.5	414.9	413.9	412.6	410.0
-0.60		415.4	415.2	414.0	412.8	411.7	409.4
-0.65		413.8	413.1	411.7	411.1	410.2	408.3
-0.70		412.4	411.8	409.7	409.4	409.0	407.3
-0.80		409.0	407.7	406.1	404.9	404.5	403.0
-0.90		403.6	402.5	400.1	398.3	398.5	397.4
-1.00		396.7	394.1	391.4	389.3	389.2	388.7
-1.10		386.4	383.3	380.3	376.4	376.6	376.9

Electrocapillary curves for mercury in aqueous KBr + 0.05416 mol l^{-1} n-Valeric Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

γ /mN m^{-1}

c/mol l^{-1}	0.01	0.025	0.05	0.10	0.50	1.00	2.00
E/volts							
-0.10		382.2	382.0	373.7	350.6		
-0.20		401.6	401.1	397.5	384.9	375.9	
-0.30		409.8	409.4	409.0	401.7	395.8	
-0.35		410.8	410.5	409.8	407.1	402.6	
-0.40		410.8	410.4	410.1	409.0	406.7	
-0.45		410.6	410.3	410.0	409.2	407.7	
-0.50		410.5	410.1	409.9	408.8	407.6	
-0.55		410.0	409.8	409.4	408.3	407.2	
-0.60		409.6	409.2	408.7	407.5	406.6	
-0.65		408.3	407.8	406.9	405.4	405.1	
-0.70		407.2	406.8	405.9	404.3	403.9	
-0.80		404.8	403.9	403.0	401.0	401.1	
-0.90		401.1	399.8	398.9	396.6	396.6	
-1.00		395.4	393.1	392.3	388.0	388.9	
-1.10		386.0	383.2	381.3	376.3	377.6	

Electrocapillary curves for mercury in aqueous KBr + 0.4687 mol l^{-1}
Propionic Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}						
c/mol l^{-1}	0.01	0.025	0.05	0.10	0.50	1.00	2.00
E/volts							
-0.10	389.5	374.7		368.0	348.4	331.9	
-0.20	404.6	396.9		393.9	383.8	374.2	
-0.30	413.7	408.8		407.2	401.0	394.2	
-0.35	415.2	412.3		411.1	406.8	400.6	
-0.40	416.1	414.2		413.3	411.1	406.2	
-0.45	416.4	414.6		414.5	413.3	409.0	
-0.50	416.2	414.8		414.4	414.2	411.0	
-0.55	415.7	414.4		414.1	414.1	411.3	
-0.60	415.2	413.9		413.3	413.4	410.8	
-0.65	414.0	412.6		411.2	411.8	410.3	
-0.70	412.4	411.1		409.6	409.8	409.0	
-0.80	407.9	406.9		404.9	405.2	403.1	
-0.90	402.2	402.3		398.8	398.0	396.1	
-1.00	395.4	395.5		390.2	388.7	387.0	
-1.10	384.3	385.3		378.7	377.2	375.4	

Electrocapillary curves for mercury in aqueous KBr + 0.6256 mol l^{-1} Acetic Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}						
c/mol l^{-1}	0.01	0.025	0.05	0.10	0.50	1.00	2.00
E/volts							
-0.10		385.2	380.6	375.6	348.7	331.9	
-0.20		404.1	401.1	398.8	383.8	374.7	
-0.30		415.4	413.4	411.2	401.8	394.7	
-0.35		419.3	417.4	415.8	407.9	401.6	
-0.40		420.9	420.0	418.9	412.9	407.6	
-0.45		422.2	421.4	420.6	415.8	412.0	
-0.50		422.5	421.7	421.1	417.9	414.8	
-0.55		422.2	421.6	420.9	418.7	416.6	
-0.60		421.6	420.7	420.0	418.5	416.9	
-0.65		419.8	418.7	418.1	417.2	415.1	
-0.70		418.2	417.1	416.0	414.9	413.2	
-0.80		412.6	411.5	410.0	408.9	407.0	
-0.90		406.1	403.9	402.0	400.4	398.4	
-1.00		396.7	394.8	392.5	390.5	387.6	
-1.10		386.0	383.8	381.0	378.4	374.7	

Electrocapillary curves for mercury in aqueous 0.1 M KBr + Isobutyric Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}					
c/mol l^{-1}	0.06597	0.08455	0.1106	0.1341	0.1617	0.1959
E/volts						
-0.10	375.6	375.7	375.7	374.4	373.8	375.8
-0.15	389.3	388.7	388.6	388.2	387.6	388.6
-0.20	398.6	398.7	398.1	397.8	397.4	398.3
-0.25	406.0	406.0	405.6	405.1	404.6	404.5
-0.30	412.0	411.4	410.5	410.7	409.8	409.1
-0.35	416.3	415.6	414.8	414.4	412.8	411.6
-0.40	419.2	418.6	416.9	415.8	414.2	412.0
-0.45	420.6	419.2	417.1	416.1	414.2	411.7
-0.50	420.8	419.2	416.8	415.7	413.7	411.1
-0.55	420.3	418.6	416.2	415.2	412.6	410.2
-0.60	419.6	417.8	415.0	414.6	411.7	409.1
-0.65	418.0	416.1	413.9	412.9	410.6	408.0
-0.70	415.6	414.5	411.8	410.9	408.8	406.6
-0.75	413.2	412.4	410.3	409.4	407.3	404.6
-0.80	410.8	409.4	407.7	407.5	405.4	402.7
-0.85	407.2	406.3	405.1	404.7	402.7	400.5
-0.90	403.1	402.7	401.6	401.4	400.6	397.8
-0.95	398.6	398.1	397.4	397.8	396.8	395.0
-1.00	393.7	393.3	392.8	393.4	393.0	391.2
-1.05	388.0	387.6	387.6	388.6	387.9	386.5
-1.10	382.3	382.2	381.8	382.4	382.4	380.8

Electrocapillary curves for mercury in aqueous 0.1 M KBr + n-Valeric Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

γ /mN m^{-1}

c/mol l^{-1}	0.02250	0.03416	0.05416	0.06166	0.07602	0.09038
E/Volts						
-0.10	373.1	373.0	373.5	373.4	373.8	372.3
-0.15	387.8	386.8	387.2	387.1	386.7	385.8
-0.20	397.4	396.0	396.7	396.4	395.9	395.3
-0.25	405.3	403.8	403.4	402.4	401.5	400.3
-0.30	411.1	408.6	408.0	405.8	403.9	402.2
-0.35	415.2	411.3	409.6	406.6	404.5	402.6
-0.40	417.4	412.1	409.9	406.9	404.6	402.7
-0.45	418.5	412.1	409.9	406.9	404.3	402.4
-0.50	418.9	411.9	409.5	406.4	404.2	401.9
-0.55	418.3	411.5	409.2	406.0	403.6	401.3
-0.60	417.1	410.9	408.5	405.4	403.1	400.5
-0.65	415.8	410.4	407.3	404.3	401.0	399.5
-0.70	414.3	409.2	406.3	403.1	400.4	398.9
-0.75	412.1	407.6	404.9	402.0	399.5	397.5
-0.80	409.4	405.7	403.2	400.5	398.0	396.0
-0.85	407.0	403.6	401.4	398.9	396.3	394.1
-0.90	403.4	400.5	398.8	396.7	394.6	392.7
-0.95	399.5	397.5	396.4	394.0	392.4	390.5
-1.00	394.5	393.2	392.5	391.2	389.7	387.9
-1.05	389.6	388.0	388.0	386.7	385.5	385.0
-1.10	383.5	382.3	382.4	382.0	381.1	380.5

Electrocapillary curves for mercury in aqueous 0.1 M KBr +
monochloroacetic Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}				
c/mol l^{-1}	0.08974	0.1231	0.1597	0.2017	0.2462
E/Volts					
-0.10	376.1	375.8	376.0	376.0	374.9
-0.15	389.6	389.2	389.4	388.9	388.0
-0.20	398.8	398.5	398.3	398.0	397.4
-0.25	406.5	405.5	405.8	404.9	404.6
-0.30	412.4	411.4	411.2	410.7	409.7
-0.35	416.8	415.7	415.5	414.8	413.7
-0.40	420.0	418.6	418.1	417.6	416.4
-0.45	421.7	420.6	419.8	418.7	417.6
-0.50	422.5	421.1	420.3	419.2	418.3
-0.55	422.5	421.1	420.2	419.2	418.0
-0.60	421.6	420.4	419.8	418.6	417.2
-0.65	420.4	419.0	418.2	415.4	415.8
-0.70	418.0	416.9	416.4	415.1	414.3
-0.75	415.5	413.8	413.5	412.5	411.8
-0.80	412.1	411.0	410.2	409.5	409.0
-0.85	408.1	407.3	406.5	406.1	405.4
-0.90	403.8	403.3	402.7	402.0	401.7
-0.95	399.5	398.9	398.1	398.0	397.1
-1.00	394.5	393.3	393.7	393.2	392.7
-1.05	388.9	388.2	388.2	388.1	
-1.10	383.0	381.9	382.0		

Electrocapillary curves for mercury in aqueous 0.1 M KBr +
dichloroacetic Acid at $33.5^{\circ}C$

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}					
c/mol l^{-1}	0.07108	0.09504	0.1230	0.1691	0.1986	0.2239
E/Volts						
-0.10	375.3	375.5	375.4	374.8	374.3	374.1
-0.15	388.9	388.9	388.2	388.1	387.5	387.1
-0.20	398.2	398.6	397.6	397.1	396.6	396.0
-0.25	406.0	405.7	404.8	404.1	403.4	402.5
-0.30	411.5	411.4	410.5	409.0	408.0	407.3
-0.35	416.0	415.6	414.0	412.6	410.8	409.9
-0.40	419.1	418.1	417.0	414.9	412.8	411.7
-0.45	420.8	419.7	418.4	415.8	414.1	412.7
-0.50	421.7	420.5	418.8	416.4	414.6	413.5
-0.55	421.7	420.5	418.8	416.4	414.6	413.5
-0.60	420.9	419.8	418.4	416.2	414.6	413.5
-0.65	419.9	418.8	417.5	415.6	414.1	413.0
-0.70	417.8	416.8	415.8	414.4	413.3	412.1
-0.75	418.9	414.2	413.5	411.8	411.4	410.6
-0.80	412.0	411.1	410.3	409.4	409.0	408.6
-0.85	407.7	407.7	406.8	406.3	405.8	405.6
-0.90	404.0	403.3	402.8	402.6		
-0.95	399.5	399.1	398.3	398.1		
-1.00	394.6	393.9	393.5	393.4		
-1.05	388.1	388.2	388.0	388.2		
-1.10	382.3	381.5	381.6	382.1		

Electrocapillary curves for mercury in aqueous 0.1 M KBr +
trichloroacetic Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}				
c/mol l^{-1}	0.05198	0.07178	0.09324	0.1485	0.1839
E/Volts					
-0.10	373.5	374.9	374.3	372.8	372.7
-0.15	387.5	388.3	387.4	386.3	386.3
-0.20	397.3	397.9	397.1	395.9	395.8
-0.25	405.0	405.4	404.5	403.2	402.6
-0.30	411.4	411.2	410.2	408.7	407.9
-0.35	416.0	415.6	414.3	417.5	411.1
-0.40	419.1	418.7	417.1	414.7	413.5
-0.45	421.5	420.1	419.1	416.1	414.6
-0.50	422.4	420.7	419.5	416.9	415.3
-0.55	422.2	420.8	419.7	417.1	415.6
-0.60	421.9	420.2	419.2	416.8	415.4
-0.65	420.0	418.9	418.0	416.0	415.0
-0.70	417.9	416.9	416.1	416.3	414.0
-0.75	415.6	414.4	413.7	413.6	412.6
-0.80	412.1	410.6	410.1	410.8	410.4
-0.85	408.0	407.2	406.4	407.6	407.2
-0.90	403.7	403.2	402.3	403.5	403.4
-0.95	399.3	398.3	397.9	399.4	399.4
-1.00	394.2	393.1	392.7	394.6	395.0
-1.05	388.7	387.6	387.1		
-1.10	382.3	380.0	381.2		

Electrocapillary curves for mercury in aqueous 0.1 M KBr + Butyric Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

γ /mN m^{-1}

c/mol l^{-1}	0.07099	0.09363	0.1214	0.1646	0.1965
E/Volts					
-0.10	374.7	374.7	374.6	375.0	375.0
-0.15	388.3	388.4	388.3	388.4	387.8
-0.20	397.6	398.1	398.2	398.0	397.0
-0.25	405.4	405.5	405.5	404.8	404.0
-0.30	411.1	410.9	410.8	409.5	408.1
-0.35	415.0	414.7	413.8	411.7	410.7
-0.40	417.7	417.0	416.0	413.6	411.2
-0.45	418.7	417.8	416.3	413.6	411.2
-0.50	419.1	417.8	416.1	413.4	410.8
-0.55	418.8	417.2	415.6	412.5	410.3
-0.60	418.3	416.1	414.8	411.7	409.6
-0.65	416.8	414.8	413.3	410.4	408.4
-0.70	414.7	413.5	412.0	409.0	407.0
-0.75	412.4	411.6	410.1	407.3	405.3
-0.80	409.7	408.8	407.5	405.0	403.0
-0.85	406.6	406.0	405.1	402.8	400.8
-0.90	402.7	402.4	401.7	399.8	398.6
-0.95	398.5	397.7	397.7	396.4	395.2
-1.00	393.5	393.8	393.5	392.1	391.5
-1.05	388.2	388.4	388.1	387.6	386.7
-1.10	382.6	382.9	382.6	381.8	381.8

Electrocapillary curves for mercury in aqueous 0.1 M KBr + Propionic Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}				
c/mol l^{-1}	0.2282	0.2961	0.4687	0.7283	0.8472
E/Volts					
-0.10	374.8	374.9	374.7	373.8	373.7
-0.15	388.3	388.5	387.9	386.7	386.4
-0.20	398.0	397.9	397.1	395.8	395.2
-0.25	405.6	405.0	403.8	402.5	401.6
-0.30	411.2	410.3	409.0	406.8	405.5
-0.35	415.1	414.3	412.3	409.7	407.3
-0.40	418.0	417.0	414.1	410.5	408.2
-0.45	419.4	417.9	414.6	410.6	408.3
-0.50	419.7	418.0	414.7	410.6	408.1
-0.55	419.5	417.7	414.2	410.0	407.5
-0.60	418.6	417.0	413.3	409.6	406.8
-0.65	417.4	415.2	411.7	408.5	405.7
-0.70	415.5	413.4	410.3	406.5	404.2
-0.75	413.0	411.2	408.3	404.6	402.2
-0.80	410.1	408.2	405.4	402.8	400.0
-0.85	406.5	405.5	402.9	399.9	397.3
-0.90	402.9	401.6	399.3	397.3	394.8
-0.95	398.6	397.2	396.0	393.8	392.1
-1.00	393.4	392.8	391.4	390.2	388.4
-1.05	388.5	387.6	386.4	385.5	384.3
-1.10	382.9	381.5	380.9	380.4	379.3

Electrocapillary curves for mercury in aqueous 0.1 M KBr + Acetic Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

	γ /mN m^{-1}						
c/mol l^{-1}	0.0669	0.1379	0.2099	0.3118	0.4528	0.6256	0.9014
E/Volts							
-0.10	376.0	376.3	376.4	376.2	376.5	376.1	375.5
-0.15	389.7	389.7	389.8	389.3	389.2	389.1	388.6
-0.20	399.6	399.4	399.6	398.6	398.7	398.8	398.1
-0.25	407.1	407.0	406.8	406.0	406.3	405.9	405.5
-0.30	413.1	413.4	413.0	412.0	411.9	411.8	410.7
-0.35	417.9	417.7	417.5	416.5	416.2	415.9	414.7
-0.40	421.4	421.0	420.6	419.9	419.5	418.8	417.5
-0.45	423.6	423.1	422.6	421.4	421.3	420.5	418.9
-0.50	424.6	424.0	423.4	422.4	421.9	421.2	419.6
-0.55	424.5	424.0	423.4	422.4	421.9	421.0	419.4
-0.60	423.7	423.2	422.5	421.5	421.0	420.4	418.6
-0.65	422.2	421.8	421.1	419.8	419.6	418.4	417.4
-0.70	419.9	418.9	418.8	413.7	417.3	416.8	415.3
-0.75	416.9	416.2	415.8	415.1	414.1	413.3	412.6
-0.80	413.2	412.5	412.0	411.6	411.1	410.7	411.6
-0.85	409.5	408.7	408.5	407.9	407.4	407.1	406.1
-0.90	405.1	404.2	403.7	403.6	403.4	403.0	402.1
-0.95	400.4	399.8	399.8	399.1	398.7	398.5	397.8
-1.00	395.2	394.6	394.6	394.4	393.7	393.7	392.9
-1.05	389.6	389.0	389.0	388.7	388.5	388.1	387.6
-1.10	383.6	382.6	382.7	382.4	382.3	381.9	381.9

Electrocapillary curves for mercury in aqueous 0.1 M KBr + Formic Acid at 33.5°C

Potentials measured with respect to a saturated calomel electrode.

Reference: M.R. Bapat, Ph.D. Thesis, Bombay, 1968.

γ /mN m^{-1}

c/mol l^{-1}	0	0.3727	0.7305	1.069	1.647	2.379	3.079
E/Volts							
-0.10	377.0	376.9	376.7	376.9	375.9	376.7	376.7
-0.15	390.1	390.2	389.3	389.5	389.1	388.6	389.2
-0.20	399.9	399.5	398.9	398.9	398.2	398.2	398.0
-0.25	407.9	407.5	406.4	406.3	405.6	405.2	404.9
-0.30	413.9	413.0	412.6	412.1	411.2	410.5	410.2
-0.35	418.4	417.9	417.0	416.3	415.7	414.9	414.1
-0.40	422.1	421.4	420.1	419.6	418.6	417.5	417.1
-0.45	424.3	423.4	422.2	421.3	420.6	419.2	418.5
-0.50	425.2	424.2	423.2	422.6	421.3	420.0	419.4
-0.55	425.0	424.2	423.2	422.6	421.3	420.0	419.2
-0.60	424.3	423.5	422.1	421.9	420.5	419.1	418.7
-0.65	422.2	421.5	420.8	420.1	419.2	417.5	417.3
-0.70	420.4	419.0	418.5	418.0	416.8	415.8	415.0
-0.75	416.7	416.2	415.0	415.0	414.3	413.0	412.7
-0.80	413.2	413.1	411.7	411.7	410.4	409.6	409.8
-0.85	409.0	409.2	407.8	407.8	406.7	406.3	406.1
-0.90	405.1	404.7	403.7	403.1	403.1	402.5	401.9
-0.95	400.5	400.1	398.8	399.0	398.6	397.6	397.8
-1.00	394.4	394.5	393.9	393.7	393.2	392.8	392.7
-1.05	389.3	389.4	388.4	388.4	387.9	387.7	387.3
-1.10	383.1	382.8	382.2	382.4	382.1	381.8	381.6

Electrocapillary data for n-pentanoic acid in 0.1 M $HClO_4$ solution.
Reference electrode: SCE (NaCl solution). T = 25°C
Reduced concentration c/c_o where c is solute concentration and c_o its saturation concentration in the base electrolyte.
γ: Interfacial tension in mN m^{-1}.

Reference: K.G. Baikerikar and R.H. Hansen. J. Colloid Interface Sci. 52 (1975) 277.

c/c_o	0.0	0.01234	0.02439	0.03614	0.04761
E/V					
-1.10	384.7	384.5	384.4	384.4	384.1
-1.05	390.5	390.2	390.0	390.0	389.7
-1.00	395.8	395.5	395.3	395.1	394.7
-0.95	400.8	400.4	400.1	399.8	399.4
-0.90	405.3	404.8	404.4	404.0	403.4
-0.85	409.6	408.9	408.3	407.8	407.1
-0.80	413.2	412.5	411.7	411.0	410.0
-0.75	416.6	415.6	414.7	413.8	412.6
-0.70	419.4	418.3	417.2	416.0	414.7
-0.65	421.8	420.5	419.3	417.9	416.4
-0.60	423.6	422.3	420.8	419.3	417.7
-0.55	424.8	423.3	422.0	420.4	418.8
-0.50	425.6	424.0	422.5	421.1	419.3
-0.45	425.5	424.1	422.7	421.3	419.7
-0.40	424.8	423.5	422.3	421.0	419.5
-0.35	423.3	422.3	421.3	420.2	418.9
-0.30	421.3	420.5	415.6	418.8	417.7
-0.25	418.6	417.9	417.3	416.7	415.9
-0.20	415.3	414.8	414.2	413.9	413.2
-0.15	411.3	411.0	410.6	410.3	409.9
-0.10	406.7	406.5	406.1	406.0	405.7
-0.05	401.6	401.4	401.1	401.1	400.8
0.00	395.8	395.6	395.4	395.4	395.3
0.05	389.5	389.4	389.3	389.3	389.1
0.10	382.6	382.5	382.3	382.3	382.2
0.15	375.0	374.9	374.9	374.9	374.8
0.20	366.7	366.7	366.7	366.7	366.6

Electrocapillary data for n-pentanoic acid in 0.1 M $HClO_4$ solution (continued)

c/c_o	0.06976	0.09090	0.1304	0.1667	0.3333	0.5000
E/V						
-1.10	384.0	383.6	383.2	381.9	378.8	375.5
-1.05	389.9	388.9	388.2	387.2	382.5	378.6
-1.00	394.3	393.4	392.4	390.7	385.3	380.9
-0.95	398.6	397.7	396.0	394.3	387.7	382.9
-0.90	402.4	401.0	398.9	396.7	389.5	384.8
-0.85	405.6	404.0	401.3	399.1	391.3	386.4
-0.80	408.3	406.3	403.2	400.7	392.9	387.8
-0.75	410.5	408.3	404.9	402.5	394.3	389.2
-0.70	412.2	409.7	406.3	403.6	395.5	390.4
-0.65	413.7	411.2	407.5	404.8	396.5	391.4
-0.60	414.8	412.3	408.3	405.8	397.6	392.5
-0.55	415.6	413.1	409.3	406.7	398.3	393.2
-0.50	416.3	413.8	409.9	407.4	399.1	394.0
-0.45	416.7	414.2	410.7	408.0	399.7	394.5
-0.40	416.8	414.4	410.8	408.3	400.1	395.0
-0.35	416.5	414.3	410.9	408.5	400.4	395.2
-0.30	415.7	413.7	410.7	408.3	400.5	395.4
-0.25	414.3	412.6	410.0	407.9	400.4	395.4
-0.20	412.1	410.7	408.7	406.8	399.9	395.0
-0.15	409.1	408.1	406.7	405.0	398.9	394.3
-0.10	405.2	404.8	403.6	402.3	397.2	393.1
-0.05	400.5	400.1	399.3	398.7	394.6	391.0
0.00	395.0	394.7	394.3	394.0	391.0	388.0
0.05	388.9	388.8	388.4	388.3	386.4	384.1
0.10	382.1	382.0	381.5	381.8	380.5	379.0
0.15	374.7	374.6	374.3	374.5	373.6	372.6
0.20	366.6	366.5	366.1	366.6	365.9	365.2

Electrocapillary data on mercury electrode in 0.2 M $NaClO_4$ containing the following molar concentration of benzoic acid. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE (NaCl). T = 25°C

Reference: J. Dojlido, M. Dmowska-Stanczak and Z. Galus.
J. Electroanal. Chem. 94 (1978) 107.

	γ /mN m^{-1}					
c/mol l^{-1}	0	0.001	0.002	0.003	0.005	0.006
-E/mV						
15	389.3	385.8	384.3	382.6	381.7	380.7
100	400.3	395.2	394.0	392.4	391.0	389.5
200	410.0	405.0	402.2	400.7	398.3	397.5
300	417.2	412.1	408.6	407.2	404.3	403.3
400	422.5	416.6	412.9	411.4	408.4	407.7
500	424.9	419.2	415.3	413.7	410.6	409.5
550	424.9	419.8	416.4	414.3	411.2	410.0
600	424.4	419.3	416.5	414.4	411.0	410.0
700		417.2	414.7	413.0	410.4	409.2
800	415.4	413.3	411.4	410.2	407.8	406.7
900		406.7	405.7	405.2	403.6	402.5
1000	400.0	398.2	398.2	398.1	397.3	396.2
1100		387.5	387.5	387.4	387.5	387.4
1200	375.8	375.9	375.8	375.9	375.7	375.5
1300		361.0	361.0	360.9	361.1	361.0
1400	346.2	346.2	346.1	346.2	346.0	346.2
1500	329.0	328.9	328.9	328.9	328.8	328.7

Electrocapillary data for mercury in solutions of LiH_2PO_4 in formic acid.

T = 25°C

Potential with respect to saturated calomel electrode in formic acid in contact with 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	γ /mN m^{-1}					
c/mol l^{-1} E/volts	0.020	0.050	0.10	0.20	0.50	1.00
0.00	381.5		380.4	380.0	380.3	379.8
-0.05	385.2		384.2	384.1	384.1	384.0
-0.10	388.2	387.7	387.6	387.5	387.3	387.7
-0.15	390.7	390.6	390.4	390.4	390.6	390.6
-0.20	393.0	392.9	392.8	393.0	392.9	393.1
-0.25	394.8	394.8	394.7	394.8	395.0	395.2
-0.30	396.1	396.1	396.2	396.4	396.6	396.8
-0.35	397.3	397.4	397.4	397.7	397.9	398.0
-0.40	398.2	398.1	398.2	398.5	398.8	398.9
-0.45	398.7	398.7	398.9	399.0	399.2	399.4
-0.50	398.9	399.0	399.2	399.2	399.4	399.6
-0.55	398.9	399.0	399.1	399.2	399.4	399.6
-0.60	398.7	398.7	398.9	398.8	399.0	399.2
-0.65	398.3	398.3	398.3	398.4	398.5	398.7
-0.70	397.8	397.7	397.6	396.5	397.5	397.8
-0.75	397.0	396.8	396.5	396.5	396.3	396.5
-0.80	396.0	396.7	395.3	395.1	394.9	395.0
-0.85	395.0	394.2	393.8	393.6	393.2	393.4
-0.90	393.5	392.6	392.1	391.8	391.4	391.2

Electrocapillary data for mercury in solutions of NaH_2PO_4 in formic acid.
T = 25°C

Potential with respect to saturated calomel electrode in formic acid in contact with 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	γ /mNm^{-1}					
c/mol l^{-1} E/volts	0.020	0.050	0.10	0.20	0.50	1.00
0.00	380.3	379.6	379.8	379.0	378.8	379.5
-0.05	384.5	383.6	384.1	383.5	383.0	383.8
-0.10	387.5	387.1	387.3	387.0	387.0	387.6
-0.15	390.1	390.0	390.6	390.3	390.3	390.6
-0.20	392.4	392.4	392.9	392.8	392.8	393.4
-0.25	394.2	394.5	395.0	394.9	395.2	395.5
-0.30	395.6	396.0	396.6	396.4	396.8	397.1
-0.35	397.0	397.3	397.8	398.0	398.1	398.4
-0.40	398.0	398.2	398.6	398.8	399.0	399.2
-0.45	398.5	398.8	399.3	399.2	399.5	399.7
-0.50	398.9	399.2	399.5	399.5	399.8	399.8
-0.55	398.9	399.2	399.5	399.4	399.8	399.8
-0.60	398.7	398.9	399.2	399.1	399.5	399.4
-0.65	398.4	398.5	398.7	398.4	398.9	398.5
-0.70	397.8	397.7	398.0	397.7	397.8	397.5
-0.75	397.0	396.9	396.8	396.7	396.7	396.3
-0.80	396.1	395.7	395.6	395.2	395.6	394.9
-0.85	394.8	394.3	394.2	393.7	393.8	393.3
-0.90	393.3	392.7	392.4	392.0	392.0	391.2

Electrocapillary data for mercury in solutions of KH_2PO_4 in formic acid.
T = 25°C

Potentials with respect to a saturated calomel electrode in formic acid in contact with a 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	γ /mN m^{-1}					
c/mol l^{-1}	0.020	0.050	0.10	0.20	0.50	1.00
E/volts						
0.00	381.6	380.7	380.6	379.5	379.2	379.1
-0.05	384.9	384.5	384.3	383.5	383.0	383.5
-0.10	388.0	387.8	387.8	387.0	387.1	387.4
-0.15	390.6	390.6	390.8	390.2	390.4	390.7
-0.20	392.8	392.8	393.1	392.6	392.9	393.3
-0.25	394.5	394.8	395.2	394.9	395.1	395.4
-0.30	395.9	396.3	396.6	396.7	396.7	397.0
-0.35	396.8	397.6	397.8	397.8	398.1	398.4
-0.40	397.7	398.4	398.6	398.6	398.8	399.1
-0.45	398.2	398.7	399.1	399.1	399.5	399.8
-0.50	398.4	399.0	399.5	399.2	399.7	399.9
-0.55	398.5	399.0	399.5	399.2	399.5	399.8
-0.60	398.4	398.8	399.3	398.9	399.1	399.6
-0.65	398.2	398.3	398.8	398.3	398.6	398.8
-0.70	397.8	397.7	397.8	397.6	397.6	397.9
-0.75	396.8	396.7	396.7	396.3	396.7	396.6
-0.80	395.9	395.5	395.6	395.2	395.0	395.2
-0.85	394.6	394.3	394.1	393.6	393.4	393.6
-0.90	393.1	392.5	392.3	391.8	391.4	391.5

Electrocapillary data for mercury in solutions of CsH_2PO_4 in formic acid.
T = 25°C

Potentials with respect to a saturated calomel electrode in formic acid in contact with a 0.2 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	γ/mN m^{-1}					
c/mol l^{-1} E/volts	0.020	0.050	0.10	0.20	0.50	1.00
0.00	382.0	380.6	380.7	380.6	379.8	380.3
-0.05	385.4	384.2	384.3	384.4	384.1	384.8
-0.10	388.4	387.7	387.9	387.9	387.5	388.1
-0.15	391.1	390.5	390.8	390.9	390.8	391.3
-0.20	393.3	392.9	393.2	393.3	393.5	394.1
-0.25	394.9	394.9	395.0	395.3	395.5	395.9
-0.30	396.3	306.6	396.7	397.0	397.1	397.7
-0.35	397.2	397.7	397.8	398.2	398.3	398.9
-0.40	398.1	398.7	398.7	398.9	399.1	399.8
-0.45	398.6	399.1	399.2	399.5	399.6	400.4
-0.50	398.9	399.4	399.6	399.7	399.8	400.6
-0.55	398.9	399.4	399.6	399.7	399.8	400.6
-0.60	398.8	399.0	399.2	399.3	399.5	400.1
-0.65	398.5	398.6	398.7	398.7	398.8	399.2
-0.70	397.9	398.0	397.9	397.8	397.8	398.1
-0.75	396.9	396.9	396.8	396.7	396.5	396.8
-0.80	395.8	395.6	395.5	395.4	395.2	395.1
-0.85	394.7	394.2	393.9	393.7	393.3	393.0
-0.90	393.3	392.5	392.0	391.9	391.1	390.8

Electrocapillary data on mercury electrode in 1 M KNO_3 + 0.00545 M $Hg_2(NO_3)_2$ + 0.0064 M HNO_3 at different temperatures. Given are the interfacial tension in mN m^{-1}, potential in V.

Reference: F.O. Koenig. Z. phys. Chem. 157 (1931) 96.

T/°C	9.3		9.3	
	E^P/V	γ/mN m^{-1}	E^P/V	γ/mN m^{-1}
	−0.012	267.2	+0.957	421.9
	0.000	277.9	0.972	422.0
	+0.012	284.9	0.984	422.1
	0.074	305.1	0.997	422.1
	0.128	319.7	1.011	422.0
	0.191	334.8	1.026	421.8
	0.245	346.5	1.037	421.7
	0.298	356.9	1.050	421.4
	0.352	366.4	1.065	421.1
	0.406	375.0	1.079	420.6
	0.459	382.9	1.091	420.2
	0.513	389.9	1.103	419.8
	0.566	396.4	1.118	419.2
	0.620	402.1	1.145	418.1
	0.674	407.2	1.172	416.5
	0.727	411.6	1.219	414.0
	0.781	415.3	1.265	410.1
	0.835	418.3	1.318	405.4
	0.859	419.4	1.372	400.3
	0.874	419.9	1.426	394.6
	0.889	420.4	1.479	388.4
	0.904	420.8	1.533	381.6
	0.918	421.3	1.587	374.4
	0.930	421.5	1.640	366.7
	0.943	421.7		

γ /mN m^{-1} (max) = 422.1

$E^{P\,max}$ / V = 0.987 ± 0.001

Electrocapillary data on mercury electrode in 1 M KNO_3 + 0.00545 M $Hg_2(NO_3)_2$ + 0.0064 M HNO_3 at different temperatures (continued)

T/°C	25		25	
	E^P/V	γ /mN m^{-1}	E^P/V	γ /mN m^{-1}
	-0.010	272.6	+0.955	420.4
	0.000	280.6	0.970	420.5
	+0.010	285.6	0.981	420.5
	0.064	302.6	0.994	420.5
	0.128	320.0	1.009	420.4
	0.191	335.0	1.023	420.1
	0.244	346.4	1.034	420.0
	0.298	356.7	1.048	419.7
	0.352	366.2	1.062	419.4
	0.405	374.7	1.077	418.9
	0.459	382.5	1.089	418.5
	0.512	389.5	1.102	418.0
	0.565	395.8	1.116	417.4
	0.618	401.5	1.142	416.2
	0.672	406.5	1.169	414.7
	0.726	410.8	1.208	412.4
	0.779	414.5	1.262	408.6
	0.818	416.6	1.315	404.1
	0.857	418.2	1.369	399.1
	0.872	418.8	1.423	393.6
	0.887	419.2	1.476	387.5
	0.902	419.5	1.530	381.1
	0.916	419.8	1.583	374.0
	0.927	420.1	1.637	366.5
	0.942	420.3		

γ /mN m^{-1} (max) = 420.5

E^P (max)/ V = 0.977 ± 0.001

Electrocapillary data on mercury electrode in 1 M KNO_3 + 0.00545 M $Hg_2(NO_3)_2$ + 0.0064 M HNO_3 at different temperatures (continued)

T/°C	40		40	
	E^P/V	γ/mN m^{-1}	E^P/V	γ /mN m^{-1}
	-0.010	275.0	+0.927	418.4
	0.000	281.8	0.939	418.5
	+0.010	286.0	0.954	418.6
	0.064	302.6	0.969	418.7
	0.117	317.1	0.981	418.6
	0.179	332.0	0.993	418.5
	0.244	345.9	1.008	418.4
	0.297	356.3	1.022	418.2
	0.351	365.6	1.035	418.0
	0.404	374.1	1.046	417.7
	0.458	381.8	1.061	417.4
	0.511	388.9	1.076	416.9
	0.565	395.1	1.088	416.5
	0.618	400.6	1.100	416.1
	0.672	405.5	1.129	414.8
	0.726	409.7	1.154	413.7
	0.779	413.2	1.207	410.6
	0.803	414.4	1.261	406.8
	0.817	415.1	1.314	402.5
	0.833	415.8	1.368	397.6
	0.847	416.4	1.421	392.3
	0.862	416.9	1.475	386.4
	0.874	417.2	1.528	380.1
	0.886	417.6	1.582	373.3
	0.901	417.9	1.635	365.9
	0.915	418.2		

γ/mN m$^{-1(\max)}$ = 418.7

$E^{P\,(\max)}$/V = 0.969 ± 0.001

Electrocapillary data on mercury electrode in 1 M KNO_3 + 0.00545 M $Hg_2(NO_3)_2$ + 0.0064 M HNO_3 at different temperatures (continued)

T/°C	55		55	
	E^P/V	γ/mN m^{-1}	E^P/V	γ/mN m^{-1}
	-0.010	276.6	+0.926	416.4
	0.000	281.9	0.938	416.4
	+0.010	285.8	0.952	416.6
	0.063	302.2	0.967	416.6
	0.117	316.5	0.979	416.5
	0.170	329.4	0.991	416.4
	0.231	342.7	1.006	416.2
	0.297	355.3	1.021	416.0
	0.350	364.6	1.033	415.8
	0.403	373.1	1.045	415.5
	0.457	380.9	1.059	415.2
	0.510	387.6	1.074	414.7
	0.563	393.9	1.086	414.3
	0.617	399.5	1.098	413.9
	0.670	404.3	1.127	412.7
	0.724	408.2	1.151	411.5
	0.762	410.7	1.207	408.6
	0.807	413.0	1.258	405.0
	0.819	413.6	1.311	400.8
	0.831	414.0	1.363	396.1
	0.846	414.6	1.417	390.9
	0.860	415.0	1.470	385.2
	0.873	415.4	1.523	379.0
	0.884	415.7	1.577	372.4
	0.899	416.0	1.630	365.3
	0.914	416.2		

γ/mN m^{-1} (max) = 416.6

$E^{P\,(max)}$/V = 0.962 ± 0.001

Electrocapillary data for mercury in solutions of Na_2SO_4 in formic acid.
T = 25°C

Potentials with respect to a saturated calomel electrode in formic acid in contact with a 0.1 M solution of the salt.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	γ/mN m^{-1}					
c/mol l^{-1}	0.020	0.050	0.10	0.20	0.50	1.00
E/volts						
0.00	382.9	381.9	381.5	381.3	380.6	380.5
-0.05	386.3	385.5	385.2	385.2	384.5	384.4
-0.10	389.3	388.7	388.6	388.5	388.0	388.2
-0.15	391.9	391.5	391.4	391.5	391.2	391.4
-0.20	394.1	393.9	393.6	393.9	393.7	394.1
-0.25	395.8	395.5	395.5	395.9	395.7	396.1
-0.30	397.0	396.9	397.0	397.5	397.5	397.8
-0.35	398.2	397.8	397.9	398.6	398.6	399.1
-0.40	398.9	398.6	398.8	399.4	399.5	399.9
-0.45	399.1	399.0	399.2	399.8	400.0	400.4
-0.50	399.2	399.2	399.4	400.0	400.3	400.5
-0.55	399.2	399.2	399.3	399.9	400.3	400.6
-0.60	399.1	398.9	399.0	399.6	399.9	400.3
-0.65	398.7	398.3	398.5	398.9	399.2	399.7
-0.70	398.0	397.6	397.6	398.2	398.3	398.7
-0.75	396.9	396.6	396.5	397.0	397.2	397.5
-0.80	395.8	395.3	395.3	395.7	395.9	396.1
-0.85	394.4	393.8	393.6	393.9	394.1	394.6
-0.90	392.8	392.1	391.8	392.2	392.2	392.0

Electrocapillary data for mercury in solutions of 0.1 M HCOONa in formic acid + water mixtures. T = 25°C

Potential with respect to a saturated calomel electrode in formic acid.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

				γ /mN m^{-1}			
Vol % HCOOH	1	5	10	30	50	70	90
E/volts							
0.05					383.8		
0.00	397.4	396.1	395.9	392.0	388.6	385.9	382.9
-0.05	402.7	401.4	401.0	397.8	392.5	389.6	386.1
-0.10	407.9	406.0	405.4	401.4	396.1	393.1	389.3
-0.15	412.0	410.3	409.6	404.4	398.9	396.0	391.9
-0.20	416.0	413.9	413.0	406.9	401.4	398.3	394.3
-0.25	419.0	416.8	415.6	408.9	403.5	400.2	396.0
-0.30	421.4	419.3	417.6	410.6	405.2	401.8	397.4
-0.35	423.2	420.7	419.1	411.7	406.4	402.8	398.7
-0.40	424.5	421.8	420.0	412.4	407.3	403.8	399.5
-0.45	424.8	422.2	420.1	412.6	407.8	404.3	400.2
-0.50	424.7	422.1	419.9	412.5	407.9	404.7	400.4
-0.55	423.6	421.3	419.0	412.1	407.8	404.3	400.2
-0.60	421.9	419.9	417.7	411.2	407.0	403.9	399.8
-0.65	419.9	418.1	416.3	410.0	406.2	403.1	399.2
-0.70	417.4	415.7	414.1	408.4	404.7	402.0	398.4
-0.75	414.5	412.8	411.4	406.2	403.1	400.6	397.4
-0.80	411.1	409.7	408.4	403.7	400.9	398.8	395.9
-0.85	407.2	405.8	404.8	400.7	398.4	396.7	394.3
-0.90	402.8	401.2	400.7	397.2	395.5	394.1	392.2
-0.95					392.0		
-1.00					388.2		
-1.05					384.0		

Electrocapillary data on mercury for 0.1 m HCl solutions in methanol-water mixtures. T = 25°.

Potential measured with respect to a hydrogen electrode in the working solution.

Given is the interfacial tension x 10^3 in $N.m^{-1}$.

Reference: M.A.V. Devanathan, Thesis, University of Londen. 1951.

	Percentage bij weight of methanol					
E / Volts	0% *	10%	20%	30%	40%	50%
-0.950	363.8	363.8	361.6	358.7	-	-
-0.900	371.3	371.3	367.4	363.6	361.2	361.8
-0.850	377.4	376.5	373.1	370.2	367.1	364.6
-0.800	383.6	382.5	378.0	375.4	371.2	368.9
-0.750	389.8	387.7	384.0	379.6	375.4	373.0
-0.700	395.0	392.7	388.3	384.1	379.4	377.4
-0.650	400.0	397.5	392.2	388.6	383.7	380.6
-0.600	404.8	401.8	396.3	392.5	387.2	384.0
-0.550	408.7	405.9	400.2	396.2	390.6	387.7
-0.500	412.8	409.8	403.5	398.9	393.3	389.7
-0.450	416.1	413.1	406.1	405.4	395.9	399.3
-0.400	419.3	415.9	408.9	404.4	398.7	395.9
-0.350	422.1	418.5	411.3	406.2	401.0	398.0
-0.300	424.2	420.6	413.2	408.3	402.8	400.0
-0.250	425.9	422.2	414.4	409.8	404.2	402.1
-0.200	426.5	423.4	415.8	411.5	405.4	403.0
-0.150	426.4	424.1	416.5	412.3	406.4	403.6
-0.100	425.9	424.2	416.6	412.5	407.2	404.4
-0.050	424.2	423.5	416.2	412.2	407.3	404.3
-0.000	421.7	421.9	414.7	411.1	406.9	404.3
+0.050	418.3	419.0	412.6	409.0	405.0	403.0
+0.100	413.1	415.1	409.6	405.6	402.1	400.5
+0.150	408.1	410.3	404.6	401.7	398.3	397.0
+0.200	401.5	405.1	398.9	395.8	393.8	391.4
+0.250	394.1	398.2	391.2	388.3	386.7	385.3
+0.300	384.7	390.3	382.8	379.9	378.8	377.5
+0.325	-	-	377.6	375.2	373.5	372.4
+0.350	374.1	380.9	372.4	369.2	368.2	365.9
+0.375	-	-	366.4	360.7	359.6	358.9
+0.400	-	370.0	-	-	-	-

* Note. The data in aqueous 0,1 m HCl solutions differs by more than the acceptable amount from that of Conway et al (p.314). Later work by Blomgren (unpublished) shows that the data of Conway et al. is more reliable. The above data seems to be inaccurate as a result of errors in the concentration which was determined conductometrically. However, this data is included for comparison with the results in methanol-water mixtures.

Electrocapillary data on mercury for HCl in methanol-water mixtures (cont.) x)

E / Volts	Percentage by weight of methanol				
	60%	70%	80%	90%	100%
-0.900	357.9	357.3	357.0	357.6	362.0
-0.850	362.0	361.9	361.3	361.6	365.6
-0.800	366.5	365.6	364.4	364.9	369.5
-0.750	371.6	369.4	364.9	369.0	372.3
-0.700	374.9	373.3	372.4	372.6	375.7
-0.650	378.7	376.6	375.9	375.6	378.4
-0.600	382.7	380.2	379.4	378.5	381.3
-0.550	385.5	383.4	381.9	381.4	383.8
-0.500	387.8	385.8	385.0	383.9	385.8
-0.450	390.8	388.4	387.1	386.5	388.1
-0.400	392.9	390.6	389.0	388.6	389.6
-0.350	395.0	392.8	391.3	390.6	391.0
-0.300	396.8	394.8	393.5	392.1	392.1
-0.250	398.4	395.9	394.8	393.6	392.7
-0.200	399.7	397.6	395.7	394.2	393.2
-0.150	400.8	398.8	396.8	395.1	392.9
-0.100	401.1	399.2	397.2	395.4	391.6
-0.050	401.3	399.5	397.4	395.3	389.5
-0.000	400.9	398.9	397.0	394.7	386.3
+0.050	400.0	397.5	396.2	393.8	381.1
+0.100	398.5	395.2	394.3	391.2	374.0
+0.150	395.4	392.2	390.4	387.6	362.0
+0.200	390.8	386.9	386.1	383.1	-
+0.250	385.7	381.2	379.1	376.3	-
+0.275	-	-	-	371.0	-
+0.300	377.3	373.0	370.6	364.0	-
+0.325	372.8	367.0	364.4	-	-
+0.350	367.4	357.7	356.2	-	-
+0.375	356.3	-	-	-	-

x) See note on page 175.

Electrocapillary data on mercury for 0.01 m HCl solutions in methanol-water mixtures. T = 25°. x)

Potential measured with respect to a hydrogen electrode in the working solution.

Given is the interfacial tension x 10^3 in $N.m^{-1}$.

Reference: M.A.V. Devanathan, Thesis, University of London, 1951.

	Percentage by weight of methanol					
E / Volts	0%	10%	20%	30%	40%	50%
-0.900	370.1	369.8	367.3	364.8	361.9	-
-0.850	376.5	375.8	372.5	369.8	366.4	364.6
-0.800	382.4	381.8	377.3	374.6	371.1	368.9
-0.750	388.6	387.1	383.0	397.7	374.9	372.8
-0.700	393.7	392.1	387.3	384.2	379.0	376.5
-0.650	398.7	397.6	391.5	387.8	382.9	380.7
-0.600	403.3	402.0	395.8	391.6	386.3	383.4
-0.550	408.0	405.6	399.2	395.3	389.3	387.0
-0.500	411.8	409.2	403.2	398.2	392.7	389.2
-0.450	415.0	412.7	406.0	400.8	395.4	392.3
-0.400	417.4	415.9	407.5	403.5	397.7	395.0
-0.350	420.4	417.8	410.7	405.6	400.3	397.3
-0.300	423.3	419.6	412.9	407.4	401.9	399.0
-0.250	424.7	421.4	414.1	409.2	403.3	400.7
-0.200	426.1	423.0	415.2	410.4	405.0	402.1
-0.150	426.5	423.8	416.8	411.3	405.7	403.3
-0.100	426.8	424.2	417.1	412.2	406.8	404.1
-0.050	426.4	424.4	417.2	412.5	407.5	404.8
-0.000	424.5	424.2	417.1	412.6	407.6	405.3
+0.050	422.9	423.6	416.5	412.4	407.5	405.5
+0.100	420.6	421.7	415.0	411.7	406.8	405.0
+0.150	417.3	419.4	413.2	409.9	405.0	404.3
+0.200	412.5	416.5	409.5	406.5	403.4	402.0
+0.250	407.0	412.3	406.2	402.4	399.8	399.3
+0.300	410.3	407.4	400.3	398.1	394.6	394.6
+0.350	394.4	400.7	392.8	391.8	388.5	388.9
+0.375	-	-	-	-	-	386.4
+0.400	385.1	393.5	385.6	384.3	380.1	382.4
+0.425	-	-	380.3	380.8	375.6	377.7
+0.450	375.5	385.0	375.6	376.8	370.1	373.4
+0.475	-	-	366.9	372.5	363.6	-

x) See note on page 175.

Electrocapillary data on mercury for HCl in methanol-water mixtures (cont.). x)

	Percentage by weight of methanol				
E / Volts	60%	70%	80%	90%	100%
-0.900	359.2	-	358.7	-	-
-0.850	363.6	361.9	361.9	361.9	365.8
-0.800	367.0	365.1	365.3	365.5	368.6
-0.750	371.0	369.2	369.7	369.6	371.7
-0.700	374.5	372.0	373.5	373.0	374.3
-0.650	378.7	376.7	376.3	375.8	377.1
-0.600	382.3	379.3	379.2	379.2	379.7
-0.550	384.7	382.8	382.3	381.9	382.3
-0.500	387.5	385.5	385.3	385.1	384.6
-0.450	390.2	388.4	387.5	387.5	386.8
-0.400	392.9	390.4	390.2	389.4	388.5
-0.350	394.6	392.6	392.0	391.1	389.8
-0.300	396.8	394.0	393.7	392.8	391.0
-0.250	398.2	395.2	395.4	393.9	391.9
-0.200	399.7	396.8	396.6	395.2	392.6
-0.150	400.8	397.8	397.5	396.0	393.1
-0.100	401.4	399.0	398.2	396.6	393.2
-0.050	402.1	399.6	398.8	397.3	392.9
-0.000	402.4	399.9	399.3	397.4	392.3
+0.050	402.6	399.9	399.4	397.2	390.1
+0.100	402.1	399.2	399.1	397.0	386.9
+0.150	400.8	398.3	398.2	395.8	383.1
+0.200	398.6	396.4	396.2	393.2	377.5
+0.250	395.0	392.6	392.7	390.0	368.5
+0.300	391.5	388.5	389.3	385.4	-
+0.350	385.8	382.6	383.3	379.7	-
+0.375	-	-	-	375.9	-
+0.400	378.3	374.1	375.8	369.7	-
+0.425	374.1	368.0	371.2	363.9	-
+0.450	369.1	361.4	364.4	-	-
+0.475	360.8	-	358.0	-	-

x) See note on page 175.

Electrocapillary data on mercury for 0.1 m HCl solutions in methanol-water mixtures. T = 45°. x)

Potential measured with respect to a hydrogen electrode in the working solution.

Given is the interfacial tension x 10^3 in $N.m^{-1}$.

Reference: M.A.V. Devanathan, Thesis. University of London, 1951.

Percentage by weight of methanol

E / Volts	0%	10%	20%	30%	40%	50%
-0.950	366.2	364.1	362.1	-	-	-
-0.900	373.0	369.8	366.8	365.2	362.3	361.7
-0.850	377.7	376.0	372.6	370.2	366.5	366.3
-0.800	383.7	381.6	377.5	374.8	372.3	371.2
-0.750	389.0	387.1	383.2	380.0	376.0	374.8
-0.700	393.3	392.0	387.1	384.4	379.8	379.3
-0.650	398.5	396.8	391.2	388.4	384.2	383.2
-0.600	403.8	401.3	394.4	392.4	387.6	386.2
-0.550	408.2	405.3	398.4	395.1	390.3	389.5
-0.500	411.7	407.9	401.8	398.5	393.9	392.4
-0.450	415.4	411.7	404.8	401.3	396.6	394.7
-0.400	417.5	414.7	407.2	403.7	399.1	397.0
-0.350	419.9	416.3	409.1	405.9	400.8	399.7
-0.300	421.9	418.1	410.8	408.2	402.9	401.2
-0.250	423.2	420.1	412.2	408.9	404.6	403.1
-0.200	423.7	420.9	413.9	410.4	405.3	404.1
-0.150	424.0	421.8	414.5	410.9	405.9	404.9
-0.100	423.6	421.7	414.7	411.1	406.6	405.1
-0.050	421.9	421.5	414.2	410.7	406.2	405.3
-0.000	420.6	419.9	412.3	409.5	405.2	404.3
+0.050	417.9	417.5	410.9	407.3	403.7	402.7
+0.100	413.5	413.9	406.8	403.6	401.3	400.0
+0.150	409.2	409.8	402.8	399.0	398.1	396.3
+0.200	403.1	403.8	396.2	393.2	392.2	391.3
+0.250	396.9	397.4	390.0	385.6	385.7	384.5
+0.300	387.0	389.6	381.0	377.1	376.4	376.2
+0.325	-	-	377.5	371.8	371.7	370.5
+0.350	379.2	379.8	377.3	365.8	366.2	364.0
+0.375	-	-	359.3	-	356.6	351.9
+0.400	366.3	366.7	-	-	-	-

x) See note on page 175.

Electrocapillary data on mercury for HCl in methanol-water mixtures (cont.). x)

	Percentage by weight of methanol			
E / Volts	70%	80%	90%	100%
-0.950	-	-	-	-
-0.900	-	356.9	359.0	363.4
-0.850	361.9	360.6	362.1	365.9
-0.800	365.3	364.6	366.2	369.1
-0.750	370.0	368.8	369.2	372.8
-0.700	372.7	372.3	372.1	375.6
-0.650	376.6	375.1	375.9	378.4
-0.600	379.8	378.3	379.2	381.0
-0.550	382.8	381.4	381.8	383.8
-0.500	385.5	383.4	384.4	385.3
-0.450	388.4	386.4	386.5	387.9
-0.400	390.0	388.0	388.4	388.8
-0.350	392.5	391.2	390.6	390.1
-0.300	394.1	392.5	392.1	391.2
-0.250	395.5	393.7	393.5	391.4
-0.200	396.9	395.0	394.2	391.6
-0.150	397.7	395.9	394.6	391.2
-0.100	398.0	396.3	395.1	390.2
-0.050	398.0	396.4	394.5	387.9
-0.000	397.5	396.0	393.7	384.2
+0.050	395.9	394.6	392.8	378.7
+0.100	393.9	392.6	390.4	370.0
+0.150	390.3	388.9	386.1	-
+0.200	385.0	384.3	380.4	-
+0.250	378.8	377.5	373.0	-
+0.275	-	373.2	367.1	-
+0.300	369.3	367.6	359.1	-
+0.325	363.6	360.4	-	-
+0.350	353.4	-	-	-

x) See note on page 175.

Electrocapillary data on mercury for 0.01 m HCl solutions in methanol-water mixtures. T = 45°. x)

Potential measured with respect to a hydrogen electrode in the working solution.

Given is the interfacial tension x 10^3 in $N.m^{-1}$.

Reference: M.A.V. Devanathan, Thesis, University of London, 1951.

E Volts	Percentage by weight of methanol					
	0%	10%	20%	30%	40%	50%
-0.900	369.2	-	368.4	366.6	-	-
-0.850	376.4	375.8	373.2	370.2	366.7	365.5
-0.800	382.2	381.1	378.6	375.1	372.0	369.5
-0.750	388.1	386.6	383.2	379.7	375.6	374.1
-0.700	393.1	391.5	387.6	383.7	379.6	377.8
-0.650	398.0	396.0	392.2	387.8	383.8	381.0
-0.600	402.5	400.8	395.7	391.5	387.1	384.5
-0.550	406.5	404.5	399.5	395.0	390.2	387.5
-0.500	410.0	408.2	402.8	397.6	393.4	390.5
-0.450	413.6	410.9	405.5	400.8	395.7	393.6
-0.400	416.0	413.3	407.5	403.2	398.3	396.0
-0.350	418.7	415.2	410.1	405.6	400.3	397.8
-0.300	420.4	417.6	411.6	407.3	402.2	399.5
-0.250	422.3	419.5	412.7	408.4	403.8	401.3
-0.200	423.2	420.6	414.2	410.4	404.7	402.2
-0.150	423.7	421.0	415.1	411.1	406.0	403.5
-0.100	423.8	421.5	415.8	411.8	406.6	404.3
-0.050	423.9	421.8	415.9	412.1	407.4	404.8
-0.000	423.6	421.6	415.8	411.9	407.3	404.7
+0.050	422.8	420.6	415.4	411.5	406.9	404.6
+0.100	420.7	419.9	414.4	410.1	406.1	404.3
+0.150	418.3	417.9	412.4	408.2	404.3	403.1
+0.200	414.8	414.6	409.0	405.8	401.7	400.7
+0.250	410.0	410.4	405.3	402.8	398.3	397.9
+0.300	405.3	405.6	400.2	397.7	393.4	392.9
+0.350	397.8	399.6	392.9	390.9	386.7	387.1
+0.400	389.9	392.7	383.8	384.2	378.5	380.3
+0.425	-	-	380.0	380.2	374.7	375.2
+0.450	380.3	383.6	375.5	374.7	369.0	370.2
+0.475	-	-	368.5	370.1	363.9	363.8
+0.500	370.4	373.5	-	-	-	-

x) See note on page 175.

Electrocapillary data on mercury for HCl in methanol-water mixtures (cont.). x)

E Volts	Percentage by weight of methanol			
	70%	80%	90%	100%
-0.900	-	357.7	-	362.0
-0.850	361.9	362.0	362.0	365.3
-0.800	365.8	365.9	366.2	368.4
-0.750	369.0	369.8	369.5	370.9
-0.700	373.1	373.3	373.1	374.3
-0.650	376.4	376.4	375.8	377.1
-0.600	379.6	379.3	379.4	379.8
-0.550	382.7	382.8	381.6	382.2
-0.500	385.0	385.1	384.6	383.8
-0.450		387.8	386.0	386.2
-0.400	389.0	390.4	388.8	387.7
-0.350	392.2	391.7	390.2	389.1
-0.300	394.2	394.1	391.5	390.5
-0.250	395.4	395.3	393.3	391.2
-0.200	396.9	396.5	394.6	391.5
-0.150	397.4	397.6	395.4	391.2
-0.100	398.1	398.3	396.5	391.8
-0.050	398.5	398.5	397.0	391.0
-0.000	398.8	398.8	397.0	390.0
+0.050	398.9	399.0	396.7	387.9
+0.100	398.5	398.4	396.4	385.7
+0.150	396.3	397.1	394.8	381.1
+0.200	394.2	395.1	392.2	375.2
+0.250	390.8	390.9	388.1	366.5
+0.300	385.6	387.1	382.1	-
+0.350	380.4	380.8	375.6	-
+0.375	-	377.4	371.5	-
+0.400	371.0	373.8	366.4	-
+0.425	366.2	369.0	359.2	-
+0.450	-	359.2	-	-

x) See note on page 175

Electrocapillary data on mercury for various concentrations of HCl in mole.l^{-1} in methanolic solution. T = 25°. x)
Potential measured with respect to a hydrogen electrode in the working solution.
Given is the interfacial tension x 10^3 in $N.m^{-1}$.
Reference: M.A.V. Devanathan, Thesis, University of London, 1951.

$\frac{E}{\text{Volts}}$	c = 1.0	0.3	0.1	0.03	0.01
-0.900	358.5	360.2	362.0	361.8	-
-0.850	363.3	364.4	365.6	364.4	365.8
-0.800	367.5	368.2	369.5	368.0	368.6
-0.750	371.4	371.6	372.3	371.8	371.7
-0.700	375.0	374.8	375.7	374.8	374.3
-0.650	377.5	377.8	378.4	378.1	377.1
-0.600	380.7	380.9	381.3	381.8	379.7
-0.550	383.1	383.0	383.8	383.5	382.3
-0.500	385.4	385.2	385.8	385.4	384.6
-0.450	387.5	387.2	388.1	387.5	386.8
-0.400	388.8	388.8	389.6	389.0	388.5
-0.350	389.8	390.4	391.0	390.5	389.8
-0.300	390.3	391.4	392.1	391.8	391.0
-0.250	390.2	391.7	392.7	392.6	391.9
-0.200	389.4	391.2	393.2	393.1	392.6
-0.150	387.4	390.1	392.9	393.2	393.1
-0.100	384.8	388.3	391.6	393.1	393.2
-0.050	379.6	385.2	389.5	392.5	392.9
-0.000	373.8	381.5	386.3	391.6	392.3
+0.025	368.9	-	-	-	-
+0.050	365.2	375.5	381.1	389.9	390.1
+0.075	-	370.9	-	-	-
+0.100	-	364.8	374.0	386.8	386.9
+0.150	-	-	362.0	381.8	383.1
+0.175	-	-	-	-	380.4
+0.200	-	-	-	375.6	377.5
+0.225	-	-	-	369.8	373.4
+0.250	-	-	-	363.9	368.5

x) See note on page 175.

Electrocapillary data on mercury in 1.0 M NH_4NO_3 in water-methanol mixtures. Potentials measured with respect to an aqueous normal calomel electrode. Given is the surface tension in mN.m^{-1}. Temp. 18°C. Reference: C. Ockrent, J. Phys. Chem., 35, 3354 (1931).

mole fraction of MeOH

E Volts	1.0	0.95	0.85	0.75	0.50	0.25	0.15	0.05	0
+0.2	332.3	333.8	333.4	333.0	334.8	335.5	336.5	339.4	339.5
0.0	367.0	367.7	367.0	367.2	369.0	370.0	371.5	374.2	375.9
-0.2	384.4	385.4	385.7	386.3	389.0	392.3	394.9	398.6	400.9
-0.4	389.3	390.4	391.2	392.3	396.4	402.0	406.4	412.5	416.6
-0.6	386.3	387.2	388.5	389.0	393.5	400.5	405.8	415.6	421.1
-0.8	377.5	379.0	380.1	380.5	385.0	392.1	397.4	407.3	412.9
-1.0	365.6	366.8	367.4	368.3	372.4	378.3	383.1	391.0	395.3
-1.2	349.3	350.2	350.7	351.6	354.6	359.2	362.4	367.2	369.6
-1.4	326.4	327.6	327.7	328.8	330.8	333.0	333.9	336.1	336.9
-1.6	295.7	296.6	296.6	272.1	296.6	296.6			

Electrocapillary data on mercury in 0.2 M NH_4NO_3 in water-ethanol mixtures. Potentials measured with respect to an aqueous normal calomel electrode. Given is the surface tension in mN.m^{-1}. Temp. 18°C. Reference: C. Ockrent, J. Phys. Chem., 35, 3354 (1931).

mole fraction of EtOH E Volts	1.0	0.95	0.75	0.5	0.25	0.15	0.05	0
+0.2			338.3	339.4	343.7	347.5	349.7	351.1
0.0	372.0	369.7	370.9	372.5	376.2	379.0	382.6	384.8
-0.2	384.8	383.3	385.3	386.9	391.2	395.5	404.0	408.9
-0.4	386.1	385.0	387.5	389.0	392.8	397.1	410.5	422.3
-0.6	383.1	381.4	383.8	384.9	388.1	392.6	407.0	424.5
-0.8	376.5	374.9	376.4	377.2	381.0	385.1	398.8	415.3
-1.0	366.2	365.1	374.5	367.2	370.4	374.4	387.0	397.9
-1.2	352.3	351.3	352.9	353.3	355.9	360.0	368.7	373.8
-1.4	332.7	331.7	332.0	332.5	335.7	338.5	341.6	343.2
-1.6				304.9	304.8	305.7	306.4	306.5

Electrocapillary data on mercury in 0.02 M LiCl in water-n-propanol mixtures. Potentials measured with respect to an aqueous normal calomel electrode. Given is the surface tension γ in mN.m^{-1}. Temperature 18°C.

Reference: C. Ockrent, J. Phys. Chem., 35, 3354 (1931).

mole fraction of n-PrOH $\frac{E}{\text{volts}}$	0.85	0.50	0.25	0.15	0.05	0.02	0.01	0.0
-0.2		381.2	384.7	385.8	389.7	401.7	407.0	411.3
-0.4		383.4	384.7	385.5	388.7	401.4	409.7	425.3
-0.6	380.6	381.6	381.8	381.7	385.1	397.6	406.7	425.6
-0.8	378.6	376.6	376.5	376.5	379.3	391.7	400.4	417.4
-1.0	371.5	369.3	369.1	369.0	372.1	383.6	391.9	403.0
-1.2		357.2	358.8	358.3	361.6	371.7	376.8	381.7
-1.4			344.3	343.8	346.2	350.2	352.0	354.3
-1.6					320.0	320.0	321.6	321.7

Electrocapillary data on mercury in 0.2 M LiCl in water-n-propanol mixtures. Potentials measured with respect to an aqueous normal calomel electrode. Temperature 18°C. Given is the surface tension in mN.m^{-1}.

Reference: C. Ockrent, J. Phys.Chem., 35, 3354,(1931).

mole fraction of n-PrOH	1.0	0.95	0.85	0.75	0.50	0.25	0.15
E/volts							
-0.2	376.9	378.1	377.7	375.3	378.9	382.8	383.8
-0.4	381.2	381.1	381.7	381.8	382.6	384.1	384.7
-0.6	379.0	378.9	379.1	379.2	379.3	380.3	373.3
-0.8	375.1	372.5	372.9	372.7	373.4	374.0	374.7
-1.0	366.1	363.5	363.6	363.7	365.3	365.0	365.7
-1.2	355.1	351.5	351.5	350.5	352.9	353.1	354.2
-1.4	340.0	335.0	332.4				
-1.6	321.2	314.1					

mole fraction of n-PrOH	0.05	0.02	0.01	0.005	0.0025	0.0
E/volts						
0.0		370.1	371.6	370.9	371.3	373.6
-0.2	387.6	398.6	402.9	404.5	405.4	406.4
-0.4	388.2	399.3	407.8	416.0	420.6	423.4
-0.6	384.3	395.6	404.5	412.6	419.2	425.1
-0.8	378.1	388.3	398.0	406.1	411.3	414.7
-1.0	369.7	380.8	388.3	393.6	396.1	397.3
-1.2	357.8	367.1	370.4	372.7	373.5	382.3
-1.4	338.1	342.8	343.0	344.2	344.2	344.1
-1.6		307.8	307.8	308.1	307.8	307.6
-1.8		265.3	265.1	265.4	264.4	264.5

Electrocapillary data on mercury in 0.1 M LiCl in water-n-butanol mixtures. Potentials measured with respect to an aqueous normal calomel electrode. Temperature 20°C. Given is the surface tension in $mN.m^{-1}$.

Reference: C. Ockrent, J. Phys. Chem., 35, 3354 (1931).

mole fraction of n-BuOH E/volts	1.00	0.95	0.85	0.01	0.0025	0.0
0.0				374.1	376.7	377.1
-0.2	367.3	376.3	374.9	393.2	406.2	408.6
-0.4	377.6	378.8	379.5	392.4	410.3	424.3
-0.6	378.6	378.7	378.1	389.3	407.0	425.1
-0.8	377.7	374.8	373.4	384.3	401.6	416.7
-1.0	374.2	367.4	365.8	376.8	392.2	398.8
-1.2	369.2	360.2	355.3	365.7	374.1	375.8
-1.4	362.3	349.1	340.5	344.8	346.5	346.4
-1.6			317.7	310.9	310.3	310.5
-1.8				268.8	268.6	268.5

Electrocapillary data for mercury in 0.1 mol l^{-1} aqueous $HClO_4$ containing 1-butanol T = 25°C

Potentials with respect to saturated NaCl calomel electrode

Reference: D.E. Broadhead, R.N. Cochran and R.S. Hansen, unpublished.

	γ /mN m^{-1}							
c/mol l^{-1} E/volt	0	0.240	0.0469	0.0895	0.197	0.236	0.493	0.938
-1.200	372.43	372.17	371.83	371.30	369.72		363.76	356.13
-1.150	378.81	378.48	378.13	377.39	375.06		367.48	359.07
-1.100	384.91	384.44	383.87	382.88	379.59	377.97	370.64	361.96
-1.050	390.62	390.06	389.36	387.98	383.52	381.50	373.61	364.70
-1.000	395.98	395.16	394.30	392.36	386.76	384.56	376.23	367.15
-0.950	400.95	399.98	398.82	396.28	389.68	387.24	378.64	369.54
-0.900	405.56	404.32	402.88	399.60	392.20	389.61	380.79	371.74
-0.850	409.67	408.22	406.44	402.46	394.43	391.74	382.86	373.71
-0.800	413.50	411.72	409.52	404.88	396.37	393.65	384.64	375.61
-0.750	416.80	414.81	412.18	406.99	398.24	395.47	386.37	377.24
-0.700	419.66	417.42	414.41	408.83	399.79	397.02	387.83	378.70
-0.650	422.04	419.68	416.35	410.38	401.21	398.40	389.25	380.12
-0.630	422.89							
-0.625		420.62	417.07		401.85		389.87	380.70
-0.600	423.90	421.37	417.78	411.76	402.40	399.48	390.39	381.33
-0.580	424.50							
-0.575		422.09	418.38		402.98		390.93	381.94
-0.550	425.07	422.57	418.98	412.82	403.54	400.66	391.49	382.43
-0.540	425.26							
-0.530	425.40							
-0.525		422.90	419.35		403.97		391.95	382.98
-0.520	425.49							
-0.510	425.54							
-0.500	425.54	423.17	419.79	413.72	404.42	401.53	392.32	383.43
-0.490	425.58							
-0.480	425.58							
-0.475		423.27	420.04		404.82		392.73	383.86
-0.450	425.34	423.21	410.20	414.46	405.13	402.30	393.10	384.27
-0.430	425.09							
-0.425		423.03	420.15		405.42		393.45	384.59
-0.400	424.55	422.70	420.18	414.85	405.71	402.87	393.68	384.90
-0.380	424.08							
-0.375					405.97			
-0.350	423.19	421.55	419.57	415.01	406.08	403.33	394.26	385.36
-0.325					406.18			
-0.300	421.08	419.72	418.33	414.59	406.27	403.54	394.44	385.66
-0.250	418.37	417.37	416.20	413.58	406.15	403.52	394.67	385.87
-0.200	415.04	414.25	413.37	411.57	405.50	403.14	394.51	385.80
-0.150	411.12	410.47	409.78	408.57	404.21	402.18	394.14	385.40
-0.100	406.57	406.15	405.61	408.66	401.81	400.33	393.20	384.56
-0.050	401.49	401.14	400.74	400.10	398.11	397.19	391.58	383.03
0.000	395.77	395.56	395.23	394.74	393.38	392.78	388.86	380.70
0.050	389.43	389.26	389.06	388.75	387.73	387.36	384.84	377.33
0.100	382.55	382.38	382.29	381.96	381.34	381.04	379.40	373.07
0.150	374.90	374.77	374.67	374.50	374.01	373.79	372.78	367.55
0.200	366.63	366.68	366.51	366.42	366.09	365.96	365.34	361.38
E^z volt	-0.486	-0.470	-0.431	-0.368	-0.292	-0.271	-0.234	-0.239
γ^z/mN m^{-1}	425.60	423.28	420.26	415.06	406.31	403.61	394.75	385.88

Electrocapillary data for adsorption of n-propanol in 1.0 M KCl.
Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ/mN m^{-1}							
c/mol l^{-1}	0.03	0.1	0.3	0.5	1.0	2.0	3.0	4.0
-E/mV								
0						338.4	334.1	331.4
50	368.1	368.2	367.5	367.4	366.0			
100	379.8	379.7	378.8	378.7	377.9	372.4	370.5	369.9
150								379.1
200	397.2	397.3	395.9	395.5	393.6	385.6	382.6	382.2
240							384.2	384.0
250					397.6	387.7		
260							384.8	
280							385.4	384.9
300	409.8	409.8	407.7	405.9	399.2	388.4	385.7	385.3
320					399.3	388.4	385.7	385.5
330				407.7				
340					399.3	388.6	385.7	385.5
350				407.9				
360					399.3	388.6	385.6	385.4
380			413.3	408.1	399.3	388.4	385.5	385.5
400	418.6	418.4	413.4	408.0	399.1	388.3	385.4	385.5
420			413.7	408.0				
440				407.8				
450	421.8	420.8	413.7		398.5			384.6
460			413.6					
480			413.5	407.4				
500	423.1	422.0	413.3		397.9	387.2	384.0	384.0
520	423.5	422.2						
540	423.6	422.2						
550			412.3	405.9				
560	423.7	422.2						
580	423.7	421.9						
600	423.4	421.5	411.0	404.8	395.6	384.9	381.8	382.0
650	422.5	420.2						
700	420.4	418.2	407.9	401.7	392.5	381.7	379.1	379.3
800	414.6	412.8	403.8	398.0	389.2	378.2	375.6	375.7
900	407.0	405.5	398.5	393.3	384.9	374.3	371.3	371.1
1000	397.1	396.5	391.6	387.5	379.9	369.0	366.7	366.6
1100	385.7	385.2	382.4	379.9	373.7	363.6	361.1	360.9
1200	372.6	372.6	370.5	369.3	366.4	357.1	354.1	354.3

Electrocapillary data for adsorption of allyl alcohol in 1.0 M KCl.
Given are the interfacial tension in mN m^{-1}, potential in mVvs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ/mN m^{-1}				
c/mol l^{-1}	0.03	0.1	0.3	0.5	1.0
-E/mV					
50	368.5	368.9	367.9	367.4	366.0
100	380.2	380.4	379.5	379.1	377.4
200	397.4	397.4	396.4	395.4	393.3
300	410.1	410.0	408.4	406.9	402.8
350					404.8
360				410.9	
400	418.9	418.4	415.3	412.4	405.4
420				412.9	405.4
440				413.3	405.4
450	421.4	420.8	417.3		
460				413.3	405.4
480			417.5	413.3	405.2
500	423.1	422.4	417.7	413.2	405.0
520	423.5	422.7	417.6		
540	423.7	422.7	417.5		
550				412.8	404.2
560	423.6	422.7	417.4		
580	423.6	422.6			
600	423.5	422.4	416.7	411.6	403.2
650	422.2	421.1	415.4		
700	420.3	419.2	413.6	408.8	400.6
800	414.3	413.4	408.9	404.6	396.7
900	406.4	405.8	402.5	399.0	391.6
1000	396.4	396.3	394.0	391.8	385.9
1100	384.9	385.0	384.0	382.2	378.2
1200	371.7	371.8	371.3	370.4	368.4

Electrocapillary data for adsorption of propargyl alcohol+) in 1.0 M KCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

			γ /mN m^{-1}		
c/mol l^{-1}	0.03	0.1	0.3	0.5	1.0
-E/mV					
50	368.9	368.3	367.1	367.1	365.1
100	380.3	380.1	378.7	378.6	376.3
200	397.6	396.9	395.3	395.0	391.8
300	410.3	409.5	406.7	406.1	402.2
400	419.3	418.2	415.2	413.8	407.9
450	421.9	420.9	417.9	415.7	409.3
480					410.1
500	423.5	422.5	419.0	417.0	410.1
520	424.0	422.8	419.3	417.1	410.1
540	424.2	423.1	419.3	417.1	410.0
560	424.2	423.1	419.2	417.0	409.9
580	424.1	423.1	419.2	416.9	
600	423.9	422.8	418.9	416.6	409.4
650	422.9	421.5	417.8	415.5	408.4
700	420.8	419.6	416.0	414.1	407.1
800	414.9	414.0	410.8	409.2	402.8
900	406.9	406.3	403.8	402.5	397.2
1000	396.9	396.6	394.6	394.2	389.8
1100	385.3	385.0	383.7	383.6	380.2
1200	371.6	371.6	370.9	371.2	368.6

+) Propargyl alcohol is 2 propyn-1-ol .

Electrocapillary data on mercury electrode in 0.1 M KCl at 25°C. Adsorption of ethylene glycol - given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: S. Trasatti. J. Electroanal. Chem. 28 (1970) 257.

	γ /mN m^{-1}								
c/mol l^{-1}	0	0.1	0.2	0.4	0.7	1.0	1.3	1.5	2.0
-E/V									
1.20	371.0	370.6	370.2	370.0	369.0	368.0	367.3	366.4	364.3
1.15	377.6	377.2	376.9	376.5	375.4	374.3	373.4	372.7	371.3
1.10	384.0	383.6	383.2	382.7	381.4	380.2	379.3	378.5	377.1
1.05	389.9	389.4	388.9	388.4	387.0	385.8	384.8	383.8	382.5
1.00	395.4	394.9	394.4	393.9	392.2	391.0	389.8	388.8	387.4
0.95	400.4	399.9	399.4	398.7	397.1	395.7	394.5	393.5	391.9
0.90	405.1	404.6	403.9	403.4	401.6	400.1	398.8	397.7	396.1
0.85	409.4	408.8	408.2	407.6	405.7	404.1	402.9	401.6	400.0
0.80	413.3	412.6	412.0	411.2	409.4	407.7	406.3	405.2	403.4
0.75	416.7	416.0	415.3	414.4	412.5	410.9	409.6	408.2	406.6
0.70	419.7	418.8	418.3	417.3	415.3	413.7	412.3	410.9	409.3
0.65	422.0	421.2	420.6	419.6	417.7	416.0	414.6	413.2	411.6
0.60	424.0	423.1	422.4	421.5	419.6	417.9	416.5	415.2	413.4
0.55	425.3	424.5	423.9	422.9	420.9	419.4	417.9	416.5	414.9
0.50	426.0	425.3	424.7	423.8	421.9	420.3	418.9	417.5	415.9
0.45	426.2	425.4	424.8	423.9	422.1	420.6	419.3	418.0	416.4
0.40	425.4	424.8	424.2	423.4	421.7	420.3	419.0	417.9	416.3
0.35	423.8	423.3	422.7	422.0	420.5	419.2	417.6	417.0	415.4
0.30	421.1	420.6	420.2	419.5	418.3	417.1	415.6	415.0	413.7
0.25	417.7	417.2	416.9	416.3	415.1	414.0	412.9	412.3	411.1
0.20	413.0	412.7	412.3	412.0	410.9	409.9	409.0	408.3	407.4
0.15	407.6	407.2	407.0	406.6	405.7	404.7	404.0	403.4	402.6
0.10	401.0	400.8	400.4	400.2	399.5	398.6	397.8	397.5	396.6
0.05	393.6	393.4	393.1	392.7	392.1	391.4	390.6	390.3	389.6
0.00	384.7	384.5	384.4	384.1	383.5	382.8	382.1	381.7	381.1

Electrocapillary data on mercury electrode in 0.1 M NaF at 25°C. Adsorption of ethylene glycol - given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: S. Trasatti. J. Electroanal. Chem. 28 (1970) 257.

	γ/mN m^{-1}								
c/mol l^{-1}	0	0.1	0.2	0.5	0.7	1.0	1.2	1.6	2.0
-E/V									
1.30	355.8								
1.25	363.4								
1.20	370.5	370.0	369.5	369.1	368.3	366.8	366.5	365.2	364.7
1.15	377.1	376.4	375.9	375.0	374.5		372.6	371.3	370.8
1.10	383.4	382.6	382.0	381.2	380.4		378.3	377.1	275.2
1.05	389.0	388.4	388.3	386.5	386.0		383.8	382.3	381.4
1.00	394.5	393.5	393.1	391.9	391.1	389.5	388.9	387.3	386.3
0.95	399.5	398.7	398.0	396.7	395.9		393.6	391.9	390.8
0.90	404.2	403.5	402.6	401.2	400.3		397.9	396.1	395.1
0.85	408.4	407.5	406.6	405.2	404.0	402.3	401.8	400.0	398.8
0.80	412.2	411.2	410.4	408.9	407.5		405.4	403.3	402.2
0.75	415.6	414.5	413.6	412.1	410.9		408.5	406.5	405.0
0.70	418.4	417.5	416.5	414.9	413.7	411.9	411.3	409.2	407.9
0.65	420.9	419.9	418.9	417.2	416.0		413.6	411.4	410.2
0.60	422.8	421.8	420.7	419.1	418.0	416.0	415.4	413.1	412.1
0.55	424.2	423.2	422.2	420.5	419.3	417.6	416.8	414.4	413.5
0.50	425.0	424.1	423.2	421.5	420.3	418.6	417.8	415.6	414.6
0.45	425.3	424.4	423.6	421.9	420.8	419.0	418.4	416.3	414.7
0.40	424.8	424.1	423.2	421.8	420.8	418.9	418.4	416.9	414.7
0.35	423.6	423.2	422.2	420.9	420.2	418.4	417.8	416.0	414.1
0.30	421.8	421.4	420.5	419.4	418.7	417.1	416.6	415.0	412.3
0.25	419.1	419.1	418.0	417.1	416.7		414.6	412.4	410.0
0.20	415.6	415.9	414.7	414.1	413.7	412.4	412.1	410.5	406.1
0.15	414.2	411.7	410.5	410.2	410.0		408.6	406.7	402.4
0.10	405.6	406.6	405.7	405.4	405.3	404.4	404.2	402.1	398.2
0.05	399.6	401.2	400.1	400.0	400.0		398.8	397.0	394.0
0.00	393.2	394.7	393.7	393.6	393.7	393.0	392.7	390.7	389.8

Electrocapillary data on mercury electrode in 0.1M KI at 25°C. Adsorption of ethylene glycol - given are the interfacial tension in mN m^{-1}, potential V vs SCE.

Reference: S. Trasatti. J. Electroanal. Chem. 28 (1970) 257.

	γ /mN m^{-1}								
c/mol l^{-1}	0	0.1	0.2	0.4	0.7	1.0	1.3	1.6	2.0
-E/V									
1.20	370.5	370.2	370.0	369.4	368.4	367.4	366.6	365.6	354.7
1.15	377.0	376.8	376.6	375.8	374.6	373.7	372.9	371.7	370.7
1.10	383.5	382.9	382.8	382.0	380.7	379.8	378.6	377.6	376.5
1.05	389.2	388.7	388.6	387.7	386.2	385.2	384.1	382.9	381.7
1.00	394.8	394.3	394.0	393.0	391.7	390.4	389.2	388.0	386.6
0.95	399.7	399.3	398.8	397.8	396.3	395.2	393.8	392.6	391.2
0.90	404.4	403.7	403.3	402.2	400.6	399.3	398.0	396.8	395.3
0.85	408.2	407.6	407.1	406.1	404.4	403.0	401.8	400.4	398.8
0.80	411.4	410.7	410.2	409.2	407.4	406.1	404.8	403.5	402.0
0.75	413.5	412.8	412.3	411.2	409.7	408.2	407.1	405.7	404.3
0.70	414.3	413.8	413.3	412.3	410.7	409.4	408.0	407.0	405.5
0.65	413.7	413.2	412.8	411.9	410.3	409.1	407.9	406.8	405.5
0.60	411.4	410.9	410.5	409.7	408.3	407.2	406.2	405.1	403.9
0.55	407.4	407.0	406.5	405.9	404.6	403.8	402.6	401.7	400.6
0.50	401.3	400.9	400.3	399.9	398.7	398.0	397.1	396.1	395.2
0.45	392.6	392.3	391.7	391.3	390.2	389.6	388.9	388.0	387.2
0.40	380.1	379.9	379.6	379.0	377.9	377.5	377.0	376.5	375.4
0.35	362.4	362.4	361.6	361.4	360.7	360.0	359.6	359.1	358.1

Electrocapillary data on mercury electrode in 0.1 M KBr at 25°C. Adsorption of ethylene glycol - given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: S. Trasatti. J. Electroanal. Chem. 28 (1970) 257.

	γ /mN m^{-1}								
c/mol l^{-1}	0	0.1	0.2	0.4	0.7	1.2	1.3	1.6	2.0
-E/V									
1.20	370.7	370.4	370.0	369.7	368.7	367.5	367.3	366.8	365.3
1.15	377.4	377.0	376.5	376.1	375.2	373.9	373.5	372.8	371.3
1.10	383.7	383.3	382.7	382.3	381.2	379.8	379.3	378.5	377.0
1.05	389.6	389.2	388.6	388.1	386.9	385.3	384.7	383.8	382.3
1.00	395.2	394.6	394.0	393.4	392.1	390.6	389.8	388.9	387.2
0.95	400.2	399.6	399.1	398.4	397.0	395.3	394.4	393.4	391.7
0.90	405.0	404.3	403.7	402.9	401.4	399.7	398.7	397.7	395.9
0.85	409.1	408.6	408.0	407.0	405.5	403.7	402.6	401.4	399.7
0.80	413.1	412.3	411.7	410.7	409.2	407.3	406.2	405.1	403.2
0.75	416.4	415.7	415.0	414.0	412.4	410.4	409.2	408.1	406.1
0.70	419.2	418.5	417.7	416.7	415.1	413.2	412.0	410.7	408.9
0.65	421.5	420.8	420.0	418.9	417.3	415.4	414.1	412.9	411.0
0.60	423.1	422.4	421.7	420.6	419.0	417.1	415.8	414.5	412.8
0.55	424.0	423.2	422.5	421.4	419.8	418.1	416.9	415.5	413.8
0.50	423.7	423.0	422.4	421.4	419.9	418.3	417.0	415.9	414.2
0.45	422.4	421.7	421.1	420.3	418.9	417.4	416.1	415.1	413.6
0.40	419.8	419.3	418.7	418.0	416.7	415.4	414.1	413.3	411.9
0.35	416.0	415.5	415.1	414.4	413.3	412.1	411.0	410.3	408.9
0.30	411.0	410.4	410.2	409.6	408.7	407.7	406.5	405.7	404.6
0.25	404.7	404.3	404.0	403.5	402.6	401.7	400.8	399.9	398.9
0.20	396.7	396.4	396.1	395.7	394.9	394.2	393.1	392.4	391.6
0.15	386.7	386.4	386.1	385.9	385.1	384.4	383.2	382.8	382.0
0.10	372.7	372.4	372.2	372.1	371.2	370.6	369.4	368.8	368.2
0.05	348.0								

Electrocapillary data on mercury electrode in 0.916 M NaF at 25^{o}C. Adsorption of 5 chloro-1-pentanol - given are the interfacial tension in mN m^{-1}, potential in V vs Orion fluoride reversible electrode.

Reference: K. Doblhofer and D.M. Mohilner. J. Phys. Chem. 75 (1971) 1698.

	γ /mN m^{-1}							
c/mol l^{-1}	0.04330	0.03436	0.02590	0.01718	0.01290	0.00859	0.00567	0.00465
E/V								
-1.40		300.8	300.6	300.6		300.5	300.8	300.8
-1.35		310.9	310.9	310.8		310.8	310.8	310.9
-1.30	320.8	320.8	320.8	320.9		320.5	320.7	320.7
-1.25	329.9	329.7	330.0	329.9		329.9	330.0	329.9
-1.20	338.5	338.9	338.8	339.0	338.7	338.8	338.8	338.8
-1.15	346.4	346.7	347.1	347.2	347.7	347.4	347.2	347.1
-1.10	353.4	354.4	354.8	355.1	355.0	355.2	355.3	355.3
-1.05	358.5	359.9	361.5	362.2	362.2	362.6	362.6	362.7
-1.00	362.3	364.4	367.3	368.5	369.1	369.4	370.0	369.9
-0.95	366.3	367.9	370.9	373.3	375.1	375.7	376.4	376.5
-0.90	369.4	371.2	374.7	377.1	379.6	381.2	382.4	382.7
-0.85	371.7	373.9	377.6	380.4	384.0	385.6	387.8	388.1
-0.80	374.4	376.8	380.7	383.2	386.4	389.3	392.9	393.7
-0.75	376.5	379.2	382.9	386.0	388.9	392.5	395.9	397.5
-0.70	378.9	381.5	385.3	388.1	391.4	395.0	399.6	400.9
-0.65	380.8	383.7	387.0	390.5	393.7	397.2	402.6	403.7
-0.60	382.8	385.4	389.1	392.5	395.4	398.9	404.6	406.0
-0.55	384.2	387.0	390.4	394.3	397.3	400.9	406.3	408.5
-0.50	385.9	388.5	391.9	395.5	398.6	402.5	407.3	410.2
-0.45	386.8	389.7	393.2	397.1	400.1	403.7	408.7	411.7
-0.40	388.3	390.8	394.1	398.0	401.0	404.6	409.5	412.8
-0.35	388.6	391.6	394.9	398.8	402.0	405.7	409.9	413.8
-0.30	389.2	392.1	395.6	399.3	402.4	406.3	410.3	414.5
-0.25	389.6	392.6	395.9	400.1	403.0	406.6	410.7	414.9
-0.20	389.9	393.0	396.2	400.0	403.4	406.7	410.8	415.3
-0.15	390.0	393.1	396.2	400.3	403.3	407.0	410.8	415.3
-0.10	389.7	393.0	396.2	400.3	403.2	406.8	410.4	414.8
-0.05	389.5	392.4	395.6	400.0	402.3	406.5	409.9	414.0
0.00	388.9	391.6	395.2	399.3	402.1	405.7	409.0	412.7
0.05	387.9	390.9	394.1	397.9	400.6	404.4	408.0	410.4
0.10	387.0	389.6	393.1	396.7	399.6	403.0	406.7	407.3
0.15	385.5	388.3	391.5	394.6	397.8	400.2	403.1	402.2
0.20	383.9	386.4	389.7	392.3	394.4	394.4	397.0	396.0
0.25	381.6	383.4	386.6	387.3	387.6	387.4	389.7	388.9
0.30	376.7	379.0	378.5	378.5	379.0	378.3	380.7	380.6

Electrocapillary data for 5 chloro-1-pentanol (continued)

	γ /mN m^{-1}					
c/mol l^{-1}	0.003487	0.002325	0.001716	0.001134	0.000429	0
E/V						
-1.40			300.5	300.7	301.1	300.6
-1.35	310.6	310.2	310.8	311.0	310.7	310.8
-1.30	320.4	320.6	320.8	320.7	320.5	320.7
-1.25	330.0	330.0	330.1	330.0	329.7	330.0
-1.20	338.9	339.0	338.9	338.9	338.7	339.0
-1.15	347.5	347.2	347.4	347.4	348.1	347.4
-1.10	355.6	355.8	355.7	355.4	355.3	355.6
-1.05	363.4	363.3	363.3	363.0	363.8	363.2
-1.00	370.3	370.6	370.2	370.2	370.4	370.5
-0.95	377.4	377.0	377.0	376.9	377.9	377.4
-0.90	383.7	384.1	383.6	383.3	383.8	383.9
-0.85	389.6	389.7	389.5	389.2	390.6	389.9
-0.80	394.4	394.8	394.5	394.5	395.6	395.5
-0.75	398.9	399.7	399.9	399.8	401.1	400.8
-0.70	402.9	404.3	404.0	404.6	405.3	405.7
-0.65	406.1	407.6	408.0	408.8	410.4	410.1
-0.60	408.5	410.9	411.0	412.4	413.5	414.1
-0.55	410.4	413.4	413.7	415.8	417.7	417.5
-0.50	412.6	415.5	415.7	418.0	419.7	420.7
-0.45	414.3	417.4	417.7	420.4	422.3	423.1
-0.40	415.3	418.3	418.9	421.9	424.1	425.1
-0.35	416.7	419.5	419.9	423.3	426.0	426.5
-0.30	417.1	419.9	420.9	423.9	426.8	427.2
-0.25	417.7	420.4	420.9	424.5	427.1	427.4
-0.20	417.8	420.4	421.0	424.5	426.3	426.8
-0.15	417.8	420.3	420.9	423.7	425.6	425.6
-0.10	417.3	419.6	420.5	422.5	423.2	423.7
-0.05	416.5	418.6	419.3	420.0	420.5	421.0
0.00	415.2	416.7	416.6	417.2	417.0	417.4
0.05	412.9	413.3	412.9	412.9	412.9	413.2
0.10	408.7	408.8	408.0	408.6	407.2	408.2
0.15	403.4	403.5	402.3	402.7	401.5	402.5
0.20	396.7	396.5	395.7	396.1	394.6	395.8
0.25	389.5	387.3		388.9		387.8
0.30	380.8	379.0		380.6		377.5

Electrocapillary data for iso-pentanol in 0.1 M $HClO_4$ solution.
Reference electrode: SCE (NaCl solution). $T = 25^{\circ}C$
γ: Interfacial tension in $mN\ m^{-1}$.
Reduced concentration c/c_0 where c is solute concentration and c_0 its concentration in the base electrolyte.

Reference: K.G. Baikerikar and R.H. Hansen. J. Colloid Interface Sci. 52 (1975) 277.

c/c_0	0.01234	0.02439	0.03614	0.04761	0.06976
E/V					
-1.10	384.6	384.5	384.4	384.3	383.9
-1.05	390.3	390.1	389.9	389.8	389.4
-1.00	395.6	395.3	395.2	394.9	394.3
-0.95	400.5	400.2	399.8	399.5	398.6
-0.90	405.1	404.5	404.2	403.7	402.5
-0.85	409.1	408.5	408.0	407.4	405.9
-0.80	412.8	412.1	411.5	410.6	408.6
-0.75	416.0	415.2	414.4	413.4	411.0
-0.70	418.8	417.8	417.0	415.7	413.0
-0.65	421.0	420.1	418.9	417.6	414.5
-0.60	422.9	421.8	420.6	419.1	415.8
-0.55	424.1	423.0	421.7	420.2	416.8
-0.50	424.7	423.7	422.5	420.9	417.6
-0.45	424.7	423.8	422.7	421.4	418.2
-0.40	424.0	423.3	422.4	421.2	418.3
-0.35	422.8	422.1	421.3	420.6	418.2
-0.30	420.8	420.3	419.7	419.2	417.2
-0.25	418.2	417.8	417.4	417.0	415.6
-0.20	415.0	414.7	414.4	414.2	413.2
-0.15	411.1	410.8	410.6	410.5	409.9
-0.10	406.5	406.3	406.2	406.2	405.7
-0.05	401.4	401.3	401.1	401.1	400.8
0.00	395.6	495.5	395.4	395.5	395.2
0.05	389.4	389.3	389.2	389.3	389.1
0.10	382.5	382.4	382.3	382.4	382.2
0.15	374.8	374.8	374.7	374.9	374.7
0.20	366.7	366.6	366.6	366.7	366.5

Electrocapillary data for iso-pentanol in 0.1 M $HClO_4$ (continued)

c/c_θ	0.09090	0.1304	0.2000	0.3333	0.500
E/V					
-1.10	383.6	383.1	381.6	379.0	375.5
-1.05	388.9	388.1	385.8	382.3	378.6
-1.00	393.6	392.4	389.6	385.4	381.2
-0.95	397.7	396.1	392.5	387.7	383.4
-0.90	401.2	399.1	395.2	390.0	385.5
-0.85	404.2	401.6	397.2	391.7	387.2
-0.80	406.7	403.7	399.1	393.6	388.9
-0.75	408.7	405.5	400.7	394.9	390.3
-0.70	410.5	407.0	402.2	396.4	391.7
-0.65	411.9	408.2	403.3	397.5	392.8
-0.60	413.2	409.4	404.5	398.6	393.9
-0.55	414.1	410.3	405.4	399.6	394.7
-0.50	415.0	411.1	406.4	400.2	395.5
-0.45	415.5	411.7	406.8	400.9	396.1
-0.40	415.9	412.1	407.5	401.3	396.7
-0.35	416.0	412.4	407.6	401.7	397.0
-0.30	415.4	412.3	407.8	401.9	397.3
-0.25	414.2	411.7	407.4	401.9	397.4
-0.20	412.2	410.3	406.6	401.7	397.1
-0.15	409.3	408.0	405.0	400.9	396.6
-0.10	405.3	404.7	402.5	399.2	395.4
-0.05	400.6	400.2	398.8	396.7	393.3
0.00	395.1	394.9	394.1	392.8	390.2
0.05	389.0	388.9	388.3	387.7	386.0
0.10	382.2	382.1	381.7	381.4	380.4
0.15	374.6	374.6	374.3	374.2	373.6
0.10	366.5	366.5	366.2	366.3	365.8

Electrocapillary data for n-butanol in 0.1 M $HClO_4$ solution.
Reference electrode: SCE (NaCl solution). T = 25°C
γ: Interfacial tension in mN m^{-1}.
Reduced concentration c/c_o where c is solute concentration and c_o its concentration in the base electrolyte.

Reference: K.G. Baikerikar and R.H. Hansen. J. Colloid Interface Sci. 52 (1975) 277.

c/c_o	0.01234	0.02439	0.03614	0.04761	0.06976
E/V					
-1.10	384.6	384.2	383.9	383.6	383.1
-1.05	390.3	389.8	389.4	389.1	388.3
-1.00	395.5	395.0	394.4	394.1	392.9
-0.95	400.4	399.7	399.1	398.5	397.1
-0.90	404.8	404.0	403.3	402.5	400.8
-0.85	408.9	408.0	407.1	406.1	404.0
-0.80	412.5	411.4	410.3	409.2	406.7
-0.75	415.7	414.5	413.2	411.9	409.1
-0.70	418.4	417.1	415.6	414.2	411.1
-0.65	420.8	419.3	417.7	416.1	412.9
-0.60	422.5	421.0	419.3	417.7	414.3
-0.55	423.8	422.3	420.6	418.9	415.5
-0.50	424.4	423.0	421.4	419.8	416.4
-0.45	424.6	423.2	421.8	420.3	417.0
-0.40	423.9	422.7	421.5	420.1	417.3
-0.35	422.7	421.6	420.6	419.6	417.1
-0.30	420.7	419.9	419.1	418.3	416.3
-0.25	418.2	417.5	416.9	416.3	414.9
-0.20	414.9	414.4	413.9	413.5	412.5
-0.15	411.1	410.6	410.3	410.0	409.3
-0.10	406.5	406.2	405.9	405.7	405.3
-0.05	401.4	401.1	400.9	400.8	400.5
0.00	395.7	395.4	395.3	395.2	395.0
0.05	389.5	389.2	389.1	389.1	388.9
0.10	382.5	382.3	382.2	382.2	382.2
0.15	374.9	374.8	374.8	374.8	374.7
0.20	366.7	366.6	366.6	366.6	366.6

Electrocapillary data for n-butanol in 0.1 M $HClO_4$ solution (continued)

c/c_o	0.09090	0.1304	0.2000	0.3548	0.5000
E/V					
-1.10	382.3	380.9	378.8	373.8	370.1
-1.05	387.4	385.6	382.8	377.2	373.0
-1.00	391.7	389.4	386.2	379.9	375.7
-0.95	395.7	393.0	389.3	382.6	378.1
-0.90	398.9	395.9	391.8	384.8	380.3
-0.85	401.9	398.5	394.1	386.9	382.4
-0.80	404.3	400.7	396.2	388.8	384.2
-0.75	406.6	402.7	398.0	390.6	385.9
-0.70	408.4	404.4	399.6	392.1	387.4
-0.65	410.0	405.9	401.0	393.5	388.8
-0.60	411.5	407.3	402.4	394.8	390.1
-0.55	412.6	408.4	403.4	395.8	391.2
-0.50	413.6	409.3	404.5	396.8	392.1
-0.45	414.2	410.1	405.2	397.6	392.9
-0.40	414.7	410.6	405.8	398.2	393.6
-0.35	414.7	410.9	406.2	398.7	394.1
-0.30	414.1	410.7	406.3	398.9	394.4
-0.25	413.0	410.0	405.9	399.0	394.5
-0.20	411.1	408.6	405.1	398.5	394.3
-0.15	408.4	406.4	403.5	397.8	393.6
-0.10	404.6	403.2	401.1	396.1	392.5
-0.05	400.0	499.1	397.6	393.7	390.6
0.00	394.6	394.0	393.1	390.1	387.6
0.05	388.6	388.2	387.6	385.6	383.8
0.10	381.8	381.5	381.2	379.9	378.6
0.15	374.4	374.2	374.0	373.1	372.3
0.20	366.4	366.2	366.2	365.5	365.0

Electrocapillary data on mercury electrode in 0.1 M $NaClO_4$ $+CH_3OH$.
Given are the surface tension in mN m^{-1}, potential in mV vs SCE (NaCl).
T = 25°C

Reference: J. Taraszewska. J.Electroanal. Chem. 49 (1974) 443.

X is mole fraction of methanol

	γ/mN m^{-1}							
X	0	0.035	0.0735	0.167	0.4171	0.568	0.8993	0.9959
-E/mV								
0	393.7	393.8	390.8	387.6	383.4	377.8	375.7	-
25	396.7	396.6	393.6	390.2	385.7	380.0	377.8	-
50	399.7	399.2	396.3	392.6	387.9	382.1	379.6	381.5
75	402.4	401.7	398.7	394.9	389.8	383.9	381.3	383.6
100	405.0	404.0	401.0	397.0	391.6	385.6	382.7	385.3
125	407.5	406.2	403.2	398.9	393.1	387.1	384.0	386.8
150	409.8	408.3	405.2	400.7	394.5	388.4	385.1	388.0
175	412.0	410.2	407.0	402.3	395.7	389.5	386.1	389.0
200	414.1	412.0	408.7	403.8	396.8	390.5	386.8	389.9
225	415.9	413.6	410.2	405.1	397.6	391.3	387.5	390.5
250	417.7	415.1	411.6	406.2	398.3	391.9	388.0	391.0
275	419.2	416.4	412.9	407.1	398.9	392.4	388.3	391.4
300	420.7	417.6	413.9	407.9	399.3	392.8	388.6	391.6
325	421.9	418.6	414.8	408.5	399.6	393.0	388.7	391.8
350	423.0	419.5	415.6	409.0	399.8	393.2	388.8	391.8
352							388.8	
371						393.2		
375	423.9	420.3	416.2	409.3	399.8	393.2	388.8	391.8
376					399.8			
400	424.3	420.8	416.7	409.5	399.9	393.2	388.7	391.6
421				409.6				
425	425.2	421.2	416.9	409.6	399.7	393.0	388.5	391.5
450	425.6	421.5	417.1	409.5	399.4	392.8	388.2	391.2
463			417.1					
475	425.8	421.6	417.1	409.3	399.1	392.5	387.8	390.8
484		421.6						
500	425.9	421.6	416.9	409.0	398.8	392.1	387.5	390.4
504	425.9							
525	425.7	421.4	416.7	408.6	398.3	391.6	387.0	389.9
550	425.4	421.2	416.3	408.1	397.7	391.1	386.4	389.3

Electrocapillary data for 0.1 M $NaClO_4$ in methanol + water mixtures (continued)

	γ/mN m^{-1}							
X	0	0.035	0.0735	0.167	0.4171	0.568	0.8993	0.9959
-E/mV								
575	425.0	420.8	415.8	407.5	397.1	390.5	385.8	388.7
600	424.4	420.2	415.2	406.9	396.5	389.8	385.1	388.0
625	423.6	419.6	414.6	406.1	395.7	389.1	384.4	387.3
650	422.8	418.9	413.8	405.3	394.9	388.3	383.6	386.5
675	421.8	418.0	412.9	404.4	394.1	387.5	382.8	385.7
700	420.6	417.1	411.9	403.4	393.2	386.6	381.9	384.8
725	419.4	416.0	410.9	402.4	392.2	385.6	380.9	383.8
750	418.0	414.8	409.8	401.3	391.1	384.6	379.9	382.8
775	416.5	413.6	408.5	400.1	390.0	383.5	378.9	381.7
800	414.8	412.2	407.2	398.9	388.9	382.4	377.7	380.6
825	413.1	410.8	405.8	397.5	387.7	381.2	376.6	379.4
850	411.3	409.3	404.3	396.1	386.4	379.9	375.3	378.2
875	409.3	407.6	402.8	394.7	385.0	378.6	374.0	376.9
900	407.2	405.9	401.1	393.1	383.6	377.2	372.7	375.6
925	405.0	404.1	399.4	391.5	382.2	375.8	371.3	374.2
950	402.8	402.2	397.5	389.8	380.7	374.3	369.9	372.7
975	400.4	400.2	395.6	388.1	379.1	372.8	368.4	371.2
1000	397.9	398.1	393.6	386.2	377.5	371.2	366.8	369.7
1025	395.3	395.9	391.5	384.3	375.8	369.5	365.2	368.1
1050	392.6	393.6	389.3	382.3	374.0	367.8	363.5	366.4
1075	389.8	391.2	387.1	380.2	372.2	366.0	361.8	364.7
1100	386.9	388.7	384.7	378.0	370.3	364.1	360.0	362.9
1125	383.9	386.2	382.2	375.8	368.3	362.2	358.1	361.1
1150	380.8	383.5	379.6	373.4	366.3	360.2	356.2	359.2
1175	377.7	380.8	377.0	371.0	364.2	358.1	354.2	357.2
1200	374.4	377.9	374.2	368.4	362.0	356.0	352.2	355.2
1225	371.0	375.0	371.4	365.8	359.7	353.8	350.1	353.1
1250	367.5	371.9	368.4	363.1	357.4	351.5	347.9	351.0
1275	363.9	368.8	365.4	360.2	354.9	349.1	345.7	348.8
1300	360.2	365.6	362.2	357.3	352.4	346.7	343.4	346.5
1325	356.4	362.2	358.9	354.2	349.8	344.1	341.0	344.2
1350	352.6	358.8	355.6	351.0	347.1	341.5	338.6	341.8
1375	348.6	355.3	352.1	347.7	344.3	338.8	336.0	339.3
1400	344.5	351.6	348.5	344.3	341.4	336.0	333.4	336.7

Electrocapillary data on mercury electrode in 1 M $NaClO_4$ + CH_3OH.
Given are the surface tension in mN m^{-1}, potential in mV vs CE in 1 M NaCl.
T = 25°C

Reference: J. Taraszewska. J. Electroanal. Chem. 49 (1974) 443.

X is mole fraction of methanol

	γ /mN m^{-1}						
X	0	0.036	0.076	0.175	0.44	0.59	0.96
-E/mV							
0	384.5	383.0	382.4	380.0	377.1	411.1	375.8
25	387.9	386.2	385.4	382.9	379.8	408.7	378.1
50	391.0	389.3	388.3	385.6	382.3	406.4	380.2
75	394.0	392.2	391.1	388.2	384.0	404.3	382.1
100	396.9	395.0	393.7	390.6	386.7	402.4	383.8
125	399.6	397.6	396.1	392.9	388.6	400.6	385.4
150	402.2	400.0	398.4	395.0	390.4	399.1	386.8
175	404.7	402.3	400.5	396.9	392.0	397.6	388.0
200	407.0	404.5	402.5	398.7	393.4	396.4	389.0
225	409.1	406.5	404.3	400.3	394.7	395.3	389.9
250	411.1	408.3	406.0	401.7	395.8	394.4	390.7
275	413.0	410.0	407.5	403.0	396.7	393.6	391.2
300	414.6	411.5	408.9	404.1	397.5	393.0	391.7
325	416.2	412.9	410.1	405.1	398.0	392.5	392.0
350	417.6	414.2	411.2	405.9	398.5	392.2	392.2
375	418.8	415.2	412.1	406.5	398.8	392.0	392.3
378							392.3
400	419.9	416.2	412.9	407.0	399.0	392.0	392.2
401						392.0	
415					399.0		
425	420.8	416.9	413.4	407.3	399.0	392.3	392.1
450	421.5	417.5	413.9	407.5	398.9	392.2	391.8
464				407.5			
475	422.0	417.9	414.0	407.5	398.7	392.5	391.6
500	422.4	418.1	414.2	407.4	398.3	392.9	391.1
502			414.2				
521		418.2					
525	422.6	418.2	414.1	407.1	397.9	393.3	390.7
537	422.6						
550	422.5	418.0	413.8	406.7	397.3	393.8	390.0
575	422.3	417.7	413.5	406.1	396.7	394.4	389.4
600	422.0	417.3	413.0	405.5	396.0	395.1	388.7

Electrocapillary data for 1 M $NaClO_4$ in methanol + water mixtures (continued)

	γ /mN m^{-1}						
X	0	0.036	0.076	0.175	0.44	0.59	0.96
-E/mV							
625	421.4	416.7	412.3	404.8	395.2	395.9	387.9
650	420.7	415.9	411.5	404.0	394.3	396.7	387.1
675	419.8	415.0	410.6	403.1	393.4	397.6	386.2
700	418.8	413.9	409.6	402.1	392.4	398.6	385.3
725	417.6	412.7	408.5	401.0	391.3	399.6	384.3
750	416.2	411.4	407.2	399.8	390.2	400.7	383.2
775	414.7	410.0	405.9	398.6	390.0	401.8	382.1
800	413.0	408.4	404.4	397.2	387.7	403.1	380.9
825	411.2	406.7	402.9	395.8	386.4	404.3	379.7
850	409.3	404.8	401.2	394.4	385.0	405.6	378.4
875	407.3	402.9	399.5	392.8	383.6	407.0	377.1
900	405.1	400.9	397.6	391.2	382.0	408.5	375.7
925	402.8	398.7	395.7	389.5	380.5	410.0	374.3
950	400.5	396.5	393.6	387.7	378.8	411.5	372.8
975	398.0	394.1	391.5	385.8	377.1	413.2	371.2
1000	395.3	391.6	389.2	383.8	375.3	414.9	369.6
1025	392.6	389.0	386.9	381.8	373.5	416.6	367.9
1050	389.8	386.3	384,4	379.7	371.6	418.4	366.2
1075	386.9	383.5	381.9	377.4	369.6	420.3	364.4
1100	383.9	380.6	379.2	375.1	367.5	422.2	362.6
1125	380.8	377.6	376.5	372.7	365.4	424.3	360.7
1150	377.6	374.5	373,6	370.2	363.2	426.4	358.7
1175	374.3	368.0	370.6	367.6	360.9	428.5	356.7
1200	370.8	364.6	367.6	364.8	358.5	430.8	354.5
1225	367.3	361.0	364.4	362.0	356.0	433.1	352.4
1250	363.7	357.4	361.1	359.1	353.4	435.5	350.1
1275	360.0	353.7	357.6	356.0	350.8	438.0	347.8
1300	356.1	349.8	354.1	352.8	348.0	440.5	345.5
1325	352.2	345.8	350.5	349.5	345.2	443.2	
1350	348.2	341.7	346.7	346.1	342.2	446.0	
1375	344.0	337.5	342.8	342.5	339.1	448.8	
1400	379.8		338.8	338.8	335.9	351.8	

Electrocapillary data on mercury electrode in 2 M $NaClO_4$ + CH_3OH.
Given are the surface tension in mN m^{-1}, potential in mV vs SCE (NaCl).
T = 25°C

Reference: J. Taraszewska. J. Electroanal. Chem. 49 (1974) 443.

X is mole fraction of methanol

	γ/mN m^{-1}					
X	0	0.038	0.079	0.184	0.49	0.92
-E/mV						
0	381.7	377.3	378.1	376.1	372.8	372.8
25	385.1	380.8	381.4	379.1	375.5	375.3
50	388.3	384.1	384.4	381.9	378.1	377.6
75	391.3	387.2	387.3	384.6	380.6	379.7
100	394.2	390.2	390.0	387.1	382.8	381.7
125	397.0	393.0	392.6	389.5	384.9	383.5
150	399.6	395.7	395.0	391.7	386.8	385.1
175	402.1	398.2	397.3	393.7	388.6	386.6
200	404.4	400.5	399.4	395.6	390.2	387.9
225	406.6	402.7	401.4	397.3	391.6	389.0
250	408.7	404.8	403.2	398.8	392.9	390.0
275	410.6	406.6	404.8	400.2	393.9	390.8
300	412.3	408.3	406.3	401.5	394.9	391.5
325	413.9	409.9	407.7	402.5	395.7	392.0
350	415.4	411.3	408.9	403.5	396.3	392.4
375	416.7	412.5	409.9	404.2	396.8	392.6
400	417.8	413.6	410.8	404.8	397.1	392.7
414						392.8
425	418.8	414.6	411.5	405.3	397.3	392.8
446					397.3	
450	419.6	415.3	412.1	405.6	397.3	392.7
475	420.2	415.9	412.5	405.8	397.3	392.5
489				405.8		
500	420.8	416.3	412.7	405.8	397.0	392.1
525	421.0	416.6	412.8	405.7	396.7	391.8
527			412.8			
549		416.6				
550	421.2	416.6	412.8	405.4	396.3	391.3
575	421.1	416.9	412.5	405.0	395.7	390.6
600	420.9	416.2	412.2	404.4	395.1	390.0

Electrocapillary data for 2 M NaClO$_4$ in methanol + water mixtures (continued)

	γ/mN m^{-1}					
X	0	0.038	0.079	0.184	0.49	0.92
-E/mV						
625	420.5	415.8	411.7	403.8	394.4	389.3
650	419.9	415.2	411.0	403.1	393.6	388.5
675	419.1	414.5	410.3	402.2	392.7	387.6
700	418.2	413.6	409.4	401.3	391.7	386.7
725	417.1	412.5	408.3	400.3	390.7	385.7
750	415.8	411.3	407.2	399.1	389.6	384.6
775	414.4	410.0	406.0	397.9	388.4	382.5
800	412.8	408.5	404.6	396.6	387.2	382.3
825	411.1	406.9	403.1	395.3	385.8	381.1
850	409.2	405.2	401.6	393.8	384.5	379.8
875	407.2	403.3	399.9	392.3	383.0	378.4
900	405.1	401.3	398.1	390.7	381.5	377.0
925	402.8	399.2	396.2	389.0	380.0	375.6
950	400.4	397.0	394.3	387.2	378.3	374.0
975	397.9	394.6	392.2	385.3	376.6	372.4
1000	395.3	392.2	390.0	383.4	374.8	370.8
1025	392.6	389.6	387.7	381.3	373.0	369.1
1050	389.7	386.9	385.3	379.2	371.1	367.3
1075	386.8	384.1	382.8	377.0	369.1	365.5
1100	383.7	381.1	380.2	374.7	367.1	363.6
1125	380.5	378.1	377.4	372.3	364.9	361.6
1150	377.3	374.9	374.6	369.8	362.7	359.6
1175	373.9	371.7	371.6	367.2	360.4	357.5
1200	370.4	368.3	368.6	364.5	358.0	355.4
1225	366.8	364.8	365.4	361.6	355.6	353.1
1250	363.1	361.2	362.1	358.7	353.0	350.8
1275	359.3	357.4	358.7	355.6	350.4	348.4
1300	355.4	353.6	355.1	352.5	347.6	346.0
1325	351.4	349.6	351.5	349.1	344.8	343.4
1350	347.3	345.5	347.7	345.7	341.8	340.8
1375	343.1	341.3	343.8	342.1	338.8	338.1
1400	338.8	336.9	339.7	338.4	335.6	335.3

Electrocapillary data for n-octanol in 0.5 mol l^{-1} H_2SO_4. Given are the interfacial tension in mN m^{-1}, potential in V with respect to the saturated calomel electrode.

Reference: R.J. Meakins. J. Appl. Chem. 15 (1965) 416.

	γ/mN m^{-1}				
			n-Octanol		
c/mol l^{-1}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	3×10^{-3}
-E/V					
0.0	388	385	385	385	384
0.1	400	402	397	393	386
0.2	412	409	402	395	387
0.3	417	410	402	393	387
0.4	419	410	402	394	388
0.5	417	409	401	393	387
0.6	416	407	401	391	386
0.7	412	405	397	391	385
0.8	408	402	395	388	381
0.9	400	397	391	385	376
1.0	378	379	382	375	373
1.1					369
1.2					
1.3					
1.4					

Electrocapillary data on mercury electrode in 0.1 M Na_2SO_4 at 25°C. Electrosorption of 2-butanol. Given are the interfacial tension in mN m^{-1}, -potential in V_{vs} corning NAS 11-18 glass sodium ion electrode

Reference: H. Nakadomari, D.M. Mohilner and P.R. Mohilner.
J. Phys. Chem. 80 (1976) 1761.

	γ /mN m^{-1}								
c/mol l^{-1}	0.000	0.013	0.010	0.020	0.025	0.031	0.037	0.046	0.058
E/V									
-1.300	354.1	353.9	353.9	354.1	354.1	353.7	353.7	353.6	353.7
-1.275	357.7	357.6	357.7	357.8	357.9	357.6	357.6	357.5	357.6
-1.250	361.7	361.4	361.4	361.6	361.6	361.3	361.3	361.2	361.2
-1.225	365.3	365.0	365.1	365.3	365.2	364.8	364.8	364.8	364.8
-1.200	368.8	368.5	368.6	368.7	368.6	368.3	368.3	368.2	368.2
-1.175	372.2	372.1	372.2	372.2	372.0	371.8	371.7	371.6	371.5
-1.150	375.7	375.4	375.4	375.6	375.4	375.1	375.1	374.9	374.8
-1.125	378.9	378.6	378.7	378.8	378.5	378.2	378.1	378.0	378.0
-1.100	381.9	381.9	381.9	381.9	381.6	381.3	381.2	381.1	381.1
-1.075	385.1	384.8	384.9	384.9	384.6	384.3	384.1	384.1	383.9
-1.050	388.1	387.8	387.8	387.8	387.5	387.1	387.0	386.9	386.8
-1.025	390.9	390.6	390.6	390.6	390.3	389.9	389.7	389.6	389.5
-1.000	393.7	393.3	393.2	393.3	392.9	392.6	392.4	392.1	391.9
-0.975	396.3	395.9	395.7	395.9	395.5	395.1	394.9	394.7	394.4
-0.950	398.8	398.3	398.3	398.3	397.9	397.6	397.2	397.1	396.6
-0.925	401.4	400.8	400.8	400.7	400.2	399.8	399.6	399.3	398.8
-0.900	403.7	403.2	403.1	403.0	402.5	402.1	401.8	401.4	400.8
-0.875	406.0	405.4	405.3	405.1	404.6	404.3	403.8	403.4	402.6
-0.850	408.1	407.4	407.4	407.2	406.7	406.2	405.8	405.2	404.5
-0.825	410.1	409.4	409.3	409.1	408.6	408.1	407.6	407.0	406.1
-0.800	412.1	411.3	411.2	411.0	410.4	409.8	409.4	408.5	407.5
-0.775	413.7	413.0	412.9	412.6	412.0	411.5	410.9	410.0	409.0
-0.750	415.6	414.6	414.4	414.2	413.5	413.0	412.3	411.4	410.1
-0.725	417.1	416.2	416.0	415.7	415.0	414.4	413.7	412.6	411.2
-0.700	418.6	417.6	417.3	417.0	416.4	415.7	414.9	413.8	412.3
-0.675	420.0	418.9	418.7	418.4	417.6	416.9	416.1	414.8	413.3
-0.650	421.2	420.1	419.9	419.5	418.8	418.1	417.1	415.8	413.9
-0.625	422.3	421.2	420.8	420.6	419.7	419.1	418.0	416.7	414.9
-0.600	423.2	422.2	421.9	421.5	420.7	419.9	418.9	417.5	415.6
-0.575	424.1	423.0	422.7	422.4	421.5	420.7	419.7	418.2	416.4
-0.550	424.8	423.7	423.5	423.1	422.3	421.5	420.4	418.9	416.9

Electrocapillary data for 2-butanol in 0.1 M Na_2SO_4 (continued)

	γ/ mN m^{-1}								
c/mol l^{-1}	0.000	0.013	0.010	0.020	0.025	0.031	0.037	0.046	0.058
E/V									
-0.525	425.4	424.4	424.0	423.7	422.8	422.1	421.0	419.5	417.5
-0.500	425.8	424.7	424.6	424.2	423.4	422.5	421.4	420.0	418.2
-0.475	426.2	425.2	424.8	424.5	423.7	423.0	421.9	420.5	418.6
-0.450	426.4	425.3	425.1	424.7	424.0	423.3	422.3	420.9	419.1
-0.425	426.4	425.5	425.2	424.9	424.1	423.5	422.5	421.2	419.4
-0.400	426.2	425.4	425.0	424.9	424.2	423.6	422.8	421.5	419.8
-0.375	425.9	425.1	425.0	424.6	424.0	423.5	422.8	421.6	420.1
-0.350	425.5	424.8	424.5	424.4	423.7	423.2	422.6	421.6	420.2
-0.325	424.9	424.1	424.0	423.8	423.2	422.9	422.3	421.6	420.3
-0.300	424.0	423.4	423.4	423.1	422.5	422.3	421.7	421.2	420.2
-0.275	423.0	422.5	422.4	422.2	421.6	421.4	421.0	420.7	419.8
-0.250	421.8	421.3	421.2	421.0	420.5	420.3	420.0	419.7	419.3
-0.225	420.3	420.0	419.9	419.7	419.2	419.1	418.8	418.6	418.3
-0.200	418.6	418.4	418.2	418.2	417.7	417.5	417.3	417.2	417.0
-0.175	416.8	416.5	416.4	416.3	415.8	415.8	415.6	415.5	415.3
-0.150	414.5	414.3	414.4	414.2	413.7	413.7	413.5	413.5	413.3
-0.125	412.2	412.1	412.0	411.9	411.3	411.3	411.2	411.2	411.3
-0.100	409.6	409.5	409.5	409.3	408.8	408.8	408.8	408.8	408.7
-0.075	406.8	406.6	406.7	406.6	406.0	406.1	406.0	406.0	406.0
-0.050	403.7	403.5	403.6	403.4	402.9	403.0	402.9	403.1	403.0
-0.025	400.4	400.3	400.4	400.2	399.6	399.7	399.7	399.8	399.7
0.000	396.8	396.8	396.7	396.6	396.0	396.2	396.2	396.3	396.2
0.025	393.0	393.1	392.9	392.9	392.4	392.5	392.5	392.6	392.5
0.050	389.0	389.1	389.0	388.9	388.4	388.5	388.5	388.6	388.5
0.075	384.8	384.9	384.7	384.7	384.2	384.2	384.2	384.4	384.3
0.100	380.3	380.3	380.1	380.1	379.6	379.7	379.8	379.8	379.8
0.125	375.5	375.4	375.3	375.4	374.8	375.0	375.1	375.1	374.9
0.150	370.5	370.6	370.4	370.3	369.8	370.0	370.1	370.2	370.0
0.175	365.2	365.3	365.1	365.0	364.6	364.7	364.8	364.9	364.8
0.200	359.7	359.6	359.5	359.5	359.1	359.2	359.4	359.4	359.3

Electrocapillary data for 2-butanol in 0.1 M Na_2SO_4 (continued)

	γ /mN m^{-1}								
c/mol l^{-1}	0.070	0.086	0.105	0.125	0.147	0.175	0.210	0.255	0.300
E/V									
-1.300	353.6	353.6	353.6	353.4	353.5	353.4	353.7	353.2	353.1
-1.275	357.4	357.4	357.5	357.2	357.3	357.3	357.3	356.9	356.7
-1.250	361.2	361.1	361.1	360.9	360.8	360.8	360.9	360.3	360.1
-1.225	364.6	364.6	364.6	364.4	364.2	364.2	364.2	363.6	363.3
-1.200	368.1	369.1	368.1	367.8	367.6	367.5	367.4	366.8	366.3
-1.175	371.4	371.4	371.2	371.0	370.8	370.6	370.5	369.8	369.0
-1.150	374.7	374.5	374.4	374.1	373.9	373.7	373.3	372.5	371.6
-1.125	377.7	377.6	377.5	377.1	376.8	376.5	375.9	374.8	373.8
-1.100	380.8	380.5	380.3	380.0	379.6	379.0	378.5	377.0	375.8
-1.075	383.6	383.4	383.2	382.6	382.1	381.5	380.6	378.9	377.6
-1.050	386.4	386.1	385.6	385.1	384.5	383.7	382.6	380.6	379.2
-1.025	389.0	388.7	388.1	387.5	386.6	385.6	384.3	382.3	380.7
-1.000	391.3	391.1	390.4	389.6	388.6	387.3	385.7	383.6	382.1
-0.975	393.8	393.3	392.4	391.5	390.2	388.9	387.4	385.0	383.4
-0.950	395.9	395.4	394.3	393.2	391.8	390.3	388.6	386.2	384.6
-0.925	398.1	397.3	396.0	394.7	393.3	391.7	389.9	387.5	385.8
-0.900	399.9	399.2	397.6	396.2	394.7	393.0	391.1	388.7	387.0
-0.875	401.6	400.7	399.0	397.5	395.8	394.1	392.3	389.7	388.0
-0.850	403.3	402.1	400.3	398.8	396.9	395.2	393.4	390.7	389.0
-0.825	404.7	403.4	401.4	400.0	399.0	396.3	394.3	391.7	390.0
-0.800	406.0	404.7	402.5	401.0	399.1	397.4	395.3	392.7	390.9
-0.775	407.3	405.8	403.6	401.9	400.0	398.2	396.2	393.6	391.9
-0.750	408.4	406.8	404.5	402.9	401.1	399.2	397.1	394.5	392.7
-0.725	409.4	407.8	405.5	404.0	401.9	400.1	398.0	395.4	393.6
-0.700	410.4	408.7	406.4	404.7	402.7	400.9	399.9	396.0	394.4
-0.675	411.2	409.5	407.1	405.5	403.5	401.7	399.6	396.9	395.1
-0.650	412.0	410.4	407.9	406.2	404.3	402.5	400.4	397.6	395.9
-0.625	412.8	411.2	408.7	407.1	405.1	403.1	401.0	398.4	396.5
-0.600	413.6	411.9	409.4	407.8	405.8	403.9	401.7	399.0	397.3
-0.575	414.2	412.5	410.1	408.5	406.4	404.5	402.3	399.6	397.9
-0.550	414.9	413.1	410.7	409.1	407.0	405.1	403.0	400.3	398.5

Electrocapillary data for 2-butanol in 0.1 M Na_2SO_4. (continued)

	γ/mN m^{-1}								
c/mol l^{-1}	0.070	0.086	0.105	0.125	0.147	0.175	0.210	0.255	0.300
E/V									
-0.525	415.5	413.8	411.3	409.8	407.7	405.8	403.6	400.9	399.1
-0.500	416.0	414.4	411.8	410.3	408.2	406.4	404.1	401.5	399.6
-0.475	416.4	414.9	412.3	410.7	408.7	406.8	404.6	402.0	400.5
-0.450	416.9	415.4	412.9	411.3	409.3	407.4	405.1	402.5	400.6
-0.425	417.3	415.8	413.3	411.6	409.8	407.8	405.7	403.0	401.1
-0.400	417.7	416.2	413.6	412.1	410.1	408.3	406.1	403.3	401.5
-0.375	418.0	416.5	414.0	412.5	410.5	408.6	406.4	403.7	402.0
-0.350	418.2	416.8	414.3	412.8	410.8	409.0	406.9	404.1	402.2
-0.325	418.4	417.0	414.6	413.1	411.1	409.3	407.3	404.4	402.6
-0.300	418.5	417.2	414.8	413.3	411.4	409.6	407.4	404.8	402.9
-0.275	418.5	417.2	415.0	413.5	411.6	409.8	407.6	405.0	403.2
-0.250	418.2	417.1	415.0	413.6	411.7	409.9	407.7	405.2	403.4
-0.225	417.5	416.7	415.0	413.5	411.8	410.1	407.9	405.4	403.5
-0.200	416.4	415.9	414.6	413.5	411.8	410.1	407.9	405.5	403.6
-0.175	415.0	414.6	413.9	413.0	411.6	410.1	408.0	405.5	403.7
-0.150	413.0	412.9	412.4	412.0	411.0	409.9	407.8	405.6	403.8
-0.125	410.9	410.8	410.5	410.2	409.8	409.2	407.6	405.5	403.6
-0.100	408.5	408.4	408.2	408.1	407.7	407.5	407.0	405.1	403.5
-0.075	405.8	405.7	405.6	405.6	405.2	405.2	404.9	404.1	403.0
-0.050	402.8	402.8	402.7	402.6	402.4	402.4	402.5	401.8	401.3
-0.025	399.6	399.6	399.4	399.4	399.2	399.3	399.4	398.9	399.6
0.000	396.2	396.1	396.0	396.1	395.8	395.9	395.9	395.6	395.4
0.025	392.4	392.4	392.3	392.4	392.2	392.3	392.3	392.2	392.0
0.050	388.5	388.5	388.4	388.5	388.3	388.5	388.4	388.2	388.1
0.075	384.3	384.3	384.2	384.3	384.1	384.3	384.3	384.1	384.0
0.100	379.7	379.7	379.6	379.7	379.6	379.7	379.8	379.6	379.6
0.125	375.0	375.0	374.9	375.1	374.9	375.0	375.1	375.0	374.8
0.150	370.0	370.0	369.9	370.1	369.9	370.1	370.1	370.0	370.0
0.175	364.8	364.8	364.7	364.8	364.7	364.8	364.8	364.8	364.2
0.200	359.3	359.2	359.2	359.3	359.2	359.3	359.3	359.2	359.2

. Electrocapillary data for 2-butanol in 0.1 M Na_2SO_4 (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.420	0.500	0.600	0.700	0.840	1.000	1.200	1.400
E/V								
-1.300	352.8	352.7	352.1	351.4	350.7	349.7	348.8	347.2
-1.275	356.2	355.8	355.2	354.4	353.3	352.3	350.9	349.3
-1.250	359.5	358.8	358.0	357.0	355.6	354.3	352.8	351.1
-1.225	362.3	361.5	360.4	359.2	357.7	356.2	354.4	352.9
-1.200	365.0	363.8	362.6	361.2	359.6	358.0	356.3	354.5
-1.175	367.3	366.0	364.5	363.0	361.3	359.6	357.9	356.0
-1.150	369.4	367.9	366.2	364.6	362.9	361.2	359.4	357.5
-1.125	371.2	369.6	367.8	366.1	364.4	362.6	360.8	358.9
-1.100	372.9	371.1	369.4	367.6	365.8	364.1	362.2	360.3
-1.075	374.4	372.6	370.8	369.0	367.2	365.4	363.5	361.6
-1.050	375.9	374.1	372.1	370.3	368.5	366.6	364.7	362.8
-1.025	377.3	375.4	373.4	371.6	369.8	367.9	366.0	364.0
-1.000	378.5	376.6	374.7	372.9	371.0	369.1	367.1	365.3
-0.975	379.8	377.9	375.9	374.0	372.2	370.2	368.3	366.4
-0.950	380.9	379.0	377.0	375.2	373.3	371.4	369.4	367.4
-0.925	382.1	380.1	378.1	376.3	374.4	372.4	370.5	368.5
-0.900	383.3	381.1	379.1	377.3	375.5	373.5	371.5	369.5
-0.875	384.3	382.4	380.2	378.4	376.5	374.3	372.5	370.6
-0.850	385.3	383.2	381.3	379.4	377.4	375.5	373.5	371.5
-0.825	386.4	384.3	382.2	380.3	378.4	376.4	374.4	372.5
-0.800	387.3	385.3	383.1	381.2	379.3	377.4	375.3	373.4
-0.775	388.1	386.0	384.0	382.1	380.2	378.2	376.1	374.2
-0.750	389.0	387.1	384.9	383.0	381.1	379.1	377.0	375.0
-0.725	389.9	387.7	385.8	383.8	381.9	379.9	377.8	375.8
-0.700	390.7	388.6	386.5	384.6	382.7	380.7	378.6	376.6
-0.675	391.4	389.4	387.3	385.4	383.5	381.4	379.3	377.3
-0.650	392.2	390.1	388.1	386.1	384.2	382.2	380.0	378.1
-0.625	392.9	390.8	388.7	386.8	384.9	382.9	380.7	378.8
-0.600	393.6	391.4	389.5	387.5	385.6	383.6	381.4	379.4
-0.575	394.3	392.1	390.1	388.1	386.2	384.2	382.0	380.1
-0.550	394.8	392.7	390.7	388.7	386.9	384.9	382.6	380.7

Electrocapillary data for 2-butanol in 0.1 M Na_2SO_4 (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.420	0.500	0.600	0.700	0.840	1.000	1.200	1.400
E/V								
-0.525	395.5	393.3	391.3	389.4	387.4	385.4	383.2	381.3
-0.500	396.0	393.8	391.9	389.9	388.0	385.8	383.8	381.8
-0.475	396.5	394.4	392.4	390.5	388.5	386.5	384.3	382.3
-0.450	397.0	394.9	392.9	391.0	389.0	387.0	384.8	382.8
-0.425	397.5	395.4	393.3	391.5	389.5	387.4	385.2	383.2
-0.400	397.9	395.8	393.7	391.9	389.9	387.9	385.6	383.8
-0.375	398.3	396.1	394.2	392.4	390.3	388.4	386.1	384.0
-0.350	398.7	396.5	394.6	392.7	390.7	388.6	386.4	384.5
-0.325	399.1	396.9	394.9	393.1	391.0	389.1	386.8	384.8
-0.300	399.4	397.1	395.3	393.4	391.4	389.4	387.0	385.1
-0.275	399.7	397.6	395.5	393.7	391.7	389.6	387.3	385.3
-0.250	399.9	397.8	395.7	394.0	391.9	389.7	387.6	385.6
-0.225	400.0	398.0	395.9	394.1	392.1	390.1	387.7	385.8
-0.200	400.2	398.0	396.1	394.3	392.2	390.2	387.9	386.0
-0.175	400.2	398.2	396.2	394.4	392.4	390.3	388.0	386.1
-0.150	400.3	398.3	396.4	394.5	392.5	390.4	388.1	386.2
-0.125	400.3	398.3	396.3	394.5	392.5	390.5	388.1	396.3
-0.100	400.2	398.3	396.3	394.5	392.4	390.5	388.2	386.2
-0.075	400.0	398.1	396.2	394.4	392.4	390.5	388.1	386.2
-0.050	399.5	397.8	396.0	394.2	392.3	390.3	388.0	386.0
-0.025	398.1	397.1	395.6	393.9	392.0	390.2	387.8	385.9
0.000	395.3	394.9	394.3	393.3	391.6	389.7	387.5	385.7
0.025	391.9	391.7	391.5	391.3	390.5	389.0	387.1	385.3
0.050	388.2	388.1	387.9	388.0	387.7	387.3	386.2	384.7
0.075	384.2	384.1	383.9	384.1	384.1	384.1	383.9	383.5
0.100	379.7	379.7	379.6	379.7	379.7	379.9	380.2	380.7
0.125	375.1	375.1	375.0	375.3	375.3	375.5	375.9	376.6
0.150	370.0	370.2	370.1	370.5	370.5	370.8	371.4	372.3
0.175	364.9	364.9	364.9	365.3	365.4	365.8	366.3	367.5
0.200	359.3	359.5	359.4	359.8	360.1	360.4	361.1	362.3

Electrocapillary data on mercury electrode in 0.1 M Na_2SO_4 at 25°C. Adsorption of t-butanol - given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: A. de Battisti, A.B. Abd el Nabey and S. Trasatti.
J. Chem. Soc. Farad. Trans. 1, 72 (1976) 2076.

	γ/mN m^{-1}							
c/mol l^{-1}	0	0.0097	0.0135	0.0189	0.0216	0.0270	0.0324	0.0432
-E/V								
0.000	396.1	395.5	395.6	394.9	395.4	396.3	396.1	395.6
0.050	402.5	401.5	401.7	402.7	402.8	402.2	402.2	402.9
0.100	408.3	407.6	408.0	408.5	408.4	408.6	408.6	408.5
0.150	413.5	413.1	413.6	413.1	412.8	413.5	414.0	414.0
0.200	417.3	417.0	417.1	417.4	417.5	418.1	417.8	416.8
0.225					419.4		419.8	419.7
0.250	420.7	420.9	420.8	421.1	421.0	421.4	421.6	421.1
0.275		421.8	422.0	422.1	422.1	422.6	422.6	421.2
0.300	423.3	422.9	423.3	423.6	423.2	423.6	423.7	423.0
0.325		423.9	423.7	424.4	424.1	424.5	424.3	424.1
0.350	425.1	424.6	424.6	425.5	424.7	425.1	424.9	424.9
0.375		425.2	425.0	425.7	425.1	425.6	425.4	425.0
0.400	425.9	425.4	425.6		425.5	425.8	425.8	425.4
0.425		425.7	425.7	425.8	425.8	426.2	425.9	425.6
0.450	426.2	425.9	425.9	426.0	425.6	425.9	425.8	425.5
0.500	426.1	425.6	425.3	425.6	425.2	425.3	425.3	424.9
0.550	425.2	424.5	424.3	424.4	424.3	423.9	424.4	423.5
0.600	423.4	422.7	423.0	423.0	422.8	423.0	422.5	421.4
0.650	421.4	420.9	420.8	420.9	420.8	420.7	420.9	419.5
0.700	418.9	418.2	418.3	418.4	418.3	418.2	418.2	417.1
0.750	415.7	415.1	415.2	415.1	415.1	415.3	415.5	414.4
0.800	412.2	411.8	412.0	412.3	411.8	411.7	411.3	411.3
0.850	408.5	408.3	408.0	408.1	407.8	407.9	408.0	407.0
0.900	404.3	403.8	403.5	403.6	403.6	403.7	403.8	403.1
0.950	399.4	399.1	399.0	399.0	398.7	399.0	398.9	398.8
1.000	394.4	393.5	394.1	394.2	393.9	393.8	393.8	393.6
1.050	388.8	388.3	388.5	388.5	388.3	388.4	388.5	388.0
1.100	382.8	382.1	382.5	382.9	382.4	382.5	382.5	382.0
1.150	376.5	376.7	376.1	376.0	375.9	376.1	376.1	376.0
1.200	369.7	369.0	369.1	369.6	369.3	369.5	369.6	369.4
1.250	362.5	362.2	362.4	362.8	362.1	362.0	362.6	362.4
1.300	355.2	354.7	354.6	354.6	354.4	354.5	354.8	355.1
1.350	347.3	346.7	346.6	346.4	346.4	346.7	347.0	346.3
1.400	338.9	338.8	338.1	338.9	337.9	338.4	338.7	338.3
1.450	330.0	329.5	329.4	329.8	329.3	328.9	330.2	329.8
1.500	320.9	320.3	320.3	320.8	320.5	320.2	320.9	320.7
1.550	311.3	310.9	310.6	310.9	310.7	310.8	311.2	310.8
1.600	301.4	300.3	300.3	300.6	300.1	301.0	300.7	300.0

Electrocapillary data for t-butanol in 0.1 M Na_2SO_4 (côntinued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.0540	0.0605	0.0702	0.0756	0.0809	0.0863	0.0917	0.0998
-E/V								
0.000	395.8	396.4	395.6	395.9	395.3	396.5	395.2	396.6
0.050	402.4	402.8	402.4	402.3	402.1	402.8	401.7	402.6
0.100	408.5	409.2	408.4	408.4	407.4	408.8	408.0	409.0
0.150	413.4	414.0	413.7	413.5	413.4	414.3	412.9	413.7
0.200	417.8	418.0	417.9	418.0	417.4	418.4	417.5	417.8
0.225	419.7		419.3	419.6	418.9	419.5	419.2	419.3
0.250	421.1	421.1	420.7	420.9	420.6	420.6	420.2	420.7
0.275	422.3	422.5	422.0	422.0	421.3	422.0	421.4	421.6
0.300	423.5	423.3	422.7	423.1	422.2	422.8	422.1	422.2
0.325	424.2	423.9	423.5	423.7	422.7	423.1	422.8	422.4
0.350	424.5	424.6	424.0	424.0	423.2	423.5	423.0	422.5
0.375	425.1	424.8	424.3	424.2	423.4	423.7	422.8	422.8
0.400	425.2	424.9	424.5	424.1	423.6	423.7	422.8	422.5
0.425	425.2	425.0	424.4	424.1	423.3	423.1	422.2	422.1
0.450	425.2	424.9	424.2	423.7	423.0	422.9	421.4	421.4
0.500	424.4	424.2	423.3	422.7	421.7	422.0	421.1	420.5
0.550	423.3	423.3	422.0	421.5	420.6	420.9	419.6	419.3
0.600	421.8	421.2	420.6	420.1	419.0	419.1	418.2	417.7
0.650	419.6	419.2	418.7	418.0	417.4	417.2	416.5	415.9
0.700	417.2	417.1	416.2	416.0	415.2	415.3	414.9	414.0
0.750	414.5	414.4	413.7	413.6	412.8	412.6	412.0	411.8
0.800	411.3	411.2	410.7	410.6	409.9	410.1	409.4	409.3
0.850	407.8	507.5	407.3	407.1	406.4	406.7	406.2	405.9
0.900	403.2	403.2	402.9	402.6	402.2	402.6	402.3	402.4
0.950	398.4	398.7	398.4	398.3	397.9	398.4	397.9	398.1
1.000	393.5	393.9	393.5	393.7	393.2	393.8	393.2	393.4
1.050	388.3	388.6	388.2	388.1	387.5	388.1	387.4	387.9
1.100	382.3	382.7	382.4	382.5	381.7	382.4	382.0	382.0
1.150	375.8	376.1	376.0	376.2	375.8	376.2	375.7	375.9
1.200	369.3	369.7	369.4	369.5	369.0	369.4	369.5	369.3
1.250	362.5	362.3	362.2	362.2	361.9	362.4	362.1	362.4
1.300	354.9	355.0	355.1	354.9	354.3	355.0	354.7	354.8
1.350	346.6	346.8	347.0	347.0	346.1	346.9	346.5	346.9
1.400	338.4	338.3	338.7	338.9	338.3	338.9	338.4	338.6
1.450	329.8	329.9	329.7	330.2	329.7	329.9	329.5	330.1
1.500	320.8	320.3	320.6	321.1	320.6	320.9	320.2	320.8
1.550	310.2	310.8	310.8	311.0	310.3	311.6	310.7	311.2
1.600	301.2	300.7	301.2	300.9	300.5	301.0	300.6	301.0

Electrocapillary data for t-butanol in 0.1 M Na_2SO_4 (continued)

	γ/mN m^{-1}						
c/mol l^{-1}	0.1160	0.1214	0.1349	0.1484	0.1651	0.1754	0.1889
-E/V							
0.000	395.5	395.0	395.4	394.4	396.0	395.7	394.8
0.050	401.9	400.8	403.1	402.0	402.4	400.8	401.0
0.100	408.0	406.8	407.8	408.0	408.6	407.6	406.5
0.150	413.1	412.9	413.0	412.9	413.1	412.3	410.8
0.200	416.9	416.8	416.9	416.2	416.3	415.9	413.8
0.225	418.4	418.2	418.3	417.6	417.3	416.4	415.3
0.250	419.7	419.5	418.5	418.7	417.8	416.8	416.8
0.275	420.6	420.3	419.9	418.3	418.3	417.2	417.1
0.300	420.9	420.5	419.5	419.0	418.8	417.4	416.9
0.325	421.1	420.5	419.5	419.1	418.6	417.0	416.6
0.350	421.1	420.6	419.4	418.7	418.2	416.9	416.1
0.375	420.6	420.4	419.1	418.3	418.0	416.6	415.9
0.400	420.2	420.3	418.8	417.9	417.5	416.4	415.2
0.425	419.9	419.5	418.6	417.7			
0.450	419.4	419.1	418.1	417.2	416.8	415.3	414.5
0.500	418.4	418.2	417.0	415.9	415.7	414.0	413.3
0.550	417.3	416.9	416.0	414.7	414.4	412.7	411.8
0.600	415.9	415.5	414.2	413.6	412.9	411.3	410.5
0.650	414.0	413.7	412.5	411.6	411.3	409.8	409.0
0.700	412.1	411.9	410.7	410.0	409.5	408.2	407.6
0.750	410.1	409.7	409.1	407.8	407.6	406.3	405.7
0.800	407.8	407.0	406.6	405.8	405.4	404.3	403.5
0.850	405.0	404.4	403.8	402.9	402.9	401.5	400.7
0.900	401.4	400.3	400.6	399.8	400.1	398.7	398.4
0.950	397.2	397.0	396.8	395.8	396.5	395.2	394.7
1.000	392.3	391.8	392.3	391.9	392.4	391.0	390.6
1.050	387.3	387.3	387.1	386.9	387.4	386.2	386.5
1.100	381.7	381.5	381.2	380.9	382.2	380.5	381.2
1.150	375.7	375.7	375.2	375.0	375.9	374.6	375.1
1.200	369.1	368.6	368.9	368.9	369.5	368.3	368.7
1.250	361.8	360.9	361.3	361.5	362.3	361.6	361.8
1.300	354.5	353.9	354.6	354.4	355.1	354.2	354.2
1.350	346.2	346.2	346.3	346.3	346.9	346.4	346.2
1.400	338.1	337.8	338.3	337.8	338.7	337.9	338.3
1.450	329.3	329.2	329.5	329.6	330.0	329.3	329.2
1.500	320.6	320.2	320.4	320.0	320.9	320.2	320.3
1.550	310.6	310.6	310.8	310.2	311.2	310.5	310.7
1.600	300.9	300.2	300.1	299.8	300.9	300.2	300.6

Electrocapillary data for t-butanol in 0.1 M Na_2SO_4 (continued)

	$\gamma/mN\ m^{-1}$						
$c/mol\ l^{-1}$	0.2159	0.2698	0.3238	0.4047	0.4857	0.5307	0.6176
$-E/V$							
0.000	394.8	394.1	394.1	394.6	394.3	393.8	
0.050	400.9	399.8	400.3	400.7	400.8	400.0	400.7
0.100	406.2	405.0	404.9	405.6	405.4	403.3	404.0
0.150	410.5	409.0	408.1	408.5	406.8	405.2	404.5
0.200	411.7	411.3	410.2	409.7	407.1	405.7	404.2
0.225	413.0	411.8	411.0	409.6	407.0	405.7	404.1
0.250	414.1	412.7	411.1	409.3	406.8	405.4	403.9
0.275	414.7	412.9	411.3	409.2	406.5	405.1	403.7
0.300	414.9		411.0	408.8	406.1	404.9	403.3
0.325	414.9	412.4	410.8	408.3	405.9	404.5	403.0
0.350	415.0		410.5	407.8	405.5	404.2	402.5
0.375	414.6	411.8	410.2	407.4	405.2	403.8	402.1
0.400	414.5		409.5	407.2	404.4	403.4	401.9
0.425	414.1	410.6		406.9			401.2
0.450	413.0	410.6	408.5	406.1	403.5	402.5	400.6
0.500	411.7	409.4	407.1	404.7	402.3	401.4	399.8
0.550	410.8	408.0	406.1	403.8	401.3	400.0	398.4
0.600	409.3	406.7	404.7	402.6	400.0	398.7	396.9
0.650	407.7	405.4	403.2	400.6	398.3	397.4	395.5
0.700	406.1	403.7	401.8	399.0	396.9	395.8	393.9
0.750	404.1	401.7	399.8	397.3	394.8	394.6	392.0
0.800	402.1	399.8	397.8	395.5	392.7	392.3	390.4
0.850	399.7	397.4	395.7	393.7	390.7	390.2	388.5
0.900	397.0	395.0	393.2	391.3	388.6	388.0	386.1
0.950	393.8	392.2	390.5	388.6	386.2	385.3	383.9
1.000	390.0	388.7	387.5	385.8	383.8	383.1	381.6
1.050	385.6	384.2	383.4	382.8	380.6	380.1	378.5
1.100	380.4	379.8	379.6	378.7	377.3	376.7	375.1
1.150	374.7	374.7	373.5	373.4	372.9	371.6	371.5
1.200	368.4	368.4	367.9	367.8	366.9	366.9	366.6
1.250	361.5	361.5	361.0	361.2	360.4	360.6	360.7
1.300	354.3	354.3	353.4	354.1	353.6	354.0	354.0
1.350	346.3	347.2	345.9	346.5	345.9	346.2	346.3
1.400	338.0	338.4	338.1	338.3	337.9	338.3	338.1
1.450	329.4	329.5	329.2	329.8	329.4	329.4	329.6
1.500	320.1	320.4	319.9	320.8	320.5	320.6	320.5
1.550	310.8	310.9	310.4	311.1	310.5	311.1	311.2
1.600	301.4	300.8	300.2	301.0	300.6	300.6	300.6

Electrocapillary data on mercury in aqueous solutions of
0.5 M Na_2SO_4 + phenol of varying concentrations.
Potentials measured with respect to a normal calomel electrode.
Temperature 20°C. Given is the interfacial tension in mN. m^{-1}.
Reference: J.A.V. Butler and C. Ockrent, J. Phys. Chem., 34,
2841 (1930).

conc. of phenol in mole l^{-1}	0	0.005	0.01	0.02	0.05	0.1
$\frac{E}{\text{volts}}$						
+0.1	366.6	366.7	364.4	362.1	356.9	353.4
0.0	388.3	383.1	379.2	376.3	369.2	363.4
-0.1	404.3	392.5	388.3	385.0	376.6	370.7
-0.2	415.2	400.0	396.2	392.3	383.6	377.1
-0.3	422.7	405.8	402.1	397.1	389.0	382.2
-0.4	426.2	410.0	406.1	400.7	393.0	386.4
-0.5	427.2	412.1	409.1	403.6	395.8	388.8
-0.6	425.4	413.9	409.9	405.0	397.1	389.4
-0.7	421.0	413.6	411.1	405.4	396.6	389.7
-0.8	414.3	411.1	409.8	404.5	396.1	388.8
-0.9	406.0	404.5	403.7	401.1	394.0	386.5
-0.10	396.0	395.3	394.9	394.0	390.0	383.3
-0.11	384.0	378.9	384.0	383.4	381.4	377.0
-0.12	367.5	370.7	370.8	370.7	369.9	368.8
-0.13	356.4	355.8	355.9	355.4	355.9	355.1
-0.14	339.2	339.2	339.5	339.1	339.2	339.1

Electrocapillary data on mercury in aqueous solutions of 1 mole.l^{-1} KNO_3 and various concentrations of quinoline. Potentials measured with respect to a normal calomel electrode. Temperature 19 ± 1 °C. Given is the interfacial tension in mN.m^{-1}.

Reference: S. Bordi, G. Papeschi, J. Electroanal. Chem. and Interfacial Electrochem., 20 (1969), 297.

Conc. of quinoline in mmole.l^{-1}	0.1	0.3	0.5	1.0	1.54	3	4	6	9	15	20	30
E volts												
0.0	374.0	372.7	371.5	369.5	367.6	365.6	364.5	362.5	359.9	356.5	353.7	346.3
	380.8	378.7	377.9	375.9	373.9	371.4	370.2	368.0	365.9	362.4	358.7	351.3
0.1	387.2	385.0	384.3	381.2	379.3	376.6	375.4	373.0	370.9	367.3	363.9	356.0
	392.7	390.0	388.6	386.1	384.3	381.2	379.5	377.4	375.1	375.1	368.2	359.5
0.2	398.2	395.0	393.5	390.8	388.3	385.0	383.3	381.1	378.8	374.5	370.9	362.7
	402.4	399.4	397.4	394.0	392.4	388.7	387.0	384.3	381.6	376.6	372.8	364.9
0.3	406.2	402.4	400.3	397.4	395.2	391.4	389.6	386.7	383.8	377.4	373.6	365.8
	409.4	405.5	403.2	400.1	397.6	393.6	391.7	388.3	384.4	376.4	372.7	365.3
0.4	412.1	407.5	405.1	402.0	399.2	395.5	393.0	389.1	382.9	375.0	371.5	364.4
	414.0	408.8	406.8	403.0	400.7	395.9	393.6	387.2	381.1	373.2	369.8	363.0
0.5	414.8	409.8	407.7	403.7	401.4	395.1	391.1	384.8	378.9	371.1	367.6	361.0
	415.3	410.8	408.2	404.0	401.1	393.0	388.5	382.4	376.5	368.7	365.2	358.8
0.6	415.0	410.1	407.6	403.4	400.4	390.8	385.7	379.8	373.8	365.9	326.3	356.2
	414.2	409.3	406.9	402.4	398.8	387.8	382.9	377.1	371.0	363.0	359.5	353.4
0.7	413.1	408.3	405.9	401.2	396.5	384.8	379.7	374.0	368.0	360.1	356.4	350.3
	411.4	406.9	404.5	399.2	393.4	381.8	376.8	370.8	364.6	356.8	353.2	347.2
0.8	409.2	405.1	402.8	396.6	390.1	378.5	373.7	367.5	361.2	353.4	349.8	344.0
	406.2	402.7	400.5	394.0	386.8	375.2	370.4	363.9	357.6	349.9	346.9	340.4
0.9	402.7	399.8	397.8	391.2	383.7	371.7	366.9	360.4	354.2	346.2	342.7	336.7
	398.5	396.3	394.6	388.1	380.4	368.5	363.3	356.6	350.5	342.4	339.1	333.1
1.0	393.8	392.4	390.8	385.0	377.0	364.2	359.5	352.8	346.6	338.6	335.3	329.2
	388.1	388.6	386.2	381.8	373.8	360.4	355.8	348.8	342.6	334.7	331.3	325.4
1.1	381.8	381.4	380.8	378.4	370.2	356.2	351.4	344.7	338.6	330.6	327.3	321.3

Electrocapillary data on mercury in aqueous solutions of 1 mole.l^{-1} KNO_3 and various concentrations of 2-methylquinoline. Potentials measured with respect to a normal calomel electrode. Temperature 19 ± 1°C. Given is the interfacial tension in mN.m^{-1}.

Reference: S. Bordi, G. Papeschi, J. Electroanal. Chem. and Interfacial Electrochem., 20 (1969) 297.

conc. of 2-MeQ in mmole.l^{-1}. $\frac{E}{\text{volts}}$	0.05	0.10	0.25	0.50	1.0	1.5	2.0	5.0	9.0
0	374.35	375.34	373.06	370.39	367.81	368.19	364.92	361.88	357.16
	380.28	381.19	377.85	376.25	372.98	371.92	370.40	367.35	362.41
0.1	385.91	387.13	383.17	381.50	377.93	377.09	375.34	371.69	367.05
	391.69	392.22	388.42	385.98	382.34	381.20	379.60	375.87	370.02
0.2	396.41	396.64	392.30	389.94	386.14	384.77	382.95	378.38	372.14
	400.74	400.59	396.18	393.21	389.26	387.66	385.61	380.81	372.91
0.3	404.32	403.63	399.07	395.88	391.54	389.94	387.89	381.50	372.37
	406.14	406.60	401.73	397.78	393.52	391.69	389.41	380.59	371.54
0.4	409.79	408.27	403.48	399.30	394.81	392.76	390.17	379.07	370.09
	411.47	409.03	404.37	400.29	395.50	392.38	388.95	377.32	368.42
0.5	412.38	409.79	405.38	400.82	394.96	391.08	387.13	375.26	366.29
	412.76	410.33	405.23	400.67	393.52	389.33	385.07	372.98	364.08
0.6	412.23	410.40	404.85	400.36	391.84	387.28	382.87	370.62	361.57
	411.54	409.95	404.24	399.53	389.87	385.00	380.13	368.04	359.06
0.7	410.55	408.96	403.33	398.46	387.58	382.49	377.47	365.30	356.48
	409.26	407.82	402.04	397.02	385.07	379.67	374.58	362.33	353.97
0.8	407.36	406.07	400.44	394.43	382.57	377.01	371.61	359.22	351.00
	405.08	403.79	398.69	391.69	379.83	374.05	368.57	356.10	347.81
0.9	401.88	400.82	396.41	388.57	376.63	371.08	365.30	352.75	344.54
	397.93	397.24	393.97	385.00	373.44	367.96	362.03	349.48	341.19
1.0	393.44	392.91	390.40	381.35	370.55	364.69	358.53	345.75	337.84
	387.89	387.66	385.99	377.62	367.20	361.27	354.88	342.25	334.35
1.1	381.50	381.65	380.28	373.74	364.01	357.85	351.23	338.45	330.77

Electrocapillary data on mercury in aqueous solutions of 1 mole.l^{-1} KNO_3 and various concentrations of 4-methylquinoline. Potentials measured with respect to a normal calomel electrode. Temperature 19 ± 1 °C. Given is the interfacial tension in mN.m^{-1}.

Reference: S. Bordi, G. Papeschi, J.Electroanal. Chem. and Interfacial Electrochem., 20, (1969), 297.

conc. of 4-MeQ in mmole.l^{-1}. E / volts	0.05	0.10	0.25	0.50	1.0	1.5	2.0	5.0	9.0
0	374.96	372.98	372.67	370.70	368.03	364.92	363.55	359.74	349.40
	380.81	379.29	378.53	376.10	372.98	370.32	368.87	364.92	354.80
0.1	385.76	385.15	383.32	381.19	378.00	375.18	373.13	369.48	358.22
	392.07	390.40	388.27	385.30	381.95	379.22	377.31	373.36	361.34
0.2	396.56	394.96	392.30	389.18	385.76	382.94	380.89	375.11	364.00
	400.05	399.07	396.18	392.45	388.72	385.98	383.02	376.17	364.92
0.3	403.93	402.42	398.99	395.11	391.16	387.89	385.22	375.18	364.38
	406.82	404.85	401.42	397.01	392.68	389.25	386.36	373.43	363.40
0.4	408.73	406.67	403.02	398.31	393.83	389.86	385.38	371.76	362.18
	410.25	407.89	404.09	399.22	394.50	388.72	384.08	369.94	360.73
0.5	411.39	408.65	404.62	399.83	393.44	387.43	382.33	367.88	385.98
	412.15	409.03	404.47	399.52	391.92	385.98	380.28	365.68	356.85
0.6	411.92	408.88	404.01	398.99	389.79	383.55	377.92	363.62	354.80
	411.39	408.50	403.40	398.15	387.66	380.96	375.49	361.34	351.99
0.7	410.70	407.59	402.64	397.47	385.30	378.30	372.98	358.68	349.32
	409.79	406.22	401.35	396.10	382.79	375.41	370.16	355.33	346.43
0.8	408.04	404.24	400.13	394.81	380.36	372.75	367.27	352.60	343.24
	405.68	401.96	398.69	393.21	377.85	369.86	364.36	349.48	340.05
0.9	401.88	399.07	396.63	391.31	374.96	367.05	361.04	346.13	336.55
	398.00	395.80	394.35	388.95	372.29	364.08	357.69	342.86	333.20
1.0	393.06	391.54	390.62	386.06	369.25	361.11	354.19	339.06	329.70
	387.13	385.76	386.36	382.94	366.51	357.77	350.69	335.63	325.75
1.1	380.89	380.05	381.04	379.52	363.70	354.50	347.58	331.91	322.02

Electrocapillary data for mercury in 0.1 mol l^{-1} aqueous $HClO_4$ containing cyclo-hexanol

T = 25°C

Potentials with respect to saturated NaCl calomel electrode

Reference: D.E. Broadhead, R.S. Hansen and G.W. Potter, Jr.
J. Colloid and Interface Sci. 31, 61 (1969)

	γ /mN m^{-1}						
c/mol l^{-1} E/volt	0	0.195	0.0977	0.0488	0.0217	0.0109	0.00543
-1.100	384.76	353.78	362.40	370.64	379.04	382.59	384.00
-1.050	390.44	356.14	364.60	372.81	381.77	387.44	389.31
-1.000	395.80	358.23	366.78	375.20	384.51	391.01	394.28
- .950	400.74	360.29	368.77	377.17	386.50	394.07	398.61
- .900	405.31	362.15	370.70	379.15	388.71	396.53	402.51
- .850	409.45	363.98	372.46	380.86	390.40	398.50	405.55
- .800	413.20	365.58	374.08	383.58	392.30	400.26	408.29
- .750	416.55	367.15	375.66	384.09	393.67	401.98	410.07
- .700	419.42	368.57	377.05	385.55	395.30	403.40	412.02
- .650	421.81	369.91	378.42	386.85	396.46	404.81	413.23
- .600	423.69	371.10	379.60	388.05	397.82	406.04	414.48
- .550	424.95	372.26	380.71	389.14	398.71	407.08	415.53
- .500	425.55	373.18	381.69	390.14	399.59	408.04	416.34
- .450	425.46	374.13	382.57	390.95	400.45	408.76	417.01
- .400	424.76	374.88	383.33	391.72	401.07	409.75	417.45
- .350	423.28	375.55	383.96	392.27	401.65	409.80	417.52
- .300	421.26	376.06	384.47	392.83	401.98	409.87	417.29
- .250	418.47	376.45	384.84	393.01	402.26	409.94	415.85
- .200	415.18	376.67	385.07	393.27	402.12	409.36	413.39
- .150	411.19	376.75	385.05	393.06	401.81	407.71	409.91
- .100	406.66	376.57	384.84	392.71	400.54	404.46	405.76
- .050	401.46	375.97	384.17	391.55	397.82	399.98	400.89
0.000	395.78	374.55	382.87	389.49	393.44	394.79	395.39
.050	389.38	371.79	380.57	385.59	387.92	388.80	389.20
.100	382.51	368.01	376.93	380.19	381.47	382.17	382.40
.150	374.93	363.10	371.51	373.44	374.33	394.73	374.93
.200	366.77	357.30	364.60	365.84	366.31	366.68	366.79
.250	357.91	350.35	356.44	357.03	357.66	357.71	358.01
.300	348.50	342.66	347.56	348.20	348.16	348.39	348.46
E^z/volt	-0.479	-0.151	-0.154	-0.176	-0.233	-0.301	-0.372
γ^z/mN m^{-1}	425.60	376.84	385.32	393.29	402.50	410.30	417.68

Electrocapillary data on mercury electrode in 0.9 M NaF.
Adsorption of ortho-cresol at 25°C - given are the interfacial tension in mN m^{-1}, potential in V vs NCE (normal calomel electrode).

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

	γ/mN m^{-1}							
c/mM l^{-1}	0	1.55	3.1	6.2	15.5	30.9	76.9	149.5
- E/V								
0.012								
0.050				381.4		371.6	361.5	341.4
0.100	408.1	395.2	392.6	386.7	379.8	375.3	366.3	353.8
0.150	413.5	401.1	396.0	391.1	384.6	379.3	368.9	357.8
0.200	417.5	405.2	399.6	396.8	388.6	381.7	372.1	362.1
0.250	421.1	410.0	403.6	399.3	391.3	384.1	374.1	364.2
0.300	423.5	413.3	406.5	401.1	393.3	385.9	376.2	366.1
0.350	425.9	413.8	408.4	403.2	395.3	387.7	377.8	368.1
0.400	427.5	415.9	410.0	405.3	396.4	388.9	379.1	369.5
0.450	428.0	416.8	413.3	406.5	397.8	389.8	380.0	370.6
0.500	428.0	418.0	415.5	407.5	398.5	390.9	380.9	371.6
0.550	427.6	417.4	415.8	408.0	399.5	391.6	381.8	372.3
0.600	425.8	417.4	417.0	408.2	399.5	392.0	381.5	372.8
0.650	423.6	417.8	416.8	408.3	399.5	391.4	381.7	373.2
0.700	421.5	416.6	416.3	408.0	398.9	391.4	381.5	373.0
0.750	418.1	415.3	414.2	408.0	398.9	391.1	381.0	373.0
0.800	415.0	412.5	412.9	407.5	398.7	391.1	380.7	373.0
0.850	411.2	409.6	408.6	405.6	397.0	389.4	379.8	372.7
0.900	406.9	405.2	405.2	402.5	396.4	389.0	379.1	371.9
0.950	402.4	401.7	401.4	398.7	394.1	387.8	378.3	371.1
1.000	397.1	397.0	396.2	394.7	392.1	386.1	376.6	370.4
1.050	391.6	392.4	392.4	390.7	388.7	384.2	374.7	368.7
1.100	385.3	385.8	385.9	386.9	384.6	381.0	373.6	367.0
1.150	379.1			378.9	379.1	377.6	371.1	364.4
1.200	372.7			372.7	373.4	373.9	368.7	362.0
1.250	365.0			365.6	365.9	365.3	363.8	359.3
1.300	358.0			358.0	358.4	357.3	357.3	355.7
1.350	349.7			349.5			349.7	349.3
1.400	341.2			341.0		341.8	341.8	341.4
1.450	332.9							333.0
1.500	323.5					322.9		323.9
1.550	313.4							313.7
1.600	303.3							

Electrocapillary data on mercury electrode in 0.9 M NaF at 25°C. Adsorption of meta-cresol - given are the interfacial tension in mN m^{-1}, potential in V vs NCE (normal calomel electrode).

Reference: G. Manohar and S. Sathyanarayana. J.Electroanal.Chem. 30 (1971) 301.

	γ /mN m^{-1}						
c/mM l^{-1}	1.5	3.0	6.0	14.9	29.7	73.9	137.9
$-E$/V							
0.10	391.2	397.0	387.6	380.7	374.6	358.4	348.5
0.15	395.4	399.0	390.7	383.7	377.1	362.9	351.6
0.20	398.9	400.7	393.6	386.7	380.6	367.0	355.3
0.25	403.8	402.1	396.0	389.9	382.4	371.0	359.2
0.30	408.0	405.7	399.6	392.0	384.7	372.9	362.3
0.35	410.7	407.1	401.3	393.7	386.6	375.9	365.3
0.40	413.0	408.9	402.4	395.6	388.3	377.6	368.0
0.45	414.8	411.8	404.4	397.1	389.7	379.1	370.0
0.50	415.3	412.5	406.3	397.7	390.5	380.3	371.2
0.55	416.2	412.8	406.1	398.1	391.5	380.9	372.7
0.60	416.5	413.5	406.8	398.6	391.7	381.6	373.0
0.65	416.1	413.5	406.5	398.6	392.0	381.4	374.2
0.70	414.4	412.6	407.0	398.6	392.0	381.3	374.3
0.75	413.2	412.1	406.1	398.6	391.3	381.0	373.6
0.80	411.8	410.2	406.0	398.0	390.7	380.4	373.7
0.85	408.5	407.8	404.0	397.0	390.0	379.9	373.0
0.90	405.4	405.0	402.3	395.6	388.6	379.1	372.6
0.95	402.4	402.4	399.6	393.7	387.0	377.6	371.0
1.00		397.4	395.7	390.4	385.4	376.7	370.2
1.05			391.4	388.6	383.7	374.9	368.9
1.10			386.1	384.6	380.6	372.7	367.4
1.15			379.9	378.3	376.3	370.7	365.2
1.20				372.6	370.9	367.6	363.3
1.25					364.9	363.3	360.4
1.30				357.9	357.3	356.6	356.4
1.35					349.5	349.3	349.7
1.40				340.7	340.7	340.9	341.9
1.45							
1.50						323.2	323.3

Electrocapillary data on mercury electrode in 0.9 M NaF at 25°C. Adsorption of para-cresol - given are interfacial tension in mN m^{-1}, potential in V vs NCE (normal calomel electrode).

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

	γ /mN m^{-1}						
c/mM l^{-1}	1.4	2.9	5.8	14.5	28.9	72.0	128.7
-E/V							
0.10	394.0	389.8	382.8	379.7	371.9	356.0	350.2
0.15	399.3	394.5	388.4	384.0	376.6	361.2	354.6
0.20	403.5	398.0	392.6	387.4	380.4	365.9	358.5
0.25	407.3	402.0	396.5	390.5	383.3	369.5	361.5
0.30	410.2	404.5	399.7	392.9	385.9	373.1	364.9
0.35	412.8	407.9	402.4	395.0	388.0	375.0	367.9
0.40	415.3	409.1	404.3	397.0	390.0	377.3	369.5
0.45	416.3	411.0	406.1	398.9	391.0	378.8	371.3
0.50	417.4	412.1	407.4	400.0	392.4	380.5	372.8
0.55	418.4	413.1	408.5	400.7	393.2	381.0	374.1
0.60	418.5	413.2	408.7	401.2	393.6	381.8	374.6
0.65	418.8	413.2	408.9	401.7	393.9	381.8	375.1
0.70	416.7	413.0	408.9	401.2	393.7	381.8	375.1
0.75	415.4	411.8	407.9	400.9	393.6	381.8	375.1
0.80	413.2	410.2	406.9	400.0	393.9	381.4	374.9
0.85	411.2	407.8	404.6	398.8	391.9	380.9	374.6
0.90	407.1	405.6	402.4	396.8	390.5	379.7	374.0
0.95	401.9	401.2	398.4	394.5	388.8	378.3	373.1
1.00	396.8	396.5	394.2	391.6	386.2	377.0	371.6
1.05			389.5	388.0	384.4	374.8	370.3
1.10			384.1	383.5	380.7	373.1	369.2
1.15					376.1	370.3	366.9
1.20			370.8	370.9	370.6	367.2	364.9
1.25						362.5	362.0
1.30				356.5	357.3	356.0	356.6
1.35							
1.40				340.2	340.2	341.3	341.9
1.45							
1.50							323.9

Electrocapillary data on mercury electrode in 1.0 M NaOH at 25°C. Adsorption of ortho-cresol - given are the interfacial tension in mN m^{-1}, potential in V vs MOE (mercury oxide electrode).

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

	γ /mN m^{-1}								
c/mM l^{-1}	0	1.55	3.1	6.2	15.5	30.9	76.9	152.5	227
-E/V									
0.100	418.5	405.4	402.4	399.6	394.3	391.0	384.7	379.8	377.4
0.150	421.5	411.1	407.0	405.3	400.0	397.3	391.4	385.8	383.5
0.200	423.5	414.6	411.5	409.0	403.8	401.7	395.4	390.5	388.8
0.250	425.0	416.7	414.4	413.0	407.8	405.7	400.7	394.8	393.6
0.300	425.5	419.5	417.8	415.1	411.0	409.0	403.0	398.5	396.8
0.350	425.5	421.1	419.0	417.5	413.3	411.1	405.3	401.0	400.0
0.400	425.0	421.6	419.5	417.5	414.7	412.4	408.5	403.3	401.7
0.450	423.5	421.2	419.0	418.0	415.0	413.0	409.3	404.3	403.4
0.500	421.9	420.3	418.2	416.9	414.7	413.4	409.2	404.8	404.0
0.550	419.4	418.0	416.3	416.3	413.7	412.4	409.1	404.7	404.2
0.600	416.1	415.3	413.9	413.6	412.2	411.0	407.1	404.3	403.6
0.650	412.5	412.3	411.5	410.5	409.5	409.0	406.3	402.5	402.6
0.700	408.4	408.2	407.9	406.9	406.3	405.9	403.3	401.0	401.2
0.750	404.0	404.2	403.7	402.8	403.8	402.0	400.8	398.5	398.7
0.800	399.5	399.6	399.3	398.1	399.0	397.8	396.7	394.8	395.8
0.850	394.4	394.5	394.3	392.4	393.7	392.7	392.3	390.8	392.0
0.900	388.7	388.6	388.7	387.5	389.0	387.0	387.5	385.7	386.7
0.950	382.1			381.7	381.7	380.9	381.8	379.8	382.0
1.000	375.6	375.7	375.6	375.4	375.3	374.4	375.4	373.8	375.8
1.050	368.8				368.9	368.9		367.3	368.4
1.100	361.6	361.6	361.5	361.5	361.6	361.6	361.5	360.0	360.9
1.150	353.9							352.5	353.3
1.200	345.5	345.5	345.5	345.4	345.5	345.5	345.4	344.0	344.9
1.250	336.7							336.0	336.3
1.300	327.8							327.7	327.4
1.350	318.4								308.1
1.400	308.3							308.2	
1.450									
1.500	286.7								
1.600	263.1								

Electrocapillary data on mercury electrode in 1.0 M NaOH at 25°C. Adsorption of meta-cresol - given are the interfacial tension in mN m^{-1}, potential in V vs MOE (mercury oxide electrode).

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

				γ/mN m^{-1}				
c/mM l^{-1}	1.5	3.0	6.0	14.9	28.7	73.9	146.7	218
-E/V								
0.10	409.9	404.6	403.3	397.4	393.3	388.6	383.5	381.6
0.15	414.5	409.4	407.8	402.4	398.7	395.0	389.0	387.6
0.20	417.8	413.4	411.5	406.8	403.4	399.7	394.2	392.7
0.25	419.6	416.2	415.2	410.5	407.8	403.5	398.4	396.4
0.30	421.0	418.9	417.3	413.6	410.2	406.5	401.4	400.1
0.35	421.9	420.0	419.5	415.3	412.9	404.0	404.4	402.1
0.40	421.9	420.6	419.7	416.3	414.1	410.8	406.2	403.5
0.45	421.2	420.2	419.8	416.6	414.8	411.2	407.4	404.4
0.50	419.8	418.9	418.6	416.3	414.5	411.7	408.0	404.8
0.55	417.4	417.5	416.9	415.1	413.6	411.2	407.7	405.0
0.60	415.2	415.3	414.8	413.5	412.1	410.1	407.0	404.5
0.65	411.8	411.6	411.8	410.7	409.8	408.2	405.4	403.8
0.70	407.9	407.3	407.8	407.0	407.1	405.7	403.4	401.7
0.75	403.2	403.1	403.5	403.3	403.3	402.1	400.7	399.8
0.80	398.6	397.7	400.0	398.3	398.6	398.4	396.9	396.4
0.85		392.8		392.9	394.0	393.9	392.6	392.2
0.90				387.5	388.5	388.5	387.8	387.9
0.95				381.4	382.2	382.6	382.4	382.4
1.00				375.1	376.1	376.0	376.1	
1.05							369.4	
1.10							362.3	
1.15							354.2	

Electrocapillary data on Mercury electrode in 1.0 M NaOH at 25°C. Adsorption of para-cresol - given are the interfacial tension in mN m^{-1}, potential in V vs MOE (mercury oxide electrode).

Reference: G. Manohar and S. Sathyanarayana. J. Electroanal. Chem. 30 (1971) 301.

	γ /mN m^{-1}							
c/mM l^{-1}	1.4	2.9	5.8	14.5	28.9	72.0	142.9	212.7
-E/V								
0.10		405.7	401.0	397.3	394.1	388.2	383.5	381.4
0.15		410.4	407.0	402.9	400.5	394.5	389.7	388.0
0.20		415.0	412.5	408.0	405.7	399.4	394.9	393.0
0.25		418.1	415.6	411.8	409.7	403.9	399.2	397.5
0.30		421.2	418.5	414.8	412.2	407.5	404.2	400.8
0.35		422.5	420.4	416.4	414.8	410.2	406.7	404.1
0.40		422.5	420.4	417.7	416.4	411.5	408.6	405.7
0.45		422.2	420.4	418.0	417.0	412.3	409.4	407.0
0.50		420.6	419.1	417.2	416.4	412.3	409.9	407.3
0.55		418.6	417.3	415.7	414.9	411.5	409.4	407.5
0.60		415.6	414.9	413.8	413.1	409.9	408.3	406.8
0.65		412.5	411.7	410.9	410.7	407.6	406.7	405.4
0.70		408.6	407.8	407.0	407.5	405.0	404.2	403.1
0.75		404.2	403.8	402.8	403.9	401.5	401.2	400.4
0.80		399.2	398.4	398.3	399.4	397.0	397.3	396.2
0.85			394.1	394.1	394.5	392.4	392.8	392.5
0.90			389.0	388.9	388.9	397.1	397.6	387.1
0.95						381.3	382.1	381.0
1.00					376.3	374.6	375.1	375.8
1.05							369.0	360.6
1.10						360.6	360.6	

Electrocapillary data on mercury in aqueous solutions of 0,1 m NaH_2PO_4 + acetanilide of concentration c_a in mole.l^{-1} T = 25^{o}. Potential measured with respect to saturated calomel electrode. Given is the interfacial tension x 10^3 in N.m^{-1}.

Reference: R. Parsons and F.G.R. Zobel, Trans. Faraday Soc. 62, 3511 (1966)

E / Volts	c_a = 2.55×10^{-3}	c_a = 5.18×10^{-3}	c_a = 1.00×10^{-2}	c_a = 2.04×10^{-2}	c_a = 3.25×10^{-2}
-1.70	281.4	282.2	281.1	281.2	281.3
-1.65	291.8	292.2	291.7	291.8	292.2
-1.60	302.4	302.8	302.2	302.5	302.4
-1.55	312.0	312.4	312.0	312.4	312.2
-1.50	321.7	322.1	321.7	321.7	321.7
-1.45	330.5	331.1	330.4	330.4	330.7
-1.40	339.8	339.9	339.4	339.1	339.2
-1.35	347.9	347.9	347.5	347.1	346.5
-1.30	355.4	355.7	354.9	353.9	352.7
-1.25	362.9	362.9	362.1	360.3	357.8
-1.20	369.7	369.5	368.3	366.0	362.2
-1.15	376.0	375.4	373.8	370.6	366.4
-1.10	382.0	380.9	378.7	374.5	370.4
-1.05	387.4	385.8	385.4	378.0	374.0
-1.00	392.1	389.7	386.6	381.2	377.7
-0.95	396.3	393.2	389.7	384.4	380.0
-0.90	399.6	396.1	392.4	387.4	382.5
-0.85	402.8	398.6	394.7	389.3	384.3
-0.80	405.6	400.9	396.9	391.2	386.9
-0.75	406.8	402.9	398.6	392.5	388.6
-0.70	408.5	404.6	399.8	393.9	389.9
-0.65	410.1	405.5	401.0	395.1	390.9
-0.60	410.6	406.1	402.2	395.9	391.9
-0.55	411.4	406.8	402.6	396.6	392.5
-0.50	411.3	407.1	405.3	396.9	393.1
-0.45	411.3	407.0	402.7	397.7	392.9
-0.40	411.0	407.2	402.2	396.9	392.8
-0.35	410.4	406.3	401.2	396.2	392.4
-0.30	408.9	405.1	401.1	395.7	392.0
-0.25	407.3	403.8	399.6	394.5	390.9
-0.20	405.9	401.8	398.0	393.2	389.7
-0.15	402.8	399.2	296.2	391.5	388.0
-0.10	400.0	396.0	393.6	389.3	386.1
-0.05	395.8	392.0	391.1	386.7	383.5
-0.00	390.5	387.3	387.2	383.4	380.4
+0.05	384.3	381.3	385.0	378.8	376.3
+0.10	376.9	373.7	376.1	373.1	371.0
+0.15	367.8	364.9	368.8	366.4	364.6
+0.20	357.4		359.8	357.6	356.3
+0.25			349.2	346.7	346.1

Electrocapillary data on mercury in aqueous solutions of $1{,}0$ m NaH_2PO_4 + acetanilide of concentration c_a in mole.l^{-1} T = 25°. Potential measured with respect to saturated calomel electrode. Given is the interfacial tension x 10^{-3} in $N.m^{-1}$.

Reference: R. Parsons and F.G.R. Zobel, Trans. Faraday Soc. 62, 3511 (1966)

E Volts	c_a =				
	1.26×10^{-3}	2.45×10^{-3}	4.89×10^{-3}	7.43×10^{-3}	1.00×10^{-2}
-1.70	273.4	270.8	271.1		270.8
-1.65	282.6	262.3	282.3	282.4	282.2
-1.60	293.9	293.1	293.4	293.5	293.1
-1.55	304.0	303.7	303.8	303.9	303.6
-1.50	314.1	313.8	313.9	314.2	313.7
-1.45	323.6	323.4	323.4	323.6	323.2
-1.40	333.0	-	332.7	333.1	332.4
-1.35	341.8	341.4	341.2	341.6	341.0
-1.30	350.2	349.5	349.4	349.8	349.0
-1.25	357.9	357.2	357.3	357.2	356.5
-1.20	365.6	364.6	364.4	364.2	363.1
-1.15	372.6	371.5	370.8	370.1	368.7
-1.10	379.0	377.7	376.6	375.4	373.3
-1.05	384.8	383.4	381.7	380.1	377.4
-1.00	389.9	388.2	386.0	384.1	381.3
-0.95	394.5	392.4	389.6	386.9	384.5
-0.90	398.9	396.2	392.5	389.7	387.2
-0.85	402.6	399.1	395.7	393.2	389.6
-0.80	405.0	401.9	398.8	395.3	392.8
-0.75	407.7	404.1	399.7	396.8	394.5
-0.70	409.5	406.2	401.1	398.3	395.6
-0.65	411.4	407.5	402.5	399.5	396.6
-0.60	412.2	408.7	403.6	400.4	397.8
-0.55	413.7	409.2	404.1	401.0	398.5
-0.50	413.6	409.5	404.7	401.0	399.1
-0.45	413.9	409.4	404.5	401.0	399.1
-0.40	413.1	409.2	404.4	400.5	399.1
-0.35	412.7	408.8	403.5	400.0	398.2
-0.30	411.2	407.3	402.5	398.9	397.7
-0.25	409.8	405.7	401.3	397.8	396.6
-0.20	407.4	403.5	399.7	396.9	395.4
-0.15	405.0	401.3	397.6	394.2	393.4
-0.10	400.8	397.9	394.7	391.5	390.9
-0.05	396.0	393.5	390.9	387.9	387.9
-0.00	389.4	387.2	386.1	383.1	383.6
+0.05	381.6	380.6	379.5	377.1	377.7
+0.10	372.5	371.9	371.4	369.6	370.1
+0.15	362.1	361.6	361.8	360.0	361.0
+0.20	350.6	350.4	350.4	348.9	350.2
+0.25	337.1	337.3	337.4	337.1	337.4

Electrocapillary data on mercury in aqueous solutions of NaH_2PO_4 (conc. in mole.l^{-1} indicated) saturated with acetanilide. T = 25^o. Potential with respect to saturated calomel electrode. Given is the interfacial tension x 10^3 in N.m^{-1}.

Reference: R. Parsons and F.G.R. Zobel, Trans. Faraday Soc. 62, 3511 (1966)

$\frac{E}{\text{Volts}}$ c =	0.0993	0.5021	1.0078	2.014	4.019
-1.70	280.7	272.3	271.1		
-1.65	291.6	285.4	282.3	281.0	
-1.60	302.2	296.1	293.5	292.5	
-1.55	311.9	306.6	303.8	302.8	
-1.50	321.4	317.5	314.0	312.5	309.7
-1.45	330.0	326.8	323.3	322.1	418.5
-1.40	338.1	325.2	332.4	331.8	
-1.35	344.8	342.8	340.8	340.3	319.9
-1.30	349.9	349.2	348.3	348.5	328.3
-1.25	354.7	354.2	353.6	355.9	337.4
-1.20	359.0	360.1	358.2	362.7	345.8
-1.15	362.7	364.6	363.0	367.7	353.7
-1.10	366.9	368.4	366.9	371.4	362.1
-1.05	370.3	372.0	370.8	374.4	370.2
-1.00	373.2	375.6	374.4	376.7	376.5
-0.95	375.8	378.6	377.4	379.9	381.9
-0.90	378.7	380.8	380.5	382.5	386.2
-0.85	380.9	383.2	383.0	385.0	388.9
-0.80	382.8	385.3	385.3	387.4	389.2
-0.75	384.5	387.0	387.1	388.5	392.6
-0.70	385.9	388.2	388.7	390.7	393.1
-0.65	387.2	389.6	390.1	391.4	394.8
-0.60	387.9	390.4	391.1	392.3	397.5
-0.55	388.5	391.0	391.8	392.8	397.0
-0.50	389.0	391.1	392.0	393.3	399.0
-0.45	389.6	391.2	392.5	393.4	396.8
-0.40	389.1	391.1	392.1	393.4	398.8
-0.35	388.8	390.8	391.8	393.5	398.4
-0.30	388.2	389.9	391.0	392.3	396.1
-0.25	387.3	389.0	390.2	391.6	396.4
-0.20	386.2	387.8	388.9	392.6	394.3
-0.15	384.7	385.9	387.0	388.9	393.1
-0.10	382.5	383.4	385.4	386.8	389.6
-0.05	379.5	380.3	382.9	384.4	387.6
-0.00	375.9	376.0	379.5	380.5	383.5
+0.05	371.2	370.0	374.7	374.5	368.8
+0.10	365.2	363.0	368.1	366.6	359.4
+0.15	357.6	353.6	359.8	357.1	348.6
+0.20	347.9		349.4	345.9	336.2
+0.25			337.0	333.1	322.7
+0.30				316.3	

Electrocapillary data for mercury in 0.1 mol l^{-1} aqueous $HClO_4$ containing chloroform

T = 25°C

Potentials with respect to saturated NaCl calomel electrode

Reference: D.E. Broadhead, R.S. Hansen and G.W. Potter, Jr.
J. Colloid and Interface Sci., 31, 61 (1969).

	γ /mN m^{-1}								
c/mol l^{-1} E/volt	0	0.0126	0.0178	0.0251	0.0316	0.0377	0.0445	0.0460	0.0494
-1.100	384.56	384.56	384.19	384.05	383.77	383.91	383.67	383.62	383.26
-1.050		390.04	389.79	389.62	389.48	389.44	389.18	389.11	388.45
-1.000	395.66	395.28	395.05	394.77	394.51	394.50	394.24	394.07	393.46
-0.950		400.18	399.93	399.62	399.39	399.26	398.66	398.59	397.62
-0.900	405.21	404.60	404.35	403.96	403.56	403.37	402.88	402.65	401.19
-0.850	409.40	408.76	408.35	407.89	407.45	407.19	406.41	405.93	403.96
-0.800	413.11	412.39	411.94	411.34	410.79	410.30	409.18	408.73	405.43
-0.775						411.52	410.33	409.58	406.17
-0.750	416.47	415.58	415.10	414.37	413.63	412.90	411.22	410.49	407.03
-0.725		417.02	416.48	415.65	414.83	413.94	411.78	411.05	407.28
-0.700	419.32	418.31	417.74	416.85	415.89	414.55	412.43	411.53	407.71
-0.675		419.44	418.87	417.96	416.77	415.07	412.93	412.00	408.32
-0.650	421.69	420.59	419.95	418.88	417.52	415.81	413.34	412.38	408.59
-0.625		421.52	420.83	419.65	418.10	416.15	413.43	412.57	408.84
-0.600	423.50	422.26	421.58	420.24	418.51	416.33	413.72	412.81	409.06
-0.575		422.83	422.19	420.76	418.76	416.54	413.95	413.02	409.20
-0.550	424.78	423.44	422.78	421.17	419.08	416.67	414.20	413.15	409.34
-0.525		423.75	423.07	421.39	419.19	416.83	414.04	413.18	409.38
-0.500	425.44	424.00	423.36	421.55	419.26	416.85	414.20	413.20	409.38
-0.475		424.00	423.39	421.66	419.17	416.76	414.15	413.18	409.43
-0.450	425.44	423.96	423.50	421.57	419.10	416.72	414.13	413.04	409.16
-0.425		423.69	423.18	421.50	419.14	416.60	413.83	412.93	409.13
-0.400	424.76	423.32	422.82	421.28	418.87	416.42	413.74	412.75	408.93
-0.375		422.76	422.28	420.92	418.58	416.08	413.50	412.50	408.77
-0.350	423.32	422.01	421.63	420.38	418.20	415.77	413.22	412.18	408.23
-0.325		421.11	420.65	419.79	417.74	415.52	412.71	411.82	407.96
-0.300	421.15	419.94	419.66	418.77	417.31	414.64	412.12	411.41	407.53
-0.275		418.78	418.26	417.73	416.52	414.36	411.80	410.69	407.08
-0.250	417.78	417.07	416.88	416.15	415.12	413.40	411.12	410.19	405.99
-0.225		415.69	415.30	414.86	414.33	412.99	410.13	409.40	405.52
-0.200	414.44	413.97	413.54	412.90	412.87	411.45	409.00	408.55	404.77
-0.175		412.08		411.59	411.19	410.51	408.80	408.09	404.19
-0.150	410.27	409.82	409.97	409.38	409.41	408.54	407.56	406.97	403.01
-0.125					407.36	406.71	406.09	405.88	402.52
-0.100	405.65	405.05	405.28	404.95	404.78	404.16	403.76	403.87	401.48
-0.050	400.97	400.52	400.47	400.14	399.77	399.67	399.18	399.36	398.63
0.000	394.77	394.49	394.62	394.54	394.15	393.91	393.72	393.87	393.55
0.050	389.41	388.85	388.93	388.92	388.66	388.36	388.10	388.09	386.92
0.100	381.65	381.44	381.73	381.79	381.15	381.16	380.81	381.19	380.83
0.150	374.83	374.74	374.83	374.75	374.83	374.39	374.13	374.42	
0.200	366.22	366.15	366.39	366.38	366.12	366.04	365.65	366.02	365.12
0.250	357.91	357.94	358.00	358.03	357.89	357.65	356.87	357.84	356.64
0.300	348.45	348.37	348.41	348.54	348.46	348.25	348.10	348.34	347.71
E^z/volt	-0.474	-0.476	-0.472	-0.468	-0.473	-0.509	-0.511	-0.515	-0.501
γ^z/mN m^{-1}	425.60	424.07	423.49	421.78	419.29	416.78	414.10	413.15	409.49

Electrocapillary data for adsorption of Furfurylamine on mercury electrode in 0.3 M HCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE.
T = 25°C

Reference: R.G. Hamilton. Ph.D. Thesis, Toronto, 1967.

	γ /mN m^{-1}					
c/mol l^{-1}	.0	0.01	0.02	0.05	0.1	0.2
-E/mV						
0	357.4	357.9	357.3	355.5	354.4	353.6
100	383.9	384.2	383.3	382.0	380.9	380.1
200	401.0	400.4	399.9	398.2	396.9	395.6
300	412.4	412.3	411.5	409.7	408.3	406.4
400	420.5	419.9	419.4	417.0	415.6	413.4
450	423.2	422.4	421.6	419.2	417.8	415.6
480				420.3	418.8	416.6
500	424.4	423.8	423.0	420.7	419.1	416.7
520	425.0	423.9	423.1	420.8	419.2	417.0
540	425.2	424.1	423.2	420.8	419.2	417.0
560	425.2	424.1	423.2	420.8	419.3	417.0
580	425.0	424.1	423.1	420.5	419.2	416.8
600	424.8	423.7	422.9	420.3	418.8	416.6
620						
650	423.8	422.7	421.7	419.0	417.7	415.4
700	421.9	420.7	419.6	417.3	415.8	413.3
800	417.0	415.3	414.2	411.6	410.2	407.7
900	409.3	408.5	406.8	404.4	402.8	400.6
1000	400.6	399.2	398.2	395.4	394.0	391.6
1100	389.7	388.5	387.6	384.7	383.2	380.9
1150						
1200	377.5	376.5	376.2	372.7	371.3	368.7

Electrocapillary data for adsorption of Furfurylamine in 0.6 M HCl.

T = 25°C

Reference: R.G. Hamilton. Ph.D. Thesis, Toronto, 1967.

	γ /mN m^{-1}						
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5
$-E$/mV							
0	354.9	354.4	353.5	352.1	351.9	350.3	349.1
100	381.1	381.7	380.3	379.3	379.0	376.9	374.9
200	397.3	397.9	396.7	395.6	395.0	392.4	390.5
300	410.3	410.2	408.9	407.7	406.9	403.8	401.3
400	418.7	418.6	417.6	415.2	414.4	411.5	407.8
450	421.5	420.8	419.4	417.4	416.6	413.3	409.6
480						414.6	410.5
500	423.5	422.7	420.9	419.2	418.0	414.6	410.7
520	424.0	423.2	421.2	419.5	418.3	415.0	411.1
540	424.2	423.4	421.4	419.8	418.7	415.0	411.1
560	424.4	423.4	421.6	419.7	418.6	415.0	411.1
580	424.6	423.4	421.4	419.7	418.5	415.0	410.9
600	424.3	422.9	421.1	419.5	418.1	414.6	410.5
620	423.9						
650	422.9	421.7	420.0	418.5	416.9	413.5	409.9
700	421.4	420.2	418.1	416.5	415.0	411.9	408.2
800	415.3	414.5	412.3	411.9	409.6	406.4	402.4
900	407.7	406.9	404.8	403.2	402.1	398.7	394.9
1000	398.1	397.2	395.6	394.0	392.9	389.9	386.4
1100	387.0	386.2	384.5	383.3	382.1	379.6	375.4
1200	374.1	373.3	372.1	370.7	370.0	366.9	363.8

Electrocapillary data for adsorption of Furfurylamine in 1.0 M HCl.

T = 25°C

Reference: R.G. Hamilton. Ph.D. Thesis, Toronto, 1967.

	γ/mN m^{-1}									
c/mol l^{-1}	0	0.01	0.02	0.03	0.05	0.1	0.2	0.3	0.5	1.0
-E/mV										
0	351.6	351.8	352.3	350.6	350.6	349.3	348.6	345.8	345.7	342.5
100	380.0	379.2	380.0	377.6	379.0	377.6	376.5	373.7	373.2	370.7
200	396.0	396.0	396.6	394.4	395.2	392.6	392.8	389.7	388.6	386.2
300	408.5	408.5	408.9	408.5	407.2	405.3	404.5	401.2	399.1	396.2
400	417.0	417.3	417.1	414.8	415.1	412.7	411.8	408.2	405.8	402.7
450	420.0	420.0	420.1	417.5	417.8	415.2	413.7	410.6	407.5	404.4
480									408.4	405.0
500	421.8	422.1	421.3	419.3	419.3	417.0	415.1	411.8	409.5	405.8
520	422.6	422.2	421.5	419.7	419.6	417.4	415.3	412.1	410.1	406.3
540	422.8	422.3	422.1	420.0	419.8	417.7	415.4	412.4	410.1	406.3
560	423.0	422.3	422.1	420.2	419.9	417.8	415.6	412.4	410.1	406.2
580	423.0	422.3	422.1	420.1	419.9	417.7	415.5	412.4	410.0	
600	422.8	422.2	422.1	420.0	419.7	417.4	415.4	412.3	409.1	405.7
620	422.5		422.1							
650	422.0	421.3	420.9	419.1	418.7	416.4	414.5	412.1	408.8	
700	419.8	419.5	418.7	417.2	416.6	414.7	412.7	411.0	406.3	402.5
800	414.2	413.5	413.3	411.8	411.2	409.2	407.0	408.9	401.0	397.6
900	406.3	406.1	405.6	404.1	404.0	401.7	399.5	403.9	394.6	390.3
1000	396.3	395.8	396.0	393.4	394.0	392.4	390.4	387.9	386.0	381.7
1100	384.8	384.5	384.7	383.9	383.3	381.9	380.4	377.6	375.5	371.2
1200	371.7	371.6	371.8	371.4	370.6	369.3	367.8	365.7	363.0	358.8

Electrocapillary data for adsorption of Furfurylamine in 3.0 M HCl.

T = 25 °C

Reference: R.G. Hamilton. Ph.D. Thesis, Toronto, 1967.

	γ/mN m^{-1}							
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5	1.0
-E/mV								
0	343.1	342.3	341.9	341.8	340.5	338.5	336.1	333.7
100	370.7	370.5	370.1	369.9	369.1	367.6	365.5	363.3
200	388.0	387.8	386.9	386.8	385.3	383.9	381.6	379.3
300	400.5	400.7	399.7	399.1	397.6	395.5	392.4	390.2
400	409.9	409.8	408.6	407.8	405.8	403.6	399.8	396.8
450	413.2	413.0	411.9	410.6	408.8	405.7	401.9	398.5
480	415.0							399.4
500	415.7	415.6	413.8	412.8	410.7	407.8	403.2	399.5
520	416.5	416.4	414.8	413.4	411.5	407.9	403.8	399.9
540	417.0	416.7	415.3	413.9	411.6	408.3	404.0	400.0
560	417.3	417.1	415.7	413.9	411.9	408.8	404.3	400.0
580	417.8	417.4	415.9	414.0	411.9	408.6	404.3	399.7
600	418.1	417.3	415.9	414.4	412.0	408.6	404.0	399.5
620	417.9	417.3	415.8	414.4	411.9	408.5		
650	417.8	417.0	415.4	414.0	411.6	408.3	403.4	398.8
700	418.5	415.8	414.2	412.8	410.2	406.6	401.7	397.3
800	411.3	410.8	409.6	407.9	405.1	402.0	397.1	392.1
900	403.1	403.1	401.9	400.6	398.5	395.2	390.5	385.8
1000	392.4	392.8	391.7	390.3	388.5	386.3	382.1	377.3
1100	380.3	379.9	380.4	379.5	378.5	376.0	371.7	368.3
1200	379.9	372.5	371.9	371.0	370.2	367.7	363.8	359.6

Electrocapillary data for adsorption of Furfurylamine in 5.0 M HCl.

T = 25°C

Reference: R.G. Hamilton. Ph.D. Thesis, Toronto, 1967.

	γ /mN m^{-1}							
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5	1.0
-E/mV								
0	336.7	334.0	337.8	338.4	337.9	336.0	332.7	328.8
100	365.2	363.7	364.8	365.7	365.6	363.6	361.7	359.6
200	382.3	381.4	381.9	382.6	381.9	379.9	378.4	376.3
300	395.3	394.6	394.6	394.6	394.5	392.0	389.6	386.3
400	405.1	404.1	403.7	403.7	402.4	400.3	396.9	393.2
450	408.4	407.3	407.2	406.4	405.6	402.8	399.1	395.3
480	410.2							
500	411.3	410.4	409.5	409.5	407.6	404.7	400.6	396.4
520	412.2							396.6
540	412.9	412.3	411.0				401.4	396.7
550				410.5	409.2	406.0		
560	413.6						401.7	396.9
580	414.1	413.1	411.7	411.1	409.5	406.3		396.8
600	414.4	413.8	412.2	411.1	409.6	406.6	401.8	396.9
620	414.5	413.9	412.5	411.4	409.7	406.6	401.7	396.6
640	414.6	413.8	412.4	411.3	409.6	406.5		
650								396.4
660	414.6	413.9	412.3	411.2	409.4	406.2		
680	414.4	413.7	412.1					
700	414.0	413.4	411.6	410.5	408.9	405.5	400.4	394.9
800	410.4	409.8	407.8	406.7	404.9	401.5	395.4	390.6
900	402.0	401.9	400.4	398.8	398.6	394.3	387.5	384.0
1000	393.2	393.0	392.2	389.1	391.2	386.2	380.5	375.7
1100	383.4	386.8	383.6	385.3	380.8	377.5	375.6	371.6
1200	379.5	377.4	380.8	377.5	376.7	375.3	371.8	365.3

Electrocapillary data for adsorption of Tetrahydrofurfurylamine in 1.0 M HCl. T = 25°C

Reference: R.G. Hamilton. Ph.D. Thesis, Toronto, 1967.

	γ /mN m^{-1}				
c/mol l^{-1}	0.01	0.03	0.1	0.3	0.5
-E/mV					
0	352.6	351.6	351.2	350.2	348.9
100	379.8	379.1	378.8	378.3	377.5
200	396.3	396.0	395.4	394.3	393.5
300	408.6	408.2	407.5	405.8	404.6
400	417.5	416.6	415.3	413.0	411.0
450	420.3	419.4	417.7	414.6	412.6
480					413.0
500	421.9	421.2	418.8	415.4	413.2
520			419.0	415.5	413.0
540	422.7	421.7	419.1	415.5	
550					412.7
560	422.8	421.7	419.1	415.2	
580	422.8	421.6	418.9	414.8	
600	422.6	421.4	418.5	414.3	411.4
620	422.3	420.6			
650	421.4	420.1	417.0	412.5	409.3
700	419.6	418.0	414,6	409,8	406.5
800	413.6	411.9	408.2	403.2	400.1
900	405.5	403.8	400.0	394.9	391.7
1000	395.7	394.0	390.3	385.3	381.7
1100	384.2	382.7	378.9	374.5	371.2
1150	378.3	376.3	372.8	368.3	365.5
1200	372.7	370.5	366.6	361.7	358.6

Electrocapillary data on mercury electrode in 1 M KCl adsorption of 3×10^{-2} M furfurylamine - given are the interfacial tension in mN m^{-1}, potential in V vs SCE (KCl). T = 25°C

Reference: E.M. Kimmerle. Ph.D. Thesis, University of Toronto, 1967.

E/V	γ/mN m^{-1}	E/V	γ /mN m^{-1}
0.025	362.6	0.625	410.0
0.050	369.3	0.650	409.0
0.075	375.3	0.675	407.9
0.100	380.0	0.700	406.7
0.125	385.6	0.725	405.5
0.150	390.3	0.750	404.1
0.175	394.3	0.775	402.8
0.200	397.9	0.800	401.3
0.225	401.1	0.825	399.7
0.250	403.8	0.850	398.1
0.275	406.2	0.875	396.4
0.300	408.1	0.900	394.5
0.325	409.8	0.925	392.5
0.350	411.1	0.950	390.4
0.375	412.1	0.975	388.2
0.400	412.8	1.000	385.7
0.425	413.3	1.025	383.1
0.450	413.5	1.050	380.4
0.475	413.5	1.075	377.4
0.500	413.3	1.100	374.3
0.525	412.9	1.125	371.1
0.550	412.4	1.150	367.7
0.575	411.7	1.175	364.3
0.600	410.9	1.200	360.9

Electrocapillary data for adsorption of furfurylamine in 0.01 M KCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T=25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

			γ/mN m^{-1}			
c/mol l^{-1}	0.01	0.02	0.05	0.1	0.2	0.5
-E/mV						
100	383.9	383.0	382.7	382.8	381.4	376.3
150				392.1	388.1	383.0
200	400.7	399.8	399.1	398.3	394.1	388.5
250						391.0
280						392.1
300	412.0	410.4	408.9	405.7	400.4	392.5
320						392.9
340					401.5	392.7
360					401.8	
380				407.3	401.9	392.5
400	418.3	416.3	412.7	408.3	401.8	392.4
420				408.5	401.7	
440				408.6		
450		417.8			401.6	
460			413.7	408.4		
480			413.8	408.3		
500	421.4	418.6	413.9	408.1	401.5	391.8
520		418.7	413.8			
540	421.8	418.7	413.7			
550				407.6		
560	421.8	418.7	413.3			
580	421.9	418.7				
600	421.9	418.6	412.9	407.0	399.9	389.6
620	421.8					
650		417.9				
700	420.9	417.2	411.0	404.2	397.6	387.1
800	418.2	414.4	408.3	401.1	394.2	383.6
900	414.9	410.2	404.2	397.0	390.0	379.6
1000	408.9	404.9	399.1	391.5	384.9	374.6
1100	399.6	398.0	393.0	384.9	378.9	368.3
1200	391.7	389.0	384.7	377.7	371.9	362.1

Electrocapillary data for adsorption of furfurylamine in 0.03 M KCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}					
c/mol l^{-1}	0.01	0.02	0.05	0.1	0.2	0.5
-E/mV						
100	384.1	383.5	381.2	380.7	378.1	375.1
105						374.7
110						375.1
150					386.9	383.1
200	400.4	400.3	398.8	397.1	394.2	389.0
250						392.1
300	411.8	410.8	408.5	405.9	401.8	393.5
320						393.5
340					402.7	393.6
350				408.0		
360					402.9	393.6
380				408.9	403.3	393.5
400	418.6	417.1	413.3	408.9	403.2	393.4
420				409.2	403.0	393.3
440				409.3	403.0	
450	420.6	418.5				392.6
460			414.3	409.3		
480		418.9	414.5			
500	422.0	419.7	414.5	409.3	402.8	392.1
520	422.2	419.7	414.5			
540	422.7	419.7	414.4	409.0		
560	422.7	419.8	414.2			
580	422.5	419.8				
600	422.4	419.6	413.6	408.1	401.2	390.1
650	421.9	418.7	412.6			
700	421.0	417.6	411.3	405.6	398.9	387.6
800	417.5	414.3	408.2	402.8	395.5	384.1
900	412.5	409.9	403.6	398.2	391.1	379.9
1000	406.4	404.0	398.1	393.2	386.1	374.8
1100	398.5	396.0	392.1	387.2	380.3	369.3
1200	388.7	387.0	383.8	379.8	373.4	361.9

Electrocapillary data for adsorption of furfurylamine in 0.1 M KCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T=25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}					
c/mol l^{-1}	0.01	0.02	0.05	0.1	0.2	0.5
-E/mV						
100	384.1	384.6	382.1	379.6	376.8	
110						372.6
220	401.0	401.2	399.2	396.1	393.6	387.0
300	412.4	411.9	409.4	404.9	401.0	392.3
340						392.7
350				407.0	402.2	
360						392.7
380					402.5	392.6
400	419.0	417.7	413.7	408.1	402.6	392.5
420				408.2	402.5	392.4
440			414.4	408.2	402.5	
450	420.7	419.1				392.0
460			414.6	408.2	402.3	
480			414.7	408.1		
500	421.5	419.8	414.7	407.9	401.9	391.1
520	421.7	419.8	414.5			
540	421.9	419.8				
550			414.2	407.1	401.0	
560	421.8	419.6				
580	421.6	419.3				
600	421.5	418.9	413.4	406.1	399.8	388.8
650	420.4	418.0				
700	419.1	416.7	410.9	403.3	396.9	385.7
800	415.2	412.8	407.4	399.7	392.9	381.7
900	409.7	408.1	402.4	394.9	388.5	377.5
1000	402.6	401.3	396.5	389.5	383.8	372.2
1100	393.2	392.7	389.3	382.6	376.9	366.0
1200	382.2	382.0	379.9	374.9	369.7	359.2

Electrocapillary data for adsorption of furfurylamine in 0.3 M KCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ/mN m^{-1}					
c/mol l^{-1}	0.01	0.02	0.05	0.1	0.2	0.5
-E/mV						
100	382.3	381.2	380.7	378.9	372.6	369.0
200	399.9	398.6	398.1	396.1	390.7	386.0
250						389.8
280						390.7
300	411.6	410.2	408.6	404.5	399.0	391.2
320						391.3
340					400.8	391.3
360					401.0	391.3
380				406.8	401.0	391.0
400	418.7	416.6	413.0	407.2	401.0	390.9
420			413.3	407.1	401.0	
440		418.0	413.5	407.1	401.0	
450	420.5					390.1
460		418.3	413.7	407.2	400.8	
480	421.1	418.5	413.7	407.0		
500	421.5	418.7	413.7	406.8	400.1	389.3
520	421.7	418.7	413.5			
540	421.7	418.7				
550			413.1	405.2		
560	421.5	418.6				
580		418.3				
600	421.0	418.0	411.8	404.4	397.6	386.4
650	420.1					
700	418.5	415.2	408.8	401.1	394.7	383.1
800	413.9	410.9	404.7	397.0	390.7	378.8
900	407.6	405.3	399.6	392.3	385.7	373.8
1000	399.6	397.7	393.1	386.2	380.2	368.0
1100	389.5	388.5	385.2	379.2	373.4	361.4
1200	377.7	377.3	375.1	370.3	365.8	353.8

Electrocapillary data for adsorption of furfurylamine in 1.0 M KCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25 °C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ/mN m^{-1}									
c/mol l^{-1}	0.001	0.003	0.01	0.02	0.03	0.05	0.1	0.2	0.3	0.5
-E/mV										
0	348.0	345.7								
50			365.6	364.0		360.7	357.9			
100	379.7	379.2	378.6	377.0	375.6	375.4	372.8	368.1	346.8	363.3
200	397.1	396.7	396.6	395.2	394.3	393.4	391.3	386.8	384.8	381.7
250										386.0
300	410.3	409.7	409.0	407.4	406.2	404.8	400.6	396.2	393.4	388.2
320										388.4
340							403.3	397.3		388.5
350							404.5		394.7	
360								397.7		388.4
380							404.5	397.8	394.8	388.2
400	418.7	418.2	417.0	414.7	413.0	410.7	405.4	397.8	394.7	388.1
420							405.5	397.8	394.4	
440					414.0	411.3	405.5	397.7		
450	421.6	420.6	419.1	416.4					394.1	387.2
460					414.5	411.6	405.5			
480				417.0	414.7	411.6				
500	423.3	422.0	420.4	417.1	414.7	411.4	404.9	396.5	392.9	386.1
520	423.5	422.4	420.7	417.1	414.7	411.1				
540	423.8	422.7	420.7	416.9						
550					414.4	410.7	404.4		392.0	
560	423.9	422.7	420.5	416.7						
580	423.8	422.5	420.2		414.0					
600	423.4	422.3	419.7	416.1		409.6	402.5	394.1		383.4
620					412.9					
650	422.4	421.1	418.6						388.9	
700	420.5	418.7	416.8	412.8	410.4	406.3	399.2	390.6		379.7
750									384.9	
800	414.6	413.3	411.8	408.4	405.9	402.0	394.8	386.4	382.6	375.4
900	406.4	405.7	404.8	402.1	399.9	396.4	389.7	381.1	377.7	370.6
1000	396.4	395.6	395.8	394.0	392.3	389.5	383.2	375.3	371.8	364.4
1100	384.8	384.1	384.8	383.6	382.8	380.9	376.3	368.4	364.9	358.0
1200	371.4	370.9	372.0	371.3	370.7	370.2	366.5	360.5	357.1	350.6

Electrocapillary data for adsorption of furfurylamine in 3.0 M KCl.
Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}					
c/mol l^{-1}	0.01	0.02	0.05	0.1	0.2	0.5
-E/mV						
0	333.3					
100	371.9	371.9	368.7	367.2	362.2	350.6
200	392.2	390.2	387.5	385.3	380.8	374.3
250						378.7
300	403.1	402.4	398.9	395.4	390.0	380.2
320						380.8
340					390.8	380.8
350			402.2	397.7		
360					390.9	380.8
380				398.0	390.9	380.7
400	411.2	409.9	404.3	398.3	390.8	380.4
420				398.2	390.8	
440			404.8	398.0		
450	413.5	411.7			390.2	379.6
460			405.0	397.7		
480		412.1	404.9			
500	415.0	412.4	404.5	396.9	389.4	378.5
520	415.1	412.4				
540	415.1	412.1				
550			403.7	396.0		
560	415.1	411.9				
580	414.8					
600	414.6	422.3	402.4	394.1	386.5	375.6
650	413.2	409.5				
700	411.4	407.8	398.8	391.1	382.9	372.2
800	406.4	403.4	394.3	386.4	378.6	367.8
900	399.6	397.6	388.9	380.9	373.8	362.7
1000	390.5	389.6	382.3	375.0	367.6	357.6
1100	379.3	379.0	374.4	367.9	360.9	350.9
1200	365.9	366.5	363.9	358.9	353.4	343.0

Electrocapillary data for adsorption of tetrahydrofurfurylamine in 1.0 M KCl - given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ/mN m^{-1}					
c/mol l^{-1}	0.003	0.01	0.03	0.1	0.3	0.5
-E/mV						
50	367.4	367.3	366.8			
100	379.2	379.3	379.6	376.9	373.7	370.6
200	396.6	396.9	397.2	395.3	393.5	391.3
250						396.8
300	409.3	409.7	409.7	406.7	403.3	399.4
340					404.7	400.1
350				409.9		
360					404.9	400.0
380				410.7	404.8	399.9
400	418.0	418.1	417.3	410.9	404.6	399.7
420				410.8		
440				410.6	404.0	
450	420.8	420.6	418.8			398.5
460			419.0			
470				410.0		
480			419.0		402.9	
500	422.4	421.8	419.0	409.2		396.9
520	422.8	422.0	418.8			
540	423.0	422.0				
550			418.0	407.7	400.3	
560	423.0	421.8				
580	422.9					
600	422.6	421.0	416.5	405.5		392.5
620					397.1	
650	421.6	419.4				
700	419.1	417.0	411.8	400.6	393.0	387.3
800	413.1	410.9	405.7	394.6	387.1	381.4
900	404.9	403.2	398.5	387.7	380.2	374.6
1000	395.0	393.8	390.0	379.9	372.5	367.1
1100	383.6	383.0	380.2	371.0	364.1	359.0
1200	370.4	370.3	368.2	361.4	355.0	349.8

Electrocapillary data for adsorption of N-methylfurfurylamine in 1.0 M KCl - given are the interfacial tension in mN m^{-1}, potential in mV. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}					
c/mol l^{-1}	0.001	0.003	0.01	0.03	0.1	0.3
-E/mV						
0	349.4					
50		367.1	365.3	361.3		
100	380.3	379.2	377.7	374.1	369.3	362.1
200	397.7	396.6	395.3	392.1	387.1	380.0
250						384.0
280						384.9
300	410.4	409.3	407.9	403.5	395.8	385.1
320						385.1
340						385.1
350				406.8	396.6	
360						385.0
380				407.6	396.5	384.7
400	419.2	417.4	414.7	407.7	396.2	384.5
420				407.5	396.0	
440				407.3		
450	421.6	419.9	416.1			383.5
460					395.4	
480			416.3	407.0		
500	423.4	421.2	416.3		394.5	382.3
520	423.6	421.5	416.0			
540	423.8	421.6	415.6	405.6		
550					393.1	
560	423.8	421.3				
580	423.7	421.1	414.8			
600	423.5	420.9		403.9	391.5	379.5
650	422.3	419.6	412.9			
700	420.3	417.5	410.7	400.5	388.2	375.9
800	414.4	412.3	406.4	396.2	384.0	371.8
900	406.6	404.8	400.6	391.1	379.1	366.9
1000	396.6	395.4	393.0	385.3	373.6	361.6
1100	385.0	384.1	383.1	378.2	367.0	355.2
1200	372.1	370.7	370.7	368.7	359.7	348.5

Electrocapillary data for adsorption of N-methyltetrahydrofurfurylamine in 1.0 M KCl - given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}					
c/mol l^{-1}	0.001	0.003	0.01	0.03	0.1	0.3
-E/mV						
50	367.7	368.1	367.7	366.5		
100	379.3	379.9	379.5	379.1	376.9	370.5
200	396.8	397.3	397.1	397.2	394.3	389.5
250						394.0
280						395.3
300	409.5	410.1	409.8	409.1	404.1	395.7
320					403.2	395.5
340					404.9	395.3
350				412.7		
360					404.7	395.1
380				413.7	404.6	
400	418.3	418.5	417.6	413.8	404.2	394.2
420				413.6		
440				413.3		
450	421.3	421.1	419.1		403.2	392.8
460				413.0		
480			419.3			
500	422.7	422.7	419.0	411.9	401.3	391.3
520	423.2	422.8	418.7			
540	423.3	422.8	418.1			
550				410.0		
560	423.3	422.8				
580	423.2	422.5	416.9			
600	422.9	422.0		408.1	397.5	387.3
650	421.6	420.5	414.0			
700	419.6	418.3	411.8	403.6	392.8	382.7
800	413.5	412.4	406.1	398.0	387.5	377.3
900	405.5	404.7	399.4	391.9	381.4	371.3
1000	395.4	395.1	391.5	384.9	374.8	364.7
1100	382.0	384.1	382.1	376.8	367.1	357.3
1200	370.7	371.1	370.0	367.3	358.7	349.3

Electrocapillary data for adsorption of furfuryl alcohol in 1.0 M KCl pH = 10
Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}					
c/mol l^{-1}	0.001	0.003	0.01	0.03	0.1	0.3
-E/mV						
0	349.1	349.4	349.1	350.6		
10					353.0	348.3
100	380.1	380.1	378.5	378.2	373.9	367.8
200	397.1	397.2	394.9	393.5	387.9	380.2
260						383.0
280						383.1
300	409.8	409.6	406.5	404.0	395.5	383.8
320						384.1
340						384.0
360						384.0
380						384.2
400	418.5	417.4	414.1	408.7	397.4	384.5
420				409.0		384.6
440				409.0	398.0	384.5
450	420.8	420.2	415.5	408.5		
460				409.0	398.0	384.6
480				408.5	397.8	
500	423.0	421.8	416.5	409.8	397.7	384.1
520	423.0	422.0	416.9			
540	423.3	422.2	416.9	409.9	397.3	
550				410.1		383.3
560	423.4	422.2	416.9			
580	423.4	422.2	416.9	410.0		
600	423.2	422.1	416.8	409.1	396.9	382.0
650	422.2	420.8	415.8			
700	420.2	419.5	414.7	407.2	394.5	379.8
800	414.5	413.9	409.7	403.8	391.5	377.1
900	406.7	406.6	403.0	399.0	387.2	373.4
1000	396.4	396.9	394.6	392.1	382.1	369.3
1100	385.4	385.1	383.3	382.6	376.1	364.7
1200	372.1	372.4	370.8	371.1	367.3	358.7

Electrocapillary data on mercury electrode for tetrahydrofurfuryl alcohol (at various molarities) in 1.0 M Na_2SO_4 at 25°C. Given are the interfacial tension in mN m^{-1}, potential in mV Potentials relative to internal Hg/Hg_2SO_4 ref. electrode.

Reference: J.M. Sedlak, unpublished data.

			γ /mN m^{-1}			
c/mol l^{-1}	0.01	0.03	0.05	0.10	0.20	0.50
-E/V						
300	372.3					
350	382.7					
375						
400	392.3	392.7	392.5	392.1	393.0	391.9
425						
450	400.7	400.8	400.8	400.5	399.6	395.4
475						
500	407.7	407.7	407.5	406.2	403.6	397.3
525						397.8
550	413.5	413.2	412.5	409.9	405.9	398.1
575				411.1		398.2
600	417.9	417.2	415.9	412.0	406.8	398.1
625				412.6	407.0	397.9
650	421.1	419.9	417.9	413.0	406.9	397.7
675				413.1	406.7	397.3
700	423.3	421.4	418.7	413.0	406.4	396.8
725	424.0	421.7	418.7	412.8	405.9	396.3
750	424.5	421.9	418.6	412.5	405.4	395.7
775	424.9	421.8	418.2	411.4	404.8	395.1
800	425.0	421.6	417.8	410.1	404.2	394.4
825	425.0	421.2	417.2			
850	424.8	420.7	416.6	410.1	402.7	392.9
875	424.5	420.0	415.8			
900	424.0	419.3	415.0	408.4	401.0	391.1
950	422.7	417.6	413.2	406.5	399.1	389.2
1000	421.0	415.6	411.1	404.4	397.0	387.0
1050	418.8	413.3	408.8	402.1	394.6	384.6
1100	416.1	410.7	406.1	399.6	392.1	382.1
1150	412.9	407.8	403.2	396.8	389.3	379.3
1200	409.3	404.6	400.0	393.9	386.4	376.4
1250	405.2	400.9	396.5	390.7	383.2	373.2
1300	400.6	396.8	392.8	387.3	379.9	370.0
1350	395.6	392.3	388.8	383.7	376.4	366.5
1400	390.1	387.5	384.5	379.9	372.7	362.9
1450	384.3	382.4	379.8	375.8	368.8	359.0
1500	378.3	377.3	374.4	371.3	364.3	354.7

Electrocapillary data on mercury electrode in 1.0 M Na_2SO_4 at 25°C. Adsorption of tetrahydrofurylamine - given are the interfacial tension in mN m^{-1}, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas and J.M. Sedlak. Electrochimica Acta, 16, (1971) 2091.

	γ /mN m^{-1}									
c/mol l^{-1}	0	0.004	0.01	0.02	0.03	0.04	0.05	0.075	0.1	0.15
-E/V										
.300	372.8									
.350	383.2									
.375		388.1	388.0	386.7						
.400	393.1	392.8	392.7	392.2	391.3	391.0	390.4	390.0	386.7	
.425					396.2	396.1	395.2	394.7	394.2	393.6
.450	401.4	400.9	400.7	400.4	400.1	399.9	399.0	398.9	397.9	397.1
.500	408.3	408.0	407.8	407.1	406.8		405.5	405.1	403.6	402.6
.525						409.2	408.2	407.3	405.2	404.1
.550	414.5	413.8	413.5	412.6	412.1		410.4	409.0	406.2	405.1
.575						413.1	411.8	410.2	407.0	405.5
.600	419.4	418.4	418.0	416.5	415.6		413.1	411.1	407.4	405.7
.625		420.4	419.7	418.1	416.3	415.4	413.6	411.5	407.5	405.7
.650	422.6	422.0	421.3	419.2	417.1	415.8	413.9	411.5	407.5	405.5
.675		423.4	422.3	419.9	417.4	415.8	414.0	411.4	407.5	405.3
.700	425.0	424.4	423.2	420.2	417.3	415.6	414.0	411.2	407.2	405.0
.725	426.0	425.3	423.7	420.2	417.3	415.4				
.750	426.6	425.8	423.9	420.1	417.0	415.0	413.1	410.4	406.4	403.8
.775	427.3	426.0	423.9	419.7	416.7	414.6				
.800	427.8	426.3	423.8	419.4	416.3	414.2	412.1	409.0	405.2	402.6
.825	428.1	426.3	423.4	419.0	415.7	413.6				
.850	428.1	426.2	423.0	418.5	415.0	412.8	411.7	407.4	403.5	400.7
.875	427.8	425.8	422.5			411.8				
.900	427.2	425.3	421.6	416.9	413.6	409.8	409.0	405.6	401.5	398.7
.950	426.2	424.1	420.0	415.1	411.5	407.7	406.9	403.5	399.3	396.7
1.000	424.5	422.1	417.9	412.6	409.4	405.0	404.8	401.0	396.9	394.3
1.050	422.5	419.9	415.5	410.2	406.8	402.3	402.2	398.4	394.5	391.5
1.100	420.0	416.9	412.4	407.6	404.0	399.4	399.5	395.7	391.6	388.8
1.150	416.7	413.3	409.2	404.6	400.8	396.3	396.6	392.7	388.9	386.1
1.200	412.7	409.5	405.9	401.3	397.6	392.8	393.3	389.4	385.9	383.1
1.250	408.3	405.7	402.2	397.7	394.2	389.2	389.8	386.0	382.6	379.7
1.300	403.5	401.7	398.3	393.8	390.4	385.1	386.0	382.7	379.0	376.1
1.350	397.7	396.9	393.9	389.6	386.2	381.0	382.1	378.6	375.2	372.3
1.400	391.8	391.6	389.0	385.2	381.9	376.7	378.1	374.6	371.2	368.2
1.450	386.0	386.0	383.9	380.3	377.1	372.0	373.4	370.6	367.1	364.0
1.500	379.1	379.9	378.0	375.2	372.2	367.3	368.7	365.9	362.6	359.8
1.550								361.0	357.8	355.5
1.600								355.6	352.9	350.6

Electrocapillary data on mercury electrode in 1.0 M Na_2SO_4, pH adjusted to 11.6 at 25°C. Adsorption of furfuryl alcohol.
Given are the interfacial tension in mN m^{-1}, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas and J.M. Sedlak. Electrochimica Acta 16 (1971) 2091.

	γ /mN m^{-1}								
c/mol l^{-1}	0	0.01	0.02	0.03	0.04	0.05	0.075	0.1	0.15
-E/V									
.350				381.9	381.1				
.375				387.0	385.8				
.400	391.2			391.1	389.6				
.425	-	395.2	395.2	395.1	393.9	394.1	392.5	390.6	388.3
.450	402.0	399.3	399.1	398.3	396.9	397.1	395.3	393.4	390.9
.475	-	-	-	-	-	-	-	395.4	392.4
.500	408.9	406.0	405.2	404.1	402.2	402.2	399.6	397.0	393.5
.550	414.7	411.3	409.8	408.6	405.9	405.8	402.2	399.1	395.5
.600	419.1	415.4	413.2	411.1	408.0	407.8	403.7	400.7	396.6
.650	422.6	418.3	415.3	412.6	409.5	409.0	404.5	401.3	397.0
.675	-	419.2	415.9	413.1	410.1	409.4	405.0	401.6	397.5
.700	425.2	420.2	416.4	413.7	410.6	409.8	405.4	401.8	397.9
.725	-	420.9	417.0	414.1	411.0	410.0	405.9	401.9	397.9
.750	426.6	421.4	417.4	414.4	411.2	410.2	406.0	402.0	397.9
.775	-	421.8	417.7	414.7	411.2	410.2	406.0	402.0	397.8
.800	427.5	421.9	418.0	414.9	411.2	410.1	405.9	402.0	397.8
.825	427.8	422.0	418.0	414.9	411.2	409.9	405.8	401.9	397.6
.850	427.9	421.9	417.9	414.6	411.1	409.6	405.6	401.8	397.4
.875	427.8	421.8	417.8	414.3	411.0	409.3	405.2	401.7	397.1
.900	427.4	421.6	417.6	413.9	410.6	409.0	404.8	401.4	396.7
.950	426.2	420.5	416.6	413.0	410.0	408.6	404.1	400.4	395.8
1.000	425.7	419.5	415.5	412.1	409.3	407.6	403.1	399.4	394.4
1.050	422.4	417.9	414.2	411.0	408.0	406.2	401.9	397.7	393.0
1.100	419.5	415.7	412.6	409.3	406.6	404.7	400.2	396.2	391.6
1.150	416.3	413.4	410.4	407.2	404.6	403.0	398.5	394.7	389.8
1.200	412.4	410.1	407.7	404.9	402.4	401.2	396.4	393.3	388.0
1.250	408.2	406.4	404.7	402.4	399.9	398.8	394.3	391.3	386.1
1.300	403.5	402.3	400.9	399.3	397.0	396.2	390.0	389.2	384.0
1.350	398.4	397.6	396.7	395.5	393.3	393.0	389.4	386.4	381.3
1.400	392.8	392.8	391.9	390.9	389.3	389.1	386.0	383.4	378.5
1.450	386.7	386.9	386.4	385.7	384.7	384.6	382.2	379.8	375.7
1.500	380.3	380.9	380.5	380.0	379.1	379.1	377.6	375.7	372.7
1.550	-	374.2	373.9	373.4	373.1	373.1	372.6	371.1	368.4

Electrocapillary data on mercury electrode in 1.0 M Na_2SO_4 at 25°C. Adsorption of furfurylamine - given are the interfacial tension in mN m^{-1}, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas and J.M. Sedlak. Electrochimica Acta, 16 (1971) 2091.

	γ /mN m^{-1}							
c/mol l^{-1}	0.004	0.01	0.02	0.03	0.05	0.1	0.2	0.3
-E/V								
.350	382.0	380.0	378.7	377.6	374.9	372.9		
.400	392.3	390.5	389.1	387.9	386.2	383.4	383.7	382.8
.425							386.7	384.3
.450	399.9	398.6	397.4	396.1	394.2	391.4	388.3	385.2
.475							389.5	385.8
.500	406.4	405.1	403.7	402.1	399.3	395.8	390.3	386.1
.525							390.4	
.550	412.4	410.5	408.3	406.3	403.2	397.5	390.5	386.1
.575							390.5	
.600	416.8	414.3	411.4	408.6	404.8	398.1	390.5	386.1
.625	418.0	415.3		409.2	405.4	398.4	390.5	386.1
.650	419.8	416.2	413.0	409.8	405.7	398.6	390.5	385.9
.675	420.9	417.1	413.5	410.2	405.7	398.7	390.2	385.6
.700	421.6	417.6	413.9	410.5	405.8	398.7	389.8	385.1
.725	422.5	417.9	414.3	410.6	405.7	398.3	389.5	384.6
.750	423.1	418.2	414.3	410.6	405.6	397.9	389.1	384.1
.775	423.3	418.4	414.1	410.4	405.3	397.5	388.8	383.7
.800	423.4	418.5	413.9	410.1	404.9	397.1	388.2	383.3
.825	423.4	418.3	413.7	409.7	404.5	396.8	387.8	382.8
.850	423.3	417.9	413.3	409.2	404.0	396.4	387.3	382.0
.875	423.1	417.4	412.9	408.8	403.6	395.8	386.7	381.2
.900	422.5	417.0	412.5	408.4	403.0	395.2	385.8	380.4
.950	421.6	416.0	411.4	407.1	401.7	393.9	384.2	378.9
1.000	420.0	414.7	410.0	405.6	400.1	392.3	382.5	377.3
1.050	418.0	412.9	408.1	403.9	398.4	390.4	380.7	375.4
1.100	415.7	410.9	406.1	401.9	396.3	388.4	378.6	373.2
1.150	412.9	408.3	403.8	399.7	394.1	386.1	376.3	370.9
1.200	409.5	405.6	401.3	397.2	391.7	383.7	374.1	368.5
1.250	405.7	402.3	398.2	394.4	389.1	381.2	371.2	365.8
1.300	401.7	398.5	394.9	391.1	386.0	378.4	368.3	363.0
1.350	397.1	394.6	391.2	387.8	382.9	375.4	365.5	359.9
1.400	392.1	389.9	387.2	384.1	379.4	372.0	362.1	356.8
1.450	386.3	384.5	382.6	380.0	375.6	368.5	358.6	353.4
1.500	380.1	379.1	377.7	375.5	371.5	364.7	355.2	

Electrocapillary data on mercury electrode in 1.0 M Na_2SO_4 at 25°C. Adsorption of tetrahydrofuran - given are the interfacial tension in $mN\ m^{-1}$, potential in V vs internal Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas and J.M. Sedlak. Electrochimica Acta 16 (1971) 2091.

	γ /mN m^{-1}							
c/mol l^{-1}	0.01	0.03	0.05	0.075	0.1	0.15	0.20	0.30
-E/V								
.375	388.2	388.2	388.5	388.7	388.4	388.2	388.2	388.1
.400	392.4	392.4	392.5	392.7	392.7	392.4	392.5	392.1
.425					388.3	396.5	396.4	395.8
.450	400.9	400.8	401.0	400.7	392.6	400.0	399.8	397.4
.475					396.9	413.2	402.2	397.7
.500	407.6	407.7	407.8	407.0	400.5	405.0	402.9	397.9
.525					403.9	406.1	403.1	397.9
.550	413.6	413.4	413.0	412.0	406.8	406.4	403.3	397.9
.575				413.9	409.5	406.7	403.3	397.9
.600	418.1	417.5	416.5	414.6	411.1	406.7	403.0	397.7
.625		419.2	418.0	414.8	411.6	406.6	402.7	397.5
.650	421.6	420.4	418.5	414.8	411.6	406.2	402.3	397.1
.675	423.0	421.5	418.5	414.7	411.7	405.7	402.1	396.7
.700	424.5	422.2	418.4	414.5	411.6	405.0	401.6	395.6
.725	425.1	422.5	418.4	414.2	411.6			
.750	425.6	422.6	418.0	413.7	411.2	403.9	400.5	394.5
.775	426.0	422.5	417.6	413.1				
.800	426.1	422.1	417.2	412.5	409.0	402.9	399.2	393.4
.825	426.2	421.5	416.5					
.850	426.2	420.9	415.8	411.4	407.4	401.5	397.6	392.2
.900	425.4	419.9	414.1	409.7	405.7	399.8	395.8	390.5
.950	424.4	418.5	412.4	407.9	403.8	398.2	394.0	388.6
1.000	422.7	416.4	410.6	406.0	401.1	396.4	391.8	386.5
1.050	420.6	414.2	408.2	403.8	399.5	394.0	389.5	384.5
1.100	418.0	411.9	405.9	401.4	397.1	391.7	387.3	382.0
1.150	414.8	409.5	403.3	398.9	394.7	389.2	384.9	379.4
1.200	411.2	406.6	400.5	396.2	392.1	386.5	382.3	376.8
1.250	407.2	403.2	397.4	393.3	389.4	383.4	379.2	373.9
1.300	402.7	399.6	394.5	390.3	386.1	380.5	376.2	370.9
1.350	397.9	395.8	391.5	386.9	383.0	377.4	373.1	367.8
1.400	392.4	391.2	387.9	383.8	379.8	374.4	370.0	364.5
1.450	386.6	385.7	383.7	380.0	376.3	370.7	366.5	361.4
1.500	380.4	379.8	378.4	375.7	372.4	367.1	363.0	357.7

Electrocapillary data for adsorption of Aniline in 1.0 M HCl

T = 25°C

Reference: R.G. Hamilton. Ph.D. Thesis, Toronto, 1967.

	γ /mN m^{-1}			
c/mol l^{-1}	0.012	0.036	0.101	0.404
-E/mV				
0	350.8	349.1	345.2	340.6
100	377.7	376.2	372.2	365.3
200	394.3	392.1	388.1	382.8
300	406.5	404.0	400.0	393.8
400	414.7	412.4	407.9	401.2
450	417.5	415.2	410.5	403.5
500	419.4	417.0	412.2	404.9
520	419.8	417.6	412.7	405.2
540	420.0	418.0	413.0	405.4
560	420.4	418.1	413.2	405.5
580	420.3	418.1	413.2	405.3
600	420.2	418.0	413.1	405.1
650	419.3	417.2	412.4	404.1
700	417.6	414.5	410.7	402.2
800	411.9	410.1	405.6	397.2
900	404.4	402.6	398.3	390.2
1000	394.9	393.2	389.1	381.8
1100	383.5	382.3	378.5	371.8
1200	371.0	370.5	366.6	360.2

Electrocapillary data on mercury electrode for adsorption of pyridine in 0.1 M KCl - given are the interfacial tension in mN m^{-1}, potential in mV vs SCE.

T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

				γ /mN m^{-1}				
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5	1.0
-E/mV								
0	364.1							
30	371.6							
50		375.4	373.8					
100	385.8	384.4	383.0	382.5	381.1	379.9	374.1	369.4
200	400.7	398.4	397.0	396.7	394.2	392.7	387.5	382.6
250								385.8
300	411.7	409.5	407.3	406.4	403.2	400.9	394.6	387.8
350								388.5
380								388.7
400	420.1	417.2	414.6	412.8	408.7	405.2	396.4	388.7
420							396.7	388.7
440						406.0	396.7	388.6
450	422.1				410.6			
460						406.3	396.5	388.4
480						406.4	396.4	
500	424.9	421.7	418.6	416.0	411.4	406.4	396.2	388.0
520					411.4	406.4		
540		422.7	419.4	416.5	411.4	406.3		
550	425.9						395.6	
560	426.5	423.1	419.7	416.7	411.4	406.1		
580	426.5	423.4	419.9	416.7	411.4			
600	426.7	423.4	419.8	416.7	411.3	405.5	394.5	386.0
620	426.8	423.4	419.8	416.6				
640	426.8							
650		423.4	419.7	416.4	410.8			
660	426.6							
700	426.3	422.9	419.1	415.7	409.8	403.4	391.4	382.7
800	423.8	420.7	417.1	413.1	406.9	399.4	386.7	378.1
900	419.3	416.9	413.1	409.5	402.7	394.4	381.1	372.6
1000	413.1	411.5	408.0	404.8	397.3	388.3	374.9	366.0
1100	404.9	403.8	401.3	398.7	390.7	381.0	367.5	358.7
1200	395.2	394.7	392.4	391.0	383.2	373.4	359.5	350.9

Electrocapillary data on mercury electrode for adsorption of pyridine in in 0.03 M KCl. Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE.

T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}							
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5	1.0
-E/mV								
0	364.5	363.7	361.5	353.6				
50					372.3	366.7		
100	386.8	386.8	386.0	383.4	382.6	379.6	374.6	368.6
200	401.7	401.3	400.4	397.6	395.8	392.5	388.1	382.5
300	413.2	412.0	410.8	407.0	404.4	400.4	394.2	387.5
360								388.2
380						403.4	395.7	388.2
400	421.4	418.9	417.2	412.5	409.2	403.8	395.7	388.2
420						404.2	395.7	388.2
440					409.7	404.2	395.5	387.9
450	423.1			413.9				
460					410.1	404.2	395.2	387.7
480					410.2	404.2		
500	425.5	422.4	420.0	414.9	410.2	404.2	394.8	386.9
520	426.0	422.6	420.3	414.9	410.2	404.0		
540	426.5	422.8	420.5	414.9	410.2	403.8		
560	426.6	422.9	420.6	414.8	410.0			
580	426.8	422.9	420.6	414.7				
600	426.8	422.9	420.6	414.5	409.9	402.8	392.3	384.4
620	426.7	422.8	420.4					
640	426.7	422.6	420.1					
650	426.1			413.8				
700	425.3	421.6	419.1	412.8	407.5	399.5	388.6	380.6
800	421.8	418.7	416.0	410.1	403.8	394.9	383.3	375.4
900	416.4	414.2	411.8	405.8	398.9	389.1	377.3	369.2
1000	409.3	407.4	405.7	400.2	392.8	382.3	370.3	362.3
1100	400.2	399.2	398.0	393.3	385.7	374.3	362.6	354.8
1200	389.6	388.9	388.1	384.8	377.1	366.1	354.4	346.7

Electrocapillary data for adsorption of pyridine in 0.1 M KCl - given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}							
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5	1.0
-E/mV								
0	361.8	359.4						
50			375.6	374.1	371.4	367.4	357.4	351.5
100	386.3	385.7	385.6	384.1	381.5	379.4	374.5	369.1
200	402.4	401.3	400.9	398.7	395.6	393.0	388.0	383.3
250								386.9
300	413.9	412.3	411.5	408.6	405.0	401.6	394.8	388.8
320								388.9
340							395.6	389.0
350						403.8		
360							395.7	389.0
380							395.8	388.9
400	421.7	419.1	417.9	414.2	409.7	404.9	395.8	388.8
420						405.2	395.7	
440						405.1		
450	424.0	421.0	419.5	415.4	410.7		395.1	388.1
460						405.1		
480					410.8	404.9		
500					410.8	404.5	394.3	387.0
520					410.7			
540					410.5			
550						403.7		
560	426.2	422.6	420.7	415.9	410.0			
580	426.2	422.5	420.6	415.7				
600	426.1	422.2	420.5	415.6	409.7	402.2	391.0	383.6
620	425.9	422.1						
650	425.3	421.5	419.5	414.5	408.1			
700	423.9	420.2	418.0	413.1	406.4	398.3	386.5	379.2
800	419.4	416.4	414.7	409.5	401.8	393.1	380.7	373.5
900	413.1	410.8	409.4	404.2	396.2	386.7	374.3	366.9
1000	405.0	403.2	402.4	398.0	389.3	379.6	366.9	359.3
1100	395.1	394.3	393.6	390.6	380.4	371.6	358.8	351.3
1200	383.5	383.1	383.0	381.2	373.0	363.1	350.6	342.6

Electrocapillary data for adsorption of pyridine in 0.3 M KCl - given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}							
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5	1.0
-E/mV								
0	357.6							
50	372.6							
100	383.8	383.4	382.7	381.3	378.9	376.9	371.8	367.5
200	400.2	399.6	398.6	396.8	393.7	391.1	385.6	381.2
300	412.4	411.1	409.5	406.8	403.2	399.7	392.6	386.5
320								386.5
340							393.7	386.6
360					406.7	402.4	393.7	386.6
380					407.3	403.0	393.7	386.5
400	420.5	418.3	416.2	412.7	408.2	403.2	393.5	386.3
420					408.4	403.4	393.3	
440				414.0	408.8	403.4		
450	423.0	420.3	418.0					385.3
460				414.4	408.9	403.2	392.5	
480				414.6	408.9			
500	424.7	421.6	418.8	414.7	408.9	402.5	391.5	383.9
520	425.1	421.8	419.0	414.7	408.8			
540	425.3	421.8	419.1	414.6				
560	425.3	421.8	419.0	414.5	408.3			
580	425.3	421.7	418.6					
600	425.2	421.6	418.5	413.6	407.3	399.5	387.5	380.0
620	424.9							
650	424.3	420.3	417.5					
700	422.4	419.0	416.1	411.0	403.7	394.9	382.5	375.0
800	417.2	414.5	411.7	406.8	398.4	388.9	376.3	368.7
900	410.0	408.0	405.7	403.0	391.7	381.8	369.3	362.0
1000	400.7	399.6	397.9	394.1	384.4	374.1	361.6	354.4
1100	390.1	389.5	388.1	385.9	376.3	366.0	353.6	346.1
1200	377.8	377.5	376.3	375.8	367.7	357.7	344.8	337.2

Electrocapillary data for adsorption of pyridine in 1.0 M KCl - given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ /mN m^{-1}						
c/mol l^{-1}	0	0.001	0.003	0.01	0.02	0.03	0.05
-E/mV							
0	349.7	349.1	349.5	349.0		347.9	
40							
50				368.1	367.5		367.0
100	379.8	379.3	379.6	379.9	379.2	378.6	378.3
200	397.4	396.7	396.9	396.8	395.7	394.7	394.1
250							
260							
280							
300	410.1	409.4	409.7	408.8	407.4	406.3	404.9
320							
340							
350							
360							
380							
400	418.9	418.2	417.9	416.6	414.8	413.1	411.4
420							
440							
450	421.7	420.9	420.6	419.0	416.8	414.8	
460							413.1
470							
480							413.4
500	423.4	422.7	422.2	420.3	416.9	415.7	413.5
520	424.0	423.2	422.7	420.7	418.0	415.8	413.4
540	424.2	423.4	422.9	420.9	418.1	415.8	413.3
550							
560	424.2	423.4	422.9	420.9	418.1	415.7	413.0
580	424.2	423.3	422.8	420.7	418.0	415.5	
600	423.9	423.0	422.6	420.4	417.5	415.1	412.1
650	422.7	421.9	421.5	419.3	416.4	414.0	410.7
700	420.7	419.8	419.3	417.4	414.5	412.1	408.6
800	414.7	414.1	413.8	412.2	409.7	407.0	403.6
900	406.6	406.0	406.0	405.0	403.1	400.1	397.0
1000	396.5	395.9	395.9	395.7	394.3	392.0	389.5
1100	384.6	384.1	384.5	384.5	383.9	382.3	381.1

Adsorption of pyridine in 1.0 M KCl (continued)

	γ /mN m^{-1}						
c/mol l^{-1}	0.1	0.2	0.3	0.5	1.0	1.75	3.50
-E/mV							
0	343.1	329.7				305.1	
40					339.1		
50							
100	376.2	373.5	372.6	368.4	363.7	359.5	354.6
200	391.8	388.5	387.0	383.4	378.9	375.5	371.6
250				387.4	382.9	379.3	
260					383.3		376.3
280					384.1		377.3
300	401.8	397.8	395.5	390.2	384.6	380.8	378.0
320				390.6	384.9	381.0	378.3
340		400.0		390.8	385.0	381.0	378.5
350	405.3		397.3				
360		400.7		390.7	385.0	381.0	378.4
380		401.1	397.6	390.6	384.8	380.9	378.2
400	407.3	401.3	397.5	390.4	384.4	380.6	378.0
420	407.8	401.3	397.3	390.1			
440	408.1	401.1	397.0				
450				389.3	383.1	379.5	376.9
460	408.3	400.8	396.2				
470							
480	408.2						
500	408.0	399.7	395.1	387.8	381.3	377.8	375.2
520							
540							
550	407.3	398.0	393.2				
560							
580							
600	405.0	395.9	390.7	383.4	376.9	373.7	370.6
650							
700	401.4	390.3	385.2	377.8	371.6	368.2	365.0
800	394.2	384.0	378.6	371.2	365.0	361.4	358.5
900	387.1	376.5	371.2	364.1	357.9	354.0	351.3
1000	379.1	368.7	363.2	356.1	349.7	346.1	343.2
1100	371.0	360.1	354.7	347.2	341.3	337.4	334.5
1200	361.8	351.2	345.3	338.2	332.1	327.3	325.4

Electrocapillary data for adsorption of pyridine in 3.0 M KCl.
Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T=25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

				γ /mN m^{-1}				
c/mol l^{-1}	0	0.01	0.02	0.05	0.1	0.2	0.5	1.0
-E/mV								
0	338.5	338.3						
50			361.0	358.1	357.5			
100	374.1	373.8	373.7	371.1	369.9	367.4	362.2	355.2
150								365.2
200	392.5	391.6	391.1	388.0	386.1	382.7	377.4	371.1
240								373.9
250						388.0	381.2	
260								374.7
280								375.5
300	406.0	404.5	403.4	399.2	396.4	391.8	382.7	375.7
320							382.9	375.8
340						393.2	382.9	375.8
350					399.8			
360						393.2	382.7	375.7
380						393.1	382.5	375.5
400	415.4	413.0	411.3	406.1	401.6	392.9	382.0	375.2
420					401.8			
440					401.7	391.8		
450	418.4	415.6	413.6	407.4			380.6	
460					401.2			
480				407.6				
500	420.8	417.2	414.8	407.6	400.1	389.5	378.6	372.4
520	421.4	417.4	415.0	407.4				
540	421.8	417.7	415.1	406.9				
560	421.9	417.7	415.0	406.1				
580	421.9	417.8	414.9					
600	421.9	417.5	414.1	404.7	395.3	384.7	373.9	368.0
620	421.6							
650	421.2	416.6	413.4	402.2				
700	419.3	414.7	411.3	399.2	389.3	378.7	368.0	362.5
800	413.1	409.7	406.4	392.6	382.3	371.8	361.4	356.0
900	404.6	402.4	399.8	385.1	374.8	364.2	353.8	348.9
1000	393.9	392.6	390.9	377.3	366.4	356.0	345.9	340.8
1100	381.5	381.0	380.2	368.4	358.0	347.3	337.1	331.8
1200	367.3	367.3	367.0	359.3	348.9	338.0	327.6	322.3

Electrocapillary data on mercury electrode for adsorption of pyridine in 1 M KCl - given are the interfacial tension in mN m^{-1}, potential in V vs CE

T = 26°

Reference: R.G. Barradas, P.G. Hamilton and B.E. Conway.
Coll. Czech. Chem. Commun, 32 (1972) 1790.

				γ /mN m^{-1}			
c/mol l^{-1}	0	0.001	0.003	0.01	0.02	0.03	0.05
-E/mV							
0.00	349.7	349.1	349.5	349.0		347.9	
0.04							
0.05				368.1	367.5		367.0
0.10	379.8	379.3	379.6	379.9	379.2	378.6	378.3
0.20	397.4	396.7	396.9	396.8	395.7	394.7	394.1
0.25							
0.26							
0.28							
0.30	410.1	409.4	409.7	408.8	407.4	406.3	404.9
0.32							
0.34							
0.35							
0.36							
0.38							
0.40	418.9	418.2	417.9	416.6	414.8	413.1	411.4
0.42							
0.44							
0.45	421.7	420.9	420.6	419.0	416.8	414.8	
0.46							413.1
0.47							
0.48							413.4
0.50	423.4	422.7	422.2	420.3	416.9	415.7	413.5
0.52	424.0	423.2	422.7	420.7	418.0	415.8	413.4
0.54	424.2	423.4	422.9	420.9	418.1	415.8	413.3
0.55							
0.56	424.2	423.4	422.9	420.9	418.1	415.7	413.0
0.58	424.2	423.3	422.8	420.7	418.0	415.5	
0.60	423.9	423.0	422.6	420.4	417.5	415.1	412.1
0.65	422.7	421.9	421.5	419.3	416.4	414.0	410.7
0.70	420.7	419.8	419.3	417.4	414.5	412.1	408.6
0.80	414.7	414.1	413.8	412.2	409.7	407.0	403.6
0.90	406.6	406.0	406.0	405.0	403.1	400.1	397.0
1.00	396.5	395.9	395.9	395.7	394.3	392.0	389.5
1.10	384.6	384.1	384.5	384.5	383.9	382.3	381.1
1.20	371.5	370.9	371.1	371.6	371.1	370.7	370.3

Electrocapillary data for pyridine in KCl (continued)

	γ/mN m^{-1}						
c/mol l^{-1}	0.1	0.2	0.3	0.5	1.0	1.75	3.5
-E/mV							
0.00	343.1	329.7				305.1	
0.04					339.1		
0.05							
0.10	376.2	373.5	372.6	368.4	363.7	359.5	354.6
0.20	391.8	388.5	387.0	383.4	378.9	375.5	371.6
0.25				387.4	382.9	379.3	
0.26					383.3		376.3
0.28					384.1		377.3
0.30	401.8	397.8	395.5	390.2	384.6	380.8	378.0
0.32				390.6	384.9	381.0	378.3
0.34		400.0		390.8	385.0	381.0	378.5
0.35	405.3		397.3				
0.36		400.7		390.7	385.0	381.0	378.4
0.38		401.1	397.6	390.6	384.8	380.9	378.2
0.40	407.3	401.3	397.5	390.4	384.4	380.6	378.0
0.42	407.8	401.3	397.3	390.1			
0.44	408.1	401.1	397.0				
0.45				389.3	383.1	379.5	376.9
0.46	408.3	400.8	396.2				
0.47							
0.48	408.2						
0.50	408.0	399.7	395.1	387.8	381.3	377.8	375.2
0.52							
0.54							
0.55	407.3	398.[illegible]	393.2				
0.56							
0.58							
0.60	405.0	395.9	390.7	383.4	376.9	373.7	370.6
0.65							
0.70	401.4	390.3	385.2	377.8	371.6	368.2	365.0
0.80	394.2	384.0	378.6	371.2	365.0	361.4	358.5
0.90	387.1	376.5	371.2	364.1	357.9	354.0	351.3
1.00	379.1	368.7	363.2	356.1	349.7	346.1	343.2
1.10	371.0	360.1	354.7	347.2	341.3	337.4	334.5
1.20	361.8	351.2	345.3	338.2	332.1	327.3	325.4

Electrocapillary data on mercury electrode in 0.5 M K_2SO_4 at 25°C. Adsorption of pyridine - given are the interfacial tension in mN m^{-1}, potential in V vs Hg/Hg_2SO_4 ref. electrode.

Reference: R.G. Barradas, unpublished results.

	γ /mN m^{-1}							
c/mol l^{-1}	0.003	0.01	0.02	0.03	0.05	0.1	0.2	0.3
-E/V								
.350	383.3							
.400	392.2	392.2	391.7	391.1	389.7	388.9	386.8	305.3
.425			395.5	394.9	393.3	392.0	389.5	387.7
.430	399.7	399.2	398.7	398.1	396.5	394.9	391.8	389.8
.475								391.1
.500	406.4	405.6	404.7	403.7	401.8	399.5	395.2	392.4
.525								393.2
.550	412.0	410.7	409.1	408.0	406.0	402.4	397.4	394.0
.575								394.4
.600	416.2	414.4	412.5	411.1	408.6	404.4	398.7	394.8
.625	418.2	415.9	413.7	412.5	409.5	405.0	399.1	394.9
.650	419.7	417.1	414.8	413.5	410.3	405.4	399.3	394.9
.675	420.9	418.2	415.8	414.3	410.9	405.7	399.3	394.9
.700	422.3	419.1	416.6	414.8	411.6	405.8	399.3	394.8
.725	423.1	419.8	416.9	415.2	411.9	405.8	399.0	394.5
.750	423.5	420.2	417.3	415.6	412.1	405.8	398.7	394.0
.775	424.0	420.6	417.8	415.7	412.1	405.7	398.3	393.5
.800	424.4	420.8	417.5	415.7	412.1	405.5	397.8	392.7
.825	424.7	420.8	417.5	415.7	412.0	405.1	397.2	392.0
.850	424.6	420.8	417.4	415.6	411.7	404.6	396.4	391.2
.875	424.5	420.8	417.1	415.2	411.2	403.9	395.6	390.2
.900	424.4	420.4	416.8	414.8	410.7	403.2	394.5	389.1
.950	423.3	419.4	416.0	413.8	409.3	401.4	392.2	387.1
1.000	421.9	418.1	414.6	412.4	407.2	399.4	389.6	384.4
1.050	419.9	416.1	412.8	410.5	405.8	396.7	387.1	381.6
1.100	417.3	413.9	410.7	408.3	403.3	393.8	384.1	378.4
1.150	414.4	411.4	408.3	405.8	400.5	390.8	380.7	376.0
1.200	411.0	408.3	405.5	403.0	397.5	387.6	377.3	371.5
1.250	406.9	404.7	402.4	399.9	394.4	384.1	373.9	367.9
1.300	402.6	401.0	398.7	396.5	390.7	380.0	369.9	364.1
1.350	397.9	396.7	394.7	392.6	386.8	376.1	366.0	360.0
1.400	392.7	391.5	389.9	388.5	382.8	372.0	361.7	355.7
1.450	386.9	385.9	384.8	383.6	373.8	367.8	357.3	351.2
1.500	380.6	379.9	379.3	378.6	374.5	363.4	353.0	346.8

Electrocapillary data for adsorption of piperidine in 1.0 M KCl.
Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE. T = 25 °C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

	γ/mN m^{-1}									
c/mol l^{-1}	0.001	0.003	0.005	0.01	0.03	0.05	0.1	0.25	0.3	1.0
-E/mV										
50	367.8	368.7								
100	379.5	380.9	379.6	380.0	376.2	373.8	372.0	359.9	367.9	353.3
140									377.6	
150							384.4			367.4
160									381.1	
180									384.0	
200	397.0	398.8	397.6	398.0	395.9	394.6	393.5	385.2	386.0	374.2
220									387.1	375.0
240									387.8	375.6
250						402.0	399.4	389.9		
260									387.9	375.5
280							400.9	390.6	387.9	375.4
300	409.9	411.3	410.1	410.5	409.0	407.3	401.3	390.8	387.7	375.2
320							401.4	390.7	387.4	
340						409.1	401.3	390.5		
350				415.6					386.8	374.3
360					414.0	409.5	401.3			
380					414.0	409.2	401.3			
400	418.7	420.3	418.8	419.2	414.7	408.9	400.4	389.2	385.7	373.0
420					414.5	408.6				
440					414.1					
450	421.4	422.9	421.5	421.7		407.9	399.9	387.8		
460				421.9	413.8					
480				422.4						
500	423.2	424.5	423.2	422.7	412.6	406.5	397.6	386.5	382.7	370.1
520	423.6	424.9	423.5	422.8						
540	423.8	425.2	423.7	422.9						
550				422.7	411.1					
560	423.8	425.1	423.7	422.7						
580	423.8	424.9	423.6							
600	423.6	424.7	423.2	421.8	409.5	402.5	394.2	383.3	379.3	366.6
650	422.4	423.4	421.8	420.0						
700	420.5	421.2	419.7	417.8	405.5	398.7	390.1	379.4	375.2	362.5
800	414.3	415.3	413.6	411.9	400.5	394.4	385.5	374.9	370.7	357.9
900	406.1	407.4	405.7	404.7	394.8	388.7	380.4	369.8	365.3	352.7
1000	396.3	397.4	396.1	395.5	387.9	382.5	374.3	364.2	359.6	346.8
1100	384.4	385.8	384.9	384.4	379.4	375.2	367.9	357.6	352.8	340.5
1200	371.4	372.6	371.6	371.6	368.6	365.8	359.3	350.3	344.9	332.8

Electrocapillary data for adsorption of morpholine in 1.0 M KCl.
Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE.
T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

			γ /mN m^{-1}		
c/mol l^{-1}	0.01	0.1	0.3	0.5	1.0
-E/V					
0	346.9				
100	380.5	379.2	377.6	376.6	374.9
200	397.3	397.4	396.2	395.2	393.0
300	410.4	410.2	407.8	406.1	403.0
350		414.3	411.5	409.4	405.2
380					405.7
400	419.2	417.8	414.1	411.1	406.0
420				411.4	406.0
440		419.5	415.0	411.7	406.1
450	421.8				
460			415.1	411.7	405.8
480		420.4	415.2	411.5	
500	423.5	420.6	415.0	411.3	405.1
520	423.7	420.7	414.9		
540	423.9	420.5			
550			414.4	410.0	403.5
560	423.9	420.3			
580	423.8				
600	423.6	419.4	412.3	408.2	401.4
650	422.1	417.8			
700	420.3	415.4	407.5	402.9	395.9
800	414.0	408.7	401.2	396.6	389.6
900	406.6	401.4	393.5	388.9	382.1
1000	396.7	392.0	384.9	380.2	373.4
1100	384.7	383.3	374.7	370.6	363.7
1200	371.7	369.7	363.3	359.6	353.0

Electrocapillary data for adsorption of dioxane in 1.0 M KCl.
Given are the interfacial tension in mN m^{-1}, potential in mV vs SCE.
T = 25°C

Reference: P.G. Hamilton. Ph.D. Thesis, University of Toronto, 1967.

			γ/mN m^{-1}			
c/mol l^{-1}	0.001	0.01	0.1	0.3	1.0	2.0
-E/mV						
0	349.1	348.5	348.3	349.8	344.9	342.6
100	396.6	396.1	395.7	395.5	390.8	387.0
200	396.6	396.1	395.7	395.5	390.8	387.0
300	409.5	409.2	408.1	406.8	399.0	391.3
340					399.7	391.6
350				410.5		
360					400.1	391.6
380					400.2	391.5
400	418.2	417.8	415.9	412.2	400.1	391.4
420					400.0	391.3
440				412.8	399.7	
450	421.1	420.6	417.7			390.9
460				412.8	399.6	
480			418.2	412.6		
500	422.8	422.4	418.3	412.5	398.8	390.2
520	423.1	422.7	418.6			
540	423.4	423.0	418.5			
550				411.5	397.9	
560	423.6	423.0	418.2			
580	423.8	422.8				
600	423.7	422.6	417.8	410.5	387.0	388.2
650	422.5	421.3	416.2			
700	420.3	419.3	414.2	407.1	393.8	385.5
800	414.4	413.3	408.8	402.4	389.9	382.1
900	406.2	405.2	401.7	396.4	385.0	377.6
1000	395.9	395.3	393.2	388.7	378.9	372.0
1100	384.4	383.9	382.2	379.4	371.3	365.2
1200	371.7	370.5	369.3	368.3	361.5	356.3

Electrocapillary data on mercury electrode in 0.25 M NaF at 25°C. Adsorption of acetonitrile. Given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: A. de Battisti and S. Trasatti. J. Electroanal. Chem. 48 (1973) 213.

					γ /mN m^{-1}				
c/mol l^{-1}	0	0.031	0.054	0.096	0.115	0.153	0.192	0.267	0.306
-E/V									
0.000	396.1	394.3	395.5	396.3	396.6	394.9	394.8	394.2	395.9
0.050	402.6	401.6	402.0	402.7	401.8	401.6	400.6	402.1	402.0
0.100	408.7	408.0	408.1	408.4	407.5	407.7	407.0	406.7	408.3
0.150	413.7	412.4	413.4	413.0	412.8	412.4	411.8	411.5	413.0
0.200	417.9	416.9	417.2	417.1	416.5	416.6	415.6	415.5	416.8
0.250	420.9	420.3	420.6	420.3	419.7	419.6	418.9	418.5	419.7
0.300	423.6	422.8	423.0	422.7	421.8	421.9	421.0	420.7	421.9
0.350	425.1	424.5	424.2	424.0	423.4	423.5	422.3	421.9	423.1
0.400	425.8	425.3	425.2	424.7	424.1	424.0	423.0	422.7	423.7
0.450	426.2	425.4	425.4	424.9	424.6	424.1	423.2	422.9	423.8
0.500	425.9	425.2	425.1	424.4	424.0	423.6	423.0	422.4	423.1
0.550	425.1	424.1	424.0	423.5	422.9	422.7	422.2	421.3	422.1
0.600	423.7	422.3	421.4	422.0	421.5	421.0	420.4	419.6	420.5
0.650	421.4	420.4	420.2	420.1	419.5	418.9	418.0	417.6	418.4
0.700	418.7	417.6	417.6	417.2	417.0	416.4	415.7	415.0	415.4
0.750	415.5	415.1	414.5	414.2	413.9	413.4	412.4	412.1	412.6
0.800	412.2	411.2	411.2	410.7	410.6	410.0	409.3	408.6	408.9
0.850	408.2	407.3	407.2	407.0	406.7	405.8	405.7	404.8	405.1
0.900	403.8	403.0	402.7	402.4	402.5	401.5	401.2	400.8	401.0
0.950	398.9	398.0	398.0	397.9	397.7	396.9	396.4	396.1	396.2
1.000	393.8	392.9	392.8	392.5	392.7	391.8	391.3	391.1	391.2
1.050	388.2	387.5	387.1	387.0	387.0	386.1	385.7	385.6	385.7
1.100	382.0	381.4	381.3	381.0	381.2	380.1	380.0	379.4	379.9
1.150	375.8	374.9	374.6	374.8	375.0	374.0	373.5	373.4	373.5
1.200	368.6	368.2	368.1	367.7	368.3	367.5	367.0	366.8	366.6
1.250	361.6	361.0	360.9	360.9	361.1	360.2	359.7	359.7	359.4
1.300	353.9	353.5	353.2	353.1	353.4	352.4	352.3	352.3	352.2
1.350	345.8	345.4	345.2	345.2	345.6	344.8	344.3	344.2	344.1
1.400	337.6	336.8	336.8	336.6	337.2	335.9	336.3	336.0	335.9

Electrocapillary data for acetonitrile in 0.25 M NaF (continued)

	γ /mN m^{-1}								
c/mol l^{-1}	0.444	0.496	0.689	0.764	0.916	0.992	1.072	1.224	1.528
-E/V									
0.000	393.3	395.6	393.9	393.7	391.1	395.3	391.2	393.6	393.8
0.050	399.0	400.7	399.7	398.9	398.9	400.8	398.2	402.8	399.7
0.100	405.9	407.0	405.6	405.6	403.5	406.5	403.7	408.2	408.9
0.150	410.4	411.5	410.0	410.2	408.2	410.7	408.5	408.2	408.9
0.200	414.2	415.0	413.6	413.2	412.3	413.8	411.6	411.6	411.6
0.250	417.4	417.6	416.3	415.9	415.0	415.8	414.4	413.1	413.4
0.300	419.2	419.6	417.9	417.7	416.4	417.3	415.6	414.8	414.2
0.350	420.3	420.4	408.2	418.4	417.4	417.9	416.1	415.0	414.5
0.400	420.7	421.3	418.8	419.6	417.7	418.1	416.2	415.1	414.6
0.450	420.7	421.0	418.6	419.0	417.4	417.8	416.2	415.1	414.1
0.500	420.3	420.7	418.0	418.2	416.8	417.2	415.7	413.9	413.0
0.550	419.0	419.3	416.7	417.0	415.9	415.7	414.4	412.9	411.7
0.600	417.4	417.8	415.2	415.6	414.2	414.2	412.9	411.4	410.0
0.650	415.2	415.5	413.1	413.7	412.2	412.2	410.6	409.4	408.0
0.700	412.9	413.3	410.5	410.9	409.6	409.7	408.2	407.0	405.5
0.750	410.0	409.7	407.7	408.2	406.6	407.0	405.4	404.2	402.9
0.800	406.5	406.7	404.6	405.3	403.8	403.8	402.4	401.2	399.7
0.850	402.8	403.2	400.9	401.8	400.1	400.2	398.8	398.1	396.5
0.900	398.7	398.8	397.2	397.4	396.1	396.3	395.1	394.2	392.9
0.950	394.1	394.2	393.0	393.0	391.7	392.1	390.7	389.7	388.8
1.000	389.1	389.4	387.9	388.3	387.2	387.4	386.4	385.1	384.3
1.050	383.7	383.9	382.9	382.7	382.3	382.4	381.6	380.3	379.5
1.100	377.9	378.3	377.4	377.6	376.8	376.7	376.0	375.0	374.5
1.150	371.6	372.0	371.5	371.7	370.5	371.0	370.1	368.9	368.7
1.200	365.0	365.2	365.1	365.2	364.5	364.5	363.7	362.8	362.8
1.250	357.9	358.3	358.3	358.3	357.6	357.8	357.0	356.1	356.2
1.300	351.7	350.9	350.6	350.8	350.1	350.6	349.8	349.2	349.3
1.350	342.7	343.1	343.1	342.7	342.6	342.8	342.2	341.5	341.8
1.400	335.2	335.0	335.1	335.0	334.4	334.5	333.9	333.6	334.0

Electrocapillary data on mercury electrode in 0.25 M NaF at 25°C. Adsorption of succinonitrile. Given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: A. de Battisti, V. Faggiano and S. Trasatti. J. Electroanal. Chem. 73 (1976) 327.

				γ /mN m^{-1}				
c/mol l^{-1}	0	0.015	0.020	0.024	0.030	0.035	0.040	0.050
-E/V								
0.002	397.8	399.0	399.4	399.4	397.6	399.7	399.7	398.4
0.050	404.5	404.8	404.9	405.1	403.6	405.6	405.6	404.1
0.100	409.5	410.3	409.7	410.2	409.1	410.9	410.9	409.7
0.150	415.6	415.2	414.3	414.7	414.1	415.5	415.5	414.3
0.200	419.1	418.9	421.2	418.4	417.7	419.2	419.2	418.2
0.250	421.9	421.7	421.2	421.3	420.7	422.1	422.1	421.0
0.300	424.1	424.0	423.4	423.3	422.7	424.0	424.0	423.0
0.350	425.5	425.5	424.8	424.8	424.3	425.2	425.2	424.6
0.400	426.6	426.1	425.7	425.6	425.2	425.9	425.9	425.1
0.450	426.7	426.1	425.9	425.6	424.9	425.8	425.8	425.0
0.500	426.2	425.7	425.4	424.9	424.3	425.2	425.2	424.3
0.550	425.5	424.8	424.6	424.1	423.4	424.1	424.1	423.3
0.600	423.5	423.3	422.7	422.6	422.2	422.5	422.5	421.4
0.650	421.5	420.8	420.7	420.4	419.6	426.8	426.8	419.3
0.700	419.0	418.7	418.0	417.9	417.4	417.6	417.6	416.8
0.750	415.9	415.3	414.6	414.8	414.0	414.7	414.7	413.7
0.800	412.2	411.9	411.6	411.3	411.0	411.1	411.1	410.5
0.850	408.3	407.8	409.5	407.6	406.6	407.3	407.3	406.7
0.900	403.9	403.3	403.4	403.2	402.3	403.0	403.0	402.1
0.950	399.0	398.8	399.2	398.3	397.7	398.2	398.2	397.3
1.000	393.8	393.5	393.3	393.3	392.5	393.2	393.2	392.3
1.050	388.1	388.0	387.4	387.8	386.8	387.6	387.6	386.6
1.100	382.0	382.0	381.6	381.6	380.1	381.6	381.6	380.7
1.150	375.6	375.6	375.4	375.6	374.4	375.2	375.2	374.3
1.200	368.7	368.7	368.9	368.3	367.8	368.3	368.3	367.6
1.250	361.5	361,5	357.8	361.1	360.6	361.1	361.1	360.6
1.300	353.9	353.9	353.4	353,5	352.9	353.0	353.0	353.1
1.350	345.9	345.9	345.6	345.6	345.2	345.6	345.6	344.9
1.400	337.4	337.4	337.0	337.3	336.5	337.2	337.2	336.4
1.450	328.5	328.5	328.5	328.5	327.7	328.3	328.3	327.8
1.500	319.3	319.3	319.3	319.3	318.6	319.1	319.1	318.4
1.550	309.5	309.5	309.5	309.5	309.3	309.6	309.6	308.5
1.600	299.4	299.4	299.4	299.4	299.0	299.7	299.7	298.3

Electrocapillary data for succinonitrile in 0.25 M NaF (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.063	0.075	0.090	0.110	0.135	0.150	0.175	0.200
-E/V								
0.002	399.8	398.4	399.1	398.4	398.2	398.7	398.0	387.2
0.050	404.2	404.5	404.9	404.0	403.8	404.0	403.8	403.6
0.100	409.8	409.4	409.8	409.5	409.3	408.8	409.2	408.7
0.150	414.4	414.0	414.6	414.2	413.7	414.0	413.7	413.0
0.200	417.9	418.0	417.6	417.4	417.0	417.1	416.7	415.8
0.250	420.9	420.9	420.8	419.9	419.6	420.7	419.1	418.8
0.300	422.9	422.8	422.6	421.7	421.2	420.9	420.4	419.4
0.350	424.1	424.0	423.4	422.8	422.0	421.6	421.2	420.0
0.400	424.6	424.5	424.0	423.1	422.1	421.7	421.2	419.9
0.450	424.6	424.4	424.0	422.8	421.9	421.3	420.6	419.4
0.500	423.9	423.7	423.4	421.9	421.1	420.5	419.8	418.5
0.550	422.8	422.5	422.2	420.5	419.9	419.3	418.4	417.0
0.600	421.2	421.0	420.4	419.0	418.3	417.3	416.7	415.3
0.650	419.0	418.7	418.3	416.8	416.0	415.3	414.4	413.2
0.700	416.2	415.9	415.5	414.4	413.4	412.7	411.8	410.5
0.750	413.3	413.1	412.6	411.5	410.4	410.0	409.0	407.5
0.800	410.0	409.6	409.2	408.1	407.1	406.6	405.6	404.1
0.850	405.9	405.6	405.3	404.3	403.5	402.8	402.0	399.4
0.900	401.6	401.4	401.3	400.2	399.1	398.7	397.8	396.7
0.950	396.9	396.6	396.5	395.6	394.6	394.2	393.4	392.3
1.000	392.0	391.6	391.6	390.7	389.8	389.2	388.7	387.4
1.050	386.5	386.1	385.8	385.3	384.4	384.0	383.4	382.3
1.100	380.5	380.4	380.1	379.5	378.5	378.3	377.7	376.7
1.150	374.5	374.0	373.7	373.2	372.3	372.0	371.7	370.6
1.200	367.5	367.5	367.2	366.6	365.8	365.7	365.2	364.2
1.250	360.5	360.2	360.0	359.5	358.9	358.7	358.3	357.4
1.300	353.0	352.8	352.7	352.1	351,4	351.3	351,2	350.2
1.350	345.0	344.8	344.6	344.3	343.6	343.6	343.5	342.6
1.400	336.7	336.6	336.4	336.0	335.4	335.3	335.2	334.6
1.450	327.5	327.8	327.5	327.4	326.7	326.5	326.6	326.0
1.500	318.6	318.6	318.4	318.4	317.7	317.5	317.4	317.0
1.550	309.0	308.9	308.9	308.7	308.1	307.9	307.9	307.4
1.600	298.8	298.9	298.8	298.6	298.2	298.0	297.9	297.6

Electrocapillary data for succinonitrile in 0.25 M NaF (continued)

	γ/mN m^{-1}								
c/mol l^{-1}	0.225	0.250	0.300	0.350	0.400	0.450	0.500	0.575	0.650
-E/V									
0.002	396.7	397.4	396.7	397.9	397.2	396.9	397.0	396.5	395.5
0.050	403.5	403.1	402.5	403.0	402.4	401.9	402.2	401.9	401.3
0.100	408.6	408.4	407.8	407.5	406.8	406.0	406.2	405.6	403.7
0.150	412.7	412.3	411.6	411.2	410.5	409.5	408.7	408.1	406.0
0.200	415.6	415.2	414.4	413.2	412.4	411.6	410.5	409.4	407.3
0.250	417.6	417.2	416.1	414.8	413.7	412.7	411.5	410.0	408.0
0.300	418.6	417.9	416.9	415.0	414.3	413.1	411.8	410.4	408.1
0.350	419.1	418.3	417.3	415.6	414.4	412.9	411.6	410.0	408.2
0.400	419.0	418.4	417.0	415.3	414.0	412.8	411.5	409.9	407.9
0.450	418.9	417.5	416.1	414.8	413.4	412.0	410.8	409.3	407.4
0.500	417.8	416.8	415.0	413.5	412.0	410.8	409.7	408.0	406.0
0.550	416.4	415.4	413.9	412.0	410.6	409.3	408.1	406.6	404.6
0.600	414.5	413.6	412.0	410.7	408.7	407.8	406.5	404.9	402.7
0.650	412.5	411.2	410.0	408.3	406.6	405.5	404.6	402.7	400.4
0.700	410.0	408.9	407.2	406.0	404.2	403.0	401.8	400.1	398.0
0.750	407.0	406.2	404.3	403.1	401.3	400.0	398.7	397.8	395.1
0.800	403.6	402.9	401.1	399.6	398.1	396.8	395.6	394.0	391.9
0.850	400.2	399.1	397.7	396.3	394.3	393.2	392.1	390.6	388.4
0.900	396.1	395.1	393.8	392.1	390.6	389.5	388.4	386.9	384.5
0.950	391.7	390.9	389.5	388.0	386.4	384.9	383.7	382.3	380.2
1.000	386.8	386.2	384.5	383.1	381.7	380.3	379.4	377.5	375.5
1.050	381.8	381.2	379.5	378.3	376.6	375.4	374.6	372.9	370.9
1.100	376.3	375.5	374.2	372,8	371,3	370.0	368.9	368.0	365.4
1.150	370.7	369.9	368.0	367.5	365.6	364.4	363.5	362.1	359.9
1.200	364.6	363.4	362.0	360.5	359.5	358.4	357.5	356.1	353.9
1.250	357.4	356.7	355.8	354.2	353.2	352.0	350.3	349,5	347.7
1.300	350.0	349.7	328.9	347.1	346.2	345.3	344.2	343.1	340.9
1.350	342.3	342.0	340.8	340.2	338.9	338.2	337.2	336.0	334.1
1.400	334.4	334.2	333.1	332.9	331.3	330.2	330.0	328.3	326.8
1.450	325.9	325.7	324.6	324.6	323.2	322.4	322.0	320.7	319.1
1.500	316.9	316.6	316.0	315.7	314.8	314.1	313.5	312.8	311.1
1.550	307.5	307.4	307.3	306.7	305.7	305.0	305.4	303.6	302.4
1.600	297.9	297.4	297.5	286.8	296.4	295.8	295.6	294.3	293.1

Electrocapillary data on mercury electrode in 0.25 M NaF, at 25°C. Adsorption of propionitrile. Given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: A.B. Abd-el Nabey and S. Trasatti. J. Chem. Soc. Farad. Trans. 1, 71 (1975) 1230.

	γ /mN m^{-1}							
c/mol l^{-1}	0	0.0283	0.0425	0.0566	0.0623	0.0850	0.0963	0.1076
-E/V								
0.000	400.0	401.3	400.8	397.8	399.8	400.0	400.4	401.5
0.050	404.7	405.9	405.6	404.3	405.9	405.4	405.5	406.4
0.100	410.3	410.6	411.1	419.5	411.3	410.8	411.0	411.2
0.150	414.9	415.4	415.3	414.1	415.4	415.0	415.4	415.6
0.200	419.0	418.9	419.0	418.1	419.0	418.0	418.5	419.6
0.250	422.1	422.1	421.4	421.0	422.0	421.5	421.3	421.7
0.300	424.2	424.0	423.7	422.7	424.0	423.4	423.4	423.4
0.350	425.6	425.2	425.0	424.4	425.2	424.4	424.7	424.5
0.400	426.6	425.8	425.6	425.2	425.5	425.1	425.0	425.0
0.450	426.7	426.0	425.8	425.3	425.7	425.2	424.8	425.3
0.500	425.2	425.3	425.1	424.7	425.0	424.5	424.1	424.9
0.550	425.0	424.2	423.9	423.9	424.2	423.2	423.1	423.0
0.600	423.1	422.7	422.3	422.3	422.4	421.4	421.5	421.3
0.650	421.3	420.4	420.1	420.0	420.1	419.5	416.4	418.9
0.700	418.9	418.0	417.5	417.4	417.6	417.2	417.1	416.8
0.750	415.7	415.2	414.6	414.8	415.0	413.9	414.0	414.0
0.800	411.9	411.6	411.1	411.2	410.9	410.2	410.7	410.2
0.850	408.1	407.8	417.7	407.2	407.2	406.7	406.2	406.5
0.900	403.5	403.3	403.0	403.0	402.7	402.5	402.5	402.5
0.950	398.6	398.5	398.5	398.2	398.4	397.8	397.9	398.4
1.000	393.4	393.2	393.1	393.3	393.0	392.6	392.8	392.8
1.050	388.2	397.7	387.2	387.8	387.4	387.0	387.4	387.6
1.100	382.2	382.0	381.5	382.9	381.5	381.2	381.5	381.5
1.150	375.7	375.2	375.2	375.7	375.5	374.8	375.3	375.3
1.200	369.1	368.5	368.5	369.0	368.4	368.1	368.5	368.4
1.250	361.9	361.5	361.1	361.7	361.4	361.1	361.5	361.7
1.300	354.2	353.9	353.7	354.0	353.5	353.4	354.1	353.8
1.350	345.7	345.9	345.9	346.0	345.5	345.6	346.0	345.6
1.400	337.2	337.4	337.2	337.9	337.1	337.1	337.4	337.4
1.450	328.5	328.7	328.6	328.9	328.7	328.3	328.4	328.5
1.500	319.4	319.6	319.2	319.5	319.6	319.2	319.8	319.4
1.550	309.7	309.4	309.2	309.9	309.7	309.6	309.9	309.8
1.600	299.3	299.8	298.8	299.5	299.4	299.4	299.4	299.2

Electrocapillary data for propionitrile in 0.25 M NaF (continued)

	γ/mN m^{-1}							
c/mol l^{-1}	0.1190	0.1275	0.1473	0.1699	0.1926	0.2124	0.2266	0.2549
-E/V								
0.000	399.7	400.1	401.3	399.2	401.4	400.0	396.7	399.9
0.050	404.9	405.0	406.4	405.4	406.6	405.0	401.8	404.9
0.100	410.1	410.5	411.1	409.9	410.8	410.6	408.2	409.9
0.150	414.5	414.5	415.4	414.6	414.8	414.4	413.2	414.1
0.200	418.0	418.2	418.5	417.8	418.0	417.7	416.8	417.2
0.250	420.7	420.8	420.9	420.6	419.8	419.7	419.1	418.8
0.300	422.7	422.7	422.5	422.4	421.6	421.2	420.4	420.0
0.350	423.5	423.4	423.6	423.2	422.0	421.9	421.3	420.4
0.400	424.0	424.0	424.1	423.5	422.2	421.9	421.1	420.2
0.450	423.9	424.0	423.4	423.0	421.6	421.2	420.4	419.4
0.500	423.1	423.0	423.0	422.2	420.6	420.4	419.4	418.6
0.550	422.0	421.9	421.6	421.0	419.5	418.9	418.6	417.0
0.600	420.3	420.4	419.8	419.2	417.7	417.3	416.7	415.1
0.650	418.4	418.2	417.8		415.8	415.3	414.6	413.5
0.700	415.6	415.6	415.5		413.3	413.0	412.3	410.9
0.750	413.1	412.5	412.7		410.6	410.0	409.7	408.2
0.800	409.7	409.5	409.0		407.4	407.3	406.7	405.2
0.850	405.9	405.9	405.6		404.0	403.5	403.3	401.6
0.900	401.9	401.4	401.5		400.1	399.7	399.4	398.0
0.950	397.5	397.1	397.1		396.0	395.6	395.3	394.3
1.000	392.5	392.0	392.2		391.1	391.2	390.6	389.4
1.050	387.0	386.7	387.0		386.1	386.1	385.7	384.9
1.100	386.9	381.0	380.7		380.3	380.6	379.9	379.3
1.150	374.7	375.0	375.0		374.5	374.5	374.0	373.4
1.200	368.0	368.0	368.0		367.8	367.8	367.6	368.1
1.250	361.0	361.0	361.0		360.8	361.0	360.6	360.4
1.300	353.5	353.5	353.8		353.5	353.6	353.3	352.3
1.350	345.6	346.0	345.5		345.7	345.5	345.2	344.9
1.400	337.1	337.5	337.5		337.2	337.6	336.8	336.5
1.450	328.5	328.9	328.4		328.8	328.6	328.2	327.7
1.500	319.5	319.7	319.2		319.0	319.3	319.1	318.6
1.550	309.4	309.9	309.7		309.3	309.7	309.6	309.1
1.600	299.6	299.3	299.1		299.6	299.4	299.2	298.5

Electrocapillary data for propionitrile in 0.25 M NaF (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.2974	0.3172	0.3399	0.3823	0.4192	0.4532	0.5098	0.5664
-E/V								
0.000	400.5	400.2	399.3	399.6	399.8	399.7	400.5	399.6
0.050	405.2	404.5	403.8	404.9	405.6	404.3	404.5	402.5
0.100	410.0	408.6	408.1	410.4	409.5	408.9	408.7	406.7
0.150	413.6	412.5	411.9	413.4	412.5	412.4	411.6	409.9
0.200	416.6	414.8	414.8	415.6	414.5	414.0	412.8	412.4
0.250	418.3	416.6	416.6	416.7	415.2	415.3	413.5	413.4
0.300	419.3	417.7	417.1	417.5	415.5	415.5	414.1	413.1
0.350	419.3	418.1	417.4	417.6	416.0	415.3	413.6	413.1
0.400	419.4	417.8	417.4	416.7	415.3	414.7	413.2	412.5
0.450	418.7	417.1	416.4	415.8	414.4	413.7	412.1	411.2
0.500	417.3	415.8	415.2	414.5	412.7	412.5	410.8	410.0
0.550	415.9	414.2	413.9	413.4	411.1	410.9	409.4	418.7
0.600	414.1	412.2	412.2	411.2	409.4	408.7	405.9	406.7
0.650	412.2	410.4	409.7	408.9	407.6	406.6	404.5	404.6
0.700	409.6	408.6	407.2	406.7	405.1	404.2	402.2	402.0
0.750	407.1	406.3	404.7	404.2	402.2	402.0	397.7	399.3
0.800	404.2	403.1	401.9	400.9	399.1	398.7	396.5	396.4
0.850	400.9	399.2	398.8	397.4	395.6	395.2	393.3	393.2
0.900	397.4	395.7	395.0	394.0	392.6	391.7	390.0	389.6
0.950	393.3	392.0	391.0	392.7	388.6	388.1	386.4	385.8
1.000	389.3	388.1	386.8	386.6	384.7	384.2	382.4	382.2
1.050	384.5	383.4	382.5	382.3	380.4	379.9	378.4	378.1
1.100	379.3	378.4	377.6	377.3	375.8	375.6	373.6	373.8
1.150	373.6	372.2	372.0	372.2	370.7	370.8	369.2	369.1
1.200	367.3	365.8	366.1	366.2	364.5	365.1	363.7	363.8
1.250	360.7	359.5	359.3	359.7	358.1	359.0	358.3	357.7
1.300	353.0	352.3	352.3	352.7	350.8	352.3	351.8	351.4
1.350	345.5	344.6	344.5	344.6	343.2	345.1	344.4	344.3
1.400	337.4	336.5	336.5	336.5	335.0	337.1	336.4	336.1
1.450	328.3	327.7	327.8	327.6	326.6	328.3	328.2	327.9
1.500	319.2	318.4	318.6	318.6	317.3	319.2	318.9	318.8
1.550	310.1	309.2	309.1	309.2	307.4	309.7	309.3	309.2
1.600	299.4	299.0	298.6	298.9	297.6	299.9	299.2	299.4

Electrocapillary data for benzene in 0.1 N $HClO_4$ solution.
Reference electrode: SCE (NaCl solution) $T = 25^{\circ}C$

Reference: K.G. Baikerikar and Robert S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

γ : Interfacial tension in mN m^{-1}

Reduced concentration = c/c_o, where c is adsorbate concentration and c_o is saturation concentration in the base electrolyte solution.

			γ/mN m^{-1}			
c/c_o	0.10	0.20	0.40	0.60	0.80	1.00
E/V						
-1.10	386.0	386.1	385.6	385.5	385.0	385.7
-1.05	391.5	391.7	391.5	391.1	390.9	390.3
-1.00	396.6	396.6	396.5	395.7	395.5	394.0
-0.95	401.1	400.9	400.7	399.9	399.3	396.3
-0.90	405.6	405.4	404.7	403.6	402.0	397.5
-0.85	409.5	409.2	408.1	406.7	404.2	398.6
-0.80	413.0	412.4	411.0	409.1	405.4	399.4
-0.75	416.1	415.3	413.6	410.4	406.4	400.2
-0.70	418.7	417.6	414.9	411.5	407.0	400.8
-0.65	420.8	419.4	416.1	412.1	407.7	401.4
-0.60	422.2	420.6	416.9	412.6	408.1	401.4
-0.55	423.3	421.3	417.2	412.8	408.6	401.8
-0.50	423.5	421.4	417.2	412.8	408.9	402.0
-0.45	423.4	421.1	416.9	412.9	409.0	402.0
-0.40	422.4	420.0	416.1	412.5	408.9	402.0
-0.35	420.7	418.5	414.9	411.6	408.4	402.0
-0.30	418.6	416.1	413.0	410.1	407.2	401.1
-0.25	415.6	413.5	410.5	408.0	405.5	400.4
-0.20	412.3	409.9	407.5	405.1	403.0	398.8
-0.15	408.2	406.4	403.7	401.8	399.8	396.4
-0.10	403.6	401.7	399.7	397.7	396.0	392.8
-0.05	398.8	397.0	395.0	393.1	391.7	389.0
0.0	393.0	391.3	389.7	388.1	386.7	384.2
0.05	386.9	385.5	384.0	382.8	381.3	379.3
0.10	380.1	378.8	377.6	376.3	375.3	373.2
0.15	372.8	371.6	370.5	369.7	368.7	366.9
0.20	364.1	363.6	362.9	362.3	361.4	359.8

Electrocapillary data for Toluene in 0.1 N $HClO_4$ solution.

Reference electrode: SCE (NaCl solution) T = 25°C

Reference: K.G. Baikerikar and Robert S. Hansen. J. Colloid Interface Sci. 61 (1977) 239.

: Interfacial tension in mN m^{-1}

Reduced concentration = c/c_o, where c is adsorbate concentration and c_o its saturation concentration in the base electrolyte solution.

	γ/mN m^{-1}						
c/c_o	0.10	0.20	0.30	0.40	0.60	0.80	1.00
E/V							
-1.10	385.2	385.0	385.0	385.1	384.9	384.7	385.0
-1.05	390.8	390.7	390.6	390.7	390.4	390.1	390.1
-1.00	396.1	396.0	395.8	395.9	395.3	394.9	394.7
-0.95	401.1	400.8	400.4	400.4	399.8	399.2	398.7
-0.90	405.5	405.2	404.7	404.5	404.0	402.0	401.1
-0.85	409.6	409.2	408.4	408.0	406.1	403.6	402.3
-0.80	413.2	412.6	411.8	411.3	407.9	404.5	403.3
-0.75	416.3	415.2	414.9	412.3	409.1	405.6	404.4
-0.70	418.9	417.4	416.1	413.5	409.6	406.6	404.8
-0.65	420.8	419.0	417.4	414.3	410.4	407.3	405.5
-0.60	421.7	419.9	417.6	415.2	410.8	407.6	405.8
-0.55	422.9	419.9	417.8	415.4	411.0	408.1	406.0
-0.50	423.1	419.8	417.8	415.4	410.9	408.4	406.2
-0.45	422.9	419.7	416.8	414.9	410.8	408.2	406.2
-0.40	421.8	418.9	416.2	413.5	410.5	407.8	406.1
-0.35	420.3	417.8	415.1	413.0	409.7	407.0	405.3
-0.30	418.4	415.9	413.5	412.1	408.6	405.8	403.7
-0.25	415.8	413.7	411.3	410.1	406.8	404.7	402.2
-0.20	412.5	410.6	408.5	407.4	404.6	402.7	399.9
-0.15	408.8	407.1	405.4	404.3	401.9	400.0	397.6
-0.10	404.5	402.8	401.5	400.6	398.2	396.8	393.8
-0.05	399.6	398.1	397.0	396.2	394.3	392.9	390.3
0.0	394.1	392.9	391.9	391.2	389.5	388.3	386.1
0.05	387.9	387.2	386.3	385.7	384.2	383.2	381.4
0.10	381.2	380.6	379.9	379.4	378.3	377.4	375.8
0.15	373.8	373.5	372.8	372.5	371.6	370.8	369.3
0.20	365.8	365.6	365.1	364.8	364.1	363.7	362.2

Electrocapillary data on mercury electrode in 1 M $NaClO_4$ at 22°C. Adsorption of acetylacetone - given are the interfacial tension in mN m^{-1}, potential in mV vs CE.

Reference: J. Dojlido, E. Galus and L.J. Jeftic. J. Electroanal. Chem. 62 (1975) 433.

	γ/mN m^{-1}							
c/mol l^{-1}	0	0.05	0.1	0.2	0.4	0.6	0.8	1.0
-E/mV								
15	386.5	386.5	385.5	384.3	382.5	380.0	378.5	375.7
100	395.0	395.2	394.5	394.0	391.9	387.9	386.2	382.2
200	405.3	405.4	402.8	401.2	398.4	396.0	392.2	388.4
300	413.2	413.2	409.5	407.4	402.4	398.5	395.3	391.2
400	418.4	418.1	412.6	408.5	402.9	399.5	395.7	392.3
500	422.4	419.8	413.5	408.6	402.4	398.2	395.0	391.7
550	423.0		412.8					
600	422.8	418.7	412.0	405.7	400.7	395.8	392.7	389.2
700	419.8	415.3	408.9	401.4	397.0	391.8	388.2	385.3
800	414.3	410.0	403.9	396.5	391.8	386.5	383.8	380.8
900	406.2	402.3	397.2	390.0	385.1	380.2	377.1	374.8
1000	396.3	393.2	388.8	382.0	377.7	373.0	369.8	367.6
1100	385.0	381.8	378.9	373.3	369.0	363.9	361.4	359.5
1200	371.5	369.2	366.8	362.9	359.0	354.4	352.0	350.0
1300	356.0	354.7	352.4	350.3	347.8	343.0	341.2	339.4
1400	337.8	337.0	336.7	336.2	335.3	331.1	329.0	328.4
1490	320.2	320.1	319.8	319.7	319.6	317.2	316.4	316.3

Differential capacity on mercury in 1 M KCl (continued)

Adsorption of C_8TPABr (continued)

	$C/\mu F\,cm^{-2}$							
	C_8TPABr							
$c/mol\ l^{-1}$	0	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
E/V								
1.645						95.8		
1.65						56.6	10.8	
1.66						25.8		
1.666								
1.68								
1.69								
1.696							96.1	
1.70		21.8	21.9	21.9	22.1	22.1	73.0	10.4
1.71							37.6	
1.73								
1.75								10.3
1.77								
1.80	23.5	24.0	24.1	24.3	24.4		24.2	33.0
1.812								48.3
1.82								36.0
1.832								
1.84								27.8
1.86								

Electrocapillary data on mercury in tetrabutylammonium iodide at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}						
c/mol l^{-1}	0.001	0.00112	0.002	0.0025	0.005	0.01	0.025
-E/V							
0.300	378.2	378.3	369.1	364.9	355.6		330.8
0.325			374.3		362.5	352.2	
0.350	387.6	386.8	378.9	376.3	368.3	358.7	
0.375			382.7			364.2	
0.400	394.0	393.0		384.1			357.2
0.450	397.9	397.0		388.9		375.0	365.1
0.475			391.9		384.0	377.2	
0.500	399.7	398.8	392.8	391.3	385.1	378.7	370.1
0.525			393.3		385.8	379.6	
0.550	399.8	398.9	393.4	391.8	386.0	380.1	372.7
0.575			393.0		385.8	380.2	
0.600	398.5	397.4	392.4	390.8	385.2	379.9	373.4
0.625			391.4		384.4	379.3	
0.650	396.2	394.8	390.2	388.8	383.4	378.4	372.4
0.675			388.7			377.3	
0.700	393.0	391.4	387.1	385.9	380.7	376.0	370.3
0.750	389.4	387.7	383.5	382.5	377.4	372.9	367.2
0.800	385.4	383.7	379.5	378.7	373.7	369.4	363.6
0.850	381.3	379.6	375.4	374.7	369.8	365.5	359.7
0.900	377.1		371.4	370.7	365.8	361.4	355.5
0.950	372.9		367.4	366.5	361.6	357.2	351.4
1.000	368.7		363.4	362.3	357.3	352.8	347.2
1.050	364.4		359.2	358.0	353.0		343.1
1.100	360.0		354.6	353.6	348.5	343.9	338.8
1.150	355.3		349.0	349.0		339.3	334.3
1.200	350.1			344.0	339.0	334.6	329.2

Electrocapillary data on mercury in tetrapropylammonium iodide at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ/mN m^{-1}						
c/mol l^{-1}	0.001	0.00105	0.002	0.00209	0.0025	0.00314	0.005
-E/V							
0.300	396.6	395.7					371.1
0.350	405.9	405.4	397.7	397.6	395.0		385.7
0.375							
0.400	412.7	412.5	405.5	405.4	403.1	400.6	395.6
0.450	417.4	417.3	411.0	411.0	408.9	406.6	402.2
0.500	420.2	420.1	414.5	414.5	412.6	410.5	406.2
0.550	421.2	421.3	416.3	416.3	414.5	412.7	408.5
0.575							
0.600	420.9	421.0	416.6	416.6	414.9	413.3	409.3
0.625							
0.650	419.3	419.5	415.5	415.6	414.0	412.6	408.8
0.675							
0.700	416.8	417.1	413.4	413.5	412.0	410.7	407.3
0.750	413.5	413.9	410.4	410.6	409.1	407.8	404.6
0.800	409.6	410.1	406.7	406.9	405.4	404.3	401.0
0.850	405.2	405.8	402.5	402.7	401.2	400.2	396.6
0.900	400.5	401.1	397.9	398.0	396.5	395.6	391.9
0.950	395.5	396.2	392.9	393.1	391.5	390.7	387.3
1.000	390.3	391.0	387.7	387.9	386.3	385.6	
1.050	385.0	385.6	382.3	382.5		380.3	
1.100	379.5	380.1	376.7	377.0		374.7	
1.150		374.3	370.9	371.3		368.8	
1.200		368.3	364.6				

Electrocapillary data for tetrapropylammonium iodide (continued)

	γ /mN m^{-1}					
c/mol l^{-1}	0.008	0.01	0.02	0.025	0.05	0.1
-E/V						
0.300						
0.350						335.6
0.375	385.4	382.3				
0.400	389.7	386.9	377.6	374.6	365.3	354.9
0.450		394.1	385.8	383.0	374.8	366.5
0.500	401.4	399.0	391.5	389.0	381.6	373.8
0.550	404.2	401.8	394.9	392.7	385.9	378.6
0.575	405.0	402.6				
0.600	405.3	403.0	396.4	394.5	388.1	381.4
0.625	405.3	403.0				
0.650	405.0	402.7	396.4	394.6	388.5	382.5
0.675	404.4	402.0				
0.700	403.5	401.2	395.1	393.3	387.5	381.6
0.750	401.0	398.7	392.6	390.8	385.4	378.9
0.800		395.5	389.3	387.6	382.3	375.4
0.850	393.9	391.7	385.2	383.7	378.6	
0.900		387.4	380.5	379.4	374.3	
0.950	384.9	382.8		374.8	369.8	
1.000	380.0	378.0		370.1	365.1	
1.050	374.9	373.0		365.1	360.2	
1.100	369.6	367.8		359.8	355.1	
1.150	364.2	362.5			349.9	
1.200		357.0			344.4	

Electrocapillary data on mercury in tetraethylammonium iodide at various concentrations - given are the interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}					
c/mol l^{-1}	0.001	0.003	0.005	0.008	0.00873	0.01
-E/V						
0.350			387.5			378.3
0.375	408.6	399.6	392.6			384.2
0.400	411.5	403.5	397.1	392.6	391.5	389.4
0.425	413.9		401.0	396.8		393.8
0.450				400.4	399.3	
0.475		412.3		403.4		400.8
0.500	418.5	414.3	409.3	405.9	404.8	403.4
0.525	419.2	415.8	411.0	407.8		405.4
0.550	419.6		412.3	409.3	408.3	407.0
0.575	419.6	417.6	413.1	410.3		
0.600	419.3	417.9	413.5	411.0	410.0	408.9
0.625	418.7	417.8	413.5	411.2		409.2
0.650	417.8	417.3	413.2	411.1	410.2	409.1
0.675	416.6	416.5	412.5	410.7		408.7
0.700	415.2	415.4	411.5	410.0	409.0	408.0
0.750	411.7		408.7	407.7	406.6	405.8
0.800	407.5	407.9	405.0	404.4	403.2	402.7
0.850	402.5			400.3	399.0	398.7
0.900			395.7		394.2	
0.950			390.4	390.3	388.8	388.9
1.000			384.9	384.6	383.0	383.3
1.050			379.2	378.6	377.1	377.3
1.100		369.5		372.7	371.0	371.0
1.150	365.8	363.1			365.1	
1.200		357.4				357.5
1.250			356.5			
1.300	347.4					
1.350	342.0		344.5			335.8

Electrocapillary data for tetraethylammonium iodide (continued)

	γ /mN m^{-1}				
c/mol l^{-1}	0.01455	0.0582	0.0873	0.1455	0.2911
-E/V					
0.350					
0.375					
0.400	384.6				
0.425					
0.450	392.9	375.7	370.1	364.0	
0.475					
0.500	399.1	383.4	378.3	372.6	363.8
0.525					
0.550	403.2	388.5	384.2	379.1	371.0
0.575					
0.600	405.5	392.2	388.1	383.4	375.9
0.625					
0.650	406.2	394.2	390.2	385.9	378.9
0.675					
0.700	405.5	394.1	390.6	386.7	380.0
0.750	403.5	393.3	389.7	385.9	379.6
0.800	400.4		387.5	383.9	377.9
0.850	396.4		384.3	380.8	374.9
0.900	391.7		380.2	376.7	370.9
0.950	386.3		375.3	371.9	366.0
1.000	380.5		369.8	366.5	360.5
1.050	374.4		363.9	360.6	354.6
1.100	368.0		357.8	354.5	348.4
1.150			351.7	348.1	342.1
1.200				341.5	
1.250					
1.300					
1.350					

Electrocapillary data on mercury in tetramethylammonium iodide at various concentrations - given are the interfacial tension in mN m^{-1}, potential in V vs SCE.
T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ/mN m^{-1}							
c/mol l^{-1}	0.001	0.0028	0.005	0.01	0.02	0.05	0.1	0.25
-E/V								
0.225		362.0	343.4					
0.250		374.1	358.0					
0.275		384.3	370.5					
0.300	404.2	392.9		369.4		331.0		
0.325	408.4	400.1	389.9	379.5		345.2		
0.350	412.1	406.0	397.3	388.1	377.0			
0.375	415.3	410.9	403.4	395.3	384.5			
0.400	417.9		408.4	401.2		376.1	365.8	349.5
0.425	420.1	417.8		406.1	396.5	383.2	372.8	358.1
0.450	421.8		415.4	410.0	401.3		378.8	365.4
0.475	423.1		417.8	413.2	405.3		380.0	371.6
0.500	424.2	423.0	419.5	415.6	408.7	398.1	388.5	376.9
0.525	424.9	423.8	420.7	417.4	411.3	401.3	392.4	381.3
0.550	425.2	424.1	421.4	418.6		403.8	395.6	385.0
0.575	425.3	424.1	421.7	419.4	414.9	405.8	398.2	388.0
0.600	425.2	423.8	421.7	419.8	416.0	407.3	400.3	390.5
0.625	424.8	423.3	421.3	419.8	416.5	408.4	402.0	392.6
0.650	424.2	422.6	420.7	419.5	416.6	409.1	403.2	394.2
0.675	423.4	421.6	419.9	418.9	416.4	409.5	404.0	395.4
0.700	422.4	420.6	418.9	418.1	415.8	409.5	404.4	396.2
0.725						409.3	404.4	396.7
0.750	420.0	418.0	416.4	415.8	413.7	408.7	404.1	396.9
0.775						407.9	403.5	396.7
0.800	416.9	414.8	413.2	412.8	410.5	406.9	402.6	396.3
0.850	413.4	411.2	409.4	409.0	406.5	403.9	400.0	
0.900	409.5	407.0	405.1	404.5			396.3	
0.950	405.2	402.3	400.2	399.5	396.6	394.6	391.7	387.4
1.000	400.5	397.1	394.7	393.8	390.9	388.3	386.3	382.4
1.050	395.4	391.5	388.9	387.7	384.8	381.0	380.3	376.7
1.100	389.9	385.5	382.7	381.3	378.5	372.9	373.7	370.9
1.150	384.1	379.3	376.5	374.9	371.9	364.7	366.9	
1.200	377.9	373.4	370.7	369.0	365.0	356.9	360.0	

Electrocapillary data on mercury in tetrabutylammonium bromide at various concentrations - given are the interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

					γ /mN m^{-1}			
c/mol l^{-1}	0.3	0.2	0.1	0.07	0.05	0.03	0.02	0.01
-E/V								
0.000		315.71			344.4			
0.050		331.2			360.8	342.9		365.4
0.100	332.1		343.8	350.3	372.7			380.1
0.150	347.6	355.5			380.8	375.5	381.4	
0.200		364.2	371.0	375.0	385.9	384.1		
0.250	366.3	370.7	377.8	381.1	388.6	389.1	392.9	399.6
0.300	371.0	375.1	381.5	384.2	389.6	391.4	394.8	400.7
0.350	373.3	377.6	382.7	385.2	389.1	391.7	394.8	400.2
0.400	373.9	378.5	382.3	384.6	387.6	390.5	393.6	398.6
0.450	373.0	377.9	380.6	382.9	385.4	388.4	391.4	
0.500	371.1	376.1	378.1	380.4	382.7	385.6	388.7	393.2
0.550	368.4	373.4	375.1	377.4	379.7	382.5	385.5	
0.600	365.3	369.9	371.8	374.2	376.4	379.2	382.2	386.4
0.650	361.8	366.0	368.4	370.7	373.0	375.8	378.7	382.9
0.700	358.2	361.9	364.9	367.2	369.4	372.3	375.1	379.3
0.750	354.4	357.7	361.3	363.5	365.8	368.7	371.4	375.7
0.800	350.6	353.6	357.7	359.8	362.1	365.1	367.7	371.9
0.850	346.8		353.9	356.0	358.4	361.3	363.9	368.1
0.900	342.9	345.9	350.0	352.0	354.4	357.3	359.9	364.1
0.950	339.0	342.3	345.9	347.8	350.4	353.1	355.8	359.9
1.000	334.9	338.7	341.6	343.4	346.2	348.7	351.5	355.5
1.050	330.7	335.0	337.1	338.9	341.9	344.1	347.0	351.0
1.100	326.3	330.8	332.5	334.4	337.5	339.5	342.3	346.3
1.150	321.6	325.7	327.9	329.9		334.9	337.7	341.6
1.200	316.5	319.3	323.5	325.6		330.6	333.1	337.2

Electrocapillary data for tetrabutylammonium bromide (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.007	0.005	0.003	0.002	0.001	0.0007	0.0005	0.0003
-E/V								
0.000			366.0	373.9		387.6	390.0	393.1
0.050						397.2	397.9	399.4
0.100	384.2	387.3	392.9	397.5	401.8	404.8	404.6	405.2
0.150		396.0	400.6	404.5	409.1	410.4	410.1	410.2
0.200	398.7	401.4	405.5	408.8	413.7	414.3	414.1	414.4
0.250	401.8	404.3	408.0	411.1	416.1	416.5	416.9	417.5
0.300	403.0	405.3	408.8	411.8	416.7	417.6	418.4	419.5
0.350	402.5	404.7	408.1	411.2	416.0	417.4	418.8	420.4
0.400	401.0	403.1	406.5	409.5	414.3	416.2	418.1	420.2
0.450	398.6	400.7	404.1	407.1	411.8	414.1	416.5	419.1
0.500	395.7	397.7	401.2	404.2	408.9	411.4	414.1	417.2
0.550	392.4	394.4	397.9	400.9	405.6	408.1	411.0	414.5
0,600	388.9	390.9	394.4	397.4	402.1	404.5	407.5	411.2
0.650	385.2	387.2	390.8	393.6	398.4	400.7	403.7	407.6
0.700	381.5	383.6	387.0	389.8	394.6	396.7	399.7	403.6
0.750	377.7	379.8	383.3	386.0	390.8	392.6	395.6	399.5
0.800	373.9	376.0	379.4	382.1	386.9	388.6	391.5	395.4
0.850	370.0	372.0	375.5	378.1	382.9	384.5	387.5	391.4
0.900	366.0	368.0	371.5	374.1	378.9	380.6	383.6	387.4
0.950	361.8	363.8	367.3	370.0	374.6	376.6	379.7	383.5
1.000	357.5	359.5	363.0	365.7	370.3	372.5	375.8	379.6
1.050	353.0	355.0	358.6	361.3	365.8	368.3	371.7	375.6
1.100	348.3	350.4	353.9	356.7	361.0	363.8	367.3	371.3
1.150		345.7	349.2	351.9	356.2	358.8	362.2	366.3
1.200		340.9	344.3	346.8	351.2	353.1	356.2	360.2

Electrocapillary data on mercury in tetrapropylammonium bromide at various concentrations - given are the interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}									
c/mol l^{-1}	1.0	0.7	0.5	0.3	0.2	0.1	0.07	0.05	0.03	0.02
-E/V										
0.000							301.9	307.1	321.5	333.1
0.050						324.6			349.3	357.9
0.100				330.3	338.6	350.6	357.0	362.4		
0.150	338.6	340.9	347.2			369.6	374.2	379.3	385.2	390.2
0.200				368.4	373.4	382.9	386.3	390.9	395.8	399.8
0.250	367.0	369.6	373.8	379.4	383.5	391.7	394.4	398.4	402.7	406.1
0.300	375.0	377.6	381.3	386.4	390.0	397.1	399.3	402.7	406.9	409.9
0.350	379.8	382.6	385.8	390.5	393.7	399.9	401.8	404.8	408.9	411.7
0.400	382.3	385.0	388.1	392.2	395.2	400.7	402.5	405.1	409.3	412.1
0.450	382.8	385.6	388.5	392.1	395.1	400.2	401.9	404.3	408.5	411.3
0.500	382.0	384.7	387.5	390.8	393.8	398.6	400.3	402.6	406.8	409.7
0.550	380.0	392.7	385.5	388.6	391.6	396.3	398.1	400.3	404.5	407.4
0.600	377.3	380.0	382.7	385.8	388.7	393.5	395.3	397.6	401.8	404.7
0.650	374.0	376.7	379.4	382.5	385.4	390.4	392.2	394.6	398.7	401.6
0.700	370.4	373.0	375.7	378.9	381.8	387.0	388.8	391.2	395.2	398.1
0.750	366.4	369.0	371.7	375.1	377.9	383.3	385.0	387.6	391.4	394.3
0.800	362.2	364.8	367.5	371.0	373.8	379.3	381.0	383.6	387.4	390.3
0.850	357.8	360.3	363.1	366.7	369.4	375.0	376.6	379.3	383.0	385.9
0.900	353.2	355.7	358.6	362.2	364.9	370.4	372.0	374.5	378.3	381.1
0.950	348.4	350.9	353.9	357.4	360.2	365.5	367.0	369.5	373.3	376.1
1.000	343.5	345.9	349.1	352.5	355.3	360.3	361.8	364.1	368.1	370.9
1.050	338.4	340.8	344.0	347.3	350.2	354.9	356.5	358.6	362.7	365.5
1.100	333.1	335.4	338.9	342.1	345.0	349.6	351.1	353.2	357.3	360.2
1.150	327.6	330.0	333.6	336.8	339.8	344.5	346.1	348.2	352.2	355.0
1.200	322.1	324.4	328.2	331.8	334.7	340.0	341.7	344.1	347.6	350.3

Electrocapillary data for tetrapropylammonium bromide (continued)

c/mol l^{-1}	γ /mN m^{-1} 0.01	0.007	0.005	0.003	0.002	0.001	0.0007	0.0005	0.0003
-E/V									
0.000	348.7	357.6	364.9	375.7	382.5	388.2	390.6	393.4	395.6
0.050				388.2	392.7	396.8	398.1	400.0	401.4
0.100	385.8	389.8	393.3	398.3	401.3	404.3	404.9	406.1	406.8
0.150	397.5	400.3	402.8	406.3	408.4	410.5	410.8	411.5	411.6
0.200	405.9	407.9	409.9	412.4	414.0	415.6	415.7	416.0	415.8
0.250	411.4	413.1	414.8	416.8	418.2	419.5	419.6	419.7	419.3
0.300	414.8	416.4	417.9	419.8	421.1	422.3	422.4	422.4	421.9
0.350		418.0	419.5	421.4	422.8	423.9	424.2	424.2	423.8
0.400	416.7	418.4	420.0	421.9	423.4	424.6	425.0	425.1	424.8
0.450	415.9	417.7	419.4	421.4	423.1	424.3	424.9	425.1	425.0
0.500	414.4	416.3	417.9	420.1	421.9	423.2	423.8	424.2	424.5
0.550	412.1	414.2	415.8	418.1	419.9	421.3	422.1	422.6	423.2
0.600	409.4	411.5	413.1	415.4	417.2	418.8	419.6	420.4	421.3
0.650	406.3	408.4	409.9	412.3	414.1	415.7	416.6	417.6	418.8
0.700	402.8	404.9	406.4	408.7	410.4	412.2	413.2	414.3	415.8
0.750	399.0	401.0	402.5	404.7	406.4	408.3	409.3	410.6	412.4
0.800	394.8	396.8	398.3	400.5	402.1	404.2	405.2	406.6	408.5
0.850	390.4	392.4	393.8	395.9	397.6	399.8	400.9	402.3	404.4
0.900	385.7	387.6	389.1	391.2	392.9	395.2	396.4	397.9	400.1
0.950	380.6	382.5	384.1	386.3	388.0	390.5	391.8	393.3	395.6
1.000	375.4	377.3	378.9	381.2	383.0	385.6	387.0	388.5	390.9
1.050	370.0	371.8	373.6	375.8	377.8	380.5	382.1	383.6	386.0
1.100	364.5	366.3	368.0	370.3	372.3	375.2	376.8	378.4	380.9
1.150	359.1	360.7	362.4	364.6	366.5	369.5	371.1	372.9	375.7
1.200	354.0	355.3	356.6	358.6	360.3	363.4	364.9	367.0	370.1

Electrocapillary data on Mercury in tetraethylammonium bromide at various concentrations - given are the interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}								
c/mol l^{-1}	1.0	0.7	0.3	0.2	0.1	0.05	0.03	0.02	0.01
-E/V									
0.000									
0.050									
0.100									392.1
0.150		331.0		363.3	372.3	383.3	388.8	393.8	401.1
0.200	351.7	350.5	369.3	375.7	383.7	393.1	397.5	401.7	408.3
0.225						397.2			411.2
0.250	364.8	365.4	379.9	385.5	392.8	400.9	404.5	408.0	413.8
0.275						404.0			416.0
0.300	375.0	376.4	388.1	393.1	399.7	406.7	409.9	412.9	417.9
0.325					402.4	408.9			419.4
0.350	382.5	384.3	394.2	398.6	404.7	410.8	413.7	416.3	420.6
0.375					406.6	412.3			421.5
0.400	387.6	389.5	398.3	402.3	407.9	413.3	416.1	418.4	422.0
0.425					408.9	414.0			422.3
0.450	390.7	392.6	400.6	404.2	409.5	414.5	417.2	419.3	422.3
0.475			401.2	404.6	409.8				422.1
0.500	391.9	393.9	401.3	404.7	409.8	414.4	417.2	419.1	421.6
0.525			401.1	404.5	409.4	413.9			
0.550	391.5	393.7	400.6	403.9	408.8	413.2	416.1	417.9	420.0
0.575			399.7	403.1					
0.600	389.7	392.2	398.6	402.0	406.8	411.1	414.0	415.8	417.7
0.650	386.9	389.6	395.6	399.1	403.8	408.1	411.1	413.0	414.6
0.700	383.1	386.2	391.8	395.5	400.2	404.5	407.6	409.5	411.0
0.750	378.6	381.9	387.3	391.2	395.9	400.4	403.5	405.4	407.0
0.800	373.5	376.9	382.4	386.4	391.2	395.7	398.9	400.9	402.5
0.850	368.0	371.3	377.1	381.3	387.1	390.7	393.9	396.0	397.7
0.900	362.2	365.1	371.6	375.8	380.6	385.4	388.6	390.7	392.6
0.950	356.2	358.5	365.9	370.0	375.0	379.8	383.0	385.2	387.2
1.000	349.9	351.6	360.1	364.1	369.1	373.9	377.1	379.4	381.6
1.050	343.6	344.6	354.2	358.0	363.0	367.9	371.1	373.3	375.7
1.100	337.1	337.8	348.0	351.7	356.7	361.6	364.7	366.9	369.5
1.150	330.3	331.3	341.6	345.2	350.2	355.0	358.0	360.1	363.0
1.200	323.4	325.6	334.7	338.4	343.4	348.2	350.8	353.0	356.0

Electrocapillary data for tetraethylammonium bromide (continued)

				γ/mN m^{-1}				
c/mol l^{-1}	0.007	0.005	0.003	0.002	0.001	0.0007	0.0005	0.0003
-E/V								
0.000				382.4	387.0	389.5	391.4	394.0
0.050			390.6	392.5	396.2	398.0	399.3	401.4
0.100		396.6	399.3	401.2	404.0	405.2	406.2	407.9
0.150	402.5	404.3	406.7	408.4	410.5	411.3	412.0	413.0
0.200	409.4	410.8	412.8	414.2	415.8	416.3	416.8	417.3
0.225								
0.250	414.8	415.9	417.6	418.8	419.9	420.2	420.6	420.7
0.275								
0.300	418.7	419.7	421.2	422.1	423.0	423.2	423.5	423.3
0.325								
0.350	421.3	422.2	423.6	424.3	425.0	425.2	425.4	425.2
0.375								
0.400	422.7	423.6	424.9	425.5	426.1	426.3	426.6	426.2
0.425								
0.450	423.0	424.0	425.2	425.7	426.4	426.6	427.0	426.6
0.475								
0.500	422.4	423.3	424.5	425.1	425.8	426.1	426.6	426.4
0.525								
0.500	420.9	421.8	423.0	423.7	424.6	425.0	425.5	425.5
0.575								
0.600	418.6	419.4	420.8	421.5	422.6	423.2	423.8	424.0
0.650	415.6	416.5	417.9	418.8	420.1	420.8	421.5	422.0
0.700	412.1	412.9	414.5	415.5	417.1	417.9	418.7	419.4
0.750	408.1	408.9	410.6	411.8	413.6	414.5	415.4	416.3
0.800	403.7	404.4	406.3	407.6	409.6	410.7	411.7	412.8
0.850	398.9	399.7	401.7	403.1	405.2	406.5	407.6	408.8
0.900	393.8	394.6	396.7	398.2	400.5	401.9	403.1	404.5
0.950	388.4	389.4	391.5	393.0	395.5	397.0	398.2	399.7
1.000	382.8	383.8	386.1	387.5	390.1	391.8	393.1	394.6
1.050	376.8	378.0	380.3	381.7	384.4	386.2	387.6	389.2
1.100	370.4	371.9	374.2	375.6	378.5	380.4	381.9	383.5
1.150	363.6	365.4	367.7	369.1	372.2	374.2	375.8	377.6
1.200	356.3	358.3	360.6	362.2	365.7	367.7	369.4	371.5

Electrocapillary data on merucry in tetramethylammonium bromide at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

					γ/mN m^{-1}				
c/mol l^{-1}	0.0005	0.0007	0.001	0.002	0.003	0.005	0.007	0.01	0.02
-E/V									
0.000	391.9	390.0	387.8	383.4	379.6	373.2	367.8		
0.050	398.9	397.8	396.4	393.1	390.8				
0.100	404.9	404.6	403.8	401.2	400.0				
0.150	410.1	410.2	409.8	408.0	407.5	405.1	403.7	402.5	398.3
0.200	414.5	414.9	414.8	413.6	413.5	411.9	411.1	409.3	405.8
0.250	418.2	418.7	418.8	418.0	418.1	417.2	416.6	414.9	412.1
0.300	421.1	421.8	421.9	421.5	421.6	421.1	420.7	419.4	417.2
0.350	423.4	424.0	424.2	424.0	424.1	423.9	423.5	422.6	420.9
0.400	425.1	425.6	425.8	425.8		425.6	425.2	424.7	423.5
0.450					426.5	426.5	426.0	425.8	424.8
0.500	426.6	426.9	426.9	426.9	426.5	426.5	426.0	426.0	425.0
0.550	426.5	426.7	426.5	426.5	425.9	425.8	425.2	425.3	424.2
0.600	425.9	425.9		425.4	424.7	424.5	423.9	423.8	422.7
0.650	424.7	424.6		423.8	422.9	422.5	421.9	421.7	420.4
0.700	423.0	422.7	422.1	421.5	420.6	420.0	419.3	418.9	417.5
0.750	420.9	420.4	419.5	418.8	417.8	417.0	416.3	415.7	414.1
0.800	418.3	417.6	416.5	415.6	414.5	413.5	412.7	412.0	410.3
0.850	415.2	414.4	413.0	411.8	410.6	409.5	408.6	407.9	406.2
0.900	411.6	410.7	409.0	407.7	406.3	405.0	404.1	403.4	401.5
0.950	407.6	406.5	404.6	403.1	401.6	400.2	399.1	398.5	396.4
1.000	403.2	402.0	399.8	398.1	396.4	394.9	393.6	393.1	390.6
1.050	398.3	397.0	394.6	392.7	390.9	389.2	387.8	387.1	383.9
1.100	393.0	391.6	389.0	387.0	385.0	383.1	381.7	380.4	
1.150	387.4	385.9	383.2	380.9	378.7	376.7			
1.200	381.4	379.8	377.1	374.6	372.2	370.0	368.7	364.0	354.5

Electrocapillary data for tetramethylammonium bromide (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.03	0.05	0.07	0.2	0.3	0.5	0.7	1.0
-E/V								
0.000								
0.050								
0.100							318.9	312.2
0.150	395.0	390.1	386.4				345.6	339.2
0.200	402.9	399.2	396.2	385.0	380.3	373.3		
0.250	409.8	406.8	404.3	394.3	390.1	384.4	380.4	375.5
0.300	415.3	412.7	410.8	401.9	398.2	393.5	391.2	387.0
0.350	419.5	417.3	415.6	407.9	404.6			395.3
0.400	422.4	420.4	419.0	412.3	409.4	405.8	404.0	401.1
0.450	424.0	422.3	421.1	415.2	412.7	409.5	407.4	405.0
0.500	424.3	423.0	421.9	416.8	414.6	411.8	409.4	407.3
0.550	423.6	422.6	421.7	417.1	415.3	412.7	410.2	408.3
0.600	422.0	421.3	420.4	416.3	414.7	412.4	410.1	
0.650	419.7	419.1	418.3	414.5	413.2	411.1	409.2	407.4
0.700	416.7	416.2	415.5	411.9	410.7	408.8	407.2	405.7
0.750	413.2	412.7	412.0	408.4	407.2	405.8	404.8	403.1
0.800	409.4	408.6	407.9	404.4	403.4	402.0	401.4	399.8
0.850	405.1	404.1	403.3	399.8	398.9	397.6	397.1	395.8
0.900	400.6	399.1	398.3	394.7	393.8	392.6	392.1	391.0
0.950	395.5	393.8	392.9	389.3	388.4	387.1	386.3	385.4
1.000	389.9	388.2	387.2	383.4	382.5	381.2	379.9	379.3
1.050	383.3	382.2	381.2	377.2	376.3	375.0	373.4	372.6
1.100	375.4	375.9	374.5	370.7	369.8	368.4	366.9	365.8
1.150		369.0	367.9	363.6	362.9	361.5	361.2	358.9
1.200	353.5				355.6	354.3		352.6

Electrocapillary data on mercury in tetrabutylammonium chloride at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976)2488.

	γ /mN m^{-1}							
c/mol l^{-1}	0.0005	0.001	0.002	0.005	0.01	0.02	0.05	0.1
-E/V								
0.000	395.4	392.8	387.3	375.0	361.4	348.0	328.2	
0.050	404.4	402.3	398.2	388.6		368.2	353.4	
0.100	411.1				389.9	382.2	371.0	361.3
0.150	415.9	414.0	411.1	404.2				372.9
0.200	418.8	416.9	414.1	407.8	402.0	396.6		380.4
0.250	420.3	418.2	415.4	409.3	404.1	399.0	392.9	384.7
0.300	420.6	418.2	415.3	409.2	404.3	399.4		386.5
0.350	419.7	417.1	414.2	407.9	403.2	398.3	392.9	386.5
0.400	418.0	415.3	412.3	405.8	401.1	396.2	390.8	385.2
0.450	415.6	412.8	409.7	403.2	398.5	393.5	388.0	383.0
0.500	412.7	409.8	406.7	400.1	395.4	390.4	384.9	380.2
0.550	409.4	406.5	403.3	396.8	392.1	387.2	381.6	377.0
0.600	405.8	403.0	399.8	393.3	388.7	383.9	378.2	373.6
0.650	402.0	399.3	396.1	389.8	385.2	380.5	374.8	370.1
0.700	398.1	395.5	392.3	386.2	381.6	377.1	371.5	366.5
0.750	394.2	391.6	388.4	382.4	378.0	373.6	368.1	362.8
0.800	390.2	387.8	384.5	378.6	374.2	369.9	364.5	359.1
0.850	386.3	383.9	380.5	374.7	370.3	366.1	360.8	355.3
0.900	382.3	379.9	376.4	370.7	366.2	362.0	356.7	351.3
0.950	378.2	375.8	372.3	366.5	361.9	357.6	352.4	347.1
1.000	374.1	371.6	367.9	362.1	357.5	353.1	347.7	342.7
1.050	369.8	367.2	363.4	357.5	352.9	348.3	342.9	338.1
1.100	365.2	362.5	358.8	352.8	348.3	343.6	338.0	333.3
1.150	360.2	357.5	353.9	348.1	343.8	339.1		328.4
1.200	354.6	351.9	348.8	343.5		335.3	329.2	323.6

Electrocapillary data on mercury in tetrapropylammonium chloride at various concentrations - given are the interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C.

Reference: H. Menard, F. Kimmerle, Can. J. Chem. 54 (1976) 2488

c/mol l^{-1}			γ/mN m^{-1}			
-E/V	0.0005	0.002	0.003	0.01	0.02	0.03
0.000	402.4	399.7	398.4	394.9	382.2	390.5
0.050	408.7	406.6	405.6	403.1	401.0	399.5
0.100	413.9	412.2	411.5	409.6	407.9	406.4
0.150	418.2	416.8	416.3	414.6	413.0	411.4
0.200	421.5	420.3	419.9	418.3	416.6	414.9
0.250	423.9	422.8	422.4	420.6	418.8	416.9
0.300	423.4	424.3	423.9	421.9	419.8	417.7
0.350	426.1	425.0	424.5	422.1	419.7	417.5
0.400	426.1	424.8	424.3	421.5	418.8	416.4
0.450	425.3	423.9	423.2	420.1	417.1	414.5
0.500	423.9	422.3	421.5	418.0	414.7	412.0
0.550	421.9	420.1	419.2	415.3	411.8	409.0
0.600	419.3	417.3	416.3	412.1	408.4	405.5
0.650	416.3	414.0	412.9	408.5	404.6	401.7
0.700	412.8	410.4	409.2	404.5	400.6	397.7
0.750	409.0	406.4	405.1	400.3	396.3	393.4
0.800	404.9	402.0	400.7	395.8	391.8	388.9
0.850	400.4	397.5	396.1	391.1	387.2	384.3
0.900	395.8	392.7	391.3	386.3	382.4	379.6
0.950	390.9	387.7	386.3	381.3	377.6	374.7
1.000	385.8	382.5	381.1	376.2	372.5	369.8
1.050	380.5	377.2	375.8	371.0	367.4	364.7
1.100	375.1	371.8	370.4	365.6	362.1	359.5
1.150	369.5	366.1	364.8	360.0	356.6	354.0
1.200	363.7	360.3	359.0	354.2	351.0	348.4
1.250	357.7	354.2	352.9	348.2	345.0	342.5
1.300	351.4	347.9	346.6	341.8	338.7	336.3
1.350	344.8	341.1	339.8	335.0	332.0	329.8
1.400	337.9	333.9	332.5	327.7	324.8	322.8

Electrocapillary data for tetrapropylammonium chloride (continued)

	γ /mN m^{-1}					
c/mol l^{-1}	0.05	0.1	0.2	0.3	0.5	1.0
-E/V						
0.000	387.5	383.0	377.8	374.7	369.6	360.8
0.050	397.3	393.1	388.6	385.8	381.3	374.1
0.100	404.5	400.6	396.4	393.8	389.8	383.7
0.150	409.7	405.9	401.9	399.4	395.6	390.3
0.200	413.1	409.3	405.3	402.8	399.3	394.4
0.250	415.0	411.1	407.0	404.6	401.1	396.4
0.300	415.6	411.6	407.4	404.8	401.5	396.8
0.350	415.2	411.0	406.6	404.0	400.6	396.0
0.400	413.8	409.4	405.0	402.2	398.9	394.1
0.450	411.7	407.2	402.6	399.8	396.4	391.6
0.500	409.1	404.4	399.6	396.8	393.4	388.5
0.550	405.9	401.1	396.3	393.3	390.0	385.0
0.600	402.4	397.5	392.6	389.6	386.2	381.3
0.650	398.5	393.6	388.7	385.6	382.3	377.3
0.700	394.4	389.5	384.6	381.5	378.2	373.2
0.750	390.2	385.2	380.4	377.3	373.9	369.0
0.800	385.7	380.8	376.0	372.9	369.6	364.7
0.850	381.2	376.3	371.6	368.5	365.1	360.3
0.900	376.5	371.6	367.0	363.9	360.5	355.7
0.950	371.7	366.9	362.3	359.2	355.9	351.0
1.000	366.8	362.0	357.4	354.4	351.0	346.1
1.050	361.7	357.0	352.4	349.5	340.1	341.0
1.100	356.5	351.9	347.3	344.3	340.9	335.7
1.150	351.1	346.5	341.9	339.0	335.6	330.2
1.200	345.6	341.0	336.4	333.6	330.0	324.5
1.250	339.7	335.3	330.7	327.9	324.3	318.6
1.300	333.7	329.3	324.7	322.0	318.5	312.8
1.350	327.3	323.1	318.7	316.0	312.6	306.9
1.400	320.6	316.7	312.5	309.9	306.6	301.3

Electrocapillary data on mercury in tetraethylammonium chloride at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}								
c/mol l^{-1}	0.0005	0.0007	0.001	0.002	0.003	0.005	0.007	0.01	0.02
-E/V									
0.000	403.0	402.4	401.7	399.9	398.6	397.6	396.0	395.2	392.2
0.050	408.8	408.3	408.0	406.6	405.5	404.5	403.4	402.7	400.2
0.100	413.7	413.4	413.3	412.2	411.3	410.5	409.7	409.0	406.9
0.150	417.9	417.6	417.6	416.9	416.1	415.5	415.0	414.3	412.6
0.200	421.2	421.0	421.2	420.7	420.0	419.5	419.1	418.5	417.1
0.250	423.8	423.6	423.9	423.5	423.0	422.5	422.3	421.7	420.5
0.300	425.7	425.6	425.8	425.5	425.1	424.7	424.5	424.0	422.9
0.350	427.0	426.8	427.0	426.8	426.4	426.0	425.8	425.3	424.3
0.400	427.6	427.3	427.5	427.2	426.9	426.4	426.2	425.8	424.7
0.450	427.5	427.3	427.3	427.0	426.6	426.1	425.9	425.4	424.3
0.500	426.9	426.6	426.5	426.0	425.7	425.1	424.8	424.3	423.0
0.550	425.7	425.3	425.2	424.5	424.1	423.4	423.0	422.5	421.0
0.600	424.0	423.5	423.2	422.4	421.9	421.1	420.6	420.0	418.4
0.650	421.8	421.2	420.8	419.7	419.1	418.2	417.6	416.9	415.1
0.700	419.1	418.3	417.9	416.6	415.8	414.8	414.1	413.3	411.4
0.750	415.9	415.1	414.5	412.9	412.1	410.9	410.1	409.3	407.1
0.800	412.3	411.3	410.7	408.9	407.9	406.6	405.7	404.8	402.5
0.850	408.2	407.2	406.5	404.5	403.4	402.0	401.0	400.0	397.5
0.900	403.8	402.7	401.9	399.7	398.5	397.0	395.9	394.8	392.2
0.950	399.0	397.8	396.9	394.6	393.3	391.7		389.4	
1.000	393.9	392.6	391.6	389.2	387.8	386.1	384.9	383.7	381.0
1.050	388.4	387.0	386.0	383.5	382.1	380.3	379.1	377.8	375.0
1.100	382.6	381.2	380.1	377.5	376.1	374.2	372.9	371.6	368.8
1.150	376.6	375.1	374.0	371.3	369.8	367.9	366.6	365.3	362.5
1.200	370.2	368.7	367.5	364.8	363.3	361.3	360.0	358.6	355.8
1.250	363.7	362.1	360.9	358.1	356.5	354.5	353.1	351.8	348.9
1.300	356.9	355.3	354.0	351.1	349.4	347.3	346.0	344.6	341.7
1.350	350.0	348.3	346.9	343.8	342.1	339.8	338.4	337.0	334.0
1.400	342.8	341.1	339.5	336.3	334.3	332.0	330.5	329.0	325.8

Electrocapillary data for tetraethylammonium chloride (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.03	0.05	0.07	0.1	0.2	0.3	0.5	1.0
-E/V								
0.000	390.2	387.9	386.0	384.3	378.8	375.2	370.5	361.9
0.050	398.5	396.6	395.0	393.5	388.8	385.8	381.7	374.5
0.100	405.5	403.9	402.6	401.2	397.2	394.5	391.0	384.9
0.150	411.4	410.0	408.8	407.6	404.0	401.6	398.5	393.1
0.200	416.0	414.8	413.7	412.5	409.3	407.2	404.3	399.4
0.250	419.6	418.4	417.4	416.3	413.2	411.2	408.5	404.0
0.300	422.0	420.8	419.9	418.7	415.8	413.8	411.2	406.9
0.350	423.4	422.2	421.2	420.1	417.1	415.2	412.6	408.4
0.400	423.8	422.6	421.5	420.4	417.4	415.4	412.8	408.5
0.450	423.4	422.0	420.9	419.7	416.6	414.6	411.9	407.5
0.500	422.0	420.6	419.4	418.2	414.9	412.8	410.1	405.5
0.550	420.0	418.4	417.1	415.8	412.4	410.3	407.4	402.6
0.600	417.3	415.6	414.2	412.8	409.2	406.9	403.9	399.0
0.650	413.9	412.1	410.6	409.1	405.4	403.0	399.9	394.8
0.700	410.0	408.1	406.5	405.0	401.0	398.6	395.3	390.0
0.750	405.7	403.7	402.0	400.4	396.3	393.7	390.3	384.9
0.800	401.0	398.9	397.1	395.4	391.2	388.5	385.0	379.4
0.850	395.9	393.8	391.9	390.1	385.8	383.1	379.4	373.7
0.900	390.6	388.3	386.5	384.6	380.1	377.4	373.7	367.8
0.950	385.0	382.7	380.8	378.9	374.3	371.5	367.8	361.8
1.000	379.2	376.9	375.0	373.0	368.4	365.5	361.8	355.7
1.050	373.2	370.9	369.0	367.0	362.3	359.4	355.7	349.5
1.100	367.0	364.7	362.8	360.8	356.1	353.2	349.5	343.2
1.150	360.6	358.3	356.4	354.5	349.8	346.9	343.1	336.8
1.200	354.0	351.7	349.8	347.9	343.2	340.4	336.7	330.3
1.250	347.1	344.8	343.0	341.1	336.4	333.6	329.9	323.5
1.300	339.8	337.6	335.8	333.9	329.3	326.6	322.9	316.4
1.350	332.2	329.9	328.1	326.2	321.8	319.1	315.4	308.9
1.400	324.0	321.6	319.8	318.0	313.7	311.1	307.3	300.9

Electrocapillary data on mercury in tetramethylammonium chloride at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}						
c/mol l^{-1}	0.001	0.002	0.003	0.005	0.007	0.01	0.02
-E/V							
0.000	401.4	399.7	398.8	397.3	396.1	394.4	391.8
0.050	408.0	406.6	405.7	404.4	403.3	401.8	399.7
0.100	413.2	412.2	411.4	410.3	409.4	408.1	406.5
0.150	417.2	416.7	416.0	415.2	414.4	413.4	412.2
0.200	420.3	420.3	419.7	419.1	418.5	417.7	416.9
0.250	422.7	423.0	422.6	422.1	421.7	421.2	420.6
0.300	424.5	425.0	424.7	424.4	424.1	423.7	423.4
0.325							
0.350	425.7	426.3	426.1	425.9	425.7	425.5	425.3
0.375							
0.400	426.6	427.1	426.9	426.8	426.7	426.5	426.4
0.425							
0.450	427.0	427.3	427.1	427.0	426.9	426.8	426.7
0.475							
0.500	427.0	427.1	426.8	426.7	426.5	426.4	426.2
0.525							
0.550	426.6	426.3	425.9	425.8	425.5	425.4	425.1
0.600	425.8	425.0	424.6	424.3	424.0	423.7	423.4
0.650	424.5	423.3	422.7	422.3	421.9	421.6	421.0
0.700	422.7	421.1	420.3	419.8	419.3	418.8	418.1
0.750	420.4	418.4	417.5	416.7	416.2	415.6	414.6
0.800	417.5	415.1	414.1	413.2	412.6	411.9	410.7
0.850	414.0	411.4	410.3	409.2	408.5	407.8	406.3
0.900	409.8	407.2	406.0	404.8	404.0	403.2	401.5
0.950	405.1	402.5	401.2	399.9	398.9	398.2	396.3
1.000	399.7	397.3	395.9	394.5	393.4	392.9	390.7
1.050	393.7	391.7	390.2	388.7	387.5	387.1	384.8
1.100	387.3	385.8	384.2	382.5	381.0	381.0	378.6
1.150	380.4	379.6	377.7	376.0	374.1	374.6	372.0
1.200		373.2	371.1	369.2	366.6	367.8	365.1
1.250	366.2					360.7	357.8
1.300	359.3					353.3	350.3
1.350	352.9					345.6	342.4
1.400	347.2					337.5	334.1

Electrocapillary data for tetramethylammonium chloride (continued)

	γ /mN m^{-1}						
c/mol l^{-1}	0.05	0.07	0.1	0.2	0.3	0.5	1.0
-E/V							
0.000	388.6	386.7	385.7	381.3	377.8	372.9	365.1
0.050	396.9	395.3	394.2	390.5	387.2	382.9	376.1
0.100	404.1	402.8	401.8	398.7	395.7	391.8	385.9
0.150	410.1	409.1	408.4	405.7	403.0	399.6	394.4
0.200	415.2	414.4	413.8	411.7	409.3	406.2	401.6
0.250	419.2	418.6	418.2	416.5	414.3	411.6	407.4
0.300	422.3	421.8	421.6	420.1	418.2	415.7	411.8
0.325			422.9	421.6	419.7	417.3	413.6
0.350	424.4	424.0	423.9	422.7	420.9	418.6	415.0
0.375			424.6	423.6	421.9	419.7	416.1
0.400	425.6	425.3	425.2	424.2	422.5	420.4	417.0
0.425			425.5	424.5	422.9	420.9	417.5
0.450	426.0	425.7	425.5	424.7	423.1	421.1	417.8
0.475			425.4	424.5	423.0	421.0	417.8
0.500	425.5	425.3	425.0	424.2	422.7	420.7	417.5
0.525			424.4	423.6	422.1	420.2	417.0
0.550	424.3	424.0	423.6	422.9	421.3	419.4	416.3
0.600	422.5	422.1	421.5	420.7	419.2	417.3	414.2
0.650	420.0	419.6	418.8	417.9	416.3	414.4	411.3
0.700	416.9	416.4	415.4	414.5	412.8	410.9	407.8
0.750	413.2	412.7	411.5	410.5	408.7	406.8	403.7
0.800	409.1	408.5	407.2	406.0	404.2	402.2	399.0
0.850	404.6	403.9	402.4	401.1	399.3	397.2	394.0
0.900	399.6	398.8	397.3	395.9	394.1	391.8	388.6
0.950	394.2	393.4	391.9	390.3	388.5	386.2	382.9
1.000	388.5	387.7	386.1	384.4	382.6	380.2	376.9
1.050	382.5	381.7	380.0	378.2	376.4	373.9	370.6
1.100	376.2	375.3	373.5	371.6	369.8	367.4	364.0
1.150	369.5	368.6	366.7	364.7	362.7	360.4	357.0
1.200	362.6	361.6	359.3	357.2	355.1	352.9	349.5
1.250	355.3	354.3					
1.300	347.6	346.6					
1.350	339.5	338.5					
1.400	331.0	329.9					

Electrocapillary data on mercury in tetrabutylammonium fluoride at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE.
T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}					
c/mol l^{-1}	0.00154	0.00339	0.00678	0.00924	0.0101	0.0339
- E/V						
-0.050	405.3	403.7				
-0.025						
0.000	411.5	410.2		407.2	406.6	405.1
0.025					409.9	
0.050	416.1	414.9	412.8	412.7	412.4	408.3
0.100	419.2	417.8	416.0	415.3	415.1	409.4
0.150	421.1	419.2	417.0	415.9	415.6	409.0
0.200	421.8	419.4	416.4	414.9	414.6	407.4
0.250	421.5	418.4	414.7	412.9	412.6	405.0
0.300	420.4	416.6	412.2	410.3	409.9	402.2
0.350	418.5	414.1	409.3	407.3	406.8	399.1
0.400	416.1	411.0	406.0	404.1	403.6	395.8
0.450	413.1	407.6	402.7	400.8	400.3	392.5
0.500	409.8	404.0	399.3	397.5	397.0	389.2
0.550	406.3	400.2	395.9	394.2	393.7	385.9
0.600	402.5		392.5	390.8		382.6
0.650	398.6		389.1	387.4	386.8	379.1
0.700	394.6		385.5	383.8	383.2	375.6
0.750	390.6	385.5	381.8	380.1	379.4	371.9
0.800	386.6	381.9	377.8	376.1	375.4	368.0
0.850	382.5	378.1	373.7	371.9	371.1	363.9
0.900	378.4	374.0	369.3	367.5	366.8	359.5
0.950	374.2	369.4	364.8	363.0	362.4	355.0
1.000	370.0	364.1	360.2	358.6	358.3	350.5
1.050	365.5		355.6	354.4		346.0
1.100	360.7		351.4			341.8

Electrocapillary data for tetrabutylammonium fluoride (continued)

	γ /mN m^{-1}		
c/mol l^{-1}	0.0678	0.101	0.339
-E/V			
-0.050	397.4		391.6
-0.025	400.5		393.1
0.000	402.8	402.3	394.2
0.025	404.4		
0.050		403.3	395.2
0.100		403.1	395.1
0.150		401.9	394.0
0.200		400.0	392.3
0.250		397.6	390.1
0.300		394.8	387.5
0.350		391.9	384.7
0.400		388.9	381.8
0.450		385.7	378.7
0.500		382.6	375.5
0.550	381.8	379.4	372.2
0.600	378.5	376.1	368.8
0.650	375.0	372.7	365.3
0.700	371.4	369.1	361.6
0.750	367.6	365.5	357.8
0.800	363.5	361.6	353.8
0.850	359.3	357.5	349.7
0.900		353.2	345.5
0.950	350.7	348.8	341.2
1.000	346.7	344.3	337.1
1.050		339.9	
1.100		335.6	

Electrocapillary data on mercury in tetrapropylammonium fluoride at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE.
T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ/mN m^{-1}						
c/mol l^{-1}	0.000983	0.00157	0.00233	0.00245	0.00393	0.0049	0.0059
-E/V							
0.000	404.9	405.2	405.3	406.1	405.2	406.7	
0.050	410.7	410.4	410.7	411.1	410.6	411.5	409.2
0.100	415.5	414.9	415.3	415.4	415.3	415.7	415.0
0.150	419.4	418.6	419.1	419.0	419.0	419.3	418.9
0.200	422.4	421.6	422.0	421.8	421.9	422.1	421.8
0.250	424.5	423.8	424.1	423.8	423.9		423.7
0.300	425.8	425.1	425.3		424.9	425.0	424.6
0.350		425.7	425.7	425.5	425.2	425.1	424.6
0.400	426.0	425.4	425.3	425.1	424.6	424.4	423.8
0.450	425.0	424.5	424.1	424.1	423.2	422.9	422.2
0.500	423.4	422.8	422.3	422.4	421.3	420.7	420.1
0.550	421.2	420.6	419.9	420.1	418.7	418.0	417.3
0.600	418.5	417.8	417.0	417.2	415.6	414.8	414.1
0.650	415.4	414.6	413.6	413.9	412.0	411.2	410.4
0.700	411.9	410.9	409.9	410.2	408.2	407.3	406.5
0.750	408.0	406.9	405.8	406.2	404.0	403.2	402.3
0.800	403.9	402.7	401.4	402.0	399.6	399.0	397.8
0.850	399.5	398.2	396.9	397.5	395.0	394.5	
0.900	394.9	393.5	392.1	392.8	390.2	389.8	388.5
0.950	390.1	388.6	387.2	388.0	385.3	384.8	383.6
1.000	385.1	383.6	382.1	383.0	380.1	379.3	378.4
1.050	379.9	378.4	376.8	377.8	374.8		373.1
1.100	374.5	372.9	371.2	372.3	369.2		
1.150	368.8	367.2	365.2	366.5	363.1		361.3
1.200	362.7	361.0	358.8	360.2	356.5		354.7

Electrocapillary data for tetrapropylammonium fluoride (continued)

	γ/mN m^{-1}							
c/mol l^{-1}	0.00983	0.0157	0.0236	0.0393	0.059	0.0983	0.157	0.236
-E/V								
0.000	402.7						398.6	397.2
0.050		408.4	408.7	406.4	406.8	405.4	405.5	404.1
0.100	414.6	414.4		412.6	412.9	411.2	410.1	408.5
0.150		418.6	418.7	416.8	416.9	414.7	412.8	411.1
0.200	421.5	421.4	421.3	419.3	419.0	416.4	414.0	412.1
0.250	423.3	422.9	422.5	420.3	419.7	416.7	413.9	411.8
0.300	424.0	423.3	422.6	420.2	419.2	415.8	412.9	410.6
0.350	423.8	422.7	421.7	419.1	417.8	414.1	411.1	408.7
0.400	422.8	421.3	420.0	417.3	415.6	411.7	408.7	406.2
0.450	421.0	419.3	417.6	414.8	413.0	408.9	405.9	403.3
0.500	418.6	416.7	414.7	411.9	409.9	405.7	402.7	400.0
0.550	415.7	413.6	411.4	408.6	406.5	402.2	399.2	396.6
0.600	412.3	410.1	407.7	404.9	402.9	398.5	395.5	392.9
0.650	408.5	406.3	403.8	401.1	399.0	394.7	391.6	389.1
0.700	404.5	402.3	399.6	397.0	395.0	390.7	387.6	385.2
0.750	400.2	398.0	395.3	392.8	390.8	386.5	383.4	381.0
0.800	395.7	393.7	390.8	388.4	386.5	382.1	379.0	376.8
0.850	391.0	388.9	386.3	383.8	381.9	377.6	374.5	372.3
0.900	386.2	384.1	381.6	379.1	377.2	372.9	369.8	367.6
0.950	381.3	379.2	376.7	374.2	372.2	368.0	365.0	362.8
1.000	376.2	374.1	371.6	369.1	367.1	362.9	359.9	357.7
1.050	370.9	368.8	366.2	363.8	361.7	357.6	354.7	352.4
1.100	365.3	363.2	360.5	358.3	356.2	352.2	349.3	347.0
1.150	359.3	357.4	354.3	352.6	350.6	346.6	343.8	341.5
1.200	352.9	351.2		346.7	344.9	341.2	338.2	336.0

Electrocapillary data on mercury in tetraethylammonium fluoride at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}					
c/mol l^{-1}	0.00108	0.00166	0.00332	0.00498	0.0054	0.00665
-E/V						
0.000	405.6	407.4	406.2	404.8	405.1	405.1
0.050	410.1	411.6	410.9	410.5	410.5	410.5
0.100	414.1	415.5	415.2	415.3	415.1	415.1
0.150	417.7	419.0	419.0	419.2	418.9	419.0
0.200	420.7	422.0	422.0	422.3	422.0	422.0
0.250	423.2	424.3	424.4	424.7	424.2	424.3
0.300	425.0	426.1	426.1	426.2	425.7	425.8
0.350	426.2	427.1	427.1	427.0	426.4	426.6
0.400	426.8	427.5	427.3	427.0	426.5	426.6
0.450	426.7	427.2	426.8	426.4	425.9	426.0
0.500	426.0	426.2	425.7	425.1	424.6	424.6
0.550	424.7	424.6	423.9	423.2	422.8	422.7
0.600	422.8	422.5	421.5	420.8	420.3	420.2
0.650	420.3	419.8	418.6	417.9	417.4	417.1
0.700	417.4	416.6	415.2	414.4	414.0	413.6
0.750	414.0	412.9	411.3	410.6	410.1	409.6
0.800	410.1	408.9	407.1	406.3	405.9	405.2
0.850	405.9	404.5	402.6	401.7	401.2	400.5
0.900	401.3	399.9	397.8	396.7	396.2	395.4
0.950	396.4	394.9	392.6	391.4	390.9	390.0
1.000	391.2	389.6	387.3	385.9	385.3	384.4
1.050	385.7	384.1	381.6	380.0	379.4	378.5
1.100	380.0	378.3	375.7	373.9	373.3	372.3
1.150	374.0	372.1	369.5	367.5	366.9	365.9
1.200	367.7	365.6	362.9	360.8	360.2	359.3
1.250	361.1	358.5	355.8	353.8	353.2	352.3
1.300	354.1			346.5		345.0

Electrocapillary data for tetraethylammonium fluoride (continued)

	γ /mN m^{-1}					
c/mol l^{-1}	0.00997	0.0108	0.0166	0.027	0.03325	0.0498
-E/V						
0.000	404.4	402.7	403.2	403.7	400.8	399.8
0.050	409.8	408.3	408.8	409.0	406.9	406.2
0.100	414.4	413.2	413.6	413.6	412.2	411.6
0.150	418.4	417.4	417.7	417.5	416.5	416.1
0.200	421.5	420.7	420.9	420.6	420.0	419.5
0.250	423.8	423.3	423.3	422.9	422.5	421.9
0.300	425.4	425.0	424.8	424.3	424.0	423.4
0.350	426.2	425.8	425.5	424.9	424.7	424.0
0.400	426.2	425.9	425.4	424.7	424.6	423.8
0.450	425.5	425.3	424.6	423.8	423.6	422.7
0.500	424.1	423.9	423.1	422.1	421.9	420.9
0.550	422.1	421.8	420.9	419.8	419.6	418.5
0.600	419.5	419.1	418.1	416.8	416.6	415.4
0.650	416.4	415.9	414.7	413.3	413.1	411.8
0.700	412.7	412.2	410.9	409.4	409.1	407.7
0.750	408.6	408.0	406.7	405.1	404.7	403.3
0.800	404.2	403.5	402.1	400.4	399.9	398.4
0.850	399.3	398.7	397.1	395.4	394.8	393.3
0.900	394.2	393.5	391.9	390.1	389.4	387.9
0.950	388.8	388.0	386.4	384.6	383.9	382.3
1.000	383.1	382.3	380.7	378.9	378.0	376.5
1.050	377.1	376.3	374.8	372.9	372.0	370.4
1.100	370.9	369.9	368.6	366.6	365.8	364.1
1.150	364.5	363.2	362.1	360.0	359.3	357.6
1.200	357.7		355.4	352.8	352.6	350.8
1.250	350.6		348.2		345.4	343.6
1.300	343.1		340.6		337.8	335.9

Electrocapillary data for tetraethylammonium fluoride (continued)

	γ /mN m^{-1}					
c/mol l^{-1}	0.054	0.0665	0.09975	0.108	0.2716	0.543
-E/V						
0.000	401.5	398.8	397.4	397.9	391.8	387.8
0.050	407.4	405.2	404.2	404.0	398.8	395.6
0.100	412.5	410.6	409.8	409.1	404.4	401.6
0.150	416.6	415.1	414.3	413.3	408.7	405.9
0.200	419.9	418.6	417.8	416.4	411.8	408.8
0.250	422.2	421.1	420.2	418.6	413.7	410.4
0.300	423.6	422.6	421.6	419.8	414.5	410.9
0.350	424.1	423.2	422.1	420.1	414.4	410.3
0.400	423.7	422.9	421.7	419.6	413.4	408.9
0.450	422.5	421.8	420.5	418.2	411.5	406.6
0.500	420.6	420.0	418.5	416.0	409.0	403.8
0.550	418.1	417.4	415.8	413.2	405.9	400.4
0.600	414.9	414.3	412.6	409.9	402.2	396.4
0.650	411.2	410.6	408.8	406.0	398.1	392.1
0.700	407.1	406.4	404.6	401.7	393.6	387.5
0.750	402.6	401.9	400.0	397.1	388.8	382.6
0.800	397.8	397.0	395.0	392.2	383.7	377.4
0.850	392.7	391.8	389.8	387.0	378.5	372.0
0.900	387.4	386.4	384.3	381.6	373.0	366.4
0.950	381.8	380.7	378.7	376.0	367.3	360.6
1.000	376.1	374.9	372.8	370.3	361.5	354.6
1.050	370.0	368.9	366.7	364.2	355.4	348.5
1.100	363.7	362.6	360.4	357.9	349.1	342.0
1.150	357.0	356.1	353.9	351.2	342.4	335.4
1.200	349.8	349.3	347.1	344.0	335.3	328.4
1.250		342.1	340.0			
1.300		334.4	332.4			

Electrocapillary data on mercury in tetramethylammonium fluoride at various concentrations - given are interfacial tension in mN m^{-1}, potential in V vs SCE. T = 25°C

Reference: H. Menard and F.M. Kimmerle. Can. J. Chem. 54 (1976) 2488.

	γ /mN m^{-1}						
c/mol l^{-1}	0.0005	0.0007	0.003	0.00956	0.0191	0.02	0.03
- E/V							
-0.100				393.4	384.7		
-0.050				400.1	394.4		
0.000	408.3	407.6		405.9	402.4		
0.050	411.8	412.0	410.8	410.9	408.9		
0.100	415.1	415.8	415.7	415.2	414.1	413.2	413.1
0.150	418.1	418.9	419.4	418.8	418.3	417.3	417.4
0.200	420.8	421.5	422.2	421.7	421.5	420.7	420.8
0.250	423.0	423.6	424.3	424.0	423.9	423.3	423.3
0.300	424.8	425.2	425.8	425.6	425.6	425.1	425.2
0.350	426.2	426.4	426.6	426.7	426.6	426.3	426.3
0.400	427.1	427.1	427.1	427.2	427.0	426.8	426.8
0.450	427.5	427.3	427.0	427.1	426.9	426.7	426.3
0.500	427.5	427.2	426.5	426.5	426.2	426.1	425.9
0.550	427.0	426.6	425.6	425.4	425.0	424.8	424.6
0.600	426.0	425.6	424.3	423.7	423.3	423.0	422.7
0.650	424.7	424.2	422.5	421.5	421.0	420.7	420.3
0.700	422.8	422.3	420.2	418.9	418.2		417.3
0.750	420.6	420.0	417.5	415.7	414.9	414.5	413.9
0.800	417.9	417.3	414.2	412.0	411.2	410.7	409.9
0.850	414.8	414.0	410.4	407.9	406.8	406.4	405.5
0.900	411.3	410.4	406.1	403.4	402.1	401.6	400.7
0.950	407.4	406.2	401.3	398.4	396.8	396.4	395.4
1.000	403.1	401.7	396.0	393.0	391.2	390.8	389.7
1.050	398.4	396.7	390.4	387.2	385.2	384.8	383.7
1.100	393.2	391.3	384.4	381.0	378.9	378.5	377.3
1.150	387.5	385.5	378.2	374.5	372.4	372.0	370.7
1.200	381.4	379.4	371.9	367.7	365.9	365.2	363.9

Electrocapillary data for tetramethylammonium fluoride (continued)

	γ /mN m^{-1}							
c/mol l^{-1}	0.0382	0.05	0.0956	0.1	0.1912	0.2	0.382	0.956
- E/V								
-0.100	386.3		379.6					
-0.050	394.7		389.6					
0.000	401.9		398.1		396.2		393.0	387.5
0.050	408.1		405.3		403.6		401.2	397.3
0.100	413.3	413.0	411.3	411.9	409.9	411.0	408.1	405.4
0.150	417.5	417.1	416.1	416.3	415.1	415.6	413.8	411.9
0.200	420.9	420.5	419.9	419.8	419.2	419.3	418.2	416.8
0.250	423.5	423.1	422.7	422.6	422.3	422.2	421.5	420.5
0.300	425.3	425.0	424.7	424.7	424.5	424.3	423.8	422.8
0.350	426.5	426.2	425.9	425.9	425.8	425.5	425.1	424.1
0.400	427.0	426.7	426.4	426.5	426.2	426.0	425.5	424.3
0.450	426.8	426.6	426.1	426.2	425.9	425.7	425.1	423.6
0.500	426.0	425.8	425.2	425.4	424.8	424.7	423.9	422.1
0.550	424.7	424.4	423.7	423.8	423.0	423.0	422.0	419.9
0.600	422.7	422.4	421.5	421.6	420.7	420.7	419.4	417.1
0.650	420.3	419.9	418.8	418.9	417.7	417.8	416.3	413.6
0.700	417.2	416.8	415.6	415.6	414.3	414.4	412.7	408.7
0.750	413.7	413.2	411.9	411.8	410.3	410.4	408.5	405.3
0.800	409.7	409.1	407.7	407.5	405.9	406.0	404.0	400.5
0.850	405.3	404.5	403.0	402.9	401.1	401.2	399.1	395.4
0.900	400.3	399.5	397.9	397.8	395.9	396.0	393.7	389.9
0.950	395.0	394.1	392.4	392.3	390.3	390.4	388.1	384.2
1.000	389.3	388.4	386.5	386.5	384.4	384.5	382.1	378.1
1.050	383.2	382.3	380.3	380.4	378.1	378.2	375.8	371.7
1.100	376.9	375.9	373.8	373.9	371.6	371.7	369.2	365.0
1.150	370.2	369.3	367.0	367.1	364.7	364.8	362.2	357.9
1.200	363.4	362.4	360.1	359.9	357.5	357.5	354.9	350.4

Electrocapillary data for mercury in solutions of 0.1 M sodium formate in formic acid containing di-n-butyl ether. T = 25°C

Slight sticking occurred in the negative branch of the four most concentrated solutions.

Potential with respect to a saturated calomel electrode in formic acid.

Reference: J. Lawrence and R. Parsons. Trans. Faraday Soc. 64 (1968) 1656.

	γ /mN m^{-1}							
c/mol l^{-1}	0.0099	0.0176	0.0277	0.0328	0.0360	0.0473	0.0665	0.0792
E/volts								
0.00	383.0	382.9	382.7	382.4	382.6	382.8	382.3	382.3
-0.05	386.3	386.4	386.2	385.9	385.8	386.0	385.7	385.5
-0.10	389.3	389.3	389.0	388.7	388.7	388.9	388.5	388.2
-0.15	391.7	391.8	391.5	391.2	391.2	391.2	390.9	390.3
-0.20	393.8	393.7	393.4	393.3	393.1	393.3	392.7	391.9
-0.25	395.4	395.4	395.1	395.0	394.7	394.7	393.9	393.4
-0.30	396.7	396.7	396.4	395.9	395.6	395.8	395.0	394.2
-0.35	397.5	397.5	397.2	396.7	396.6	396.4	395.5	394.4
-0.40	398.3	398.3	397.5	397.2	397.0	396.6	395.7	394.4
-0.45	398.6	398.5	397.6	397.3	397.1	396.5	395.6	394.2
-0.50	398.8	398.6	397.5	397.2	397.0	396.2	395.2	393.6
-0.55	398.8	398.5	397.4	396.9	396.4	395.6	394.3	392.8
-0.60	398.4	398.1	396.7	396.1	395.7	394.7	393.3	391.8
-0.65	397.8	397.4	395.8	395.2	394.7	393.7	392.2	390.6
-0.70	397.1	396.3	394.8	393.9	393.5	392.5	390.9	389.3
-0.75	396.9	395.0	393.5	392.7	392.1	391.0	389.4	387.9
-0.80	394.6	393.7	392.0	391.1	390.6	389.5	387.8	386.2
-0.85	393.2	392.3	390.2	389.4	389.0	387.8	386.1	384.4
-0.90	391.4	390.3	388.4	387.5	387.1	385.8	384.3	382.6

Electrocapillary data on mercury for aqueous solutions of narcotine of concentration c_N in mole.l^{-1} in 1 N HCl. T = 25°. Potential measured with respect to a hydrogen electrode in 1 N HCl. Given is the interfacial tension x 10^3 in N.m^{-1}.

Reference: B.E. Conway, J.O'M. Bockris and B. Lovrecek,
Proc. 6 th Meeting of CITCE, 207 (1955).

E / Volts	$c_N = 0$	E / Volts	10^{-6}M	5 x 10^{-6}M	10^{-5}M	5 x 10^{-5}M	10^{-4}M
-0.900	374.3	-0.892	373.2	373.8	371.9	356.7	353.3
-0.850	381.3	-0.842	379.9	380.9	378.7	361.9	360.2
-0.800	387.4	-0.792	386.6	386.7	385.4	366.9	367.0
-0.750	393.5	-0.742	391.7	392.2	391.9	372.5	373.4
-0.700	399.3	-0.692	397.3	397.7	396.7	376.5	378.9
-0.650	403.6	-0.642	402.4	403.5	401.5	381.7	381.2
-0.600	408.2	-0.592	406.5	407.1	405.4	384.9	383.8
-0.550	411.3	-0.542	410.5	410.8	409.3	388.1	385.8
-0.500	414.5	-0.492	415.1	414.3	413.6	391.5	387.4
-0.450	418.2	-0.442	417.2	417.1	416.2	394.8	390.3
-0.400	420.6	-0.392	419.3	419.3	418.2	396.5	391.1
-0.350	421.9	-0.342	420.4	421.0	419.4	397.9	393.4
-0.325	422.2	-0.292	421.5	421.6	419.9	399.8	394.4
-0.300	422.6	-0.267	421.9	421.8	420.3	-	-
-0.275	422.1	-0.242	421.2	421.3	419.7	404.6	395.3
-0.250	421.8	-0.192	420.2	419.7	418.8	402.8	396.0
-0.200	421.0	-0.142	418.6	418.3	417.1	404.1	396.5
-0.150	418.8	-0.092	414.6	416.0	414.3	402.1	396.4
-0.100	415.9	-0.042	410.9	411.9	409.5	400.4	395.1
-0.050	411.4	+0.008	405.5	407.3	404.4	398.9	392.2
-0.000	406.5	+0.058	399.5	401.5	399.0	398.3	389.3
+0.050	401.4	+0.108	392.3	394.6	391.6	392.5	384.9
+0.100	393.3	+0.158	384.2	384.6	383.9	383.9	377.3
+0.150	385.1	+0.208	375.0	376.9	374.3	374.5	370.0
+0.200	375.5	+0.258	362.8	364.7	362.1	362.0	358.9
+0.250	363.3						

Electrocapillary data on mercury for aqueous solutions of narcotine of concentration c_N in mole.l^{-1} in 1 N HCl. (cont.)

$\frac{E}{\text{Volts}}$	5×10^{-4}M
-0.884	343.8
-0.834	350.9
-0.784	356.9
-0.734	362.2
-0.684	365.8
-0.634	369.1
-0.584	372.1
-0.534	375.6
-0.484	379.0
-0.434	382.3
-0.384	384.7
-0.334	387.0
-0.284	388.3
-0.234	389.9
-0.184	390.8
-0.159	391.4
-0.134	391.6
-0.109	391.4
-0.084	390.8
-0.034	389.4
+0.016	386.6
+0.066	383.3
+0.116	379.2
+0.166	373.0
+0.216	365.2
+0.266	353.9

Electrocapillary data on mercury for aqueous solutions of quinine (concentration c_Q in mole.l^{-1} indicated) in 0.1 m HCl. T = 25°.

Potential measured with respect to hydrogen electrode in 0.1 m HCl. Given is the interfacial tension x 10^3 in N.m^{-1}.

Reference: B.E. Conway, J.O'M. Bockris and B. Lovrecek,
Proc. 6 th Meeting of CITCE, 207 (1955).

$\frac{E}{\text{Volts}}$	$c_Q = 0$[*]	10^{-6}	10^{-5}	10^{-4}	10^{-3}	10^{-2}
-0.900	369.3	369.5	368.8	341.6	327.3	315.7
-0.850	376.7	376.0	375.3	361.4	332.7	322.9
-0.800	382.7	382.1	381.8	357.0	337.6	328.3
-0.750	388.3	388.1	387.9	363.9	342.9	332.7
-0.700	393.7	393.3	393.2	375.8	349.5	338.6
-0.650	399.3	398.5	398.2	377.9	356.5	346.3
-0.600	404.3	403.1	402.6	376.2	364.4	355.3
-0.550	408.5	407.9	407.2	387.3	369.7	362.0
-0.500	412.0	411.2	410.9	394.5	379.3	370.4
-0.450	415.3	414.6	414.5	404.2	387.7	379.0
-0.400	418.6	417.5	416.3	406.9	397.0	388.2
-0.350	420.9	420.2	419.4	409.3	401.9	393.1
-0.300	422.7	422.0	421.5	411.2	404.3	396.2
-0.250	424.2	423.6	422.6	412.9	406.2	398.2
-0.200	425.5	424.4	424.2	415.1	407.4	399.8
-0.185	425.7	424.8				
-0.170	425.6	424.8				
-0.150	425.4	424.6	424.4	417.0	408.5	400.6
-0.100	424.9	424.1	423.6	416.0	408.8	400.8
-0.050	423.3	422.5	422.2	415.2	407.7	400.1
0.000	420.7	420.2	419.8	412.4	406.5	399.1
+0.050	416.7	416.3	415.9	408.6	404.2	397.3
+0.100	412.0	411.8	411.7	406.7	401.5	394.1
+0.150	406.4	406.9	406.2	403.0	396.9	390.4
+0.200	400.3	400.3	400.5	397.4	391.7	385.3
+0.250	392.9	393.2	393.2	391.9	385.8	379.4

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[*] See discussion remark on p. 175.

Electrocapillary data on mercury in aqueous solutions of 0.5 M Na_2SO_4 + caffeine of varying concentrations. Potential measured with respect to a normal calomel electrode. Temperature 20°C. Given is the interfacial tension in $mN.m^{-1}$. Reference: J.A.V. Butler and C. Ockrent, J. Phys. Chem., 34, 2841 (1930).

conc. of caffeine in mole l^{-1}. $\frac{E}{volts}$	0.0033	0.01	0.02
+0.1	367.0	368.0	367.5
0.0	387.3	384.4	382.3
-0.1	394.3	389.8	386.5
-0.2	396.4	390.8	388.3
-0.3	397.3	392.5	389.1
-0.4	397.2	392.1	388.8
-0.5	396.1	391.2	387.5
-0.6	394.6	388.9	385.9
-0.7	391.4	386.7	383.2
-0.8	388.8	383.3	379.6
-0.9	384.7	380.0	375.4
-1.0	380.6	374.9	370.6
-1.1	376.9	369.1	364.5
-1.2	369.2	363.9	357.7
-1.3	354.7	351.9	347.0
-1.4	338.3	336.9	334.6

Electrocapillary data on mercury in aqueous solutions of 0.5 M Na_2SO_4 + 0.01 M caffeine + phenol of various concentrations. Potential measured with respect to a normal calomel electrode. Temperature 20°C. Given is the interfacial tension in mN. m^{-1}. Reference: J.A.V. Butler and C. Ockrent, J. Phys. Chem., 34 2841 (1930).

conc. of caffeine in mole l^{-1}.	0.005	0.01	0.02	0.05	0.1
E / volts					
+0.1	366.8	364.2	361.9	356.8	353.5
0.0	380.9	375.9	366.2	367.1	362.6
-0.1	386.2	383.8	380.7	373.9	367.0
-0.2	390.2	388.4	385.6	380.0	372.5
-0.3	392.0	391.0	388.8	383.8	377.4
-0.4	392.1	391.4	389.5	385.6	379.8
-0.5	391.1	390.8	389.2	386.4	380.5
-0.6	389.0	389.2	387.4	385.9	380.5
-0.7	386.5	386.6	385.0	384.6	379.9
-0.8	383.3	383.6	382.7	381.2	377.6
-0.9	380.0	378.4	378.6	378.3	374.1
-0.10	374.7	474.8	374.9	374.1	370.3
-0.11	369.1	369.8	369.3	369.2	366.6
-0.12	364.2	364.1	364.6	363.7	364.1
-0.13	352.0	353.4	353.1	351.9	353.3
-0.14	337.0	337.2	337.1	336.9	337.3

Electrocapillary data on mercury in aqueous solutions of 0.5 M Na_2SO_4 + 0.05 M phenol + caffeine of two concentrations. Potential measured with respect to a normal calomel electrode. Temperature 20°C. Given is the interfacial tension in $mN.m^{-1}$. Reference: J.A.V. Butler and C. Ockrent, J. Phys. Chem., 34 2841 (1930).

conc. of caffeine in mole l^{-1}.	0.0033	0.02
$\frac{E}{\text{volts}}$		
+0.1	356.9	356.6
0.0	369.0	366.9
-0.1	376.3	372.4
-0.2	382.3	377.9
-0.3	386.5	381.4
-0.4	388.8	382.6
-0.5	389.5	383.1
-0.6	389.2	382.3
-0.7	387.3	380.1
-0.8	385.4	377.5
-0.9	382.2	374.0
-0.10	379.8	369.2
-0.11	376.5	364.1
-0.12	369.1	357.3
-0.13	354.7	347.1
-0.14	337.9	334.3

Electrocapillary data on mercury electrode in 0.2 M HCl.
Adsorption of hyamine+), given are the interfacial tension in mN m^{-1},
potential in mV (SCE). T = 25°C

Reference: B. Lovrecek, N. Bon, Lj. Duic, M. Bravar. Tenside 8 (1971) 140.

	γ /mN m^{-1}				
c/mol l^{-1} E/mV	+10^{-3}	+10^{-4}	+10^{-5}	+10^{-6}	0
0	362.2	380.6	393.0	394.0	396.9
-80	374.4	385.4	399.8	403.7	406.0
-160	376.9	385.4	411.0	412.5	413.4
-240	376.9	385.6	416.5	418.4	419.3
-320	374.4	385.4	418.4	422.0	420.7
-400	372.0	381.8	420.8	422.0	424.2
-480	369.6	379.3	420.8	422.0	424.2
-560	361.7	374.4	419.0	420.8	421.8
-640	359.2	369.6	418.4	418.4	419.3
-720	352.5	366.7	410.0	411.0	413.0
-800	342.6	362.0	398.8	405.7	407.7
-880	336.7	354.4	398.5	398.8	401.4
-960	328.0	347.6	386.5	391.5	396.1
-1040	320.8	342.7	377.4	381.6	386.3
-1120	307.5	333.0	368.7	374.6	372.5
-1200	298.3	379.4	357.4	361.0	366.2

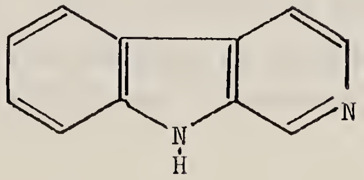
+) Hyamine is $(CH_3)_3CCH_2C(CH_3)_2C_6H_4(OC_2H_4)_2N^+(CH_3)_2CH_2C_6H_5\ Cl^-$.

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4. Adsorption of polynuclear bases from 5×10^{-5} m solutions. Given are the interfacial tension in mN m^{-1}, potential in V vs NCE. T = 12°C

Reference: J.A.V. Butler and W.O. Kermack. Proc. Roy. Soc. Edinburgh, 49 (1929) 300.

-E/V	5:6-benz-4-carboline methosulphate	4-methyl-5:6-benz-4-carboline methosulphate	3-ethyl-5:6-benz-4-carboline methosulphate	3-methyl-5:6-benz-4-carboline methosulphate	5:6-benz-4-carboline hydrochloride	1-methyl-5:6-benz-4-carboline hydrochloride	1:3-dimethyl-5:6-benz-4-carboline hydrochloride
0.0	-	-	368	-	366	372	379
0.1	384	-	378	354	373	-	390
0.2	388	-	381	368	380	381	397
0.3	388	377	387	-	386	-	406
0.4	-	386	392	376	391	381	412
0.5	403	393	397	388	396	-	416
0.6	408	400	400	401	398	397	419
0.7	413	410	400	406	400	400	420
0.8	414	414	401	414	400	400	414
0.9	405	-	-	406	-	394	-
1.0	394	395	395	395	394	392	394
1.1	382	-	-	-	-	-	-
1.2	368	369	369	369	368	368	368
1.3	353	-	-	-	-	-	-
1.4	335	335	336	336	335	335	335
1.5	316	-	-	-	-	-	-
1.6	295	296	296	-	294	295	295

4-carboline is



Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 at 12°C.
Given are the interfacial tension in mN m^{-1}, potential in V vs NCE.

Reference: J.A.V. Butler and C. Ockrent. J. Phys. Chem. 34 (1930) 2297.

E/V	γ /mN m^{-1}								
	0.5 M sodium sulphate (C)	0.05 M sodium cinnamate	0.05 M sodium cinnamate + 0.05 M sodium trichloracetate	0.05 M sodium cinnamate + 0.05 M potassium benzosulphonate	0.05 M sodium terephthalate	0.05 M sodium cinnamate + 0.05 M sodium terephthalate	0.05 M sodium cinnamate + 0.05 M sodium phthalate	0.05 M sodium toluate	0.05 M sodium salicylate + 0.05 M sodium toluate
+0.1	366.9	345.2					344.3	358.3	341.8
0.0	387.7	360.3	359.8	360.2	365.7	359.1	360.3	373.3	357.8
-0.1	403.8	370.9	370.7	371.6	379.9	370.8	370.7	385.2	369.2
-0.2	415.4	379.7	379.8	379.8	391.5	379.5	379.7	394.3	378.2
-0.3	422.5	387.1	387.4	387.0	401.0	387.0	387.3	401.6	387.1
-0.4	425.7	392.4	392.7	392.4	409.2	392.0	391.6	408.0	393.5
-0.5	426.3	397.1	397.5	397.5	413.9	397.3	395.1	411.6	399.1
-0.6	424.7	399.5	399.4	399.5	416.7	399.5	397.8	413.2	401.6
-0.7	420.2	400.5	400.5	400.4	416.2	400.7	398.3	412.4	403.2
-0.8	414.1	399.2	399.1	399.0	412.5	399.2	398.4	408.3	402.7
-0.9	406.0	395.5	395.4	395.3	405.0	394.9	395.4	402.4	399.2
-1.0	395.5	389.4	389.4	389.1	395.3	389.1	389.4	393.5	392.0
-1.1	383.4	380.7	381.0	380.9	383.5	380.6	380.3	382.2	382.0
-1.2	370.6	369.1	369.3	368.9	370.7	368.8	369.2	370.0	370.1
-1.3	355.6	355.1	355.3	354.9	356.0	353.5	355.2	355.7	355.7
-1.4	338.6	338.7	338.7	338.2	339.6	337.8	338.5	339.1	339.5
-1.5	320.8	320.7	321.0	320.5	321.5	320.1	320.4	320.3	321.8
-1.6	300.3	300.8	300.9	300.5	301.6	300.4	301.1	300.4	301.3
-1.7	278.5	279.2	279.3	278.9	279.7	278.7	279.2	278.9	279.9
-1.8	254.3	255.8	255.5	255.5	256.2	255.0	255.5	255.0	256.1
-1.9	228.7								
-2.0	200.3								

Electrocapillary data in 0.5 M Na_2SO_4 (continued)

E/V	γ /mN m^{-1}							
	0.05 M potassium benzene sulphonate + 0.05 M sodium trichloracetate	0.05 M sodium terephthalate + 0.05 M sodium trichloracetate	0.05 M tetra methylammonium hydroxide	0.05 M tetra methylammonium hydroxide + 0.05 M p-toluene sulphonic acid	0.05 M phenol	0.05 M lactose	0.05 M phenol + 0.05 M lactose	0.01 M caffeine
+0.1					352.8	366.3	353.2	368.2
0.0	374.8	365.6			362.3	383.7	361.6	382.8
-0.1	386.0	379.6		382.5	372.8	395.9	372.4	388.1
-0.2	395.4	391.1		391.7	381.1	400.9	381.2	390.3
-0.3	402.4	401.1	418.8	399.5	387.9	406.6	387.8	391.2
-0.4	407.6	408.2	423.0	405.0	391.8	409.1	391.9	391.2
-0.5	410.8	412.4	424.6	409.0	394.2	409.8	394.5	390.1
-0.6	413.9	415.9	422.2	410.2	395.9	409.0	395.0	388.4
-0.7	413.5	416.5	417.4	408.1	396.0	406.6	395.7	386.0
-0.8	410.5	412.5	409.9	403.3	395.0	402.7	394.9	382.8
-0.9	404.1	405.2	400.8	395.9	392.2	396.2	391.9	379.1
-1.0	395.1	395.8	389.8	387.4	388.2	387.7	386.2	375.5
-1.1	383.1	383.9	378.1	376.5	380.8	377.1	377.0	369.0
-1.2	370.8	371.4	364.3	364.4	369.2	365.4	365.7	363.1
-1.3	355.7	356.4	349.4	350.3	355.2	351.9	352.0	353.2
-1.4	339.4	339.9	333.6	334.5	339.4	336.6	336.8	337.7
-1.5	321.6	321.6	316.1	316.7	321.5	319.5	319.6	320.5
-1.6	301.3	301.9	297.3	297.9	301.9	299.9	300.0	301.5
-1.7	279.7	280.1	276.6	276.9	280.2	278.3	278.3	279.8
-1.8	256.3	256.5	254.7	254.2	256.8	255.7	256.0	256.3
-1.9			229.4	229.2	230.9	230.5	230.7	230.8
-2.0			202.0		203.2	203.0	203.2	202.9

Electrocapillary data in 0.5 M Na_2SO_4 (continued)

E/V	γ /mN m^{-1} 0.01 M caffeine + 0.05 M lactose	0.01 M caffeine + 0.05 M phenol	0.01 M caffeine + sat. p.toluidine	0.05 M aniline	0.05 M aniline + 0.05 M phenol	0.05 M phenol + 0.05 M sodium benzoate
+0.1		352.4		347.0	345.2	340.8
0.0	382.1	360.7	363.1	365.7	360.4	356.9
-0.1	387.1	372.4	373.6	378.7	369.3	368.1
-0.2	389.9	379.6	377.9	384.6	375.6	376.4
-0.3	390.7	383.7	383.9	387.4	378.8	382.8
-0.4	391.1	385.7	380.3	389.0	381.1	387.1
-0.5	390.0	386.3	380.3	388.7	381.9	389.7
-0.6	388.3	386.1	379.5	388.2	381.8	391.3
-0.7	386.3	384.6	377.8	386.1	380.7	392.1
-0.8	383.0	382.1	376.1	384.4	379.1	391.0
-0.9	379.1	377.9	373.1	380.9	376.5	389.3
-1.0	375.3	374.8	368.6	377.9	373.2	387.0
-1.1	369.9	368.7	364.8	373.1	368.2	380.3
-1.2	363.1	363.3	359.2	365.9	362.3	369.2
-1.3	350.6	352.0	352.4	354.0	352.7	355.0
-1.4	335.5	337.5	337.9	338.3	337.7	339.1
-1.5	319.8	320.2	320.4	320.4	319.9	320.9
-1.6	299.9	301.0	301.2	300.2	299.9	301.5
-1.7	278.6	279.1	279.5	287.6	278.4	279.4
-1.8	255.9	255.9	255.9	254.5	254.3	255.7
-1.9	230.5	230.1	230.4	228.7	228.7	229.9
-2.0	203.0	202.4	202.9	200.3	200.3	201.8

Electrocapillary data for mercury electrode in 0.5 mol l^{-1} Na_2SO_4 (A, B, C, D, E, or F); adsorption of various compounds at 18 °C. Given are the interfacial tension in mN m^{-1}, potential in V vs Hg/Hg_2SO_4, 0.5 mol l^{-1} Na_2SO_4 electrode. The solutions indicated by capital letters served as base solutions for measurements in the presence of organic substances x).

Reference: G. Gouy. Ann. Chim. Phys. 8 (8) (1906) 291.

γ/mN m^{-1}

-E/V	0.5 mol l^{-1} Na_2SO_4					
	A	B	C	D	E	F
0.1	326.3	326.0	326.8	325.4	325.1	324.7
0.2	353.1	352.8	353.7	352.4	352.0	351.8
0.3	375.3	375.1	375.9	374.8	374.6	374.4
0.4	393.5	393.4	394.1	392.3	392.8	392.7
0.5	407.8	407.7	408.3	407.6	407.2	407.2
0.6	418.0	417.9	418.4	417.9	417.5	417.2
0.7	424.2	423.9	424.5	424.4	424.1	424.2
0.8	427.1	426.9	426.7	427.1	427.1	426.7
0.9	426.9	427.0	427.0	427.0	427.1	427.0
1.0	424.2	424.4	424.0	424.4	424.4	424.4
1.1	419.3	419.3	419.0	419.3	419.5	419.4
1.2	412.1	412.4	412.0	412.4	412.5	412.5
1.3	403.3	403.6	403.1	403.7	403.7	403.8
1.4	392.9	393.1	392.6	393.3	393.2	393.3
1.5	380.7	381.0	380.5	381.2	381.1	381.3
1.6	367.0	367.4	366.8	367.6	367.4	367.7
1.7	351.8	356.4	351.4	352.3	352.2	352.5
1.8	334.6	335.1	334.5	335.4	335.2	335.5
1.9	316.0	316.6	315.8	316.9	316.4	317.0
2.0	295.5	296.0	295.2	296.3	296.0	296.9
2.1	-	-	272.9	274.2	273.6	-
2.2	-	-	248.3	250.2	249.1	-
2.3	-	-	221.8	224.1	222.6	-
2.4	-	-	193.2	196.2	194.0	-

x) Gouy's measurements of electrocapillary curves of organic compounds in 0.5 mol l^{-1} Na_2SO_4 were made over a number of years. During this period he found some variation in the results obtained with the base electrolyte which he attributed to variations in the mercury-mercurous sulphate reference electrode. He therefore recommends that the curves with organic additives should be compared with the measurements on a base solution made about the same time. Therefore, these base solutions are identified by the letters A, B, C, D, E, F and the appropriate base is indicated at the top of each column in the following tables.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)$^{x)}$

γ/mN m^{-1}

-E/V	sat trimethyl ethylene C	1 mol l^{-1} methyl alcohol C	1 mol l^{-1} ethyl alcohol F	1 mol l^{-1} normal propyl alcohol A	1 mol l^{-1} isopropyl alcohol C	0.1 mol l^{-1} normal butyl alcohol A
0.1	-	326.1	323.6	324.9	325.6	326.1
0.2	-	353.0	351.0	352.0	352.7	353.1
0.3	-	-	373.0	374.6	375.0	375.5
0.4	393.8	392.4	391.6	391.0	392.6	393.3
0.5	407.8	-	405.1	395.7	404.2	406.0
0.6	416.9	415.7	413.5	396.3	406.0	408.1
0.7	418.0	-	417.0	396.0	405.2	408.4
0.8	417.8	424.1	417.5	394.7	403.3	407.6
0.9	-	-	415.8	392.9	401.1	406.0
1.0	415.2	420.9	412.9	390.3	394.5	404.0
1.1	-	-	408.7	387.4	395.2	401.2
1.2	409.1	409.5	403.3	383.6	391.0	397.5
1.3	-	-	396.5	379.3	385.3	392.9
1.4	392.0	390.9	388.0	374.1	379.5	387.1
1.5	-	-	377.7	367.4	372.2	378.1
1.6	367.0	361.6	365.6	359.7	362.4	365.9
1.7	-	-	351.2	348.4	349.5	351.0
1.8	334.7	334.2	335.0	333.2	333,8	334.3
1.9	316.1	315.7	316.7	315.1	315.6	315.5
2.0	-	295.1	296.6	294.9	295.4	295.2

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)$^{x)}$

	γ/mN m^{-1}				
-E/V	0.1 mol l^{-1} secondary butyl alcohol C	0.1 mol l^{-1} tertiary butyl alcohol C	0.1 mol l^{-1} isobutyl alcohol D	0.1 mol l^{-1} tertiary amyl alcohol C	sat normal heptyl alcohol C
0.1	326.7	326.7	325.4	326.8	327.0
0.2	353.6	353.3	348.1	353.6	353.8
0.3	375.8	375.8	374.2	375.8	376.1
0.4	393.6	393.8	393.1	393.8	393.5
0.5	407.2	407.6	406.5	404.9	393.1
0.6	413.7	417.0	412.0	404.5	394.1
0.7	413.3	419.1	412.8	403.4	394.0
0.8	412.0	417.7	412.0	402.0	393.8
0.9	410.0	415.8	410.6	400.3	394.7
1.0	407.5	413.0	408.8	398.5	393.8
1.1	404.6	409.6	406.0	395.3	393.9
1.2	400.2	405.5	402.7	392.0	392.8
1.3	395.5	399.6	397.9	388.3	391.4
1.4	388.6	390.8	390.3	383.8	388.7
1.5	378.5	379.5	379.7	377.7	378.8
1.6	365.8	366.3	367.0	365.8	366.4
1.7	350.9	351.2	352.1	351.1	351.4
1.8	334.2	334.2	335.4	334.4	334.6
1.9	315.8	315.5	316.7	315.9	315.9
2.0	-	-	-	-	-

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E, or F.); adsorption of various compounds (continued)$^{x)}$

	γ/mN m^{-1}					
-E/V	1 mol l^{-1} allyl alcohol A	sat benzyl alcohol E	sat anisyl*) alcohol E	1 mol l^{-1} ethylene glycol D	1 mol l^{-1} isobutyl glycol C	sat pinacol+) hydrate A
0.1	325.8	322.8	323.8	-	325.5	325.5
0.2	351.9	346.7	347.1	351.6	352.2	352.6
0.3	374.1	361.9	359.8	-	373.9	375.1
0.4	390.2	367.7	365.3	391.5	390.6	392.6
0.5	399.4	370.6	367.7	-	398.2	405.8
0.6	402.5	380.6	368.9	413.9	401.6	408.5
0.7	403.1	372.6	365.5	-	401.9	408.2
0.8	402.6	372.5	370.1	420.5	401.3	406.7
0.9	401.2	371.8	369.8	-	399.5	404.4
1.0	398.7	370.5	368.5	416.8	397.4	401.6
1.1	395.5	368.4	366.5	-	394.3	398.1
1.2	391.3	365.9	363.9	405.7	390.0	394.4
1.3	386.5	362.7	360.4	-	385.0	389.1
1.4	380.2	358.9	356.0	388.3	378.9	383.9
1.5	372.2	350.3	350.6	-	371.4	376.6
1.6	362.1	349.1	344.3	364.5	361.5	364.9
1.7	348.9	342.8	337.1	-	348.7	350.8
1.8	333.2	332.2	327.6	333.6	333.4	334.3
1.9	315.2	315.6	314.8	-	315.5	315.6
2.0	-	295.1	295.6	295.6	295.7	295.1

x) See footnote on the page containing the base solutions.

+) Pinacol is 2,3 dimethyl-2,3 butane diol.

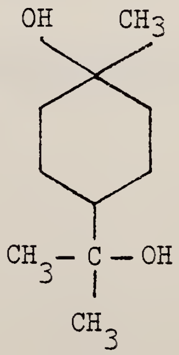
*) Anisyl alcohol is p-methoxy benzyl alcohol.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)x)

	γ/mN m^{-1}					
-E/V	sat terpin hydrate +) A	1 mol l^{-1} glycerol F	1 mol l^{-1} erythritol B	0.1 mol l^{-1} mannitol C	0.1 mol l^{-1} galactitol B	0.1 mol l^{-1} perseitol C
0.1	326.3	323.5	323.8	326.0	325.2	324.8
0.2	-	350.4	350.0	352.4	351.6	351.5
0.3	393.2	372.1	371.3	374.3	373.4	307.4
0.4	397.8	389.4	386.0	392.0	391.2	390.9
0.5	403.3	402.3	400.9	405.6	405.1	404.8
0.6	402.8	411.0	409.5	415.4	415.0	414.6
0.7	402.0	415.8	414.7	421.3	421.1	420.6
0.8	-	417.9	417.2	424.1	423.9	423.4
0.9	398.0	417.7	417.2	424.1	424.1	423.3
1.0	-	415.4	415.3	421.6	419.2	421.2
1.1	392.8	411.0	410.9	417.0	417.2	416.9
1.2	-	404.8	404.8	410.4	410.5	410.3
1.3	385.2	396.9	396.8	401.9	402.0	402.0
1.4	-	387.2	387.1	391.7	391.8	391.7
1.5	375.8	376.1	375.7	379.6	379.8	379.7
1.6	-	363.3	362.7	366.0	366.2	366.2
1.7	350.8	348.8	348.0	350.8	351.3	350.8
1.8	-	332.6	331.7	333.8	334.3	333.9
1.9	315.6	314.5	313.6	315.2	315.8	315.5
2.0	-	294.8	293.7	295.0	295.5	294.9

x) See footnote on the page containing the base solutions.

+) (cis-) terpin hydrate is terpinol:



Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)x)

	γ/mN m^{-1}					
-E/V	0.1 mol l^{-1} quercitol +) C	0.1 mol l^{-1} inositol *) E	1 mol l^{-1} xylose E	1 mol l^{-1} glucose E	0.1 mol l^{-1} galactose C	1 mol l^{-1} saccharose E
0.1	326.4	-	318.4	317.6	326.4	320.7
0.2	352.8	350.9	348.7	347.4	352.6	345.8
0.3	374.5	-	-	-	374.3	365.6
0.4	392.1	392.1	384.8	383.9	391.5	380.6
0.5	405.8	-	-	-	404.4	391.4
0.6	415.6	415.1	404.6	404.2	413.3	398.0
0.7	421.7	-	-	-	418.3	401.9
0.8	424.5	423.8	410.8	411.8	420.4	403.4
0.9	424.6	-	-	-	420.2	403.2
1.0	422.1	422.0	409.5	410.3	418.3	401.5
1.1	417.5	-	-	-	414.3	398.2
1.2	410.7	411.0	402.4	401.8	408.1	393.6
1.3	402.6	-	-	-	400.0	387.4
1.4	-	392.1	387.0	385.8	389.8	380.0
1.5	380.0	-	-	-	378.1	370.7
1.6	-	366.5	364.1	362.4	364.8	359.4
1.7	351.1	-	-	-	349.7	346.3
1.8	-	334.4	333.6	331.9	333.0	331.0
1.9	315.7	-	-	-	314.6	313.5
2.0	-	295.3	-	294.6	294.4	294.3

x) See footnote on the page containing the base solutions.

+) Quercitol is 1,2,3,4,5 cyclohexane pentol.

*) Inositol is 1,2,3,4,5,6 cyclohexane hexol.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A.B,C,D,E or F); adsorption of various compounds (continued)$^{x)}$

	γ/mN m^{-1}				
-E/V	sat lactose C	0.1 mol l^{-1} maltose C	0.1 mol l^{-1} raffinose C	0.1 mol l^{-1} phenol E	sat isobutyl phenol D
0.1	323.7	325.6	325.7	318.2	324.2
0.2	348.3	351.0	351.7	340.1	350.0
0.3	367.9	-	372.1	356.2	367.7
0.4	383.0	386.5	387.2	367.9	372.2
0.5	393.6	-	396.9	376.5	376.3
0.6	399.8	402.8	402.2	382.3	380.6
0.7	403.4	-	404.7	386.2	381.8
0.8	404.5	407.3	405.2	388.2	386.7
0.9	404.0	-	404.3	389.7	387.8
1.0	401.6	406.0	402.0	390.0	389.4
1.1	397.6	-	398.3	389.5	392.0
1.2	392.4	398.7	393.2	388.3	394.0
1.3	385.1	-	386.8	386.0	395.4
1.4	376.3	385.1	379.0	382.2	390.9
1.5	366.0	-	369.5	376.1	380.6
1.6	353.9	362.5	358.6	365.7	367.6
1.7	340.3	-	345.8	351.5	352.5
1.8	324.8	332.4	330.7	335.0	335.9
1.9	307.6	-	313.8	316.6	317.8
2.0	288.5	394.4	294.2	296.3	-
2.1	267.4				
2.2	244.2				
2.3	218.8				
2.4	191.4				

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)x)

	γ/mN m^{-1}			
-E/V	sat naphthol β C	0.1 mol l^{-1} pyro-catechine C	0.1 mol l^{-1} resorcinol C	0.1 mol l^{-1} hydro-quinone C
0.1	-	309.1	305.7	312.0
0.2	334.2	330.6	326.4	331.5
0.3	-	348.0	343.8	348.0
0.4	360.3	362.5	358.4	362.1
0.5	-	374.3	366.0	373.7
0.6	372.4	383.2	379.8	383.2
0.7	-	389.6	387.1	390.2
0.8	379.0	393.8	392.2	395.3
0.9	-	396.3	395.4	398.5
1.0	383.7	397.5	397.3	399.9
1.1	-	397.3	397.5	399.9
1.2	385.7	395.4	395.9	398.0
1.3	-	391.6	392.3	394.0
1.4	383.7	385.9	386.4	387.5
1.5	-	377.3	377.7	378.2
1.6	365.8	365.5	365.8	365.9
1.7	-	351.0	351.8	351.8
1.8	334.1	334.4	334.4	334.3
1.9	-	315.8	315.8	315.7
2.00	294.8	295.4	295.5	295.4

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x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)x)

-E/V	sat dioxynphtaline D	1 mol l^{-1} pyrogallol C	sat coniferyl alcohol E	sat saligenin+) C	sat paraoxybenzoic aldehyde C	sat protocatechic aldehyde E
	γ/mN m^{-1}					
0.1	-	283.8	316.3	309.9	325.0	-
0.2	-	303.6	335.8	329.6	346.1	328.6
0.3	303.5	322.5	-	344.3	359.1	343.5
0.4	319.7	338.7	359.3	354.8	368.4	355.1
0.5	337.3	352.5	-	362.4	374.7	363.5
0.6	346.9	363.5	370.3	367.6	379.8	370.3
0.7	355.4	372.3	372.1	370.8	383.3	374.7
0.8	360.7	378.7	374.7	372.4	385.9	378.0
0.9	365.7	383.0	375.8	373.1	387.1	380.0
1.0	367.6	385.6	375.2	373.0	387.6	381.3
1.1	369.2	386.4	373.1	372.0	386.6	380.8
1.2	370.2	385.3	370.4	370.1	384.5	379.4
1.3	369.4	382.3	-	367.4	381.2	376.2
1.4	367.4	377.2	363.1	363.4	376.6	372.7
1.5	364.1	369.9	-	358.8	369.9	362.9
1.6	358.3	360.3	352.4	352.2	361.1	359.0
1.7	349.5	347.9	-	344.0	349.7	349.6
1.8	335.4	332.8	333.4	331.9	334.0	334.8
1.9	318.1	315.2	-	314.8	315.5	316.7
2.0	-	295.2	295.9	294.8	295.3	-

x) See footnote on the page containing the base solutions.

+) Saligenin-β is toluene,α,2-dihydroxyglucoside.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A.B,C,D,E or F); adsorption of various compounds (continued) x)

	γ/mN m^{-1}						
-E/V	0.01 mol l^{-1} vanilline B	sat paraoxy-benzoic acid E	sat tyrosine C	sat proto-catechic acid E	sat vanillic acid E	sat gallic acid C	sat ethyl ether D
0.1	325.1	315.2	-	298.3	314.8	-	323.4
0.2	349.6	337.1	351.7	316.6	-	310.4	350.9
0.3	365.3	352.3	-	337.4	349.7	327.5	373.2
0.4	374.8	-	386.2	350.9	-	342.4	391.7
0.5	381.3	373.7	-	362.5	369.2	355.0	398.2
0.6	386.0		403.8	371.4	-	365.6	396.8
0.7	388.9	387.7	-	379.3	381.6	374.9	395.0
0.8	391.6	-	414.2	385.4	-	382.2	392.4
0.9	392.6	396.1	-	389.7	388.6	387.4	387,0
1.0	393.0	-	415.7	393.0	-	390.9	-
1.1	392.6	397.5	-	393.9	389.2	392.6	378.6
1.2	390.0	-	408.7	393.2	-	392.2	-
1.3	386.7	393.2	-	390.3	385.8	389.4	368.7
1.4	382.8	387.4	391.8	385.3	-	384.4	-
1.5	375.2	378.7	-	377.2	374.8	376.6	354.9
1.6	364.3	-	362.2	366.0	365.1	365.3	-
1.7	351.1	352.0	-	347.5	-	350.8	340.7
1.8	334.7	-	334.3	335.6	335.3	334.2	-
1.9	316.0	316.4	-	316.8	-	315.7	316.7
2.0	295.7	-	295.4	297.2	-	295.3	-

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)$^{x)}$

γ/mN m^{-1}

-E/V	sat anisol	sat phenetol	sat ethyl formiate	sat ethyl acetate	1 mol l^{-1} glyceric monochlorhydrine	sat piperonal
	D	E	C	C	C	A
0.1	325.0	-	323.2	325.9	-	326.1
0.2	-	351.6	350.3	353.0	349.4	352.8
0.3	373.2	-	372.1	375.4	368.1	372.4
0.4	388.3	390.0	387.5	393.2	379.2	376.9
0.5	-	-	390.3	394.7	383.8	381.0
0.6	392.0	406.8	390.7	393.9	386.4	380.6
0.7	-	-	390.3	392.9	388.2	-
0.8	395.2	408.4	389.3	391.2	388.7	381.9
0.9	-	-	387.7	389.1	387.7	-
1.0	395.9	414.0	385.3	386.1	386.6	380.2
1.1	-	-	382.2	382.3	384.1	-
1.2	392.4	405.7	378.1	378.1	380.4	374.9
1.3	-	-	373.4	373.4	375.5	-
1.4	388.3	390.9	367.9	368.2	370.5	366.0
1.5	380.0	-	361.5	362.3	363.3	360.8
1.6	367.3	367.2	354.1	355.9	354.5	354.2
1.7	-	-	345.1	348.1	343.0	348.7
1.8	336.2	335.6	333.0	333.5	329.2	333.4
1.9	-	-	-	315.4	312.6	315.5
2.0	-	-	-	295.2	293.3	295.4

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)$^{x)}$

	γ /mN m^{-1}						
-E/V	aceto-nitrile D	sat propio-nitrile D	sat capro-nitrile D	1 mol l^{-1} acetic aldehyde F	1 mol l^{-1} propionic aldehyde C	sat isobutylic aldehyde C	sat valeric aldehyde C
0.1	321.7	327.3	324.7	324.2	326.0	326.4	326.7
0.2	-	354.2	-	351.4	353.1	353.1	353.1
0.3	372.5	375.9	374.2	373.3	375.1	375.2	374.9
0.4	390.4	393.2	392.9	391.8	392.4	391.3	392.4
0.5	401.5	394.6	397.4	405.2	396.3	392.8	393.0
0.6	403.9	395.4	397.9	414.2	396.9	393.1	397.1
0.7	-	-	397.4	417.3	396.2	392.7	396.8
0.8	403.8	393.7	395.9	417.0	393.8	391.5	396.1
0.9	-	-	394.9	414.4	391.1	389.1	393.4
1.0	398.5	387.5	390.6	410.0	387.8	386.0	390.3
1.1	-	-	388.8	404.5	383.7	383.1	387.0
1.2	388.1	377.2	384.3	398.1	379.3	378.4	382.2
1.3	-	-	381.3	391.2	374.0	373.8	376.9
1.4	373.6	363.1	377.9	363.6	368.2	368.5	371.8
1.5	-	-	372.6	374.0	361.1	361.2	365.6
1.6	354.8	345.9	364.3	363.5	360.3	353.6	358.9
1.7	-	-	-	348.8	338.7	339.2	347.9
1.8	330.8	326.3	335.9	332.8	323.9	323.2	331.3
1.9	-	314.1	-	316.1	311.2	311.9	314.4
2.0	295.4	-	-	-	-	-	294.2

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x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na$_2$SO$_4$ (A,B,C,D,E or F); adsorption of various compounds (continued)x)

	γ/mN m^{-1}						
-E/V	sat heptanol B	sat furfural C	sat paral-dehyde C	1 mol l^{-1} glyoxal C	sat acetal C	1 mol l^{-1} acetone B	sat butyrone+) C
0.1	-	326.2	326.1	324.9	325.0	324.7	327.2
0.2	353.5	352.2	353.1	350.9	352.2	351.8	353.6
0.3	-	370.2	375.5	371.9	374.5	374.5	375.7
0.4	393.6	375.2	392.3	388.3	392.0	392.5	393.5
0.5	397.4	377.6	391.9	400.2	401.3	405.3	392.0
0.6	398.3	379.3	389.8	408.9	401.9	411.0	395.8
0.7	-	380.1	387.2	414.3	401.1	412.3	394.8
0.8	397.4	379.5	384.1	417.4	398.8	411.3	392.9
0.9	-	378.1	380.1	417.8	395.8	408.8	390.3
1.0	395.8	375.3	375.7	416.3	392.1	404.6	386.6
1.1	-	371.7	370.0	412.7	387.4	399.4	383.2
1.2	390.5	366.9	363.9	407.4	382.1	393.2	379.3
1.3	-	361.0	357.5	399.9	376.4	386.2	374.3
1.4	384.6	353.3	-	390.5	370.2	378.1	367.9
1.5	-	344.6	342.5	379.2	363.0	369.0	361.8
1.6	366.3	335.2	-	365.9	355.5	358.9	355.3
1.7	-	328.8	325.4	-	345.6	347.0	348.1
1.8	334.7	-	-	-	332.7	332.4	334.1
1.9	-	-	304.5	-	315.4	315.1	315.6
2.0	-	-	291.4	-	295.6	295.5	-

x) See footnote on the page containing the base solutions.

+) Butyrone is 4-heptanone.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued) x)

$\gamma/mN\ m^{-1}$

-E/V	sat acetyl-acetone	sat diethyl-acetone	sat pinacolon +)	1 mol l^{-1} methyl-ethyl-acetone	sat ethyl bromide	sat hydrated butyl-chloral *)
	C	D	D	C	E	C
0.1	326.6	-	324.6	325.3	-	326.4
0.2	352.5	-	-	352.4	-	351.7
0.3	374.8	374.0	374.1	375.9	370.0	368.0
0.4	384.2	388.9	392.5	392.3	382.3	374.7
0.5	386.5	389.4	394.4	394.8	390.6	379.4
0.6	387.2	388.6	393.8	394.9	395.6	383.0
0.7	386.9	387.1	392.5	394.1	-	386.3
0.8	386.0	384.3	390.5	391.5	400.2	388.1
0.9	384.1	381.3	387.7	388.4	-	390.2
1.0	381.4	377.2	384.1	383.2	401.9	390.2
1.1	377.8	372.2	380.0	377.8	-	389.1
1.2	372.9	366.7	375.2	371.9	398.8	388.6
1.3	366.3	360.3	370.2	365.1	-	387.9
1.4	359.5	353.4	364.8	358.6	389.9	383.6
1.5	351.6	346.1	359.1	350.0	-	377.7
1.6	342.9	337.9	352.4	341.2	366.3	366.6
1.7	332.9	329.1	344.5	332.1	-	351.8
1.8	322.1	320.2	335.6	322.2	335.0	335.0
1.9	309.8	310.9	317.8	311.4	-	-
2.0	-	296.6	-	296.7	-	-

x) See footnote on the page containing the base solutions.

+) Pinacolone is 3,3-dimethyl-2-butanone.

*) Butylchloral is 2,2,3-trichlorobutanol.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued) x)

	γ/mN m^{-1}					
-E/V	sat chloroform	acetic acid	1 mol l^{-1} monochlor- acetic acid	1 mol l^{-1} propionic acid	sat normal butyric acid	sat isovaleri- anic acid
	C	C	B	C	C	C
0.1	-	326.7	325.0	-	326.1	-
0.2	-	353.2	349.3	353.1	352.2	-
0.3	-	374.6	367.3	374.6	369.9	-
0.4	393.5	391.0	380.6	390.1	378.4	381.2
0.5	-	402.5	388.6	399.0	382.2	385.1
0.6	404.3	409.7	390.9	402.9	384.1	386.3
0.7	-	413.6	399.2	404.3	384.4	386.3
0.8	406.1	415.2	401.7	404.0	384.0	386.2
0.9	-	414.5	402.8	402.8	383.3	385.3
1.0	406.7	413.0	402.4	400.9	392.0	384.1
1.1	-	409.6	400.9	398.1	380.0	382.3
1.2	404.1	404.7	397.8	394.3	377.3	379.9
1.3	-	397.9	393.2	389.8	374.0	377.0
1.4	390.9	389.1	386.7	383.8	370.0	372.6
1.5	-	378.4	378.1	375.5	365.1	369.1
1.6	366.1	361.5	-	364.4	358.7	363.4
1.7	-	351.0	-	350.6	349.0	350.7
1.8	334.2	334.7	-	334.4	338.9	334.3
1.9	-	316.4	-	316.4	316.4	315.9
2.0	-	-	-	-	-	295.7

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued) x)

	γ/mN m^{-1}					
-E/V	sat normal caproic acid E	sat crotonic acid E	1 mol l^{-1} lactic acid C	sat α oxybutyric acid D	1 mol l^{-1} glyceric acid C	0.1 mol l^{-1} oxalic acid C
0.1	-	324.6	325.5	324.3	325.7	325.6
0.2	351.8	351.3	351.9	350.4	352.0	352.1
0.3	372.5	371.7	372.8	368.2	373.3	373.4
0.4	379.6	383.9	388.7	378.4	389.6	390.4
0.5	384.4	390.2	394.7	383.2	401.5	404.0
0.6	385.8	393.3	397.2	384.6	408.9	414.1
0.7	-	-	397.6	-	412.7	421.4
0.8	385.3	394.3	396.5	383.6	413.9	425.3
0.9	-	-	394.4	-	413.1	426.0
1.0	383.2	391.5	391.8	379.3	410.8	423.6
1.1	-	-	388.3	-	406.9	418.7
1.2	380.1	385.7	384.1	374.0	401.4	411.9
1.3	-	-	379.3	-	394.4	403.3
1.4	374.3	377.6	373.7	363.4	385.9	392.8
1.5	370.4	371.6	367.3	-	375.7	380.7
1.6	365.0	363.7	359.4	351.7	363.6	367.1
1.7	-	-	349.4	-	349.7	352.2
1.8	335.6	335.2	334.6	334.8	333.6	-
1.9	-	-	317.2	-	-	-
2.0	-	-	-	-	-	-

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued) x)

	γ/mN m^{-1}						
-E/V	1 mol l^{-1} malonic acid	pyro-tartaric acid	sat suberic acid +)	sat fumaric acid	1 mol l^{-1} tartaric acid	0.1 mol l^{-1} citric acid	1 mol l^{-1} taurine
	C	D	C	D	C	C	C
0.1	325.6	323.7	-	-	322.1	325.4	-
0.2	351.2	-	-	351.0	348.5	352.5	-
0.3	371.7	370.3	-	371.6	369.0	374.4	355.3
0.4	387.4	384.3	386.8	386.3	384.4	391.3	380.8
0.5	398.6	393.0	-	395.8	395.7	403.7	397.1
0.6	406.3	397.9	401.4	402.5	404.1	411.9	409.2
0.7	411.7	400.5	-	408.6	410.3	416.9	417.4
0.8	414.7	401.3	407.6	414.9	414.1	419.3	422.2
0.9	415.7	400.8	-	417.3	415.9	419.6	423.2
1.0	414.7	399.0	405.0	417.2	415.4	417.9	421.2
1.1	411.7	396.3	-	415.2	412.6	414.2	416.7
1.2	406.8	-	399.2	411.3	407.4	408.6	409.9
1.3	399.9	387.5	-	403.5	399.9	401.0	401.3
1.4	390.9	-	388.3	-	390.7	391.3	390.7
1.5	379.9	373.8	-	-	379.5	380.0	378.5
1.6	367.0	363.1	366.2	-	366.5	366.5	364.6
1.7	352.4	350.6	-	-	351.8	351.3	349.0
1.8	336.1	335.4	334.4	-	335.5	335.7	331.8
1.9	-	-	-	-	-	315.8	312.9
2.0	-	-	295.4	-	-	-	291.5

x) See footnote on the page containing the base solutions.

+) Suberic acid is octanedioic acid.

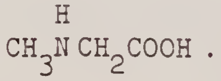
Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued)[x)]

γ/mN m^{-1}

-E/V	1 mol l^{-1} ethylene glycol A	1 mol l^{-1} sarcosine +) C	1 mol l^{-1} alanine C	sat leucine C	sat asparagine A	0.01 mol l^{-1} benzoic acid A
0.1	317.5	321.0	320.4	-	320.0	325.3
0.2	348.9	350.3	349.7	349.8	350.9	351.0
0.3	372.7	373.4	372.9	-	373.7	369.9
0.4	391.1	391.7	391.0	385.2	391.8	382.3
0.5	405.6	405.7	404.9	-	405.6	391.1
0.6	416.0	415.2	424.5	403.8	415.3	397.5
0.7	422.7	421.0	421.0	-	421.5	401.5
0.8	425.8	423.6	423.9	411.7	424.4	404.4
0.9	426.1	423.2	423.9	-	424.7	404.5
1.0	423.2	420.3	421.3	411.7	422.3	404.0
1.1	418.0	415.0	416.4	-	417.6	402.3
1.2	410.8	407.9	409.7	403.8	410.7	399.0
1.3	392.7	399.0	401.1	-	401.9	395.3
1.4	391.4	388.6	390.4	388.9	391.4	389.0
1.5	379.1	376.3	378.6	-	379.2	378.4
1.6	365.3	362.7	366.2	365.1	365.3	366.3
1.7	349.7	347.5	349.3	350.2	349.7	351.4
1.8	332.4	330.5	332.1	333.4	332.4	334.4
1.9	313.2	311.7	313.3	-	309.2	315.9
2.0	285.3	291.3	292.6	294.8	293.0	295.5

x) See footnote on the page containing the base solutions.

+) Sarcosine is:



$CH_3\overset{H}{N}CH_2COOH$.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued) x)

	$\gamma/mN\ m^{-1}$					
-E/V	sat para-toluic acid C	0.01 mol l^{-1} ortho-phthalic acid B	sat. 2-furan carbonic acid C	valero-acetone E	1 mol l^{-1} formamide C	1 mol l^{-1} acet-amide F
0.1	-	322.5	323.5	323.1	326.8	323.9
0.2	353.3	350.0	345.5	350.9	352.6	350.8
0.3	-	368.2	360.2	373.1	-	372.6
0.4	388.3	381.0	371.7	382.3	392.1	389.8
0.5	-	391.0	380.2	384.5	-	402.4
0.6	403.8	398.5	386.6	385.0	414.7	410.6
0.7	-	404.4	391.3	384.3	-	415.3
0.8	410.1	408.4	393.7	382.5	422.7	417.3
0.9	-	410.5	394.6	379.4	-	417.1
1.0	413.1	411.6	394.0	375.0	420.1	415.0
1.1	-	410.4	392.1	369.7	-	410.9
1.2	409.5	407.2	389.4	363.3	408.9	405.1
1.3	-	401.0	384.0	356.6	-	-
1.4	393.0	391.8	379.2	349.0	390.3	397.9
1.5	-	380.1	372.7	340.6	-	388.1
1.6	363.5	366.5	363.3	331.2	365.0	377.0
1.7	-	351.2	350.5	321.3	-	364.3
1.8	335.9	334.4	334.7	310.7	333.1	334.8
1.9	-	315.4	-	299.2	-	315.7
2.0	296.3	294.7	-	287.1	294.1	295.7

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued) x)

	γ/mN m^{-1}					
-E/V	sat propioamide C	sat normal butyramide F	sat acetanilide C	1 mol l^{-1} urea C	1 mol l^{-1} urethane E	methyl urea D
0.1	325.0	323.5	-	318.4	324.2	320.2
0.2	351.5	350.2	349.8	344.2	351.1	343.7
0.3	371,8	370.7	-	365.1	372.9	362.9
0.4	385.7	381.9	378.3	382.5	387.7	378.1
0.5	393.2	387.9	-	396.6	394.7	390.0
0.6	396.7	391.0	387.6	407.2	397.3	398.7
0.7	397.8	392.3	-	414.7	398.2	404.5
0.8	397.4	392.1	390.3	418.7	398.1	401.5
0.9	395.7	390.0	-	419.7	398.1	407.9
1.0	392.6	386.9	389.2	418.1	394.5	406.4
1.1	388.9	382.6	-	414.2	391.2	402.9
1.2	382.5	377.5	383.7	407.7	386.8	397.4
1.3	375.7	371.3	-	399.9	381.6	390.0
1.4	367.4	363.9	373.9	389.8	375.5	380.5
1.5	357.7	355.5	-	378.0	368.5	369.4
1.6	347.0	346.1	360.2	364.6	360.0	356.8
1.7	335.0	335.6	-	349.3	349.0	341.9
1.8	321.8	323.9	334.1	332.5	334.1	325.7
1.9	307.0	310.9	-	314.0	316.6	308.1
2.0	290.9	294.3	295.5	293.7	296.4	289.7

x) See footnote on the page containing the base solutions.

Electrocapillary data for mercury in 0.5 mol l^{-1} Na_2SO_4 (A,B,C,D,E or F); adsorption of various compounds (continued) x)

	γ /mN m^{-1}				
-E/V	1 mol l^{-1} parabanic acid E	sat hypoxanthine E	sat biuret C	0.1 mol l^{-1} amygdaline F	0.1 mol l^{-1} salicin +) C
0.1	317.3	-	-	324.2	322.6
0.2	344.3	-	340.4	347.7	345.2
0.3	-	369.1	358.3	363.4	362.1
0.4	379.4	382.7	371.5	374.1	373.9
0.5	-	391.1	379.4	381.2	381.2
0.6	402.1	397.9	385.7	385.3	385.1
0.7	-	401.3	390.8	386.6	387.4
0.8	412.6	404.0	391.2	386.7	388.7
0.9	415.5	405.9	395.5	384.6	389.0
1.0	415.7	405.3	399.3	381.7	388.4
1.1	414.1	403.8	401.5	378.1	386.2
1.2	410.5	400.3	400.6	373.6	382.8
1.3	402.6	395.8	395.7	368.0	378.3
1.4	391.8	388.3	388.3	361.4	371.7
1.5	381.3	378.4	377.9	353.9	364.7
1.6	-	366.0	365.1	345.7	355.4
1.7	-	351.4	350.4	334.8	344.2
1.8	-	334.8	333.8	322.7	330.6
1.9	-	316.3	315.1	308.7	314.1
2.0	-	296.3	294.8	292.2	294.6

x) See footnote on the page containing the base solutions.

+) Salicin is toluene,α,2-dhydroxy glucoside.

Electrocapillary data on mercury electrode in various salts at 18°C. Given are the interfacial tension in mN m^{-1}, potential V vs KCl calomel electrode.

Reference: G. Gouy. Ann. Chim. Phys. 8 (8) (1906) 291.

-E/V	γ/mN m^{-1} 0.5 mol l^{-1} Na_2SO_4	 0.5 mol l^{-1} Na_2SO_4 + 0.1 mol l^{-1} tertiary amyl alcohol	 1 mol l^{-1} KNO_3	 1 mol l^{-1} KNO_3 + 0.1 mol l^{-1} tertiary amyl alcohol	 1 mol l^{-1} KNO_3 + 0.1 mol l^{-1} phenol
0.1	326.7	326.4	-	-	-
0.2	353.6	353.1	-	-	-
0.3	375.8	375.3	367.3	367.1	350.9
0.4	394.0	392.6	382.9	382.5	362.9
0.5	408.2	395.4	395.6	395.2	372.6
0.6	418.2	395.7	406.0	405.2	380.1
0.7	424.4	395.6	414.1	406.3	385.8
0.8	427.1	395.0	419.6	405.1	389.4
0.9	426.9	393.8	422.2	403.1	391.5
1.0	424.0	391.9	421.6	400.8	392.5
1.1	418.9	389.8	417.8	398.1	392.5
1.2	412.1	387.3	411.4	394.7	390.9
1.3	403.1	384.0	402.9	391.1	388.2
1.4	392.6	380.1	392.4	386.4	384.0
1.5	380.6	374.9	380.3	378.6	376.8
1.6	366.8	364.8	366.4	365.8	364.9
1.7	351.6	350.7	351.0	350.8	350.7
1.8	334.6	334.2	334.0	333.7	333.7
1.9	316.0	308.2	315.6	315.3	314.1
2.0	295.7	295.2	-	-	-

Electrocapillary data on mercury in various salts (continued)

	γ /mN m^{-1}			
-E/V	1 mol l^{-1} KNO$_3$ + 0.1 mol l^{-1} caffeine	0.33 mol l^{-1} (NH$_4$)$_2$HPO$_4$	0.33 mol l^{-1} (NH$_4$)$_2$HPO$_4$ 0.1 mol l^{-1} tertiary amyl alcohol	0.33 mol l^{-1} (NH$_4$)$_2$HPO$_4$ 0.1 mol l^{-1} phenol
0.1	-	-	-	-
0.2	346.8	-	-	-
0.3	364.2	371.9	371.4	356.6
0.4	374.6	393.2	392.7	370.0
0.5	380.6	408.7	407.4	379.0
0.6	383.8	418.7	407.4	385.2
0.7	385.2	424.2	406.0	389.1
0.8	385.0	426.9	405.2	391.5
0.9	383.7	426.5	403.2	393.0
1.0	381.9	424.1	401.1	393.4
1.1	378.8	419.2	398.3	392.9
1.2	375.1	412.4	394.9	391.3
1.3	370.4	403.6	391.3	388.7
1.4	364.4	393.2	386.7	384.5
1.5	357.5	381.1	379.1	377.4
1.6	349.0	363.1	366.8	366.2
1.7	338.8	352.2	351.9	352.2
1.8	326.0	334.6	334.8	334.9
1.9	310.8	315.8	316.0	316.1
2.0	293.7	295.0	295.0	295.1

Electrocapillary data on mercury in various salts (continued)

	γ /mN m^{-1}				
-E/V	0.1 mol l^{-1} KI	0.1 mol l^{-1} KI + 0.1 mol l^{-1} phenol	1 mol l^{-1} KI + 0.1 mol l^{-1} aniline	0.5 mol l^{-1} H_2SO_4	0.5 mol l^{-1} H_2SO_4 + 0.1 mol l^{-1} resorcinol
0.1	-	-	-	320.0	300.3
0.2	-	-	-	346.1	321.8
0.3	-	-	-	367.4	340.1
0.4	-	-	-	385.1	-
0.5	-	-	-	399.3	367.6
0.6	-	-	-	410.7	-
0.7	-	-	-	419.1	385.4
0.8	336.2	334.5	325.7	424.5	-
0.9	370.4	364.0	353.1	426.4	395.3
1.0	388.7	378.5	368.3	425.2	-
1.1	397.9	385.9	376.9	412.2	398.5
1.2	400.7	388.7	379.7	415.4	-
1.3	397.9	387.7	378.0	407.6	395.1
1.4	390.4	383.9	374.2	398.2	-
1.5	379.3	376.6	369.1	387.2	383.2
1.6	365.9	364.9	361.5	374.6	-
1.7	350.4	349.9	348.7	360.8	360.3
1.8	333.1	332.9	332.4	346.1	-
1.9	313.9	313.8	313.5	-	-
2.0	292.9	292.8	292.5	-	-

Electrocapillary data on mercury electrode in various salts, bases and acids at 18°C. Given are the interfacial tension in mN m^{-1}, potential in V vs KCl calomel electrode. x)

Reference: G. Gouy. Ann. Chim. Phys. 8 (9) (1906) 75.

-E/V	γ/mN m^{-1} 0.05 mol l^{-1} Na_2SO_4	0.5 mol l^{-1} H_2SO_4	1 mol l^{-1} H_3PO_4	1 mol l^{-1} HBr
0	-	283.4	298.4	-
0.1	337.5	319.6	328.9	-
0.2	361.8	345.7	351.9	-
0.3	381.9	367.0	371.1	-
0.4	398.5	384.7	387.3	-
0.5	411.0	399.1	401.1	-
0.6	419.7	410.5	412.0	353.4
0.7	424.8	419.2	419.8	381.5
0.8	426.8	424.3	424.6	397.8
0.9	426.5	426.1	426.0	408.6
1.0	424.1	425.1	424.7	415.1
1.1	419.8	421.5	421.1	416.3
1.2	413.3	415.5	415.6	413.6
1.3	405.1	407.8	407.9	407.1
1.4	395.3	398.5	398.8	398.0
1.5	383.9	387.5	387.9	387.1
1.6	366.6	375.0	375.8	374.8
1.7	356.4	361.4	362.2	360.7
1.8	340.2	347.1	347.8	-
1.9	322.3	-	-	-
2.0	303.0	-	-	-

x) These solutions are base solutions for the data on p. 348-362.

Electrocapillary data on mercury in various salts. bases and acids x) (continued)

	γ/mN m^{-1}				
-E/V	0.1 mol l^{-1} tetramethylammonium hydrate	0.05 mol l^{-1} H_2SO_4 + 0.1 mol l^{-1} tetramethylammonium hydrate	0.5 mol l^{-1} H_2SO_4 + 0.1 mol l^{-1} tetramethylammonium hydrate	0.1 mol l^{-1} tetraethylammonium hydrate	0.1 mol l^{-1} tetraethylammonium hydrate + 0.05 mol l^{-1} H_2SO_4
0	-	-	284.9	-	283.7
0.1	-	337.4	320.7	-	319.9
0.2	-	362.1	346.8	-	346.3
0.3	-	381.9	368.1	-	367.7
0.4	-	398.0	385.7	-	385.1
0.5	410.4	410.7	400.1	410.0	399.5
0.6	419.4	419.4	411.4	419.3	410.5
0.7	424.4	424.5	419.7	422.8	418.3
0.8	426.0	426.0	424.7	421.9	421.2
0.9	424.7	425.2	426.3	417.7	419.4
1.0	420.8	421.4	424.3	411.1	414.7
1.1	414.5	415.3	419.7	402.8	407.3
1.2	406.3	407.1	412.6	393.2	398.3
1.3	396.1	397.2	403.7	382.5	388.3
1.4	384.7	385.7	393.4	370.5	372.2
1.5	372.0	373.2	381.5	358.0	359.7
1.6	357.8	359.1	368.2	344.0	346.7
1.7	342.3	343.8	353.8	329.2	331.7
1.8	325.8	327.3	338.8	314.3	316.8
1.9	307.9	305.3	-	298.3	300.7
2.0	289.0	290.7	-	281.2	283.8

x) For base solutions, see p. 347.

Electrocapillary data on mercury in various salts, bases and acids x) (continued)

	γ/mN m^{-1}				
-E/V	0.1 mol l^{-1} tetraethylammonium hydrate + 0.5 mol l^{-1} H_2SO_4	0.1 mol l^{-1} tetraethylammonium hydrate + 0.05 mol l^{-1} H_3PO_4	0.1 mol l^{-1} tetraethylammonium hydrate + 1 mol l^{-1} H_3PO_4	0.1 mol l^{-1} tetraethylammonium hydrate + 1 mol l^{-1} HBr	0.5 mol l^{-1} tetraethylammonium hydrate
0	283.7	-	-	-	-
0.1	338.6	-	331.6	-	-
0.2	346.3	-	352.3	-	-
0.3	367.7	-	371.7	-	-
0.4	385.1	-	388.2	-	-
0.5	399.5	409.9	401.7	-	407.1
0.6	410.5	419.1	412.6	339.3	416.7
0.7	418.3	422.9	419.8	376.1	418.2
0.8	421.3	422.4	422.2	391.5	414.2
0.9	419.7	418.0	420.1	400.5	405.0
1.0	414.7	411.2	414.5	404.1	401.3
1.1	407.3	404.1	406.9	402.3	392.3
1.2	398.3	394.3	397.6	395.9	381.9
1.3	388.3	384.0	387.2	386.8	370.5
1.4	377.2	372.1	375.8	376.1	354.0
1.5	365.4	359.6	363.5	364.7	345.3
1.6	352.7	346.2	350.6	352.3	331.7
1.7	339.0	331.8	337.1	339.6	317.3
1.8	325.0	316.5	-	-	302.1
1.9	-	298.9	-	-	285.8
2.0	-	283.3	-	-	268.9

x) For base solutions, see p. 347.

Electrocapillary data on mercury in various salts, bases and acids x) (continued)

	γ/mN m^{-1}					
-E/V	0.5 mol l^{-1} tetraethylammonium hydrate + 0.25 mol l^{-1} H_2SO_4	0.1 mol l^{-1} choline	0.1 mol l^{-1} choline + 0.05 mol l^{-1} H_2SO_4	0.1 mol l^{-1} neurine	0.1 mol l^{-1} neurine + 0.05 mol l^{-1} H_2SO_4	0.1 mol l^{-1} neurine + 1 mol l^{-1} H_3PO_4
0	-	-	-	-	-	-
0.1	-	-	332.6	-	337.5	328.9
0.2	-	-	360.2	-	361.8	351.9
0.3	-	-	381.1	-	382.0	371.5
0.4	-	-	397.2	-	398.0	388.0
0.5	407.9	407.7	409.4	409.5	410.7	401.7
0.6	416.5	417.8	418.1	419.4	419.2	411.8
0.7	418.6	422.5	423.5	424.2	424.4	420.3
0.8	416.2	427.8	424.9	424.4	426.0	424.4
0.9	410.4	421.9	422.9	423.5	424.4	424.8
1.0	402.5	417.5	417.8	419.7	420.3	422.1
1.1	393.5	410.7	412.2	413.1	413.6	416.5
1.2	383.2	402.5	403.3	404.6	405.4	408.9
1.3	372.1	392.7	388.9	394.6	395.5	399.5
1.4	360.1	381.3	382.0	383.2	384.0	388.6
1.5	347.6	368.7	368.8	370.5	371.4	376.6
1.6	329.8	354.6	354.8	356.4	357.4	363.0
1.7	319.8	339.2	339.0	341.2	342.2	348.6
1.8	304.9	322.3	322.5	324.6	325.9	-
1.9	289.0	304.5	304.2	307.1	308.3	-
2.0	272.7	285.6	285.2	288.3	289.7	-

x) For base solutions, see p. 347.

Electrocapillary data for mercury in various salts, bases and acids (continued)

x)

	γ/mN m^{-1}				
-E/V	0.1 mol l^{-1} tetramethyl-arsonium hydrate	0.1 mol l^{-1} tetramethyl-arsonium hydrate + 0.05 mol l^{-1} H_2SO_4	0.1 mol l^{-1} tetramethyl-arsonium hydrate + 1 mol l^{-1} H_3PO_4	0.1 mol l^{-1} trimethyl-sulfine hydrate	0.1 mol l^{-1} trimethyl sulfine hydrate + 0.05 mol l^{-1} H_2SO_4
0	-	-	-	-	-
0.1	-	-	328.9	-	-
0.2	-	359.3	351.4	-	357.7
0.3	-	380.2	370.9	-	380.4
0.4	-	396.5	387.3	-	397.0
0.5	-	409.3	401.2	409.3	409.9
0.6	419.2	418.6	412.4	419.3	418.9
0.7	424.1	423.7	420.3	424.4	424.4
0.8	426.3	426.0	424.5	426.3	426.3
0.9	424.4	425.4	425.5	424.8	425.2
1.0	-	421.3	423.1	420.5	421.3
1.1	415.1	416.0	418.6	413.9	414.8
1.2	-	408.2	411.3	405.3	406.4
1.3	392.9	398.6	402.4	394.9	396.4
1.4	-	388.1	391.9	383.2	384.5
1.5	374.0	375.4	380.2	370.0	371.5
1.6	-	361.9	367.1	355.4	356.9
1.7	345.4	347.3	353.6	339.5	341.2
1.8	-	331.6	-	322.1	324.0
1.9	313.6	314.5	-	303.6	305.4
2.0	-	296.3	-	283.8	285.8

x) For base solutions, see p. 347.

Electrocapillary data for mercury in various salts, bases and acids x) (continued)

	γ/mN m^{-1}					
-E/V	0.1 mol l^{-1} trimethyl-sulfine hydrate + 0.5 mol l^{-1} H_2SO_4	0.1 mol l^{-1} trimethyl-sulfine hydrate + 1 mol l^{-1} HBr	0.5 mol l^{-1} trimethyl-sulfine hydrate	0.5 mol l^{-1} trimethyl-sulfine + 0.25 mol l^{-1} H_2SO_4	0.1 mol l^{-1} trimethyl-sulfine	0.1 mol l^{-1} trimethyl-sulfine + 0.05 mol l^{-1} H_2SO_4
0.1	-	-	-	-	-	-
0.2	347.2	-	-	-	-	354.3
0.3	367.6	-	-	372.9	-	380.3
0.4	385.3	-	-	391.5	-	396.6
0.5	395.9	-	405.8	405.9	-	407.8
0.6	411.9	343.7	417.0	416.2	413.3	413.5
0.7	419.5	377.9	422.4	422.3	418.4	418.3
0.8	424.9	394.4	423.5	424.4	419.7	420.1
0.9	426.3	405.1	422.0	422.7	417.2	418.0
1.0	423.4	411.0	415.6	417.8	411.6	412.5
1.1	419.5	411.9	408.4	410.7	404.0	405.2
1.2	412.4	408.2	398.0	402.0	395.0	395.8
1.3	403.3	401.3	388.2	391.3	384.3	385.6
1.4	392.4	391.7	375.8	379.2	372.6	374.0
1.5	380.2	380.2	362.1	365.9	359.5	360.5
1.6	366.7	367.1	347.0	351.1	345.3	346.9
1.7	351.9	352.8	330.5	335.0	330.1	332.1
1.8	336.5	-	312.7	317.7	314.0	315.7
1.9	-	-	293.7	299.0	296.5	298.8
2.0	-	-	273.1	279.1	278.6	280.4

X) For base solutions, see p. 347.

Electrocapillary data for mercury in various salts, bases and acids x) (continued)

	γ/mN m^{-1}					
-E/V	0.1 mol l^{-1} trimethylsulfine + 1 mol l^{-1} H_3PO_4	1 mol l^{-1} NH_4OH + 0.5 mol l^{-1} Na_2SO_4	1 mol l^{-1} methylamine + 0.5 mol l^{-1} Na_2SO_4	1 mol l^{-1} methylamine + 0.5 mol l^{-1} H_2SO_4 + 0.5 mol l^{-1} Na_2SO_4	1 mol l^{-1} trimethylamine + 0.5 mol l^{-1} H_2SO_4	0.1 mol l^{-1} trimethylamine
		C	D	D	E	
0.1	328.9	-	-	-	327.2	-
0.2	392.9	-	-	347.5	349.3	-
0.3	370.0	-	-	-	375.5	-
0.4	387.0	-	-	389.0	393.4	391.5
0.5	400.5	406.8	-	-	407.3	395.4
0.6	411.2	417.2	415.2	415.2	417.3	393.2
0.7	418.8	422.8	420.2	-	423.1	391.5
0.8	421.2	424.9	421.2	426.1	424.8	389.2
0.9	419.4	424.0	419.3	-	423.2	387.9
1.0	414.8	420.9	415.2	423.8	419.0	383.9
1.1	408.2	415.6	-	-	412.4	379.9
1.2	399.7	408.5	402.6	411.2	404.0	377.0
1.3	390.0	399.9	-	-	394.1	372.2
1.4	378.6	389.4	384.8	391.2	382.9	366.7
1.5	366.3	377.6	-	-	370.3	361.4
1.6	352.9	364.0	361.6	364.7	356.0	354.2
1.7	338.3	349.1	-	-	340.6	345.6
1.8	-	332.5	331.9	331.6	323.5	333.8
1.9	-	314.2	-	-	304.7	319.0
2.0	-	294.3	294.4	-	284.1	303.0

x) For base solutions, see p. 347.

Electrocapillary data for mercury in various salts, bases and acids x) (continued)

γ/mN m^{-1}

-E/V	0.1 mol l^{-1} triethyl-amine + 0.05 mol l^{-1} H_2SO_4	0.1 mol l^{-1} triethyl-amine + 1 mol l^{-1} H_3PO_4	1 mol l^{-1} isobutyl-amine + 0.5 mol l^{-1} Na_2SO_4 C	0.1 mol l^{-1} isobutyl-amine + 0.05 mol l^{-1} H_2SO_4	0.08 mol l^{-1} triisobutyl-amine + 0.04 mol l^{-1} H_2SO_4
0.1	337.9	329.0	-	337.9	336.9
0.2	362.3	351.9	-	362.1	361.2
0.3	382.4	371.7	-	381.9	381.0
0.4	398.4	388.0	-	398.0	396.2
0.5	410.8	401.7	384.3	410.4	406.3
0.6	419.4	412.4	383.9	419.1	407.1
0.7	427.8	419.9	382.7	424.2	404.2
0.8	423.9	423.5	380.2	425.5	400.2
0.9	421.2	422.2	377.1	424.4	396.8
1.0	416.0	418.2	373.8	402.7	390.4
1.1	408.9	412.0	370.0	415.3	383.6
1.2	400.2	404.0	365.2	407.9	376.2
1.3	390.6	394.7	360.0	398.9	368.2
1.4	379.8	384.0	353.8	388.6	359.5
1.5	367.7	372.5	347.0	377.0	350.7
1.6	354.4	359.8	338.7	363.6	340.4
1.7	340.3	346.1	328.7	349.0	329.7
1.8	324.8	-	316.8	332.8	317.8
1.9	319.0	-	302.9	315.2	306.2
2.0	290.8	-	286.8	296.2	294.4
2.1	-	-	268.4	-	-
2.2	-	-	247.2	-	-
2.3	-	-	220.3	-	-
2.4	-	-	191.3	-	-

x) For base solutions, see p. 347.

Electrocapillary data for mercury in various salts, bases and acids x) (continued)

	γ/mN m^{-1}				
-E/V	0.1 mol l^{-1} triisobutyl-amine + 1 mol l^{-1} H_3PO_4	0.1 mol l^{-1} triisobutyl-amine + 1 mol l^{-1} HBr	sat isoamyl-amine + 0.5 mol l^{-1} Na_2SO_4 E	1 mol l^{-1} isoamyl-amine + 0.5 mol l^{-1} H_2SO_4 E	0.1 mol l^{-1} isoamyl-amine
0.1	329.1	-	-	320.9	-
0.2	352.3	-	-	349.0	-
0.3	371.7	-	-	371.5	-
0.4	388.1	-	375.5	387.7	-
0.5	401.5	-	379.0	395.0	397.2
0.6	410.5	337.9	378.6	396.3	399.0
0.7	407.1	371.7	377.1	394.8	398.5
0.8	402.2	384.9	-	392.8	393.3
0.9	396.3	389.5	372.4	390.2	394.6
1.0	390.0	387.2	-	387.1	-
1.1	383.2	381.5	365.2	383.3	388.0
1.2	376.0	375.0	-	378.6	-
1.3	368.2	367.5	354.2	372.6	379.1
1.4	359.8	359.3	-	365.3	-
1.5	350.6	351.0	338.1	356.2	366.7
1.6	341.2	342.0	-	345.3	-
1.7	331.2	332.5	315.0	332.1	349.5
1.8	-	-	301.0	316.9	-
1.9	-	-	284.7	299.2	323.3
2.0	-	-	224.7	278.7	-
2.1	-	-	247.3	-	-
2.2	-	-	226.6	-	-
2.3	-	-	201.8	-	-
2.4	-	-	178.7	-	-

x) For base solutions, see p. 347.

Electrocapillary data for mercury in various salts, bases and acids x) (continued)

-E/V	γ /mN m^{-1} 0.1 mol l^{-1} isoamylamine + 0.05 mol l^{-1} H_2SO_4	0.1 mol l^{-1} isoamylamine + 1 mol l^{-1} H_3PO_4	0.1 mol l^{-1} diisoamylamine + 1 mol l^{-1} H_2SO_4	0.1 mol l^{-1} diisoamylamine + 1 mol l^{-1} H_3PO_4	0.1 mol l^{-1} heptylamine + 0.05 mol l^{-1} H_2SO_4	1 mol l^{-1} allylamine + 0.5 mol l^{-1} Na_2SO_4 E
0.1	338.1	328.9	337.8	328.9	338.0	-
0.2	362.5	351.9	362.0	351.9	362.2	-
0.3	382.4	371.3	381.9	371.3	381.8	-
0.4	398.2	387.7	397.6	387.7	396.8	386.0
0.5	410.4	401.3	405.9	401.1	403.8	399.0
0.6	418.6	412.0	405.1	406.7	405.8	399.2
0.7	422.7	419.4	402.5	404.9	405.3	398.5
0.8	423.2	422.6	398.9	401.7	403.4	396.4
0.9	421.1	422.1	394.7	397.9	400.9	393.2
1.0	416.8	418.8	389.9	393.4	397.3	389.4
1.1	411.1	413.9	384.8	388.3	393.1	384.8
1.2	403.8	407.0	378.5	382.4	387.9	379.2
1.3	395.2	398.5	371.7	375.2	381.6	373.2
1.4	384.9	389.0	363.7	367.4	374.0	366.1
1.5	373.6	377.9	354.3	358.9	364.4	358.9
1.6	360.7	365.6	344.6	349.0	353.2	349.1
1.7	346.4	351.9	332.9	337.4	340.3	339.0
1.8	330.5	-	319.9	-	325.7	327.0
1.9	312.9	-	305.0	-	308.6	312.2
2.0	293.5	-	288.6	-	289.4	293.8
2.1	-	-	-	-	-	272.9
2.2	-	-	-	-	-	248.4
2.3	-	-	-	-	-	222.4
2.4	-	-	-	-	-	-

x) For base solutions, see p. 347.

Electrocapillary data for mercury in various salts, bases and acids x) (continued)

	γ/mN m^{-1}				
-E/V	1 mol l^{-1} allyl-amine + 0.5 mol l^{-1} H_2SO_4 + 0.5 mol l^{-1} Na_2SO_4 E	0.1 mol l^{-1} allyl-amine + 0.5 mol l^{-1} H_2SO_4	0.1 mol l^{-1} allyl-amine + 1 mol l^{-1} H_3PO_4	1 mol l^{-1} ethylene-diamine + 0.5 mol l^{-1} Na_2SO_4 C	1 mol l^{-1} piperazine + 0.5 mol l^{-1} Na_2SO_4 C
0.1	319.9	338.1	329.3	-	-
0.2	347.3	362.5	352.3	-	-
0.3	369.8	382.4	371.7	-	-
0.4	387.9	398.5	388.2	394.7	384.9
0.5	402.3	410.9	401.8	403.6	400,4
0.6	412.4	419.8	412.8	410.8	409.5
0.7	419.1	424.5	420.5	413.6	412.5
0.8	421.2	426.5	424.7	412.5	411.4
0.9	420.4	425.2	425.5	408.6	408.1
1.0	-	422.2	423.5	404.0	403.1
1.1	411.8	416.7	418.8	397.6	396.7
1.2	-	409.5	412.3	390.2	389.1
1.3	395.8	400.8	403.8	381.6	380.6
1.4	-	390.5	393.9	372.3	370.9
1.5	374.0	378.6	382.7	361.5	360.4
1.6	-	365.3	369.8	349.6	348.8
1.7	345.9	350.3	355.8	336.2	336.1
1.8	-	333.8	-	321.5	321.9
1.9	311.3	315.5	-	305.3	306.1
2.0	-	296.0	-	287.2	288.0
2.1	-	-	-	266.9	267.8
2.2	-	-	-	244.0	244.4
2.3	-	-	-	218.0	218.6
2.4	-	-	-	189.5	-

x) For base solutions, see p. 347.

Electrocapillary data on mercury in various salts, bases and acids x) (continued)

-E/V	γ/mN m^{-1} 1 mol l^{-1} guanidine + 0.5 mol l^{-1} Na_2SO_4 E	0.01 mol l^{-1} caffeine + 0.5 mol l^{-1} Na_2SO_4 F	0.01 mol l^{-1} caffeine + 0.5 mol l^{-1} H_2SO_4	sat pyrrol + 0.5 mol l^{-1} Na_2SO_4 C	0.1 mol l^{-1} aniline + 0.5 mol l^{-1} Na_2SO_4 F
0	-	-	284.1	-	-
0.1	-	324.7	320.2	301.7	-
0.2	-	351.8	346.2	328.1	327.9
0.3	-	373.4	367.3	345.3	349.9
0.4	-	385.4	382.8	356.9	366.8
0.5	391.2	390.0	391.6	364.1	375.3
0.6	401.1	391.7	395.8	369.0	378.4
0.7	410.8	392.8	398.1	371.7	380.1
0.8	416.6	392.4	398.4	373.5	380.7
0.9	418.4	391.0	397.9	375.5	380.4
1.0	417.0	388.8	393.7	375.8	379.2
1.1	413.8	385.6	392.1	376.0	377.4
1.2	407.9	381.9	387.0	375.8	374.8
1.3	400.2	376.9	381.2	374.8	371.5
1.4	390.4	370.9	374.3	372.3	367.3
1.5	378.7	363.7	365.0	368.2	362.6
1.6	365.6	355.0	355.2	361.2	356.2
1.7	350.3	344.4	386.3	348.9	347.5
1.8	333.4	331.2	332.8	333.2	333.5
1.9	314.8	315.4	-	314.9	315.5
2.0	294.5	296.1	-	294.8	295.4

x) For base solutions, see p. 347.

Electrocapillary data on mercury in various salts, bases and acids x) (continued)

	γ /mN m^{-1}				
-E/V	0.01 mol l^{-1} aniline + 0.5 mol l^{-1} Na_2SO_4 A	sat meta-toluidine + 0.5 mol l^{-1} Na_2SO_4 C	sat para-toluidine + 0.5 mol l^{-1} Na_2SO_4 E	0.01 mol l^{-1} para-toluidine + 0.5 mol l^{-1} H_2SO_4 E	sat meta-xylidine + 0.5 mol l^{-1} Na_2SO_4 C
0	-	-	-	284.1	-
0.1	311.7	-	-	320.2	-
0.2	339.6	327.9	330.1	346.1	335.7
0.3	362.0	348.4	351.4	367.3	353.7
0.4	379.6	362.1	367.3	384.7	363.3
0.5	393.1	366.4	378.5	398.6	366.5
0.6	402.0	368.3	384.1	409.5	368.1
0.7	-	369.4	386.3	417.2	368.1
0.8	409.6	369.4	386.6	421.8	368.8
0.9	-	369.7	386.0	423.1	368.5
1.0	409.2	369.1	385.3	421.2	369.1
1.1	-	368.3	382.7	417.2	368.2
1.2	402.8	366.6	378.6	411.5	366.8
1.3	-	363.3	375.3	403.4	365.9
1.4	389.2	360.1	371.8	394.5	363.0
1.5	-	357.1	366.0	383.9	357.5
1.6	366.0	352.2	359.8	372.0	352.7
1.7	-	345.2	349.9	359.2	345.8
1.8	334.1	332.9	334.6	345.0	333.3
1.9	-	315.2	-	-	315.3
2.0	295.0	294.9	296.2	-	294.9

x) For base solutions, see p. 347.

Electrocapillary data on mercury in various salts, bases and acids x) (continued)

	$\gamma/\mathrm{mN\ m^{-1}}$				
-E/V	0.01 mol l^{-1} meta-xylidine + 0.5 mol l^{-1} H_2SO_4	sat benzyl-amine + 0.5 mol l^{-1} Na_2SO_4 E	0.1 mol l^{-1} benzyl-amine + 0.05 mol l^{-1} H_2SO_4	sat phenyl hydrazine + 0.5 mol l^{-1} Na_2SO_4 E	sat naphthyl-amine + 0.5 mol l^{-1} H_2SO_4
0	284.2	-	-	-	282.3
0.1	320.2	-	336.5	-	316.4
0.2	346.2	-	359.7	-	341.1
0.3	367.3	-	378.6	-	361.1
0.4	384.6	-	392.7	-	376.4
0.5	398.0	366.0	402.8	-	388.9
0.6	408.2	366.0	408.6	354.7	397.8
0.7	415.1	365.0	411.3	358.7	404.4
0.8	418.8	363.1	411.9	360.0	404.1
0.9	419.4	-	409.6	359.7	-
1.0	417.1	358.2	406.8	358.6	396.5
1.1	413.0	-	402.1	357.1	-
1.2	407.4	350.5	395.9	354.7	387.0
1.3	400.2	-	388.2	351.1	-
1.4	391.5	340.8	379.3	346.7	375.5
1.5	381.6	-	369.0	341.6	368.6
1.6	370.6	325.3	356.9	335.2	362.7
1.7	358.0	-	342.9	327.6	356.0
1.8	344.9	299.3	327.3	318.2	344.0
1.9	-	-	309.4	306.4	-
2.0	-	264.6	-	291.2	-
2.1	-	245.8	-	271.6	-
2.2	-	221.9	-	248.5	-
2.3	-	215.3	-	-	-

x) For base solutions, see p. 347.

Electrocapillary data on mercury in various salts, bases and acids x) (continued)

-E/V	γ/mN m^{-1} 1 mol l^{-1} pyridine + 0.5 mol l^{-1} Na_2SO_4 F	1 mol l^{-1} pyridine + 0.5 mol l^{-1} H_2SO_4 + 0.5 mol l^{-1} Na_2SO_4 F	0.1 mol l^{-1} pyridine + 0.5 mol l^{-1} Na_2SO_4 D	0.1 mol l^{-1} pyridine + 0.05 mol l^{-1} H_2SO_4 + 0.5 mol l^{-1} Na_2SO_4 D	sat picoline + 0.5 mol l^{-1} Na_2SO_4 C
0.1	-	320.2	-	-	-
0.2	-	347.8	350.4	351.9	-
0.3	367.7	370.2	373.1	-	369.1
0.4	377.8	387.7	389.2	392.4	375.5
0.5	381.0	400.4	399.4	-	376.6
0.6	381.9	408.7	405.0	416.0	376.3
0.7	380.9	412.6	406.8	-	374.8
0.8	378.7	413.2	406.8	423.1	371.7
0.9	374.9	410.8	404.8	-	367.4
1.0	369.9	406.0	401.0	417.7	362.1
1.1	363.9	399.5	395.6	-	355.9
1.2	357.0	-	389.2	404.2	349.1
1.3	349.3	-	381.5	-	341.8
1.4	341.1	-	373.2	383.8	333.9
1.5	332.2	-	363.5	-	325.7
1.6	322.8	-	355.1	357.9	316.9
1.7	313.0	-	345.3	-	307.8
1.8	302.6	-	334.5	328.4	298.0
1.9	291.7	-	316.5	-	287.9
2.0	280.6	-	296.2	-	277.5
2.1	268.0	-	-	-	266.6
2.2	249.6	-	-	-	248.7
2.3	-	-	-	-	-
2.4	-	-	-	-	-

x) For base solutions, see p. 347.

Electrocapillary data on mercury in various salts, bases and acids x) (continued)

	γ/mN m^{-1}			
-E/V	1 mol l^{-1} piperidine + 0.5 mol l^{-1} Na_2SO_4 C	0.1 mol l^{-1} pilocarpine + 0.5 mol l^{-1} Na_2SO_4 E	sat cocaine + 0.5 mol l^{-1} Na_2SO_4 E	sat codeine + 0.5 mol l^{-1} Na_2SO_4 E
0.1	-	-	-	-
0.2	-	349.9	349.9	-
0.3	-	365.5	370.8	368.8
0.4	-	375.3	383.3	372.9
0.5	373.3	379.8	386.1	374.0
0.6	373.3	381.9	386.1	-
0.7	371.7	382.3	384.8	372.0
0.8	369.1	381.9	383.9	-
0.9	366.0	379.6	381.6	366.0
1.0	362.3	376.0	377.4	-
1.1	358.4	370.8	373.3	358.3
1.2	353.6	364.7	369.1	-
1.3	348.2	357.7	364.0	346.1
1.4	342.3	349.6	358.4	-
1.5	335.6	340.3	352.1	333.6
1.6	327.8	330.6	341.7	-
1.7	318.9	319.5	332.5	317.5
1.8	307.7	308.4	323.2	308.2
1.9	294.2	296.3	308.2	296.9
2.0	278.8	283.3	294.3	295.3
2.1	-	-	273.7	-
2.2	-	-	250.0	-
2.3	-	-	-	-
2.4	-	-	-	-

x) For base solutions, see p. 347.

Electrocapillary data on mercury electrode in 1 M KCl at 20°C.
Given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: J.R. Meakins. J. Appl. Chem. 17 (1967) 156.

Adsorption of C_{12}TMABr= n-dodecyltrimethylammonium bromide

	γ /mN m^{-1}				
			C_{12}TMABr		
c/mol l^{-1}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}
-E/V					
0.0	360	354	355	350	335
0.1	387	379	379	370	367
0.2	395	388	382	377	372
0.3	395	391	386	379	372
0.4	395	389	384	377	373
0.5	389	385	381	374	369
0.6	386	380	377	369	365
0.8	375	369	365	357	354
1.0	359	355	351	344	337
1.2	342	339	332	325	315
1.4	310	310	307	299	290
1.6	271	271	270	266	259
1.8	222	221	222	222	217

Electrocapillary data for quaternary ammonium salts in KCl (continued)

Adsorption of C_{16}TMABr = n-hexadecyltrimethylammonium bromide

			γ/mN m^{-1}			
			C_{16}TMABr			
c/mol l^{-1}	10^{-4}	1.5×10^{-4}	2×10^{-4}	3×10^{-4}	5×10^{-4}	10^{-3}
-E/V						
0.0		357	359	359	358	361
0.1	379	374	368	362	357	361
0.2	392	377	372	367	364	358
0.3	394	379	372	369	365	360
0.4	393	381	374	367	364	362
0.5	388	375	371	364	360	359
0.6	387	372	369	362	359	356
0.8	377	364	357	350	349	342
1.0	366	354	347	337	334	331
1.2	347	339	331	319	316	308
1.4	313	313	311	298	290	282
1.6		270	271	269	256	249
1.8		223	222	217	215	207

Electrocapillary data for quaternary ammonium salts in 1 M KCl (continued)

Adsorption of C_{12}TEABr= n-dodecyltriethylammonium bromide

	γ /mN m^{-1}					
	C_{12}TEABr					
c/mol l^{-1}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}
-E/V						
0.0	359	361	357	344	330	313
0.1	383	383	378	371	365	358
0.2	393	390	384	380	372	367
0.3	395	390	386	378	371	370
0.4	393	386	383	374	370	365
0.5	386	381	376	372	365	363
0.6	383	376	372	365	360	357
0.8	372	365	360	353	348	340
1.0	355	350	344	338	332	327
1.2	339	333	325	318	310	305
1.4	320	307	302	294	285	279
1.6	284	280	272	264	256	251
1.8	233	234	234	228	219	213

Electrocapillary data for quaternary ammonium salts in 1 M KCl (continued)

Adsorption of C_{16}TEABr = n-hexadecyltrimethylammonium bromide

	γ/mN m^{-1}					
	C_{16}TEABr					
c/mol l^{-1}	1.0×10^{-4}	1.5×10^{-4}	2×10^{-4}	3×10^{-4}	5×10^{-4}	10^{-3}
-E/V						
0.0	370	368	369	358		347
0.1	381	372	369	366	362	358
0.2	393	383	380	377	375	373
0.3	390	381	380	380	375	373
0.4	387	380	378	376	373	374
0.5	385	377	374	373	369	369
0.6	382	373	369	368	365	363
0.8	375	362	358	355	352	352
1.0	363	353	345	340	336	333
1.2	351	337	327	322	315	315
1.4	331	322	309	296	289	288
1.6	290	285	282	265	260	258
1.8	238	238	238	232	224	218

Electrocapillary data on mercury electrode in 1 M KCl - adsorption of C_8TBABr. Given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: R.J. Meakins and R. Driver. Unpublished results.

	γ/mN m^{-1}				
	C_8TBABr				
c/mol l^{-1}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}
-E/V					
0	361	361	350	340	326
0.1	385	383	373	367	355
0.15			379		
0.2	392	387	380	374	365
0.25	393	386	381	375	366
0.3	391	384	380	373	368
0.35	389				368
0.4	387	379	376	372	368
0.45					364
0.5	383	377	373	368	361
0.55					
0.6	378	371	364	362	356
0.7	371	366	362	356	351
0.8	364	359	356	350	344
0.9	358	352	348	342	337
1.0	352	344	339	334	329
1.1	345	337	332	326	322
1.2	334	328	324	319	312
1.4	315	309	306	300	295
1.6	282	281	279	278	269
1.8	237	235	235	236	235

Electrocapillary data on mercury electrode in 1 M KCl (continued)

Adsorption of C_8TEABr.

γ /mN m^{-1}

				C_8TEABr			
c/mol l^{-1}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
-E/V							
0	361	361	359	355	344	327	287
0.1	384	386	384	381	376	368	349
0.15							
0.2	401	401	399	395	389	381	367
0.25		405	402	399	392	384	
0.3	411	407	404	400	393	387	375
0.35	411	406	402	399	392	388	378
0.4	408	406	400	397	390	386	377
0.45	408						
0.5	404	401	396	393	386	380	374
0.55							
0.6	399	392	389	384	378	373	367
0.7	392	384	380	377	371	365	359
0.8	383	377	373	368	365	359	353
0.9	376	371	365	362	356	350	347
1.0	368	361	355	352	349	347	337
1.1	356	350	348	343	340	336	330
1.2	347	338	337	333	330	325	318
1.4	323	319	315	312	308	303	295
1.6	284	286	286	286	280	277	269
1.8	237	237	238	237	236	238	239

Electrocapillary data on mercury electrode in 1 M KCl (continued)

Adsorption of C_8TMABr

	γ /mN m^{-1}						
				C_8TMABr			
c/mol l^{-1}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
-E/V							
0	361	361	361	357	347	332	286
0.1	384	385	383	383	380	375	354
0.15							
0.2	399	400	400	397	394	391	372
0.25							
0.3	413	413	411	406	402	397	384
0.35		415	412	407	404	398	387
0.4	418	414	411	407	403	398	389
0.45	417	414		406	401	398	387
0.5	416	411	408	406	401	396	386
0.55	415						
0.6	411	406	402	400	396	391	381
0.7	406	400	399	394	388	384	375
0.8	399	394	390	387	381	377	369
0.9	390	385	383	379	372	371	361
1.0	381	378	375	369	365	362	354
1.1	370	368	363	360	356	351	345
1.2	360	356	351	349	344	341	333
1.4	328	327	326	325	321	318	310
1.6	286	287	286	286	286	286	280
1.8	238	236	239	236	236	237	238

Electrocapillary data on mercury electrode in 1 M KCl (continued)

Adsorption of C_8TPABr

				γ /mN m^{-1}			
				C_8TPABr			
c/mol l^{-1}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
-E/V							
0	359	361	362	350	333	312	310
0.1	383	383	383	377	365	356	339
0.15				382			
0.2	398	394	391	384	376	370	356
0.25	400	396	391	383	377	372	362
0.3	400.5	396	390	383	376	371	364
0.35	399	394	388	384	375	370	365
0.4	398	392	387	381	372		363
0.45							
0.5	393	387	381	377	370	365	361
0.55							
0.6	386	381	377	373	367	361	354
0.7	378	374	371	367	359	356	349
0.8	373	368	365	361	353	350	342
0.9	365	361	358	353	347	343	336
1.0	358	353	349	346	339	336	329
1.1	350	346	342	339	332	328	322
1.2	341	336	334	330	322	320	311
1.4	321	316	313	311	305	300	292
1.6	285	285	284	287	283	276	266
1.8	236	231	239	237	239	239	236

Electrocapillary data on mercury electrode in 1 M KCl (continued)

Adsorption of C_{12}TBABr

γ/mN m^{-1}

	C_{12}TBABr					
c/mol l^{-1}	10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}
-E/V						
0	367	366	368	364	352	353
0.1	388	388	385	377	365	359
0.2	403	398	387	380	372	369
0.25		400	387	381	374	369
0.3	417	399	386	379	373	369
0.35		399				
0.4	424	397	382	376	369	366
0.45	426					
0.5	426	392	379	372	366	359
0.55	425					
0.6	424	388	374	366	358	355
0.7	420	386	368	361	353	348
0.8	411	382	363	356	347	341
0.9	403	378	356	347	340	334
1.0	392	372	351	340	333	326
1.1	379	362	341	333	323	318
1.2	363	356	333	324	314	305
1.4	329	328	317	303	291	285
1.6	289	289	289	279	268	257
1.8	242	241	242	236	233	225

Electrocapillary data on mercury electrode in 1 M KCl (continued)

Adsorption of C_{12}TPABr

γ/mN m^{-1}

	C_{12}TPABr						
c/mol l^{-1}	10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-3}	10^{-3}	10^{-2}
-E/V							
0	367	367	375	370	348	328	325
0.1	386	387	392	386	373	364	350
0.2	403	406	406	395	385	375	365
0.25			406	396	385	379	370
0.3	415	419	405	395	386	382	371
0.35		426			385	380	370
0.4	425	430	402	390.5	383	380	368
0.45	427	432					
0.5	427	430	401	388	379	374	364
0.55							
0.6	421	430	394	383.5	373	368	358
0.7	417	426	392	378	369	364	355
0.8	409	415	387	372	364	358	347
0.9	401	406	383	364	357	350	340
1.0	391	392	378	358	347	342	332
1.1	377	378	372	351	341	335	324
1.2	361	364	364	343	332	327	310
1.4	327	328	330	322	312	306	289
1.6	289	290	290	289	287	281	266
1.8	240	241	241	241	241	241	237

Electrocapillary data on mercury electrode in 1 M KCl (continued)

Adsorption of C_8pyBr

γ /mN m^{-1}

	C_8pyBr					
c/mol l^{-1}	2 x 10^{-4}	5 x 10^{-4}	10^{-3}	2 x 10^{-3}	5 x 10^{-3}	10^{-2}
-E/V						
0	366	363	364	360	348	325
0.1	390	387	386	387	382	370
0.2	405	401	400	401	395	388
0.25						
0.3	414	411	408	404	402	395
0.35	416	413	409	406	402	395
0.4	416	412	409	405	402	395
0.45						
0.5	413	409	406	402	400	394
0.55						
0.6	407	403	401	395	392	387
0.7	401	397	393	390	385	381
0.8	393	389	385	383	378	374
0.9	384	381	377	373	370	363
1.0	374	370	366	363	360	354
1.1	363	360	355	352	349	342
1.2	354	348	346	343	336	336
1.3						
1.4						
1.6						
1.8						

Electrocapillary data on mercury electrode in 1 M KCl (continued)

Adsorption of C_{12}PyBr

γ /mN m^{-1}

				C_{12}pyBr			
c/mol l^{-1}	10^{-5}	3×10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}
-E/V							
0	363	360	361	360	354	335	316
0.1	386	385	385	378	376	367	354
0.2	400	399	395	389	385	378	367
0.25				391	387	379	
0.3	413	404	398	393	387	379	372
0.35			400	392	387		371
0.4	421	409	399	393	387	380	371
0.45	423	410	398				
0.5	423	408	397	390	384	375	368
0.55							
0.6	422	407	393	383	379	368	363
0.7	415	402	386	381	375	365	357
0.8	407	395	379	373	367	358	352
0.9	398	388	373	364	360	349	342
1.0	386	382	365	359	351	342	334
1.1	375	370	358	352	349	345	339
1.2	360	360	355	351	352	352	342
1.3	344			345	346	345	346
1.4	326	325	328	328	328	328	329
1.6	285	285	285	284	285	285	285
1.8	235	235	236	235	235	235	235

Electrocapillary data on mercury electrode in 1 M KCl (continued)
Adsorption of $C_{10}iQBr$

γ /mN m^{-1}

	$C_{10}iQBr$					
c/mol l^{-1}	10^{-5}	2×10^{-5}	5×10^{-5}	10^{-4}	2×10^{-4}	10^{-3}
-E/V						
0	361	363	363	355	340	308
0.1	386	384	381	376	371	348
0.2	402	396	389	383	380	364
0.25				387	380	370
0.3	415	408	393	389	381	372
0.35			392	387	382	372
0.4	422	412	392	386	382	371
0.45	424	423	391			
0.5	425	424	391	384	377	366
0.55	423	423				
0.6	421	413	387	377	371	361
0.7	417	409	380	372	364	356
0.8	409	403	375	366	361	351
0.9	401	394	381	376	368	352
1.0	389	387	382	384	380	364
1.1	376	377	376	377	375	373
1.2	360	361	363	363	362	360
1.4	328	328	326	327	327	327
1.6	287	287	285	287	288	287
1.8	237	237	238	238	237	237

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 - given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Adsorption of C_8TBABr

Reference: R.J. Meakins and R. Driver. Unpublished data.

	γ /mN m^{-1}					
	C_8TBABr					
c/mol l^{-1}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}
-E/V						
0	389	393	391	390	388	380
0.1	403	405	405	403	400	393
0.15			413	407	403	399
0.2	414	415	414	407	404	398
0.25			414	406	404	398
0.3	423	421	411	405	400	395
0.35	423					
0.4	423	425	408	399	397	391
0.45	425	428				
0.5	425	429	403	394	394	384
0.55	425	429				
0.6	424	428	399	391	389	380
0.7	421	422	393	384	381	376
0.8	414	417	386	377	375	369
0.9	406	407	380	370	369	362
1.0	397	398	372	362	361	355
1.1	388	388	364	355	353	345
1.2	382	380	358	347	344	337

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 (continued)

Adsorption of C_8TEABr

	γ /mN m^{-1}						
	C_8TEABr						
c/mol l^{-1}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
-E/V							
0	389	386	378	369	350	328	287
0.1	403	403	400	393	381	370	356
0.15							
0.2	417	417	411	403	396	388	377
0.25	421	417	412	410	402	393	
0.3	421	417	412	411	403	398	385
0.35	422		412	409	403	399	386
0.4	420	414	412	409	402	399	388
0.45							387
0.5	417	409	407	403	397	392	386
0.55							
0.6	409	403	399	396	391	385	377
0.7	401	395	391	388	385	381	370
0.8	393	386	384.5	383	375	372	363
0.9	385	381	377	374	370	363	353
1.0	378	373	367	364	361	352	345
1.1	369	362	358	355	350	344	337
1.2	359	352	350	345	337	334	328

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 (continued)

Adsorption of C_8TPABr

	γ/mN m^{-1}						
	C_8TPABr						
c/mol l^{-1}	2×10^{-4}	5×10^{-4}	10^{-3}	2×10^{-3}	5×10^{-3}	10^{-2}	3×10^{-2}
-E/V							
0	391	384	376	364	340	323	329
0.1	404	402	394	386	377	369	349
0.15		406	400	396			
0.2	414	409	405	400	389	381	370
0.25	415	409	404	400	392	383	375
0.3	414	408	403	398	393	388	379
0.35					391	387	380
0.4	411	404	402	395	388	386	376
0.45							
0.5	406	399	397	389	384	380	372
0.55							
0.6	400	393	389	383	377	374	366
0.7	394	388	383	377	374	369	358
0.8	389	382	377	367	366	361	355
0.9	381	374	367	363	358	356	346
1.0	373	366	361	358	351	349	338
1.1	363	356	351	348	344	340	329
1.2	355	348	342	353	337	331	320

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 (continued)

Adsorption of $(C_8TBA)_2SO_4$

	γ /mN m^{-1}					
	$(C_8TBA)_2\ SO_4$					
c/mol l^{-1}	5×10^{-5}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}
-E/V						
0	389	391	390	391	387	381
0.05		398	396	393	387	381
0.1	406	402	396	391	388	381
0.15	405	401	395.5		385	379
0.2	406	400	394	390	383	378
0.25	405					
0.3	403	397	391	386	381	375
0.4	397	393	386	383	379	371
0.45						
0.5	393	388	382	379	373	367
0.6	391	383	377	375	368	361
0.7	384	378	371	368	364	356
0.8	376	371	365	362	355	352
0.9	371	365	360	356	350	343
1.0	364	358	352	352	345	337
1.1	355	349	343	341	335	330
1.2	347	343	337	332	323	320

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 (continued)

Adsorption of $[C_8TEA]_2SO_4$ = n octyltriethylammonium sulfate.

	γ /mN m^{-1}						
	$[C_8TEA]_2SO_4$						
c/mol l^{-1}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
-E/V							
0	391	391	391	389	390	391	390
0.05							
0.1	406	403	404	406	402	401	398
0.15		411	410	412	406	404	397
0.2	416	416	414	412	407	404	395
0.25	421	417	414	411	407	402	
0.3	420	416	414	410	405	401	393
0.4	416	412	409	406	399	395	390
0.45							
0.5	412	406	404	401	394	390	383
0.6	406	400	397	394	387	383	378
0.7	398	392	389	386	382	377	371
0.8	391	385	383	380	373	371	365
0.9	382	376	374	369	365	361	353
1.0	372	366	365	361	356	351	346
1.1	364	357	354	351	345	342	337
1.2	352	349	346	340	336	333	327

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 (continued)

Adsorption of $[C_8TMA]_2SO_4$

	γ /mN m^{-1}						
	$[C_8TMA]_2SO_4$						
c/mol l^{-4}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-2}	1.5×10^{-2}
-E/V							
0	389	392	391	391	388	391	392
0.05							
0.1	404	405	403	406	403	405	403
0.15						412	407
0.2	414	417	415	417	414	415	407
0.25		419	419	421	417	416	406
0.3	425	425	422	421	416	415	407
0.4	427	425	421	420	414	405	404
0.45	427						
0.5	425	423	418	418	411	405	400
0.6	418	417	414	409	405	400	394
0.7	413	409	406	405	397	395	390
0.8	406	404	397	396	391	385	383
0.9	397	395	391	388	385	380	374
1.0	388	386	381	380	374	370	366
1.1	378	378	374	372	364	363	355
1.2	370	367	363	362	355	353	347

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 (continued)

Adsorption of $[C_8TPA]_2SO_4$

	γ /mN m^{-1}						
	$[C_8TPA]_2SO_4$						
c/mol l^{-1}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
-E/V							
0	391	391	392	394	389	388	382
0.05			398	399	394	389	383
0.1	406	405	404	401	393	388	381
0.15	411	409	405	399			
0.2	414	407	404	398	391	386	379
0.25	413						
0.3	411	406	400	396	386	384	376
0.4	405	402	396	391	385	379	373
0.45							
0.5	402	397	391	387	381	374	368
0.6	397	390	387	382	374	370	364
0.7	391	385	379	378	370	367	358
0.8	385	379	373	370	364	359	350
0.9	376	372	365	362	356	354	346
1.0	367	361	359	356	346	344	338
1.1	359	352	350	347	340	338	326
1.2	350	346	343.5	338	331	328	322

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4

Adsorption of $(C_{16}TMA)_2SO_4$

γ/mN m^{-1}

	$(C_{16}TMA)_2SO_4$						
c/mol l^{-1}	2.5×10^{-5}	3.5×10^{-5}	5×10^{-5}	7.5×10^{-5}	1.5×10^{-4}	3×10^{-4}	5×10^{-4}
-E/V							
0	390	392	390	387	381	379	379
0.05			394	389	384	382	383
0.1	403	401	394	392	387	380	382
0.15		401	396	393	384	380	
0.2	404	399	394	392	384	380	381
0.25	406						
0.3	404	402	394	394	379	378	381
0.4	403	399	392	386	378	374	378
0.45							
0.5	401	398	389	386	376	371	373
0.6	400	394	386	384	369	368	370
0.7	395	391	382	378	365	364	363
0.8	389	385	379	371	359	355	357
0.9	385	382	373	367	355	354	352
1.0	374	374	368	359	348	349	343
1.1	371	368	362	355	340	344	335
1.2	365	360	352	347	337	335	329

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)

Adsorption of $(C_8TBA)_2SO_4$

	γ /mN m^{-1}				
	$(C_8TBA)_2SO_4$				
c/mol l^{-1}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}
-E/V					
-0.1				383	380
-0.05	398	397	397	391	382
0	410	405	400	387	381
0.05	407	402	399	387	380
0.1	407	400	389	383	377
0.15					
0.2	398	391	385	381	376
0.25					
0.3	391	384	379	378	372
0.35					
0.4	387	379	374	372	368
0.45					
0.5	381	372	368	366	364
0.6	374	366	364	362	359
0.7	369	360	357	354	353
0.8	362	355	350	347	347
0.9	354	347	342	341	339
1.0	347	339	335	331	331
1.1	339	330	324	324	322
1.2	330	321	315	317	313
1.4	311	303	297	295	292
1.6	286	278	272	271	265
1.8	237	237	237	239	233

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)

Adsorption of $(C_8TEA)_2SO_4$

	γ/mN m^{-1}						
	$(C_8TEA)_2SO_4$						
c/mol l^{-1}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
-E/V							
-0.1							
-0.05							
0	407	410	415	417	418	416	413
0.05				427	424	417	407
0.1	417	416	426	426	429	420	410
0.15		423	426	425	428	420	409
0.2	431	427	426	423	424	417	410
0.25	431	427					
0.3	432	426	423	421	416	411	406
0.35	430						
0.4	427	423	418	415	412	407	401
0.45							
0.5	423	413	412	408	407	401	393
0.6	413	407	404	401	397	390	386
0.7	405	397	396	392	389	385	380
0.8	396	390	386	386	378	374	372
0.9	387	381	378	375	373	368	364
1.0	380	373	370	368	365	359	354
1.1	370	364	360	357	354	351	343
1.2	359	354	349	348	344	340	333
1.4	336	330	327	324	321	317	309
1.6	300	298	298	295	290	289	283
1.8	249	251	249	251	251	252	246

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)

Adsorption of $(C_8TMA)_2SO_4$

	γ /mN m^{-1}						
	$(C_8TMA)_2SO_4$						
c/mol l^{-1}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
-E/V							
-0.1							
-0.05							
0	410	406	404	404	414	419	416
0.05							
0.1	417	415	415	415	423	429	426
0.15					425	426	425
0.2	427	430	430	430	427	426	425
0.25		431	434	431	427	424	
0.3	436	434	436	429	428	423	421
0.35	438	434	434		427		
0.4	437	433	432	427	423	419	419
0.45	436						
0.5	436	430	427	419	417	415	412
0.6	428	422	419	414	411	405	403
0.7	421	414	412	406	403	400	398
0.8	415	406	404	401	399	393	392
0.9	405	398	395	393	387	385	383
1.0	397	389	385	384	378	378	374
1.1	384	380	374	372	367	370	365
1.2	371	367	364	361	360	359	351
1.4	341	337	337	336	333	333	326
1.6	300	299	299	301	301	298	295
1.8	290	250	250	250	250	252	250

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)

Adsorption of $(C_8TPA)_2SO_4$

	γ/mN m^{-1}						
	$(C_8TPA)_2SO_4$						
c/mol l^{-1}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-3}	1.5×10^{-2}
-E/V							
-0.1							
-0.05				390	401	401	398
0	406	409	405	400	404	404	400
0.05		412	413	404	401	402	397
0.1	419	418	412	406	400	401	396
0.15	418	417	412	402			
0.2	419	414	408	401	398	398	398
0.25	417						
0.3	415	408	402	396	395	395	392
0.35							
0.4	409	404	399	393	392	393	389
0.45							
0.5	404	398	393	387	390	386	382
0.6	400	392	386	382	380	382	378
0.7	392	385	382	376	375	376	371
0.8	384	377	374	369	366	368	364
0.9	379	372	364	360	361	361	360
1.0	368	361	355	352	350	352	349
1.1	358	354	349	343	341	343	339
1.2	347	345	338	333	332	330	328
1.4	325	322	316	311	311	310	308
1.6	296	297	294	289	285	283	280
1.8	248	248	246	248	248	252	245

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_{10}TMA)_2SO_4$

	γ/mN m^{-1}						
	$(C_{10}TMA)_2SO_4$						
c/mol l^{-1}	2.5×10^{-5}	5×10^{-5}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1.5×10^{-3}	5×10^{-3}
-E/V							
0	404	402	407	408	412	414	405
0.05					411	418	407
0.1	411	412	413	421	422	417	403
0.15			420	421	419		
0.2	422	424	424	423	419	411	396
0.25			424	422			
0.3	430	426	425	420	418	406	394
0.35	429	425					
0.4	429	423	421	418	415	404	391
0.45	427						
0.5	424	418	415	413	409	397	386
0.6	416	412	407	403	400	392	383
0.7	412	404	403	397	391	388	373
0.8	404	398	395	388	384	379	369
0.9	395	387	386	382	376	371	361
1.0	387	379	378	370	369	364	352
1.1	379	370	369	365	361	354	342
1.2	366	361	358	352	352	343	331
1.4	336	334	334	331	327	322	306
1.6	291	294	294	294	294	288	267
1.8	246	246	246	246	246	244	

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)
Adsorption of $[C_{12}TMA]_2SO_4$

				γ /mN m^{-1}			
				$[C_{12}TMA]_2SO_4$			
c/mol l^{-1}	1.5×10^{-5}	2.5×10^{-5}	5×10^{-5}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	5×10^{-3}
-E/V							
0	404	398	406	406	404	401	399
0.05				412	405	403	399
0.1	413	410	412	408	402	398	394
0.15			411				
0.2	421	418	411	405	395	392	387
0.25		417					
0.3	428	416	409	403	392	389	387
0.35	427						
0.4	427	415	407	401	391	385	383
0.45							
0.5	425	409	403	397	386	380	380
0.6	420	404	398	389	380	375	375
0.7	413	400	392	386	376	370	369
0.8	406	392	386	380	368	365	362
0.9	397	384	380	372	360	357	354
1.0	389	377	371	365	354	345	344
1.1	379	370	363	358	342	336	335
1.2	365	359	353	346	332	326	326
1.4	333	332	330	321	306	299	296
1.6	293	293	293	285	271	266	263
1.8	245	245	244	243	233	222	219

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)
Adsorption of $[C_{14}TMA]_2SO_4$

	γ /mN m^{-1}						
	$[C_{14}TMA]_2SO_4$						
c/mol l^{-1}	2.5×10^{-5}	3.5×10^{-5}	5×10^{-5}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}
-E/V							
-0.05						390	392
0	399	395	394	396	393	393	393
0.05		402	395	395	393	393	395
0.1	404	401	399	394	389	389	392
0.15	404	401	399	392			
0.2	404	399	396	393	380	389	377
0.3	404	399	395	392	386	386	377
0.4	402	395	393	389	381	380	374
0.45							
0.5	399	393	392	386	378	378	369
0.6	396	386	387	377	365	374	366
0.7	390	380	378	373	364	364	359
0.8	387	377	367	366	356	361	352
0.9	380	371	362	355	348	352	351
1.0	372	364	360	345	342	343	345
1.1	365	358	351	338	335	332	335
1.2	356	347	336	332	326	323	324
1.4	329	326	317	303	300	299	299
1.6	289	290	287	278	268	268	269
1.8	246	246	244	240	227	225	223

Electrocapillary data on mercury electrode in 0.5 M Na_2SO_4 (continued)
Adsorption of $(C_{16}TMA)_2SO_4$

	γ/mN m^{-1}						
	$(C_{16}TMA)_2SO_4$						
c/mol l^{-1}	2.5×10^{-5}	3.5×10^{-5}	5×10^{-5}	7.5×10^{-5}	1.5×10^{-4}	2.5×10^{-4}	3×10^{-4}
-E/V							
0	401	393	397	397	395	394	395
0.05			400	398	397	398	396
0.1	413	407	395	399	393	393	394
0.15	417	408	396	401	393		
0.2	417	408	396	398	392	392	392
0.25	420	407.5					
0.3	419	406	402	397	393	389	389
0.4	415	404	401	395	389	389	389
0.45							
0.5	411	401	398	392	385	384	382
0.6	407	401	393	389	381	380	377
0.7	402	396	393	384	377	370	369
0.8	395	390	387	379	369	366	365
0.9	389	386	383	375	364	356	356
1.0	384	379	377	368	355	350	349
1.1	374	371	369	360	345	340	339
1.2	369	360	356	348	333	333	331
1.4	334	329	331	324	310	304	302
1.6	294	294	294	294	279	271	270
1.8	247	246	247	247	243	231	229

Electrocapillary data on mercury electrode in 0.5 M H_2SO_4 at 20°C. Given are the interfacial tension in mN m^{-1}, potential in V vs SCE.

Reference: R.J. Meakins. J. Appl. Chem. 15 (1965) 416.

Adsorption of C_8TMABr = n-octyltrimethylammonium bromide

	γ /mN m^{-1}					
	C_8TMABr					
c/mol l^{-1}	3×10^{-4}	10^{-3}	3×10^{-3}	10^{-2}	3×10^{-2}	10^{-1}
-E/V						
0.0	386	381	366	335	295	
0.1	404	401	396	381	363	329
0.2	413	412	408	397	386	367
0.3	423	419	414	405	395	382
0.4	423	420	415	407	401	387
0.5	421	414	411	405	397	389
0.6	414	411	406	399	396	388
0.7	408	404	401	394	387	380
0.8	400	396	394	386	383	374
0.9	392	390	385	380	375	365
1.0	384		376	370	364	357
1.1			366	362	359	348
1.2				351	343	339
1.3				341	336	326
1.4						

Electrocapillary data for quaternary ammonium salts in 0.5 M H_2SO_4 (continued)

Adsorption of C_{12}TMABr = n-dodecyltrimethylammonium bromide.

	γ /mN m^{-1}					
	C_{12}TMABr					
c/mol l^{-1}	5×10^{-5}	10^{-4}	3×10^{-4}	10^{-3}	10^{-2}	3×10^{-2}
-E/V						
0.0		391	386	360	386	
0.1	402	400	394	380	361	
0.2	404	404	398	389	373	
0.3	407	406	395	391	376	374
0.4	404	404	394	384	377	376
0.5	400	400	393	384	375	376
0.6	399	397	388	380	369	369
0.7	393	393	381	374	363	365
0.8	387	387	377	368	359	360
0.9	379	377	368	362	353	356
1.0	373	371	359	355	344	345
1.1	367	364	352	345	335	334
1.2	357	351	344	337	324	322
1.3		342	333	324	311	311
1.4		336	324	318	300	298

Electrocapillary data for quaternary ammonium salts in 0.5 M H_2SO_4 (continued)

Adsorption of C_{16}TMABr = n-hexadecyltrimethylammonium bromide.

	γ/mN m^{-1}					
	C_{16}TMABr					
c/mol l^{-1}	10^{-4}	2×10^{-4}	5×10^{-4}	10^{-3}	3×10^{-3}	10^{-2}
-E/V						
0.0	388	378	369	368	376	
0.1	398	378	381	378	370	354
0.2	392	379	375	371	371	364
0.3	391	380	379	375	371	371
0.4	388	378	375	374	372	371
0.5	388	375	370	370	369	368
0.6	387	370	366	366	367	362
0.7	382	366	362	359	361	359
0.8	375	361	355	355	353	353
0.9	372	356	349	348	346	346
1.0	364	349	343	340	339	339
1.1	356	341	334	331	331	332
1.2	349	334	324	322	320	319
1.3	343	323	314	310	309	309
1.4	332	316	304	301	298	296

Electrocapillary data for quaternary ammonium salts in 0.5 M H_2SO_4 (continued)
Adsorption of $(C_{12}TMA)_2SO_4$ = n-dodecyltrimethylammonium sulphate.

	γ /mN m^{-1}							
	$(C_{12}TMA)_2SO_4$							
c/mol l^{-1}	2.5×10^{-5}	5×10^{-5}	1×10^{-4}	2.5×10^{-4}	5×10^{-4}	1×10^{-3}	2.5×10^{-3}	5×10^{-3}
-E/V								
0.0	388	380	383	384	384	380	373	359
0.1	401	401	397	394	389	382	375	366
0.2	409	402	399	392	389	384	368	367
0.3	405	401	399		387	381	364	365
0.4	405	399	393	389	384	378	362	361
0.5	396	390	386	380	374	361	352	353
0.6	383	380	374	369	364	351	343	341
0.7	370	365	358	354	351	335	330	328
0.8	350	342	342	336	331	323	308	303
0.9	324	317	319	309	301	290	283	280
1.0								
1.1								
1.2								
1.3								
1.4								

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11. ABSTRACT *(A 200-word or less factual summary of most significant information. If document includes a significant bibliography or literature survey, mention it here)*

This paper contains data on the electrical double layer on a pure mercury electrode in contact with a solution which contains at least one electrolyte and, in some cases, a non-electrolyte as well. The quantities covered are the double-layer capacity and the interfacial tension.

The search of the literature was carried through 1978. The tables have been made as complete as possible subject to certain criteria of reliability relating to the purity of substances used, the probability that equilibrium had been reached, the control of ambient conditions and geometrical and electrical conditions specific to the type of measurement. Estimates are given for the reliability of the data.

The classes of solutions covered are: solutions of single electrolytes, solutions containing mixture of electrolytes, electrolyte solutions containing nonionic organic additives, and solutions containing substances with ionic surfactant character. The data are presented in primary form with a minimum of subsequent processing. There is relatively little overlap between data sets.

12. KEY WORDS *(Six to twelve entries; alphabetical order; capitalize only proper names; and separate key words by semicolons)*

Data compilation; critically evaluated data; electrical double layer; interfacial tension; mercury electrode.

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